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# Combination of organocatalytic oxidation of alcohols and organolithium chemistry (RLi) in aqueous media, at room temperature and under aerobic conditions†

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**A tandem protocol to access tertiary alcohols has been developed which combines the organocatalytic oxidation of secondary alcohols to ketones followed by their chemoselective addition by several RLi reagents. Reactions take place at room temperature, under air and in aqueous solutions, a trio of conditions that are typically forbidden in polar organometallic chemistry.**

As Nature's solvent in biological processes, the use of water as a medium in organic synthesis<sup>1</sup> is widely recognized as an important strategy towards developing sustainable chemical processes that implement some of the most significant Principles of Green Chemistry.<sup>2</sup> Water is abundant, non-toxic, cost-effective, non-flammable, renewable, and non-volatile<sup>3</sup> and it has already shown an enormous potential as a versatile reaction medium in many cornerstone organic transformations such as oxidations, reductions, reductive aminations and pericyclic reactions to name just a few,<sup>1</sup> as well as in a myriad of transition-metal-catalyzed reactions.<sup>4</sup> Furthermore, in many cases water is not only a greener alternative to toxic organic solvents but can also be a superior medium, improving reactivities and selectivities, simplifying work up procedures and facilitating catalyst recycling.<sup>5</sup> Nevertheless, despite these key advantages, until recently, water has shown very little promise as a solvent to carry out organolithium chemistry with only a few examples in the literature.<sup>6</sup> This is probably as a consequence of the common assumption that these polar reagents rapidly decompose in contact with this solvent. In fact, the use of protective inert atmosphere conditions and stringently dry organic solvent along with strict control of the reaction temperatures (usually of the order

of  $-78\text{ }^{\circ}\text{C}$ ) are usually mandatory to avoid fast degradation and control the exceptionally high reactivity of organolithium reagents.<sup>7</sup> However, what was thought impossible just a few years ago, an aerobic sustainable polar organometallic chemistry, is now starting to look distinctly possible.<sup>6,8</sup> Thus, Capriati has reported the chemoselective addition of organolithiums to carbonyl compounds<sup>8d</sup> and to nitriles and imines<sup>8h</sup> under on water conditions as well as palladium catalyzed direct cross-couplings of RLi and hetero(aryl) halides.<sup>8i</sup> Furthermore bioinspired Deep Eutectic Solvents (DESSs) which are usually made up of a combination of biorenewable hydrogen bond acceptors [*e.g.*, ammonium salts like choline chloride (*ChCl*)] and hydrogen bond donor (*e.g.*, water, glycerol)<sup>9</sup> have also shown great promise as non-conventional media for air and moisture compatible organolithium synthesis as demonstrated recently by our groups for nucleophilic arylation/alkylation of ketones, imines and nitriles<sup>10</sup> as well as anionic polymerization of styrenes.<sup>11,12</sup> Remarkable, in some of these examples, higher yields and selectivities were achieved than when using inert atmosphere protocols. Expanding even further the synthetic applications of polar organometallic chemistry in water, herein we report a new protocol that pairs these operational breakthroughs with an organocatalytic transformation, uncovering a novel one-pot tandem approach to access highly-substituted tertiary alcohols.

By eliminating the need of isolation and purification of reaction intermediates (and subsequently minimizing chemical waste generation, reaction times and energy consumption), one-pot tandem processes are highly useful and efficient green approaches in modern synthesis for the production of organic molecules.<sup>13</sup> More often than not, these protocols tend to use the same type of transformational tools (metal-, main-group element-, bio- or organo-catalysts) throughout all the tandem process, whereas the number of examples which combine different methodologies is scarce, especially those that also employ sustainable solvents. In this regard recently we have successfully combined transition-metal-catalysis in water or DESSs with (Scheme 1a): (i) chemoselective addition of organolithium

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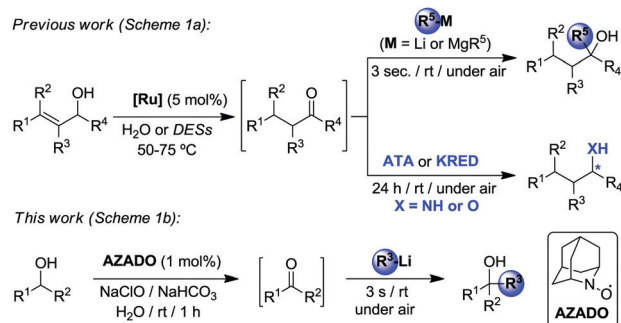
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**Scheme 1** Design of one-pot tandem combinations that rely on different organic synthetic tools in water or *DESs*.

reagents ( $\text{RLi}$ ) to ketones as transient intermediates;<sup>14</sup> and (ii) enzymatic enantioselective bioamination (transaminases, ATA) or bioreduction (ketoreductases, KRED) of *in situ* generated prochiral ketones.<sup>15</sup> Breaking new ground in this evolving field, we combine for the first time organocatalysis and organolithium chemistry in water as a new methodology to prepare highly substituted tertiary alcohols. The approach employed comprises the organo-catalyzed oxidation of secondary alcohols (furnishing ketones *in situ* as intermediates) followed by subsequent chemoselective and ultrafast addition of organolithium reagents affording the relevant tertiary alcohols. Reactions were carried out under one-pot conditions in aqueous media, under air and at room temperature (Scheme 1b).

Building on previous studies reported by some of us, in which we have uncovered the efficient organocatalytic activity of AZADO (2-azaadamantane *N*-oxyl, **1a**) in combination with an aqueous solution of  $\text{NaClO}$  (**2a**, acting as external oxidant) as a straightforward and operationally simple method for the selective oxidation of secondary alcohols;<sup>16</sup> we selected, as a model reaction, the oxidation of the benzylic-type alcohol **3a** (see Table 1). Firstly, we assessed if the use of toxic organic co-solvents such as  $\text{PhCF}_3$ ,  $\text{CH}_2\text{Cl}_2$  or  $\text{CH}_3\text{CN}$  was actually required in order to achieve quantitative yields of the desired ketone **4a** (entries 1–3, Table 1).<sup>16</sup>

**Table 1** Oxidation of 1-phenylpropan-1-ol (**3a**) into propiophenone (**4a**) organocatalyzed by AZADO (**1a**) or minoxidil (**1b**) in combination with different external oxidants (**2a–c**) in aqueous media, at room temperature and under air<sup>a</sup>

Entry	Organocat.	Oxidant	Co-solvent	Yield <sup>b</sup> (%)
1	AZADO ( <b>1a</b> )	aq. $\text{NaClO}$ ( <b>2a</b> )	$\text{PhCF}_3$	95
2	AZADO ( <b>1a</b> )	aq. $\text{NaClO}$ ( <b>2a</b> )	$\text{CH}_2\text{Cl}_2$	99
3	AZADO ( <b>1a</b> )	aq. $\text{NaClO}$ ( <b>2a</b> )	$\text{CH}_3\text{CN}$	99
4	AZADO ( <b>1a</b> )	aq. $\text{NaClO}$ ( <b>2a</b> )	—	99
5	AZADO ( <b>1a</b> )	$\text{KMnO}_4$ ( <b>2b</b> )	—	69
6	AZADO ( <b>1a</b> )	$\text{I}_2$ ( <b>2c</b> )	—	13
7	Minoxidil ( <b>1b</b> )	aq. $\text{NaClO}$ ( <b>2a</b> )	—	17
8	Minoxidil ( <b>1b</b> )	$\text{KMnO}_4$ ( <b>2b</b> )	—	42
9	Minoxidil ( <b>1b</b> )	$\text{I}_2$ ( <b>2c</b> )	—	4

<sup>a</sup> General reaction conditions are described in ESI. <sup>b</sup> Isolated yields.

Pleasingly, we found that 1-phenylpropan-1-ol (**3a**) can be quantitatively converted into the desired propiophenone (**4a**, entry 4, Table 1) under the same reaction conditions (1 h, rt, 270 mM loading of substrate) in the absence of these toxic organic co-solvents. For completeness of this parametrization, we also studied the use of other co-oxidants for AZADO (**1a**) like  $\text{KMnO}_4$  (**2b**, entry 5, Table 1) or  $\text{I}_2$  (**2c**, entry 6, Table 1) finding in both cases reduced activity (69–13%). Interestingly, we found that when AZADO was replaced by a different amine *N*-oxide such as minoxidil (**1b**, 6-piperidin-1-ylpyrimidine-2,4-diamine 3-oxide) in combination with the aforementioned sacrificial oxidants ( $\text{NaClO}$ , entry 7;  $\text{KMnO}_4$ , entry 8; and  $\text{I}_2$ , entry 9; Table 1) significant lower conversions were observed (4–42% yields), highlighting the high efficiency of the organocatalytic/oxidation  $\text{NaClO}$ /AZADO system while operating in aqueous media.

Encouraged by these initial results, which show that we can quantitatively prepare ketones from secondary alcohols in aqueous media at room temperature and in the absence of any organic co-solvents, we next pondered if this transformation could also be part of a modular combination, where the relevant synthesized ketones could subsequently undergo chemoselective alkylation using a one-pot approach while operating in aqueous media under air. Towards this target, it should be noted that *s*-block polar organometallic reagents such as Grignard ( $\text{RMgX}$ ) and organolithium ( $\text{RLi}$ ) reagents are widely used for addition reactions to ketones, being a fundamental and versatile method to generate C–C bonds to access tertiary alcohols.<sup>17</sup> Notwithstanding, the chemoselectivity of these processes can be compromised by the formation of undesired reduction and/or enolization products, resulting from competing  $\beta$ -hydride elimination and deprotonation reactions respectively.<sup>18</sup> Recently our groups have shown using *DESs* which combine ammonium salt choline chloride (*ChCl*) with water or glycerol (*Gly*) as reaction media, Grignard and organolithium reagents can promote room temperature chemoselective ketone alkylation/arylation reactions.<sup>10</sup> Along with greater selectivities and milder reaction conditions, these reactions are compatible with the presence of air.<sup>10</sup> However, *DESs* are not compatible with the organocatalytic oxidation system AZADO/ $\text{NaClO}$ , as the addition of  $\text{NaClO}$  to the eutectic reaction media produced the formation of  $\text{Cl}_2$  gas through the expected comproportionation reaction between the  $\text{Cl}^-$  and  $\text{ClO}^-$  anions.

We next assessed whether this tandem oxidation/alkylation approach could be compatible with aqueous media. Thus, we selected as a model reaction the AZADO/ $\text{NaClO}$  organocatalyzed oxidation of 1-phenylpropan-1-ol (**3a**) into ketone **4a** followed by the direct addition of *n*-BuLi (Table 2). Firstly, the reaction mixture containing the alcohol **3a** (270 mM) and the organocatalytic oxidation system (AZADO/ $\text{NaClO}$ ) was allowed to react in aqueous media, at room temperature and in the presence of air for 1 h to trigger the expected oxidation. As soon as the total conversion of **3a** into propiophenone (**4a**) was observed (1 h, GC analysis), *n*-BuLi was directly added to the reaction mixture (without any intermediate step of isolation or purification of ketone **4a**), in the absence of protecting atmosphere and at room temperature (conditions usually forbidden for  $\text{RLi}$  reagents).<sup>17</sup> Astonishing, the almost instantaneous formation (3 s) of the relevant tertiary alcohol **5a**

**Table 2** Direct conversion of 1-phenylpropan-1-ol (**3a**) into tertiary alcohols **5a–d** through the one-pot combination of the organocatalyzed (AZADO/NaClO) oxidation of alcohols and the chemoselective addition of RLi/RMgX in aqueous media, at room temperature and in the presence of air<sup>a</sup>

Entry	R-M	Equivalents	Product	Yield <sup>b</sup> (%)
1	<i>n</i> -BuLi	1	<b>5a</b>	23
2	<i>n</i> -BuLi	2	<b>5a</b>	47
3	<i>n</i> -BuLi	2.5	<b>5a</b>	62
4	<i>n</i> -BuLi	3	<b>5a</b>	76
5	MeLi	3	<b>5b</b>	70
6	<i>sec</i> -BuLi	3	<b>5c</b>	83
7	PhLi	3	<b>5d</b>	35
8	<i>n</i> -BuMgCl	3	<b>5a</b>	—

<sup>a</sup> General reaction conditions are presented in ESI. <sup>b</sup> Isolated yields.

was observed, finding a significantly improved yield of 76% when three equivalents of the organolithium reagent were employed (entry 4, Table 2).<sup>19</sup> However, gradually decreasing the number of equivalents of *n*-BuLi (from 3 to 1) produces a concomitant diminution of the nucleophilic addition of *n*-BuLi into the *in situ* formed ketone **4a**, from 76 to 23% yield of the final tertiary alcohol **5a** (entries 1–4, Table 2). At this point, it is important to note that in all cases (entries 1–4, Table 2) the one-pot tandem construction of the tertiary alcohol **5a** occurred with total chemoselectivity, as no side products were observed in the crude reaction mixture (containing only the desired final product **5a** and the unreacted ketone **4a**). We next decided to expand the scope to other organolithium (RLi) or organomagnesium (RMgX) reagents.

Interestingly we found that other aliphatic organolithium reagents such as MeLi (entry 5, Table 2); *sec*-BuLi (entry 6, Table 2) can selectively add across the C=O double bond of **4a** under the above mentioned conditions, affording alcohols **5b** and **5c** in excellent yields (70 and 83% respectively) whereas using less nucleophilic PhLi, **5d** is obtained in a 35% yield (entry 6, Table 2). The high conversions and chemoselectivities observed while operating at room temperature are particularly noteworthy for *sec*-BuLi which has a greater tendency to undergo  $\beta$ -hydride elimination, and has to be usually stored at low temperatures (2–8 °C) to avoid degradation *via* this side reaction. Furthermore, Ishihara has shown that when *n*-BuLi is employed as an alkylating reagent towards enolizable ketones variable amounts of the aldol condensation products are obtained even when operating at –78 °C in THF under strict inert atmosphere conditions.<sup>18b</sup> On the other hand, this method does not seem compatible with Grignard reagents and when *n*-BuMgCl was employed as an alkylating reagent, alcohol **5a** could not be detected (entry 8, Table 2). This lack of reactivity is consistent with previous studies in the addition of Grignard reagents to carbonyl compounds in water that have shown that *n*-BuMgCl undergoes complete protonation in competition with carbonyl addition.<sup>20</sup> At this point we would like to highlight the use of commercially available solutions of RMgX and RLi reagents containing different organic solvents, namely hexane for *n*-BuLi, diethyl ether for MeLi, and

dibutyl ether for PhLi. We hypothesize that besides the intrinsic reactivity of each reagent, both the nature and the amount of the organic accompanying solvent is critical for the outcome of the addition reaction, based on the premise that the reactions are taking place under “on water” conditions.<sup>21</sup> Accordingly, the more water-immiscible the RLi solvent and the smaller concentration of the commercially available organolithium solution, the more favoured that interface and therefore the better performance. In good agreement, the highest yield was obtained in cyclohexane (*sec*-BuLi, entry 6), the most hydrophobic solvent of the series and with the smaller concentration (1.4 M), followed by hexane (*n*-BuLi, 1.6 M, entry 4) and finally the ether solvent (entries 5 and 7). In the same line, another interesting finding was that when acetophenone was reacted with *n*-BuLi using water as a reaction medium under exactly the same conditions depicted in Table 2, the alkylation reaction is suppressed. Whereas, with propiophenone (**4a**), alcohol **5a** was obtained in a 76%. A plausible explanation could be the different solubilities in water of both ketones, with acetophenone showing a greater solubility in water than **4a**, again in agreement with the possibility that these addition reactions are taking place under “on water” conditions.<sup>21</sup> “On water” reactions are thought to occur at the organic/liquid water interface with water insoluble reactants. In this regard Capriati *et al.* have recently postulated that interfacial H-bonding to water insoluble carbonyl compounds embedded at the organic/water interphase may play a key role in favouring “on-water” nucleophilic additions promoted by RMgX and RLi reagents.<sup>8d,h</sup> Once the scope of the *s*-block highly-polar organometallic reagents was established, we decided to further expand this one-pot tandem oxidation/RLi addition in aqueous media to a series of secondary alcohols (**3a–f**, Table 3) and *n*-BuLi or *sec*-BuLi as alkylating reagents.

We found that this method tolerates a variety of functional groups in the aromatic scaffold of the starting alcohols **3a–f** being compatible with the presence of either electron-withdrawing (**3b–c**, entries 3–6) or electron-donating groups (**3d–f**, entries 7–12) giving rise to the desired tertiary alcohols **5a,c,e–n** in good yields (up to 95%), and without the need of any halfway step of isolation or purification of the intermediate ketones **4a–f**, which were always formed in quantitative yields after 1 h of reaction (GC analysis). In general, better yields of the final tertiary alcohols were obtained when *n*-BuLi (even entries in Table 3) was used as alkylating reagent. This experimental observation is in good agreement with the expected lower nucleophilic activity of the more sterically demanding *sec*-BuLi when compared with linear *n*-BuLi.<sup>17</sup> At this point, it is important to note that although electronic properties of the substituents in the aromatic ring have only a modest effect, the presence of an *ortho*-substituent (entries 11 and 12, Table 3) produces a considerable decrease in the yield of the final alcohols **5m–n**. This experimental observation seems to indicate that the electronic effects have a smaller impact than steric hindrance which can be associated to the virtually instantaneous nature of the addition reactions in aqueous media. Interestingly, other possible competing processes with the addition reaction or RLi to ketones like: (i) metalation of benzylic position in substrates **3d–f**; or (ii) Li-halogen exchange processes in alcohols **3b–c**, were not detected.

**Table 3** AZADO/NaClO organocatalyzed oxidation of secondary alcohols (**3a–f**) combined with the chemoselective addition of organolithium reagents in a one-pot process, at room temperature, in the presence of air and in aqueous media<sup>a</sup>

Entry	R <sup>1</sup>	R <sup>2</sup> -M	Product	Yield <sup>b</sup> (%)
1	H ( <b>3a</b> )	<i>n</i> -BuLi	<b>5a</b>	76
2	H ( <b>3a</b> )	<i>sec</i> -BuLi	<b>5c</b>	83
3	<i>p</i> -Cl ( <b>3b</b> )	<i>n</i> -BuLi	<b>5e</b>	82
4	<i>p</i> -Cl ( <b>3b</b> )	<i>sec</i> -BuLi	<b>5f</b>	54
5	<i>p</i> -Br ( <b>3c</b> )	<i>n</i> -BuLi	<b>5g</b>	95
6	<i>p</i> -Br ( <b>3c</b> )	<i>sec</i> -BuLi	<b>5h</b>	73
7	<i>p</i> -Me ( <b>3d</b> )	<i>n</i> -BuLi	<b>5i</b>	90
8	<i>p</i> -Me ( <b>3d</b> )	<i>sec</i> -BuLi	<b>5j</b>	66
9	<i>p</i> -MeO ( <b>3e</b> )	<i>n</i> -BuLi	<b>5k</b>	40
10	<i>p</i> -MeO ( <b>3e</b> )	<i>sec</i> -BuLi	<b>5l</b>	74
11	<i>o</i> -Me ( <b>3f</b> )	<i>n</i> -BuLi	<b>5m</b>	68
12	<i>o</i> -Me ( <b>3f</b> )	<i>sec</i> -BuLi	<b>5n</b>	39

<sup>a</sup> General reaction conditions are described in ESI. <sup>b</sup> Isolated yields.

In conclusion, edging closer towards the development of air and water compatible organolithium chemistry, these findings have uncovered the potential of using water as an alternative sustainable reaction medium for tandem protocols that combine organocatalysis with the use of organolithium reagents. Interestingly water not only enables more environmentally benign reaction conditions (under air, room temperature) but also greater selectivities than when using conventional organic solvents under inert atmosphere conditions.

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## Conflicts of interest

There are no conflicts to declare.

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