


Cite this: *CrystEngComm*, 2021, 23, 6755

Received 11th August 2021,
Accepted 30th August 2021

DOI: 10.1039/d1ce01075a

rsc.li/crystengcomm

Novel TCNQ-stacking motifs in (12-crown-4)-complexes of alkali metal TCNQ salts†

Bingjia Yan,  ‡ Peter N. Horton,  Simon C. Weston,  §
Andrea E. Russell  and Martin C. Grossel  *

An investigation of the solid-state behaviour of five 12-crown-4 alkali metal TCNQ complexes, (12-crown-4)₂LiTCNQ (1), (12-crown-4)₂NaTCNQ (2), (12-crown-4)₂Li(TCNQ)₂ (3), (12-crown-4)₂Na(TCNQ)₂ (4), and (12-crown-4)₂K(TCNQ)₂ (5), reveals an unusual “cross-stitch” packing motif with the extended face-to-face π -stacked TCNQ^{•−} columns present in complexes 1 and 2. The effect of the presence of additional neutral TCNQ⁰ has also been explored.

The ability of 7,7',8,8'-tetracyanoquinodimethane (TCNQ) to act as a good one-electron acceptor results in charge-transfer complexes,^{1–7} which can exhibit interesting solid-state electrical and magnetic properties,^{8,9} arising from the presence of unpaired electrons.^{10–16} The TCNQ^{•−} moieties have a tendency to form face-to-face π -stacked dimers, which can assemble into infinite columns within a crystal.^{17–19} Furthermore, neutral TCNQ⁰ can be incorporated into these anion stacks resulting in a considerable change in the solid-state electronic properties of such materials.²⁰ The combination of the potential packing arrangements of the TCNQ units and the size and nature of the counter-cation are the dominant factors controlling the solid-state architectures of such complexes.²¹

We have previously reported detailed studies of the solid-state behaviour of a range of ionophore-complexed alkali metal-TCNQ^{•−} salts including 15-crown-5 and 18-crown-6 complexes of Li⁺, Na⁺, K⁺, Tl⁺ and Rb⁺TCNQ^{•−},^{1,17,19,22–25} and 15-crown-5 complexes of Li⁺ and Na⁺ (TCNQ^{•−})(TCNQ⁰).²⁵ Recently, Akutagawa and co-workers^{26,27} have reported the crystal structure of Li⁺(12-crown-4)₂(TCNQ)₂ and Na⁺(12-crown-4)₂(TCNQ)₂, but there is limited information to the solid-state behaviour of other 12-crown-4 alkali metal TCNQ complexes.

In this paper, we report the synthesis and crystallographic studies of five 12-crown-4 alkali metal TCNQ complexes, two simple TCNQ salts: (12-crown-4)₂LiTCNQ (1), and (12-crown-4)₂NaTCNQ (2); and three related complex TCNQ salts: (12-crown-4)₂Li(TCNQ)₂ (3), (12-crown-4)₂Na(TCNQ)₂ (4), and (12-crown-4)₂K(TCNQ)₂ (5) (see Scheme 1).

Experimental procedure

General consideration

All experiments were conducted under a nitrogen atmosphere unless stated otherwise. Acetonitrile was dried over calcium hydride and distilled before use. Diethyl ether was dried over sodium wire and was freshly decanted before use. All chemicals were commercially available and used as received, unless otherwise stated.

Physical measurements

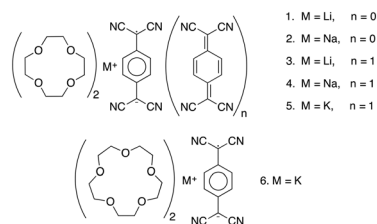
Melting points were measured on an electrothermal melting point apparatus and are uncorrected. Elemental analyses were performed by Medac Ltd., Chobham Business Centre, Chertsey Road, Chobham, Surrey, GU24 8JB. IR (KBr discs) spectra were obtained using a Golden Gate sampling attachment on a Mattson Satellite 3000 FTIR. Raman spectra

School of Chemistry, University of Southampton, Highfield, Southampton, SO17 1BJ, UK. E-mail: M.C.Grossel@soton.ac.uk; Fax: +44 (0)23 8059 3781; Tel: +44 (0)23 8059 3153

† Electronic supplementary information (ESI) available: Including additional figures and key crystallographic information. CCDC 2101940–2101944. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/d1ce01075a

‡ Current address: Leibniz-Forschungsinstitut für Molekulare Pharmakologie, 13125 Berlin, Germany.

§ Current address: Corporate Strategic Research, ExxonMobil Research & Engineering Company, Annandale, NJ 08801, USA.



Scheme 1 Structures of TCNQ complexes (1–5) investigated in this study.



were collected using a Renishaw In-Via system with a high powered near Infrared (HPNIR) 785 nm laser and microscope using a 50× objective. ESI mass spectra were obtained using a solarix (Bruker Daltonics, Bremen, Germany) mass spectrometer equipped with a 4.7 T magnet and FT-ICR cell.

Single crystal X-ray diffraction data collection and processing

A suitable crystal was selected and measured following a standard method²⁸ for each compound. For **1** on a *Rigaku 007HF* diffractometer equipped with Varimax confocal mirrors and an *AFC11* goniometer and an enhanced sensitivity (HG) *Saturn 944+* detector at 100 K. For **2** on a *Rigaku SPIDER RAPID* diffractometer at 120 K with an image plate detector. For **3**, **4**, and **5** on a *Rigaku AFC12* goniometer equipped with HG *Saturn724+* detector mounted at the window of a *FR-E+ SuperBright* molybdenum rotating anode generator with either VHF *Varimax* optics (70 μm focus) (**3** and **4**) or HF *Varimax* optics (100 μm focus) (**5**) at 100 K. Cell determination and data collection were carried out using *CrystalClear*.²⁹ Data reduction, cell refinement, and absorption correction were carried out using either *CrystalClear*²⁹ (**2**) or *CrysAlisPro* (**1**, **3**, **4**, **5**).³⁰ Structures were solved with Olex2³¹ using the *ShelXT*³² and refined using *ShelXL*.³³

General procedure for preparation of simple TCNQ salts (**1**) and (**2**)

A solution of the alkali metal TCNQ salt (1 mol equiv.) and 12-crown-4 (2 mol equiv.) in anhydrous acetonitrile (50 ml) was boiled for 5 minutes, and filtered whilst hot. The mixture was then left to cool. The solvent was allowed to evaporate slowly over a period of several days during which a crystalline solid formed which was separated by filtration. This solid was washed with dry diethyl ether (50 ml), and was then dried in vacuum to give the resulting small crystals.

Preparation of (12-crown-4)₂LiTCNQ (**1**)

Reaction of LiTCNQ (0.424 g, 2 mmol) and 12-crown-4 (0.700 g, 4 mmol) achieved small crystals of **1** as black chunks (0.437 g, 78%). MS (solution) (MeCN) (ESI[−]) *m/z*: 204.1 (TCNQ[−]). (ESI⁺) *m/z*: 375.2 (2 × crown + Na⁺), 199.2 (crown + Na⁺). IR $\nu_{\max}/\text{cm}^{-1}$ 2968, 2915, 2875 (saturated C–H stretch), 2173, 2151 (C≡N stretch), 1589 (C=C(CN)₂ stretch), 1505 (C=C ring stretch), 1361 (C–H bend), 1177 (C–CN and C–C ring stretch), 986 (C–C ring stretch), 819, 718 (C–H out of plane bend). Raman $\nu_{\max}/\text{cm}^{-1}$ 2214 (C≡N stretch), 1604 (C=C ring stretch), 1384 (C–CN stretch), 1208 (C=C–H bending). M.p. 234 °C (dec.). Elemental analysis: calculated: C: 59.68%, H: 6.44%, N: 9.94%. Found: C: 59.80%, H: 6.15%, N: 10.36%.

Preparation of (12-crown-4)₂NaTCNQ (**2**)

Reaction of NaTCNQ (0.454 g, 2 mmol) and 12-crown-4 (0.700 g, 4 mmol) afforded **2** as small shiny black needles (0.269 g,

47%). MS (solution) (MeCN) (ESI[−]) *m/z*: 204.1 (TCNQ[−]). (ESI⁺) *m/z*: 375.2 (2 × crown + Na⁺), 199.2 (crown + Na⁺). IR $\nu_{\max}/\text{cm}^{-1}$ 2962, 2916, 2875 (saturated C–H stretch), 2172, 2150 (C≡N stretch) (lit.³⁴ 2177, 2156), 1567 (C=C(CN)₂ stretch), 1505 (C=C ring stretch), 1361 (C–H bend), 1177 (C–CN and C–C ring stretch), 986 (C–C ring stretch), 820, 718 (C–H out of plane bend). Raman $\nu_{\max}/\text{cm}^{-1}$ 2257 (C≡N stretch), 1609 (C=C ring stretch), 1389 (C–CN stretch), 1206 (C=C–H bending). M.p. 226 °C (dec.) (lit.³⁴ 185–187 °C). Elemental analysis: calculated: C: 58.03%, H: 6.22%, N: 9.67%. Found: C: 57.64%, H: 6.16%, N: 9.48%.

Preparation of (12-crown-4)₂Li(TCNQ)₂ (**3**)²⁷

Reaction of LiTCNQ (0.0211 g, 0.1 mmol), neutral TCNQ⁰ (0.0204 g, 0.1 mmol) and 12-crown-4 (0.0352 g, 0.2 mmol) gave **3** as dark green crystalline solid (0.02122 g, 28%). MS (solution) (MeCN) (ESI[−]) *m/z*: 204.7 (TCNQ[−]). (ESI⁺) *m/z*: 375.6 (2 × crown + Na⁺), 199.4 (crown + Na⁺). IR $\nu_{\max}/\text{cm}^{-1}$ 2970, 2940 (saturated C–H stretch), 2215, 2188 (C≡N stretch), 1558 (C=C(CN)₂ stretch), 1507 (C=C ring stretch), 1327 (C–H bend), 1134 (C–CN and C–C ring stretch), 953 (C–C ring stretch), 861, 694 (C–H out of plane bend). Raman $\nu_{\max}/\text{cm}^{-1}$ 2216 (C≡N stretch), 1606 (C=C ring stretch), 1385 (C–CN stretch), 1207 (C=C–H bending). M.p. 312 °C (dec.).

Preparation of (12-crown-4)₂Na(TCNQ)₂ (**4**)²⁶

Reaction of NaTCNQ (0.0227 g, 0.1 mmol), neutral TCNQ⁰ (0.0204 g, 0.1 mmol) and 12-crown-4 (0.0352 g, 0.2 mmol) gave **4** as black crystalline solid (0.01085 g, 14%). MS (solution) (MeCN) (ESI[−]) *m/z*: 204.5 (TCNQ[−]). (ESI⁺) *m/z*: 375.7 (2 × crown + Na⁺), 199.5 (crown + Na⁺). IR $\nu_{\max}/\text{cm}^{-1}$ 2969, 2935, 2881 (saturated C–H stretch), 2223, 2201, 2177 (C≡N stretch), 1560 (C=C(CN)₂ stretch), 1506 (C=C ring stretch), 1302 (C–H bend), 1136 (C–CN and C–C ring stretch), 951 (C–C ring stretch), 848, 689 (C–H out of plane bend). Raman $\nu_{\max}/\text{cm}^{-1}$ 2212 (C≡N stretch), 1602 (C=C ring stretch), 1382 (C–CN stretch), 1206 (C=C–H bending). M.p. 310 °C (dec.).

Preparation of (12-crown-4)₂K(TCNQ)₂ (**5**)

Reaction of KTCNQ (0.0972 g, 0.4 mmol), neutral TCNQ⁰ (0.0816 g, 0.4 mmol) and 12-crown-4 (0.1408 g, 0.8 mmol) afforded **5** as a black crystalline solid (0.01942 g, 24%). MS (solution) (MeCN) (ESI[−]) *m/z*: 204.4 (TCNQ[−]). (ESI⁺) *m/z*: 375.6 (2 × crown + Na⁺), 199.4 (crown + Na⁺). IR $\nu_{\max}/\text{cm}^{-1}$ 2965, 2926, 2877 (saturated C–H stretch), 2195, 2179, 2167 (C≡N stretch), 1560 (C=C(CN)₂ stretch), 1508 (C=C ring stretch), 1363 (C–H bend), 1181 (C–CN and C–C ring stretch), 956 (C–C ring stretch), 828, 700 (C–H out of plane bend). Raman $\nu_{\max}/\text{cm}^{-1}$ 2217 (C≡N stretch), 1608 (C=C ring stretch), 1390 (C–CN stretch), 1204 (C=C–H bending). M.p. >300 °C (dec.). Elemental analysis: calculated: C: 60.06%, H: 5.04%, N: 14.00%. Found: C: 59.82%, H: 4.44%, N: 15.74%.



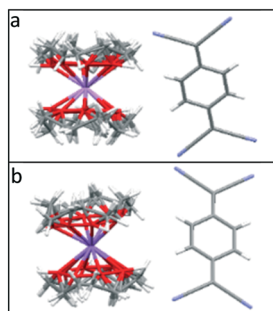


Fig. 1 Unit cell of TCNQ complexes (a). (12-crown-4)₂LiTCNQ (**1**) and (b). (12-crown-4)₂NaTCNQ (**2**). In each case, the 12-crown-4 ether ligands are disordered.

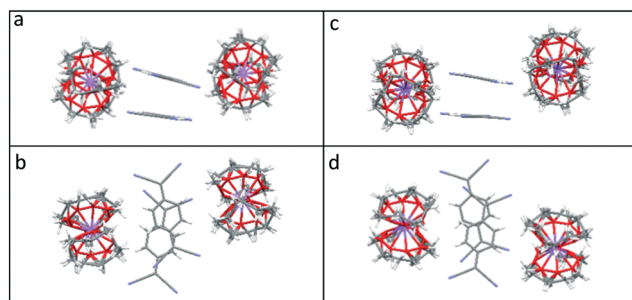


Fig. 2 Side (a) and top (b) views of the TCNQ^{••} dimer in (12-crown-4)₂LiTCNQ (**1**) in respect to the views (c and d) of (12-crown-4)₂NaTCNQ (**2**).

Result and discussion

Complexes involving “simple” TCNQ^{••} salts (**1** and **2**)

Crystallisation of Li or NaTCNQ from anhydrous acetonitrile in the presence of two equivalents of 12-crown 4 afforded black single crystals of the bis-crown ether complexes (**1**) and (**2**) suitable for X-ray structural study. It should be noted that crystallisation of LiTCNQ with 1 equivalent of 12-crown-4 also only produced that bis-crown ether complex (**1**). Attempts to isolate crystalline samples of (12-crown-4)₂KTCNQ also failed, however the salt (12-crown-4)₂K(TCNQ)₂ (**5**) was successfully prepared instead.

For (12-crown-4)₂LiTCNQ (**1**) and (12-crown-4)₂NaTCNQ (**2**), some difficulties were encountered during the X-ray crystal structure solution as the 12-crown-4 moieties are disordered. The atomic occupancy of the disordered 12-crown-4 units in **1**

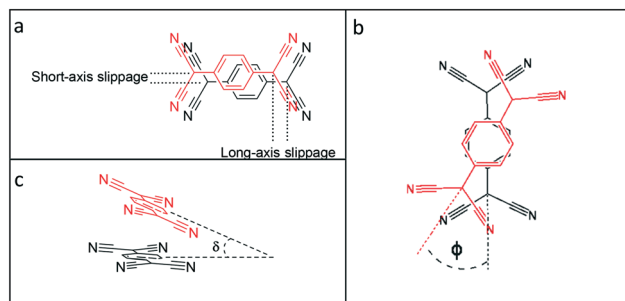


Fig. 3 Definition of short and long-axis slippage (a), twist angle ϕ (b), and tilt angle δ (c) in this study (see Table 1).

is 0.504(2): 0.496(2) and 0.5: 0.5 in **2**, respectively. Complexes **1** and **2** are iso-structural. In each case, the alkali metal cation (Li⁺ and Na⁺) is sandwiched between two 12-crown-4 units and is thus completely screened from the TCNQ^{••} counter anion (see Fig. 1). This situation is similar to that seen previously for the (15-crown-5)₂KTCNQ (**6**),¹ in which one K⁺ cation is sandwiched between two disordered 15-crown-5 ether ligands at room temperature, the presence of the ionophore sandwich preventing direct cation–anion contacts.¹

The TCNQ^{••} anions are face-to-face π -stacked in both **1** and **2** (see Fig. 2), in which neighbouring anion planes are tilted in respect to each other (6.27° in **1** and 6.89° in **2**) forming a novel “cross-stitch”-like stack (see Fig. 2a and c). A top view of adjacent TCNQ^{••} units reveals that neighbouring TCNQ^{••} anions are significantly long-axis slipped and slightly short-axis slipped (see Table 1 and Fig. 3 for the key structural details). Furthermore, adjacent TCNQ^{••} anions are mutually twisted in respect to each other (26.02° in **1** and 26.00° in **2**) (see Fig. 2b and d). To our knowledge, this “cross-stitch”-like twisted packing motif of neighbouring TCNQ^{••} units is a novel packing motif for alkali metal TCNQ complexes.

In the solid state, both **1** and **2** the TCNQ^{••} moiety and the cation-crown ether complex (12-crown-4)₂M⁺ (M⁺ = Li⁺ and Na⁺) form mixed 1:2 alternating sheets, in a manner similar to that previously reported for **6**¹ (see Fig. 4a and b).

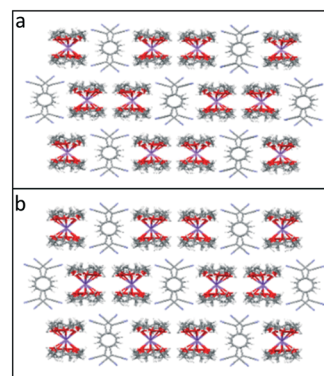


Fig. 4 A top view of the solid-state structure of (12-crown-4)₂LiTCNQ (**1**, a) and (12-crown-4)₂NaTCNQ (**2**, b).

Table 1 Key structural data for TCNQ dimers in the complexes (**1–5**) discussed here with comparative literature data (see Fig. 3 for definitions of long- and short-axis slippage, twist angle, and tilt angle)

Complex	π - π separation/Å	Long-axis slippage/Å	Short-axis slippage/Å	Ref.
1 . (at 100 K)	—	2.40	0.26	This work
2 . (at 100 K)	—	2.34	0.23	This work
3 . (at 100 K)	3.153	1.99	0.48	This work
4 . (at 100 K)	3.123	2.04	0.33	This work
5 . (at 100 K)	3.125	2.03	0.30	This work



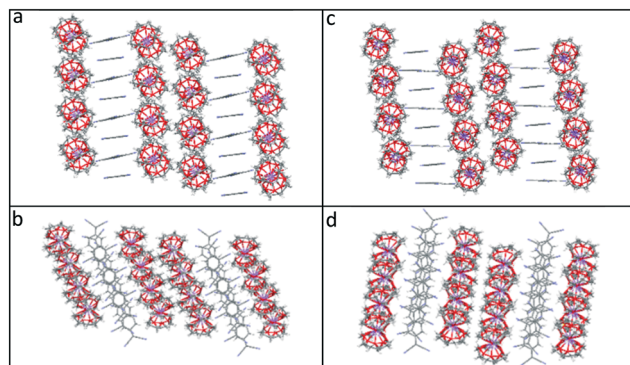


Fig. 5 Comparison of the side and top views of the extended columns formed by (12-crown-4)₂LiTCNQ (1), (a) and (b) respectively, and by (12-crown-4)₂NaTCNQ (2) (c) and (d) respectively.

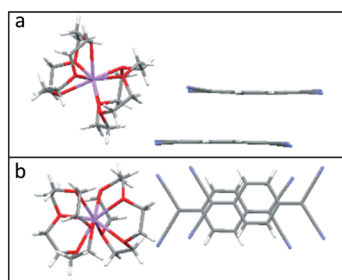


Fig. 6 Side (a) and top (b) views of the (TCNQ)₂²⁻ dimer formed in (12-crown-4)₂Li(TCNQ)₂ (3).

Viewed from the side, both the TCNQ^{•-} dimers and the cation barrels formed by the (12-crown-4)₂M⁺ (M⁺ = Li⁺ and Na⁺) moieties in 1 and 2 are assembled further into infinite extended columns (see Fig. 5). The centroid-centroid distance between adjacent TCNQ^{•-} units within the column is 4.155 Å in 1 and 4.154 Å in 2. This packing motif in both 1 and 2 is clearly different from that seen for 6. In the latter case, TCNQ^{•-} anion dimers are isolated from each other, being surrounded by (15-crown-5)₂K⁺ cation barrels.¹

Complex TCNQ salts (3, 4, and 5)

Crystallisation of Li⁺, Na⁺, and K⁺TCNQ^{•-} from anhydrous acetonitrile in the presence of 12-crown-4 and one molar

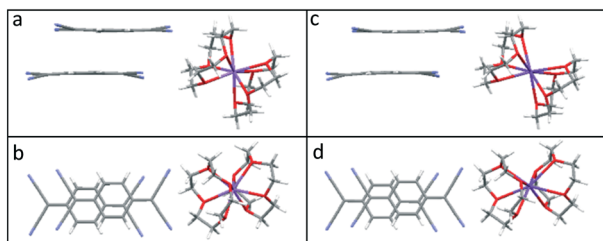


Fig. 7 Side and top views of the (TCNQ)₂²⁻ dimer formed in (12-crown-4)₂Na(TCNQ)₂ (4) (a and b) and in (12-crown-4)₂K(TCNQ)₂ (5) (c and d).

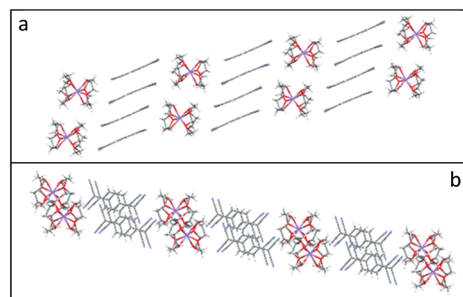


Fig. 8 Side (a) and top (b) views of the sheets of face-to-face π -stacked columns of TCNQ dimers formed in (12-crown-4)₂Li(TCNQ)₂ (3), revealing the long-axis slipped character of the TCNQ dimer.

equivalent of neutral TCNQ⁰ afforded crystals of the complex salts of (12-crown-4)₂M(TCNQ^{•-})(TCNQ⁰) (M⁺ = Li⁺, Na⁺, or K⁺) 3, 4, and 5, respectively. These three complex TCNQ salts (3, 4, and 5) are iso-structural.

In the solid-state, the alkali metal cation (Li⁺, Na⁺, or K⁺) in each of the complex TCNQ salts (3–5) is coordinated to and sandwiched between two 12-crown-4 units. The resulting cation “barrels” of (12-crown-4)₂M⁺ (M⁺ = Li⁺, Na⁺, and K⁺) inhibit any direct interaction between alkali metal cations and the TCNQ units (M⁺–NC), with the cation complexes sitting in channels between the infinite columns of face-to-face π -stacked (TCNQ)₂^{•-} moieties (see Fig. 6 for example).

In 3–5, neighbouring (TCNQ)₂^{•-} units form extended face-to-face π -stacked columns. Within the (TCNQ)₂^{•-} units, the TCNQ moieties are significantly long-axis slipped (see Table 1 and Fig. 3 for the key structural details) with a vertical π – π separation of 3.153 Å in 3, 3.123 Å in 4, and 3.125 Å in 5, respectively. Additionally, individual TCNQ units adopt a shallow boat conformation, as previously reported by Akutagawa for Na⁺[(12-crown-4)₂(TCNQ)₂] at 100 K²⁶ (see Fig. 7).

Furthermore, in 3–5, two pairs of TCNQ dimer units assemble into extended tetramers forming parallel sheets, which are separated by the corresponding cation barrels of (12-crown-4)₂M⁺ (M⁺ = Li⁺, Na⁺, and K⁺) (see Fig. 8). This observation of tetramer packing motif is consistent with that reported by Akutagawa and co-workers for Li⁺[(12-

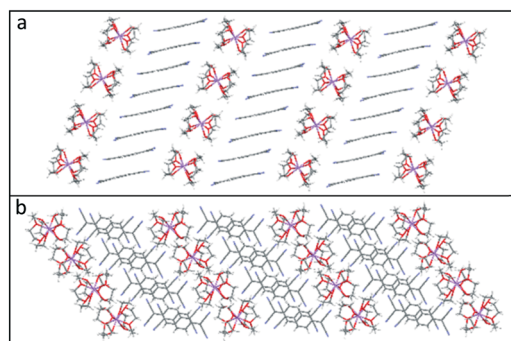


Fig. 9 Side (a) and top (b) views of the packing motif of the alternating infinite columns of TCNQ units and cation barrels in (12-crown-4)₂Li(TCNQ)₂ (3).



Table 2 Infrared and Raman data for the TCNQ complexes (1–5) reported in this work together with selected comparative literature data

Compound	Infrared data/cm ⁻¹		Raman data/cm ⁻¹			Ref.
	C≡N stretch	C=C stretch	C≡N stretch	C=C stretch	C–C≡N stretch	
TCNQ ⁰	2228, 2225	1545	2225	1600	1450	36
TCNQ ⁰	2224, 2220	1545	2230	1603	1454	This work
LiTCNQ	2201, 2171	1501				This work
NaTCNQ	2197, 2184, 2160	1505	2204	1602	1384	This work
KTCNQ	2215, 2152	1505				This work
1	2173, 2151	1505	2214	1604	1384	This work
2	2172, 2150	1505	2206	1609	1389	This work
3	2222, 2198, 2165	1507	2216	1606	1385	This work
4	2223, 2201, 2177	1506	2212	1602	1382	This work
5	2195, 2179, 2167	1508	2217	1608	1390	This work

crown4)₂(TCNQ)₂, and Na⁺[(12-crown4)₂(TCNQ)₂], in which two kinds of TCNQ units (**A** and **B**) form (**A**⋯**B**) face-to-face π -stacked dimers at 100 K with additional interdimer interactions ⋯(**A**⋯**B**)⋯(**B**⋯**A**)⋯.^{26,27} Fig. 8 illustrates this behaviour for **3** (please refer to ESI† (Fig. S1 and S2) for the corresponding information in **4** and **5**).

This repetitive pattern of tetramers further assembles into infinite TCNQ columns (see Fig. 9). The similarity of the bond lengths within the TCNQ columns makes it difficult to distinguish between the TCNQ^{•-} and neutral TCNQ⁰ components [please refer to ESI† (Fig. S3 and S4) for the corresponding figures for **4** and **5**].

Vibrational spectra of ionophore-encapsulated alkali metal TCNQ complexes (1–5)

Vibrational spectroscopy provides a convenient probe into the chemical nature of TCNQ species.³⁵ Table 2 presents key data for TCNQ complexes 1–5 together with literature data for comparison. In the infrared (IR) region, two bands are seen which are associated with the C≡N stretches characteristic of the TCNQ^{•-} anion in **1** and **2**. An extra band (at ~2200 cm⁻¹) indicates the presence of neutral TCNQ⁰ in **3**–**5**. In the spectra, the C=C and C≡N stretching bands of TCNQ units will generate obvious Raman peaks because the π electrons are strongly polarisable.³⁷ For the TCNQ complexes (1–5), the peaks in the region 2217 and 2206 cm⁻¹ are ascribed to stretching bands for cyano groups in TCNQ molecule.³⁷ The peak between 1390 and 1380 cm⁻¹ reflects electron transfer between neutral TCNQ⁰ and TCNQ^{•-}.^{37,38}

Conclusions

In this paper, we have explored the solid-state architecture of five ionophore-encapsulated alkali metal TCNQ complexes (1–5). The “simple” TCNQ^{•-} salts **1** and **2** are iso-structural. The 12-crown-4 moiety in both complexes is disordered with one alkali metal cation sandwiched between two crown ethers. Neighbouring TCNQ^{•-} anions are twisted with respect to each other and form face-to-face π -stacks with further assembling into infinite columns instead of isolated dimers as observed in (15-crown-5)₂K⁺ analogue.¹ In the solid-state,

the three complex TCNQ salts (**3**, **4**, and **5**) consist of (TCNQ)₂^{•-} units, which are packed into repetitive tetramers throughout the crystal structures and separated by cation barrels of (12-crown-4)₂M⁺ moieties.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

We thank the EPSRC UK National Crystallography Service at the University of Southampton for the collection of the crystallographic data.²⁸

References

- M. C. Grossel and S. C. Weston, *Chem. Mater.*, 1996, **8**, 977–980.
- M. C. Grossel and S. C. Weston, *Contemp. Org. Synth.*, 1994, **1**, 367–386.
- V. F. Kaminskii, R. P. Shibaeva and L. O. Atovmyan, *J. Struct. Chem.*, 1974, **14**, 1014–1019.
- V. F. Kaminskii, R. P. Shibaeva and L. O. Atovmyan, *J. Struct. Chem.*, 1974, **14**, 646–650.
- R. P. Shibaeva and L. O. Atovmyan, *J. Struct. Chem.*, 1972, **13**, 514–531.
- R. Shibaeva, L. Atovmyan, V. Ponomarev, O. Philipenko and L. Rozenberg, *Tetrahedron Lett.*, 1973, **14**, 185–188.
- M. C. Grossel, P. B. Hitchcock, K. R. Seddon, T. Welton and S. C. Weston, *Chem. Mater.*, 1994, **6**, 1106–1108.
- L. R. Melby, R. J. Harder, W. R. Hertler, W. Mahler, R. E. Benson and W. E. Mochel, *J. Am. Chem. Soc.*, 1962, **84**, 3374–3387.
- J. G. Vegter, T. Hibma and J. Kommandeur, *Chem. Phys. Lett.*, 1969, **3**, 427–429.
- D. Chesnut and W. Phillips, *J. Chem. Phys.*, 1961, **35**, 1002–1012.
- D. Chesnut and P. Arthur, *J. Chem. Phys.*, 1962, **36**, 2969–2975.
- P. Nordio, Z. Soos and H. McConnell, *Annu. Rev. Phys. Chem.*, 1966, **17**, 237–260.



- 13 T. Hibma, P. Dupuis and J. Kommandeur, *Chem. Phys. Lett.*, 1972, **15**, 17–20.
- 14 T. Hibma and J. Kommandeur, *Phys. Rev. B: Solid State*, 1975, **12**, 2608.
- 15 A. Silverstein and Z. Soos, *Chem. Phys. Lett.*, 1976, **39**, 525–530.
- 16 J. R. Morton, K. F. Preston, M. D. Ward and P. J. Fagan, *J. Chem. Phys.*, 1989, **90**, 2148–2153.
- 17 M. C. Grossel and S. C. Weston, *J. Phys. Org. Chem.*, 1992, **5**, 533–539.
- 18 M. C. Grossel, F. A. Evans, J. A. Hriljac, K. Prout and S. C. Weston, *J. Chem. Soc., Chem. Commun.*, 1990, 1494–1495.
- 19 R. C. Hynes, J. R. Morton, K. F. Preston, A. J. Williams, F. Evans, M. C. Grossel, L. H. Sutcliffe and S. C. Weston, *J. Chem. Soc., Faraday Trans.*, 1991, **87**, 2229–2233.
- 20 J. Lu, X. Qu, G. Peleckis, J. F. Boas, A. M. Bond and L. L. Martin, *J. Org. Chem.*, 2011, **76**, 10078–10082.
- 21 J. Huang, S. Kingsbury and M. Kertesz, *Phys. Chem. Chem. Phys.*, 2008, **10**, 2625–2635.
- 22 M. C. Grossel, F. A. Evans, J. A. Hriljac, J. R. Morton, Y. LePage, K. F. Preston, L. H. Sutcliffe and A. J. Williams, *J. Chem. Soc., Chem. Commun.*, 1990, 439–441.
- 23 M. C. Grossel, F. A. Evans, J. A. Hriljac, K. Prout and S. C. Weston, *J. Chem. Soc., Chem. Commun.*, 1990, 1494–1495.
- 24 M. C. Grossel and S. C. Weston, *J. Chem. Soc., Chem. Commun.*, 1992, 1510–1512.
- 25 B. Yan, P. N. Horton, A. E. Russell, C. J. Wedge, S. C. Weston and M. C. Grossel, *CrystEngComm*, 2019, **21**, 3273–3279.
- 26 K. Sambe, N. Hoshino, T. Takeda, T. Nakamura and T. Akutagawa, *Cryst. Growth Des.*, 2020, **20**, 3625–3634.
- 27 K. Sambe, N. Hoshino, T. Takeda, T. Nakamura and T. Akutagawa, *J. Phys. Chem. C*, 2020, **124**, 13560–13571.
- 28 S. J. Coles and P. A. Gale, *Chem. Sci.*, 2012, **3**, 683–689.
- 29 *CrystalClear*, Rigaku Corporation, The Woodlands, Texas, U.S.A., 2008–2014.
- 30 *CrysAlisPro Software System*, Rigaku Oxford Diffraction, 2021.
- 31 O. V. Dolomanov, L. J. Bourhis, R. J. Gildea, J. A. Howard and H. Puschmann, *J. Appl. Crystallogr.*, 2009, **42**, 339–341.
- 32 G. M. Sheldrick, *Acta Crystallogr., Sect. A: Found. Adv.*, 2015, **71**, 3–8.
- 33 G. M. Sheldrick, *Acta Crystallogr., Sect. C: Struct. Chem.*, 2015, **71**, 3–8.
- 34 S. C. Weston, *PhD thesis*, University of London, Royal Holloway and Bedford New College, 1992.
- 35 R. Ramanathan, S. Walia, A. E. Kandjani, S. Balendran, M. Mohammadtaheri, S. K. Bhargava, K. Kalantar-Zadeh and V. Bansal, *Langmuir*, 2014, **31**, 1581–1587.
- 36 T. Takenaka, *Bull. Inst. Chem. Res., Kyoto Univ.*, 1969, **47**, 387–400.
- 37 E. Faulques, A. Leblanc, P. Molinié, M. Decoster, F. Conan, J. E. Guerchais and J. Sala-Pala, *Spectrochim. Acta, Part A*, 1995, **51**, 805–819.
- 38 C. Ye, G. Cao, F. Fang, H. Xu, X. Xing, D. Sun and G. Chen, *Micron*, 2005, **36**, 461–464.

