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Advances in the synthesis of Fe-based bimetallic electrocatalysts for CO₂ reduction

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Achieving carbon neutrality and slowing down global warming requires research into the electrochemical CO_2 reduction reaction (CO_2RR), which produces useful compounds. Utilizing renewable energy to meet carbon-neutral energy goals produces single-carbon (C_1) and multi-carbon (C_{2+}) goods. Efficient and selective electrocatalysts are essential to advancing this revolutionary technology; bimetallic Fe-based catalysts work better than their monometallic counterparts because multiple metals work synergistically to reduce CO_2 levels. A thorough summary of recent developments in the synthesis of Fe-X bimetallic catalysts will be provided in this review, with an emphasis on key performance indicators like stability, faradaic efficiency, potential, current density, and primary product production. In addition, this analysis will look at representative instances of Fe bimetallic catalysts that are well-known for their selectivity in generating particular alcohols and hydrocarbons, clarifying the mechanics behind CO_2 reduction, pointing out existing difficulties, and examining the potential of electrosynthesis processes in the future.

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1. Introduction

Over the past 170 years, human activity has caused a sharp rise in CO₂ emissions, which has resulted in ocean acidification and global climate change.^{1,2} Terrestrial ecosystems only absorb around 30% of CO₂ produced by humans annually, which is insufficient to offset anthropogenic emissions, and the amount



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of CO_2 in the atmosphere has dramatically increased in recent decades, reaching 400 ppm for the first time in human history. Since the late 1950s, the Mauna Loa Observatory in Hawaii has been continually monitoring atmospheric CO_2 , and as of 2024,



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CO₂ concentrations are approximately 417 ppm. One of the main causes of climate change, this represents a sharp rise from pre-industrial levels of about 280 ppm.³⁻⁵ According to predictions, if present emission trends continue, CO₂ levels might surpass 450 ppm by 2030. If immediate mitigation measures are not taken, some models forecast significantly higher concentrations.⁶ Over 450 ppm could dramatically raise the probability of catastrophic climate consequences, including more severe and frequent heatwaves, rising sea levels, and disturbances to ecosystems and food security.⁷

Although CO_2 is a very stable molecule that is usually inert, it can undergo electrochemical activation and be transformed into reduced products through the CO₂ reduction reaction (CO₂RR) with the help of protons in solution and appropriate cathodic reduction potentials.8,9 Several methods have been suggested to change CO2 into value-added products, such as chemical transformation,^{10,11} reduction by photocatalysis,¹²⁻¹⁴ reduction by electrocatalysis,15,16 thermal catalysis,17,18 photothermal catalysis,19 as well as biological conversion.20,21 The simplicity, mild reaction conditions, environmental compatibility, and possible integration with energy from renewable sources of electrocatalytic CO2 reduction (eCO2RR) make it stand out among these applications. Since CO₂ has a linear, symmetrical structure with a zero-dipole moment, it is stable and challenging to activate in electrocatalytic reduction. Because of the molecule's symmetry, the opposing dipoles of the C=O bonds cancel each other out, resulting in a low electron density around the molecule's center, especially close to the carbon atom. The molecule's general lack of a dipole moment is a result of the decreased electron density surrounding the carbon center. It is more difficult to activate CO2 due to its low electron density. Effective electrocatalytic reduction requires the catalyst to capture and hold CO₂ molecules long enough for the reaction. However, CO2 lacks a strong dipole and other reactive properties, making it more difficult to activate for the reduction process. This demands extremely effective catalysts that can change CO2's electron density or aid in breaking its strong C=O bonds.22,23 A significant overpotential is needed for CO2 activation due to the dissociation energy needed to break the C=O bond, which is more than 750 kJ mol⁻¹, leading to poor energy efficiency and high operational costs.²⁴ The shale gas revolution of the last twenty years, especially the large amounts of ethane (3-12% fraction) from shale gas, has changed the dynamics of the global energy market and produced an excess of ethane despite its relatively low market price, which is especially noticeable in the United States.^{25,26} Studies on CO₂ reduction for CO and carbon-based energy sources²⁷⁻²⁹ have gained increasing attention owing to rising atmospheric CO2 levels and expanding energy demands, with a focus on finding inexpensive, efficient, and selective catalysts.30-32 These catalysts include homo-bimetallic sites (Fe-Fe, Co-Co, Ni-Ni, Cu-Cu) that indicate enhanced reactivity in comparison to monometallic counterparts, while heterobimetallic catalysts remain relatively overlooked.33-41

Efficient CO₂ reduction catalyst design is hard due to stability, huge potentials,^{29,42} restricted solubility, competing HER, and slow kinetics.^{43–46} Reduced activation barriers are

critical for improving electrocatalyst efficiency and selectivity. Despite the beneficial features of iron-group metallic alloys and compounds, their efficacy as catalysts in CO2 reduction remains insufficient due to low activity and stability.47,48 Because Fe is so readily available, it is essential to build highly efficient Fe-N-C catalysts. With the abundance of iron, there is a need to produce these catalysts. Fe-porphyrins treated with phenolic groups showed remarkable CO faradaic yields exceeding 90% without degradation, emphasizing Fe-N4 sites in macrocycles as active centers.49 The incorporation of Fe atomically into nitrogendoped carbon substrates, such as Fe-N-C catalysts, has demonstrated remarkable catalytic reactivity towards CO₂ reduction to CO, with Fe-N4 sites largely studied as active sites in several investigations.⁵⁰⁻⁵³ Fe species provide dynamic surface manipulation, which is critical for understanding structure dynamics and rational catalyst designing in CO₂ electroreduction reaction (CO2ERR).54-56 Recent studies reveal that heteroatom inclusion in carbon support alters the electrical environment, allowing tailored Fe sites to lower the energy of activation limitations in electrocatalysis.57 Fe or Cu-based metals/alloys are widely used as catalysts in CO2 reduction; Fe has significant catalytic activity and a minimal energy barrier, whereas Cu has excellent CO2 adsorption traits and resistance to coking.58-61 In addition to heteroatom inclusion, the use of a second metal atom in Fe-based materials also termed dualatomic catalysts (DACs) improves catalytic activity synergistically. CO2 reduction relies on DACs, while Ru, Fe, Mn-based homogeneous, and Cu-based heterogeneous catalysis provide viable alternatives.62-72 The adsorbate-metal surface interaction in Fe-N-C single-atom catalysts (also known as SACs) is influenced by the shift of the d-band center.73 Determining the intensity and kind of these interactions between molecules is largely dependent on this change. Consequently, it has a major effect on the catalytic activity. Integrating heteronuclear metal atoms such as Ni, Co, and Zn permits electronic structure adjustment, which facilitates adsorbate absorption as well as desorption on the surface of the catalyst.74-76 Recently, multiple reviews have investigated the CO2ERR, spanning diverse catalysts like as copper-palladium nanoalloys,9 Cu-based nanocrystals,77 bimetallic chalcogenides,78 bimetallic catalysts with atomic sites,⁷⁹ Bi-based,⁸⁰ Ni-based,⁸¹ Sn-based,⁸² and carbidebased bimetallic catalysts.83 Fe-based bimetallic electrocatalysts are superior to other metals because they are more affordable, widely available, and have better selectivity for CO2 reduction products. When it comes to stability, efficiency, and scalability, they can perform better than single-metal catalysts like Cu. There is currently no review that provided in-depth analysis of the research on Fe-based bimetallic electrocatalysts for CO2 reduction. Given the growing importance of electrolytic CO2 reduction, the performance of Fe-based bimetallic catalysts merits a thorough examination. This research focuses on the production, implementation, and mechanistic understanding of these catalysts in CO₂ electrocatalysis, covering a wide spectrum of product forms. Furthermore, the study highlights the problems and opportunities in developing and comprehending Fe-based bimetallic electrocatalysts, which offer useful insights for future research paths in this sector.



Fig. 1 (A) Proposed paths for reduction of CO_2 into CO over Fe/Ni–N–C. (B) Using DFT, the electrochemical reduction of CO_2 to CO on Fe–N–C, Ni–N–C, and Fe/Ni–N–C with and without adsorbed *CO was represented by a free energy diagram.

2. CO₂ reduction pathways over Febased bimetallic electrocatalysts

Fe/Ni-N-C materials were used as electrocatalytic reduction (ECR) catalysts by Huiying Tian and coworkers. These substances were utilized to speed up the electrochemical processes that produce CO.84 Fig. 1A illustrates the CO2 reduction into CO over the Fe/Ni-N-C catalyst. For the catalytic reduction of CO2 into CO, the Fe/Ni-N-C catalyst provides sites for CO2 to bind, and Fe/Ni acts as active centers for the reduction. For binding, CO2 accepts electrons and protons from the electrolyte solution and converts them into intermediate CO precursors such as CHOOH. This reaction is typically completed in two steps $CO_2 \rightarrow COOH \rightarrow CO^*$ as mentioned in Fig. 1. The Fe/Ni-N-C electrocatalyst achieved an impressive (faradaic efficiency of CO) FE_{CO} of 92.9% at -0.677 V vs. (reversible hydrogen electrode) RHE, indicating great efficiency. The system retained an elevated current density and faradaic efficiency for the generation of CO (FE_{CO}) when applied in a continuous flow cell at scale, holding onto over 89% shortly after 40 hours of electrolysis. Because of the binary metals combined effect, charge transfer rates were increased, resulting in favorable kinetics and long-term, effective electrochemical performance.57

To thoroughly investigate the combined effect of Ni-N-C, Fe-N-C, and Fe/Ni-N-C on ECR, DFT studies were performed utilizing the computational hydrogen electrode technique.85 In electrocatalytic CO2 reduction, the Fe atom's electrical properties and catalytic action are largely determined by its spin orientation.86,87 Fe can exist in both high and low spin states in bimetallic Fe-based catalysts, which affects the electron distribution in the d-orbitals and changes the CO₂ and intermediate adsorption strength.⁸⁸ While a low-spin level may produce more stable, less reactive configurations, a high-spin state can increase the activation of CO₂ by offering more accessible electron states.89 The impact of these spin arrangements on the CO₂ reduction reaction mechanism and efficiency is clarified using DFT simulations. The major catalytic sites, Me-N₄ motifs, were used as single-site models based on prior investigations.^{51,90} The electrocatalytic reduction (ECR) process involved typical two-electron and two-proton transfer reactions,

culminating in the creation of *COOH and *CO intermediates. The symbol asterisk (*) represents the active site. Ni-N-C, Fe-N-C, and Fe/Ni-N-C optimized geometries served as computational models for the investigation.91,92 On Fe-N-C sites, the rate-determining step was $*CO \rightarrow CO_{(g)}$, while on Ni–N–C sites, it was $CO_{2(g)} \rightarrow *COOH$. Fe/Ni-N-C generated *COOH intermediates easily, while *CO desorption was difficult. Fe/Ni-N-C adsorbed with *CO intermediates had a much smaller free energy shift for the rate-determining phase $*CO \rightarrow CO_{(g)}$, indicating easier desorption. This shows that Fe/Ni-N-C provides more active sites by efficiently combining the benefits of Ni-N-C and Fe-N-C sites, increasing CO generation catalytic activity. CO2 adsorption on Fe-Ni bimetallic sites, electron and proton transfer pathways to form *COOH within *CO intermediate, and subsequent CO(g) desorption to regenerate Fe-Ni-N-C (*CO) are the suggested ECR reaction routes on Fe/Ni-N-C. This highlights the increased catalytic activity seen in the studies. The Fe, Ni bimetallic nitrogen-doped carbon successfully lowered the energy barriers of *COOH intermediate production and *CO-to-CO, improving favorable kinetics and increasing ECR activity, verified by the DFT calculations. According to the findings of the calculations above, the suggested ECR reaction pathways of CO2 to CO on Fe/Ni-N-C and their energy diagram using DFT calculation are shown in Fig. 1.⁸⁴

3. Bimetallic graphene catalysts: mechanistic pathways for CO₂ reduction and CH₄ production

Previous studies have shown that single-atom-doped graphene is exceptionally efficient for catalyzing CO₂RR.⁹³⁻⁹⁶ Researchers have studied diverse doping techniques for adding transition metals to the graphene, demonstrating that bimetal singleatom-doped catalysts have greater catalytic performance than standard single-atom-doped catalysts.⁹⁷⁻¹⁰⁰ Run Zhang *et al.*, employed DFT calculations for examining CO₂RR on three bimetal-doped graphene catalysts, Cu–Ni/DG, Cu–Fe/DG, and Fe–Ni/DG. Different reduction pathways yield various products

Table 1 Summary of adsorption energy of different reaction intermediates for the production of CH₄ on Cu–Fe/DG and Fe–Ni/DG catalysts

	Potential energy	(eV)
Reaction intermediates	Cu–Fe/DG	Fe–Ni/DG
*CO	-1.7	-2.81
*HCOOH	0.78	-0.27
*CH ₂ O	-1.46	-2.05
*CH ₃ OH	0.97	-0.44
$*CH_4$	0.70	-0.76

such as CH₄, CH₃OH, HCOOH, and CO. In the initial stages of CO₂RR on doped graphene, CO₂ adsorption occurs, characterized by analysis of E_{ads} , electron density difference, density of state (DOS), as well as Integrated Crystal Orbital Hamiltonian Population (ICOHP). Compared to Cu–Fe/DG, CO₂ reacts more strongly with iron-based Fe–Ni/DG. The catalytic performance of the material is improved by this greater contact. Electron density difference, DOS, and ICOHP studies reveal more robust interactions between certain dopants (Fe and Ni) and CO₂.^{84,101} Table 1 summarizes the adsorption energies of process intermediates on bimetal-doped Fe catalysts. When CO₂ is first protonated, it produces *COOH or *OCHO, which can then be hydrogenated again to generate CO or HCOOH. These changes proceed in many ways:

$$COOH \rightarrow *CO + H_2O \rightarrow * + CO$$

and

OCHO \rightarrow *HCOOH \rightarrow * + HCOOH

On Cu–Fe/DG, Cu–Ni/DG, and Fe–Ni/DG catalysts, the high adsorption of CO and HCOOH encourages continued reduction as intermediates. Nevertheless, significant free energy barriers prevent CO and HCOOH from being desorbed from the catalyst's surface, which presents problems for product release. In the CO₂ reduction reaction, CH₃OH is a potential product. Four pathways for CO₂ \rightarrow CH₃OH involve *CO or *HCOOH as intermediates. *CO undergoes the following reactions:

 $*CO \rightarrow *CHO \rightarrow *CH_2O$ (exothermic)

or

$$*CHOH \rightarrow CH_3OH$$

*CHO is exothermically converted to $*CH_2O$ on Cu–Ni/DG and Cu–Fe/DG. On the other hand, *CHO into *CHOH conversions on Fe–Ni/DG are endothermic.

*CHO \rightarrow *CHOH (endothermic)

This distinction draws attention to the different energetics of different catalysts for these reactions.

The thermodynamic favorability of $*CH_2O$ formation is highlighted by its reduced variance in free energy, which is why it is preferred over *CHOH formation. This preference highlights the role that energetics play in identifying the paths of reactions. Six routes for $CO_2 \rightarrow CH_4$ reduction were investigated, using *CO or *HCOOH as intermediaries. *COundergoes the following conversions:

$$*CO \rightarrow *CHO$$

or

$$*CO \rightarrow *COH$$

It resulted in *CHO or *COH and finally CH₄ by additional hydrogenation. *CO prefers *CHO production because of the lower free energy fluctuation. On Cu–Ni/DG and Cu–Fe/DG,

*CHO
$$\rightarrow$$
 *CH₂O (exothermic)
*CHO \rightarrow *CHOH (endothermic)

While, Fe-Ni/DG prefers,

$$*CHO \rightarrow *CH_2O$$

Path 3 is the best route from Path 1 to Path 5, demonstrating its effectiveness and fit for the intended change. Nevertheless, route 6 (*OCHO \rightarrow *HCOOH) doesn't work as the best route for Fe–Ni/DG due to significant free energy fluctuation. The optimized pathway is Path 6, which is exothermic on Cu–Fe/DG and Cu–Ni/DG.¹⁰¹

$$*CHO \rightarrow *CH_2OH \rightarrow *OCH_3 \rightarrow *CH_4$$
 (exothermic)

Because CO_2 interacts with Fe or Ni atoms more strongly than its interaction with Cu, Fe–Ni/DG is more stable than Cu– Ni/DG and Cu–Fe/DG. The catalytic potential of Fe–Ni/DG for CO_2 conversion reactions is highlighted by its improved stability. Fe–Ni/DG becomes more prominent in bimetal-doped graphene systems because of its increased stability, which implies that it can support effective and long-lasting catalytic activity. Graphene doped with Cu, Fe, and Ni shows significant selectivity for CO_2 reduction over hydrogen evolution (HER), suggesting that these materials are effective catalysts for CO_2 conversion processes, with various product outcomes seen for the initial protonation step of CO_2 on these catalysts.

4. FeCo-Pc catalysts for multi-carbon (C₂) product formation

FeCo-Pc catalyst with dual metal–nitrogen active sites for efficient CO₂RR. FeCo-Pc overcomes the challenge of C–C coupling seen in single-atom catalysts,^{102,103} enabling the production of

Table 2 Summary of reaction intermediates generates during the formation of C_{2+} products with their possible potential energy over FeCo-Pc catalyst

Compound	Reaction intermediate	Potential energy (eV)
C ₂ H ₄	COOH*	0.13
2 1	CO*	0.23
	CHO*	0.68
	COCHO*	0.70
	CHOCHO*	0.38
	CHOHCHO*	0.55
	CHCHO*	0.27
	CH_2CHO^*	0.17
	CH ₂ CHOH*	0.06
	CH_2CH^*	0.51
C_2H_5OH	CHOHCH ₂ O*	0.41
	CHCH ₂ O*	0.21
	$CH_2CH_2O^*$	0.40
	$C_2H_2OH^*$	0.71
CH ₃ COOH	COCHOH*	0.57
	$COCH_2O^*$	1.59
	COHCH ₂ O*	0.26
	CH ₂ COOH*	0.60
CH ₂ OHCH ₂ OH	CHOHCH ₂ OH*	0.16
	CH ₂ OHCH ₂ OH*	0.20

 C_2 products, as shown in Table 2. These C_2 products include C₂H₄, CH₂OHCH₂OH, C₂H₅OH, and CH₃COOH with enhanced selectivity, due to the cumulative effects of Fe and Co dual active sites anchored within phthalocyanine (FeCo-Pc).104 The computations are performed using the Gaussian 09 program, PBE exchange-correlation functional, and 6-31G* basis sets.105,106 C-C coupling processes in CO2 reduction are critical for comprehending multi-carbon product generation.107,108 Water plays an important role in the reduction of CO₂ because it influences intermediate hydration, provides protons for product production, and functions as a solvent for ion transport. Additionally, through the oxygen evolution reaction (OER), it competes with CO2 reduction and affects catalyst behavior.^{109,110} CO was identified as a crucial intermediate, discovered by in situ spectroscopy.^{111,112} PCETs generate C₁ intermediates like CHO* and COH*, with thermodynamics and kinetics assessed on FeCo-Pc surfaces. CO to CHO* (formyl group) is produced by further reducing *CO and adding an extra proton and electron. Usually, this process produces more valuable chemicals such as alcohols and aldehydes. In the case of COH* protonation step is required for the formation of COH*, but not the complete reduction required to generate *CHO. In the synthesis of other C1 products, such as ethanol or



Scheme 1 Possible electrochemical reaction pathways of CO_2 over Fe-based bimetallic catalyst into C_{2+} products.

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methane, this is frequently a transitional stage. CO dimerization, proposed as the initial step to C₂₊ products, and carbene (CH_2^*) coupling with CO* to form COCH₂^{*} or CH₂CO*, are investigated.42,113 However, the generation of "dead-end" intermediates, such as COCO* and other C-C linked intermediates, tightly bound to FeCo-Pc, incurs high energy consumption, creating kinetic obstacles and surface contamination. Electrolytes C-C coupling pathways are influenced by pH,¹¹⁴ with high pH favoring CO* to COH* conversion and CO* coupling with CHO* to generate COCHO*, boosting C2 synthesis.115,116 COCHO* formation has a lower activation barrier compared to CHO*-CO* precursors, demonstrating its thermodynamic favorability and potential to provide more C2 products in CO2RR on FeCo-Pc surfaces.117 In an aqueous solution, the relative stability of COOH* as well as OCHO* in CO2RR against H* in HER affects the competing processes of the H₂ evolution reaction (HER).¹¹⁸ FeCo-Pc prefers CO₂RR over HER due to larger free energy changes for COOH* and OCHO* production against adsorbed H*. COOH* has a smaller overpotential than OCHO*, indicating FeCo-Pc prefers CO2RR.119,120 The electro-conversion of CO₂ to C₂H₄ is important for the C₂H₄ industry, however, it faces challenges with an elevated overpotential and multielectron transfer processes.121,122

FeCo-Pc catalysts increase CO₂RR by favoring COOH* over H*, resulting in C_2H_4 generation. During CO reduction, CHO* formation takes precedence over COH* formation. The ratelimiting step for C_2H_4 generation is coupling CHO* with CO* to generate COCHO*.²² FeCo-Pc surfaces aid in producing C_2H_4 by reducing COCHO* to glycolaldehyde and hydrogenating further. FeCo-Pc catalysts provide a viable avenue for the electrochemical process to transform CO₂ into ethanol (C_2H_5OH), a critical commodity chemical, *via* C-C coupling reactions. The procedure is optimized by hydrogenating typical intermediates with ethylene (C_2H_4). CHOHCHO* is found as a selectivitydetermining molecule. Thermodynamically, C_2H_5OH is produced through the optimum process of CHOHCHO* hydrogenation to CHOHCH₂O*.^{117,123,124}

The rate-limiting step (RLS) for C₂H₅OH creation involves the hydrogenation process of CH₂CH₂O* to C₂H₄OH*, which has a greater barrier than the formation of C₂H₄. The preference for CHCHO* or CHOHCH₂O* production during CHOHCHO* reduction determines the selectivity of C₂H₄ and C₂H₅OH. Increasing potential increases the feasibility of producing C₂H₄ and C2H5OH on FeCo-Pc, with all fundamental stages downward energetically at -0.66 V-RHE. Previous investigations have discovered ethylene glycol to be a negligible product in CO₂RR utilizing catalysts such as Au, Ru, and Cu.125-127 Calvinho et al., recently proved that CO₂RR may be converted to ethylene glycol (CH₂OHCH₂OH) using a transition-metal phosphide catalyst. CHOHCHO*, like C₂H₄ and C₂H₅OH, determines the selectivity of CH₂OHCH₂OH production. Protonation of CHOHCH₂O* results in the formation of CHOHCH₂OH*, which is preferred over CHCH₂O*. This is then transformed into ethylene glycol. Geometry optimization demonstrates that CH₂OH is not chemisorbed, indicating that it prefers the formation pathway over C2H5OH. Both CH2OHCH2OH and C2H5OH have an identical kinetic barrier for CHOHCH₂O* production, resulting in

 C_2H_4 selectivity.^{104,128-130} Possible electrochemical reduction pathways for CO_2 into C_2 products are shown in Scheme 1.

5. Fe-based bimetallic electrocatalysts: advanced pathways for CO₂ reduction

As previously reported, nitrogen-doped carbon nanotubes with Fe/Fe_3N nanoparticles improve the catalytic performance of the oxygen reduction process (ORR) by exposing active areas and enabling electron transport.^{131–134} Before pyrolysis, Fe-doped zinc-imidazole frameworks (ZIF-8) were changed with phosphomolybdic acid hydrate (PMo), resulting in the formation of Fe nanoparticles contained within molybdenum and nitrogenco-doped carbon scaffolds (Fe-NP/MNCF). In CO₂ electrolysis powered by a Zn–air battery (ZAB), Fe-NP/MNCF served as a dual-functional catalyst during ORR and CO₂RR. To synthesize, $Fe(NO_3)_3 \cdot 9H_2O$, $Zn(NO_3)_2 \cdot 6H_2O$, and PMo were dispersed in 2.5 mL of water that was deionized by ultrasound. The precursor was calcined at 900 °C in an argon environment for 2 hours to produce Fe-NP/MNCF, as illustrated in Fig. 2A.^{135,136}

Nitrogen-doped carbon was obtained by annealing ZIF-8 at 950 °C in an argon environment. The final product, called ZIF-NC, has a three-dimensional porous structure.¹³⁷ FeNi@N-CNTs catalysts were developed by wet impregnation and thermal processing. The heat was applied at 1100 °C to produce FeNi@N-CNTs-*X*, where *X* is the temperature at which they were annealed (Fig. 2B). By encasing the FeNi alloy in N-CNTs, this synthesis method enhances the catalytic properties for potential CO₂ reduction uses. Better CO₂ER activity and stability were demonstrated by FeNi@N-CNTs-1100, which showed over 90% CO faradaic efficiency spanning a wide potential range (-0.47 to -0.97 V ν s. RHE). Optimized *COOH adsorption and *CO desorption were credited with improving catalytic activity and CO selectivity.^{138,139}

Xiao Han et al. employed a solution approach to create FeNi precursors, which were subsequently transformed into a variety of FeNi-NC catalysts via one-step pyrolysis. Dissolved $Ni(NO_3)_2 \cdot 6H_2O$, $Zn(NO_3)_2 \cdot 6H_2O$, and $Fe(NO_3)_3 \cdot 9H_2O$ in 80 mL of methanol and stirred thoroughly. Separately, another 80 mL of methanol was used to dissolve 2-methylimidazole and added to the metal nitrate solution. The resultant mixture was agitated constantly for 8 hours to produce the catalyst precipitate. The resulting precipitate washed away with the solvent methanol centrifuged, evaporated at 60 °C, and powdered to produce Fe₃Ni₇-ZIF samples with different Fe/Ni ratios, as illustrated in Fig. 2C. These catalysts attain about 100% overall Faraday efficiency by promoting CO₂ electroreduction into CO and H₂. Furthermore, we discovered that altering the applied potential across a large range throughout the procedure makes it simple to change the syngas ratio from 1:1 to 6:1 (CO/H₂). Because of its versatility, syngas can be utilized to manufacture fuels and raw materials for chemicals.140

The unique features of doped Cu in Fe–N–C catalysts, including its numerous oxidation states, which facilitate fast electron transfer,^{87,141} ability to particularly manufacture C₂₊



Fig. 2 (A) Synthesis of Fe-NP/MNCF (B) synthesis of FeNi@N-CNTs and (C) synthesis of Fe₃Ni₇-N-C.

products,^{142–144} and improved interaction with CO₂ to limit hydrogen development, have sparked great interest in CO₂RR. The Fe/Cu–N–C catalyst, which was produced by adding a copper promoter to a mixture of iron and carbon sources and then pyrolyzing it, has outstanding CO₂ reduction efficiency with more than 90% CO faradaic productivity (FE_{CO}) in a broad potential range (-0.5 to -0.7 V) and remarkable stability, with FE_{CO} maintained after 10 hours of electrolysis. To make the Fe/ Cu–N–C catalyst, Shulin Zhao *et al.*, mixed tris(2,4-pentanedionato)iron(m), Cu-acetylacetonate, along with meso-tetra(4methoxyphenyl) porphin in CHCl₃ and stirred it at 60 °C for the period of 3 h. Rotational evaporation was used to extract the solvent from the mixture after 30 minutes of sonication following the addition of zinc oxide. After the powder was produced, it was heated to 900 °C in an argon atmosphere for two hours, then it was leached for six hours at 80 °C in 0.5 M H₂SO₄ and allowed to dry overnight.¹⁴⁵ Table 3 shows the comparative analysis of heteronuclear Fe-based catalysts for CO₂ electroreduction.

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	parative analysis of neteronuclear Fe-bas	ed catalysts for UO_2 electroreduction			
Catalyst	Synthesis method	Key features	CO ₂ reduction efficiency	Stability	Reference
Fe-NP/MNCF	Fe-doped ZIF-8 modified with PMo, pvrolvsis at 900 °C	Molybdenum and nitrogen co-doped carbon scaffold	The dual-functional catalyst for ORR and CO,RR	Used in Zn-air battery-powered CO ₂ electrolysis	135 and 136
ZIF-NC	Annealing ZIF-8 at 950 °C	Three-dimensional porous structure		, I	137
FeNi@N-CNTs- X	Wet impregnation and thermal processing at 1100 °C	Encapsulated FeNi alloy in N-CNTs, enhanced CO ₂ reduction	>90% CO faradaic efficiency (-0.47 to -0.97 V vs. RHE)	High stability	138 and 139
Fe ₃ Ni ₇ -ZIF	Solution approach, pyrolysis	Various Fe/Ni ratios to adjust performance	\sim 100% faraday efficiency, tunable syngas ratio (1:1 to 6:1 CO/H $_2$)	High stability, broad potential range	140
Fe/Cu-N-C	Cu promoter added to Fe/carbon mixture, pyrolysis at 900 °C, acid leaching	Enhanced electron transfer, improved CO2 adsorption	>90% CO faradaic efficiency (–0.5 to –0.7 V)	Stable after 10 hours of electrolysis	145
C-Fe-Co-ZIF	Impregnation of ZIF-8 with Fe and Co, pyrolysis	Bimetallic Co-Fe catalyst for CO ₂ electroreduction	+10% CO faradaic efficiency <i>vs.</i> pure Co- ZIF	H_2 /CO ratios tunable (0.8 to 4.2), 93% FE _{CO} + H ₂ over 10 hours	146
Fe/Mn-N-C	Potassium citrate calcination, Fe and Mn doping, pyrolysis at 800 °C	Atomic dispersion of Fe and Mn for CO selectivity	94% CO faradaic efficiency at -0.5 V (RHE)	>80% FE _{CO} after 12 hours	146 and 147

The production of Fe/Mn-N-C, a unique bimetallic catalyst consisting of iron and manganese atomic dispersion, involved the elevated temperatures calcination of an organic carbonbased porous precursor. The solution of potassium citrate monohydrate was initially calcined for an hour at 800 °C in a nitrogen atmosphere to create porous black carbon compounds. The resulting solid was dried in the oven for 12 hours at 80 °C after being rinsed with deionized water and a 1 M H₂SO₄ solution until it attained a neutral pH. A mixture consisting of carbon material, Fe(NO₃)₃·9H₂O, and MnCl₂·4H₂O in deionized H₂O was ultrasonically treated for an hour, centrifuged, and dried afterward. The resultant solid was combined with melamine in a particular mass ratio and then calcined at 800 °C, over a nitrogen environment for two hours to generate the Fe/Mn-N-C catalyst,¹⁴⁶ Fig. 3A. At a -0.5 V overpotential (RHE), the Fe/Mn-N-C catalyst produced a 94% Faraday efficiency (FE) for CO in the 0.1 M KHCO₃ electrolyte. This shows that, in these electrochemical circumstances, the catalyst has a high selectivity for CO synthesis. The catalyst's performance is notable when compared to previously published iron-based and manganese-based electrocatalysts, which include FeMn-N-C (FE_{CO} 80% at -0.5 V RHE), NFe-CNT/CNS (FE_{CO} 69% at -0.6 V RHE), and Mn-N-C (FE_{CO} 70% at -0.6 V RHE).^{53,148,149} Following just 12 hours of uninterrupted catalysis, the FE_{CO} was above 80%, suggesting good stability. Density functional theory (DFT) calculations show that the interaction of neighboring Fe-Mn centers lowers the potential for COOH* production and CO desorption.146

Development of atomically 6. distributed Co-Fe catalysts for CO₂ reduction

Bimetallic Co-Fe catalysts that are atomically distributed were developed in two steps. Using this method, the catalysts were synthesized with accurate atomic-level dispersion of iron and cobalt through a series of synthesis steps. To ensure a successful yield without interference in the crystallization of Co-ZIF, Fe-Co-ZIF precursors were generated by an impregnation process that modified ZIF-8 into Co-ZIF and absorbed Fe source. Pyrolysis was then used to manufacture the final catalysts (C-Fe-Co-ZIF) for CO₂ electro-reduction,¹⁵⁰ shown in Fig. 3B. The bimetallic catalysts produced more CO, with an additional 10% in CO Faradaic efficiency (FE) when compared to pure C-Co-ZIF. Adjustable H_2/CO ratios (0.8 to 4.2) reached across a wide potential range, with a high overall FE CO + H₂ of 93% over 10 hours, showing the catalyst's capacity for efficient syngas production from CO2.147

Graphene oxide-based catalysts for 7. CO₂ reduction

Graphene oxide (GO) was produced with graphite using the modified Hummers' method.¹⁵¹ GO suspension (2 mg mL⁻¹) was made by sonicating it in deionized water for 5 hours. Iron and nickel nitrates were introduced to the GO solution, which



Fig. 3 Methodology for the synthesis of (A) Fe/Mn–N–C and (B) C–Fe–Co-ZIF catalysts

was sonicated for three more hours. The resulting slurry was heated to 180 °C in an autoclave lined with Teflon for 12 h before being freeze-dried to generate a columnar product. Subsequently, a chemical vapor deposition (CVD) process at 1000 °C with Ar and NH3 was used to synthesize the H-NiFe/NG composite, followed by annealing with hot steam. A novel method involving steam-assisted chemical vapor deposition introduces surface oxygen vacancies (Vo) into Ni-Fe BM NPs, creating electron-rich centers that activate CO2 molecules.152,153 This method reduces the energy barrier for creating COOH* intermediates, increasing the reduction of carbon dioxide to CO while maintaining a faradaic efficiency of as high as 94% at -0.80 V (vs. RHE) along with excellent stability. Surface V₀modified atoms of nickel have a vital role in increasing the electrocatalytic efficacy of reduction of CO₂ to CO, according to density functional theory simulations.154

8. Molecular catalyst-based heterostructures for CO₂ reduction

The design and synthesis of a molecular catalyst-based heterostructure for the reduction of CO₂ is still a serious issue. Molecular catalysts with transition-metal elements (Co, Ru, Fe, Ni, Cu) and ligands made of organic compounds (phthalocyanine, polypyridine, porphyrin) provide precise active sites and structural tunability for researching CO₂ER processes.^{155–159} These catalysts facilitate detailed investigations into CO₂ reduction catalysis. A crystalline bimetallic phthalocyanine heterostructure electrocatalyst (CoPc/FePc HS) was developed for CO₂ reduction, achieving a remarkable CO₂ to CO conversion efficiency of 99% at the potential of -0.87 vs. RHE and demonstrating outstanding stability over 10 h of electrocatalysis. Different Co/Fe molar ratios (3:1, 1:1, 1:3)¹⁶⁰ of CoPc/FePc heterostructures, along with CoPc and FePc controls were synthesized by dispersing a mixture of CoPc and FePc in DMF and subjecting it to solvothermal treatment at 180 °C for 24 hours. Precipitates in the shape of purple microrods were gathered and cleaned with ethanol. They were then calcined for three hours at 450 °C in an Ar environment. CoPc/FePc heterostructures were formed as a consequence of this technique. This method provides a controlled approach to tailor the composition of bimetallic phthalocyanine heterostructures for CO₂ reduction applications.¹⁶¹

Cu–Fe–N₆–C: a highperformance diatomic site catalyst for CO₂ reduction

Metal-nitrogen-carbon (M–N–C) catalysts have great potential for CO₂ electrocatalytic reduction because of their abundance of active sites and low-cost raw ingredients.^{52,162–165} Cu–Fe–N₆–C,

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a new diatomic site catalyst coordinated with nitrogen and embedded into a carbon matrix, was developed, and synthesized. Cu-Fe-N₆-C was synthesized in two primary stages. First, PcCu-Fe-ZIF-8 is created by combining PcCu, zinc nitrate, iron nitrate, and 2-Me-imidazole, resulting in a blue precipitate that indicates uniform dispersion of Cu and Fe species inside the framework. PcCu-Fe-ZIF-8 becomes Cu-Fe bimetallic sites distributed on a nitrogen-doped carbon framework upon annealing at 1000 °C under Ar. The necessity for extra acid leaching treatment is eliminated by this technique. For a variety of processes, the resulting catalyst structure improves catalytic performance (Fig. 4). This catalyst outperformed individual Cu-N-C and Fe-N-C catalysts thanks to synergistic effects at bimetallic sites. Cu-Fe-N6-C demonstrated outstanding CO selectivity, with an exceptional faradaic efficiency of 98% at -0.7 V, and maintained selectivity after 10 hours of electrolysis. Experimental and theoretical investigations revealed that the combined catalysis of several metallic sites increased CO₂ adsorption enthalpy, and lowered activation energy, resulting in enhanced selectivity, activity, and stability, as well as decreased impedance in CO₂ hydrogenation.¹⁶⁶

For CO₂ conversion, some Na-promoted Co–Fe bimetallic catalysts ranging in proximity and compositions were investigated. These catalysts are designed to use the strong selectivity of iron for olefins during CO₂ hydrogenation, along with the high activity and reducibility of cobalt. The goal of this combination is to improve CO₂ conversion operations' overall efficiency and selectivity.^{167–170} Co-precipitation was used to produce Co–Fe bimetallic catalysts, which were then hydrothermally treated. The manufacture of uniform catalysts with regulated compositions and architectures is made easier by this technique.^{171,172} Co(NO₃)₂·6H₂O and Fe(NO₃)₃·9H₂O were

dissolved sequentially in deionized water to achieve a $[Co]^{2+}$ + $[Fe]^{3+}$ concentration of 0.09 M, followed by the addition of 5 mol L⁻¹ NaOH solution until pH 11 was reached. The resultant hydroxide precipitates were hydrothermally treated at 150 ° C for 24 hours before being centrifuged, washed, and dried at 80 °C. The dried products were calcined at 400 °C for 3 hours to produce Co–Fe catalysts with various Co/Fe molar ratios (1/4, 1/2, 1/1, 2/1, and 4/1), designated as Co1Fe4, Co1Fe2, Co1Fe1, Co2Fe1, and Co4Fe1, respectively. The Co1Fe2 catalyst, having a Co/Fe molar proportion of 1/2 and proximity, permitted the quick reduction of CoFe₂O₄ to Co_xFe_y alloy and subsequently carbonization to χ -(Co_xFe_{1-x})₅C₂ alloy carbide. It demonstrated improved stability and performance in olefin production without deactivation over 500 h on-stream.¹⁷³

10. Fe/Ni–N–C catalysts with 3D carbon-based structures for CO₂ reduction

A 3D carbon-based material was produced, featuring bimetallic centers¹⁷⁴ that include NiNC and FeNC, which demonstrated synergistic effects advantageous to the CO₂RR. The synthesis procedure involved numerous steps to produce various catalyst materials. Tripotassium citrate monohydrate was cooked at 800 °C under nitrogen, and then treated with sulfuric acid and water to create a porous carbon material. Next, a mixture containing carbon, nickel nitrate, iron nitrate, and glucose in water was processed using ultrasound and then combined with melamine. This mixture was heated at 800 °C under nitrogen to produce the NiNC/FeNC catalyst.¹⁷⁵ Further, specific catalysts like FePc@NiNC and NiPc@FeNC were prepared by treating



Fig. 4 Synthesis of Cu-Fe-N₆-C.



Fig. 5 Methodology for the synthesis of FePc@NiNC catalyst.

NiNC or FeNC with *N*,*N*-dimethylformamide and adding iron phthalocyanine (FePc) or nickel phthalocyanine (NiPc), respectively (Fig. 5). Each stage required precise chemical reactions and thermal treatments to generate the correct catalyst compositions. DFT models and observations show¹⁷⁶ Fe atoms are reactive and adsorption sites for CO₂RR, while substantial CO* adsorption reduces stability. By adding Ni atoms, CO* adsorption on Fe is decreased, changing the energy barriers and improving stability. The Fe–N₄ and Ni–N₄ sites work in concert to facilitate the rate-limiting processes (CO_{2(g)} \rightarrow COOH*, +0.95 eV) in FePc@NiNC. Flexible syngas composition is made

possible by this synergy while high catalytic activity is maintained.¹⁷⁷

11. Ni/Fe-N-C: diatomic metalnitrogen catalysts for CO₂ reduction

A ZIF-8 was used to create a catalyst consisting of isolated diatomic metal–nitrogen species. Initially, Fe-doped ZIF-8 was made by combining zinc nitrate, iron nitrate, and 2-methyl-imidazole, maintaining that Fe ions were chemically bound to the organic ligand rather than being physically absorbed.¹⁷⁸ Fe-



Fig. 6 Methodology for the synthesis of Ni/Fe-N-C catalyst.



Fig. 7 Synthesis of 57 FeCu and 57 FeAg NPs for CO₂ERR.

doped ZIF-8 was dissolved in *n*-hexane, and nickel nitrate methanol solution was added gradually. Nickel was contained within ZIF-8's tiny cavities using this method. Nickel was well incorporated into the framework owing to the steady infusion.^{90,179} After thermal treatment at 1000 °C, the resulting catalyst, Ni/Fe–N–C, containing nitrogen-coordinated diatomic Ni–Fe species, was obtained (Fig. 6). For comparison, crystalline Ni–N–C and Fe–N–C catalysts were synthesized similarly. After 30 hours, the Ni/Fe–N–C catalyst retains 99% selectivity and over 90% CO faradaic efficiency from -0.5 to -0.9 V, which ended at 98% at -0.7 V. Synergistic Ni–Fe interactions lower CO₂ reduction reaction barriers and cause structural changes upon CO₂ adsorption, according to DFT research, improving the catalyst's performance.¹⁸⁰

12. Micelle-encapsulated Fe-based nanoparticles for CO₂ reduction

Inverse micelle encapsulation was used to produce size-selected nanoparticles (NPs) of Fe, FeCu, FeAg, Ag, and Cu. The poly(styrene)-block-poly(2-vinylpyridine) (PS-P2VP) diblock copolymer, which was obtained from Polymer Source Inc., was used in this procedure. Metallic salts (FeCl₂, AgNO₃, CuCl₂, FeCl₃) and the copolymer were dissolved in toluene. By taking advantage of the micelles encapsulating attributes, this technique made controlled nanoparticle manufacturing easier.181,182 Specifically, isotopically enriched ⁵⁷FeCl₂ salt was employed for NRIXS measurements, prepared from iron foil with 95% 57Fe isotopic enrichment using adapted literature procedures.¹⁸³ Following NP synthesis, the samples were soaked with carbon black powder and then treated using N2-plasma to eliminate the polymer, resulting in clean NP surfaces. The NPs were subsequently distributed into an ethanol/Nafion solution enabling electrode deposition, accompanied by further N2-plasma treatment to remove any remaining polymer before electrochemical evaluation. Fig. 7 shows that the production of ⁵⁷Fe NPs involves mixing PS-P2VP in toluene to create reverse micelles, which were then added to 57FeCl2 salt and stirred for 72 h. Similar methods were utilized to create 57FeCu and 57FeAg NPs by changing the ratios of 57 FeCl₂ to CuCl₂ or AgNO₃ in the micellar solution. FeAg NPs had 36% CO faradaic selectivity at -1.1 V vs. RHE in 0.1 M KHCO₃, similar to pure Ag NPs, but FeCu NPs prefer H₂ evolution, similar to pure Fe NPs.¹⁸⁴ Table 4 summary of recently reported Fe-based bimetallic electrocatalysts for CO2 reduction.

13. Summary and outlook

Possibilities for the advancement of sustainable energy technology look promising for future studies on electrochemical CO_2 reduction with bimetallic catalysts. Optimizing catalyst compositions and structures to increase selectivity and efficiency in the production of CO, syngas, and other multi-carbon

Table 4	Summar	y of recently	/ reported	Fe-based	bimetallic	electrocatal	ysts for	CO ₂ reducti
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Catalyst	Electrolyte	Major product	FE (%)	Potential (V)	Current density (mA cm ⁻²)	Stability
Fe-NP/MNCF	0.5 M KHCO ₃	CO, H_2	87.50%	-0.7	10	36 h
FeNi@N-CNTs	0.5 M KHCO ₃	CO	90%	-0.47 to -0.97	20.18	35 h
Fe ₃ Ni ₇ -ZIF	0.5 M KHCO ₃	CO, H_2	81.30%	-0.9	-22.5	Good
Fe/Cu–N–C	0.5 M KHCO ₃	CO	97%	-0.6	74	10 h
Fe/Mn-N-C	0.1 M KHCO ₃	CO	94%	-0.5	-83.5	12 h
Fe/Ni-ZIF-8	0.5 M KHCO ₃	CO	89%	-0.677	26.92	40 h
FePc@NiNC	0.5 M KHCO ₃	CO, H_2	100%	-0.8	260	18 h
Ni/Fe–N–C	0.5 M KHCO ₃	CO	98%	-0.7	7.4	30 h
H-NiFe/NG	0.1 M KHCO ₃	CO	94%	-0.8	18.2	20 h
PcCu-Fe-ZIF-8	0.1 M KHCO ₃	CO	98%	-0.7	7	10 h
⁵⁷ FeAg NPs	0.1 M KHCO ₃	CO	36%	-1.1	0.35	2.5 h
Fe-Co-ZIF	0.5 M KHCO ₃	CO, H_2	93%	-0.55	8	10 h
CuFe/OG	_	CH_4	_	0.97	_	_
FeNi/DG	_	CH_4	_	-0.44	_	_
FeCo-Pc	_	C_{2^+}	—	-0.66	_	

products is a crucial field of research. Fe-Ni, Fe-Ag, Fe-Mo, Fe-Co, Cu-Fe, and Fe/Mn-N are examples of novel metal combinations that present the potential for enhanced catalytic performance. The main goals of the research will be to comprehend the fundamental structure-activity correlations and stability of these bimetallic catalysts in practical working environments. The scalability of bimetallic catalysts for largescale commercial applications is limited by their typical synthesis, which involves intricate deposition, pyrolysis, and reduction methods. The development of more affordable, optimized synthesis techniques with improved loading capacities is necessary to meet this challenge and permit Fe-based bimetallic catalysts to be widely used in renewable energy systems. Moreover, a major challenge presented by the chemical instability of these catalysts is the reduction of active sites and changed performance caused by corrosion of the carbon substrate. Under practical circumstances, Fe-based bimetallic electrocatalysts for CO2 reduction encounter difficulties such as low catalytic activity, poor selectivity for target products, and restricted stability. Controlling the chemical intermediates, improving the electrical and geometric properties, and interpreting the synergistic effects between metals are still major challenges. Precise control of Fe-based bimetallic catalysts' shape, structure, and atomic coordination is required to strike a compromise between stability and catalytic activity. The aim is to design specialized bimetallic catalysts with enhanced stability features and active sites outperforming existing catalysts. This will facilitate the development of scalable CO₂ conversion technologies for use in sustainable energy applications, assisting in the shift to a world without carbon emissions. The development of effective CO₂ electroreduction catalysts will be speed up by collaborative, multidisciplinary research that combines theoretical and experimental methods. Future research should concentrate on investigating novel bimetallic combinations that improve performance and customizing catalyst structures by nano-structuring. Enhancing these catalysts' scalability for industrial applications is also essential. Developments in reaction mechanism research, computational modeling, and in situ characterization methods will improve catalyst design and propel more effective CO₂ conversion systems.

Data availability

No primary research results, software, or code have been included, and no new data were generated or analyzed as part of this review.

Author contributions

Ayesha Zafar: writing – original draft. Adnan Majeed: writing – review & editing and software. Abdul Ahad: formal analysis. Muhammad Adnan Iqbal: conceptualization, resources, supervision. Tanveer Hussain Bokhari: validation. Zanira Mushtaq: data curation, validation. Shahzaib Ali: visualization.

Conflicts of interest

The authors declare no conflict of interest.

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