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2024, 12, 7562Anti-solvent engineering enables efficient
ambient-processed halide perovskite solar cells†Ivy M. Asuo, ^{*ab} Arezo Mahdavi Varposhti, ^c Enrique D. Gomez ^{ad} and
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Organic–inorganic metal halide perovskite thin film formation is one of the major challenges for solution-processed perovskite optoelectronic and photovoltaic devices, in particular, in ambient conditions due to their hygroscopic nature. Therefore, to improve the quality, optical properties, and performance of ambient processed metal halide perovskite thin films, it is necessary to have control over the fabrication process. We demonstrate how compositional and morphological control of ambient-processed perovskite films can be achieved using anti-solvents. We compare how exposing the film to anti-solvents, including dichlorobenzene, ethanol, and chlorobenzene, affects the crystallization and device performance of ambient-processed planar heterojunction perovskite solar cells. We also analyzed the charge carrier dynamics of the devices and found that including dichlorobenzene leads to reduced charge carrier recombination. The incorporation of dichlorobenzene results in densely packed grains without voids, leading to the best-performing device with a reproducible power conversion efficiency of ~20%. These findings open the possibility of developing low-cost, highly reliable perovskite solar cells for commercial applications in the future.

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Introduction

Metal halide perovskites have recently gained traction in research due to their unique properties. These include optoelectronic properties such as high absorption coefficients, adjustable band-gap, long charge carrier diffusion lengths, low exciton binding energy, and a cost-effective manufacturing process.^{1–3} Their distinct characteristics render them exceptionally appropriate for various applications, spanning lighting,^{4,5} lasers,⁶ sensing and imaging,^{7,8} and photovoltaics⁹ (PVs). Within the field of PVs, a notable increase in the efficiency of perovskite solar cells (PSCs) has been recorded in recent years. In particular, the power conversion efficiency (PCE) of PSCs has experienced a substantial rise, from a modest 3.8% to a remarkable certificated PCE value of 26.1%.^{10–12}

The microstructural morphology and uniformity of the perovskite thin film have been identified as critical parameters in boosting the performance of PSCs.¹³ Numerous researchers are pioneering diverse techniques to fabricate perovskite thin films of superior quality through additive incorporation,¹⁴ vacuum flash annealing,^{15,16} gas blowing,¹⁷ and anti-solvent treatment.¹⁸ The anti-solvent treatment, simple and compatible with various halide perovskites, has become a prevalent method in producing high-efficiency PSCs¹³ and substantially propelled further development. Seok *et al.*¹⁹ first demonstrated anti-solvent treatment in 2014, employing toluene in perovskite thin film solar cells with a notably enhanced PCE of 16.2% and minimal hysteresis. This breakthrough has since played a pivotal role in advancing PSC technology. Anti-solvents are miscible with other solvents, do not dissolve perovskite materials, and thus, are strategically used to induce supersaturation. This is achieved by diminishing the solubility of the perovskite layer. This effectively curtails the rapid crystallization and subsequent formation of perovskite thin films, a phenomenon attributable to the inherently low solubility of the perovskite precursor in solvents.²⁰

Implementing anti-solvent engineering has been demonstrated as a straightforward and efficient strategy for modulating nucleation and crystal growth processes, thereby enhancing the fabrication of efficient and stable high-quality perovskite solar cells.^{18,21} Various anti-solvents, such as toluene, chloroform, ethanol (EtOH), chlorobenzene (CB), diethyl ether, and mixed solvent–antisolvents, have been studied to understand

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their impact on the crystallinity and morphology of perovskite thin films.^{22,23} Gedamu *et al.*²² devised a new method using mixed solvents under ambient conditions to produce well-oriented perovskite crystallites with controlled sizes and minimal pinhole densities, critical for efficient PSCs reaching up to 14.0%. Lee *et al.*²⁴ examined the influence of various anti-solvents with different dipole moments and dielectric constants on the crystallinity and surface coverage of the thin films, resulting in 11.25 cm² perovskite sub-modules exhibiting a negligible current density–voltage (*J*–*V*) hysteresis and a PCE of above 11.5%. In another study, Prochowicz *et al.*²⁵ demonstrated the effect of dipping perovskite absorber films in antisolvents with different polarities using CB and isopropanol (IPA). The perovskite films treated with IPA exhibit a more uniform morphology with larger grains than CB-treated films. However, the IPA-treated devices exhibit higher hysteresis during *J*–*V* scans. Chen *et al.*²⁶ explored various volatile anti-solvents for controlled PbI₂ film uniformity and porosity to attain a perovskite film free of residual PbI₂, which is crucial for high-efficiency PSCs. Their findings demonstrate that ethanol-treated perovskite films exhibit high density and smooth surfaces, resulting in devices with a PCE of 16.8%. A common trend observed in much of the research conducted on anti-solvent engineering is the enhancement of solar cell performance. This improvement is greatly influenced by several factors associated with the film microstructure, including grain size, grain boundaries, pinhole density, crystalline film quality, and orientation.²²

While most literature reports CF, CB, EtOH, IPA, toluene, *etc.*, as anti-solvents for perovskite thin film treatment in

controlled environments, little is known about dichlorobenzene (DCB). Kim *et al.*²⁷ reported the low boiling point CF- and high boiling point DCB-treated single-cation inverted PSCs made in a controlled environment. They showed a visibly poor morphology with noticeable pinholes for DCB-treated PSCs, leading to a PCE of 8.1%, and nicely compact and dense film microstructure for CF-based PSCs with a PCE of 14.2%. They ascribed the difference in performance to the vaporization of the anti-solvent upon annealing and concluded that DCB is a poor anti-solvent for PSCs. Similarly, Taylor *et al.*²⁸ studied triple cation inverted PSCs fabricated in a controlled environment and treated with a host of anti-solvent treatments, including DCB and the green solvent EtOH, yielding about 18.5% and 17% efficiencies, respectively. The limited number of reports on DCB as an anti-solvent for PSCs have relied on costly controlled environments, such as gloveboxes. Thus, to the best of our knowledge, there is no report on DCB anti-solvent-treated ambient-processed double cation perovskite thin films for PV applications. Meanwhile, there is a strong emphasis on devising novel strategies to alleviate manufacturing costs associated with PSCs for their commercialization. Thus, developing new preparation methods for perovskite films under ambient conditions effectively reduces the economic cost and operational complexity of commercializing PSCs.

In this study, we employ a cost-effective approach to fabricate planar heterojunction (PHJ) PSCs under ambient conditions, where DCB is used as an anti-solvent for enhanced crystallinity and improved quality of the perovskite thin film microstructure. We assess the impact of using DCB as an anti-

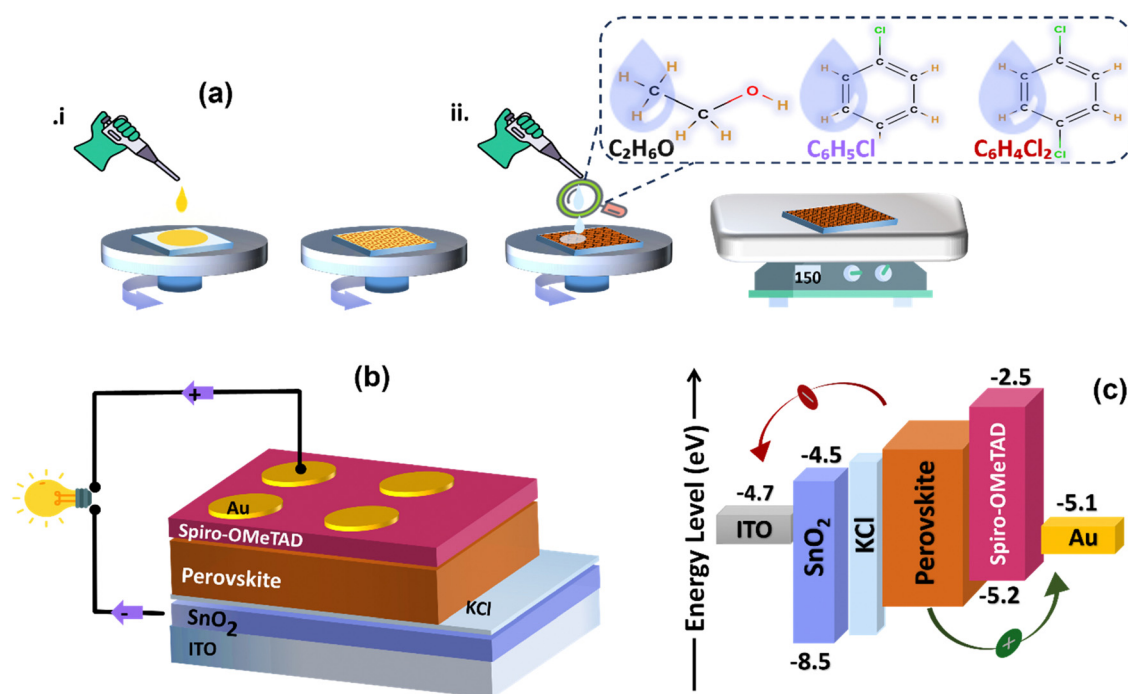


Fig. 1 The schematic representation of (a) fabrication process of solution-processed planar heterojunction perovskite solar cells PHJ PSCs by applying different anti-solvents (EtOH, CB, and DCB). (b) The final device architecture of PHJ PSCs. (c) Energy bands alignment between electrodes, charge transport layers, and halide perovskites facilitate efficient electron injection into SnO₂/KCl and hole injection into spiro-OMeTAD layers.



solvent and compare it to commonly used CB and EtOH anti-solvents on the PV performance of ambient-processed PHJ PSCs, a combination unexplored in the existing literature. We examined and related the performance of the devices to the charge carrier dynamics through device physics. Our results show that the DCB anti-solvent treatment led to better densely packed grains without pinholes than EtOH and CB treatments. The best DCB-treated devices achieved a reproducible PCE of around 20%, almost 2% and 3% higher than EtOH- and CB-treated PSCs, respectively. This approach demonstrates the potential for fabricating solar cells in ambient conditions, with improved stability compared to EtOH-treated PSCs, warranting further investigation in future research to realize the low-cost, highly reliable PSCs for commercialization.

Results and discussions

We fabricate thin films and subsequently planar heterojunction solar cells from perovskite precursor solutions by spin coating in ambient conditions. Fig. 1(a) depicts how the solution-processed layers were spun-cast for this study, and Fig. 1(b) illustrates the schematic diagram of the PSCs configuration. Fig. 1(c) shows the energy band diagram of the stacked thin layers on the glass substrate, depicting their interfacial energy level offsets. Upon illumination and facilitated by its energy band alignment with SnO₂ and spiro-OMeTAD, charge carriers are generated within the perovskite layer, followed by efficient transport of electrons and holes through the ETL and HTL, respectively, and extraction at the back and top electrode.

To attain high-efficiency perovskite solar cells, control of the crystallization and growth of tightly packed perovskite film

microstructures is imperative. This is particularly crucial due to the observed shrinkage and porosity of the perovskite film during one-step CH₃NH₃PbI_{3-x}Cl_x deposition, leading to large pinhole formation ($\sim 5\ \mu\text{m}$).²⁹ In this study, we used a formamidinium (FA)-based perovskite solution in ambient to investigate the crystalline microstructure under three distinct anti-solvent treatments (absolute EtOH, CB, and DCB). The perovskite thin films are obtained through a two-step spin coating process. We initially spin cast the perovskite precursor solution at 4000 rpm, followed by a second spin combined with solvent dripping at 5000 rpm and annealing to complete crystallization.

Fig. 2(a–c) displays the SEM images of three distinct perovskite thin films on glass/ITO under ambient conditions and treated with pure EtOH, CB, and DCB anti-solvents, respectively. The perovskite grain size and film smoothness are not the primary determinants of the overall device performance. Fig. 2(a) exhibits the microstructure of the perovskite film treated with pure EtOH. The measured average grain size is around 0.87 μm , with a smooth surface exhibiting an average roughness of 5.9 nm, as shown in Fig. 2(d and e). Despite the smooth surface and relatively large grain sizes, pinholes are observed between the grains, and a significant gap is evident around the grain boundaries in various areas of the SEM image, depicted by the broken yellow circles. The presence of pinholes in the perovskite films could significantly impede charge extraction, decreasing the short circuit current (J_{sc}) and overall PCE.²⁸ The perovskite thin films treated with CB anti-solvent show a slightly different microstructure than films treated with EtOH, with better surface coverage without pinholes. Fig. 2(b) shows that the average grain size measures approximately 0.48 μm , with a roughness of 9.3 nm. The SEM image also reveals that the perovskite thin film does not show a fully

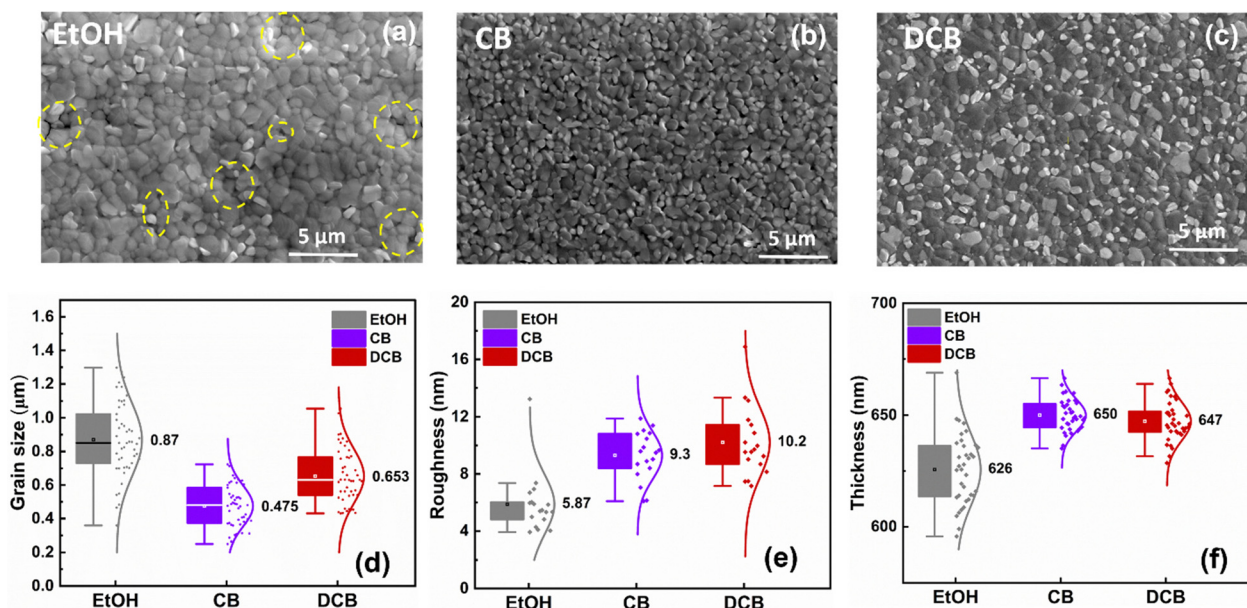


Fig. 2 SEM images of the perovskite layers deposited on an ITO substrate using different solvent treatments: (a) EtOH, (b) CB, and (c) DCB. Frequency distribution plots illustrating the (d) grain sizes, (e) roughness of perovskite thin films, and (f) thickness of perovskite thin films treated with the different anti-solvents.



continuous film morphology compared to the DCB samples. Remarkably, employing pure DCB results in a densely packed perovskite thin film devoid of evident pinholes and gaps at grain boundaries, in contrast to previous observations made for DCB-treated single cation perovskite thin film fabricated in a controlled environment.²⁷ Additionally, the SEM cross-sectional images in Fig. S1 (ESI†) revealed that DCB-treated PSCs exhibit better substrate/perovskite thin film contact formation without pronounced gaps between the substrate/perovskite interface compared to EtOH- and CB-treated PSCs. The average grain size for the perovskite active layer treated with DCB is approximately 0.65 μm , with a roughness of around 10.2 nm. All anti-solvent treated perovskite thin films exhibit an average thickness of around 650 nm, as shown in Fig. 2(f).

Fig. 3(a) displays XRD patterns of thin films treated with EtOH, CB, and DCB anti-solvents and post-annealed at 70 °C for 5 minutes and 150 °C for 10 minutes. All XRD patterns are characterized by peaks positioned at $2\theta \approx 14.1^\circ$ and 28.3° , indicating the formation of (110) and (220) crystal planes of the $\alpha\text{-FAPbI}_3$ phase attributed to desirable perovskite thin films.³⁰ The DCB anti-solvent treated perovskite thin film exhibits more intense diffraction peaks at the known angles attributed to the tetragonal crystalline perovskite, which are $2\theta = 14.08^\circ$ and $2\theta = 28.32^\circ$. These peaks confirm the successful formation of a highly crystalline perovskite structure. The presence of residual PbI_2 peak is apparent in all three films. There are conflicting reports about the role of residual PbI_2 in the performance of PSCs.^{31,32} Some studies show that lower PbI_2 content enables devices to perform better, while others have found that excess

PbI_2 results in the best-performing PSCs.^{32–34} However, the presence of more PbI_2 is known to catalyze the degradation of perovskite films.³⁵

We also measured the coherence length, often assumed to be equivalent to the average crystallite size, of the (110) reflection of the anti-solvent-treated perovskite thin films using Scherrer's eqn (1).

$$D = \frac{K\lambda}{B \cos \theta} \quad (1)$$

Here, K represents the Scherrer constant, which we assume to be 0.9, D is the coherence length, B denotes the full width at half maximum (FWHM) of the measured diffraction peak, θ represents the diffraction angle, and λ stands for the X-ray wavelength (0.154 nm). Fig. 3(b) shows the average coherence length derived from the FWHM of the samples. It is evident from the graph that the CB-treated samples exhibit the largest coherence length compared to the DCB and EtOH-treated samples, while the FWHM is more sharply defined for the EtOH-treated sample. Further optoelectronic characterizations revealed that the densely packed microstructure of the DCB-treated perovskite thin films could influence the charge carrier transport and device performance.

To elucidate the optical properties and charge transport dynamics of PSCs, we conducted absorption, photoluminescence (PL), and time-resolved PL (TRPL) measurements on thin films deposited onto glass substrates. In Fig. 3(c), we show the similar absorption spectra of the three films. The optical bandgap exhibits almost no change (1.55 eV for EtOH and CB

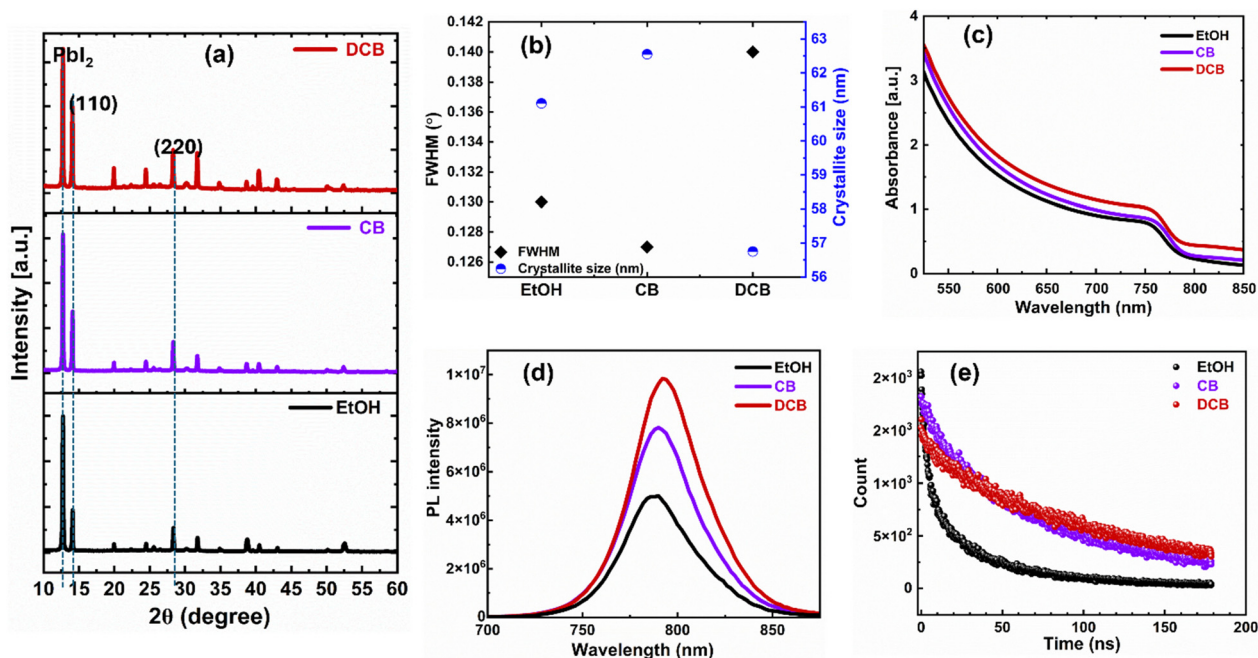


Fig. 3 (a) XRD pattern of the perovskite thin films on ITO substrates corresponding to each anti-solvent treatment. (b) Plot showing the calculated full width at half maximum (FWHM) and crystallite size values at $2\theta = 14.1^\circ$. (c) UV-Vis-NIR absorption spectra of the thin films. (d) Steady-state photoluminescence (PL) spectra of the perovskite thin films deposited on glass substrates. (e) Time-resolved photoluminescence (TRPL) profiles of the perovskite thin films deposited on glass substrates.

and 1.54 eV for DCB). Therefore, the anti-solvent treatment does not affect the absorption of the perovskite thin films. As shown in Fig. 3(d), the intense PL intensity from DCB-treated perovskite films indicates that anti-solvent exposure helps form a higher quality film and effectively suppresses non-radiative charge carrier recombination.³⁶ Fig. 3(e) shows the normalized TRPL for the perovskite coated with various antisolvent treatments. The carrier lifetimes from TRPL profiles, determined *via* bi-exponential fittings, are presented in the ESI,[†] Table S1. In general, the slow decay component (τ_1) is attributed to non-radiative recombination from surface defects and traps at the grain boundaries, and τ_2 is the fast decay component, linked to the quenching of charge carriers at the interface and the bulk.^{29,30} The TRPL analysis reveals an improvement in carrier lifetime τ_2 for the DCB- and CB-treated films of 110 ns and 80 ns, respectively, in contrast to the EtOH-based perovskite films with a carrier lifetime of 35 ns. The extended carrier lifetime is due to higher crystallinity and suppression of defects in the DCB-treated perovskite films.³⁶ Thus, the microstructural crystallinity and morphology, PL, and TRPL characteristics of the DCB-treated ambient perovskite films point to better performance over CB-treated, especially EtOH-treated devices.

We then fabricated planar PSCs with the following structure: glass/ITO/SnO₂/KCl/perovskite/spiro-OMeTAD/Au. A comprehensive description of the procedure, from perovskite solution preparation to device fabrication steps, is provided in the Experimental section. We measured the *J*-*V* characteristics test under AM 1.5G irradiation to determine the performance of the devices. From the findings shown in Fig. 4(a), DCB-treated PSCs are generally the most efficient of the three types of devices at 20.1% for highest performing solar cells, followed by CB-treated devices at 18.2% and then EtOH-treated devices at 17.2% with device area of 0.092 cm². Though all three devices yielded similar short circuit current (*J*_{sc}) of 24.5 mA cm⁻², 24.4 mA cm⁻², and 24.9 mA cm⁻², respectively, for EtOH-, CB-, and DCB-treated PSCs, DCB-based PSC exhibit the highest open circuit voltage (*V*_{oc}) of 1.131 V (compared to 1.049 V for CB-based PSC and 1.062 V for EtOH-based PSC) and the highest fill factor (FF) at 71.1% compared to 70.9% and 66.2% respectively for CB- and EtOH-based devices. Table 1 shows the best values and similar trends for the average values of PCEs obtained from the statistics of the reproducible figures of merit (from PCE, *J*_{sc}, *V*_{oc}, and FF), as depicted in Fig. 4(b–e). We ascribe the better performance of DCB-treated PSCs to the overall effect of DCB anti-solvent in ambient on the perovskite films.

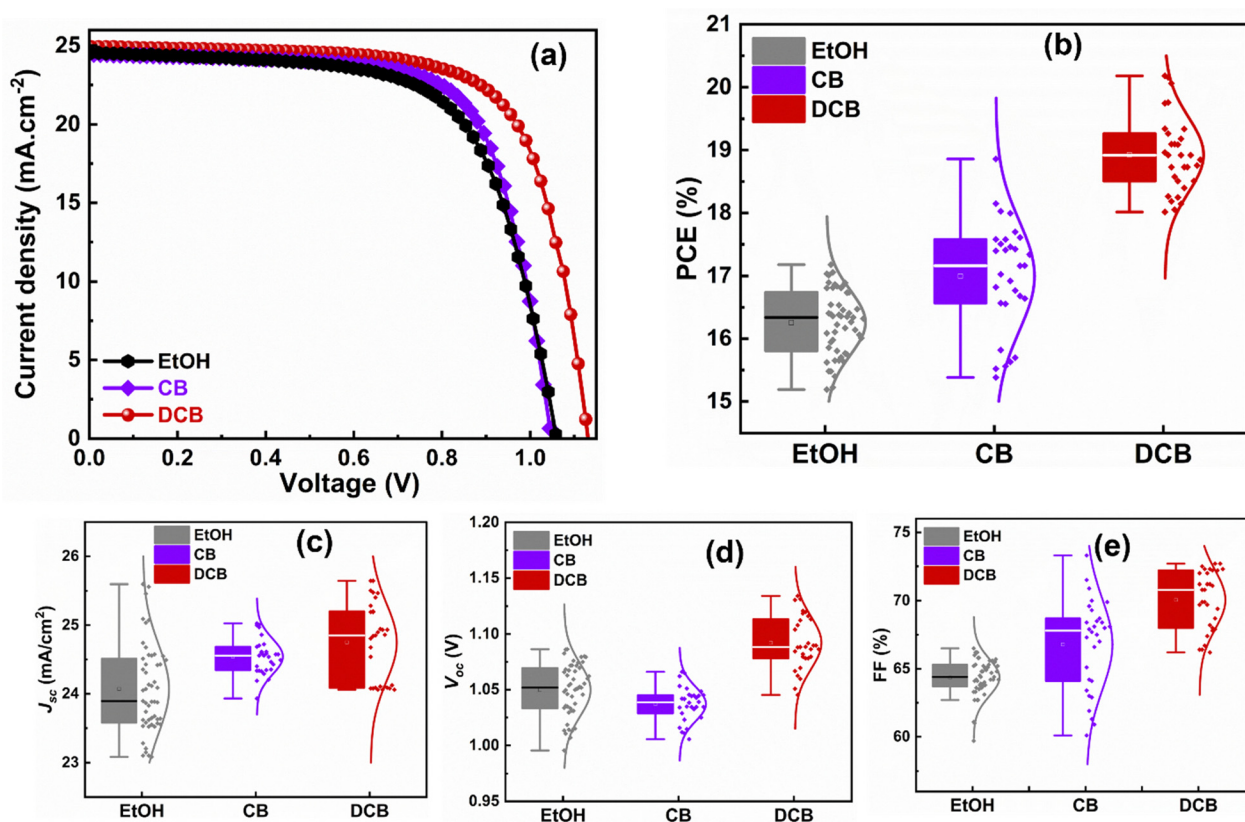


Fig. 4 (a) Current density–voltage (*J*-*V*) characteristics of the champion devices treated using EtOH, CB, and DCB antisolvents with a device area of 0.092 cm². (b) Frequency distribution plot of the planar heterojunction perovskite solar cells (PSCs) treated with corresponding anti-solvents (b) power conversion efficiency (PCE) showing DCB-treated PSCs averaging efficiencies 11% and 16% higher than those of CB- and EtOH-treated PSCs respectively, (c) short-circuit current density, (*J*_{sc}) with marginally similar average currents between DCB- and CB-treated PSCs (d) open-circuit voltage (*V*_{oc}) showing DCB-treated PSCs averaging values of 3–5% higher than those of CB- and EtOH-treated PSCs, (e) fill-factor (FF) showing DCB-treated PSCs averaging values 4% and 11% higher than those of CB- and EtOH-treated PSCs respectively.



Table 1 The PV parameters of the various anti-solvent-treated PSCs are shown in Fig. 4. The number of devices is 51 for EtOH-treated, 30 for CB-treated, and 31 for DCB-treated PSCs

Anti-solvent treatment	J_{sc} (mA cm^{-2})	V_{oc} (V)	FF (%)	Best PCE (%)	Avg. PCE \pm SD (%)
EtOH	24.5	1.062	66.2	17.2	16.3 ± 0.5
CB	24.4	1.049	70.9	18.2	17.0 ± 0.9
DCB	24.9	1.131	71.1	20.1	18.9 ± 0.6

The exposure to DCB leads to densely packed perovskite crystallites in the thin film that are devoid of pinholes and gaps at grain boundaries, reasonable average grain sizes, higher film quality, effective suppression of nonradiative charge carrier recombination, and long charge carrier lifetimes, when compared to CB-treated PSCs that show intermediate performance and EtOH anti-solvent-treated PSCs that exhibit the lowest, yet still substantial, performance. The 17.2% efficiency from green solvent-treated devices is reasonable compared to chlorinated solvent-based PSCs (approximately 1% less than CB and 3% less than DCB), their microstructure and short charge carrier lifetimes indicate the presence of defects and recombination that could impede stability (see early stability study in Fig. S3, ESI†). Indeed, the EtOH-treated PSCs showed more loss in PCE in our stability study under continuous illumination, followed by the DCB-treated PSCs, and CB-treated PSCs as the most stable devices. Also, the presence of more PbI_2 , as shown in Fig. 3a, can catalyze device degradation. A deeper investigation of the long-term stability of DCB-based perovskite solar cells is required to fully assess the role of PbI_2 in these films. A detailed stability

study is beyond the scope of this work and will constitute future work.

To gain more insight into the differences caused by the three antisolvent treatments of the perovskite layers on the devices, we examine charge carrier transport and dynamics in all three systems. We measured and analyzed the transient photocurrent (TPC) and transient photovoltage (TPV) of the PSCs using the Fluxim AG Paios set-up. Fig. 5(a) shows the TPC for the various PSCs. The TPC decay times obtained after fitting are 1.17 μs , 1.12 μs , and 0.98 μs for EtOH, CB, and DCB-treated devices, respectively. The decay time decreased to 0.98 μs for DCB, suggesting an excellent charge carrier extraction process in the device, explaining their better performance compared to CB- and EtOH-based PSCs.³³ In Fig. 5(b), we plotted the TPV data measured under open-circuit conditions. The graph displays two distinct time regimes: the first, associated with trap-assisted recombination, and the second, linked to carrier recombination at the interface.^{37,38} The carrier recombination lifetimes obtained after fitting the TPV data are 46.82 μs for EtOH-treated, 48.62 μs for CB-treated, and 63.09 μs for DCB-treated PSCs, suggesting similar lifetimes for both EtOH and CB-treated devices. These results indicate that nonradiative charge recombination was minimized in the DCB-based solar cells compared to the other two. Particularly, recombination is substantially minimized for DCB-treated devices, which translates to higher V_{oc} and FF, and thus, better performance.^{36,38} Furthermore, we measured dark currents for the three devices to corroborate the reduced charge recombination lifetime observed for the DCB-treated devices. The DCB-based PSCs

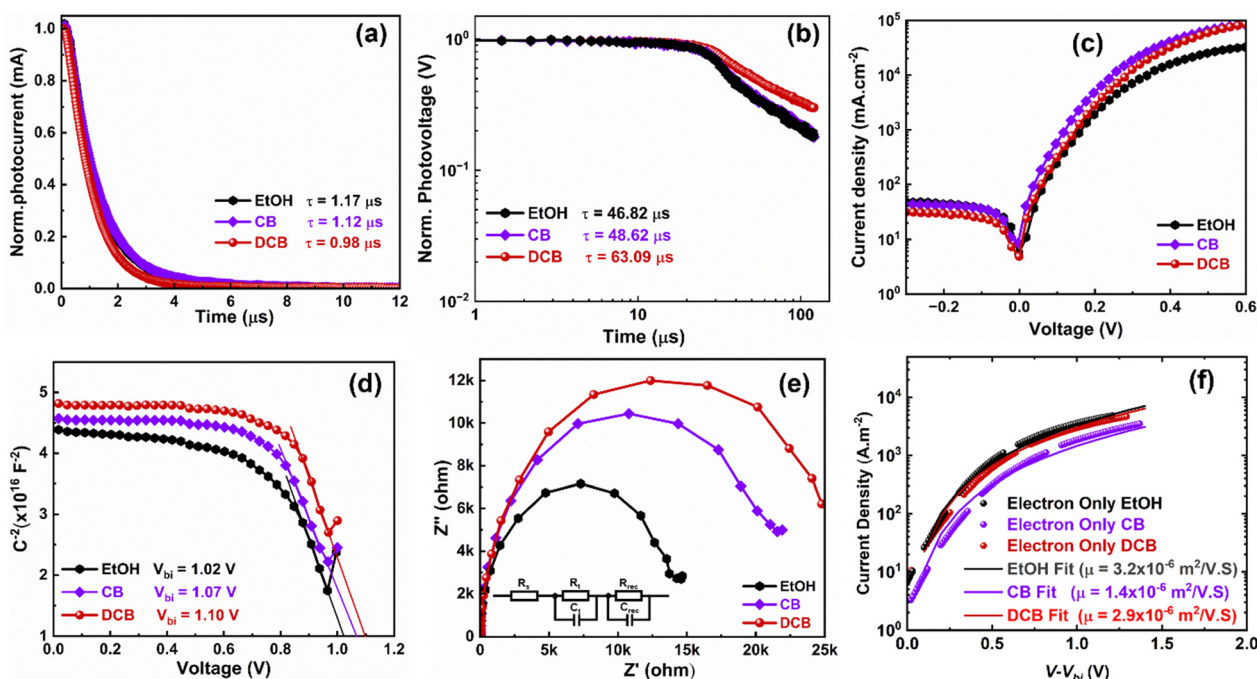


Fig. 5 (a) Transient photovoltaic current (TPC), and (b) transient photovoltage (TPV) plots of the PHJ PSCs treated using EtOH, CB, and DCB anti-solvents. (c) Dark J - V curves of PHJ PSCs using corresponding anti-solvent treatments. (d) Capacitance-voltage plot measured at room temperature at a fixed frequency of 10 kHz for each solvent treatment. (e) Nyquist plots of three different PHJ PSCs with the inset showing the equivalent circuit for fitting the impedance spectroscopy (EIS) data (insert: equivalent circuit for fitting of data). (f) The fitted J - V curve to the modified Mott-Gurney equation.



consistently exhibit slightly lower leakage currents than the other two, as shown in Fig. 5(c). Such lower leakage currents suggest lower background carrier density and better diode behavior of the PSCs.³⁹

Furthermore, a Mott-Schottky analysis is performed to better understand the improved V_{oc} . Fig. 5(d) shows the capacitance-voltage plot measured at room temperature at a fixed frequency of 10 kHz. The built-in potential (V_{bi}) of 1.10 V is obtained for DCB-treated devices compared to 1.02 V and 1.07 V for EtOH and CB-treated devices, respectively. These results again corroborate the enhanced V_{oc} observed for DCB-treated PSCs compared to CB- and EtOH-treated PSCs.^{39,40} A similar deduction on the effectively suppressed charge recombination in the DCB-treated devices can be drawn from the impedance spectroscopy (EIS) measurement depicted in the Nyquist plots for the three devices in Fig. 5(e). To fit the EIS data, an $R-R||C-R||C$ equivalent circuit (shown as an insert in Fig. 5(e)) is used, which includes the series resistance R_s , the charge transfer resistance R_t , and the interfacial charge recombination resistance R_{rec} . The arc at a low-frequency regime is associated with R_{rec} .^{41,42} It has been observed that DCB suppresses the charge recombination, resulting in improved device performance. Likewise, electron-only single carrier devices were made with the following structure: glass/ITO/SnO₂/perovskite/PC₆₁BM/Ag. The fitted JV to the modified Mott-Gurney equation¹⁶ is displayed in Fig. 5(f) for the EtOH-, CB-, and DCB-treated devices. The extracted mobilities for DCB- and EtOH-treated devices are around $3 \times 10^{-6} \text{ m}^2 \text{ V}^{-1} \text{ s}^{-1}$, and that of the CB-treated device is around $1.4 \times 10^{-6} \text{ m}^2 \text{ V}^{-1} \text{ s}^{-1}$. This may partly explain the slightly better J_{sc} of their PSCs compared to CB-treated devices. However, this may not have readily translated into (i) efficient

charge extraction as shown earlier through measurements of charge carrier decay time *via* TPC, (ii) balanced charge carriers, leading to loss in the FF, or (iii) minimized losses through charge carrier recombination as alluded earlier from SEM images and TPV with the lowest charge carrier recombination lifetime for the EtOH-based devices. Therefore, our experimental and device physics analyses are evidence of the improved performance of PSCs treated with DCB anti-solvent.

Additionally, we used high-resolution X-ray photoelectron spectroscopy (XPS) to characterize the chemical states (the location of the maximum valence band in relation to the Fermi level energy E_f) of the all-ambient solution-processed perovskite thin films deposited on ITO/SnO₂ substrates. The XPS survey spectra acquired from the perovskite thin films presented in Fig. 6(a) show peaks at $138.5 \pm 0.2 \text{ eV}$, associated with the Pb 4f_{7/2} core level indicative of Pb-X (X = I, Cl) bonds. Additional core level peaks such as C 1s, N 1s, and I 3d_{5/2} were observed at $288.4 \pm 0.1 \text{ eV}$, $400.8 \pm 0.1 \text{ eV}$, and $619.2 \pm 0.2 \text{ eV}$, respectively, attributed to the FA_{1-x}Cs_xPbI₃ perovskite.^{43,44} Fig. S2 (ESI†) displays the individual carbon species attributed to the C 1s core level, and Fig. 6 depicts the Pb 4f_{7/2} core levels (b, c, d). As specified in Table S2 (ESI†), a reduction of Pb²⁺ to metallic lead (Pb⁰) is observed. This reduction is attributed to a multi-origin degradation mechanism of the perovskite thin films, as discussed by Bechu *et al.*⁴³

To assess the band alignment between the perovskite layer and charge transport layers, we conducted ultraviolet photoelectron spectroscopy (UPS) measurements on the FA_{1-x}Cs_xPbI₃ perovskite thin films on ITO substrates. The work function (ϕ) of the thin films was determined using the secondary electron cut-off (SECO) method, as shown in Fig. 6(e). The

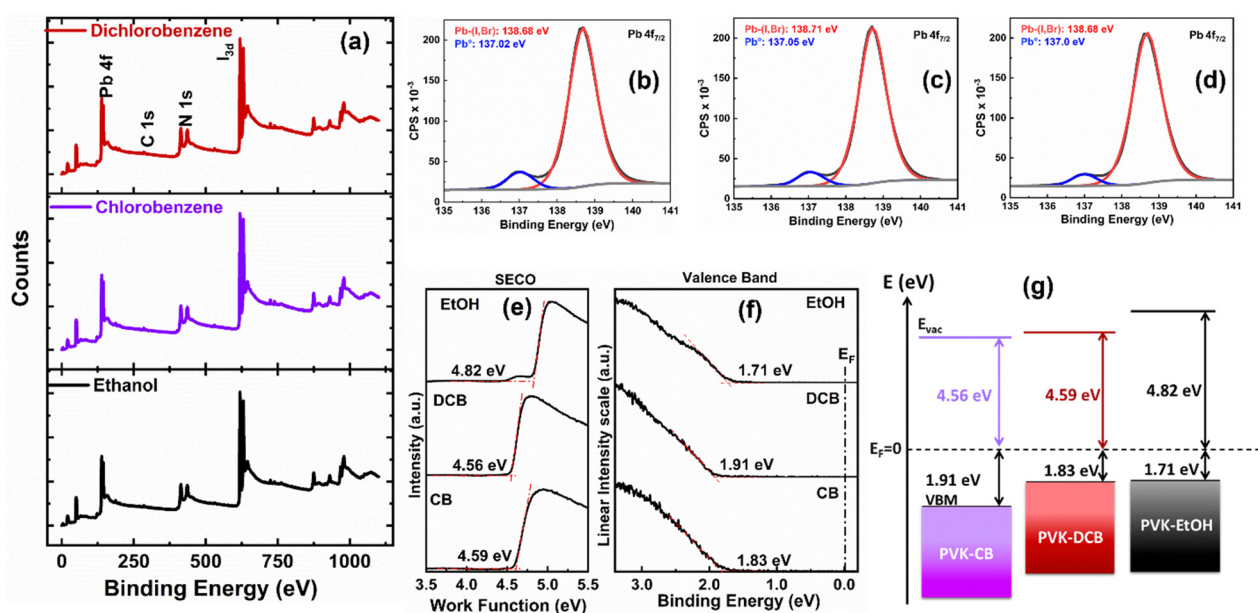


Fig. 6 (a) X-ray photoelectron spectroscopy (XPS) spectra showing Pb 4f, C 1s, N 1s, and I 3d core levels of FAPbI₃ after anti-solvent treatments. XPS spectra of the Pb 4f_{7/2} core level attributed to FAPbI₃ of the (b) EtOH, (c) CB, and (d) DCB antisolvents treated perovskite thin films. UPS spectra of the perovskite thin films with (e) SECO showing the work function value (ϕ) and (f) valence band spectra of the treated-perovskite thin films. (g) The energy-level scheme of the treated-perovskite thin films.



ϕ values are 4.59 eV, 4.56 eV, and 4.82 eV for the DCB-, CB-, and EtOH-treated films, respectively. It is observed that the work function of the EtOH-treated film is higher than that of both the DCB and CB-treated films. The valence band maxima (VBM) were determined based on the spectrum's onset, as shown in Fig. 6(f). For our thin films, the E_{VBM} is measured at 1.83 eV for DCB-treated samples, 1.91 eV, and 1.71 eV for CB and EtOH-treated samples, respectively. Variations in the E_{VBM} values among our samples can be attributed to the distinct functional groups present in the applied anti-solvents on the perovskite thin films. The presence of OH groups in the EtOH anti-solvent likely induces a surface dipole,⁴⁵ thereby elevating the work function in EtOH-treated samples compared to the chlorinated ones. We also presented an energy-level alignment scheme in Fig. 6(g) for the perovskite layers, and the calculated values are shown in Table S3 (ESI†). The DCB-treated sample exhibited the smallest E_{VB} value, facilitating hole extraction. This suggests a lower energy barrier for hole-carrier collection at the interface of the perovskite thin films and spiro-OMeTAD. It is known that lower energy barriers are linked to reduced surface recombination rates, potentially increasing V_{oc} .^{45,46} Indeed, these findings support the charge transport dynamics explained using device physics.

Conclusions

We have demonstrated the fabrication of perovskite thin films under ambient conditions by treating films using an anti-solvent, either one of two chlorinated anti-solvents or a green solvent. We show that dichlorobenzene performs well as an anti-solvent, similarly to chlorobenzene, and improves the crystal quality of halide perovskite thin films and their optoelectronic properties. Additionally, using a green solvent, ethanol, results in a promising device efficiency of 17.2%. Surface defects are reduced with DCB treatment, which leads to an enhanced solar cell performance with a PCE of 20.1%. The higher efficiency of DCB-based solar cells is attributed to suppressed non-radiative recombination of devices and lower leakage current. Overall, this work reveals that DCB solvent treatment facilitates the ambient processing of efficient perovskite solar cells.

Experimental methods

Materials

Except otherwise specified, most materials and solvents were supplied by Sigma-Aldrich. In addition, lead(II) iodide (PbI_2), lead(II) chloride (PbCl_2), cesium iodide (CsI), and potassium chloride (KCl) were purchased from TCI America, while formamidine iodide (FAI) and FK209 were supplied by Greatcell solar materials and spiro-OMeTAD from Ossila. Absolute ethanol, chlorobenzene, and dichlorobenzene were purchased from VWR.

Formulation of perovskite precursor solution

To prepare the perovskite precursor solution, the commercially available FAI, CsI, and PbI_2 were dissolved in a mixed solvent of *N,N*-dimethylformamide (DMF), and *N*-methyl-2-pyrrolidone

(NMP). A double cation perovskite of formulation $\text{Cs}_{0.17}\text{FA}_{0.83}\text{-PbI}_3$ was prepared in the mixed solvents. The precursor solution was then stirred for about 5 hours at 700 rpm to obtain a yellowish transparent perovskite precursor solution. Before spin coating, the perovskite solution was filtered with a 0.45 μm PTFE filter (Ossila).

Device fabrication

All perovskite layers were deposited under the same ambient environmental conditions ($\text{RH} \approx 15\text{--}35\%$), and anti-solvents were applied during the second step of the spin coating process.

Solar cells. The following multilayer stacking order with different anti-solvent treatments is used in fabricating a planar heterojunction solar cell device: glass/indium-doped tin oxide (ITO)/tin oxide (SnO_2)/potassium chloride (KCl)/perovskite/spiro-OMeTAD/Au. The ITO-coated glass substrate ($1.1 \times 20 \times 15$ mm, $12\text{--}20 \Omega \text{sq}^{-1}$, sourced from Colorado concept coating LLC) underwent a cleaning process including sequential ultrasonication in a mixture of Hellmanex III detergent and deionized water, followed by acetone, and isopropyl alcohol, each for 15 minutes. The cleaned glass/ITO substrates were then dried and placed in an oven at 140°C for 20 minutes to remove any remaining solvents. Finally, the dried glass/ITO substrates underwent UV-ozone treatment for 20 minutes to cleanse the surface before spin-coating the SnO_2 as the ETL layer. A thin layer of SnO_2 was deposited through the spin coating of the SnO_2 solution at 3000 rpm for 30 s and annealed at 150°C for 30 minutes in air. A layer of aqueous KCl solution (3 mg mL^{-1}) was spin coated at 5000 rpm for 20 s and annealed in air at 100°C for 20 minutes, followed by UV-ozone treatment for 10 min. Subsequently, the perovskite precursor solution was deposited using a two-step spin-coating process under open-air conditions to form a thin layer of perovskite film under a relative humidity of 19–22% at a speed of 4000 rpm for 10 s and then at 5000 rpm for 15 s. During the 5000-rpm spin-coating step, 200 μL of each EtOH, CB, and DCB anti-solvents were dynamically dispensed on the substrates, facilitating the formation of the intermediate phase film.¹² This was followed by a two-step annealing process, initially at 70°C for 5 minutes and 150°C for 10 minutes. Next, the spiro-OMeTAD solution was prepared as previously reported^{16,22} and was spin-coated atop the perovskite layer to form the hole transport layer (HTL) and stored in a desiccator overnight. The solar cells were finished by thermally evaporating an 80 nm layer of gold atop the spiro-OMeTAD layer.

Electron-only devices. Electron-only devices were fabricated by subsequently spin-coating the perovskite solution on the glass/ITO/ SnO_2 layer and [6,6]-phenyl- C_{61} -butyric acid methyl ester (PC_{61}BM) solution (15 mg mL^{-1} in chlorobenzene) at 2000 rpm for 30 s, followed by thermally evaporated silver electrodes. The devices were measured under dark conditions. The measured $J\text{--}V$ curves were fitted using a modified Mott–Gurney equation.^{47,48}

Device and thin film characterization

To examine the morphology and microstructural evolution of perovskite thin films, we employed field-emission scanning



electron microscopy (FE-SEM) with a Verios G4 XHR SEM (Thermo Scientific). For the X-ray diffraction (XRD) analysis of the films, a Malvern Panalytical Empyrean IV, PIXcel 3D, and 1Der detector was utilized to study crystallinity and ascertain the phase composition of the perovskite films. Time-resolved photoluminescence (TRPL) and steady-state photoluminescence (PL) spectra were obtained from the FLS1000 photoluminescence spectrometer (Edinburgh Instrument, FLS1000). The J - V characteristics were measured using a Keithley 4200 source meter for the PSCs in the dark and under simulated 100 mW cm⁻² AM 1.5 G solar radiation. This radiation was generated by a 450 W xenon lamp (Oriel Sol 2A Class ABA), calibrated with a reference silicon cell covered by KG5 filtered glass. The short-term stability curves of the PSCs were recorded by collecting their J - V response under 100 mW cm⁻² irradiance generated by a steady-state LED lamp (Ossila Solar Simulator Class AAA). Fluxim AG Paios system with white LED light was used to carry out the transient photocurrent (TPC), transient photovoltage (TPV), capacitance-voltage (Mott-Schottky analysis), and electrochemical impedance spectroscopy (EIS) measurements.

The XPS experiments were conducted with a VersaProbe III instrument from Physical Electronics, outfitted with a concentric hemispherical analyzer beside a monochromatic Al K α X-ray source ($h\nu = 1486.6$ eV). Charge neutralization was omitted in the procedure. Calibration of the binding energy axis involved the use of sputter-cleaned Cu (Cu 2p_{3/2} = 932.62 eV, Cu 3p_{3/2} = 75.1 eV) and Au foils (Au 4f_{7/2} = 83.96 eV). Measurements were conducted at a takeoff angle of 45° relative to the sample surface plane, resulting in a typical sampling depth of 3–6 nm (where 95% of the signal originated from this depth or shallower). Peaks were calibrated with reference to the CH_x band in the carbon 1 s spectra at 284.8 eV. The analysis size was approximately 200 μ m in diameter. Quantification was achieved using instrumental relative sensitivity factors (RSFs), which consider the X-ray cross-section and inelastic mean free path of electrons. Major elements (> 5 atom%) on homogeneous samples typically exhibit less than 3% standard deviations, whereas minor elements may show significantly higher deviations.

Conflicts of interest

The authors declare no conflict of interest.

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References

- 1 Y. Cheng, Q. Wei, Z. Ye, X. Zhang, P. Ji, N. Wang, L. Zan, F. Fu and S. Frank Liu, *Sol. RRL*, 2022, **6**, 2200418.
- 2 G. Li, Z. Wang, Y. Wang, Z. Yang, P. Dong, Y. Feng, Y. Jiang, S.-P. Feng, G. Zhou, J.-M. Liu and J. Gao, *Small*, 2023, **19**, 2301323.
- 3 J. Jiao, C. Yang, Z. Wang, C. Yan and C. Fang, *Results Eng.*, 2023, **18**, 101158.
- 4 K. Lin, J. Xing, L. N. Quan, F. P. G. de Arquer, X. Gong, J. Lu, L. Xie, W. Zhao, D. Zhang, C. Yan, W. Li, X. Liu, Y. Lu, J. Kirman, E. H. Sargent, Q. Xiong and Z. Wei, *Nature*, 2018, **562**, 245–248.
- 5 X. Guo, I. M. Asuo, A. Pignolet, R. Nechache and S. G. Cloutier, *Opt. Mater. Express*, 2022, **12**, 153–165.
- 6 J. Moon, Y. Mehta, K. Gundogdu, F. So and Q. Gu, *Adv. Mater.*, 2023, 2211284.
- 7 Y. Zhao, X. Yin, P. Li, Z. Ren, Z. Gu, Y. Zhang and Y. Song, *Nano-Micro Lett.*, 2023, **15**, 187.
- 8 I. M. Asuo, D. Gedamu, I. Ka, L. F. Gerlein, F.-X. Fortier, A. Pignolet, S. G. Cloutier and R. Nechache, *Nano Energy*, 2018, **51**, 324–332.
- 9 R.-J. Jin, Y.-H. Lou, L. Huang, K.-L. Wang, C.-H. Chen, J. Chen, F. Hu and Z.-K. Wang, *Adv. Mater.*, 2024, 2313154.
- 10 A. Kojima, K. Teshima, Y. Shirai and T. Miyasaka, *J. Am. Chem. Soc.*, 2009, **131**, 6050–6051.
- 11 M. A. Green, E. D. Dunlop, J. Hohl-Ebinger, M. Yoshita, N. Kopidakis and X. Hao, *Prog. Photovoltaics Res. Appl.*, 2021, **29**, 657–667.
- 12 X. Cao, G. Su, L. Hao, J. Zhou, Q. Zeng, X. He and J. Wei, *J. Mater. Chem. C*, 2024, **12**, 1631–1639.
- 13 J. Sun, F. Li, J. Yuan and W. Ma, *Small Methods*, 2021, **5**, 2100046.
- 14 X. Yuan, R. Li, Z. Xiong, P. Li, G. O. Odunmbaku, K. Sun, Y. Deng and S. Chen, *Adv. Funct. Mater.*, 2023, **33**, 2215096.
- 15 X. Li, D. Bi, C. Yi, J.-D. Décoppet, J. Luo, S. M. Zakeeruddin, A. Hagfeldt and M. Grätzel, *Science*, 2016, **353**, 58–62.
- 16 I. M. Asuo, D. Gedamu, N. Y. Doumon, I. Ka, A. Pignolet, S. G. Cloutier and R. Nechache, *Mater. Adv.*, 2020, **1**, 1866–1876.
- 17 K. Geistert, S. Ternes, D. B. Ritzler and U. W. Paetzold, *ACS Appl. Mater. Interfaces*, 2023, **15**, 52519–52529.
- 18 W. Zhang, T. Zhang, L. Qin, S.-Z. Kang, Y. Zhao and X. Li, *Chem. Eng. J.*, 2024, **479**, 147838.
- 19 N. J. Jeon, J. H. Noh, Y. C. Kim, W. S. Yang, S. Ryu and S. I. Seok, *Nat. Mater.*, 2014, **13**, 897–903.
- 20 O. A. Khaleel and D. S. Ahmed, *Opt. Mater.*, 2022, **126**, 112148.
- 21 J. Li, R. Yang, L. Que, Y. Wang, F. Wang, J. Wu and S. Li, *J. Mater. Res.*, 2019, **34**, 2416–2424.
- 22 D. Gedamu, I. M. Asuo, D. Benetti, M. Basti, I. Ka, S. G. Cloutier, F. Rosei and R. Nechache, *Sci. Rep.*, 2018, **8**, 12885.



- 23 H. Yang, H. Wang, J. Zhang, J. Chang and C. Zhang, *Coatings*, 2019, **9**, 766.
- 24 K.-M. Lee, C.-J. Lin, B.-Y. Liou, S.-M. Yu, C.-C. Hsu, V. Suryanarayanan and M.-C. Wu, *Sol. Energy Mater. Sol. Cells*, 2017, **172**, 368–375.
- 25 D. Prochowicz, M. M. Tavakoli, A. Solanki, T. W. Goh, K. Pandey, T. C. Sum, M. Saliba and P. Yadav, *J. Mater. Chem. A*, 2018, **6**, 14307–14314.
- 26 Y. Chen, W. Wu, R. Ma and C. Wang, *J. Mol. Struct.*, 2019, **1175**, 632–637.
- 27 G.-H. Kim, J. Jeong, H. Jang, J. W. Kim and J. Y. Kim, *Thin Solid Films*, 2018, **661**, 122–127.
- 28 A. D. Taylor, Q. Sun, K. P. Goetz, Q. An, T. Schramm, Y. Hofstetter, M. Litterst, F. Paulus and Y. Vaynzof, *Nat. Commun.*, 2021, **12**, 1878.
- 29 X. Gong, M. Li, X.-B. Shi, H. Ma, Z.-K. Wang and L.-S. Liao, *Adv. Funct. Mater.*, 2015, **25**, 6671–6678.
- 30 W. Luo, J. Zhong, L. Zhang, Y. Peng, J. Lu, M. Xu and Z. Ku, *Mater. Sci. Semicond. Process.*, 2023, **164**, 107603.
- 31 A. D. Taylor, Q. Sun, K. P. Goetz, Q. An, T. Schramm, Y. Hofstetter, M. Litterst, F. Paulus and Y. Vaynzof, *Nat. Commun.*, 2021, **12**, 1878.
- 32 M. Saliba, T. Matsui, J.-Y. Seo, K. Domanski, J.-P. Correa-Baena, M. K. Nazeeruddin, S. M. Zakeeruddin, W. Tress, A. Abate, A. Hagfeldt and M. Grätzel, *Energy Environ. Sci.*, 2016, **9**, 1989–1997.
- 33 Q. Chen, H. Zhou, T.-B. Song, S. Luo, Z. Hong, H.-S. Duan, L. Dou, Y. Liu and Y. Yang, *Nano Lett.*, 2014, **14**, 4158–4163.
- 34 T. J. Jacobsson, J. P. Correa-Baena, E. Halvani Anaraki, B. Philippe, S. D. Stranks, M. E. Bouduban, W. Tress, K. Schenk, J. Teuscher, J. E. Moser, H. Rensmo and A. Hagfeldt, Unreacted PbI₂ as a Double-Edged Sword for Enhancing the Performance of Perovskite Solar Cells, *J. Am. Chem. Soc.*, 2016, **138**(32), 10331–10343.
- 35 P. Fassl, V. Lami, A. Bausch, Z. Wang, M. T. Klug, H. J. Snaith and Y. Vaynzof, *Energy Environ. Sci.*, 2018, **11**, 3380–3391.
- 36 M. Kim, G.-H. Kim, T. K. Lee, I. W. Choi, H. W. Choi, Y. Jo, Y. J. Yoon, J. W. Kim, J. Lee, D. Huh, H. Lee, S. K. Kwak, J. Y. Kim and D. S. Kim, *Joule*, 2019, **3**, 2179–2192.
- 37 X. Zhang, S. Eurelings, A. Brancesco, W. Song, S. Lenaers, W. Van Gompel, A. Krishna, T. Aernouts, L. Lutsen, D. Vanderzande, M. Creatore, Y. Zhan, Y. Kuang and J. Poortmans, *ACS Appl. Mater. Interfaces*, 2023, **15**, 46803–46811.
- 38 H. Bi, Y. Fujiwara, G. Kapil, D. Tavgeniene, Z. Zhang, L. Wang, C. Ding, S. R. Sahamir, A. K. Baranwal, Y. Sanehira, K. Takeshi, G. Shi, T. Bessho, H. Segawa, S. Grigalevicius, Q. Shen and S. Hayase, *Adv. Funct. Mater.*, 2023, **33**, 2300089.
- 39 L. Chen, E. Kinsa Tekelenburg, K. Gahlot, M. Pitaro, J. Xi, A. Lasorsa, G. Feraco, L. Protesescu, P. C. A. van der Wel, G. Portale, P. Rudolf, C. J. Brabec and M. Antonietta Loi, *Energy Environ. Sci.*, 2023, **16**, 5315–5324.
- 40 S. Shao, Y. Cui, H. Duim, X. Qiu, J. Dong, G. H. ten Brink, G. Portale, R. C. Chiechi, S. Zhang, J. Hou and M. A. Loi, *Adv. Mater.*, 2018, **30**, 1803703.
- 41 S. Yuan, T. Zhang, H. Chen, Y. Ji, Y. Hao, H. Zheng, Y. Wang, Z. David Chen, L. Chen and S. Li, *Chem. Eng. J.*, 2022, **433**, 133227.
- 42 H. Zheng, X. Peng, T. Chen, T. Zhang, S. Yuan, L. Wang, F. Qian, J. Huang, X. Liu, Z. David Chen, Y. Zhang and S. Li, *J. Energy Chem.*, 2023, **84**, 354–362.
- 43 S. Béchu, M. Ralaïarisoa, A. Etcheberry and P. Schulz, *Adv. Energy Mater.*, 2020, **10**, 1904007.
- 44 A. Bensekhria, I. M. Asuo, I. Ka, R. Nechache and F. Rosei, Improved Performance of Air-Processed Perovskite Solar Cells via the Combination of Chlorine Precursors and Potassium Thiocyanate, *ACS Appl. Mater. Interfaces*, 2023, **15**(48), 56413–56423.
- 45 J. Tirado, M. Vásquez-Montoya, C. Roldán-Carmona, M. Ralaïarisoa, N. Koch, M. K. Nazeeruddin and F. Jaramillo, *ACS Appl. Energy Mater.*, 2019, **2**, 4890–4899.
- 46 K. Taek Cho, S. Paek, G. Grancini, C. Roldán-Carmona, P. Gao, Y. Lee and M. Khaja Nazeeruddin, *Energy Environ. Sci.*, 2017, **10**, 621–627.
- 47 N. Y. Doumon, M. V. Dryzhov, F. V. Houard, V. M. Le Corre, A. Rahimi Chatrri, P. Christodoulis and L. J. A. Koster, *ACS Appl. Mater. Interfaces*, 2019, **11**, 8310–8318.
- 48 P. N. Murgatroyd, *J. Phys. D: Appl. Phys.*, 1970, **3**, 308.

