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Fabrication of heterostructure multilayer devices through the optimization of Bi-metal sulfides for high-performance quantum dot-sensitized solar cells

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In this work, a titanium dioxide and lead sulfide (TiO₂/PbS) nano-size heterostructure with tin sulfide was fabricated and coated *via* a two-step direct deposition process. Its microstructure, morphology, elemental composition, optical absorption, and photochemical activity were investigated. Linear sweep voltammetry and cyclic voltammetry curves substantiated its catalytic activity, indicating quantum dot effects of a well-developed space charge domain on the surface of the hybrid structure. These give rise to electron–hole recombination suppression and a high charge mobility rate. Moreover, direct stabilization was identified in current density, corresponding to the hybrid structures limiting the diffusion current process. Higher J_{SC} values observed were substantiated by the role of quantum dot-size effects and enhanced crystalline structures, leading to a reduction in series resistance and an improved conversion efficiency of 10.05%. Overall, theoretical analyses and empirical findings indicated that the seamless migration of photoexcited electrons across the interfaces of SnS and PbS is linked to quantum dot effect synergy. This is facilitated by the space charge region, which serves as a conduit for efficient electron transfer between the respective materials.

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Introduction

The need for clean energy sources to mitigate the impact of global climate change has made clean energy conversion devices such as solar photovoltaic (PV) cells a highly sought-after solution.^{1–3} Third-generation solar cells such as dye-sensitized solar cells (DSSCs) have emerged with tremendous prospects because of their easy fabrication, low cost, realistic power conversion efficiencies, environmental friendliness, and ease of sandwich build-up.^{3–6} To improve the conversion efficiencies of DSSCs in the past two decades, studies have been carried out with stable and efficient modern dyes, fast electron transfer, excellent photon scattering ability, and a new high-surface-area photoanode electrode.^{7–9} Organic sensitizers have shown pathways for highly efficient electron transfer, which is important for DSSCs for donating and accepting electrons, along with their good chemical, photochemical, and long-term stability.¹⁰ To further enhance the efficiency of this technique, great attention has been paid to conventional liquid electrolytes

with over 14% power conversion efficiency (PCE).¹¹ To achieve the theoretical values of 44% in practical applications of DSSCs, semiconductor quantum dot (QD) engineering is of great importance in increasing the PCE of quantum dot-sensitized solar cells (QDSSCs). Tin sulfide (SnS) is another interesting material whose *I*–*V* measurements with several metals form a Schottky and ohmic contact in Al–SnS–In and Ag–SnS–In systems.^{3,12} SnS is chemically stable, nontoxic, naturally abundant, cheap, and easy to fabricate, making it suitable for QDSSC applications.^{13–15} Moreover, its high absorption coefficient and optical band gap (<2 eV) with a p-type nature give rise to better charge migration, which is crucial for QDSSC devices.¹⁶ Moreover, according to density functional theory calculations, SnS has emerged as a highly useful material in QDSSCs and hydrogen production.^{3,17}

Previous studies have shown that higher efficiencies can be achieved by combining QDs with other materials.^{18,19} Nanostructure materials such as Sb₂S₃, PbS, CdSe, PbSe, and CdS have been used as absorbers in QDSSCs. Among these semiconductors, lead sulfide (PbS) QDs have a tunable band gap of 0.41 eV with a high absorption coefficient and could be explored as co-absorbers with other metal sulfides such as SnS to enhance the PCE of QDSSCs.¹⁸ This is because the theoretical efficiency, multiple excitons, and quantized discrete band structures that are above the Shockley–Queisser limit are in favour of PbS absorbers.²⁰ The low performance of PbS

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absorbers in QDSSCs as compared with other metal sulfide QD materials^{21,22} paves the way for more studies on the development and exploration of PbS with other semiconductors. The passivation of the PbS layer directly above the SnS-absorber surface of the TiO₂ electrode can reduce the photoinduced recombination of electrons with the polysulfide electrolyte and improve the PCE of QDSSCs.¹⁹ This PCE can be achieved by combining PbS and SnS *via* step-wise modification of band-edge levels.²³

Several established deposition techniques such as thermal evaporation,²⁴ atomic layer deposition,²⁵ electrochemical deposition,²⁶ chemical vapor deposition,²⁷ sputtering,²⁸ chemical bath deposition,²⁹ spray pyrolysis,³⁰ and E-beam evaporation³¹ have been reported for the coating of SnS and PbS semiconductors onto the surface of TiO₂ electrode substrates. All the above-mentioned approaches are not cost-effective because they suffer from slow deposition rates or vacuum-based technology. Presently, most attention is focused on the fabrication of nanostructure materials through optimization to form uniform sizes and shapes since their physicochemical and optoelectronic properties are strongly influenced by the optimization. Therefore, the formation of controlled metal sulfide nanostructures with good optoelectronic properties will benefit both the PCE QDSSCs and their commercialization.³¹ The single-source precursor (SSP) approach is one of the most attractive, cost-friendly, and large-scale production, which is accurate and easy to control, with additional benefits over metal sulfide nanostructured fabrication. Herein, we report on the indirect loading of PbS and SnS solution-process-controlled synthesis of semiconductor metal sulfides nanostructured using SSPs on the surface of TiO₂ glass substrates in QDSSCs.

Experimental

Materials and methods

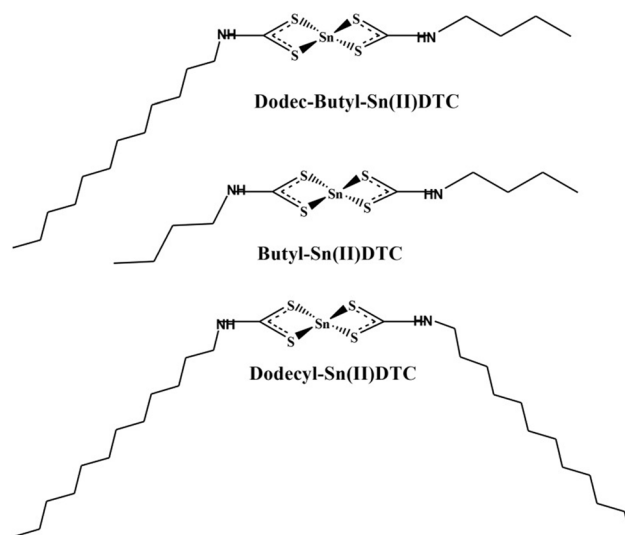
All chemical reagents are of high-purity grade, and were used as received from Sigma Aldrich. Ligands such as *N*-butyldithiocarbamate and *N*-dodecyldithiocarbamate were synthesized according to the procedure earlier reported in ref. 32. The following chemicals were purchased: methanol, oleic acid, tin(II) chloride dihydrate (Cl₂Sn·2H₂O (<https://www.ereztech.com/product/stannous-chloride-dihydrate-10025-69-1/>)), oleylamine (ODE), tri-*n*-octylphosphine oxide (TOPO), ammonium, carbon disulfides (CS₂), diethyl ester, butylamine, and dodecylamine. Bis(*N*-diisopropyl-*N*-octyldithiocarbamate) (OCT) Pb(II) complexes were obtained from Merck, which has been reported in our previous study.²² Testing kits containing fluorine-doped tin oxide (FTO) platinum electrode (drilled) (size: 20 × 20 mm; active area: 6 × 6 mm), TiO₂ FTO electrode (opaque) (size: 20 × 20 mm; active area: 6 × 6 mm), Iodolyte HI-30 polysulfide electrolyte, gaskets, masks, and hot seal were purchased from Solaronix. Water, PbS and SnS nanoparticles were obtained from dithiocarbamate molecular precursors.

[Dodecylidtc] yield: 11.6789 g (69.80%) ¹H NMR (DMSO) δ 3.23 (s, 2H-NH), 0.82 (t, 3H-CH₃), 1.53–1.56 (s, 2H-CH₂), 1.25–1.38 (s, 1H-SH). ¹³C NMR (DMSO) δ 22.6–26.3 (–CH₂), 198 (–CS₂),

14.4 (–CH₃), 45.2 (–C-NH). Selected IR (cm^{–1}) 1412 ν(C–N), 1218 ν(C–S), 3454 ν(N–H), 2936 ν(CH₃). UV-vis (CH₃OH solution, nm): 236. [Butyldtc] yield: 5.3103 g (93.33%) ¹H NMR (DMSO) δ 3.24–3.33 (s, 2H-NH), 0.87 (t, 3H-CH₃), 1.65–2.51 (s, 2H-CH₂), 1.28 (s, 1H-SH). ¹³C NMR (DMSO) δ 19.7–29.5 (–CH₂), 198.5 (–CS₂), 13.6 (–CH₃), 44.9 (–C-NH). Selected IR (cm^{–1}) 1485 ν(C–N), 1097 ν(C–S), 3276 ν(N–H), 2908 ν(CH₃). UV-vis (CH₃OH solution, nm): 369.

Synthesis of bis(*N*-butyl-*N*-dodecyldithiocarbamate)Sn(II) complexes [Dodec-butyldtc] (1)

The complexes were synthesized following the procedure previously reported in ref. 33 with slight modification, as shown in Scheme 1. *N*-Butyldithiocarbamate and *N*-dodecyldithiocarbamate ligands (0.4158 g, 2.5 mmol and 0.6964 g, 2.5 mmol) were each dissolved separately in 15 mL of distilled water. SnCl₂·2H₂O salts (0.5641 g, 2.5 mmol) were then mixed in a ratio of 1/1/1 together with continuous stirring for 2 h. The mixtures gave rise to yellow precipitates, which were filtered and thoroughly rinsed with cold methanol and dried. Formulated as [Sn(Dodec-butyldtc)] (C₁₈H₃₈N₂S₄Sn) yield: 3.7091 g (46.23%) ¹H NMR (DMSO) δ 3.25–3.68 (s, 2H-NH), 0.85–0.92 (t, 3H-CH₃), 1.55–1.64 (s, 2H-CH₂), 1.25–1.38 (s, 1H-SH). ¹³C NMR (DMSO) δ 19.6–31.8 (–CH₂), 200 (–CS₂), 13.6–14.4 (–CH₃), 39.2 (–C-NH). Selected IR (cm^{–1}) 1515 ν(C–N), 944 ν(C–S), 3408 ν(N–H), 580 ν(M–S), 2851 ν(CH₂), 2918 ν(CH₃). UV-vis (CH₃OH solution, nm): 280. Same steps were followed for single ligands: *N*-butyldithiocarbamate and *N*-dodecyldithiocarbamate with SnCl₂·2H₂O salts in ratio 2/1 and formulated as [Sn(Dodecylidtc)] and [Sn(Butyldtc)]. [Sn(Dodecylidtc)] (2) (C₂₆H₅₀N₂S₄Sn) yield: 3.5284 g (44.12%) ¹H NMR (DMSO) δ 3.65–3.68 (s, 2H-NH), 0.87–0.92 (t, 3H-CH₃), 2.51–2.76 (m, 2H-CH₂), 1.28–1.65 (s, 1H-SH). ¹³C NMR (DMSO) δ 19.7–31.7 (–CH₂), 201 (–CS₂), 13.6–14.2 (–CH₃), 40 (–C-NH). Selected IR (cm^{–1}) 1517 ν(C–N), 936 ν(C–S), 3306 ν(N–H), 541 ν(M–S), 2851 ν(CH₂), 2919 ν(CH₃). UV-vis (CH₃OH solution, nm): 293. [Sn(Butyldtc)] (3) (C₁₀H₂₀N₂S₄Sn)



Scheme 1 Structures of Sn(Dodec-butyldtc), Sn(Dodecylidtc) and Sn(Butyldtc) complexes.



yield: 2.0176 g (33.16%) ^1H NMR (DMSO) δ 3.23–3.65 (s, 2H–NH), 0.84–0.92 (t, 3H–CH₃), 1.53–1.56 (s, 2H–CH₂), 1.24–1.25 (s, 1H–SH). ^{13}C NMR (DMSO) δ 22.6–31.8 (–CH₂), 200 (–CS₂), 14.4 (–CH₃), 40 (–C–NH). Selected IR (cm^{−1}) 1503 ν (C–N), 1232 ν (C–S), 3173 ν (N–H), 406–611 ν (M–S), 2843 ν (CH₂), 2927 ν (CH₃). UV-vis (CH₃OH solution, nm): 280.

Formation of tin sulfide SnS nanoparticles

The synthesis of SnS was done with slight modifications in line with the procedure already reported in the literature.^{32,33} First, 0.2 g of [Dodec-butylidtc] (1) was dispersed in 4 mL of oleylamine and TOPO in a three-necked bottom flask (250 mL). The resulting mixture was degassed for 10 min at about 20–30 °C, heated up to 250 °C and then maintained for 1 h in an inert atmosphere. Afterwards, the resulting solution was obtained by naturally cooling to 70 °C followed by the addition of 30 mL of methanol for the precipitation of the SnS nanoparticles. The obtained precipitates were washed severally with methanol and centrifuged to obtain the final end product as SnS nanoparticles, which is labelled as SnS1. Similar steps were taken for [Dodecylidtc] (2) and [Butylidtc] (3), labelled as SnS2 and SnS3. The bis(*N*-diisopropyl-*N*-octyldithiocarbamate) Pb(II) complex was used to synthesize PbS as reported previously.²²

Fabrication of QDSCs

The TiO₂ electrode was soaked in 0.010 g PbS dissolved in 10 mL of distilled water for 24 hours, rinsed with distilled water, and then soaked in 0.010 g SnS in 10 mL of distilled water for 24 hours. The sensitized FTO-TiO₂/PbS/SnS films were used as photoanodes, with a FTO-Pt film as the counter electrode. The Iodolyte HI-30 electrolyte solution was injected between the photoanode and counter electrode through a siphonic effect. The QDSSC device was assembled as TiO₂/PbS/SnS/HI-30/Pt with an active area of 6 × 6 mm for the three devices. The assembled devices were sealed under a hot press and labelled as P/S#1, PS#2, and PS#3. To the best of our knowledge, both materials synthesized from these molecular precursors have not been reported as co-absorbents in QDSCs. These parameters were used to calculate the conversion efficiency (%) of the prepared devices using eqn (1)–(4):

The efficiency of the cells developed in this work was estimated using eqn (1)–(5):⁵

$$J_{\text{L-ideal}} = q \int_{E_g}^{\infty} N_0(E) dE \quad (1)$$

$$J(V) = J_{\text{L-ideal}} - J_{0,\text{rad}} \left(\exp\left(\frac{V}{V_T}\right) - 1 \right) \quad (2)$$

$$V_{\text{OC-ideal}} = V_T \ln \left[\left(\frac{J_{\text{L-ideal}}}{J_{0,\text{rad}}} \right) + 1 \right] \quad (3)$$

$$\eta_{\text{limit}} = \left(\frac{J_{\text{L-ideal}} V_{\text{OC-ideal}}}{P_{\text{in}}} \right) \text{FF}_{\text{ideal}} \quad (4)$$

$$\text{FF}_{\text{ideal}} = P_{\text{max}} / J_{\text{L-ideal}} V_{\text{OC-ideal}} \quad (5)$$

P_{max} is the maximum power density as: $P(V) = VxP(V)$ curve.

Physical characterizations

X-ray diffraction was performed within a rotation anode high-power X-ray diffractometer (XRD) (Bruker D8 Advance) with Cu-K α radiation ($k = 1.5406 \text{ \AA}$) at 40 kV and 40 mA. Surface morphology and the chemical constituents of the fabricated DSSC cells were analyzed by field emission scanning electron microscopy and X-ray energy-dispersive spectroscopy using FESEM, JEOL, JSM-6700F, 5 kV with a Nanotrace EDX detector (Thermo Electron) and transmission electron microscopy (TEM, JEOL JEM-2010) at 200 kV. Functional group identification was achieved by Fourier transform infrared (FTIR) spectroscopy using a PerkinElmer Spectrum 1. The optical properties of the cells were measured using a UV-visible and PL spectrometer by Perkin Elmer-Lambda 45. We measured the surface's roughness at a scan rate of 0.8 Hz using an atomic force microscope (JPK NanoWizard II AFM, JPK). Keithley 2400 with Source Meter (AM) 1.5G (1000 Wm^{−2}) was used to measure the IV properties of the as-prepared devices. A Gamry 10101E electrochemical workstation equipped with a three-electrode device was employed to evaluate the electrocatalytic performance of the three samples. The electrochemical impedance spectra (EIS) were analyzed while operating at a frequency range of 1 Hz–10⁷ Hz at 10 mV. Cyclic voltammetry (CV) and linear sweep voltammetry (LSV) analyses were performed at scan rates of 100 and 10 mV s^{−1}. Chronoamperometry (CA) was evaluated to study the effect of passivation on the stability of the cells.

Results and discussion

AFM and SEM for SnS quantum dot photon absorbers

Fig. 1a–f display the AFM images of 2-D and 3-D for the three samples. The obtained values for the surface roughness are 324 nm, 408 nm and 454 nm for SnS1, SnS2 and SnS3 respectively. The increases in the surface roughness can be linked to the quantum-sized effects.³⁴ The low surface roughness is due to the increase in size distribution. Moreover, the optical

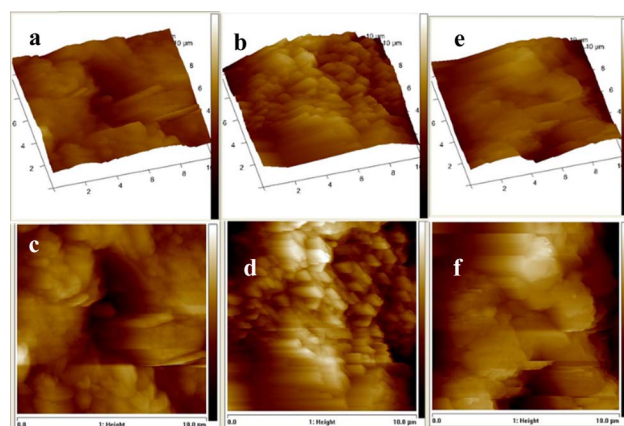


Fig. 1 Three- and two-dimensional AFM images of SnS1 (a and c), SnS2 (b and d) and SnS3 (e and f) nanoparticles.



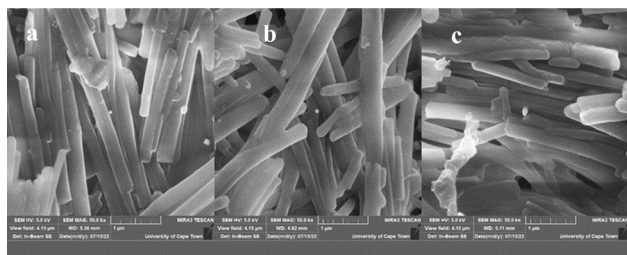


Fig. 2 SEM images of SnS1 (a), SnS2 (b) and SnS3 (c) nanoparticles.

properties and conversion efficiency of QDSCs rest on the low surface roughness. Fig. 2a–c reveal the morphology of the samples where SnS nano-rods with smooth surfaces are densely packed together with a good crystal feature for all the samples. SEM micrographs follow a similar trait as with the AFM results obtained in this study.^{35,36} Co-absorbance of PbS and SnS nanorods blend into an electron transporting layer like TiO₂ is important for solar cell applications and allows for fast electron transport at the electrolyte interface.

The thermal stability of SnS was identified by thermogravimetric (TGA) analysis, as shown in Fig. 3. The stability of the SnS QDs was found at 354–362 °C, with complete loss of organic species from the structure of the molecular precursor, as reported in our previous work.⁴ The high thermal stability of SnS makes it a potential material for the effective optimization of various optoelectronic properties, which pave the way for their applications and improved conversion efficiency obtained in the present study.

XRD, Raman and FTIR spectra of P/S#1, P/S#2, and P/S#3

The structure of the prepared devices was characterized using an XRD instrument, as illustrated in Fig. 4a. The Crystal Impact software was used for the identification and quantification of phase indexes and the elemental weights of the assembled electrodes as shown in Table 1. The XRD peaks show the preferential orientation of the modified cells, which corresponds with JCPDS card no. 96-152-7227 for orthorhombic SnS. This is oriented to the (110), (131), (002) and (112) directions for SnS. The XRD peaks for PbS agree with the (200), (210), (310), (311), and (314) orientations with JCPDS card no. 96-901-4779 for orthorhombic PbS. These agree with the literature report in ref.

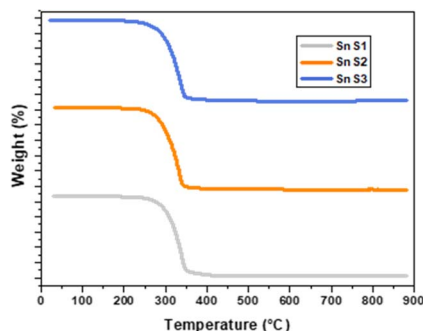


Fig. 3 TGA analysis of SnS1, SnS2 and SnS3 nanoparticles.

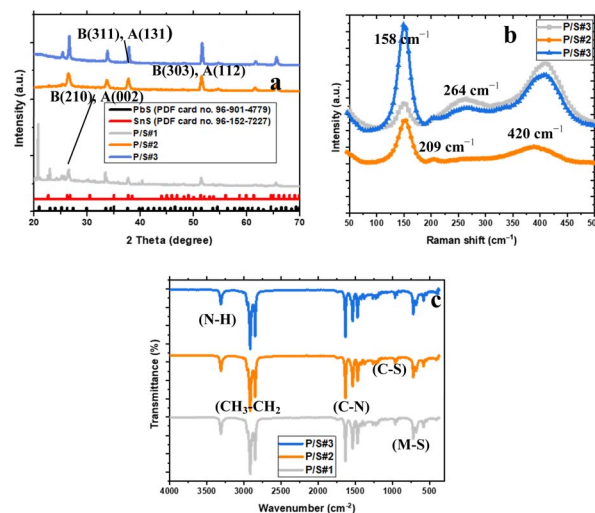


Fig. 4 XRD patterns (a), Raman spectra (b) and FTIR spectra (c) of P/S#1, P/S#2 and P/S#3 modified devices.

37 and 38. To further assess the crystalline quality, phase, and dopant information of the FTO electrodes, Raman spectroscopy measurements were employed for the P/S#-1, P/S#-2, and P/S#-3 modified cells, as seen in Fig. 4b. Raman peaks at 264 cm⁻¹ and 420 cm⁻¹ indicate the presence of PbS, which is consistent with the observed vibrational shifts reported previously.³⁹ These peaks are ascribed to the phonon mode of first-order LO and the two-phonon scattering 2LO mode.³⁸ A sharp lattice vibration peak at 158 cm⁻¹ and small vibrations at 209 cm⁻¹ are indexed to the SnS phase of A_g and B_g vibration modes.^{40,41} The presence of a known mode relies on incidence and scattering radiation directions parallel to any of the three axis (*a*, *b*, *c*), scattered photons, and the electrical polarization of the incidence.^{41,42} The functional groups of the modified devices as P/S#1, P/S#2, and P/S#3 FTIR spectra are shown in Fig. 4c.

In the three devices, the ν(C–N) bands were detected at 1470, 1536, and 1633 cm⁻¹, while the band at 964 cm⁻¹ corresponds to ν(C–S). The strong peaks at 2918 and 2851 cm⁻¹ for P/S#1, P/S#2, and P/S#3 are ascribed to the asymmetric CH₃ and CH₂ correlated with the capping agent of hexadecylamine. The ν(N–H) bands observed at 3257 cm⁻¹ are assigned to the hexadecylamine vibrations.⁴³ The ν(M–S) peaks at 730, 578, and 413 cm⁻¹ were assigned to the Sn–S and Pb–S of the metal and sulfur ions. The FTIR studies indicate the successful loading of SnS and PbS on TiO₂, which is confirmed by the presence of titanium, oxygen, and carbon in the EDS spectra.^{4,20,44}

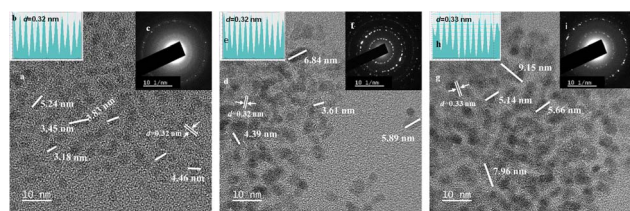
HRTEM images of P/S#1, P/S#2, and P/S#3

The interfacial activities of the TiO₂/PbS/SnS/HI-30/Pt-modified electrodes were studied by high-resolution transmission electron microscopy. The devices were examined by HRTEM, as seen in Fig. 5a–l. The HRTEM images presented in Fig. 5a, d and g display hexagonal shapes for P/S#1, P/S#2, and P/S#3 devices, which consist of the deposition of PbS and SnS on the TiO₂. The HRTEM images of P/S#1, P/S#2, and P/S#3 devices indicate a high crystallinity heterostructure of TiO₂/PbS/SnS. The lattice



Table 1 XRD parameters for P/S#1, P/S#2, and P/S#3 devices with orthorhombic (B) PbS and orthorhombic (A) SnS phases

					Amounts calculated using the reference intensity ratio			
Samples	2θ (deg.)	FWHM (deg.)	Miller indices (hkl)	Crystallite size (nm)	Phase index amount (%)	Element amount (weight%)		
P/S#1	20.76	0.1050	B (200)	80.37	SnS 34.4, PbS 65.6	Pb 56.8	Sn 27.1	S 16.1
	22.68	0.1050	A (110)	80.63				
	25.78	0.2100	B (210), A (002)	40.55				
	33.42	0.1575	B (301)	55.03				
	37.67	0.2625	B (311), A (131)	33.41				
	51.42	0.2100	B (303), A (112)	43.87				
			Average 55.64					
P/S#2	26.15	0.2625	B (210), A (002)	32.45	SnS 44.7, PbS 55.3	47.9	35.2	16.9
	37.30	0.2625	B (311), A (131)	33.37				
	51.42	0.2100	B (303), A (112)	43.87				
			Average 36.56					
P/S#3	26.64	0.2625	B (002), A (002)	32.50	SnS 41.6, PbS 58.4	50.6	32.7	16.7
	37.83	0.2625	B (311), A (131)	33.43				
	51.61	0.2100	B (303), A (112)	43.91				
			Average 36.61					

**Fig. 5** HRTEM image (a, d, and g), lattice spacing (b, e, and h) and SAED pattern (c, f, and i) of P/S#1, P/S#2 and P/S#3 modified devices.

spacing and their crystalline plane for the three devices are (303) and 0.32 nm for P/S#1, (002), (301) and 0.33 nm for P/S#2, and (002), (301), (314) and 0.33 nm for P/S#3, to the orthorhombic PbS and SnS (JCPDS data 96-901-4779 and 96-152-7227). The selected area electron diffraction (SAED) pattern corresponds to the PbS and SnS (SAED) pattern in the inserted Fig. 5d, h and i shows the apparent ring diffraction pattern comes from the cubic PbS and hexagonal SnS crystalline nature. The average particle sizes are 4.03 nm, 5.18 nm, and 6.97 nm. The average particle size distribution of SnS QDs was found within 2.22–17.11 nm, while PbS QDs reveal a grain size average of 1.82–5.95 nm.⁴² The size distribution shows that the average particle size is within the quantum dot region.⁴⁵

SEM, AFM and EDS images of P/S#1, P/S#2, and P/S#3

Fig. 6d–l show the SEM, AFM and EDS images of the P/S#1 device, displaying a polyhedral morphology.⁴⁶ The microstructure of the TiO₂ films changed after the deposition of the binary structure of PbS/SnS, as seen in Fig. 6b, for the P/S#2 device. The porosity of the P/S#3 cell decreased with the addition of PbS and SnS. SnS covered the surface of the PbS cell and filled in the pores generated by TiO₂ particle packing.⁴⁷ The elemental composition of P/S#1, P/S#2, and P/S#3 modified devices is

studied by the EDS images illustrated in Fig. 6m–o. The uniform distribution of elements in P/S#1 and P/S#2 devices was confirmed in comparison to the P/S#3 device. Atomic force microscopic (AFM) images of the multilayer devices are displayed in Fig. 6a–c. A detailed analysis of the surface profile of the three modified cells in terms of root mean square (RMS) and average roughness (R_a) was obtained. The smooth topography has RMS roughness values of 31.2 nm, 29.6 nm, and 29.9 nm, while the R_a values are around 23.3–24.5 nm for P/S#1, P/S#2, and P/S#3, respectively. This is an indicator of the role of both PbS and SnS quantum dot size effects of the modified cells, which account for their high current density and electron transfer speed, supported by the EIS and the I - V results.

UV-vis and Tauc plot spectra of P/S#1, P/S#2, and P/S#3

The optical properties of P/S#1, P/S#2, and P/S#3 are presented in Fig. 7a by the UV-vis absorption spectrum. P/S#3 reveals a high intensity with an absorption edge over the entire range. This can be linked to the relatively narrow band gap of PbS and SnS. It is worth noting that P/S#2 has a similar absorption edge with less intensity compared to the P/S#1 cell device. Moreover, both P/S#2 and P/S#3 cell devices display absorption covering the entire visible region. The relation of $(\alpha h\nu)^2$ vs. $h\nu$ could be obtained by calculating the UV-vis data, as shown in Fig. 6b. The empirical Tauc formula is as follows:

$$(\alpha h\nu) = A'(E - E_g)^n \quad (6)$$

where $h\nu$ stands for the energy of the incident photon and the absorption coefficient, where A and n are constants. The E_g value of the three samples could be valued from the $(\alpha h\nu)^2$ vs. $h\nu$ curves by the extrapolating method. E_g of P/S#3 at 2.75 eV and P/S#3 has the values at 2.0 eV and 1.0 eV as first and last order, while that of P/S#1 at 2.5 eV, confirming the heterostructure combination of TiO₂/PbS/SnS, which are closely related to the

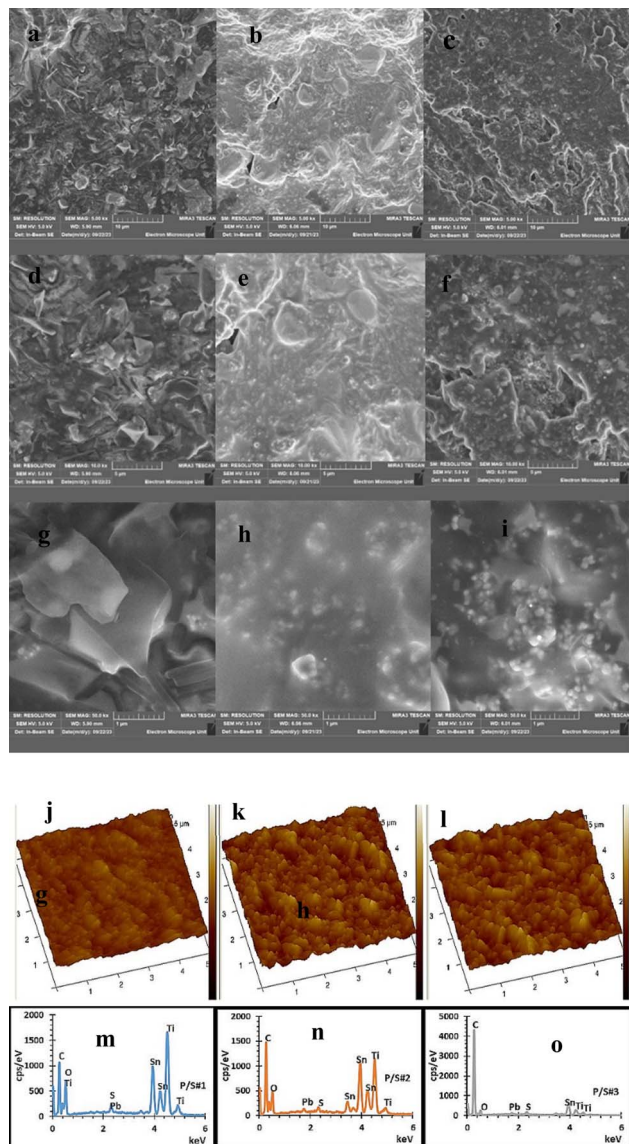


Fig. 6 AFM images (a–c), SEM images (d–l) and EDS spectra (m–o) of P/S#1, P/S#2 and P/S#3 modified devices.

previous studies by ref. 48–50. The spectral response falls in the visible regions for the blue shift, which agrees with the HRTEM quantum dot size distribution; this area is beneficial for the enhancement of the photochemical response under visible light.^{51,52} Moreover, these results are supported by the $J-V$ results of this study. To further assert the optical properties of the produced cells, PL analysis was performed at an excitation wavelength of 350 nm, as shown in Fig. 7c and d. This is an important technique for understanding the recombination of electron-hole photogenerated in P/S#1, P/S#2, and P/S#3 devices. Emissions are detected in the blue region at 410–460 nm for three samples confirming their quantum-sized effects. These emissions correspond to defects and emission bands; radiation decays from the conduction band to the valence band. A decrease in P/S#2 and P/S#3 intensity was associated with a slower recombination rate and increased

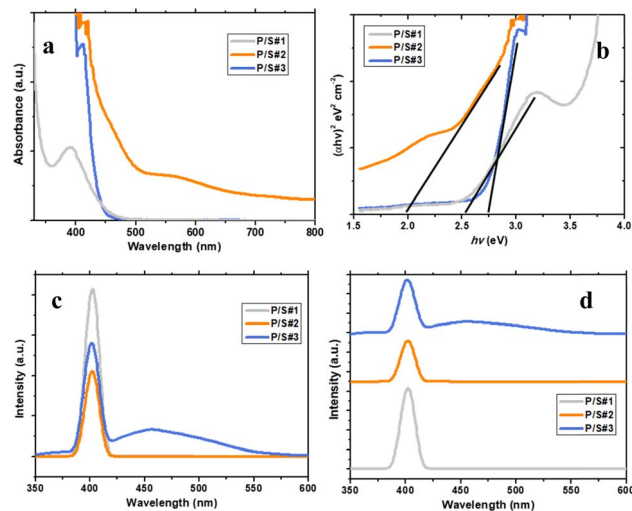


Fig. 7 UV-vis absorption spectra (a), Tauc plot (b), PL spectra (c) and PL spectra superimposed (d) of P/S#1, P/S#2 and P/S#3 modified devices.

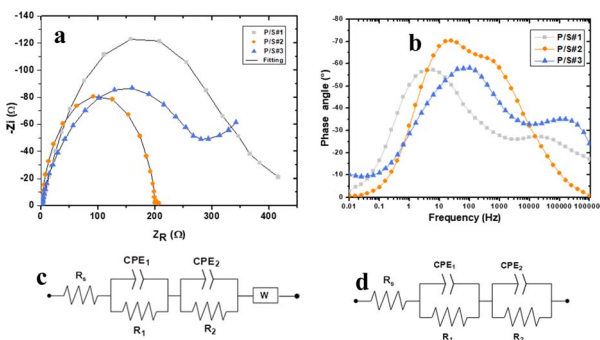


Fig. 8 EIS Nyquist plot (a), Bode plot (b) and equivalent model circuit (c) for P/S#1, P/S#3 and (d) P/S#2 modified devices.

charge separation, which correlate with the EIS observation. This occurrence also suggests that there is multiple exciton generation from the absorber (PbS and SnS), which may generate more than one electron-hole pair for each absorbed photon with energy at least twice that of the absorber band gap. The high intensity observed in P/S#1 indicates that electrons are not excited more efficiently from the VB to the CB. This inference implies that the cells have lower photocatalytic activity and charge separation compared to P/S#2 and P/S#3 which is supported by the $I-V$ results in this study.

EIS and Bode plot analysis of P/S#1, P/S#2, and P/S#3

To understand the electron activity at the interfaces of P/S#1, P/S#2, and P/S#3, the EIS analysis was performed, as seen in Fig. 8a. For all devices (compositions), the diameter of the semicircular impedance spectrum (high-frequency time-constant region) is considerably comparative to the electron-transfer resistance at the electrode-electrolyte interface, revealing the kinetics of charge transfer from the redox reaction at the electrode-electrolyte interface. The Nyquist plots for the 3



Table 2 EIS fitted parameters for P/S#1, P/S#2 and P/S#3 devices^a

Parameters	P/S#1	P/S#2	P/S#3
R_s (Ω)	2.745	1.667	1.501
R_1 (Ω)	56.640	16.308	12.317
R_2 (Ω)	375.489	200.514	310.659
σ ($\Omega s^{1/2}$)	—	—	0.0015

^a LSV and CV curves of P/S#1, P/S#2, and P/S#3.

devices and the equivalent circuit model used in fitting the parameters are shown in Fig. 8a–d and Table 2. The models used in describing the Nyquist and Bode plots are similar to those reported in the literature.^{53,54}

The overall impedance, Z_T , of the Nyquist plots of P/S#1 and P/S#2 in Fig. 8a is given by eqn (7), while that of P/S#3 is given by eqn (8).

$$Z_T = R_s + R_1 \| Z_{CPE1} + R_2 \| Z_{CPE2} \quad (7)$$

$$Z_T = R_s + R_1 \| Z_{CPE1} + R_2 \| Z_{CPE2} + Z_w \quad (8)$$

where R_s is the series resistance of the cell and is evaluated from the real impedance axis shift of the Nyquist plot from the origin, Z_{CPE1} and Z_{CPE2} describe the impedance to mobile charge carriers from the different bulk transport layers of the specimen. From eqn (8) Z_w , represents the slow diffusion process observed on P/S#3 as an incomplete arc starting at an approximate angle of 45° .⁵⁵ This region of the Nyquist plot is commonly referred to as Warburg region. The Warburg element's impedance was calculated as follows:

$$Z_w = \frac{\sigma}{\sqrt{2\pi f}} - j \frac{\sigma}{\sqrt{2\pi f}} \quad (9)$$

where σ is the Warburg coefficient defined as follows:

$$\sigma = \frac{RT}{n^2 F^2 \sqrt{2}} \left(\frac{1}{C_O(D_O)^{1/2}} + \frac{1}{C_R(D_R)^{1/2}} \right) \quad (10)$$

where D represents the diffusion coefficient, and R , T and F are the ideal gas constant, absolute temperature, and Faraday constant respectively. Subscripts “o” and “r” represent the oxidation and reduction species. “C” and “D” are the concentration and diffusivities of the species, while “n” is the number of electrons transferred.

The EIS response of the P/S#1 and P/S#3 devices, as shown in Fig. 8a, supplies information about the bi-metal layer electrode and the carrier selective layer in terms of their charge transport.⁵⁵ The modified hybrid device of P/S#2 reveals an arc with a smaller radius, illustrating a higher transport efficiency, which is in agreement with the disparity of shunt resistance (R_{sh}) and the series resistance (R_s) lower value as an indicator of better transport performance. P/S#2 devices show a heterojunction hybrid structure of PbS and SnS QD distributed layers of TiO₂/PbS/SnS/HI-30, which is favorable for carrier transport.⁵⁶ This implies that bimetal with quantum dots size is additive with the hybrid anode, which reduces non-radiative recombination more efficiently. Therefore, it improves current density

with a better effect on electron transport in the conduction band of the hybrid anode. A Bode phase plot is a characteristic tool with the ability to indicate the predominant electrical behavior of a device at a given frequency range. It provides easier visualization of the different electron diffusion processes (time-constants). Fig. 8b shows the Bode phase plots of the three devices. The Bode plot of P/S#1 first knee observed and P/S#3 more evident are linked to the capacitive–resistive transition.⁵⁷ The plots show lower capacitive characteristics than that of P/S#2 caused by a more resistive PbS/electrolyte interface as well as the high porosity of the SnS. This suggests that the co-absorbent of PbS and SnS improves charge transfer while inhibiting charge recombination. The increase in J_{SC} , which can be observed in Table 4 J – V parameter, supports the P/S#2 and P/S#3 devices. This is further confirmed by the AFM and LSV results. The Bode plot for the P/S#2 device indicates a highly capacitive device acting as an efficient cooperative bi-metal material.⁵⁷

Fig. 9a and b show the Linear Sweep Voltammetry (LSV) curves of P/S#1, P/S#2, and P/S#3 modified devices at a scan rate of 10 mV s^{-1} . The PbS-coupled SnS binary heterostructure current density for P/S#3 devices is higher than that of P/S#1 and P/S#2 devices within the measured potential range. The P/S#1 and P/S#2 low current density could be linked to a large band gap, resulting in poor response in charge transfer ability, which can be correlated with the Tauc plot result in Fig. 7b. The improved P/S#3 hybrid structure agrees with the band gap narrowing ability of chalcogenides SnS and PbS, as seen in the Tauc plot. This combination benefits from being able to reduce the band gap, leading to the entire absorption of the solar cell spectrum and synergistic interaction between both metal sulphides, thus creating better paths or conduits for effective interfacial charge transfer. This also implies that the quantum dot size effect plays a vital role in the formation of heterostructure metal sulfides and the higher current density obtained by P/S#2 and P/S#3 modified devices in comparison to P/S#1. Moreover, a higher surface area promotes more active sites, increases the transport rate of photoinduced charge carriers, and notably increases the efficiency of the catalyst, which accounts for the observed enhancement by P/S#2 and P/S#3 devices. These LSV results elucidated the influence of lower band gaps on the P/S#2 and P/S#3 modified devices, which is favorable for efficient and better mobility of charge carriers and photogenerated charge carrier separation for improved

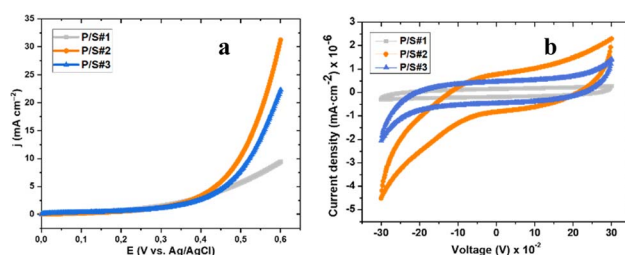


Fig. 9 LSV curves (a) and CV curve (b) of P/S#1, P/S#2 and P/S#3 modified devices.



conversion output.⁵⁸ The electrochemical properties were taken using the potential window and scan rate as stated above for the LSV to analyse the cyclic voltammetry (CV) curves of P/S#1, P/S#2, and P/S#3 modified devices. Fig. 9b confirms the catalytic behaviors of the three cells. P/S#3 and P/S#2 devices show a better current density response as compared to the P/S#1 device, implying a well-developed space charge region at the surface of the binary hybrid structure. These give rise to electron-hole recombination suppression and a high charge mobility rate. Moreover, the easy flow of photogenerated electrons through SnS and PbS quantum dot superior synergy and the space charge region produced channels for the electronic transfer of electrons between them, which led to improved conversion efficiency. Therefore, the electrocatalytic and photocatalytic performance of the modified devices is linked to the high active surface area and quantum size of the binary heterostructure.⁵⁸

CA and *I*-*V* analysis of P/S#1, P/S#2, and P/S#3

Fig. 10a and Table 3 display the chronoamperograms of P/S#1, P/S#2, and P/S#3-modified devices. The three samples show the gradual increase in current density with time until reaching a maximum (I_{\max} , t_{\max}), represented using the Cottrell eqn (11).

$$i(t) = \frac{nFA D_0^{1/2} C_0^*}{\pi^{1/2} t^{1/2}} \quad (11)$$

where $i(t)$ is the current, n is the number of electrons transferred, F is Faraday's constant, A is the electrode surface area, $D_0^{1/2}$ is the diffusion coefficient, C_0^* is the initial concentration, and t is the time. The rise in the current portion relates to a new phase of development through the increasing number of nuclei, while the final is caused by the flow of diffusion ions to the surface of the working electrode to stabilize the change in potential. The growing nuclei of the binary hybrid structures form successive overlapping and diffusion regions around themselves under mass transfer, leading to a planar surface

Table 3 Maximum parameters (I_{\max} , t_{\max}) for P/S#1, P/S#2 and P/S#3 devices

Parameters	P/S#1	P/S#2	P/S#3
I_{\max} (mA cm ⁻²)	2.7	18	0.9
t_{\max} (s)	311	411	0.2

corresponding to t_{\max} and the later decay in current density as a function of time. Instant decay is observed in P/S#3 modified devices, while direct stabilization is identified in the current density values of P/S#1 modified devices, corresponding to the binary hybrid structures limiting the diffusion current process.⁵⁹ Moreover, the I_{\max} for the P/S#2 cells is longer than that of the P/S#1 device, which could be justified by the size distribution of quantum dot variation and the growth kinetics of the binary hybrid structures.⁶⁰ An additional procedure to identify the impact of binary structure on the photovoltaic properties of P/S#1, P/S#2, and P/S#3 modified devices by taking their *J*-*V* measurement is seen in Fig. 10b. The solar cell parameters J_{SC} , V_{OC} , FF, and η are shown in Fig. 10b and Table 4. The increase in J_{SC} observed in P/S#2 and P/S#3 devices is due to their small quantum dot-sized and better crystallinity, leading to a reduction in series resistance in the modified cells. The decrease in V_{OC} observed in P/S#1 is caused by E_g reduction in SnS and shunt resistance in the modified device. According to this result, the electrocatalytic ability of Iodolyte HI-30 polysulfide electrolyte ions and the related device performances might be directly influenced by the elemental proportion of bimetal on the TiO₂/PbS/SnS electrodes. This is largely anchored on the molecular precursor variation of PbS and SnS paving the way for the effective optimization of optoelectronic properties. The high J_{SC} and the P/S#2 and P/S#3 devices' electrochemical performance correlate with the bimetal phase and elemental ratio, which is consistent with other studies⁶¹⁻⁶⁴ with improved electrocatalytic characteristics. Better electrocatalytic activity and favorable surface morphology interactions with the Iodolyte HI-30 polysulfide electrolyte ions resulted in greater J_{SC} for P/S#2 and P/S#3.⁶⁵⁻⁶⁷ The high efficiency of 9.99–8.95% reported in the present study is significantly higher than that reported in previous research studies that used co-sensitization, including Li *et al.* (8.72%),⁶⁸ Xu *et al.* (8.28%),⁶⁹ and Yang and Zhong (7.24%),⁷⁰ as seen in Table 5.

Eqn (12) compares the incident photon-to-current conversion efficiency (IPCE) measurement values for the best-performing devices that were collected between 350 and 800 nm.⁶⁴

$$\text{IPCE} = \frac{1240 \times J_{\text{SC}}}{\lambda \times I} \times 100\% \quad (12)$$

where I stands for the incident light power at λ , and J_{SC} is the measured short circuit current in the device. The P/S#2 device's IPCE value, which was 57% at 410 nm, was higher than that of the P/S#3 device, which was 54% at 410 nm, and the P/S#1 device, which produced an IPCE value of 34% at 410 nm. Because of better light absorption and less recombination, the

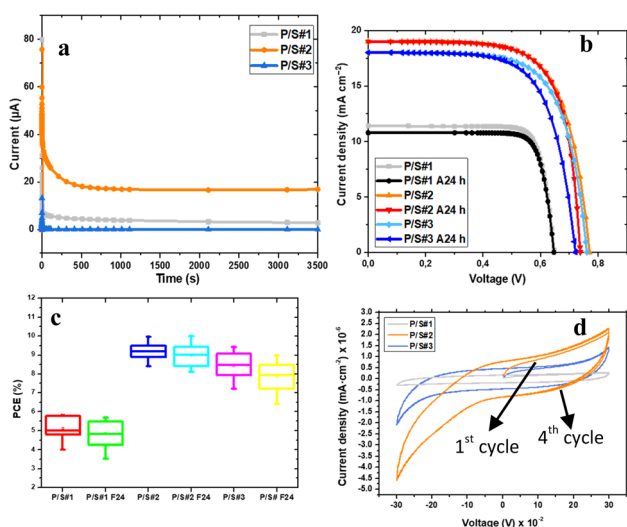


Fig. 10 Chronoamperograms (a), *I*-*V* curve (b), PCE (c) and multi-cycle CV test (d) of P/S#1, P/S#2 and P/S#3 modified devices.



Table 4 I – V parameters for P/S#1, P/S#2 and P/S#3 devices

Samples	V_{OC} (V)	J_{SC} (mA cm ⁻²)	FF	η (%)	P_{max} (W)
P/S#1	0.64 ± 00	11.4 ± 01	0.81 ± 01	5.90 ± 01	5.97
P/S#1 F 24 h	0.64 ± 01	10.8 ± 01	0.82 ± 02	5.67 ± 02	5.67
P/S#2	0.77 ± 00	19 ± 02	0.68 ± 01	9.95 ± 01	10.05
P/S#2 F 24 h	0.73 ± 02	19 ± 02	0.72 ± 03	9.99 ± 01	10.03
P/S#3	0.76 ± 01	18 ± 01	0.69 ± 00	9.44 ± 02	9.45
P/S#3 F 24 h	0.72 ± 03	18 ± 02	0.69 ± 01	8.95 ± 01	8.86

Table 5 Previously reported I – V parameters in the literature on co-sensitization^a

Samples	V_{OC} (V)	J_{SC} (mA cm ⁻²)	FF	η (%)	Ref.
ZnO/NR@/ZnO/IO/CdS/CdSe	0.66	38.60	0.44	11.24	71
TiO ₂ /ZCISe	0.59	26.43	0.56	8.72	68
TiO ₂ /CdS/CdSe	0.62	32.29	0.40	8.28	69
TiO ₂ /CdS/CdSe NS	0.58	15.07	0.51	4.42	72
TiO ₂ /CuInS ₂	0.68	11.33	0.41	3.13	73
TiO ₂ /CdS/CdSe	0.572	15.8	0.59	5.28	74
TiO ₂ /CdTe/CdSe _x S _{1-x}	0.62	16.58	0.69	7.24	70
TiO ₂ /CdS/CdSe/ZnSe	0.58	20.11	0.55	6.39	75
TiO ₂ /CdS/CdSe	0.66	14.39	0.60	5.70	76
TiO ₂ /CdTe/CdS/CdS	0.61	20.19	0.51	6.32	77
TiO ₂ /CdS/CdSe	0.60	14.98	0.54	4.85	78
TiO ₂ /CdSe/CdS	0.42	13.46	0.41	2.31	79
TiO ₂ /CdS/CdSe	0.59	14.23	0.55	4.62	80
P/S#1	0.64 ± 00	11.4 ± 01	0.81 ± 01	5.90 ± 01	TW
P/S#1 F 24 h	0.64 ± 01	10.8 ± 01	0.82 ± 02	5.67 ± 02	TW
P/S#2	0.77 ± 00	19 ± 02	0.68 ± 01	9.95 ± 01	TW
P/S#2 F 24 h	0.73 ± 02	19 ± 02	0.72 ± 03	9.99 ± 01	TW
P/S#3	0.76 ± 01	18 ± 01	0.69 ± 00	9.44 ± 02	TW
P/S#3 F 24 h	0.72 ± 03	18 ± 02	0.69 ± 01	8.95 ± 01	TW

^a TW = this work.

increased IPCE value for P/S#2 shows an effective contribution of absorbed photons to the photocurrent density.

A stability assessment of the three prepared modified electrodes was done to assess their stability. After 24 hours, the three fabricated devices' PCE η trends in each cell were studied following the procedure reported in the literature,⁶² as seen in Fig. 10c. According to these results, P/S#2 provides QDSSC devices with better stability than P/S#1 and P/S#3 after 24 hours. The PCE η of the P/S#1 4.01% (5.90%) device showed drops that after 24 hours reached 3.51% (5.67%) of their initial performance, whereas the P/S#3 7.19% (9.44%) device showed consistent drops that resulted in 6.40% (8.95%) decrease. On the other hand, the P/S#2 8.40% (9.95%) device showed superior stability with a reduction of 8.10% (9.99%) from its initial performance, and the J_{SC} values remained unchanged over 24 hours. The strong corrosion-resistant property of the bimetallic sulfide, which may limit the corrosion interaction between the HI-30 polysulfide electrolytes and Pt counter electrode, may be directly responsible for the P/S#2 device's excellent stability improvement.^{61,62} This further confirms that measuring the chronoamperogram current as a function of time leads to a direct stabilization of J_{SC} from the first seconds for P/S#1, and a burn-in reduction during the first 400 s for P/S#2 before the

current stabilizes. The stability performance of the electrode materials in the polysulfide redox couple is often assessed using the multi-cycle CV test.⁶¹ The surface alterations of the P/S#1, P/S#2, and P/S#3 electrodes following the multi-CV treatment are shown in Fig. 10d. Four cycles of the multi-cycle CV curves were performed at a scanning rate of 10 mV s⁻¹. Following the first cycle and the final cycling treatment, there is no discernible change between the three samples, suggesting that they demonstrated exceptional stability in the polysulfide electrolyte system.

Conclusions

In this work, a nano-size hybrid structure has been successfully synthesized and loaded using a two-step deposition technique. The photoelectrochemical performance was evaluated using various characterization techniques. P/S#2 and P/S#3 modified cells exhibit improved photoelectrochemical performance and photocatalytic activity compared to the P/S#1 device. Moreover, the P/S#2 cell possesses the highest J_{SC} value and the best stability. The excellent properties are attributed to the easy flow of photogenerated electrons through SnS and PbS quantum dot synergy, and the space charge region produces channels for the

electronic transfer of electrons between them. More importantly, this study delivers a high-conversion nano-size hetero-structure for QDSSCs.

Data availability

The data supporting the findings of this study are available from the corresponding author upon reasonable request.

Author contributions

Conceptualization, methodology, validation, formal analysis, investigation, resources, data curation, writing (original draft preparation), writing (review and editing), monitoring, and funding acquisition. All authors have read and agreed to the manuscript.

Conflicts of interest

The authors declare that the research was conducted in the absence of any commercial or financial relationships that could be construed as a potential conflict of interest.

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