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Electrochemically switchable rotaxanes: recent strides in new directions

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Are they still electrifying? Electrochemically switchable rotaxanes are well known for their ability to efficiently undergo changes of (co-)conformation and properties under redox-control. Thus, these mechanically interlocked assemblies represent an auspicious liaison between the fields of molecular switches and molecular electronics. Since the first reported example of a redox-switchable molecular shuttle in 1994, improved tools of organic and supramolecular synthesis have enabled sophisticated new architectures, which provide precise control over properties and function. This perspective covers recent advances in the area of electrochemically active rotaxanes including novel molecular switches and machines, metal-containing rotaxanes, non-equilibrium systems and potential applications.

Introduction

Rotaxanes, a class of mechanically interlocked molecules, consist in their most basic form of a macrocyclic wheel threaded on a linear axle molecule.¹ Bulky stopper groups placed at the ends of the molecular thread prevent deslipping of the wheel. Encouraged by synthetic advances made during the last decades, for example in template synthesis,^{2–4} self-sorting^{5,6} or

dynamic combinatorial chemistry,^{7–9} complex (supra)molecular architectures in a large structural complexity and variety^{10,11} have become more easily accessible. Conformational flexibility and dynamic nature of rotaxanes, given by the mechanical bond between their components, make them perfectly suited to undergo substantial structural changes when manipulated by external stimuli.^{12,13} Based on that, bi- and multi-stable rotaxanes arose as promising candidates for artificial molecular switches and machines over the last decades.^{14–16}

Electrochemically switchable rotaxanes bear functional groups, which can undergo reversible redox reactions. These rotaxanes provide (co-)conformational control and can reversibly generate large-amplitude molecular motions by

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German Research Foundation (DFG) in the group of Prof. A. James Link at Princeton University, USA, with focus on mechanically interlocked peptides. His current research interests include intertwined molecules, molecular switches and machines, organic synthesis, and electrochemistry.



Christoph Schalley studied chemistry at the University of Freiburg/Germany and graduated in 1993 from the Technical University Berlin. In his PhD work with Helmut Schwarz at TU Berlin, he received his thorough education as a mass spectrometrists and gas-phase chemist. After postdoctoral work on supramolecular chemistry with Julius Rebek, Jr. at The Scripps Research Institute, La

Jolla/USA, he combined both topics when he started his own research group as a habilitand at the University of Bonn. In 2005, he was appointed professor of organic chemistry at Free University Berlin. Meanwhile the Schalley group's research topics are diverse and comprise supramolecular chemistry in the gas phase as well as in solution and at interfaces.



electrochemical stimuli. In other words, they can convert electrochemical energy into directed movements at the molecular scale. Electrochemical energy is considered to be a simple, “clean” (no chemical waste production), selective and fast way to initiate molecular switching processes.^{17,18} Electrochemistry also offers a rich toolbox of methods to probe structural and electronic properties of even complex supramolecular ensembles.¹⁹ Valuable information including a qualitative and quantitative assessment of the available oxidation states,²⁰ electrochemical reaction mechanisms,²¹ diffusion constants,²² orbital energies,²³ and much more can be extracted by electro-analytical methods. Additionally, methods of supramolecular electrochemistry²⁴ provide rare access to thermodynamic and kinetic parameters of intramolecular (co-)conformational changes, which are often out of the scope of classical analytic methods in supramolecular chemistry.^{25,26}

Exactly 25 years ago, the first electrochemically switchable molecular shuttle based on a two-station rotaxane has been reported,²⁷ raising high hopes for the field of molecular switches and machines. Although these rotaxanes are among pioneering examples, systems have evolved from rather simple molecular shuttles to advanced supramolecular architectures bringing multiple functions together. Recent systems come with a precise electrochemical control of molecular motions, properties and functions. Additionally, new types of metal-organic rotaxanes provide potential applications as catalysts, sensors, and single-molecule electronic devices. Undoubtedly, the 2016 Nobel Prize in Chemistry^{28–30} “for the design and synthesis of molecular machines” marks the outstanding significance of these developments. On the other hand, it redirects the trajectory of the field of rotaxanes towards more functionalized and applicable systems.

Numerous excellent overviews covering switchable rotaxanes,^{1,31,32} metal-containing rotaxanes^{2,33,34} and electrochemically active supramolecular systems^{24,35,36} have been

published. This perspective is focused on recent developments of redox-active (pseudo)rotaxanes—especially those reported after the 2016 Nobel Prize—and applications arising from their electrochemical activity on the basis of a non-exhaustive list of illustrative examples. In particular, this summary gives a differentiated view on recent advances beyond classical molecular shuttles. It covers (i) rotaxanes generating molecular motions; (ii) metal-containing rotaxanes; (iii) applications of rotaxanes including sensing, catalysis, molecular electronics and molecular machines; and (iv) non-equilibrium systems.

Rotaxanes generating molecular motions

Fundamentals of redox-controlled molecular shuttles

Archetypical examples of electrochemically switchable rotaxanes are two-station molecular shuttles,²⁷ in which the position of a wheel on a molecular thread can be manipulated by reduction or oxidation (Fig. 1). Electroanalytic methods, such as commonly applied cyclic voltammetry (CV), can provide qualitative evidence for co-conformational changes through altered half-wave ($E_{1/2}$) or peak (E_p) potentials and can even quantify the distribution of isomers in different redox-states.^{24,35} The isomer distribution is given by the dimensionless co-conformational equilibrium constant K_x (Fig. 1, eqn (1)) for each redox-state (x). The efficiency of a molecular switch can be expressed as the binding enhancement of a station which reflects the ratio between K_x values in two redox-states (Fig. 1, eqn (2)). In recent years, supramolecular construction motifs have been continuously optimized to gain more control over co-conformational switching.³⁷

Electrochemical experiments (*e.g.* CV) at different scan rates or chrono methods can additionally help to identify reversibility, to elucidate reaction mechanisms, and to investigate the kinetics³⁸ of conformational changes. Furthermore, digital simulations³⁹ are

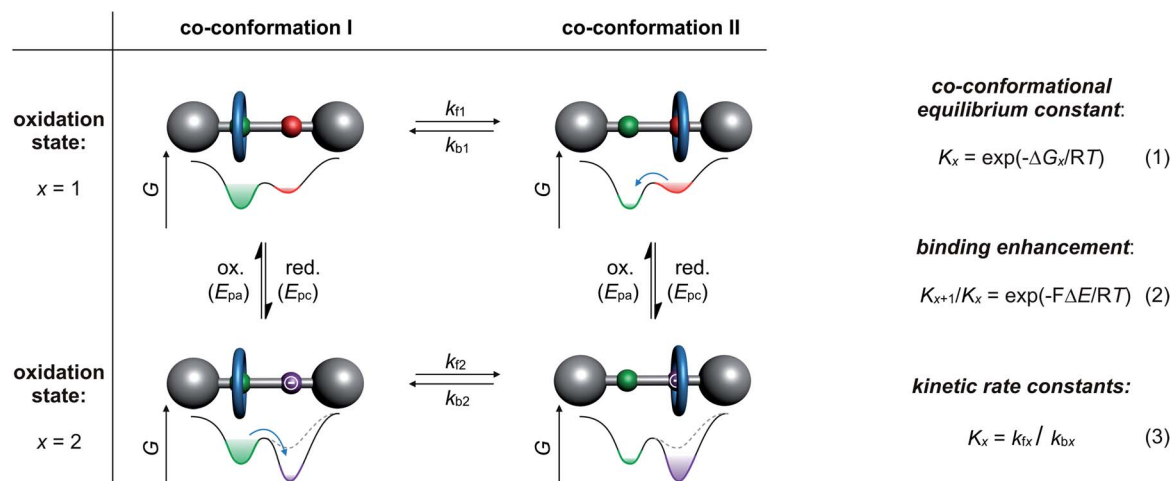


Fig. 1 A general square scheme for an electrochemically switchable [2]rotaxane. It illustrates co-conformational isomers and the corresponding energy landscapes of wheel translation in two stable redox-states ($x = 1$ or 2). The eqn (1)–(3) describe the thermodynamics and kinetics of the switching processes, where G_x is the Gibbs free energy in a given oxidation state, k a first-order rate constant, F the Faraday constant, T the temperature, ΔE a potential difference between two redox-reactions and R the ideal gas constant.



often used as a complementary tool not only to predict voltammetric responses of supramolecular systems but also to fit experimental data. Digital simulation provides access to thermodynamic and kinetic parameters (K_x , k_{fx} and k_{bx}) as a self-consistent set of values (Fig. 1, eqn (3)). Naturally, the accuracy of the derived values depends on the correct choice of reaction mechanism, the quality of experimental data and the input variables. Improvement in accuracy can be achieved by comparing with data, for example, diffusion or binding constants derived from independent analytical methods or model systems.

For instance, an electrochemically controllable molecular shuttle with a highly selective co-conformational distribution was reported in a 2003 landmark publication of Leigh, Paolucci and co-workers.⁴⁰ The thread of hydrogen-bonded [2]rotaxane **1** (Fig. 2) bears a naphthalimide unit, which serves simultaneously as a stopper for the wheel and as a redox-active recognition side. In the neutral state, the tetralactam wheel preferably encircles the diamide over the naphthalimide station. Electrochemical one-electron reduction generates a stable naphthalimide radical anion ($E_{pc} = -1.40$ V vs. Fc/Fc⁺, Fc = ferrocene), which is now a more favorable hydrogen bond acceptor. By CV experiments and digital simulations, a binding enhancement of $>10^8$ for the naphthalimide station after reduction was determined leading to a translation of the wheel towards it. In comparison to the free thread, re-oxidation ($E_{pa} = -0.89$ V) is anodically shifted reflecting stabilization of the radical anion by the hydrogen-bonded wheel. Variable-temperature experiments in THF revealed the kinetic parameters $\Delta H^\ddagger = 59$ kJ mol⁻¹ and $\Delta S^\ddagger = 36$ kJ mol⁻¹ for wheel translation from the diamide to the naphthalimide station after reduction, resulting in a shuttling time of ~ 50 μ s at room temperature. In a recent report, the shuttling mechanism was studied in more detail by transient two-dimensional infrared spectroscopy.⁴¹

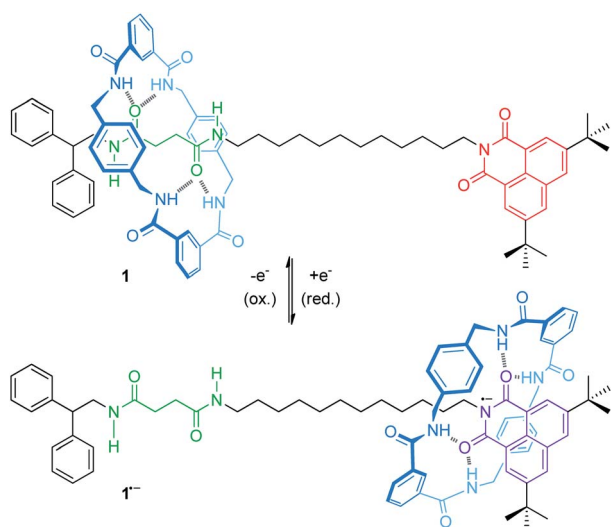


Fig. 2 Electrochemically switchable molecular shuttle **1**. In the neutral state, the tetralactam wheel preferably encircles the diamide station. One-electron reduction to the **1**^{•-} state moves the wheel to the naphthalimide radical anion site.

Rotaxane shuttles enabling rare interactions

The mechanical bond prevents dissociation of non-covalently bound molecular entities and can stabilize interactions, which are otherwise unfavorable.⁴² Thus, rotaxanes offer an opportunity to study these rare interactions between subcomponents or even utilize them for (co-)conformational control.

For example, Stoddart and co-workers recently reported on the tristable [2]rotaxane **2**⁶⁺ (Fig. 3) with three different recognition sites on the thread, namely two triazoles, two tetrafluorobenzenes and one bipyridinium unit.⁴³ In the initial state, the quadruply charged cyclobis(paraquat-*p*-phenylene) (CBPQT⁴⁺) wheel preferably encircles one of the relatively π -electron-rich triazoles by forming a donor-acceptor complex. The wheel's binding affinity can then be altered by stepwise reduction. Cyclovoltammetry reveals an initial two-electron reduction ($E_{pc1} = -0.14$ V vs. Ag/AgCl) to **2**^{2+/4+} closely followed by a one-electron reduction ($E_{pc2} = -0.22$ V). The resulting tri-radical species **2**³⁽⁺⁾ is stabilized by radical-pairing interactions between the doubly reduced wheel CBPQT²⁽⁺⁾ and the singly reduced bipyridinium station. A further three-electron reduction ($E_{pc3} = -0.69$ V) to **2**⁽⁰⁾ neutralizes wheel and bipyridinium station. Now, the CBPQT⁽⁰⁾ ring migrates to one of the π -electron-poor tetrafluorobenzene moieties. This supramolecular charge-transfer complex—which is enabled through spatial pre-organization of the wheel—is rather weak and cannot be easily studied in “the context of a host-guest complex”,⁴³ unless stabilized by a mechanical bond.

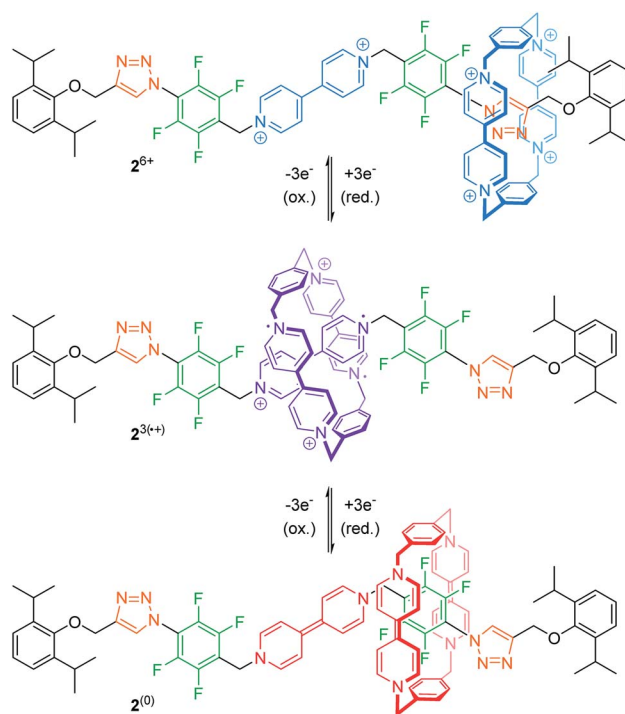


Fig. 3 Tristable [2]rotaxane **2**⁶⁺ in its three different redox-states and resulting co-conformations. The unusual CBPQT⁽⁰⁾/tetrafluorobenzene charge-transfer complex is stabilized by the mechanical bond.



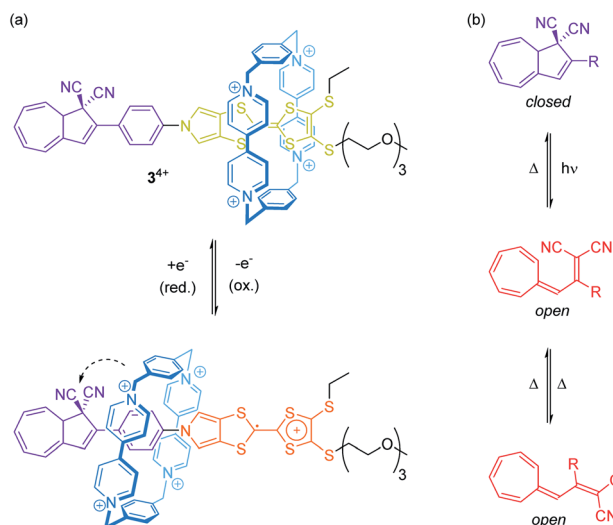


Fig. 4 (a) Redox-active [2]rotaxane 3^{4+} bearing a photo- and thermoswitchable dihydroazulene unit as stopper. The dashed arrow illustrates through-space interactions between CBPQT $^{4+}$ wheel and dihydroazulene site. (b) Photo- and thermoswitching of the 1,1-dicyano-dihydroazulene moiety.

Rare through-space interactions also play a role in the orthogonally switchable CBPQT $^{4+}$ -rotaxane recently reported by Jeppesen, Nielsen and co-workers (Fig. 4a).⁴⁴ The CBPQT $^{4+}$ wheel of [2]rotaxane 3^{4+} is threaded on a redox-active monopyrrolo-tetrathiafulvalene axle, which is terminated on one side with a 1,1-dicyano-dihydroazulene moiety, a molecular photo- and thermoswitch (Fig. 4b). In both photoisomeric forms, the closed dihydroazulene and open vinylheptafulvene form, deslipping of the wheel is sterically hindered. The presence of the wheel anodically shifts both tetrathiafulvalene (TTF) oxidation potentials as compared to the free thread. However, the first oxidation already induces a movement of the wheel towards the dihydroazulene/phenylene site forming weak interactions between the cyano groups and the wheel's bipyridinium α -protons. These rather weak interactions between wheel and axle significantly accelerate the thermal back-conversion of the dihydroazulene switch (open to closed form) after photoirradiation.

Weak through-space interactions were also found to affect TTF redox-potentials of our recently reported rotaxane shuttles consisting of a TTF-decorated crown ether wheel and a neutral axle.⁴⁵ The TTF/TTF $^{+}$ oxidation was cathodically shifted compared to the free TTF-wheel, indicating significant stabilization of the TTF $^{+}$ state by interactions between the mechanically bound axle and the TTF $^{+}$ moiety.

Oligorotaxanes and daisy chain rotaxanes

Linear main-chain oligo- and polyrotaxanes, in which several wheels are threaded on an axle, have been used as supramolecular scaffolds to generate large-amplitude motions.^{46,47} The resulting structural changes are reminiscent of molecular folding and contraction processes of biopolymers, *e.g.* protein folding or muscle contraction.

The Stoddart group presented an oligo[3]rotaxane, in which axle and wheels as a whole undergo a redox-triggered folding/unfolding along the trajectory of the axle.⁴⁸

Our lab recently reported on oligorotaxanes, in which inter-wheel distances can be controlled by radical-cation interactions between TTF-decorated wheels (Fig. 5a).⁴⁹ In the initial state, the 24-crown-8 ether wheels of oligo[4]rotaxane 4^{3+} are bound to secondary ammonium stations by hydrogen bonding. Reversible oxidations of the wheels' TTF units weaken the wheel-axle interactions through counteracting Coulomb repulsion. As shown by electrochemistry, UV/Vis-NIR and electron paramagnetic resonance spectroscopy, the TTF $^{+}$ moieties can form inter-wheel mixed-valence ($4^{+/4+}$) and radical-cation ($4^{3+/6+}$) interactions, which consequently lead to a contraction of the ensemble in both redox-states. Further oxidation (TTF $^{+}$ to TTF $^{2+}$) to the 4^{9+} state creates strong charge repulsion between the TTF $^{2+}$ units and leads to a maximization of the inter-wheel distances. The resulting accordion-like motion (Fig. 5b) is independent of the molecular thread and, thus, promising for the construction of piston-type rotaxane motors.

Electrochemically switchable daisy chain rotaxanes are ideal prototypes to generate a muscle-like contraction. Moulin and Giuseppone introduced a [c2]daisy chain rotaxane (5^{4+} , Fig. 6) with two different endgroups, where the mechanical actuation is controlled by acid/base addition.⁵⁰ The crown ether/

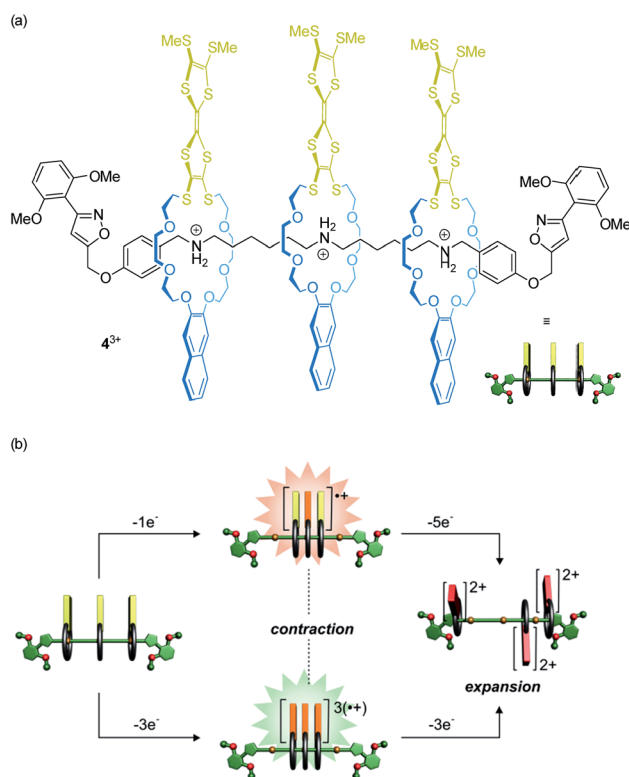


Fig. 5 (a) Structure of the electrochemically switchable oligo[4]rotaxane 4^{3+} . (b) Graphical summary of the redox-controlled generation of an accordion-like contraction/expansion motion. Partly adapted with permission from ref. 49. Copyright (2019) John Wiley & Sons.



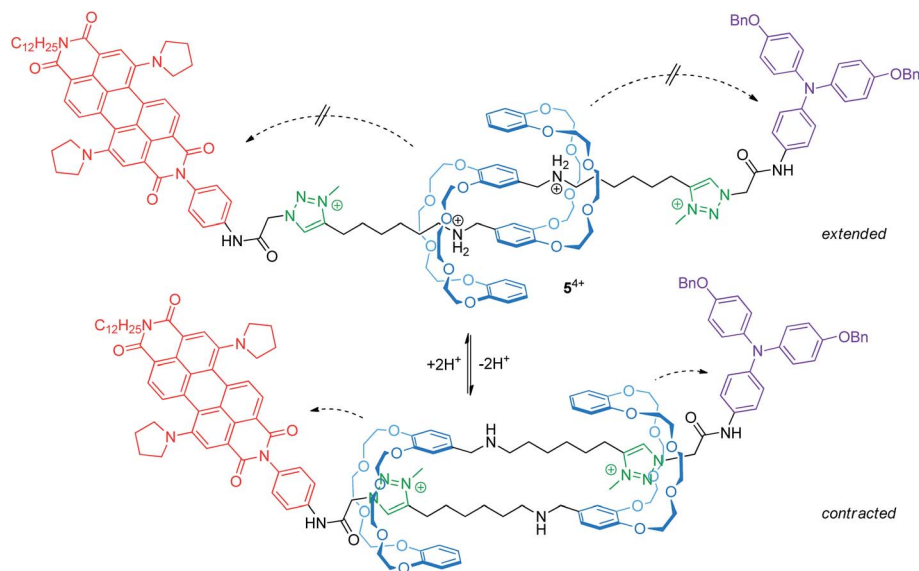


Fig. 6 Switching of the acid/base-responsive [c2]daisy chain rotaxane 5^{4+} between its extended and contracted form. Both forms show altered electrochemical properties. The dashed arrows illustrate electronic through-space interactions between the wheels and the redox-centers.

ammonium rotaxane is decorated with two redox-active stoppers, a 1,7-dipyrrolidinyl-substituted perylene diimide and a triarylamine unit. In the extended form (protonated), three reversible electrochemical processes are observed by cyclic voltammetry: a two-electron reduction ($E_{1/2}^1 = -1.30$ V vs. Fc/Fc^+) centered on the perylene diimide, an oxidation wave (+0.19 V) resulting from a superposition of two one-electron oxidations centered on the triarylamine and the perylene diimide and a second one-electron oxidation of the perylene diimide unit (+0.31 V). The cyclic voltammogram is a superposition of those of the corresponding individual reference compounds; thus, both redox-active groups are electrochemically unaffected in the extended form. In the contracted form (deprotonated), however, the two-electron reduction wave is broadened and the peak current of the re-oxidation is significantly decreased indicating a lower chemical stability. At the same time, the second oxidation of the perylene diimide unit is shifted to lower potentials. The authors have attributed both effects to the proximity of the crown ethers to the electro-active stoppers after contraction.

Other types of redox-controlled motion

Electrochemical control of wheel–axle interactions opens up possibilities for other types of controlled motion. For example, electrochemically switchable pseudorotaxanes consisting of redox-active threads and cyclodextrins⁵¹ or cucurbiturils^{52–54} as macrocyclic compounds have been used to create stimuli responsive polymers and materials.

Recently, Stoddart and co-workers built the pseudo[1]rotaxane 6^{6+} based on their well investigated CBPQT⁴⁺ wheel covalently attached to a viologen thread (Fig. 7a).⁵⁵ Complex 6^{6+} efficiently switches between a rope-like and a lasso-type conformation upon reduction. CV and differential pulse voltammetry show that the first reduction wave ($E_{\text{pc}} = -0.33$ V vs. Ag/AgCl) is a combination of a two- and a one-electron reduction leading to triradical species $6^{3(++)}$. An anodic shift ($\Delta E_{\text{pc}} =$

+0.10 V relative to a suitable control compound), indicates the stabilization of the self-inclusion conformation to result from the strong, intramolecular radical-pairing interactions between

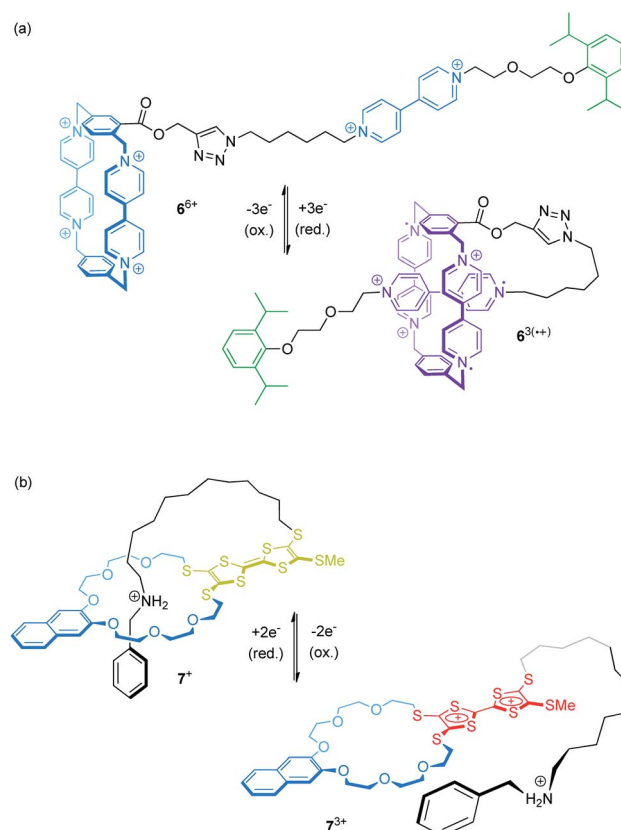


Fig. 7 Two examples for electrochemically controlled pseudo[1] rotaxanes which are convertible between a rope-like and a lasso-type conformation: (a) 6^{6+} undergoes self-inclusion after reduction, whereas (b) 7^{7+} is preferentially threaded in the initial state and unthreads upon oxidation.



Besides shuttling, controlling the rotation of wheels around an axle (pirouetting) is a worthwhile but challenging goal.^{33,37,58} Crown ether/ammonium [3]rotaxane **8**²⁺ which enables electrochemically controlled pirouetting (Fig. 8) was recently reported.⁵⁹ The wheels are decorated with redox-switchable TTF units. A structurally similar control with only one TTF-wheel displays two reversible one-electron oxidations at $E_{1/2} = 0.16$ V

and $E_{1/2}^2 = 0.40$ V vs. Fc/Fc⁺ for the TTF/TTF^{•+} and TTF^{•+}/TTF²⁺ transition, respectively. In contrast, rotaxane **8**²⁺ exhibits a splitting of the TTF/TTF^{•+} wave ($E_{1/2}^1 = 0.08$ V and $E_{1/2}^2 = 0.19$ V) and a two-electron oxidation at $E_{1/2}^3 = 0.48$ V. The splitting indicates an intramolecular mixed-valence complex (TTF₂)^{•+}, whose stability is reflected in a comproportionation constant of $K_c = 73$. The second one-electron oxidation $E_{1/2}^2$ yields a stable radical-cation dimer (TTF^{•+})₂, which was confirmed by electron paramagnetic resonance and UV/Vis-NIR spectroscopy. Supported by digital simulations and density functional theory, it could be shown that the wheels in the oxidation states 2TTF, (TTF₂)^{•+} and (TTF^{•+})₂ adopt a syn co-conformation, whereas an anti co-conformation is found for the 2TTF²⁺ state.

Great effort has been made in recent years to understand and control the kinetics of redox-induced translational motion in rotaxane shuttles.³⁸

Lately, the dependence of the shuttling kinetics of an electrochemically switchable CBPQT⁴⁺-rotaxane on the electrolyte concentration was investigated by Jeppesen and co-workers.⁶⁰ For electrochemical experiments, supporting electrolytes are frequently required in high concentrations to increase solution conductivity. It was found that tetraalkylammonium salts ($R_4N^+X^-$) in general lead to an increased switching rate, whereas the size of the R_4N^+ cation has no significant effect on wheel shuttling. However, a increased rate was observed when using ClO_4^- instead of PF_6^- as the electrolyte's anion.

The Stoddart group has recently investigated shuttling rates in a redox-switchable ring-in-ring rotaxane, where the axle itself incorporates a redox-active macrocycle.⁶¹ Electrochemically induced shuttling was found to be significantly slower compared to their previous systems.⁶² The slow motion results in an electrochemical hysteresis for switching as shown by CV and digital simulations. Furthermore, the slow shuttling could be used to characterize a diamagnetic tetraradical complex by ¹H NMR spectroscopy, whose lifetime is otherwise too short.

Metal-containing rotaxanes

Since Sauvage's seminal work^{33,63} on Cu-containing catenanes, metal-ligand interactions have always played a key role in the synthesis of mechanically interlocked molecules, for example in a templating pre-organization of the precursor molecules.³⁴ More recently, the active template approach² evolved, in which metal ions serve as templates and as mediators for the formation of mechanical bonds. However, metals are also integral subcomponents of rotaxanes bringing with them versatile properties and functions. Metals can serve as persistent redox-centers, which change conformational or optoelectronic properties of a rotaxane with their oxidation state.

Rotaxanes bearing covalently bound metal complexes

The group of Korybut-Daszkiewicz reported recent progress on their switchable rotaxanes with a cyclidene/crown ether⁶⁴ motif. Fig. 9 shows [2]rotaxane **9⁴⁺** which consists of an axle with a 14-

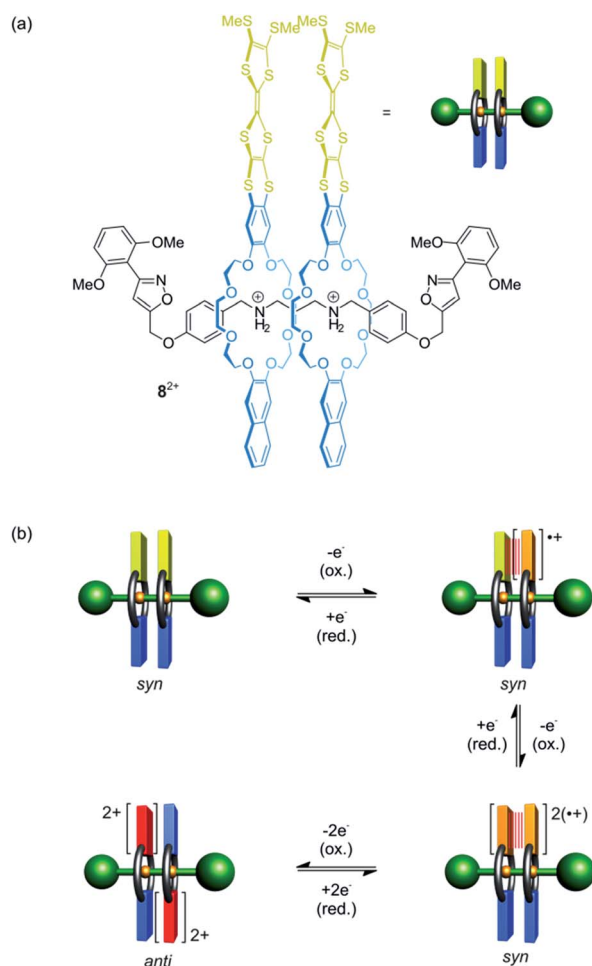


Fig. 8 (a) Electrochemically switchable crown ether/ammonium [3] rotaxane 8^{2+} , which allows to control relative wheel pirouetting motions. (b) Graphical summary of the most-stable co-conformations of 8^{2+} in its four stable redox-states.

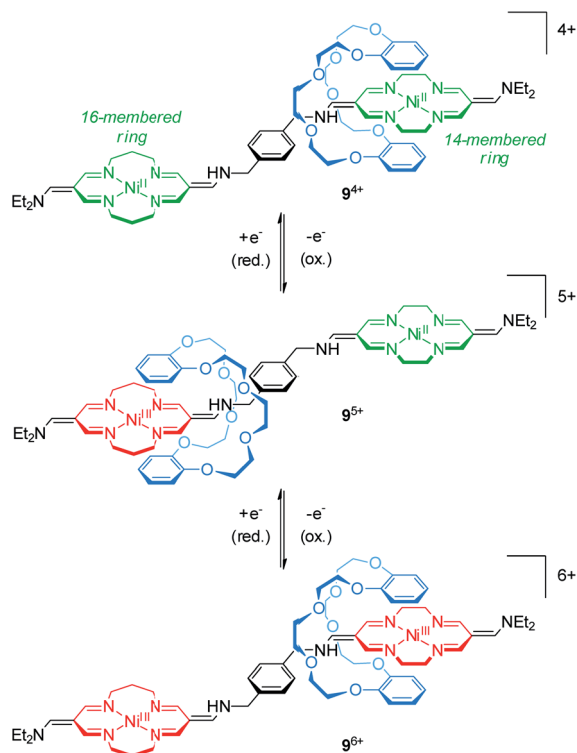


Fig. 9 Redox-controlled molecular shuttle 9^{4+} based on the cyclodene/crown ether motif in its three stable oxidation states ($\text{Ni}^{\text{II}}/\text{Ni}^{\text{II}}$, $\text{Ni}^{\text{III}}/\text{Ni}^{\text{II}}$ and $\text{Ni}^{\text{III}}/\text{Ni}^{\text{III}}$).

and a 16-membered Ni^{II} -complexing tetraazamacrocycle and a threaded dibenzo-24-crown-8 wheel.⁶⁵ In the initial state, the crown ether preferentially encircles the 14-membered ring since it is a stronger electron acceptor. However, the 16-membered Ni^{II} complex shows a less positive oxidation potential at 1.15 V vs. Ag/AgCl as shown by differential pulse voltammetry. One-electron oxidation ($\text{Ni}^{\text{II}}/\text{Ni}^{\text{III}}$) converts the 16-membered ring to a better electron acceptor and the wheel translates to the Ni^{III} binding site. Subsequent one-electron oxidation of the 14-membered ring complex at higher potentials induces a reversed shuttling motion of the wheel. Interestingly, a splitting of the signal for the $\text{Ni}^{\text{II}}/\text{Ni}^{\text{III}}$ oxidation is observed indicating that wheel shuttling takes place on the time scale of the experiment.

The team of Ma synthesized a pH-controlled molecular shuttle based on a crown ether/ammonium [2]rotaxane.⁶⁶ The crown ether macrocycle is decorated with a Pt^{II} porphyrin moiety, which allows following wheel shuttling by changes of room temperature phosphorescence. Additionally, CV indicates electronic interaction between the Pt^{II} porphyrin and a redox-active viologen unit, which is additionally incorporated in the axle.

Delavaux-Nicot, Nierengarten, Maisonhaute and others have reported the pillar[5]arene rotaxane **10** bearing ten peripheral Fc units attached to the wheel to investigate electron transfer processes (Fig. 10).⁶⁷ One of the stoppers is substituted with two lipoic acid esters, which allow grafting the rotaxane onto gold microelectrodes. In solution, CV and Osteryoung squarewave voltammetry showed a single ten-electron wave ($E_{1/2} = 0.59$ vs.

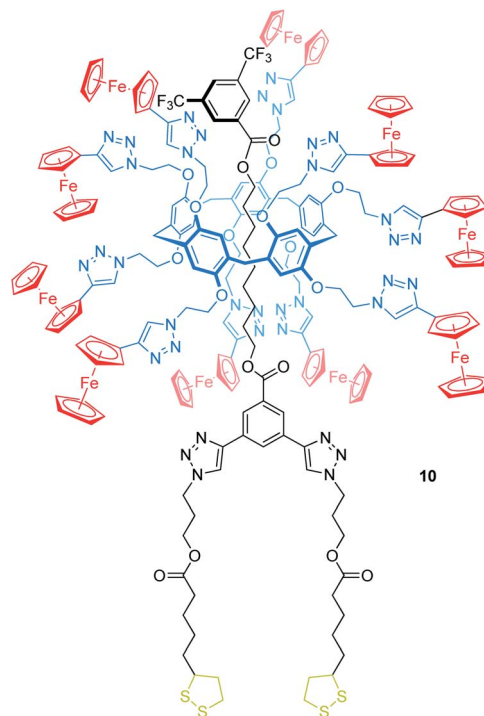


Fig. 10 Structure of the ferrocene-decorated rotaxane **10**, which has been grafted onto a gold microelectrode.

SCE, SCE = saturated calomel electrode) due to the Fc/Fc^+ oxidations. A set of control experiments confirmed that no other redox-process occurs in this potential window. After grafting the rotaxane onto a gold microelectrode, a self-assembled monolayer of **10** was probed by ultrafast voltammetry revealing the rates of electron transfer k_{ET} in dependence of different scan rates and surface coverages. A low coverage of 19 nm² per molecule leads to a k_{ET} of $5.8 \times 10^5 \text{ s}^{-1}$, whereas a denser coverage decreases k_{ET} (e.g. 3.1 nm² per molecule, k_{ET} of $1.7 \times 10^5 \text{ s}^{-1}$). The authors suggest that the electron transfer between Fc units (hopping) and from Fc units to the electrode surface (heterogeneous electron transfer) is fast when the molecules are isolated. In the presence of neighboring molecules, however, intermolecular interactions restrict the motion towards the crowded electrode surfaces which decreases the k_{ET} values.

Ulfkjær *et al.* described a metal-containing [2]rotaxane, which is stoppered with gold nanoparticles (Fig. 11).²² A diammonium thread with a central Fc unit and cucurbit[7]uril self-assemble into the very stable ($K_a = 10^{13} \text{ M}^{-1}$) pseudo[2]rotaxane **11**²⁺ in water. Basic hydrolysis gives **12**²⁺, whose thiol groups are able to coordinate onto gold nanoparticles to form a mechanically interlocked assembly. CV experiments were performed with the pseudo[2]rotaxanes **11**²⁺ and **12**²⁺ indicating typical anodic shifts of the Fc/Fc^+ transition after encapsulation by cucurbit[7]uril ($\Delta E_{1/2} = +0.17$ and $+0.19 \text{ V}$, respectively). Scan-rate-dependent measurements reveal a diffusion-controlled electrochemical process for **11**²⁺. Normally, a diffusion coefficient smaller than that of the free thread would be expected as an indication of the increased bulkiness of the assembly. However, dithiol **12**²⁺ displays an abnormal diffusion coefficient



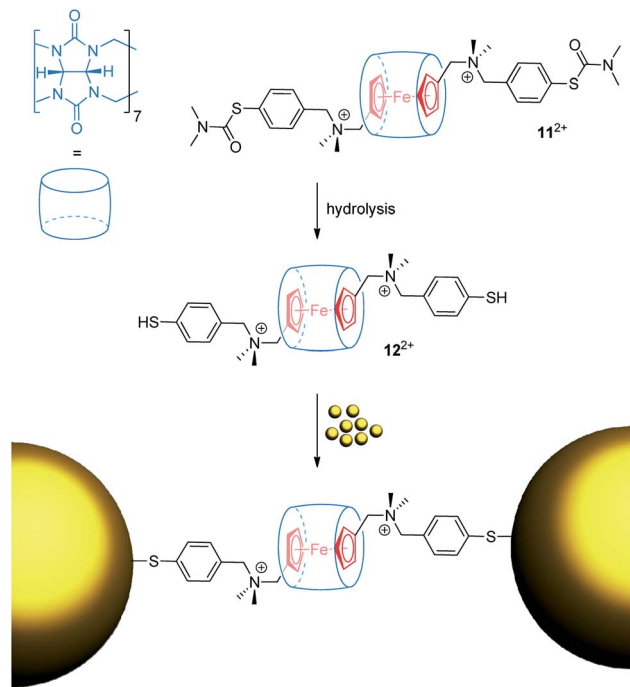


Fig. 11 Reaction scheme for the formation of a gold nanoparticle-stoppered rotaxane from the pseudo[2]rotaxane precursor 12^{2+} .

and interfacial electron transfer rate constant. The authors suggest that thiol groups bind more strongly to the surface of the used glassy carbon electrode than thiocarbamate groups; thus, pseudorotaxane 12^{2+} displays interfacial rather than diffusion-controlled electron transfer kinetics.

Osakada and co-workers described a series of electro-active, Fe-containing [1–4]rotaxanes that were prepared from a common crown ether/ammonium pseudo[2]rotaxane precursor by copper-catalyzed azide–alkyne cycloaddition.⁶⁸ The triazole groups of the redox-switchable [3]rotaxane can be additionally used for metal complexation. The [3]rotaxane forms 1 : 1 or 1 : 2 coordination complexes depending on the concentration of added Pd^{II} ions.

An interesting approach of stabilizing sensitive metal complexes by rotaxation was lately reported by the group of Terao.⁶⁹ They used an α -cyclodextrin-based [1]rotaxane as insulating shell to stabilize a Ni^{II} -bis(dithiobenzoate) complex. CV demonstrates a reversible and quasi-reversible reduction at -1.26 and -1.83 V vs. Fc/Fc^+ . This represents the first example of reversible bis(dithiobenzoate) transition metal complex reduction. The same strategy was later used to stabilize a redox-active $\text{Pt}(\text{II})$ bis(dithiobenzoate) complex.⁷⁰

Rotaxanes as ligands for metal-coordination

Rotaxanes can provide unique coordination sites for metal ions. Goldup, Roessler and co-workers have described a library of heteroleptic metal/rotaxane complexes (Fig. 12).⁷¹ An active template synthesis followed by addition of a transition metal salt yields the heteroleptic metal/rotaxane complexes 13–15. CV and EPR reveal a reversible, metal-centered oxidation ($\text{Cu}^{\text{I/II}}$) at

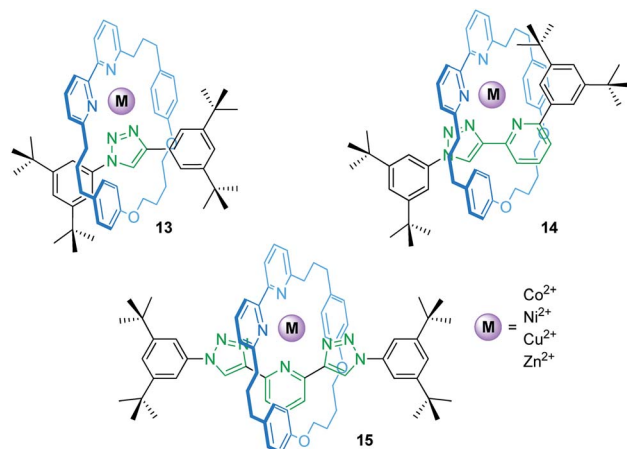


Fig. 12 Heteroleptic transition metal/rotaxane complexes 13–15 by Goldup, Roessler and co-workers.

+0.96 V and +0.56 V vs. SHE (standard hydrogen electrode) and a reversible ligand-centered reduction at -1.50 and -1.44 V for the copper species $\text{Cu}(\mathbf{14})$ and $\text{Cu}(\mathbf{15})$, respectively. The potential difference for the $\text{Cu}^{\text{I/II}}$ oxidation ($\Delta E_{1/2} = 0.40$ V) results from the fact that $\mathbf{14}$ provides an ideal binding site for Cu^{I} , whereas $\mathbf{15}$ provides a flexible enough coordination sphere to sufficiently stabilize Cu^{I} and Cu^{II} . For $\text{Cu}(\mathbf{13})$, only an irreversible process is observed suggesting unstable Cu^{I} and Cu^{II} complexes. CV control experiments with a non-interlocked complex display two semi-reversible steps for the $\text{Cu}^{\text{I/II}}$ transition caused by ligand disproportionation. Thus, the mechanical bond in $\mathbf{14}$ and $\mathbf{15}$ prevents a structural reorganization and allows redox-reactions without degradation of the complexes which would otherwise not be achievable.

The groups of Armaroli and Diederich synthesized a Cu^{I} bis-phenanthroline rotaxane by using tetracyano-*p*-quinodimethane as the stopper unit in a $[2 + 2]$ cycloaddition-retro-electrocyclization.⁷² Electrochemical experiments revealed for the resulting rotaxane a reversible $\text{Cu}^{\text{I/II}}$ oxidation and two reversible one-electron reductions for the dicyanoquinodimethane moiety, which was generated during the stoppering reaction.

An efficient operation of multi-ring rotaxanes, such as in the case of palindromic [3]rotaxanes, requires an ingenious design of a switchable supramolecular binding motif. The group of Flood developed a novel construction motif for a Cu^{I} -containing pseudo[3]rotaxane (Fig. 13).⁷³ Their multi-component system consists of a phenanthroline-bearing macrocycle, Cu^{I} ions, a 5,5'-dimethyl-2,2'-bipyridine ligand, and a redox-switchable 3,6-bis(5'-methyl-2-pyridyl)-1,2,4,5-tetrazine ligand. In the initial state, Cu^{I} favors coordination by one macrocycle and one bipyridine ligand. Thus, the potential for single-electron reduction of the tetrazine ligand ($E_{\text{pc}} = -0.95$ V vs. Ag/AgCl) is mainly unaffected. Reduction of the tetrazine to its radical anionic form, however, dramatically enhances its ligand strength and initiates ligand exchange with the bipyridines to form the pseudo[3]rotaxane $\mathbf{16}^+$. As shown by an anodically shifted re-oxidation potential ($E_{\text{pa}} = -0.25$ V), the pseudo[3]



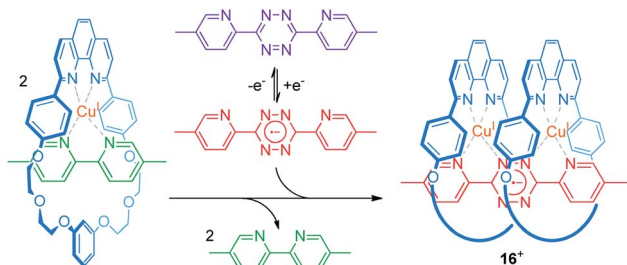


Fig. 13 Electrochemical switching of a multi-component system to reversibly form the palindromic pseudo[3]rotaxane 16^+ .

rotaxane state is energetically stabilized. Digital simulations and control experiments confirm a stepwise mechanism. The first step (exchange with the first bipyridine/Cu^I/macrocycle complex) is significantly faster than the second step, due to increased bulkiness of the intermediate pseudo[2]rotaxane. However, the second step displays a higher thermodynamic driving force (19 vs. 25 kJ mol⁻¹) in accordance with a positive thermodynamic cooperativity. Flood recently reported a similar observation for non-covalently coupled “cyanostar” macrocycles in a metal-free pseudo[3]rotaxane.⁷⁴

Applications of redox-active rotaxanes

Electrochemically switchable rotaxanes are obviously appealing for integration into electronic devices, for example by interfacing them with electrodes.^{75,76} Starting more than a decade ago, switchable rotaxanes on solid supports have been intensively studied in self-assembled monolayers.^{77–81} This research greatly expanded the fundamental understanding of molecular switching on solid support, but also disclosed that organic structures often suffer from low long-term persistence and structural integrity.^{82,83} However, in recent years also other applications arose from access to novel construction motifs for redox-active rotaxanes.

Beer developed a halogen bonding [2]rotaxane⁸⁴ which can serve as an electrochemical bromide sensor, even in water-containing media (Fig. 14). Rotaxane 17^+ provides a halogen-bonding anion binding cavity and a ferrocene unit as an electrochemical probe. ¹H NMR titration experiments in an organic solvent mixture with 10 vol% D₂O give association constants for the anions Cl⁻, Br⁻, I⁻ and SCN⁻, which the authors report to be 101, 340, 1125 and 102 M⁻¹, respectively. The halide anions nicely follow the Hofmeister series; however, the SCN⁻ anion, which should exceed I⁻, displays a rather poor binding indicating that geometry and size complementarity play a crucial role. The electrochemical response to Cl⁻, Br⁻ and SCN⁻ was further probed by square-wave voltammetry showing cathodic shifts for the Fc/Fc⁺ transition. Whereas addition of 20 equivalents of Cl⁻ and SCN⁻ display negligibly small shifts ($\Delta E_{1/2} < 5$ mV), the addition of Br⁻ generates more pronounced shifts (*e.g.* 20 equiv. result in $\Delta E_{1/2} = -10$ mV). Control experiments in dry and “wet” solvents indicate that hydration strongly contributes to the observed effects. This is in line with the more pronounced voltammetric response for Br⁻, since Br⁻ is expected to display

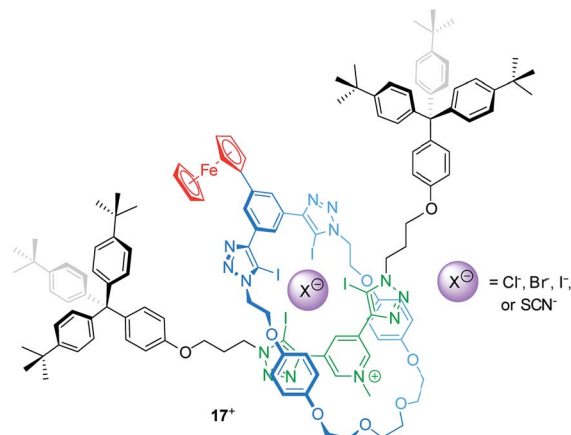


Fig. 14 Structure of the bromide-sensing [2]rotaxane 17^+ , which provides a halogen bonding anion binding cavity.

lower hydration energy. In another recent report, Beer presented a Fe-decorated [3]rotaxane shuttle which utilizes anion recognition to induce wheel translation.⁸⁵

The Tanaka group has described porphyrin-phthalocyanine catalyst 18^{4+} for O₂ to H₂O reduction,⁸⁶ which is based on their previously reported fourfold threaded [2]rotaxane motif⁸⁷ (Fig. 15a). The heterodimeric scaffold incorporates a μ -oxo-dinuclear iron complex (Fe^{III}-O-Fe^{III}). Experiments using a rotating ring disc electrode showed a fairly high positive onset potential of 0.78 V *vs.* RHE (reversible hydrogen electrode) indicating a high catalytic activity for oxygen reduction.

Furthermore, the determined number of 3.8 transferred electrons (at 0.47 V) and the percentage of H₂O generation (88%) confirm that a four-electron reduction of O₂ to H₂O is the dominant reaction. As shown by CV, the E_{pc} value of a redox-wave at $E_{1/2} = 0.60$ V, corresponding to a one-electron reduction of the Fe^{III} phthalocyanine complex, significantly increases in the presence of O₂. The proposed mechanism, in which protonation and Fe^{III} phthalocyanine reduction result in an initial water loss, is shown in Fig. 15b. Subsequently, the resulting mixed-valence species and O₂ form a peroxo-Fe^{III}

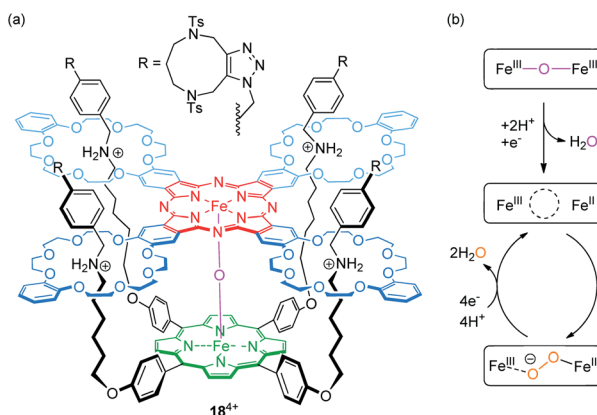


Fig. 15 (a) Structure of the fourfold threaded [2]rotaxane catalyst 18^{4+} , which incorporates a μ -oxo-dinuclear iron complex. (b) Suggested mechanism for the reduction of oxygen to water.



complex ($\text{Fe}^{\text{III}} \cdots \text{O}^- - \text{O} - \text{Fe}^{\text{III}}$), which releases, after four-electron reduction and protonation, two H_2O molecules. The authors attribute the high catalytic activity to the flexible rotaxane scaffold and its inherent structural adaptability. It provides a well suitable environment to accommodate the longer O_2 molecule as well as the shorter bridging O^{2-} ligand.

Conjugated molecular chains that conduct electric current—so-called molecular wires—are attracting considerable attention due to their potential use for nanoscale electronic devices. Fig. 16 shows rotaxane **19**, consisting of an oligoyne thread with 6 conjugated triple bonds, which is assembled within a scanning tunneling microscope break junction.⁸⁸ Both, the presence of the 3,5-diphenylpyridine endcapping groups and the encircled phenanthroline containing macrocycle, chemically stabilize the oligoyne by steric shielding and yield a kind of insulated molecular wire. Molecular conductance measurements showed a slight increase of conductance for the rotaxane **19** compared to the free thread suggesting a rather small effect of the wheel on the electronic structure of the thread. Rotaxation also comes with a drawback: due to the possibility of inhibiting molecular bridging between wheel and gold contacts, the presence of the threaded wheel reduces the probability for successful molecular junction formation.

A series of heteroleptic Cu^{I} pseudo[2]rotaxanes featuring bis[2-(diphenylphosphino)phenyl] ether and macrocyclic phenanthroline ligands have been recently investigated as materials for organic light emitting diodes (OLEDs).⁸⁹ In a poly(methyl methacrylate) film (1 wt%), they exhibit excellent photophysical properties with photoluminescence quantum yields up to 44% ($\lambda_{\text{exc}} = 380 \text{ nm}$), which makes them promising for optoelectronic devices. An OLED device based on one of the Cu^{I} pseudorotaxanes shows intense green electroluminescence at 10 V with a maximum intensity at $\sim 520 \text{ nm}$. It is noteworthy that the pseudorotaxane displays a higher electroluminescence and enhanced device stability compared to a reference compound with a non-macrocyclic phenanthroline ligand. The authors attribute the improved performance to the pseudorotaxane structure that prevents complex disassembly and, thus, chemical degradation.

Transport and controlled release of substrates are vital functions of natural molecular machines. Credi and co-workers have reported a synthetic molecular transporter based on the [2]rotaxane **20**⁵⁺ (Fig. 17).⁹⁰ The crown ether

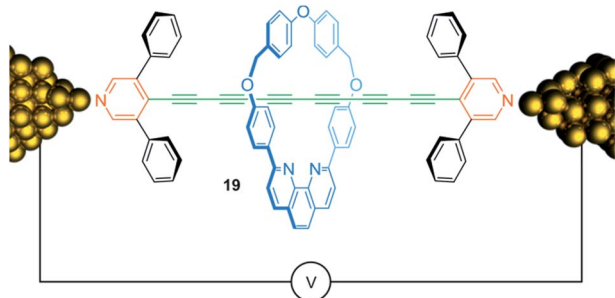


Fig. 16 Schematic representation of oligoyne rotaxane **19** within in a single molecule break junction.

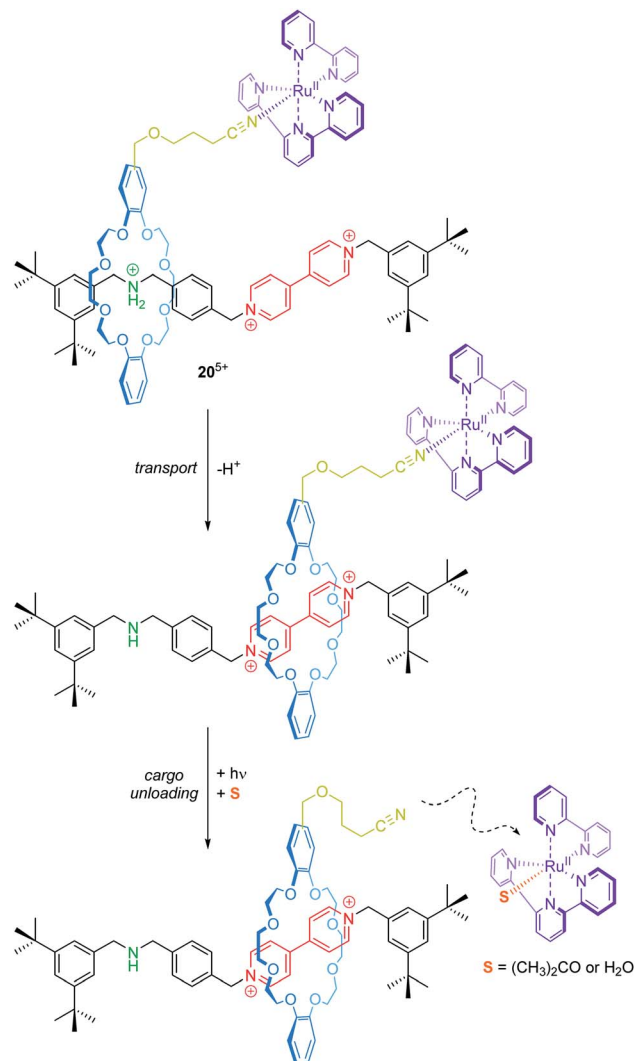
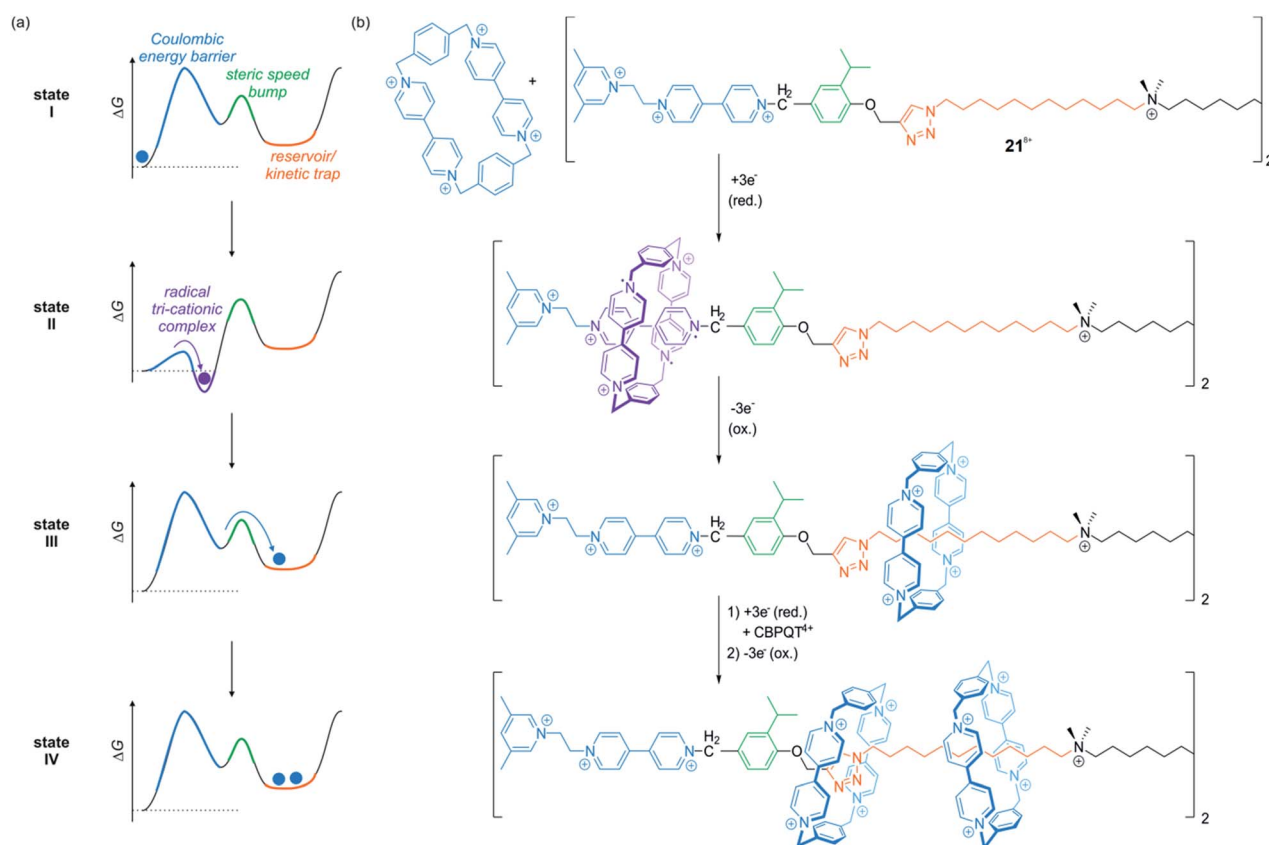


Fig. 17 Operation sequence of rotaxane-based molecular transporter **20**⁵⁺ by orthogonal use of acid/base addition and light as external stimuli. The dashed arrow illustrates the release of the formerly bound Ru^{II} complex.

wheel of the rotaxane carries a ruthenium complex, which acts as a molecular cargo. The position of the Ru^{II} complex on the thread can be manipulated by acid/base addition. Irradiation with visible light triggers a ligand exchange and the Ru^{II} complex dissociates from the rotaxane (unloading). Here, electrochemistry serves as a tool to monitor *in situ* operation of the transporter. In the protonated state, two one-electron reduction processes for the bipyridinium unit at -0.35 and -0.75 V vs. SCE and a metal-centered oxidation at $+1.34 \text{ V}$ are observed. Deprotonation leads to wheel shuttling (transport) towards the viologen station and cathodically shifts both reduction potentials of it, whereas the $\text{Ru}^{\text{II/III}}$ oxidation remains unaffected. After irradiation, however, the signal at $+1.34 \text{ V}$ disappears indicating a successful photodissociation. Although the transport distance is rather small, this proof-of-principle nicely illustrates that artificial systems are capable of reversibly performing transport and cargo loading/unloading processes.

In a series of recent papers, the Stoddart group developed variants of electrochemically driven molecular pumps which operate in non-equilibrium states.^{98–100} Through an elegantly designed sequence of steric “speed bumps” and redox-switchable binding sites implemented into a molecular axle,

Fig. 18 shows the operation sequence of their latest system,¹⁰⁰ which involves the symmetric axle 21^{8+} consisting of two molecular pumping units at its ends and a central reservoir to accumulate CBPQT⁴⁺ wheels. In the ground state, the tetra-cationic wheel and the 3,5-dimethylpyridinium/viologen moiety repel each other (state I). Upon reduction, however, repulsion is reduced and a CBPQT²⁽⁺⁾ ring from the bulk solution threads on the axle to form a tri-radical complex with the radical cationic viologen unit (state II). Re-oxidation of the ensemble destabilizes the complex and drives the wheel over the steric barrier given by the isopropylphenyl moiety. The wheel slips into the reservoir domain where it is kinetically trapped (state III). After one redox cycle, it is possible to repeat the process and collect another wheel on each half of the symmetric thread to yield a [5]rotaxane (state IV). Consecutive chemical reduction/oxidation is experimentally difficult due to formation of waste products.¹⁰⁰ However, the ratchet mechanism can also be powered by bulk electrolysis cycles in which the working electrode is alternately held at a constant reductive or oxidative potential.



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