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Graphene/polyaniline@carbon cloth composite as a high-performance flexible supercapacitor electrode prepared by a one-step electrochemical co-deposition method†

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Increasing numbers of flexible universal energy storage devices are required for wearable and portable products. A typical representation of energy storage devices are supercapacitors. However, their applications are greatly limited due to their low conductivity, which is strongly restricted by the effect of the binder. We designed and prepared a flexible and binder-free supercapacitor based on a graphene–polyaniline@carbon cloth composite by a one-step electrochemical co-deposition process. The composite exhibited excellent electrochemical performance with a high specific capacitance of 793 F g^{−1} at a current density of 0.5 A g^{−1} and favorable cycle stability (retention of 81% of initial specific capacitance after 10 000 cycles). The flexible symmetric device also showed satisfactory specific capacitance of 512 F g^{−1} at a current density of 0.5 A g^{−1}, desirable cycling stability (retention of 85% specific capacitance over 10 000 cycles) and considerable energy density of 11.38 W h kg^{−1} at a power density of 199.80 W kg^{−1}. We believe that this flexible and binder-free supercapacitor with excellent electrochemical properties will meet the demands in roll-up displays, electronic paper and wearable electronic products.

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1. Introduction

Due to rapid technological developments, there has been a huge demand for universal flexible energy storage devices because of their extraordinary potential application in wearable and portable electronic products such as roll-up displays, electronic paper, and wearable systems for personal multimedia.^{1,2} Flexible supercapacitors (SCs) are a typical representation of flexible energy storage devices. They can meet the needs of flexible energy storage due to their high power density, high operational safety, fast charge/discharge capacity, excellent cycle stability and acceptable energy density.^{3,4} The key components of a flexible SC are the electrodes, which make a significant contribution to the electrochemical performance of the device. Therefore, there is a huge requirement for designing and preparing electrode materials with high flexibility, moderate mechanical strength, large specific capacitance and excellent cycle stability. In recent years, plenty of works have

concentrated on using various carbon-based materials such as carbon nanotubes,^{5,6} carbon anions,⁷ activated carbon,⁸ carbon fibers⁹ and graphene¹⁰ to fabricate SCs. Several studies have demonstrated that to construct SCs with high flexibility and excellent electrochemical properties not only requires constitutive materials to be equipped with high mechanical integrity and outstanding electrical performance, but also relies on their controlled assembly into functional devices.^{11–15}

Among various materials, pseudocapacitive materials such as conducting polymers and transition metal oxides exhibit great potential in flexible SCs electrodes because of their fast and reversible redox processes as well as excellent electrochemical behaviors. In particular, polyaniline (PANI), as one type of conducting polymer, has been intensely studied in recent years due to its ultrahigh theoretical capacitance, favorable environmental stability, controlled structure and morphology, special conducting mechanism and doping/dedoping chemistry.¹⁶ PANI has become one of the most promising candidates for energy storage device materials. However, the volumetric changes and mechanical degradation of PANI electrodes during charging/discharging processes result in poor cycling stability and rate capability, and hinder their practical applications severely.^{17,18} Furthermore, the low conductivity of neutral PANI also greatly impedes its electrochemical properties when it is used directly as electrode

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material in SCs. Introduction of carbon-based materials, for example, graphene hydrogel, into PANI, means that these electrode composite materials have greatly improved conductivity, higher specific capacitance and longer cycling life.¹⁹ However, it is necessary to add binders and conductive additives not only to improve the adhesion between electrode materials and substrates, but also to enhance the conductivity of the electrodes, which inevitably and severely reduces the electronic performance of these devices.²⁰ Thus, two main problems must be solved to accelerate the development of flexible SCs: (i) search for electrode materials with high electrochemical performance; (ii) enhance the connection between electrodes with collectors to improve the conductivity of the device. Carbon cloth could be used as a conductive and flexible substrate material for flexible SCs electrodes due to its high conductivity, flexibility and fast electron/ion transport.

Herein, we successfully prepared a flexible, binder-free graphene/PANI electrode based on carbon cloth (GP@cc) by a simple one-step electrochemical co-deposition method. This electrode exhibited an impressive specific capacitance of 793 F g⁻¹ at 0.5 A g⁻¹ in 1 M H₂SO₄ aqueous electrolyte in a three-electrode system. Moreover, the GP@cc composite showed excellent rate capability and favorable cycling stability (retained 81% of initial specific capacitance after 10 000 cycles at 5 A g⁻¹). Carbon cloth could serve as collector due to its high conductivity and fast electron/ion transport. Thus, we directly used GP@cc composites as electrodes and collectors to fabricate a flexible symmetric SC using 1 M H₂SO₄ as electrolyte without adding any other binder or conductive additives. The integrated device not only showed extraordinary mechanical flexibility and high conductivity, it also exhibited the desired specific capacitance (512 F g⁻¹ at 0.5 A g⁻¹) and acceptable cycling stability (retained 85% of initial specific capacitance after 10 000 cycles at 5 A g⁻¹).

2. Experimental section

2.1 Synthesis of graphene oxide (GO)

On the basis of the modified Hummers' method, GO was synthesized through the oxidation of natural graphite powder.²¹ The obtained GO dispersion was then centrifuged for 30 min at 4000 rpm to remove aggregates. Then, it was purified by dialysis for 1 week to remove remaining salts impurities. Finally, the resulting GO dispersion was adjusted to 0.36 mg mL⁻¹ for subsequent experiments.

2.2 Preparation of graphene/polyaniline composite

Carbon cloth (1 cm × 2 cm) was soaked in acetone and distilled water, sonicated for 30 min to remove grease and impurities on the surface. This procedure was repeated several times, respectively, and then the cloth was dried under vacuum at 60 °C. Then, 135 μL aniline was added to 10 mL distilled water, and sonicated at room temperature until the solution became transparent. The transparent solution mentioned above was added to 1 mL 1.0 M Na₂SO₄ in 15 mL GO dispersion, sonicated for 10 min, and the resulting mixed solution used as the

electrolyte for electrochemical co-deposition. The latter was conducted in a three-electrode system (carbon cloth as the working electrode, platinum (Pt) foil as the counter electrode, and Ag/AgCl as the reference electrode) using cyclic voltammetry at a scan rate of 2 mV s⁻¹ with a potential window from -1.2 to 0.8 V and with 20 cyclic voltammetry (CV) cycles. As a comparison, PANI@cc was prepared *via* cyclic voltammetry with a potential window from 0 to 0.8 V at 2 mV s⁻¹; RGO@cc was also obtained by CV with a potential window from -1.2 to 0 V at 2 mV s⁻¹. These samples were cleaned by distilled water and the procedure repeated several times to remove the absorbed monomer or remaining salts, and they were noted as RGO@cc, PANI@cc, and GP@cc, respectively.

2.3 Device fabrication

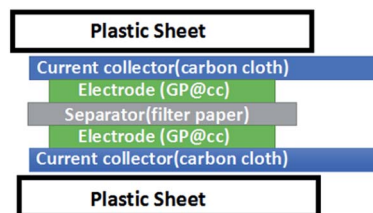
Two pieces of graphene/PANI@cc composites were soaked in 1.0 M H₂SO₄ aqueous electrolyte for 12 h with constant stirring. These were used as both electrodes of the flexible SC device, and were separated by filter paper also soaked with 1.0 M H₂SO₄ aqueous electrolyte. Carbon cloth was directly used as a current collector without any binders or other additives. Then, they were assembled into a layered structure by "sandwiching" the entire SC device between two plastic sheets using clips and sealed by parafilm for subsequent electrochemical tests. The structure of the flexible SC device is shown in Scheme 1.

2.4 Characterization

The morphologies of RGO@cc, PANI@cc, and GP@cc were examined on a Zeiss Ultra 55 field emission scanning electron microscope. Raman spectra were collected on an XploRA laser Raman spectrometer using a 532 nm laser beam. X-ray photoelectron spectra (XPS) were recorded on PerkinElmer PHI 5300 using an Mg K α source (1253.6 eV). XRD analyses were done with a Rigaku D/Max 2500 X-ray diffractometer with a Cu K α radiation source ($k = 1.54 \text{ \AA}$).

2.5 Electrochemical characterization and analyses

All electrochemical experiments were carried out using a CHI 760E electrochemical workstation (Shanghai CH Instrument Co., China). The electrochemical measurements of the device were conducted in a two-electrode system, whereas those for the graphene/PANI@cc electrode were performed in a three-electrode system in 1.0 M H₂SO₄ aqueous solution electrolyte at room temperature, with Pt and Ag/AgCl as counter and reference electrodes, respectively. The potential ranges of CV



Scheme 1 Structure of the flexible supercapacitor device.



and galvanostatic charge/discharge (GCD) tests were 0.0 to 0.8 V, and the GP@cc was directly employed as the working electrode. Electrochemical impedance spectroscopy measurements were carried out at open circuit potential with a sinusoidal signal over a frequency range from 100 kHz to 0.01 Hz at an amplitude of 5 mV. The cycle life of the GP@cc electrode and the device were conducted by GCD measurements at a current density of 5 A g⁻¹. The specific capacitances (C_{wt}) derived from galvanostatic charge curves were calculated based on the following formula: $C_{wt} = I\Delta t/m\Delta V$, where I is the constant discharge current, Δt is the time for full discharge, m is the mass of one electrode, and ΔV represents the voltage drop upon discharging (excluding the IR drop). In symmetric SCs, the specific capacitance (C_{sc}) of the device was calculated according to $C_{sc} = 2C_{wt}$. The energy density of the symmetric device was calculated based on the following formula: $E_{wt} = C_{wt}V^2/8$, where V is the operating voltage. Power densities (P) were calculated using the following formula: $P = E/\Delta t$, where Δt is the discharge time.

3. Results and discussion

The graphene/PANI@cc composite was successfully prepared *via* a simple one-step electrochemical co-deposition method using GO and aniline mixed solution as the electrolyte in a three-electrode system. The mass transfer of GO from the bulk phase of the dispersed system to the surface of the electrode leads to the absorption of GO to the surface of the electrode. The high negative potential (<-0.75 V, *vs.* Ag/AgCl) can overcome the energy barriers for the reduction of oxygen functionalities ($-OH$, $C-O-C$ on the plane and $-COOH$ on the edge). Moreover, cathodic peak potentials of GO electro-reduction shift negatively as pH increased. Protonation in the electro-reduction process occurs and is facilitated at lower pH values. Thus, we could use the following equation to explain the reduction mechanism of GO during the deposition, indicating the important role of hydrogen ions in this process.^{22,23}



Therefore, the reduction of GO during the deposition could be achieved in various electrolytes, such as solutions of acid, alkali and saline. Carbon cloth was used as working electrode and served as a flexible and highly conductive substrate to synthesize graphene/PANI composites due to its favorable mechanical strength, excellent conductivity and high flexibility. Moreover, we could adjust the area of the carbon cloth to control the size of the graphene/PANI composite. A scanning electron micrograph of pure carbon cloth is shown in Fig. S1,[†] and reveals that carbon cloth was weaved by abundant carbon fibers and had a highly smooth surface. Fig. 1a–f illustrates the morphology and structure of the resultant RGO@cc, PANI@cc and GP@cc composites by scanning electron microscopy (SEM). Fig. 1a shows the well-defined and interconnected graphene framework completely coating the surface of the carbon cloth. Fig. 1b illustrates PANI also totally coating the surface of carbon cloth. Fig. 1c–f show SEM images of the GP@cc composite.



Fig. 1 SEM image of RGO@cc (a), PANI@cc (b) and SEM images of GP@cc (c–f).

From Fig. 1c we can see that graphene and PANI are distributed on the surface of carbon cloth uniformly. Fig. 1d clearly shows the surface of carbon cloth to be tightly coated by graphene nanosheets and that PANI nanowires are uniformly dispersed on graphene nanosheets. From Fig. 1e we can see more clearly that PANI nanowires are distributed on graphene nanosheets uniformly and connected to nearby carbon fibers with void spaces. Graphene nanosheets not only interconnected with each other to form a thin graphene layer but also interacted with PANI nanowires through the large surface of carbon cloth. Fig. 1f demonstrates the graphene nanosheets interconnected and formed 3D porous frameworks with pore sizes from sub-micrometers to several micrometers to facilitate fast access of ions to electrolytes.

To explore the interaction between graphene with PANI more clearly, Raman spectroscopy was performed. Fig. 2a illustrates the Raman spectra of RGO@cc, PANI@cc and GP@cc. There are two obvious peaks at 1349 and 1600 cm⁻¹ in the spectrum of RGO@cc that correspond to the characteristic peaks of carbon materials, D and G bands, respectively.^{24,25} The Raman spectrum of PANI@cc displays the characteristic peaks of PANI. The 1599 cm⁻¹ peak corresponds to the C–C stretching of benzenoid units, the Raman peaks at 1348 cm⁻¹ and 1186 cm⁻¹ are



Fig. 2 (a) Raman spectra of RGO@cc, PANI@cc and GP@cc; (b) XPS spectra of GP@cc; (c and d) C 1s (c) and N 1s (d) XPS spectra of GP@cc.



connected with radical cation C–N⁺ stretching and C–H in-plane bending, respectively. Compared with neat PANI@cc, the GP@cc composite shows an obvious difference in sharpness and intensities of peaks within low wave number regions. The peak at 412 cm^{−1} is related to the out-of-plane C–H wag. The peak corresponding to a phenazine-like segment appears at around 607 cm^{−1} and the peak around 813 cm^{−1} can be assigned to C–H deformation.^{26,27} According to the spectra of the GP@cc composite, we could confirm that doped PANI and reduced GO coexisted in the composite. Furthermore, Fig. 2b–d illustrates the XPS of the GP@cc composite in a wide scan. Fig. 2b reveals that the resultant graphene/PANI@cc composite was composed of the four elements C, N, O and S, confirming the existence of PANI and that PANI was doped by SO₄^{2−}. Fig. 2c shows the XPS spectrum of the N 1s region. We could conclude that its existence mainly included three electronic states: the binding energy (BE) at around 399.4 eV related to the benzenoid amine; the BE at around 398.8 eV assigned to the quinoid amine; and the BE at around 401.4 eV associated with the nitrogen cationic radical (N⁺).²⁸ Fig. 2e presents the XPS spectrum of the C 1s region, which could be divided into four peaks with a BE of 284.5 eV (C–C/C=C), 285.6 eV (C–N), 286.6 eV (C–O/C=O), and 290.2 eV (O–C=O), respectively.²⁹ According to this spectrum, we could conclude that the oxygenic functional groups showed much weaker relative intensities than GO, indicating the high reduction of GO. The Raman spectra and XPS spectra results discussed above and XRD (shown in S2†) indicated that we prepared the graphene/PANI composite on the surface of carbon cloth successfully.^{30,31}

The electrochemical performances of the as-prepared GP@cc, RGO@cc and PANI@cc composites were compared with each other, as shown in Fig. 3. Fig. 3a illustrates the CV curves of the RGO@cc, PANI@cc and the GP@cc composites within a potential range from 0 to 0.8 V at a scan rate of 5 mV

s^{−1}. The specific capacitance of these samples could be represented by the area of closed CV curves. According to Fig. 3a, the GP@cc electrode showed much larger specific capacitance than those of RGO@cc and PANI@cc electrodes, which could be attributed to the favorable connection with each other leading to excellent conductivity. Fig. 2b shows the GCD tests within a potential range from 0 to 0.8 V at a current density of 0.5 A g^{−1}, and indicated that the GP@cc electrode had the longest discharge time. The specific capacitances of RGO@cc, PANI@cc and GP@cc could be calculated from GCD curves. The specific capacitances of RGO@cc, PANI@cc and GP@cc were measured to be 86, 249 and 793 F g^{−1} at a current density of 0.5 A g^{−1}, respectively, indicating that the GP@cc electrode provided much larger specific capacitance than those of RGO@cc and PANI@cc, which was consistent with the results of CV curves. Furthermore, we calculated the specific capacitances of the GP@cc, RGO@cc and PANI@cc electrodes at different current densities, as shown in Fig. 3c. The specific capacitances of the GP@cc electrode at various current densities were all much larger than those of the others. Moreover, the GP@cc electrode retained a specific capacitance of 444 F g^{−1} even though the current density increased to 5 A g^{−1}. This result indicated its excellent rate capability, which was due to the high conductivity of carbon cloth and the excellent contact between carbon cloth, PANI and graphene. Electrochemical impedance spectroscopy is a very useful approach to evaluate various kinds of resistance during charge/discharge. Fig. 3d shows the Nyquist plots of RGO@cc, PANI@cc and GP@cc electrodes within the frequency range from 100 kHz to 0.05 Hz. In general, a Nyquist plot is composed of a semicircle in a high-frequency region corresponding to charge transfer resistance (*R*_{ct}) and a straight line in a low-frequency region associated with capacitive behavior.³² The points extrapolating the straight line to intersect *Z'* axes are assigned to electrolyte solution resistance, intrinsic resistance of active materials, and the interfacial contact resistance between active materials and the current collector.^{33–35} According to the magnified plots (inset of Fig. 3d), the point of GP@cc was much nearer *−Z''* axes than those of RGO@cc and PANI@cc, indicating the excellent contact between RGO, PANI and carbon cloth. Furthermore, *R*_{ct} of GP@cc corresponding to the diameter of an inconspicuous loop was much smaller than that of RGO@cc and PANI@cc, indicating that the synergistic effect of RGO and PANI improved the conductivity of GP@cc. Moreover, the slope of the low-frequency region for GP@cc was much larger than those of RGO@cc and PANI@cc, suggesting faster ion diffusion of the electrolyte. On the basis of the results discussed above, we can conclude that the GP@cc electrode had much larger conductivity and better ion diffusion behavior than those of RGO@cc and PANI@cc, leading to its excellent electrochemical performances.

To further study the advantage of the flexible and binder-free GP@cc electrode for SC application, we used the GP@cc composite as the working electrode directly in a three-electrode system to evaluate its electrochemical properties. Fig. 4a shows the CV curves of the GP@cc electrode, which exhibited nearly rectangular shapes at scan rates from 5 to 200 mV s^{−1}, which indicated that GP@cc possessed low resistance and ideal



Fig. 3 Electrochemical characterization of RGO@cc, PANI@cc and GP@cc electrodes in a 1 M H₂SO₄ electrolyte, three-electrode system. (a) CV curves at a scan rate of 5 mV s^{−1}. (b) Galvanostatic charge/discharge curves at a current density of 1 A g^{−1}. (c) Specific capacitances versus current densities. (d) Nyquist plots with inset showing close-up views of the high-frequency regime.





Fig. 4 Electrochemical performance of the GP@cc composite. (a) CV curves at different scan rates in a potential range from 0 to 0.8 V and (b) Galvanostatic charge/discharge curves of the GP@cc composite at various current densities in 1 M H₂SO₄ aqueous solution in a three-electrode system; (c) specific capacitance at different current densities; (d) capacitance retention over 10 000 cycles at 5 A g⁻¹.

supercapacitive properties. In addition, the transition of PANI leads to the appearance of two redox peaks in CV curves.³⁶ The CV curves of RGO@cc and PANI@cc at various scan rates are shown in Fig. S3 and S4,[†] respectively. Fig. 4b displays the GCD curves of the GP@cc electrode at different current densities. The GCD curves of rGO@cc and PANI@cc are shown in Fig. S5 and S6,[†] respectively. The discharge curves are nearly symmetrical to the relative charge curves, indicating the well-deserved combination between pseudocapacitive properties with a double-layer contribution. The specific capacitances of the GP@cc electrode at different current densities were calculated from the GCD curves, as shown in Fig. 4c. The specific capacitance of GP@cc was as high as 793 F g⁻¹ at a current density of 0.5 A g⁻¹, which is much larger than those of other graphene/PANI composites reported previously.^{30,37–42} Moreover, about 56% of capacitance was retained as the current density increased to 5 A g⁻¹, indicating its favorable rate capacity. In addition, the cycle life is a significant factor of SCs for practical applications. The cycle life of the GP@cc electrode was measured by repeated GCD tests in a voltage range from 0 to 0.8 V at a current density of 5 A g⁻¹. The results are displayed in Fig. 4d. Its specific capacitance retained about 81% after 10 000 cycles, which indicated that the GP@cc electrode exhibited excellent cycling stability. The remarkably improved electrochemical performance originates in the special electrode design, which avoids the use of binders and other conductive additives, thus enhancing the contact between graphene, carbon cloth and PANI. In addition, the use of carbon cloth greatly improves the conductivity of the composite.

The flexible and binder-free symmetric SC device was fabricated using two pieces of GP@cc composites with 1.0 M H₂SO₄ as the electrolyte, and a digital picture of the flexible symmetric SC device is shown in Fig. S7.[†] The CV curves for a single symmetric device maintained the quasi-rectangular shape at

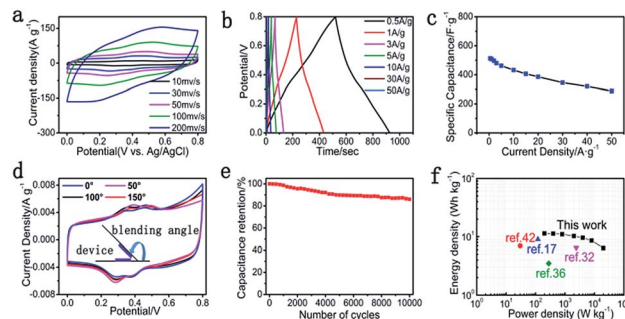


Fig. 5 (a) CV and (b) GCD curves of the symmetric SC in 1.0 M H₂SO₄ electrolyte in a two-electrode system. (c) Specific capacitances at various current densities. (d) CV curves of the GP@cc-based flexible SC at different bending angles. (e) Cycling behavior of the SC device at a current density of 5 A g⁻¹. (f) Ragone plot of the SC device.

various scan rates from 10 to 500 mV s⁻¹, as shown in Fig. 5a. The GCD curves of the device shown in Fig. 5b display a straight line and symmetric shape within a potential window from 0 to 0.8 V at different current densities from 0.5 to 50 A g⁻¹. These characteristics suggest the ideal capacitance of the symmetric device. According to Fig. 5c, the specific capacitance of the symmetric device achieved 512 F g⁻¹ at a current density of 0.5 A g⁻¹. Moreover, the device still retained 56.3% of the initial specific capacitance as the current density increased to 50 A g⁻¹, indicating its favorable rate capacity. Furthermore, the GP@cc-based flexible symmetrical SC also showed high flexibility in bending measurements. Fig. 5d illustrates the CV curves of the device at different bending angles. The CV curve of the device showed little change even as the bending angle increased to 150°, indicating that the flexible SC based on GP@cc composite exhibited extraordinary flexibility and favorable mechanical properties even in a highly bent state. Fig. 5e illustrates the cycling behavior of the symmetric device at a current density of 5 A g⁻¹. The symmetric device still retained about 85% of the initial specific capacitance after 10 000 cycles, indicating its excellent cycle stability. This can be ascribed to the unique electrode design of the flexible, binder-free GP@cc composite. Fig. 5f is a Ragone diagram of the symmetric device. The energy and power densities of the symmetric device were calculated from the GCD curves. According to Fig. 5f, the symmetric device exhibited a maximum energy density of 11.38 W h kg⁻¹ at a power density of 199.80 W kg⁻¹, and a maximum power density of 20 000 W kg⁻¹ at an energy density of 6.4 W h kg⁻¹ at an open-circuit voltage of 0.8 V. This work shows a higher energy density than those of devices reported previously,^{17,32,36,42} which can be attributed to the synergistic effect of all components in GP@cc composite materials.

4. Conclusions

We successfully prepared a flexible and binder-free GP@cc composite using a facile and scaled-up one-step electrochemical co-deposition method, and studied its practical application in a flexible symmetric SC. Owing to the synergistic effects of different components in the GP@cc composite material and the



high flexibility and excellent conductivity of carbon cloth, the symmetric SC device based on a flexible, binder-free GP@cc electrode exhibited acceptable energy density, remarkable flexibility and excellent cycle stability. These features endow its great potential to be applied in smart, wearable and portable electronic products.

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Notes and references

- 1 J. M. Tarascon and M. Armand, *Nature*, 2001, **414**, 359.
- 2 F. Beguin, K. Szostak, G. Lota and E. Frackowiak, *Adv. Mater.*, 2005, **17**, 2380.
- 3 Y. Xu, Z. Lin, X. Huang, Y. Wang, Y. Huang and X. Duan, *Adv. Mater.*, 2013, **25**, 5779.
- 4 P. Simon and Y. Gogotsi, *Nat. Mater.*, 2008, **7**, 845.
- 5 M. Kaempgen, C. K. Chan, J. Ma and Y. Cui, *Nano Lett.*, 2009, **9**, 1872–1876.
- 6 H. Chen, Y. Li, Y. Feng, P. Lv, P. Zhang and W. Feng, *Electrochim. Acta*, 2012, **60**, 449–455.
- 7 D. Pech, M. Brunet, H. Durou, P. Huang, V. Mochalin, Y. Gogotsi, P. L. Taberna and P. Simon, *Nat. Nanotechnol.*, 2010, **5**, 651–654.
- 8 F. C. Wu, R. L. Tseng, C. C. Hu and C. C. Wang, *J. Power Sources*, 2004, **138**, 351–359.
- 9 C. Kim, Y. O. Choi, W. J. Lee and K. S. Yang, *Electrochim. Acta*, 2004, **50**, 883–887.
- 10 L. L. Zhang and X. S. Zhao, *Chem. Soc. Rev.*, 2009, **38**, 2520–2531.
- 11 K. Wang, W. Zou, B. Quan, A. Yu, H. Wu, P. Jiang and Z. Wei, *Adv. Energy Mater.*, 2011, **1**, 1068–1072.
- 12 X. Xiao, T. Li, P. Yang, Y. Gao, H. Jin and W. Ni, *ACS Nano*, 2012, **6**, 9200–9206.
- 13 B. G. Choi, S. J. Chang and H. W. Kang, *Nanoscale*, 2012, **4**, 4983–4988.
- 14 Y. J. Kang, S. J. Chun, S. S. Lee and B. Y. Kim, *ACS Nano*, 2012, **6**, 6400–6406.
- 15 B. G. Choi, J. Hong, W. H. Hong, P. T. Hammond and H. Park, *ACS Nano*, 2011, **5**, 7205–7213.
- 16 D. Tobjork and R. Osterbacka, *Adv. Mater.*, 2011, **23**, 1935–1961.
- 17 Y. F. Wang, X. W. Yang, L. Qiu and D. Li, *Energy Environ. Sci.*, 2013, **6**, 477–481.
- 18 Q. Q. Zhang, Y. Li, Y. Y. Feng and W. Feng, *Electrochim. Acta*, 2013, **90**, 95–100.
- 19 Q. Wu, Y. X. Xu, Z. Y. Zhao, A. R. Liu and G. Q. Shi, *ACS Nano*, 2009, **3**, 1745–1752.
- 20 H. Y. Yu, J. H. Wu, L. Q. Fan, K. Q. Xu, X. Zhong, Y. Z. Lin and J. M. Lin, *Electrochim. Acta*, 2011, **56**, 6881.
- 21 W. S. Hummers and R. E. Offeman, *J. Am. Chem. Soc.*, 1958, **80**, 1339.
- 22 M. Zhou, Y. L. Wang, Y. M. Zhai, J. F. Zhai and W. Ren, *Chem.–Eur. J.*, 2009, **15**, 6116–6120.
- 23 Y. Y. Shao, J. Wang and M. Engelhard, *J. Mater. Chem.*, 2010, **20**, 743–748.
- 24 K. X. Sheng, Y. Q. Sun, C. Li, W. J. Yuan and G. Q. Shi, *Sci. Rep.*, 2012, **2**, 247.
- 25 Y. R. Li, K. X. Sheng, W. J. Yuan and G. Q. Shi, *Chem. Commun.*, 2013, **49**, 291–293.
- 26 H. L. Wang, Q. L. Hao, X. J. Yang, L. D. Lu and X. Wang, *Nanoscale*, 2010, **2**, 2164–2170.
- 27 Y. Zhao, H. Bai, Y. Hu, Y. Li, L. T. Qu, S. W. Zhang and G. Q. Shi, *J. Mater. Chem.*, 2011, **21**, 13978–13983.
- 28 M. G. Han, S. K. Cho, S. G. Oh and S. S. Im, *Synth. Met.*, 2012, **126**, 53.
- 29 L. Tang, Y. Wang, Y. Li, H. Feng, J. Lu and J. Li, *Adv. Funct. Mater.*, 2009, **19**, 2782.
- 30 X. M. Feng, R. M. Li, Y. W. Ma, R. F. Chen, N. E. Shi, Q. L. Fan and W. Huang, *Adv. Funct. Mater.*, 2011, **21**, 2989–2996.
- 31 T. Lee, T. Yun, B. Park, B. Sharma, H. K. Song and B. S. Kim, *J. Mater. Chem.*, 2012, **22**, 21092–21099.
- 32 L. Wang, Y. J. Ye, X. P. Lu, Z. B. Wen, Z. Li, H. Q. Hou and Y. H. Song, *Sci. Rep.*, 2013, **3**, 3568.
- 33 F. Huang and D. Chen, *Energy Environ. Sci.*, 2012, **5**, 5833–5841.
- 34 M. D. Stoller and R. S. Ruoff, *Energy Environ. Sci.*, 2010, **3**, 1294–1301.
- 35 M. Hughes, M. S. P. Shaffer, A. C. Renouf, C. Singh and G. Z. Chen, *Adv. Mater.*, 2002, **14**, 382–385.
- 36 Y. G. Wang, Q. H. Li and Y. Y. Xia, *Adv. Mater.*, 2006, **18**, 2619–2623.
- 37 S. P. Zhou, H. M. Zhang, Q. Zhao, X. H. Wang, J. Li and F. S. Wang, *Carbon*, 2013, **52**, 440–450.
- 38 J. W. Luo, W. B. Zhong, Y. B. Zou, C. L. Xiong and W. T. Wang, *J. Power Sources*, 2016, **319**, 73–81.
- 39 Z. Niu, L. Liu, L. Zhang, Q. Shao, W. Zhou, X. Chen and S. Xie, *Adv. Mater.*, 2014, **26**, 3681–3687.
- 40 F. Yang, M. Xu, S. J. Bao, H. Wei and H. Chai, *Electrochim. Acta*, 2014, **137**, 381–387.
- 41 Z. Tai, X. Yan and Q. Xue, *J. Electrochem. Soc.*, 2012, **159**, A1702–A1709.
- 42 Y. F. Xu, M. G. Schwab, A. J. Strudwick, I. Hennig, X. L. Feng and Z. S. Wu, *Adv. Energy Mater.*, 2013, **3**, 1035–1040.

