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# Recent advances in developing nanoscale electro-/photocatalysts for hydrogen production: modification strategies, charge-carrier characterizations, and applications

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For clean hydrogen (H<sub>2</sub>) production, electrocatalysis and photocatalysis are widely regarded as promising technologies to counter the increasing energy crisis. However, developing applicable catalysts with high H<sub>2</sub> production performances still poses a challenge. In this review, state-of-the-art nanoscale electrocatalysts for water electrolysis and photocatalysts for water splitting, tailored for different reaction environments, including acidic electrolytes, alkaline electrolytes, pure water, seawater, and hydrohalic acids, are systematically presented. In particular, modification approaches such as doping, morphology control, heterojunction/homo-junction construction, as well as the integration of cocatalysts and single atoms for efficient charge transfer and separation are examined. Furthermore, the unique properties of these upgraded catalysts and the mechanisms of promoted H<sub>2</sub> production are also analyzed by elucidating the charge carrier dynamics revealed by photophysical and photoelectrochemical characterization methods. Finally, perspectives and outlooks on future developments for H<sub>2</sub> production using advanced electrocatalysts and photocatalysts are proposed.

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## 1. Introduction

The production of hydrogen fuel through water electrolysis, photocatalytic water splitting, and photocatalytic hydrohalic acid (HX) splitting offers a promising green synthesis approach for the development of renewable energy.<sup>5–13</sup> Nevertheless, approximately 95% of global hydrogen production relies on fossil fuels, resulting in carbon emissions, whereas approximately only 4% of global hydrogen production depends on water electrolysis.<sup>19,20</sup> Current research efforts are focused on

water electrolysis and photocatalysis to mitigate the reliance on fossil fuels for hydrogen energy production.<sup>5,7–9,22–24</sup>

Electrochemical water splitting comprises two half-cell reactions: the hydrogen evolution reaction (HER) at the cathode, which involves the transfer of two electrons, and the oxygen evolution reaction (OER) at the anode, which necessitates the participation of four electrons. These reactions occur in acidic or alkaline media. Presently, noble metal electrocatalysts such as Ir/Ru-based compounds for the OER and Pt for the HER are considered to be promising electrocatalysts for water electroly-

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sis.<sup>27</sup> Furthermore, the OER electrocatalysts currently utilized for water electrolysis in acidic media are both costly and scarce.<sup>27,29</sup> In contrast, utilizing electrocatalysts in an alkaline environment can significantly lower the overall expense of water electrolyzer systems.<sup>12,30,31</sup> Consequently, it is crucial to focus on the development of electrocatalysts specifically designed for water electrolysis under alkaline conditions.

Nevertheless, the use of bifunctional catalysts for water electrolysis may lead to the mixing of gases ( $H_2$  and  $O_2$ ) during collection, introducing additional costs for gas separation, which represents a potential drawback. In this context, proton exchange membrane (PEM) water electrolysis in acidic environments could mitigate the issue of gas mixing to a certain extent. Nonetheless, this approach necessitates the development of acid-stable HER and OER catalysts for efficient PEM water electrolysis.<sup>33</sup> Therefore, developing electrocatalysts for water electrolysis in an acidic environment is essential.

Water electrolysis technologies, including proton exchange membranes (PEMs) and alkaline electrolyzers, traditionally rely on high-purity water as a feedstock.<sup>34,35</sup> Given that Earth-abundant seawater constitutes approximately 96.5% of the world's water reserves, its utilization as a feedstock for water electrolysis could significantly reduce the reliance on scarce freshwater resources. This shift is particularly critical for regions with high freshwater demands, such as coastal countries, arid zones, and islands.<sup>23,36</sup> However, the adoption of seawater electrolysis for industrial applications is currently limited due to the complex composition of seawater electrolytes.<sup>37,38</sup> Consequently, there is a pressing need to develop electrocatalysts tailored to the seawater electrolysis environment to overcome these challenges. Photocatalytic  $H_2$  production *via* water splitting using photocatalysts is a promising route to convert solar energy into clean chemical fuels.<sup>11</sup> Designing photocatalysts with suitable bandgaps for the efficient absorption of sunlight, excellent charge separation and transfer efficiency, and satisfactory surface reactions is essential. Besides, photocatalytic  $H_2$  production using hydrohalic acid (HX) is a promising strategy, where value-added chemicals (such as  $X_2/X_3^-$ ) can also be generated along with clean, renewable  $H_2$  fuel.<sup>5,9,10,49</sup> Developing photocatalysts that are both resistant to strong acids and active under visible light, especially eco-friendly, lead-free options for hydrogen production *via* hydrohalic acid (HX) splitting, presents a significant challenge. Thus, advancing the creation of such materials is crucial for efficient photocatalytic  $H_2$  generation.

The Nyquist plot, from electrochemical impedance spectroscopy (EIS) data, is a simple and useful tool for developing electrocatalysts for hydrogen production, where comparing the quantified charge transfer resistance of the electrocatalyst with its counterparts can provide valuable information about the conductivity of the electrocatalysts. Generally, a wide range of electrocatalysts demonstrate high activity and low charge transfer resistance for the hydrogen evolution reaction (HER) and/or the oxygen evolution reaction (OER) across various environments, including alkaline, acidic, and seawater.<sup>1,4,40,50</sup> However, exceptions exist,<sup>39,51</sup> as the catalytic activity for the HER or the OER is not solely determined by conductivity. It also relies on

several critical factors: the availability of abundant active sites, optimal adsorption energies for HER/OER intermediates, a suitable electronic structure, a large electrochemically active surface area, and the ability to facilitate gas evolution.<sup>30,33,36,52,53</sup> Xi *et al.*<sup>4</sup> observed that Cu  $RuO_2^{activated}$  exhibited a lower charge transfer resistance ( $R_{ct1}$ :  $0.27\text{ W cm}^{-2}$ ;  $R_{ct2}$ :  $0.09\text{ W cm}^{-2}$ ) for the HER compared to  $RuO_2^{activated}$  ( $R_{ct1}$ :  $4.49\text{ W cm}^{-2}$ ;  $R_{ct2}$ :  $0.44\text{ W cm}^{-2}$ ), and it also demonstrated a higher activity ( $\eta$  of  $-19\text{ mV}$  at  $-10\text{ mA cm}^{-2}$ ) and larger electrochemical surface area for the HER in  $1\text{ M KOH}$  compared to  $RuO_2^{activated}$ . Thus, doping Cu into  $RuO_2$  facilitates the partial formation of metallic Ru as active sites for the HER during electrochemical activation, leading to enhanced conductivity and activity for the HER.

Liu *et al.*<sup>40</sup> reported that Ru- $Ni_3N/NiO$  exhibited a lower charge transfer resistance ( $R_{ct}$ :  $1.52\text{ }\Omega\text{ or }\Omega\text{ cm}^{-2}$ ) and higher activity ( $\eta$  of  $-190\text{ mV}$  at a huge current density of  $-1000\text{ mA cm}^{-2}$ ) for the HER than those of Ru- $Ni_3N$  and Ru- $NiO$ , and it also exhibited a much higher activity for the OER ( $\eta$  of  $385\text{ mV}$  at a huge current density of  $1000\text{ mA cm}^{-2}$ ) and overall water splitting (potential of  $1.74\text{ V}$  at a huge current density of  $1000\text{ mA cm}^{-2}$ ) in  $1\text{ M KOH}$ . Theoretical calculations suggested that the structural stability, orbital coupling, and electronic distributions of Ru- $Ni_3N/NiO$  could be enhanced by the  $Ni_3N/NiO$  heterostructures when compared to those of Ru- $Ni_3N$  and Ru- $NiO$ ; this enhancement could lower the energy barriers and increase the activity for the HER and the OER. In addition, Zhao *et al.*<sup>1</sup> reported that strained  $Mn_{0.98}Ir_{0.02}O_2$  exhibited a higher activity ( $\eta$  of  $105\text{ mV}$  at  $10\text{ mA cm}^{-2}$ ; mass activity:  $300\text{ mV at }5681\text{ A g}^{-1}$ ) and lower charge transfer resistance than those of unstrained  $Mn_{0.98}Ir_{0.02}O_2$  for the OER in  $0.5\text{ M H}_2\text{SO}_4$ . The incorporation of Ir into  $MnO_2$  can induce strain, leading to an elongated Mn-O bond length and pushing the d-band center nearer to the Fermi level, while the holes in the O 2p orbitals can more readily transfer to the catalyst surface, leading to acceleration in charge transfer, and thereby modify the electronic structure, leading to improved activity for the OER. Moreover, Wang *et al.*<sup>50</sup> observed that Co-N,P-HCS (HCS: hollow carbon spheres) exhibited a higher activity and lower charge transfer resistance than those of HCS, Co-HCS, and Co-N-HCS for the HER and the OER in seawater with  $1\text{ M KOH}$ , and it exhibited high stability for 1000 h. Co-N,P-HCS is composed of an asymmetric Co- $N_3P_1$  structure, which is implanted onto hollow carbon spheres. DFT calculations reveal that asymmetric Co- $N_3P_1$  sites with symmetry-breaking electronic structures can be formed due to P-doping of Co-N-C, which enhances the affinity of strong oxygen-containing intermediates with weak  $Cl^-$  adsorption and moderate H adsorption. On the other hand, Seenivasan *et al.*<sup>39</sup> observed that  $Nb_4N_5$  exhibited a significantly higher activity ( $\eta$  of  $-7\text{ mV}$  at  $-10\text{ mA cm}^{-2}$ ;  $\eta$  of  $-281\text{ mV}$  at a huge current density of  $-1000\text{ mA cm}^{-2}$ ) and lower charge transfer resistance ( $R_{ct}$ :  $0.892\text{ }\Omega$ ; Table 1) than those of  $Fe_4N/Nb_4N_5$  ( $R_{ct}$ :  $2.165\text{ }\Omega$ ) for the HER in  $1\text{ M KOH}$ , while  $Fe_4N/Nb_4N_5$  exhibited a significantly higher activity ( $\eta$  of  $438\text{ mV}$  at a huge current density of  $1000\text{ mA cm}^{-2}$ ) and lower charge transfer resistance ( $R_{ct}$ :  $0.206\text{ }\Omega$ ) than those of  $Nb_4N_5$  ( $R_{ct}$ :  $1.369\text{ }\Omega$ ) for the OER in  $1\text{ M KOH}$ . This obviously demonstrates the role of con-

**Table 1** Overpotential ( $\eta$ ) to achieve  $-10 \text{ mA cm}^{-2}$  for the hydrogen evolution reaction (HER) of various electrocatalysts in 1 M KOH alkaline electrolyte ( $\eta$  to achieve greater than  $-10 \text{ mA cm}^{-2}$  are indicated with their corresponding current densities), where  $R_{\text{ct}}$  or  $R_{\text{ct1}}$  and  $R_{\text{ct2}}$  ( $R_{\text{ct}}$  or  $R_{\text{ct1}}$  or  $R_{\text{ct2}}$ : charge transfer resistance) are acquired from electrochemical impedance spectroscopy (EIS)

| HER electrocatalysts                     | $\eta$ at $-10 \text{ mA cm}^{-2}$ (mV) | $R_{\text{ct}}$ or $R_{\text{ct1}}$       | $R_{\text{ct2}}$              | Stability                                                                | Ref. |
|------------------------------------------|-----------------------------------------|-------------------------------------------|-------------------------------|--------------------------------------------------------------------------|------|
| Nb <sub>4</sub> N <sub>5</sub>           | -7                                      | 0.892 $\Omega$                            | NA                            | NA                                                                       | 39   |
| Ru-Ni <sub>3</sub> N/NiO                 | -15                                     | 1.52 $\Omega$ or $\Omega \text{ cm}^{-2}$ | NA                            | 1000 h ( $500 \text{ mA cm}^{-2}$ )                                      | 40   |
| Cu-RuO <sub>2</sub> <sup>activated</sup> | -19                                     | 0.27 $\Omega \text{ cm}^{-2}$             | 0.09 $\Omega \text{ cm}^{-2}$ | 50 h                                                                     | 4    |
| FeCoCrCuO <sub>x</sub>                   | -40                                     | NA                                        | NA                            | 100 h ( $500 \text{ mA cm}^{-2}$ )                                       | 41   |
| Zr <sub>0.1</sub> Ni <sub>1.9</sub> P    | -68                                     | 5 $\Omega$                                | NA                            | 100 h                                                                    | 42   |
| P/Fe-NiSe <sub>2</sub>                   | -74                                     | 4.78 $\Omega$                             | NA                            | Less than 5% current density loss over 24 h ( $100 \text{ mA cm}^{-2}$ ) | 43   |
| Fe@Co/Se <sub>2</sub>                    | -78                                     | 10.80 $\Omega$                            | NA                            | 0.3% loss of over potential after 20 h ( $10 \text{ mA cm}^{-2}$ )       | 14   |
| DH-CuCo-P@NC/CC                          | -85                                     | 2.4 $\Omega$                              | NA                            | 12 h                                                                     | 44   |
| cNiFe/aNiFeOOH@NiMo                      | -91.9 at $-20 \text{ mA cm}^{-2}$       | 0.946 $\Omega$                            | NA                            | 100 h ( $100 \text{ mA cm}^{-2}$ )                                       | 45   |
| 1T-Mn-VS <sub>2</sub> @Co <sub>2</sub> P | -98                                     | 2 $\Omega$                                | NA                            | 100 h ( $500 \text{ mA cm}^{-2}$ )                                       | 46   |
| CoP <sub>3</sub> -Nb <sub>2</sub> P/PCC  | -111                                    | 3.5 $\Omega$                              | NA                            | 6.4% loss of current density after 50 h ( $100 \text{ mA cm}^{-2}$ )     | 47   |
| NiCoSe <sub>2</sub> /MoSe <sub>2</sub>   | -244 at $-50 \text{ mA cm}^{-2}$        | 13.1 $\Omega$                             | NA                            | 72 h                                                                     | 48   |

HER: hydrogen evolution reaction;  $\eta$ : overpotential;  $R_{\text{ct}}$  or  $R_{\text{ct1}}$  or  $R_{\text{ct2}}$ : charge transfer resistance; NA: not applicable.

ductivity on the activity of the HER or the OER, where the conductivity and activity of Nb<sub>4</sub>N<sub>5</sub> for the HER are higher than those of Fe<sub>4</sub>N/Nb<sub>4</sub>N<sub>5</sub>, whereas the conductivity and activity of Fe<sub>4</sub>N/Nb<sub>4</sub>N<sub>5</sub> for the OER are higher than those of Nb<sub>4</sub>N<sub>5</sub>. The electrochemical reconstruction processes in Nb<sub>4</sub>N<sub>5</sub> and Fe<sub>4</sub>N/Nb<sub>4</sub>N<sub>5</sub> are the underlying reason. Specifically, Nb<sub>4</sub>N<sub>5</sub> remains stable under HER conditions. However, under OER conditions, the electrochemical reconstruction of Nb<sub>4</sub>N<sub>5</sub> leads to the formation of inactive KNbO<sub>3</sub> species. Additionally, the presence of Fe ions in Fe<sub>4</sub>N/Nb<sub>4</sub>N<sub>5</sub> alters the electrochemical reconstruction process, resulting in the generation of highly active FeOOH species that contain high-valence Fe<sup>3+</sup> ions.

A comparison of charge transfer resistance across various photocatalysts provided by EIS reveals crucial insights into their charge separation efficiencies.<sup>5,7,54</sup> High activity in hydrogen evolution depends on a range of factors beyond just charge separation resistance.<sup>17,55,56</sup> These include effective charge separation and transfer, minimized recombination of photogenerated charges, rapid diffusion rates of charges, prolonged lifetimes of excitons, efficient use of visible light, appropriate bandgaps, and the optimization of charge distributions at interfaces. Additionally, ensuring efficient electron transport for proton reduction and providing active sites with low energy barriers for hydrogen production are essential.<sup>5,8,9,18,32,54</sup> Moon *et al.*<sup>7</sup> found that the Pt/g-C<sub>3</sub>N<sub>4</sub>/TiO<sub>2</sub>/IrO<sub>x</sub> (PCTI) hollow sphere photocatalyst demonstrated a lower charge transfer resistance ( $R_{\text{ct}} = 1.52 \text{ k}\Omega$ ) compared to that of other co-catalysts on g-C<sub>3</sub>N<sub>4</sub>/TiO<sub>2</sub>, indicating more rapid charge transfer in PCTI. This configuration significantly enhances the transient photocurrent density, which is attributed to efficient surface charge separation where photogenerated electrons are trapped by Pt on the CT's external surface and photoexcited holes are trapped by IrO<sub>x</sub> internally. This arrangement boosts the water redox reaction performance, illustrating that strategic spatial separation of co-catalysts on a Z-scheme heterostructure can accelerate surface reaction rates, ensuring effective charge separation and transfer, and thus, improving photocatalytic H<sub>2</sub> evolution. Tang *et al.*<sup>5</sup> observed that the bismuth-based perovskite heterojunction MA<sub>3</sub>Bi<sub>2</sub>I<sub>9</sub>/DMA<sub>3</sub>BiI<sub>6</sub>

(BBP-5; MA: methylammonium; DMA: dimethylammonium) achieved a remarkable H<sub>2</sub> evolution rate of  $198.2 \mu\text{mol h}^{-1} \text{ g}^{-1}$  and superior apparent quantum efficiency, outperforming its counterparts without noble metal co-catalysts under visible light ( $\lambda \geq 420 \text{ nm}$ ). BBP-5's lower charge transfer resistance ( $R_{\text{ct}}$ :  $1.15 \text{ k}\Omega \text{ cm}^2$ ) underscores its enhanced charge separation capabilities. This discovery suggests that the *in situ* formation of such heterojunctions can not only facilitate charge separation and reduce charge recombination but also extend exciton lifetimes and promote efficient electron transport for H<sup>+</sup> ion reduction, culminating in heightened photocatalytic H<sub>2</sub> evolution *via* HI splitting. Furthermore, Guo *et al.*<sup>17</sup> noted that the NiCo<sub>2</sub>S<sub>4</sub>/ZnIn<sub>2</sub>S<sub>4</sub> hollow nanocages (NCS/ZIS-HNCs) showcased a lower charge transfer resistance and a higher transient photocurrent density than those of ZIS-HNCs alone, indicating effective photogenerated electron migration at the NCS/ZIS-HNCs interface. Despite NCS exhibiting a lower charge resistance than NCS/ZIS-HNCs, its photocatalytic activity is negligible, whereas NCS/ZIS-HNCs exhibit a significantly higher H<sub>2</sub> production activity. This highlights that optimal photocatalytic performance relies on more than just low charge resistance, including the photothermal effect, which is markedly enhanced in NiCo<sub>2</sub>S<sub>4</sub>/ZnIn<sub>2</sub>S<sub>4</sub> without cooling water, achieving a H<sub>2</sub> evolution rate 5.98 times greater than that with cooling. This enhancement is attributed to improved electron dynamics facilitated by interfacial S atoms with vacancies, indicating that efficient carrier transfer and the localized surface plasmon resonance (LSPR) effect, leading to the generation of hot electrons, are critical for maximizing photocatalytic H<sub>2</sub> evolution.

This review embarks on an in-depth exploration of nanostructured electrocatalysts and photocatalysts, showcasing their pivotal roles in the hydrogen evolution reaction (HER) and the oxygen evolution reaction (OER) across varied environments including alkaline, acidic, and seawater media. It particularly emphasizes diverse material compositions such as oxides, hydroxides, sulfides, phosphides, nitrides, and innovative carbon-based, alloy, and metal-organic frameworks (MOFs), alongside the strategic implementations of heterojunctions, co-catalyst loading, defect engineering, and elemen-

tal doping. The narrative extends to evaluating photocatalytic strategies for hydrogen production, highlighting advancements in photocatalyst design for both water splitting and hydrohalic acid (HX) splitting, with a focus on the integration of lead-free and perovskite-based materials. Through this review, we aim to encapsulate the latest developments, draw comparisons among various catalytic systems, and identify promising directions for future research in the quest for efficient and sustainable hydrogen production technologies.

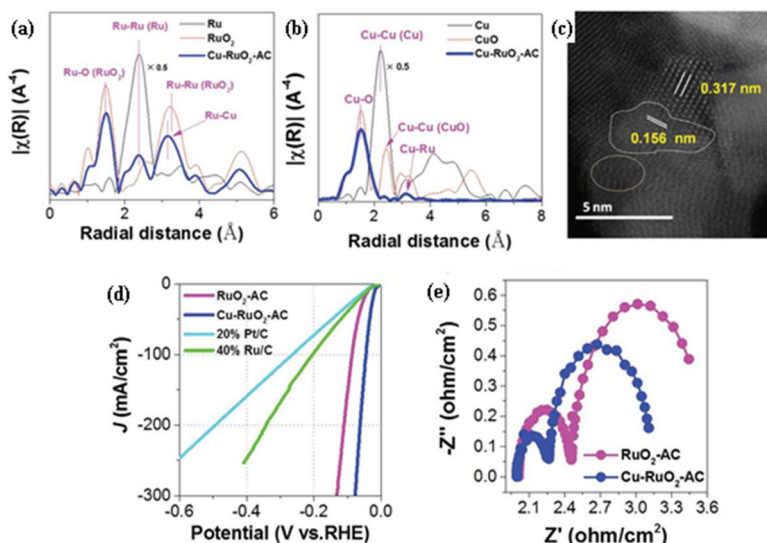
## 2. Water electrolysis in an alkaline environment

### 2.1. Oxide/hydroxide/layered double hydroxide based nano-electrocatalysts

Cu doping of  $\text{RuO}_2$  has been shown to facilitate the formation of metallic Ru active sites for the hydrogen evolution reaction (HER) during electrochemical activation, enhancing both conductivity and HER activity. Xi *et al.*<sup>4</sup> observed that Cu-doped  $\text{RuO}_2$ , prepared through a hydrolysis method followed by calcination at 300 °C for 2 h and electrochemically activated with five cycles of cyclic voltammetry (CV) from 0 to 1.03 V *versus* RHE for 7.5 min, exhibited significantly higher HER activity. The presence of  $\text{Cu}^{2+}$  sites, due to their lower valence state than that of  $\text{Ru}^{4+}$ , leads to oxygen vacancies and the disruption of Ru–O octahedra, thereby reducing adjacent  $\text{Ru}^{4+}$  under certain voltages during CV activation. This process generates  $\text{Ru}^0$  sites on the surface, as evidenced by the extended X-ray absorption fine structure (EXAFS) spectrum at the Ru K-edge, showing distinct interaction peaks for Ru–O, Ru–Ru, and Ru–Cu in Cu-doped  $\text{RuO}_2$ , indicative of altered Ru–O octahedral structures (Fig. 1(a)). Similarly, as shown

in Fig. 1(b), the EXAFS spectrum at the Cu K-edge revealed changes in Cu–O and the presence of Cu–Ru metallic coordination, confirming the activation effect of Cu doping. High-angle annular dark-field scanning transmission electron microscopy (HAADF-STEM) images further confirmed the formation of  $\text{Ru}^0$  and highlighted lattice distortions and defects post-activation (Fig. 1(c)). Density functional theory (DFT) calculations supported the activation mechanism, showing that  $\text{Cu}^{2+}$  doping activated adjacent  $\text{Ru}^{4+}$ . Notably, Cu-doped  $\text{RuO}_2$  displayed a significantly lower charge transfer resistance and higher electrochemical surface area for the HER compared to undoped  $\text{RuO}_2$ , underlining its superior activity and conductivity (Fig. 1(d)). In 1 M KOH, this catalyst achieved an overpotential of –19 mV at –10  $\text{mA cm}^{-2}$ , showcasing exceptional activity, and maintained negligible decay over 93 h, indicating remarkable stability. For overall water splitting in the same medium, the HER performance of Cu– $\text{RuO}_2$ //Cu– $\text{RuO}_2$  demonstrated very high activity, requiring potentials of 1.66 V at 100  $\text{mA cm}^{-2}$  and 1.78 V at 200  $\text{mA cm}^{-2}$ . This comprehensive analysis underscores the significant impact of Cu doping on enhancing the electrocatalytic performance of  $\text{RuO}_2$  for efficient hydrogen production.

Besides, Triolo *et al.*<sup>57</sup> found that the high-entropy spinel oxide  $(\text{Cr}_{1/5}\text{Mn}_{1/5}\text{Fe}_{1/5}\text{Co}_{1/5}\text{Ni}_{1/5})_3\text{O}_4$  nanofibers exhibited higher activity and lower charge transfer resistance for the OER than those of  $(\text{Cr}_{1/5}\text{Mn}_{1/5}\text{Fe}_{1/5}\text{Co}_{1/5}\text{Zn}_{1/5})_3\text{O}_4$  and  $(\text{Cr}_{1/5}\text{Mn}_{1/5}\text{Fe}_{1/5}\text{Ni}_{1/5}\text{Zn}_{1/5})_3\text{O}_4$ . This could be attributed to the  $e_g$  filling with nearer to its optimal value, lower outer 3d-electron number, and greater occupation of the 16d sites by the most redox-active species. Moreover, Liu *et al.*<sup>41</sup> observed that  $\text{FeCoCrCuO}_x$  exhibited higher activity and lower charge transfer resistance for the HER than those of  $\text{FeCoO}_x$ ,  $\text{FeCoCrO}_x$ , and  $\text{FeCoCuO}_x$ , and it exhibited high activity and stability for overall



**Fig. 1** (a) EXAFS spectrum obtained at the Ru K-edge of Cu– $\text{RuO}_2^{\text{activated}}$  in comparison with those of Ru and  $\text{RuO}_2$ ; (b) EXAFS spectrum obtained at the Cu K-edge of Cu– $\text{RuO}_2^{\text{activated}}$  in comparison with those of Cu and CuO; (c) HAADF-STEM image of Cu– $\text{RuO}_2^{\text{activated}}$ , where the oval pink area reveals the obvious lattice distortion; (d) LSV curves for the HER of Cu– $\text{RuO}_2^{\text{activated}}$  in comparison with those of  $\text{RuO}_2$ -AC, 20% Pt/C and 40% Ru/C; (e) EIS plot for the HER of Cu– $\text{RuO}_2^{\text{activated}}$  in comparison with that of  $\text{RuO}_2$ -AC obtained at –0.1 V *versus* RHE (reproduced with permission from ref. 4. Copyright 2023, Wiley-VCH GmbH).



water splitting in 1 M KOH. In addition, it only requires 3.18 V to achieve  $>30$  A ( $\sim 1.84$  A cm $^{-2}$ ) on a large-sized electrode (16.5 cm $^2$ ) for alkaline water electrolysis. The superior performance of FeCoCrCuO $_x$  is attributed to the synergistic interaction between multiple metals. Additionally, surface activation and robust Cr–OH interactions upon activation also play a crucial role in enhancing the HER performance.

Integration of the heterostructured nanosheet array with Ru nanoparticles could provide optimal metal–support interactions, enhance conductivity, improve orbital coupling, provide optimal adsorption energy with intermediates, and facilitate gas evolution, which could enhance the performance for the HER and OER. Liu *et al.*<sup>40</sup> fabricated a Ru–Ni $_3$ N/NiO catalyst with an optimal metal–support interaction, where Ru nanoparticles were uniformly dispersed on a 2D Ni $_3$ N/NiO heterostructured nanosheet array, and it exhibited significantly higher activity and stability for the HER and OER. Theoretical calculations suggested that the structural stability, orbital coupling, and electronic distributions of Ru–Ni $_3$ N/NiO could be enhanced by the Ni $_3$ N/NiO heterostructures when compared to those of Ru–Ni $_3$ N and Ru–NiO, leading to lower energy barriers and increased activity for the HER and OER. Ru–Ni $_3$ N/NiO demonstrates enhanced electrochemical performance, exhibiting lower charge transfer resistance ( $R_{ct}$ : 1.52  $\Omega$  cm $^{-2}$ ) and superior hydrogen evolution reaction (HER) activity compared to both Ru–Ni $_3$ N and Ru–NiO individually. In 1 M KOH, this catalyst achieves an impressive overpotential ( $\eta$ ) of 385 mV at a substantial current density of 1000 mA cm $^{-2}$  for the oxygen evolution reaction (OER), highlighting its remarkable activity. Similarly, for the HER under the same conditions, it reaches an overpotential of  $-190$  mV at a current density of  $-1000$  mA cm $^{-2}$ , further underscoring its exceptional performance. When applied to overall water splitting in 1 M KOH, Ru–Ni $_3$ N/NiO maintains a potential of 1.74 V at a current density of 1000 mA cm $^{-2}$ , reflecting its high activity. Moreover, it shows outstanding stability with negligible decay over 1000 h at a current density of 500 mA cm $^{-2}$ , demonstrating its robustness for long-term applications.

The activation of Ni(OH) $_2$  followed by electrolysis in 1 M KOH with the addition of Fe(III) could enable the formation of highly active species for the OER, which could enhance the kinetics for the OER in 1 M KOH with Fe(III) compared to that in 1 M KOH. Huang *et al.*<sup>16</sup> observed that Ni(OH) $_2$  exhibited a lower charge transfer resistance (Fig. 2(d) and (e)) and higher activity (Fig. 2(b)) for the OER in 1 M KOH with 500  $\mu$ M of Fe(III) than that in 1 M KOH. The Ni(OH) $_2$  nanosheet arrays were obtained on Ni foam through a one-step ultrasonic process, where Ni foam was sonicated for 30 min in 5 wt% HCl, while Ni(OH) $_2$  was activated in 1 M KOH or 1 M KOH with 500  $\mu$ M of Fe(III). Moreover, Fig. 2(a) depicts a schematic diagram of the synthesis process for the Ni(OH) $_2$  nanosheet arrays and subsequent structural evolution during the surface activation process by a chronopotentiometry test in 1 M KOH in the presence of 500  $\mu$ M Fe(III) at different time intervals. The morphology, structure, and *in/ex situ* electrochemical investigations suggest that Fe(III) acts as the active site and enhances the kinetics for the OER in the following two ways (Fig. 2(c)):

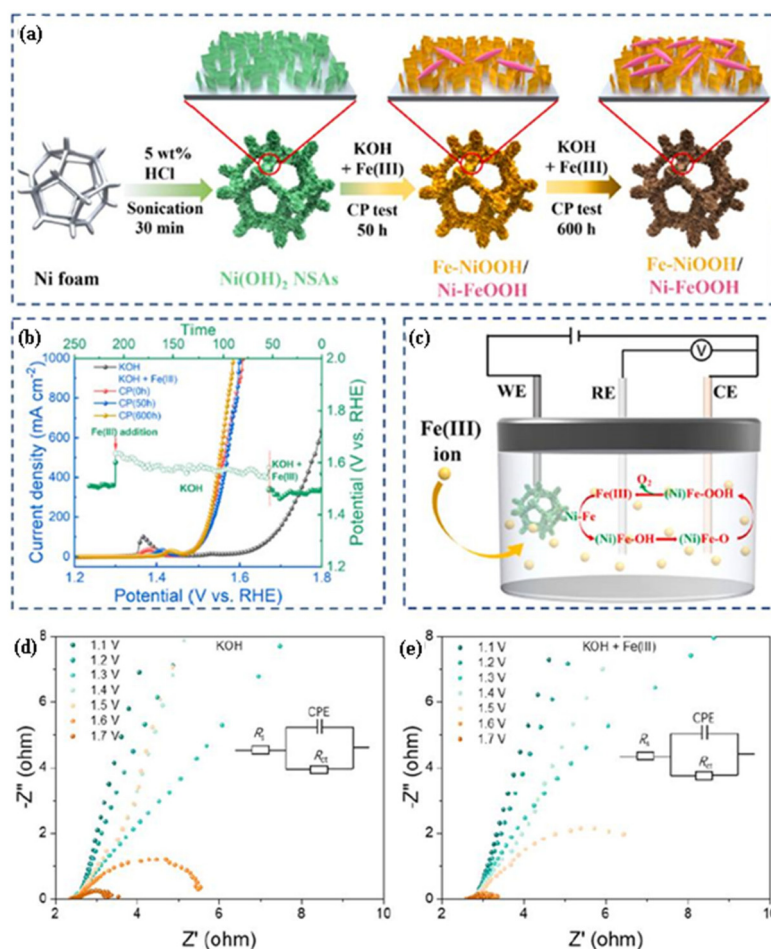
(i) the formation of several intermediates in a cyclic process on the electrode/electrolyte interface such as Fe(III), (Ni)Fe $^{x+}$ , (Ni)Fe–OH, (Ni)Fe–O, (Ni)Fe–OOH, followed by Fe(III) (+O $_2$ ); (ii) growth of the (oxy)hydroxide interface (Ni–FeOOH/Fe–NiOOH) on the surface of Ni(OH) $_2$ . In 1 M KOH, Ni(OH) $_2$  exhibits an impressive overpotential ( $\eta$ ) of 616 mV at a current density of 1000 mA cm $^{-2}$ , indicating its outstanding activity for the oxygen evolution reaction (OER). When supplemented with 500  $\mu$ M of Fe(III), the performance of Ni(OH) $_2$  is further improved, achieving an overpotential of 367 mV at the same current density, which not only highlights its enhanced efficiency but also its high stability under these conditions.

The loading of Ce single atoms on NiV LDH (LDH: layered double hydroxide) can tune the band gap from semiconductor to metallic behavior, to afford an optimal adsorption energy with intermediates, and this could enhance the performance for the OER. Zeng *et al.*<sup>58</sup> observed that Ce/NiV LDH exhibited a lower charge transfer resistance ( $R_{ct}$ : 0.6  $\Omega$ ), higher electrochemically active surface area, and higher activity than that of NiV LDH for the OER in 1 M KOH. The Ce/NiV LDH is composed of Ce single atoms, which are anchored on monolayer nickel–vanadium layered double hydroxide. Theoretical calculations and *in situ* electrochemical impedance spectroscopy investigations reveal that the immobilization of monatomic Ce on NiV LDH narrows the band gap, transforming its semiconducting properties into metallic behavior. This transformation strengthens the density of states near the Fermi level, providing an optimal adsorption energy with intermediates and reducing the charge transfer barrier. Consequently, this enhances the performance for the OER.

## 2.2. Sulfide/selenide based nano-electrocatalysts

The doping of Ni $_3$ S $_2$  with Co could enhance the conductivity and optimize the adsorption energy with intermediates, thereby improving the performance for the OER. Additionally, this doping can reduce the cell voltage for hydrogen production by substituting the OER with sulfion oxidation or urea oxidation, further enhancing the overall efficiency of the system. Li *et al.*<sup>59</sup> reported that Co-doped Ni $_3$ S $_2$  demonstrated superior activity and reduced charge transfer resistance ( $R_{ct}$ : 33.12  $\Omega$  or W cm $^{-2}$ ) for the OER compared to pure Ni $_3$ S $_2$ . This enhanced performance is achieved through the preparation of Co–Ni $_3$ S $_2$  on Ni mesh, involving an electrodeposition step followed by a hydrothermal-sulfurization reaction. Furthermore, the utilization of a sulfion-containing electrolyte has enabled a significant reduction in the cell voltage to 0.80 V at 50 mA cm $^{-2}$ , which is about 1.20 and 0.94 V lower than that of water splitting and urea electrolysis, respectively. The enhanced sulfion oxidation is attributed to direct electron transfer to the surface active sites, while the improved urea oxidation is ascribed to the formation of a suitable oxidation state of the Ni $^{III}$  on the surface, which facilitates a shortened oxidation pathway.

The construction of bimetallic sulfide nanoarray surfaces has been shown to promote gas evolution, improve conductivity, and subsequently boost the HER and OER performance. Findings from He *et al.*<sup>60</sup> showed that the FeNiS nanoarray exhibited higher activity and lower charge transfer resistance



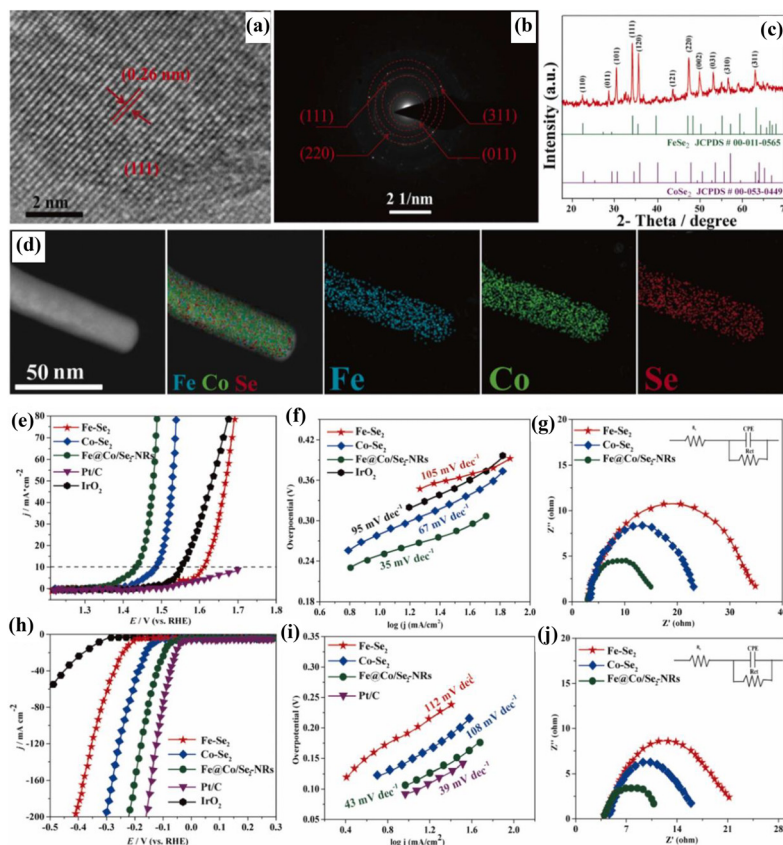
**Fig. 2** (a) Schematic diagram representing the synthesis process for Ni(OH)<sub>2</sub> nanosheet arrays and their subsequent structural evolution during a chronopotentiometry test in 1 M KOH in the presence of 500 μM Fe(III) at different time intervals. (b) The LSV and chronopotentiometry plot for the OER of Ni(OH)<sub>2</sub> in KOH with Fe(III) in comparison with those of Ni(OH)<sub>2</sub> in KOH. (c) Proposed mechanism for the OER on Ni(OH)<sub>2</sub> in alkaline electrolyte with Fe(III); CE, RE, and WE are the counter electrode, reference electrode, and working electrode, respectively. Nyquist plot for the OER of (d) Ni(OH)<sub>2</sub> in KOH with Fe(III) in comparison with that of (e) Ni(OH)<sub>2</sub> in KOH (reproduced with permission from ref. 16. Copyright 2023, Elsevier B.V.).

( $R_{ct}$ : 17.1  $\Omega$  cm<sup>-2</sup>) for the OER than those of NiS, while it exhibited high activity and stability for overall water electrolysis in 1 M KOH. The FeNiS nanoarray is composed of FeS nanosheets, which are anchored on Ni<sub>9</sub>S<sub>8</sub> nanorods. Their phase-field simulations revealed that the flow field of the arrayed electrode surface focused the shear force directly onto the evolved bubbles, leading to enhancement in gas evolution, which could reduce the shielding effect by passage of flow through channels and voids. Moreover, the flow field exerts a more significant influence on the arrayed surface compared to a nanoscale smooth surface, effectively reducing the size of the evolving bubbles.

Integration of a transition metal phosphide with sulfide could enhance the conductivity, to afford an optimal adsorption energy with intermediates, and this could enhance the performance for the HER and OER. It was demonstrated by Dhakal *et al.*<sup>46</sup> that 1T-Mn-VS<sub>2</sub>@Co<sub>2</sub>P exhibited higher activity and lower charge transfer resistance ( $R_{ct}$ : 2  $\Omega$ ) for the HER than those of 1T-VS<sub>2</sub>@Co<sub>2</sub>P, while it exhibited remarkable

activity and stability for overall water electrolysis in 1 M KOH. It is composed of a HER-active hierarchical 2D–2D tangled 3D heterostructure of Mn-doped 1T VS<sub>2</sub>, which is assembled with OER-active cobalt phosphide.

Incorporation of transition metal multi-cation species into nanostructured selenides could modify the electronic structure and enhance the conductivity, which could boost the performance for the HER and OER. Ibraheem *et al.*<sup>14</sup> observed that the Fe@Co/Se<sub>2</sub> nanorods exhibited higher activity and lower charge transfer resistance (for the HER,  $R_{ct}$ : 10.80  $\Omega$  (Fig. 3j); for the OER,  $R_{ct}$ : 15.04  $\Omega$  (Fig. 3g)) than those of Fe–Se<sub>2</sub> and Co–Se<sub>2</sub> for the HER and OER in 1 M KOH. Fe@Co/Se<sub>2</sub> contains FeSe<sub>2</sub> and CoSe<sub>2</sub> phases (Fig. 3(c)), while the HRTEM image (Fig. 3(a)) and the corresponding SAED pattern (Fig. 3(b)) depict inter-planar distances of about 2.6 Å, 1.40 Å, 1.86 Å, and 3.07 Å consistent with the (1 1 1), (3 1 1), (2 2 0) and (0 1 1) lattice planes of the FeSe<sub>2</sub> and CoSe<sub>2</sub> phases. It contains Fe, Co, and Se, which are homogeneously distributed (Fig. 3(d)). By means of a Fe-coordinated Co-bridged bond, the Fe<sup>2+/3+</sup>/Co<sup>2+</sup>



**Fig. 3** (a) HRTEM image, (b) the corresponding SAED pattern, (c) XRD pattern, and (d) HAADF-STEM image (with elemental mapping images of Fe, Co and Se) of Fe@Co/Se<sub>2</sub>; (e) LSV curve, (f) Tafel slope, and (g) Nyquist plot for the OER of Fe@Co/Se<sub>2</sub> in comparison with those of Fe–Se<sub>2</sub> and Co–Se<sub>2</sub> in 1 M KOH; (h) LSV curve, (i) Tafel slope, and (j) Nyquist plot for the HER of Fe@Co/Se<sub>2</sub> in comparison with those of Fe–Se<sub>2</sub> and Co–Se<sub>2</sub> in 1 M KOH (reproduced with permission from ref. 14. Copyright 2021, Elsevier B.V.).

species was highly coupled with selenide, suggesting a modified electronic structure.

Creating nanostructured heterojunctions from transition metal selenides can alter their electronic structures and improve conductivity, thereby enhancing their efficiency in both the hydrogen evolution reaction (HER) and the oxygen evolution reaction (OER). Meshesha *et al.*<sup>48</sup> observed that NiCoSe<sub>2</sub>/MoSe<sub>2</sub> exhibited higher activity and lower charge transfer resistance ( $R_{ct}$ : 13.1  $\Omega$ ) than those of NiCoSe<sub>2</sub> and MoSe<sub>2</sub> for the HER in 1 M KOH, while it exhibited high activity and stability for overall water electrolysis. The enhanced performance is attributed to the formation of a heterojunction of nickel–cobalt selenide with metal phase MoSe<sub>2</sub> nanospheres.

Wang *et al.*<sup>43</sup> observed that ultrathin P/Fe–NiSe<sub>2</sub> nanosheets exhibited higher activity and lower charge transfer resistance ( $R_{ct}$ : 4.78  $\Omega$ ) than those of Fe–NiSe<sub>2</sub> for the HER in 1 M KOH, while they exhibited high efficiency for hydrazine-assisted water splitting and a Zn-hydrazine battery, achieving a remarkable current density of 100 mA cm<sup>−2</sup> at potentials of −168 and 200 mV for the HER and HzOR, respectively. P/Fe–NiSe<sub>2</sub> has a much lower apparent activation energy value (38.9 kJ mol<sup>−1</sup>) than those of Fe–NiSe<sub>2</sub> (68.3 kJ mol<sup>−1</sup>) and NiSe<sub>2</sub> (74.9 kJ mol<sup>−1</sup>), indicating that the introduction of P atoms can signifi-

cantly reduce the kinetic barrier and facilitate the HER process. This superior performance is credited to the ultrathin nanosheet structure and the synergistic effect of P/Fe codoping in NiSe<sub>2</sub>, which collectively enhance the electronic structure, conductivity, and interactions with intermediates.

### 2.3. Phosphide/nitride based electrocatalysts

The incorporation of high-valence Zr into nickel phosphide can optimize the adsorption energy with intermediates and enhance conductivity, improving the performance for the HER and OER. Fan *et al.*<sup>42</sup> observed that Zr<sub>0.1</sub>Ni<sub>1.9</sub>P exhibited higher activity and lower charge transfer resistance ( $R_{ct}$ : 5  $\Omega$ ) than those of Ni<sub>2</sub>P for the HER in 1 M KOH, while it exhibited high activity and stability for overall water electrolysis. Their DFT calculations reveal that the incorporation of high-valence Zr can reduce the free energy change for the potential-determining step ( $O^* \rightarrow OOH^*$ ) to enhance the alkaline OER, and it can improve the dissociation and adsorption processes of H<sub>2</sub>O to facilitate the HER process.

The synergistic effect of bimetallic phosphide could enhance the conductivity and provide an optimal adsorption energy with intermediates, which could enhance the performance for the HER. Xiang *et al.*<sup>47</sup> observed that CoP<sub>3</sub>–Nb<sub>2</sub>P/PCC

exhibited a significantly higher activity ( $\eta$  of  $-375$  mV at a huge current density of  $-1000$  mA cm $^{-2}$ ) and lower charge transfer resistance ( $R_{ct}$ :  $3.5$   $\Omega$ ) than those of CoO–NbO/PCC for the HER in  $1$  M KOH (PCC: plasma-defect-engineered carbon cloth). Their *in situ* Raman spectroscopy and DFT calculations reveal that the outstanding performance for the HER is attributed to the synergistic interaction between Co and Nb phosphide. Furthermore, the Nb–Nb vibration weakens the P–H $_{ads}$  bond, ultimately leading to efficient H $_2$  production.

Ir single atom incorporation into and Zr doping of cobalt phosphide could modify the electronic structure, improve the conductivity, optimize the adsorption energy with intermediates, and enhance the performance for the HER and OER. Ngo *et al.*<sup>61</sup> observed that Ir@Zr–CoP exhibited higher activity and lower charge transfer resistance ( $R_{ct}$ :  $1.97$   $\Omega$ ) than those of Zr–CoP for the OER in  $1$  M KOH, while it exhibited a cell voltage of  $1.88$  V at a huge current density of  $1000$  mA cm $^{-2}$  for a large, scalable, anion exchange membrane water electrolyzer with high stability. It was obtained by hydrothermal treatment of carbon cloth followed by wet impregnation and phosphidation (Fig. 4(a)).

Furthermore, creating nanoporous bimetallic alloy/OOH and metal oxynitride heterolamellas enhances the OER performance by exposing numerous active sites, improving conductivity, optimizing the adsorption energy with intermediates, and facilitating gas evolution. Zeng *et al.*<sup>21</sup> synthesized FeCo/CeO $_{2-x}$ N $_x$  through a method involving alloying and thermal nitridation. FeCo/CeO $_{2-x}$ N $_x$  exhibits a significantly higher activity ( $\eta$  of  $360$  mV at a huge current density of  $3940$  mA cm $^{-2}$ ) and lower charge transfer resistance ( $R_{ct}$ :  $1.3$   $\Omega$ ;  $R_p$ :  $1.6$   $\Omega$ ) than those of FeCo/Ce–O for OER in  $1$  M KOH, while it delivers  $\sim 1900$  mA cm $^{-2}$  for about  $1000$  h, suggesting its robust stability.

The insulating hexagonal boron nitride with an optimal thickness (one-atom-thick) can be used as a co-catalyst to enhance the activity for the OER. Lu *et al.*<sup>62</sup> reported that hexagonal BN–NiFeO $_x$ H $_y$  exhibited a significantly higher activity ( $\eta$  of  $230$  mV at  $10$  mA cm $^{-2}$ ;  $\eta$  of  $490$  mV at a huge current density of  $2000$  mA cm $^{-2}$ ) for the OER in  $1$  M KOH compared to NiFeO $_x$ H $_y$ , and it afforded a low charge transfer resistance for the OER. The high activity of hexagonal BN–NiFeO $_x$ H $_y$  is attributed to the covering of NiFeO $_x$ H $_y$  with one-atom-thick hexagonal boron nitride as a co-catalyst, where the oxygen-containing intermediates can adsorb insulating boron nitride, and localized electrons can enhance the deprotonation processes at the electrode. However, the catalytic activity of multi-layer hexagonal boron nitride is significantly reduced due to its insulating properties.

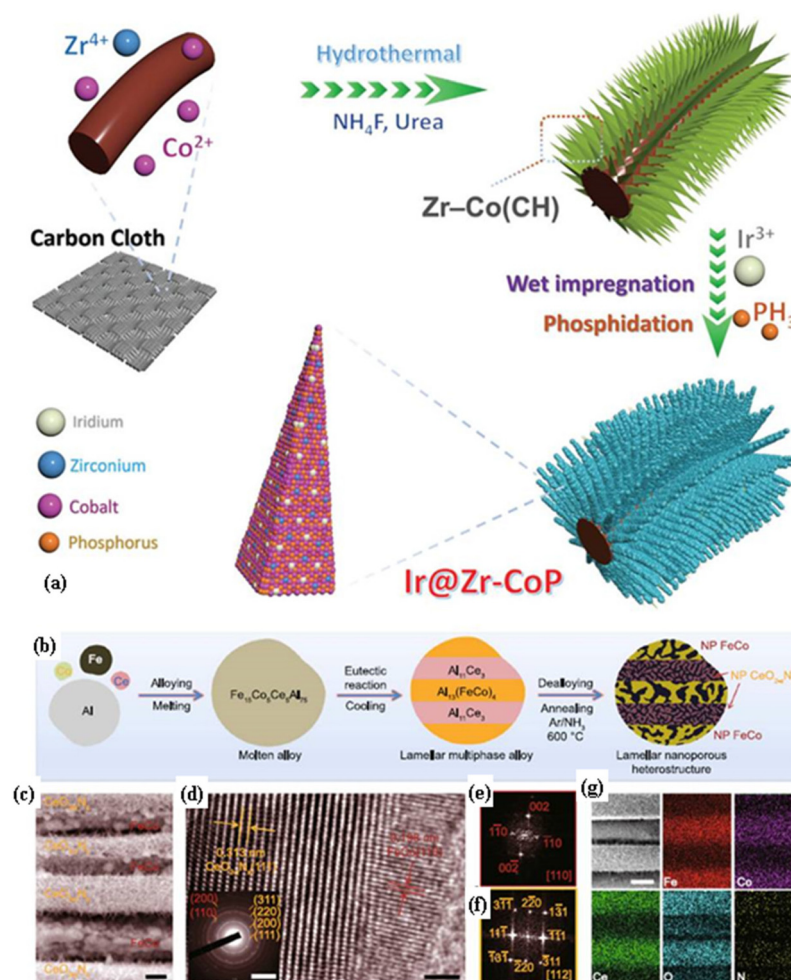
#### 2.4. Carbon/alloy/MOF/organic based nano-electrocatalysts

Fabricating pyrrolic-N-doped graphitized porous carbon loaded with Ru nanoparticles could enhance the conductivity and provide an optimal adsorption energy with intermediates, and boost the performance for the HER. Shin *et al.*<sup>63</sup> observed that Ru–N-GPC (pyrrolic-N-doped graphitized porous carbon loaded with Ru nanoparticles) exhibited a much higher activity

( $\eta$  of  $-9.6$  mV at  $-10$  mA cm $^{-2}$ ) and lower charge transfer resistance than those of Ru–N-rGO (rGO: reduced graphene oxide) for the HER in  $1$  M KOH; it was obtained by pyrolyzing g-C $_3$ N $_4$  in the presence of Mg metal. The improvement in the Tafel step can be attributed to the strengthened coupling between the Ru nanoparticles and the pyrrolic-N moieties present in the support, which promote an optimal adsorption energy with intermediates on the Ru nanoparticles.

Cu doping can amplify the Kirkendall effect, and the fabrication of double-shell hollow Cu-doped CoP nanoparticles anchored to carbon arrays can optimize the adsorption energy with intermediates and improve the conductivity, thereby enhancing the HER and OER performance. Wang *et al.*<sup>44</sup> found that DH–CuCo–P@NC/CC outperformed H–CuCo–P@NC/CC in terms of the HER activity and charge transfer resistance ( $R_{ct}$ :  $2.4$   $\Omega$ ) in  $1$  M KOH; it also showed notable activity for water electrolysis. This improvement is attributed to the structure of DH–CuCo–P@NC/CC, featuring double-shell hollow Cu-doped CoP nanoparticles supported on carbon arrays, which intensify the Kirkendall effect and optimize the adsorption energy and conductivity for better HER and OER outcomes. Furthermore, integrating high-polarity F, CoFe alloy, and N-doped carbon nanotubes enhances the conductivity and adsorption with intermediates, boosting the OER performance. Kuang *et al.*<sup>64</sup> demonstrated that CoFe@NCNTs–F exhibited superior activity and lower charge transfer resistance ( $R_{ct}$ :  $31$   $\Omega$ ) compared to CoFe@NCNTs in  $1$  M KOH. The enhanced properties are due to the CoFe alloy/fluoride composition interconnected by bamboo-like N-doped carbon nanotubes, which improve the conductivity and adsorption, thus elevating the OER efficiency. Constructing Mn-doped bimetallic hydroxide nanostructures intertwined with carbon nanotubes modifies the electronic structure and optimizes the adsorption energy, leading to increased conductivity and improved OER performance. Wu *et al.*<sup>65</sup> observed that  $1.5\text{Mn–NiCo}$  HNS/CNT showed higher activity and lower charge transfer resistance than both NiCo HNS/CNT and  $1.5\text{Mn–NiCo}$  HNS in  $1$  M KOH. The composition of Mn-doped NiCo hydroxide nanosheets intertwined with carbon nanotubes enhances the electronic structure and adsorption energy, significantly improving the OER efficiency. The fabrication of nanostructured transition bimetallic alloy/OOH heterostructures can offer complex synergistic effects, rapid mass and charge transfer, and abundant exposed active sites, leading to superior performance for the HER and OER. Lv *et al.*<sup>45</sup> observed that cNiFe/aNiFeOOH@NiMo (a: amorphous; c: crystalline) exhibited higher activity and lower charge transfer resistance (for the HER,  $R_{ct}$ :  $0.946$   $\Omega$ ; for the OER,  $R_{ct}$ :  $0.962$   $\Omega$ ; Table 2) than those of cNiFe/aNiFeOOH and NiMo for the HER and OER in  $1$  M KOH. It was prepared by a two-step electrodeposition process. cNiFe/aNiFeOOH@NiMo is composed of ultrafine NiFe nanocrystals with numerous edges and exceptional electronic conductivity, amorphous NiFeOOH enriched with defects and nanosized pores, and NiMo alloy as a firm grasp for nanostructured cNiFe/aNiFeOOH, which could provide complex synergistic effects,





**Fig. 4** (a) Schematic diagram illustrating the synthesis process for Ir@Zr-CoP. (b) Schematic diagram with microstructures depicting the synthesis process for the lamellar nanoporous FeCo/CeO<sub>2-x</sub>N<sub>x</sub> composite electrode. (c) SEM image of the FeCo/CeO<sub>2-x</sub>N<sub>x</sub> composite (scale bar: 200 nm), where the hierarchical nanoporous FeCo alloy and nanoporous CeO<sub>2-x</sub>N<sub>x</sub> lamellas alternate and are periodically arranged. (d) HRTEM image of FeCo/CeO<sub>2-x</sub>N<sub>x</sub> depicts the heterostructured interface with a scale bar of 1 nm, while the inset shows the corresponding SAED patterns with a scale bar of 5 nm<sup>-1</sup>. FFT patterns of (e) FeCo alloy and (f) CeO<sub>2-x</sub>N<sub>x</sub> obtained from the selected areas in the FeCo/CeO<sub>2-x</sub>N<sub>x</sub> heterostructure in (d). (g) SEM image with the corresponding EDS elemental mapping images of FeCo/CeO<sub>2-x</sub>N<sub>x</sub> for N, O, Ce, Fe, and Co elements with a scale bar of 200 nm (reproduced with permission from ref. 21. Copyright 2023, Springer Nature).

fast mass and charge transfer, and expose/provide abundant active sites, leading to enhanced performance for the HER and OER.

Cr doping of FeMOF with Ru nanoclusters could manipulate the d orbitals of the Ru sites and enhance the conductivity, which could enhance the performance for the HER and OER. Zhao *et al.*<sup>66</sup> observed that Ru@Cr-FeMOF (MOF: metal-organic frameworks) exhibited higher activity and lower charge transfer resistance than those of Ru@FeMOF for the HER and OER in 1 M KOH, while it exhibited a cell voltage of 1.72 V at a huge current density of 1000 mA cm<sup>-2</sup> for overall water electrolysis, suggesting its excellent activity. Ru@Cr-FeMOF is composed of Ru nanoclusters anchored onto Cr-doped Fe-metal-organic frameworks. This unique structure enhances electron transfer and enables the manipulation of the highest occupied d orbitals and d-band centers

of the Ru sites, thereby boosting the catalytic activity for both the HER and OER.

An organic coating on Ni foam can significantly improve the conductivity and optimize the adsorption energy with intermediates, thereby boosting the hydrogen evolution reaction (HER) activity. Das *et al.*<sup>67</sup> found that Ni foam coated with a dual-layer of *N,N*-dimethylaminopropylamine (DMAPA) and 2-(dimethylamino)ethyl acrylate (DMAEA) outperformed the HER performance of bare Ni foam in 1 M KOH, exhibiting a superior current density of 1270 mA cm<sup>-2</sup> (at -0.9 V), an overpotential value of -250 mV (at 100 mA cm<sup>-2</sup>) and reduced charge transfer resistance. This enhancement is attributed to the conductivity and adsorption properties of the DMAPA and DMAEA coating, which provides a non-adhesive superaerophobicity with a nucleated bubble adhesion force of 4.6 ± 0.3 μN, which is favorable for the HER.

**Table 2** Overpotential ( $\eta$ ) to achieve  $10 \text{ mA cm}^{-2}$  for the oxygen evolution reaction (OER) of various electrocatalysts in 1 M KOH alkaline electrolyte ( $\eta$  to achieve greater than  $-10 \text{ mA cm}^{-2}$  are indicated with their corresponding current densities; electrolyte other than 1 M KOH is indicated on the electrocatalyst), where  $R_{\text{ct}}$  or  $R_{\text{ct1}}$  and  $R_{\text{ct2}}$  or  $R_{\text{p}}$  ( $R_{\text{ct}}$  or  $R_{\text{ct1}}$  or  $R_{\text{ct2}}$ : charge transfer resistance;  $R_{\text{p}}$ : pore resistance) are acquired from electrochemical impedance spectroscopy (EIS)

| OER electrocatalysts                                                                                                     | $\eta$ at $10 \text{ mA cm}^{-2}$<br>(mV) | $R_{\text{ct}}$ or $R_{\text{ct1}}$        | $R_{\text{p}}$ or<br>$R_{\text{ct2}}$ | Stability                                                                        | Ref. |
|--------------------------------------------------------------------------------------------------------------------------|-------------------------------------------|--------------------------------------------|---------------------------------------|----------------------------------------------------------------------------------|------|
| cNiFe/aNiFeOOH@NiMo                                                                                                      | 133.2 at $20 \text{ mA cm}^{-2}$          | $0.962 \Omega$                             | NA                                    | 100 h ( $100 \text{ mA cm}^{-2}$ )                                               | 45   |
| Ru-Ni <sub>3</sub> N/NiO                                                                                                 | 181                                       | NA                                         | NA                                    | 1000 h ( $500 \text{ mA cm}^{-2}$ )                                              | 40   |
| FeNiS                                                                                                                    | 191                                       | $17.1 \Omega \text{ cm}^{-2}$              | NA                                    | 550 h ( $500 \text{ mA cm}^{-2}$ )                                               | 60   |
| Fe@Co/Se <sub>2</sub>                                                                                                    | 200                                       | $15.04 \Omega$                             | NA                                    | 20 h ( $10 \text{ mA cm}^{-2}$ )                                                 | 14   |
| Ce/NiV LDH                                                                                                               | 209                                       | $0.6 \Omega$                               | NA                                    | 20 h ( $50 \text{ mA cm}^{-2}$ )                                                 | 58   |
| Fe <sub>4</sub> N/Nb <sub>4</sub> N <sub>5</sub>                                                                         | 224                                       | $0.206 \Omega$                             | NA                                    | NA                                                                               | 39   |
| Hexagonal BN-NiFeO <sub>x</sub> H <sub>y</sub>                                                                           | 230                                       | NA                                         | NA                                    | Current density varies less than 10% over 150 h<br>( $2000 \text{ mA cm}^{-2}$ ) | 62   |
| CoFe@NCNTs-F                                                                                                             | 231                                       | $31 \Omega$                                | NA                                    | 20 h ( $10 \text{ mA cm}^{-2}$ )                                                 | 64   |
| FeCoCrCuO <sub>x</sub>                                                                                                   | 255                                       | NA                                         | NA                                    | 165 h ( $500 \text{ mA cm}^{-2}$ )                                               | 41   |
| Ni(OH) <sub>2</sub> Fe(III)                                                                                              | 269 at $100 \text{ mA cm}^{-2}$           | NA                                         | NA                                    | 600 h ( $1000 \text{ mA cm}^{-2}$ )                                              | 16   |
| Ir@Zr-CoP                                                                                                                | 292                                       | $1.97 \Omega$                              | NA                                    | 50 h ( $100 \text{ mA cm}^{-2}$ )                                                | 61   |
| FeCo/CeO <sub>2-x</sub> N <sub>x</sub>                                                                                   | 297 at $1000 \text{ mA cm}^{-2}$          | $3.3 \Omega$                               | $1.6 \Omega$                          | 1000 h ( $1900 \text{ mA cm}^{-2}$ )                                             | 21   |
| Cu-RuO <sub>2</sub> activated                                                                                            | 316 at $100 \text{ mA cm}^{-2}$           | NA                                         | NA                                    | NA                                                                               | 4    |
| (Cr <sub>1/5</sub> Mn <sub>1/5</sub> Fe <sub>1/5</sub> Co <sub>1/5</sub> Ni <sub>1/5</sub> ) <sub>3</sub> O <sub>4</sub> | 360                                       | NA                                         | NA                                    | 3 h ( $40 \text{ mA cm}^{-2}$ )                                                  | 57   |
| Co-Ni <sub>3</sub> S <sub>2</sub> 1 M NaOH                                                                               | 440                                       | $33.12 \Omega$ or $\Omega \text{ cm}^{-2}$ | NA                                    | 24 h ( $50 \text{ mA cm}^{-2}$ )                                                 | 59   |

OER: oxygen evolution reaction;  $\eta$ : overpotential;  $R_{\text{ct}}$  or  $R_{\text{ct1}}$  or  $R_{\text{ct2}}$ : charge transfer resistance;  $R_{\text{p}}$ : pore resistance; NA: not applicable.

### 3. Water electrolysis in an acidic environment

#### 3.1. Oxide based nano-electrocatalysts

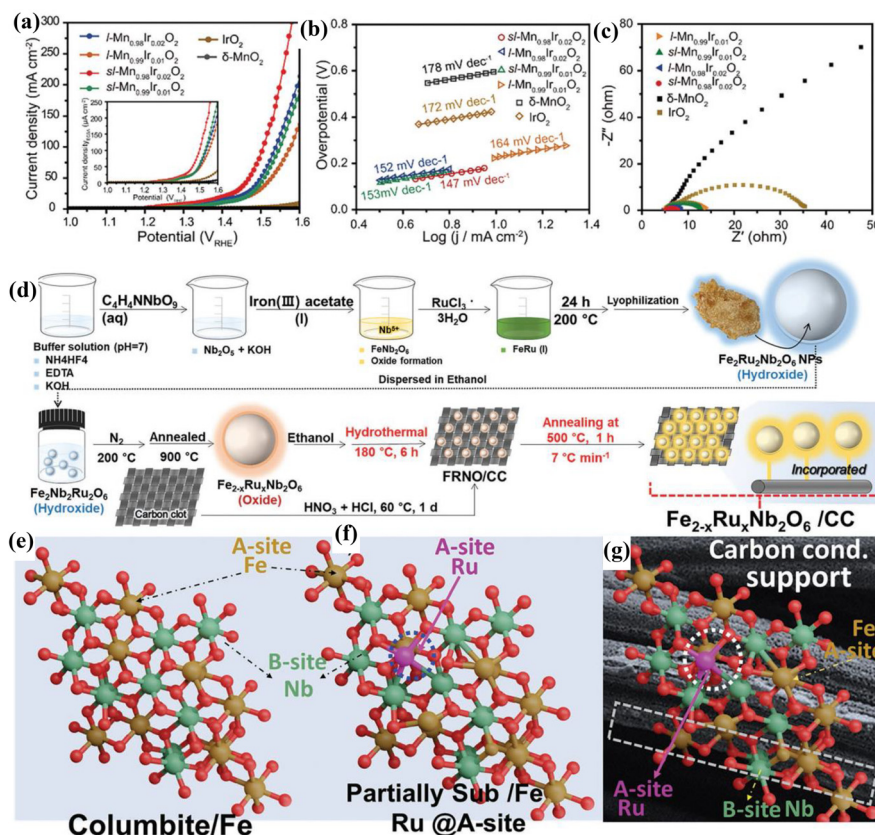
The fabrication of strained Ir incorporated into MnO<sub>2</sub> could modify the electronic structure, facilitate the conductivity, and provide an optimal adsorption energy with intermediates, which could enhance the performance for the OER. Zhao *et al.*<sup>1</sup> successfully prepared strained-Mn<sub>0.98</sub>Ir<sub>0.02</sub>O<sub>2</sub> through a hydrothermal method and observed that it exhibited higher activity ( $\eta$  of 105 mV at  $10 \text{ mA cm}^{-2}$  (Fig. 5(a)); mass activity: 300 mV at  $5681 \text{ A g}^{-1}$ ) and lower charge transfer resistance (Fig. 5(c)) than those of unstrained-Mn<sub>0.98</sub>Ir<sub>0.02</sub>O<sub>2</sub> for the OER in 0.5 M H<sub>2</sub>SO<sub>4</sub>, while it exhibited high stability for 168 h. The introduction of Ir into MnO<sub>2</sub> induces strain, resulting in an elongated Mn–O bond length and a shift of the d-band centre closer to the Fermi level. This modified electronic structure facilitates the transfer of holes from the O 2p orbitals to the catalyst surface, accelerating charge transfer and enhancing the catalytic activity for the OER.

Similarly, doping Fe<sub>2</sub>Nb<sub>2</sub>O<sub>6</sub> with Ru, as detailed by Bacirhonde *et al.*,<sup>26</sup> alters its electronic structure to enhance both the hydrogen evolution reaction (HER) and OER performances. The product Fe<sub>2-x</sub>Ru<sub>x</sub>Nb<sub>2</sub>O<sub>6</sub> (FRNO) demonstrates superior activity (with an overpotential of  $-30 \text{ mV}$  at  $-10 \text{ mA cm}^{-2}$  for the HER and  $200 \text{ mV}$  at  $10 \text{ mA cm}^{-2}$  for the OER) compared to its undoped counterparts in 0.5 M H<sub>2</sub>SO<sub>4</sub>, alongside lower charge transfer resistances, as shown in Tables 3 and 4. This enhancement is credited to a low-temperature hydrogel crosslinking and hydrothermal encapsulation process, which not only boosts intrinsic conductivity but also optimizes the adsorption

energy and facilitates water dissociation on its surface through effective charge transfer. DFT calculations and *operando* X-ray adsorption studies reveal that Fe<sub>2-x</sub>Ru<sub>x</sub>Nb<sub>2</sub>O<sub>6</sub>, with its high intrinsic conductivity, enhances the adsorption and dissociation of H<sub>2</sub>O on its surface. This is achieved by modulating the charge distribution, which facilitates charge transfer to the coordinated surface oxygen. This can stabilize the adsorption of the \*OOH intermediate on Ru and Fe sites, and further enhances the activity for the OER.

Developing acid stable OER electrocatalysts for high performance water splitting is of great significance but still remains challenging. To this end, a Ru/MnO<sub>2</sub> electrocatalyst where patches of Ru atom arrays were supported on  $\alpha$ -MnO<sub>2</sub> was prepared by Lin *et al.*<sup>68</sup> The catalyst follows an oxygen evolution mechanism involving only \*O and \*OH species as intermediates, which enables direct O–O radical coupling for O<sub>2</sub> evolution. This results in high activity ( $161 \text{ mV}$  at  $10 \text{ mA cm}^{-2}$ ) and outstanding stability, with minimal degradation after 200 hours of operation, making Ru/MnO<sub>2</sub> one of the best-performing acid-stable OER catalysts. *Operando* vibrational and mass spectroscopy measurements validate the reaction intermediates and gaseous products, confirming the OER pathway. First-principles calculations revealed a cooperative catalysis mechanism with a reduced energy barrier. Additionally, time-dependent elemental analysis demonstrated an *in situ* dynamic cation exchange reaction during the OER, which triggered the reconstruction of Ru atoms into a highly durable ordered array.

Besides, research by Chen *et al.*<sup>69</sup> indicates that Ru@MoO(S)<sub>3</sub> exhibits higher activity and lower charge transfer resistance than those of MoO(S)<sub>3</sub> in acidic medium for the HER and OER. In this structure, Ru nanoparticles are



**Fig. 5** (a) LSV curves (inset plot depicts LSV curves normalized by ECSA), (b) corresponding Tafel slopes, and (c) Nyquist plots of  $\text{Ir}_{0.98}\text{Mn}_{0.02}\text{O}_2$  and its counterparts in 0.5 M  $\text{H}_2\text{SO}_4$  (reproduced with permission from ref. 1. Copyright 2023, Wiley-VCH GmbH). (d) Schematic diagram depicting the synthesis process for  $\text{Fe}_{2-x}\text{Ru}_x\text{Nb}_2\text{O}_6/\text{carbon cloth}$ . Chemical structures of (e)  $\text{FeNb}_2\text{O}_6$ , (f)  $\text{Fe}_{2-x}\text{Ru}_x\text{Nb}_2\text{O}_6$ , and (g)  $\text{Fe}_{2-x}\text{Ru}_x\text{Nb}_2\text{O}_6/\text{carbon cloth}$  depicting the site occupancies in columbite (reproduced with permission from ref. 26. Copyright 2023, Wiley-VCH GmbH).

harmoniously encapsulated within a sulfur-modulated molybdenum oxide ( $\text{MoO}(\text{S})_3$ ) shell. This all-in-one integration enables Ru to maintain its intrinsic catalytic activity and rapid mass transfer capabilities while preventing electrochemical oxidation, thereby ensuring recyclability. Notably,  $\text{Ru}@\text{MoO}(\text{S})_3$  facilitates water splitting in 0.5 M  $\text{H}_2\text{SO}_4$  and 1 M KOH at a current density of  $10 \text{ mA cm}^{-2}$ , requiring cell voltages of only 1.522 V and 1.526 V, respectively, with nearly 100% faradaic efficiency. Moreover,  $\text{Ru}@\text{MoO}(\text{S})_3$  also demonstrates satisfactory stability during a 24 h continuous water splitting test. In another study, Wu *et al.*<sup>70</sup> synthesized La-doped  $\text{RuO}_2$  ( $\text{La-RuO}_2$ ) nanocrystals with oxygen vacancies.  $\text{La-RuO}_2$  exhibits a remarkable electrocatalytic performance, with ultra-low overpotentials of 208 mV for the OER and 71 mV for the HER at a current density of  $10 \text{ mA cm}^{-2}$ , and demonstrates robust long-term stability in 0.5 M  $\text{H}_2\text{SO}_4$ . Furthermore, the  $\text{La-RuO}_2$  nanocrystals require only 1.53 V to drive overall water splitting, maintaining exceptional stability. DFT calculations reveal that the synergistic effects of La doping and oxygen vacancy engineering tune the d-band center of Ru active sites and optimize the Gibbs free energy of intermediate states, enhancing the catalytic activity of  $\text{RuO}_2$ .

In a recent study, Song *et al.*<sup>71</sup> utilized a lattice-matching strategy to construct a  $\text{RuO}_2\text{-CeO}_2$  electrocatalyst for the OER. The interfacial  $\text{Ru-O-Ce}$  bridge structure facilitates electron transfer between Ru and Ce, creating lattice stress that distorts the local structure of  $\text{RuO}_2$ . This  $\text{RuO}_2\text{-CeO}_2$  catalyst demonstrated exceptional stability with negligible decay after 1000 hours of the OER in 0.5 M  $\text{H}_2\text{SO}_4$ , and high activity with an overpotential of only 180 mV at  $10 \text{ mA cm}^{-2}$ . *In situ* attenuated total reflectance surface-enhanced infrared absorption spectroscopy (ATR-SEIRAS), *in situ* differential electrochemical mass spectrometry (DEMS), and density functional theory (DFT) calculations were employed to reveal that the interface and noninterface  $\text{RuO}_2$  sites enabled an oxide path mechanism (OPM) and an enhanced adsorbate evolution mechanism (AEM-plus), respectively, during the OER. This dual-pathway approach, guided by lattice matching, enhances electrocatalyst design for the OER in acidic media.

It is claimed that current reliance on iridium-based OER catalysts is unsustainable due to their high cost and limited availability. Recently, Wu *et al.*<sup>72</sup> fabricated a nickel-stabilized ruthenium dioxide ( $\text{Ni-RuO}_2$ ) catalyst as a promising alternative, demonstrating significant improvements in both



**Table 3** Overpotential ( $\eta$ ) to achieve  $-10 \text{ mA cm}^{-2}$  for the hydrogen evolution reaction (HER) of various electrocatalysts in 0.5 M  $\text{H}_2\text{SO}_4$  acid electrolyte ( $\eta$  to achieve greater than  $-10 \text{ mA cm}^{-2}$  are indicated with their corresponding current densities), where  $R_{\text{ct}}$  ( $R_{\text{ct}}$ : charge transfer resistance) is acquired from electrochemical impedance spectroscopy (EIS)

| HER electrocatalysts                                          | $\eta$ at $-10 \text{ mA cm}^{-2}$ (mV) | $R_{\text{ct}}$ | Stability                         | Ref. |
|---------------------------------------------------------------|-----------------------------------------|-----------------|-----------------------------------|------|
| $\text{Pd}_2\text{RuO}_x$                                     | -12                                     | NA              | 20 h ( $10 \text{ mA cm}^{-2}$ )  | 74   |
| $\text{RuP}_2\text{-C@RuP}_2\text{-C}$                        | -17                                     | 9.10 $\Omega$   | 20 h ( $10 \text{ mA cm}^{-2}$ )  | 77   |
| $\text{Ru/Co-N-C}$                                            | -17                                     | NA              | 20 h ( $100 \text{ mA cm}^{-2}$ ) | 78   |
| $\text{PtCo@PtSn}$                                            | -21                                     | NA              | 100 h ( $80 \text{ mA cm}^{-2}$ ) | 79   |
| $\text{Ru-30}$                                                | -24                                     | NA              | 10 h ( $10 \text{ mA cm}^{-2}$ )  | 76   |
| $\text{IrIn}_2/\text{C}$                                      | -25                                     | NA              | NA                                | 80   |
| $\text{IrCo@CNT}$                                             | -26                                     | 3.78 $\Omega$   | 90 h ( $10 \text{ mA cm}^{-2}$ )  | 81   |
| $\text{Mo}_8\text{O}_{26}\text{-NbN}_x\text{O}_y/\text{NG}$   | -27                                     | 3.79 $\Omega$   | 100 000 s                         | 73   |
| $\text{Fe}_{2-x}\text{Ru}_x\text{Nb}_2\text{O}_6$             | -30                                     | 2.72 $\Omega$   | 80 h ( $10 \text{ mA cm}^{-2}$ )  | 26   |
| $\text{Co@CNTs Ru}$                                           | -32                                     | 51.46 $\Omega$  | NA                                | 82   |
| $\text{Ni}_2\text{P-Ni}_{12}\text{P}_5\text{@Ni}_3\text{S}_2$ | -46                                     | NA              | 24 h ( $10 \text{ mA cm}^{-2}$ )  | 83   |
| $\text{Ru-g-CN}$                                              | -54.5                                   | NA              | 120 h ( $10 \text{ mA cm}^{-2}$ ) | 84   |
| $\text{Ru@MoO(S)}_3$                                          | -63                                     | NA              | 24 h                              | 69   |
| $\text{Mn}_2\text{P-MnP/PNC}$                                 | -63                                     | NA              | 96 h                              | 85   |
| $\text{NCAG/Ru-3}$                                            | -65                                     | NA              | 30 h ( $10 \text{ mA cm}^{-2}$ )  | 86   |
| $\text{La-RuO}_2$                                             | -71                                     | NA              | 50 h ( $50 \text{ mA cm}^{-2}$ )  | 70   |
| $\text{Pb-CoSe}_2\text{-DETA}$                                | -74                                     | 4.8 $\Omega$    | 20 h ( $10 \text{ mA cm}^{-2}$ )  | 87   |
| $\text{Ni}_2\text{P/Ni}_5\text{P}_4$                          | -78                                     | NA              | 100 h ( $50 \text{ mA cm}^{-2}$ ) | 3    |
| $\text{N-IrP}_2$                                              | -80 mV at $-100 \text{ mA cm}^{-2}$     | NA              | 10 h ( $60 \text{ mA cm}^{-2}$ )  | 28   |
| $\text{WC-W}_2\text{C-HCDs}$                                  | -96                                     | 20.1 $\Omega$   | 150 h ( $10 \text{ mA cm}^{-2}$ ) | 88   |
| $\text{Co}_2\text{P/CoP@Co@NCNT}$                             | -136                                    | 10.3 $\Omega$   | 12 h ( $10 \text{ mA cm}^{-2}$ )  | 51   |
| $\text{Holey-VO}_2\text{ (R)}$                                | -184                                    | NA              | NA                                | 89   |
| $\text{V-Ti}_4\text{N}_3\text{T}_x$                           | -330                                    | 70 $\Omega$     | NA                                | 90   |

HER: hydrogen evolution reaction;  $\eta$ : overpotential;  $R_{\text{ct}}$ : charge transfer resistance; NA: not applicable; DETA = diethylenetriamine; HCDs: hollow carbon dodecahedrons.

**Table 4** Overpotential ( $\eta$ ) to achieve  $10 \text{ mA cm}^{-2}$  for the oxygen evolution reaction (OER) of various electrocatalysts in 0.5 M  $\text{H}_2\text{SO}_4$  acid electrolyte, where  $R_{\text{ct}}$  ( $R_{\text{ct}}$ : charge transfer resistance) is acquired from electrochemical impedance spectroscopy (EIS)

| OER electrocatalysts                                                  | $\eta$ at $10 \text{ mA cm}^{-2}$ (mV) | $R_{\text{ct}}$ | Stability                         | Ref. |
|-----------------------------------------------------------------------|----------------------------------------|-----------------|-----------------------------------|------|
| Strained- $\text{Mn}_{0.98}\text{Ir}_{0.02}\text{O}_2$                | 105                                    | NA              | 48 h ( $10 \text{ mA cm}^{-2}$ )  | 1    |
| $\text{N-IrP}_2$                                                      | 160                                    | NA              | 50 h ( $100 \text{ mA cm}^{-2}$ ) | 28   |
| $\text{Ru-30}$                                                        | 181                                    | NA              | 10 h ( $10 \text{ mA cm}^{-2}$ )  | 76   |
| $\text{GB-Ta}_{0.1}\text{TM}_{0.1}\text{Ir}_{0.8}\text{O}_{2-\delta}$ | 198                                    | NA              | 500 h ( $10 \text{ mA cm}^{-2}$ ) | 91   |
| $\text{Fe}_{2-x}\text{Ru}_x\text{Nb}_2\text{O}_6$                     | 200                                    | 4 $\Omega$      | 24 h ( $200 \text{ mA cm}^{-2}$ ) | 26   |
| $\text{La-RuO}_2$                                                     | 208                                    | NA              | 50 h ( $10 \text{ mA cm}^{-2}$ )  | 70   |
| $\text{Ru@MoO(S)}_3$                                                  | 226                                    | NA              | 24 h                              | 69   |
| $\text{Ru/Co-N-C}$                                                    | 232                                    | NA              | 20 h ( $10 \text{ mA cm}^{-2}$ )  | 78   |
| $\text{IrCo@CNT}$                                                     | 241                                    | NA              | 90 h ( $10 \text{ mA cm}^{-2}$ )  | 81   |
| $\text{IrIn}_2/\text{C}$                                              | 254                                    | NA              | 55 h ( $10 \text{ mA cm}^{-2}$ )  | 80   |
| $\text{CeO}_2\text{-Co}_2\text{NiP}_{0.03}\text{O}_x$                 | 262                                    | 7.7 $\Omega$    | 30 h ( $5 \text{ mA cm}^{-2}$ )   | 75   |
| $\text{Holey-VO}_2\text{ (R)}$                                        | 282                                    | NA              | NA                                | 89   |

OER: oxygen evolution reaction;  $\eta$ : overpotential;  $R_{\text{ct}}$ : charge transfer resistance; NA: not applicable.

activity and durability for the acidic OER in PEM-WE applications. The Ni-RuO<sub>2</sub> catalyst showed exceptional stability, maintaining its performance for over 1000 hours at a water-splitting current of  $200 \text{ mA cm}^{-2}$ . In contrast, pristine RuO<sub>2</sub> degraded rapidly under similar conditions. The incorporation of nickel into the RuO<sub>2</sub> lattice significantly stabilized both surface Ru and subsurface oxygen, extending the catalyst's durability by more than an order of magnitude. The

improved stability was confirmed by density functional theory (DFT) studies and *operando* differential electrochemical mass spectroscopy (DEMS), which highlighted the critical role of Ni dopants in enhancing the lattice stability and supporting the adsorbate-evolving mechanism (AEM) on Ni-RuO<sub>2</sub>. It is revealed by DFT calculations that the Ni dopant prefers to replace the coordination bridge Ru site of the RuO<sub>2</sub>(110) surface. Moreover, the limiting potential for the OER following an AEM is 1.87 eV lower than that of the lattice oxygen mechanism (LOM).

Moreover, Yang *et al.*<sup>73</sup> observed that Mo<sub>8</sub>O<sub>26</sub>-NbN<sub>x</sub>O<sub>y</sub>/NG (G: graphene) afforded an  $\eta$  of  $-27 \text{ mV}$  at  $-10 \text{ mA cm}^{-2}$  for the HER in 0.5 M  $\text{H}_2\text{SO}_4$ , suggesting both high activity and low charge transfer resistance (3.79  $\Omega$ ). It is composed of heterogeneous Mo<sub>8</sub>O<sub>26</sub>-NbN<sub>x</sub>O<sub>y</sub>, which is supported on N-doped graphene. It is prepared by hydrothermal treatment followed by a nitridation process. In Mo<sub>8</sub>O<sub>26</sub>-NbN<sub>x</sub>O<sub>y</sub>/NG, the O-exposed Mo<sub>8</sub>O<sub>26</sub> clusters are covalently bound to NbN<sub>x</sub>O<sub>y</sub> nanodomains, leading to a unique interface configuration with an optimal electronic structure, and the totally exposed multiple active sites provide superior HER activity in acidic medium, which is almost the same as that of a commercial Pt/C catalyst at  $-75 \text{ mA cm}^{-2}$ . DFT calculations reveal that facilitated initial H<sub>2</sub>O adsorption and near-optimal hydrogen adsorption energy are achieved due to the Mo<sub>8</sub>O<sub>26</sub>-NbN<sub>x</sub>O<sub>y</sub> interface with a modified electronic structure. The terminal O atoms of the Mo<sub>8</sub>O<sub>26</sub> clusters interact with the Nb atoms to enhance the initial H<sub>2</sub>O adsorption, and then decrease the H<sub>2</sub>O dissociation energy, facilitating the entire HER kinetics.



In addition, Do *et al.*<sup>74</sup> observed that Pd<sub>2</sub>RuO<sub>x</sub> afforded an  $\eta$  of  $-12$  mV at  $-10$  mA cm<sup>-2</sup> for the HER in 0.5 M H<sub>2</sub>SO<sub>4</sub>, showing satisfactory activity and durability (negligible decay after 10 000 cycles of CV at  $-140$  mA cm<sup>-2</sup>), as well as a low charge transfer resistance. The Pd<sub>2</sub>RuO<sub>x</sub> material is comprised of atomically thin Pd-PdO nanodomains that are encapsulated within amorphous Ru metallene oxide (RuO<sub>2</sub>). This multi-component heterostructure, along with the atomically thin nanosheet morphology and coexistence of defective crystalline and amorphous phases, collectively regulate the electronic structure, provide an optimal adsorption energy for the intermediates at the active sites, expose/provide abundant active sites, and significantly enhance the HER performance.

Additionally, Liu *et al.*<sup>75</sup> discovered that the precious-metal-free CeO<sub>2</sub>-Co<sub>2</sub>NiP<sub>0.03</sub>O<sub>x</sub> electrocatalyst delivered an exceptional OER performance, achieving an overpotential of 262 mV at 10 mA cm<sup>-2</sup> in 0.5 M H<sub>2</sub>SO<sub>4</sub>, along with notable stability and a low charge transfer resistance of 7.7  $\Omega$ . This innovation, leveraging a CeO<sub>2</sub> nanorod core with a Co-Ni-P oxide shell, benefits from the addition of trace P atoms and heterojunction interfaces that effectively adjust Co valence states to +2.51. This modification enhances the transformation of surface Co<sup>2+</sup>-O sites into more reactive Co<sup>3+</sup>-O sites, improving water molecule deprotonation and OER kinetics. The enhanced durability of CeO<sub>2</sub>-Co<sub>2</sub>NiP<sub>0.03</sub>O<sub>x</sub> is further reinforced by an increased vacancy formation energy at the Co site, reducing its dissolution in acidic environments.

Similarly, the study by Yan *et al.*<sup>76</sup> highlights Ru-30's outstanding OER and HER performance in 0.5 M H<sub>2</sub>SO<sub>4</sub>, achieving overpotentials of 181 mV and  $-24$  mV at 10 mA cm<sup>-2</sup>, respectively. This performance is coupled with substantial stability over 10 h and sustained durability, demonstrating minimal degradation after extensive cyclic voltammetry (500 cycles for the OER and 5000 cycles for the HER). For overall water splitting, Ru-30 achieves a potential of 1.468 V at 10 mA cm<sup>-2</sup>, maintaining its effectiveness over 16 h. Ru-30, a partially oxidized 3D porous ruthenium aerogel, boasts a distinct Ru/RuO<sub>2</sub> interface and a unique porous architecture. This unique combination facilitates rapid mass and electron transfer channels, offering an abundance of electrochemically active sites. Furthermore, it modifies the electronic structure and optimizes the adsorption energy with water splitting intermediates. Consequently, Ru-30 significantly enhances the performance for the hydrogen evolution reaction (HER) and the oxygen evolution reaction (OER) in acidic environments.

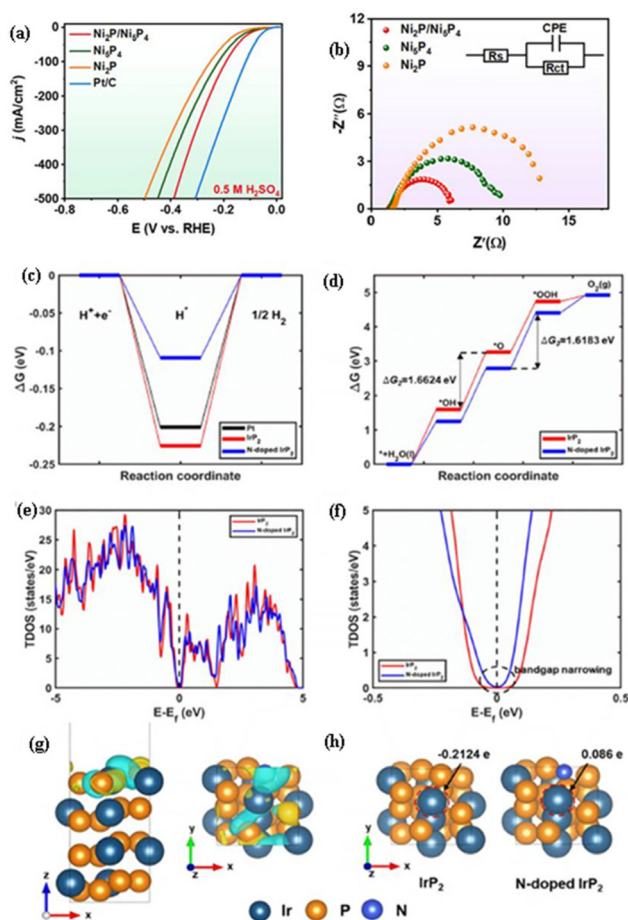
Najafi *et al.*<sup>89</sup> reported that holey-VO<sub>2</sub> (R) achieved an overpotential of 282 mV at 10 mA cm<sup>-2</sup> for the OER in 0.5 M H<sub>2</sub>SO<sub>4</sub>, highlighting its exceptional activity, along with displaying reasonable stability over 150 h. For the HER, this material reaches an overpotential of  $-184$  mV at  $-10$  mA cm<sup>-2</sup>, underscoring its high activity and maintaining robust stability for 120 h. Moreover, it demonstrates a high mass activity of 972.3 A g<sup>-1</sup> at  $-0.300$  V *versus* RHE and features a low charge transfer resistance, showcasing its efficiency at electrochemical water splitting. As-prepared VO<sub>2</sub> (R) is composed of rutile VO<sub>2</sub>

nanosheets, while it is a 2D metallic non-layered material and stable at room temperature. It is prepared under a reducing Ar-H<sub>2</sub> atmosphere through topochemically transforming liquid-phase exfoliated VSe<sub>2</sub>, whereas the 2D holey-VO<sub>2</sub> (R) nanosheets with a porous morphology were obtained through pretreating the VSe<sub>2</sub> nanosheets using O<sub>2</sub> plasma. Metallic holey-VO<sub>2</sub> (R) with a high specific surface area and defective sites exhibits accelerated electron/mass transfer, enhanced conductivity, and results in good performance for the HER and OER. Moreover, Hao *et al.*<sup>91</sup> observed that torsion-strained GB-Ta<sub>0.1</sub>Tm<sub>0.1</sub>Ir<sub>0.8</sub>O<sub>2- $\delta$</sub>  (GB: abundant grain boundaries) afforded an  $\eta$  of 198 mV at 10 mA cm<sup>-2</sup> for the OER in 0.5 M H<sub>2</sub>SO<sub>4</sub>, showing a laudable activity, robust stability, and low charge transfer resistance. X-ray absorption spectroscopy, microstructural analyses, and DFT calculations reveal that abundant trigeminal grain boundaries are formed due to the doping induced ligand effect and torsion-strained Ir-O bonds, which could modulate the adsorption energy of the intermediates thereby improving the activity for the OER. A proton exchange membrane electrolyser delivers a remarkably high current density of 1.5 A cm<sup>-2</sup> for 500 h at a low cell voltage of 1.766 V using GB-Ta<sub>0.1</sub>Tm<sub>0.1</sub>Ir<sub>0.8</sub>O<sub>2- $\delta$</sub>  as a catalyst for the OER at a low mass loading of 0.2 mg cm<sup>-2</sup> with an estimated cost of US\$1 per kg of hydrogen, which is much lower than the target of US\$2 per kg per H<sub>2</sub> (fixed by the US Department of Energy).

### 3.2. Phosphide/sulfide/selenide/nitride based nanoelectrocatalysts

The nanoscale interface-engineered transition metal phosphide heterostructure can enhance the conductivity, modify the electronic structure, and expose/afford abundant active sites for the HER. Lyu *et al.*<sup>3</sup> observed that the Ni<sub>2</sub>P/Ni<sub>5</sub>P<sub>4</sub> heterostructure nanosheets exhibited a higher activity ( $\eta$  of  $-78$  mV at  $-10$  mA cm<sup>-2</sup>; Fig. 6(a)) and lower charge transfer resistance (Fig. 6(b)) than those of Ni<sub>2</sub>P and Ni<sub>5</sub>P<sub>4</sub> for the HER in 0.5 M H<sub>2</sub>SO<sub>4</sub>. It was obtained by a solvothermal method and phosphorization treatment. The interface-engineered Ni<sub>2</sub>P/Ni<sub>5</sub>P<sub>4</sub> heterostructure with a porous nanosheet morphology can enhance the conductivity, and expose/afford abundant active sites for the HER. DOS simulations and UPS characterization reveal that a built-in electric field is observed at the heterostructure interface region, where electrons have been transferred to Ni<sub>5</sub>P<sub>4</sub> from Ni<sub>2</sub>P. Theoretical calculations reveal that the lowest water dissociation energy barrier and optimal H\* adsorption free energy are observed for the active sites at the interface domain of Ni<sub>2</sub>P/Ni<sub>5</sub>P<sub>4</sub> when compared to that of single-phase Ni<sub>5</sub>P<sub>4</sub> and Ni<sub>2</sub>P, which could optimize the electronic structure and redistribute the electrons, leading to an enhancement of the activity of Ni<sub>2</sub>P/Ni<sub>5</sub>P<sub>4</sub> for the HER.

The introduction of N dopants into IrP<sub>2</sub> effectively modifies the charge distribution on the IrP<sub>2</sub> surface and adjusts the band (d and p) centers of P and Ir. This modulation narrows the bandgap, enhances conductivity, and modifies the electronic structure. Additionally, it provides an optimal adsorp-



**Fig. 6** (a) LSV curve and (b) Nyquist plot for the HER of  $\text{Ni}_2\text{P}/\text{Ni}_5\text{P}_4$  in 0.5 M  $\text{H}_2\text{SO}_4$  in comparison with those of  $\text{Ni}_2\text{P}$  and  $\text{Ni}_5\text{P}_4$  (reproduced with permission from ref. 3. Copyright 2023, Elsevier B.V.). The Gibbs free energy difference ( $\Delta G$ ) for (c) the HER process on N-doped  $\text{IrP}_2$  in comparison with those of Pt and  $\text{IrP}_2$ , and (d) the OER process on N-doped  $\text{IrP}_2$  in comparison with that of  $\text{IrP}_2$ . (e) The TDOS of N-doped  $\text{IrP}_2$  in comparison with that of  $\text{IrP}_2$ ; (f) enlarged view of the TDOS of N-doped  $\text{IrP}_2$  in comparison with that of  $\text{IrP}_2$  near the Fermi level, where the bandgap of N-doped  $\text{IrP}_2$  is narrowed when compared to that of  $\text{IrP}_2$ , while the bandgap had been calculated from TDOS having a tolerance of 0.1 states per eV. (g) The difference in the charge densities between  $\text{IrP}_2$  and the N-doped  $\text{IrP}_2$  structure, where cyan and yellow colors indicate electron density depletion and enrichment with an iso-value of  $0.01 \text{ e Bohr}^{-3}$ . (h) Bader charge analysis of  $\text{IrP}_2$  and N-doped  $\text{IrP}_2$ , where the charge value of the Ir atom increases from  $-0.2124 \text{ e}$  in  $\text{IrP}_2$  to  $0.086 \text{ e}$  in N-doped  $\text{IrP}_2$  as indicated by the red dashed circle (reproduced with permission from ref. 28. Copyright 2022, Elsevier B.V.).

tion energy for intermediates, ultimately leading to improved performance for the HER and OER. Roy *et al.*<sup>28</sup> observed that N- $\text{IrP}_2$  exhibited higher activity and lower charge transfer resistance than those of  $\text{IrP}_2$  for the HER and OER in 0.5 M  $\text{H}_2\text{SO}_4$ , while it exhibited a low cell voltage of 1.41 V at  $10 \text{ mA cm}^{-2}$  (Table 5) for overall water electrolysis, suggesting its markedly higher activity in an acidic environment. As shown in Fig. 6(c), the Gibbs free energy differences ( $\Delta G$ ) for the HER processes of all the catalysts were almost zero, while N-doped  $\text{IrP}_2$  exhibited a much lower  $\Delta G$  (much closer to zero) than those of Pt

**Table 5** Potential to achieve  $10 \text{ mA cm}^{-2}$  for overall water splitting of various kinds of bifunctional electrocatalysts in 0.5 M  $\text{H}_2\text{SO}_4$  acid electrolyte

| Bifunctional electrocatalysts for HER and OER     | Overall water splitting (V at $10 \text{ mA cm}^{-2}$ ) | Stability                         | Ref. |
|---------------------------------------------------|---------------------------------------------------------|-----------------------------------|------|
| N- $\text{IrP}_2$                                 | 1.41                                                    | 50 h ( $100 \text{ mA cm}^{-2}$ ) | 28   |
| Ru-30                                             | 1.468                                                   | 10 h ( $10 \text{ mA cm}^{-2}$ )  | 76   |
| Ru/Co-N-C                                         | 1.49                                                    | 20 h ( $10 \text{ mA cm}^{-2}$ )  | 78   |
| $\text{IrCo@CNT}$                                 | 1.5                                                     | 90 h ( $10 \text{ mA cm}^{-2}$ )  | 81   |
| $\text{IrIn}_2/\text{C}$                          | 1.51                                                    | 55 h ( $10 \text{ mA cm}^{-2}$ )  | 80   |
| $\text{Ru@MoO(S)}_3$                              | 1.522                                                   | 24 h                              | 69   |
| La- $\text{RuO}_2$                                | 1.53                                                    | 50 h ( $10 \text{ mA cm}^{-2}$ )  | 70   |
| $\text{Fe}_{2-x}\text{Ru}_x\text{Nb}_2\text{O}_6$ | 1.55                                                    | 24 h ( $200 \text{ mA cm}^{-2}$ ) | 26   |

and  $\text{IrP}_2$ , suggesting high activity for N-doped  $\text{IrP}_2$ , which was corroborated by the experimental results. Fig. 6(d) depicts the  $\Delta G$  values of N-doped  $\text{IrP}_2$  in comparison with  $\text{IrP}_2$ . The OER process is composed of four elementary reaction steps, where the rate determining step for the overall OER process is determined by the largest step value of  $\Delta G$  among the four steps. In  $\text{IrP}_2$ , the H dissociation step ( $\text{OH} \rightarrow \text{O}$ ), with a  $\Delta G$  value of 1.6624 eV, serves as the rate-determining step, while in N-doped  $\text{IrP}_2$ , it shifts to OH adsorption by the O atom ( $\text{O} \rightarrow \text{OOH}$ ) with a slightly lower  $\Delta G$  of 1.6183 eV. Thus, N doping induces a shift towards a different rate-limiting process with a reduced  $\Delta G$  value, which aligns well with the experimental results. To investigate the bandgap, the total density of states (TDOS) of N-doped  $\text{IrP}_2$  in comparison with  $\text{IrP}_2$  is shown in Fig. 6(e), with an enlarged view shown in Fig. 6(f). The N-doped  $\text{IrP}_2$  catalyst exhibits a narrower bandgap (0.1381 eV) when compared to that of  $\text{IrP}_2$  (0.2194 eV), suggesting enhanced electron transfer, leading to an enhancement in the activity of N-doped  $\text{IrP}_2$  for the HER and OER. Fig. 6(g) depicts the difference in charge densities between  $\text{IrP}_2$  and the N-doped  $\text{IrP}_2$  structure. DFT calculations revealed that the N dopant in  $\text{IrP}_2$  could modify the charge distribution of the  $\text{IrP}_2$  surface and the band (d and p) centers of P and Ir, leading to high activity for water electrolysis. The Bader charge analysis presented in Fig. 6(h) reveals a significant increase in the positive charge of Ir atoms on the surface of N-doped  $\text{IrP}_2$ , from  $-0.2124 \text{ e}$  to  $0.086 \text{ e}$ . This positive charge augmentation endows the Ir atoms with enhanced activity for the chemisorption of OER intermediates. Consequently, N doping of  $\text{IrP}_2$  leads to a reduction in  $\Delta G$  and subsequent enhancement of the OER activity.

Besides, the work of Wu *et al.*<sup>85</sup> proved that  $\text{Mn}_2\text{P-MnP/PNC}$  exhibited a higher activity than those of  $\text{MnP/PNC}$  and  $\text{Mn}_2\text{P/PNC}$  for the HER in 0.5 M  $\text{H}_2\text{SO}_4$  and 1 M KOH, while it exhibited a lower charge transfer resistance than those of  $\text{MnP/PNC}$  and  $\text{Mn}_2\text{P/PNC}$  for the HER in 1 M KOH. Moreover, Luo *et al.*<sup>77</sup> observed that  $\text{RuP}_2\text{-C@RuP}_2\text{-C}$  exhibited a higher activity than that of  $\text{RuP}_2$  for the HER in 0.5 M  $\text{H}_2\text{SO}_4$ , while it exhibited a low charge transfer resistance for the HER in 0.5 M  $\text{H}_2\text{SO}_4$ . In contrast, Lu *et al.*<sup>51</sup> observed that  $\text{Co}_2\text{P/CoP@Co@NCNT}$  exhibited a higher activity than those of

Co<sub>2</sub>P@Co@NCNT and CoP@Co@NCNT for the HER in 0.5 M H<sub>2</sub>SO<sub>4</sub> and 1 M KOH, whereas Co<sub>2</sub>P@Co@NCNT ( $R_{ct}$ : 8.9  $\Omega$ ) exhibited a lower charge transfer resistance than those of Co<sub>2</sub>P/CoP@Co@NCNT ( $R_{ct}$ : 10.3  $\Omega$ ) and CoP@Co@NCNT ( $R_{ct}$ : 12.4  $\Omega$ ) for the HER in 1 M KOH. The decrease in the relative metallic cobalt content and increase of the P content in Co<sub>2</sub>P/CoP@Co@NCNT can be attributed to the decrease in the electrical conductivity of Co<sub>2</sub>P/CoP@Co@NCNT compared to that of Co<sub>2</sub>P@Co@NCNT.

Moreover, Yang *et al.*<sup>83</sup> found that Ni<sub>2</sub>P–Ni<sub>12</sub>P<sub>5</sub>@Ni<sub>3</sub>S<sub>2</sub> demonstrated a remarkable HER activity with an overpotential of –46 mV at –10 mA cm<sup>–2</sup> in 0.5 M H<sub>2</sub>SO<sub>4</sub>, alongside notable stability over 24 h and minimal charge transfer resistance. This heterostructure, featuring Ni<sub>2</sub>P–Ni<sub>12</sub>P<sub>5</sub> nanorod arrays on Ni<sub>3</sub>S<sub>2</sub> film, is crafted through simultaneous corrosion, sulfidation, and phosphidation. *In/ex situ* analyses and DFT studies reveal strong interfacial coupling between Ni<sub>12</sub>P<sub>5</sub> and Ni<sub>2</sub>P, with the Ni<sub>2</sub>P–Ni<sub>12</sub>P<sub>5</sub> nanorods' stability being significantly bolstered by the Ni<sub>3</sub>S<sub>2</sub> film, thereby optimizing water adsorption/dissociation energies and hydrogen adsorption, enhancing electron/proton transfer, and ultimately, boosting the HER efficiency.

Similarly, Wu *et al.*<sup>87</sup> highlighted Pb–CoSe<sub>2</sub>–DETA's exceptional HER performance, showing an overpotential of –74 mV at –10 mA cm<sup>–2</sup> in 0.5 M H<sub>2</sub>SO<sub>4</sub>, excellent stability for 20 h, and a low charge transfer resistance of 4.8  $\Omega$ . This material, a Pb single-atom-doped CoSe<sub>2</sub>–DETA nanobelt, is closest to the peak of the HER activity in the “volcano plot,” indicating an optimal Co–N to Co–Se bond ratio. The unique Co atomic configuration, induced by Pb doping, optimizes the adsorption energy with intermediates, significantly enhancing the HER performance. Yu *et al.*<sup>84</sup> reported that Ru–g-CN exhibited remarkable HER activity in 0.5 M H<sub>2</sub>SO<sub>4</sub>, achieving an overpotential of –54.5 mV at –10 mA cm<sup>–2</sup>. It demonstrates a high TOF (12.9 s<sup>–1</sup> at 100 mV), mass activity (24.55 A mg<sup>–1</sup> at  $\eta$  of 100 mV), apparent current density, and stability (lasting 120 h). Furthermore, it displays a low charge transfer resistance and a high photocatalytic H<sub>2</sub> evolution rate of 489.7 mmol H<sub>2</sub> g<sup>–1</sup> h<sup>–1</sup>. Ru–g-CN is composed of Ru single atoms anchored on mesoporous graphitic carbon nitride. Theoretical calculations reveal that anchoring Ru single atoms on g-CN can tune the electronic structure, decrease the bandgap, facilitate electrical conductivity, enhance electron transfer, reduce the Gibbs free energy, and thereby enhance the HER performance.

Djire *et al.*<sup>90</sup> noted V–Ti<sub>4</sub>N<sub>3</sub>T<sub>x</sub>'s superior HER activity in 0.5 M H<sub>2</sub>SO<sub>4</sub>, with an overpotential of –330 mV at –10 mA cm<sup>–2</sup> and a low charge transfer resistance of 70  $\Omega$ . V–Ti<sub>4</sub>N<sub>3</sub>T<sub>x</sub> is prepared from pristine exfoliated Ti<sub>4</sub>N<sub>3</sub>T<sub>x</sub> MXene with V salt through a solution-based method, while the mixed transition metal nitride MXene with rich metal loading optimizes the electrochemistry, increases the conductivity, and enhances the HER performance. The electrochemical activity of individual MXene flakes was investigated by scanning electrochemical microscopy (SECM), which revealed that the basal plane of MXene exhibited much higher HER activity than the flake

edge. This may be ascribed to the high exposure of metallic sites for proton adsorption.

### 3.3. Carbon/alloy based electrocatalysts

The engineering of the electronic structure of single-atom Ru sites through Co–N<sub>4</sub> sites can provide an optimal adsorption energy with intermediates, enhance the conductivity, and promote the performance for the HER and OER. Rong *et al.*<sup>78</sup> reported that Ru/Co–N–C exhibited higher activity, reduced charge transfer resistance, and improved stability compared to that of Ru–N–C. DFT calculations reveal that the introduction of Co–N<sub>4</sub> sites into Ru/Co–N–C fine-tunes the electronic structure of Ru by enhancing its electron density and strengthening the Ru–O covalency. This modulation results in an optimal adsorption energy for HER and OER intermediates on Ru sites, thereby boosting the performance for both reactions. Additionally, the incorporation of Co–N<sub>4</sub> sites prompts electron redistribution around Ru–N<sub>4</sub>, thereby enhancing the corrosion resistance of Ru/Co–N–C during acidic electrolysis.

The fabrication of IrIn<sub>2</sub> intermetallics having a low Ir content with an unconventional face-centered orthorhombic phase could modify the electronic structure, enhance the conductivity, and provide optimal adsorption of intermediates, which could enhance the performance for the HER and OER. He *et al.*<sup>80</sup> observed that IrIn<sub>2</sub>/C exhibited higher activity and lower charge transfer resistance than those of Ir/C for the HER and OER in 0.5 M H<sub>2</sub>SO<sub>4</sub>, and it also showed high activity and stability for overall water electrolysis in 0.5 M H<sub>2</sub>SO<sub>4</sub>. IrIn<sub>2</sub> intermetallics having a low Ir content of 33% with an unconventional face-centered orthorhombic phase were obtained by a high-temperature annealing reduction method. Theoretical calculations reveal that an upward shift in the d-band center of IrIn<sub>2</sub> is observed when compared to that of metallic Ir, leading to an enhancement in the adsorption of intermediates. Moreover, the ordered Ir–In bonds in IrIn<sub>2</sub> intermetallic compounds could enhance charge transfer and enable resistance to dissolution and atom migration during electrolysis.

Moreover, Chen *et al.*<sup>82</sup> observed that Co@CNTs/Ru afforded an  $\eta$  of –32 mV at –10 mA cm<sup>–2</sup> for the HER in 0.5 M H<sub>2</sub>SO<sub>4</sub>, demonstrating excellent activity and stability as well as a low charge transfer resistance (51.46  $\Omega$ ). Co@CNTs/Ru is composed of Co nanoparticles, which are *in situ* confined inside bamboo-like carbon nanotubes, where an ultralow Ru loading (~2.6  $\mu$ g per electrode area ~ cm<sup>–2</sup>) is homogeneously deposited on their exterior walls. DFT calculations and atomic-scale structural analysis reveal that outer loaded Ru and inner confined Co can induce charge redistribution and synergistic electron coupling, which can provide an optimal adsorption energy of H intermediates ( $\Delta G_{H^+}$ ) and enhance electron/mass transfer, finally improving the HER performance.

In addition, Wang *et al.*<sup>81</sup> observed that IrCo@CNT afforded an  $\eta$  of 241 mV at 10 mA cm<sup>–2</sup> for OER in 0.5 M H<sub>2</sub>SO<sub>4</sub>, suggesting high activity and stability over 90 h. IrCo@CNT demonstrates an overpotential of –26 mV at

$-10 \text{ mA cm}^{-2}$  for the hydrogen evolution reaction (HER) in  $0.5 \text{ M H}_2\text{SO}_4$ , highlighting exceptional activity and maintaining high stability over 90 h with a low charge transfer resistance of  $3.78 \Omega$ . In the context of overall water splitting in the same acidic medium, IrCo@CNT achieves a potential of  $1.5 \text{ V}$  at  $10 \text{ mA cm}^{-2}$ , utilizing an ultra-low Ir loading of  $0.027 \text{ mg cm}^{-2}$ , which underlines its remarkable efficiency. Additionally, its performance remains consistent without any decay throughout the 90-hour test period, confirming its durable stability. It was obtained on carbon cloth through a MOF-templated and dicyandiamide-assisted pyrolysis strategy. IrCo@CNT is composed of IrCo nanoparticles, which are encapsulated by high-density carbon nanotubes (CNTs). The superior performance of IrCo@CNT can be attributed to the synergistic interaction between Ir and Co, coupled with the protective role played by CNTs. Notably, the IrCo nanoparticles within IrCo@CNT remain intact after the HER, whereas they transform into Co-doped  $\text{IrO}_2$  nanoparticles following the OER. Theoretical calculations reveal that the energy barriers for the HER and OER can be lowered due to the optimal electronic structure obtained through Co doping of Ir and  $\text{IrO}_2$ , respectively.

Sun *et al.*<sup>88</sup> reported that WC- $\text{W}_2\text{C}$ -HCDs (HCDs: hollow carbon dodecahedrons) exhibited a remarkable HER activity in  $0.5 \text{ M H}_2\text{SO}_4$ , achieving an overpotential of  $-96 \text{ mV}$  at  $-10 \text{ mA cm}^{-2}$ . Additionally, they demonstrate high stability over 150 h and a low charge transfer resistance of  $20.1 \Omega$ . WC- $\text{W}_2\text{C}$ -HCDs are composed of dual-phase WC- $\text{W}_2\text{C}$  nanocrystals, which are encapsulated in hollow carbon dodecahedrons. The heterointerfaces between WC and  $\text{W}_2\text{C}$  facilitate strong electronic coupling, optimizing the adsorption energy for reaction intermediates. Concurrently, the hollow carbon dodecahedrons enhance the catalytic activity and stability for the HER by exposing abundant active sites and enhancing the electrical conductivity.

He *et al.*<sup>86</sup> discovered that NCAG/Ru-3 delivered an overpotential of  $-65 \text{ mV}$  at  $-10 \text{ mA cm}^{-2}$  for the hydrogen evolution reaction (HER) in  $0.5 \text{ M H}_2\text{SO}_4$ , highlighting its outstanding activity and stability, alongside a low charge transfer resistance. NCAG/Ru-3, a composite of carbon aerogels doped with atomically isolated Ru nanoclusters and  $\text{RuN}_x$  moieties, was synthesized through the pyrolysis of biomass hydrogels. DFT calculations suggest that the high activity originates from the  $\text{RuN}_x$  sites located at zigzag edges and nanowrinkles. Similarly, Chen *et al.*<sup>79</sup> reported that PtCo@PtSn achieved an overpotential of  $-21 \text{ mV}$  at  $-10 \text{ mA cm}^{-2}$  for the HER in  $0.5 \text{ M H}_2\text{SO}_4$ , demonstrating exceptional activity and remarkable durability, with negligible decay observed after 50 000 cycles of cyclic voltammetry (CV), and maintaining a low charge transfer resistance. The superior performance of PtCo@PtSn is credited to the presence of a Pt-based bimetallic alloy and heterojunction, with DFT calculations indicating that interface engineering optimizes the surface electronic structure of Pt. This optimization enhances the water dissociation capability and reduces the bond strength of Pt-H, achieving an optimal  $\text{H}^*$  Gibbs free energy ( $\Delta G_{\text{H}^*}$ ), thereby boosting HER efficiency.

## 4. Electrolysis in seawater

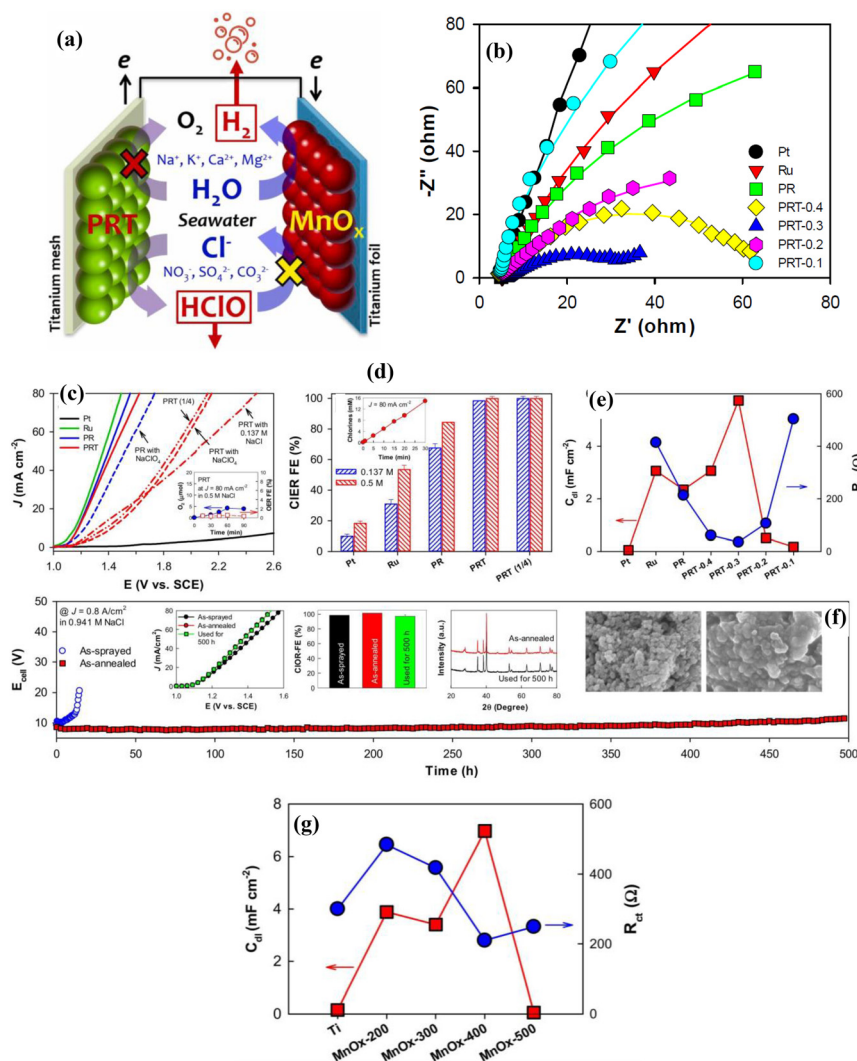
### 4.1. Oxide/hydroxide/(oxy)hydroxide/layered double hydroxide based nano-electrocatalysts

Liang *et al.*<sup>25</sup> demonstrated highly stable seawater electrolysis lasting 100 h, employing a  $\text{MnO}_x$  cathode and a PtRuTiO<sub>x</sub> (PRT) anode in a single-compartment cell without the need for a membrane. As schematically depicted in Fig. 7(a), the use of a selective chlorine oxidation reaction (ClOR) anode and a HER-selective cathode effectively eliminates the chlorine reduction reaction (ClRR) and the oxygen evolution reaction (OER). Consequently, pure  $\text{H}_2$  gas was produced, and the chlorine generated remained in the aqueous phase. The PtRuTiO<sub>x</sub> catalyst exhibits superior performance in saline water, boasting a high faradaic efficiency (FE) of approximately 100% for the chlorine oxidation reaction (ClOR) (Fig. 7(d)), as well as reduced charge transfer resistance (Fig. 7(b and e)). Its high activity (Fig. 7(c)) compared to other catalysts indicates a selectivity towards the ClOR over the oxygen evolution reaction (OER). The as-annealed PRT catalyst exhibits high stability for the ClOR for 500 h at a higher current density of  $800 \text{ mA cm}^{-2}$  (Fig. 7(f)). The  $\text{MnO}_x$  catalyst demonstrates a remarkable performance for the hydrogen evolution reaction (HER). In saline water with chlorine species ( $\text{HClO}/\text{ClO}^-$ ), it exhibits a high faradaic efficiency (FE) of approximately 100% (Fig. 7(g)) and reduced charge transfer resistance (Fig. 7(g)), outperforming other catalysts. Its high activity further underscores its selectivity towards the HER over the chlorine reduction reaction (ClRR), making it an ideal candidate for efficient seawater electrolysis.

The design of an electrocatalyst with a dual interface can reinforce the built-in electric field, enhance the conductivity, and provide an optimal adsorption energy for intermediates, which could enhance chlorine-free seawater oxidation. Zhang *et al.*<sup>92</sup> observed that the  $(\text{Ni}, \text{Fe})\text{OOH}@ \text{Ni}_x\text{P} \parallel \text{Ni}_x\text{P}$  electrolyzer exhibited higher activity ( $1.71 \text{ V}$  at  $100 \text{ mA cm}^{-2}$  and  $1.81 \text{ V}$  at  $500 \text{ mA cm}^{-2}$ ) and stability, and a lower charge transfer resistance than those of  $(\text{Ni}, \text{Fe})\text{OOH} \parallel \text{Ni}_x\text{P}$  and  $\text{Ni}_x\text{P} \parallel \text{Ni}_x\text{P}$  for overall seawater splitting in alkaline natural seawater. Moreover, almost no hypochlorite was detected for a long time in the seawater electrolyte at a cell voltage of  $\geq 1.72 \text{ V}$ , where this operating voltage was larger than that of the theoretical ClOR potential.  $(\text{Ni}, \text{Fe})\text{OOH}@ \text{Ni}_x\text{P}$  is comprised of a dual-interface between  $(\text{Ni}, \text{Fe})\text{OOH}$  and  $\text{Ni}_{12}\text{P}_5/\text{Ni}_2\text{P}$ . This interface enhances the intrinsic driving force, facilitating interfacial electron transport and generating a robust built-in electric field with asymmetric charge distribution at the  $(\text{Ni}, \text{Fe})\text{OOH}/\text{Ni}_{12}\text{P}_5$  boundary. This configuration boosts the kinetics for the oxygen evolution reaction (OER) while mitigating the adsorption capacity of chloride ions ( $\text{Cl}^-$ ), thus optimizing the catalyst's performance in electrolysis.

Moreover, Ren *et al.*<sup>93</sup> reported that Cu- $\text{MoO}_2$  achieved an overpotential of  $-330.2 \text{ mV}$  at  $-10 \text{ mA cm}^{-2}$  for the hydrogen evolution reaction (HER) in artificial seawater (neutral pH), indicating exceptional activity. Furthermore, it demonstrates significant durability, with minimal decay observed after 1000 cycles of cyclic voltammetry (CV), and maintains a low charge





**Fig. 7** (a) Schematic diagram illustrating seawater electrolysis using the MnO<sub>x</sub> cathode and PtRuTiO<sub>x</sub> (PRT) anode without using a membrane in a single compartment cell, where the ClOR selective anode and HER selective cathode could eliminate the ClRR, while pure H<sub>2</sub> gas could be formed, where the chlorine generated could be in the aqueous form. (b) Nyquist plot of Pt<sub>0.06</sub>Ru<sub>0.24</sub>Ti<sub>0.7</sub>O<sub>y</sub> in comparison with those of its counterparts for ClOR in simulated seawater electrolyte. (c) LSV curves for ClOR of PRT in comparison with those of its counterparts in 0.5 M NaCl (pH 6), 0.137 M NaCl and 0.5 M NaClO<sub>4</sub> (pH 6), where PRT (1/4) indicates a mass loading of 0.83 mg cm<sup>-2</sup>, while the inset depicts the ClOR and faradaic efficiency (FE) of PRT at 80 mA cm<sup>-2</sup> in 0.5 M NaCl. (d) FE of PRT in comparison with those of its counterparts for ClOR at 80 mA cm<sup>-2</sup>, where the inset depicts the production of free chlorine over time at 80 mA cm<sup>-2</sup>. (e) R<sub>ct</sub> and C<sub>dl</sub> values obtained from the Nyquist plot of Pt<sub>0.06</sub>Ru<sub>0.24</sub>Ti<sub>0.7</sub>O<sub>y</sub> in comparison with those of its counterparts in simulated seawater electrolyte. (f) Long-term stability test for ClOR of as-annealed PRT in comparison with that of as-sprayed PRT at 80 mA cm<sup>-2</sup> in 0.941 M NaCl, where the inset depicts the LSV curves and FE for ClOR of the as-sprayed and as-annealed materials and PRT used for 500 h at 80 mA cm<sup>-2</sup> in 0.5 M NaCl, while the inset further shows the XRD patterns of the as-annealed material and PRT used for 500 h, and the inset further depicts the SEM images of the as-annealed material (left) and PRT used for 500 h (right). (g) R<sub>ct</sub> and C<sub>dl</sub> values obtained from the Nyquist plot of MnO<sub>x</sub>-400 in comparison with those of its counterparts for the HER in simulated seawater electrolyte (reproduced with permission from ref. 25. Copyright 2022, Elsevier B.V.).

transfer resistance. It was *in situ* obtained from polyoxometalate-based metal-organic frameworks (POMOFs). Cu-MoO<sub>2</sub> is a hybrid material with a Schottky heterojunction, which could modulate the electronic structure, provide an optimal adsorption energy with intermediates, and enhance the performance for the HER.

In addition, Liu *et al.*<sup>94</sup> observed that Fe-Mo-Co-O afforded an  $\eta$  of 365 mV at 50 mA cm<sup>-2</sup> for the OER in 1 M KOH + seawater, suggesting its very high activity, while it exhibited low

charge transfer resistance. It affords an  $\eta$  of -250 mV at -50 mA cm<sup>-2</sup> for the HER in 1 M KOH + seawater, suggesting its very high activity. For overall water splitting in 1 M KOH + seawater, Fe-Mo-Co-O//Fe-Mo-Co-O affords a potential of 1.59 V at 10 mA cm<sup>-2</sup>, suggesting its very high activity, while it affords high stability for 180 h. Fe-Mo-Co-O is composed of Fe-doped Co<sub>2</sub>Mo<sub>3</sub>O<sub>8</sub>/MoO<sub>3</sub>/Co<sub>3</sub>O<sub>4</sub> nanosheets, where the enhanced performance is attributed to the combination of highly active Fe dopant and corrosion-resistant Mo-O species.

Jin *et al.*<sup>95</sup> highlighted that  $\text{CoCr}_{0.7}\text{Rh}_{1.3}\text{O}_4$  reached an overpotential of 273 mV at  $10 \text{ mA cm}^{-2}$  for the OER in 1 M KOH + 0.5 M NaCl, indicating outstanding activity and stability for 12 h, along with minimal charge transfer resistance. It is composed of nanofibers of single-phase trimetallic spinel  $\text{CoCr}_{0.7}\text{Rh}_{1.3}\text{O}_4$ . DFT calculations reveal that the synergistic effects of Co, Cr, and Rh can decrease the energy barriers for water dissociation, which could enhance the performance for the OER.

Wen *et al.*<sup>96</sup> revealed that  $\text{Fe}_1/\text{Mn-NiO}$  and  $\text{Pt}_1/\text{Mn-NiO}$  exhibited overpotentials of 290 mV for the OER and -129 mV for the HER, respectively, at  $100 \text{ mA cm}^{-2}$  in 0.5 M KOH + seawater, both maintaining stability for 30 h. The introduction of Mn into low polarized NiO nanosheets can alter the surface polarization, which can regulate the electronic metal-support interactions between the NiO support and anchored Pt or Fe single atoms, while the resulting change in coordination structure and charge redistribution in  $\text{Fe}_1/\text{Mn-NiO}$  and  $\text{Pt}_1/\text{Mn-NiO}$  can decrease the Gibbs free energy for the rate-determining steps in the OER and HER, respectively, which can enhance the performance.

In addition, Hemmati *et al.*<sup>97</sup> observed that NRAHM-NiO afforded an  $\eta$  of 70 mV at  $10 \text{ mA cm}^{-2}$  for the OER in 1 M KOH + 0.5 M NaCl, suggesting its very high activity, while it exhibited a low charge transfer resistance ( $R_{\text{ct}}$ :  $3.4 \Omega$ ;  $R_p$ :  $0.5611 \Omega$ ). NRAHM-NiO affords an  $\eta$  of -51 mV at  $-10 \text{ mA cm}^{-2}$  for the HER in 1 M KOH + 0.5 M NaCl, suggesting its very high activity, while it exhibits a low charge transfer resistance ( $R_{\text{ct}}$ :  $3.015 \Omega$ ;  $R_p$ :  $0.4327 \Omega$ ). For overall water splitting in 1 M KOH + 0.5 M NaCl, NRAHM-NiO//NRAHM-NiO affords a potential of 1.66 V at  $100 \text{ mA cm}^{-2}$ , suggesting its very high activity, while it affords high stability for 168 h; the integrated photolysis system powered by a single solar cell exhibits a high solar-to-hydrogen (STH) efficiency of 9.9%. It exhibits high selectivity (99% faradaic efficiency) for the OER in seawater without generating any chlorine species. NRAHM-NiO is composed of hydrophilic urchin-like nanorod array-based hierarchical NiO microspheres, which can facilitate gas evolution, enhance the charge transfer ability, afford high intrinsic electroactive sites, and improve the corrosion resistance, and thus could enhance the performance for the HER and OER.

Furthermore, Zhuo *et al.*<sup>98</sup> reported the synthesis of Ni-Fe hydroxides through a cost-effective, ultrafast method. The catalyst demonstrates low OER overpotentials of 240 mV and 254 mV at  $10 \text{ mA cm}^{-2}$  in 1 M KOH and alkaline seawater, respectively, while maintaining excellent electrochemical stability. *In situ* Raman spectroscopy and other physical characterization studies reveal that the incorporation of Fe and the transformation of  $\text{Ni(Fe)(OH)}_2$  into  $\text{Ni(Fe)OOH}$  are key to the enhanced OER performance. Additionally, the Ni-Fe hydroxide catalyst can be synthesized in just 1 min and scaled up easily. Even at a size of  $2000 \text{ cm}^2$ , the catalyst remains electrochemically uniform.

A study by Du *et al.*<sup>99</sup> showed that  $\text{Co,Cl-Cu(OH)}_2$  demonstrated an overpotential of 310 mV at  $100 \text{ mA cm}^{-2}$  for the OER and  $\text{Ru,Cl-Cu(OH)}_2$  achieved -37 mV at  $-10 \text{ mA cm}^{-2}$  for

the HER in 1 M KOH + seawater, highlighting exceptional activity for both reactions. Moreover, both catalysts show high stability and low charge transfer resistance, indicating their efficiency at electrolytic water splitting. For overall water splitting in 1 M KOH + seawater,  $\text{Co,Cl-Cu(OH)}_2^{\text{OER}}//\text{Ru,Cl-Cu(OH)}_2^{\text{HER}}$  affords a potential of 1.55 V at  $10 \text{ mA cm}^{-2}$ , suggesting its very high activity, while it affords high stability for 100 h. It was obtained by *in situ* corrosive engineering and cation/anion exchange at ambient temperature.  $\text{Co,Cl-Cu(OH)}_2$  is composed of Cu-based hollow nanotubes, where the rough surfaces are coupled with a chloridion-enriched environment. The enriched chloride surface can inhibit the attack of chlorine species, while the hollow nanotubes with rough surfaces can expose abundant active sites and facilitate gas evolution, which could enhance the performance for the OER.

Sun *et al.*<sup>100</sup> identified that  $\text{Pt}_2/\text{Ni(OH)}_2$ , derived from a MOF, catalyzed the HER with an overpotential of -19 mV at  $-10 \text{ mA cm}^{-2}$  in 1 M KOH + seawater, showcasing high activity, low charge transfer resistance, and a solar-to-hydrogen efficiency of 13.5% when powered by perovskite solar cells. A low cell voltage of 1.46 V at  $10 \text{ mA cm}^{-2}$  is achieved from the anion exchange membrane (AEM) seawater electrolyzer with a  $\text{Pt}_2/\text{Ni(OH)}_2$  cathode and Ni foam anode, while it exhibits low energy consumption to generate  $1 \text{ m}^3$  of  $\text{H}_2$  at  $200 \text{ mA cm}^{-2}$  (3.8 kW h).  $\text{Pt}_2/\text{Ni(OH)}_2$  is composed of  $\text{Ni(OH)}_2$  nanosheets, which are decorated with Pt nanoparticles. It is derived from a MOF. It exhibits SMSIs (strong metal-support interactions) of Pt-O-Ni bonds at the interface, a high superhydrophilic surface area, and the synergistic effect of Pt-O vacancy ( $\text{O}_v$ ) sites, which could enhance the performance for the HER. Furthermore, DFT calculations reveal that SMSIs at the interface and  $\text{O}_v$  can alter the electronic structure and localized charge density of  $\text{Ni(OH)}_2$ . The synergistic effect of Pt- $\text{O}_v$  sites can afford optimal adsorption and desorption of  $\text{H}^*$ , while adsorption of  $\text{H}_2\text{O}$  with high efficiency can be achieved by  $\text{O}_v$ , which can regulate the spillover of hydrogen to Pt from  $\text{Ni(OH)}_2$ , and this could enhance the performance for the HER.

Pan *et al.*<sup>101</sup> found that s-NiFeCrMo-OH achieved an overpotential of 242 mV for the OER at  $100 \text{ mA cm}^{-2}$  in 1 M KOH + 0.5 M NaCl, with exceptional stability and a low charge transfer resistance. A low cell voltage of 1.82 V at a huge current density of  $500 \text{ mA cm}^{-2}$  over 1000 h at ambient temperature was achieved in an overall alkaline seawater electrolyzer using s-NiFeCrMo-OH as the anode, while the floating solar seawater splitting device exhibited a high STH efficiency of  $20.61 \pm 0.77\%$ . s-NiFeCrMo-OH is a 2D quaternary metal hydroxide nanosheet, which is composed of Ni, Fe, Cr, and Mo elements; partial Mo element leaching morphologically transforms the catalyst by *in situ* electrochemical activation. The existence of a metal with high valence states and abundant O vacancies is attributed to the enhanced catalytic performance and corrosion resistance.

Moreover, Zhou *et al.*<sup>102</sup> observed that  $\text{Ru/NiFeOOH}$  afforded an  $\eta$  of 330 mV at a huge current density of  $500 \text{ mA cm}^{-2}$  for the OER in 1 M KOH + seawater, suggesting its very high activity, and it exhibited high stability for >95 h with a

low charge transfer resistance. Ru/NiFeOOH is prepared on NiFe foam through a one-step bimetallic substrate induction strategy to construct supported metal catalysts (SMCs) integrally as a binder-free electrode; this synthesis strategy (bimetallic substrate induction) can also be applied using seawater rather than pure water. The NiFe foam substrate is a source of Ni and Fe metals, acts as a reducing agent for the *in situ* growth of NiFeOOH nanosheets with Ru nanoparticles, and affords strong interactions between the NiFe substrate and SMCs. Experimental studies and DFT calculations reveal that *in situ* formed metal-support interactions can provide optimal electronic states and increase the chemical stability.

In addition, Zhang *et al.*<sup>109</sup> demonstrated that the Ni based anode suffered more corrosion by Br<sup>−</sup> in seawater than that of Cl<sup>−</sup> in seawater due to the faster corrosion kinetics in bromine. Simulated and experimental studies reveal that extensive etching of Br<sup>−</sup> leads to the formation of shallow-wide pits, while localized corrosion of Cl<sup>−</sup> leads to the generation of narrow-deep pits, which can be ascribed to the lower reaction energy of Br<sup>−</sup> and the fast diffusion kinetics of Cl<sup>−</sup> in the passivation layer. Furthermore, NiFe LDH exhibits rapid performance degradation due to the extensive spalling of the catalyst layer induced by Br<sup>−</sup>, suggesting the importance of designing anti-Br<sup>−</sup> corrosion anodes along with anti-Cl<sup>−</sup> corrosion anodes for seawater electrolysis.

Enkhtuvshin *et al.*<sup>106</sup> observed that c-NiFe/a-NiFe LDH (crystalline and amorphous layered double hydroxide) demonstrated an overpotential of 300 mV at 100 mA cm<sup>−2</sup> for the oxygen evolution reaction (OER) in 1 M KOH + 0.5 M NaCl, showcasing exceptional activity and maintaining its stability over 100 h. Similarly, for the hydrogen evolution reaction (HER) under the same conditions, it achieves an overpotential of −200 mV at −100 mA cm<sup>−2</sup>, again proving its high activity and enduring stability, coupled with a low charge transfer resistance of approximately 5 Ω. In applications involving overall water splitting in 1 M KOH + seawater, the c-NiFe/a-NiFe LDH

system requires about 1.69 V at 10 mA cm<sup>−2</sup> to operate, maintaining its performance and stability for 100 h. This LDH structure, comprising a Ni<sub>3</sub>Fe alloy phase within a 2D network topped by a ~2 nm ultrathin, low-crystalline NiFeOOH phase, facilitates enhanced catalytic activity for both cathodic and anodic reactions, alongside robust corrosion resistance. Notably, c-NiFe/a-NiFe LDH can be efficiently scaled up to 25 cm<sup>2</sup>, demonstrating its potential for hydrogen production from saline water using a low-intensity, lab-made photovoltaic cell, highlighting its scalability and efficiency in practical applications.

#### 4.2. Phosphide/sulfide/carbon/alloy based nano-electrocatalysts

Creating phosphorus vacancies in cobalt phosphide embedded in N-doped carbon can modify the electronic structure, provide an optimal adsorption energy with intermediates, enhance the conductivity, and enable the rapid reconstruction of active species for the OER, which could enhance the performance for the HER and OER. Wang *et al.*<sup>15</sup> reported that Co<sub>x</sub>P<sub>v</sub>@NC demonstrated superior activity and lower charge transfer resistance compared to Co<sub>x</sub>P@NC, Co<sub>x</sub>P, and Co<sub>x</sub>P<sub>v</sub> for both the HER and OER in seawater containing 1 M KOH. Specifically, it achieves an impressive HER performance with an  $\eta$  of −260 mV at a high current density of −1000 mA cm<sup>−2</sup> (Table 6) and an OER  $\eta$  of 386 mV at 1000 mA cm<sup>−2</sup> (Table 7). Furthermore, Co<sub>x</sub>P<sub>v</sub>@NC exhibits remarkable activity and stability for overall seawater electrolysis, achieving a voltage of 1.88 V at 1000 mA cm<sup>−2</sup> (Table 8; Fig. 8(j)) and maintaining its stability for 100 h (Fig. 8(k)). It was obtained on Ni foam by the following steps (Fig. 8(a)): first, the MOF (Co(OH)F@ZIF-67) was prepared by hydrothermal treatment followed by immersion; then, Co<sub>x</sub>P@NC was obtained by phosphidation; finally, Co<sub>x</sub>P<sub>v</sub>@NC was obtained by NaBH<sub>4</sub> solution treatment. Co<sub>x</sub>P<sub>v</sub>@NC possesses a nanowire morphology, where tiny lamella structures are observed on the surface (SEM image;

**Table 6** Overpotential ( $\eta$ ) to achieve −10, −50, −100, −500, or −1000 mA cm<sup>−2</sup> for the hydrogen evolution reaction (HER) of various electrocatalysts in seawater based electrolyte, where  $R_{ct}$  and  $R_p$  ( $R_{ct}$ : charge transfer resistance;  $R_p$ : polarization resistance) are acquired from electrochemical impedance spectroscopy (EIS)

| HER electrocatalysts                        | $\eta$ (mV) | $j$ (mA cm <sup>−2</sup> ) | $R_{ct}$ | $R_p$    | Electrolyte                      | Stability                         | Ref. |
|---------------------------------------------|-------------|----------------------------|----------|----------|----------------------------------|-----------------------------------|------|
| Pt <sub>2</sub> /Ni(OH) <sub>2</sub>        | −19         | −10                        | NA       | NA       | 1 M KOH + seawater               | 200 h                             | 100  |
| Ru,Cl-Cu(OH) <sub>2</sub>                   | −37         | −10                        | NA       | NA       | 1 M KOH + seawater               | 12 h (500 mA cm <sup>−2</sup> )   | 99   |
| NRAHM-NiO                                   | −51         | −10                        | 3.015 Ω  | 0.4327 Ω | 1 M KOH + 0.5 M NaCl             | 100 h (500 mA cm <sup>−2</sup> )  | 97   |
| Ni@CNTs-Mo <sub>x</sub> C/Ni <sub>2</sub> P | −65.9       | −10                        | NA       | NA       | 1 M KOH + seawater               | 50 h (50 mA cm <sup>−2</sup> )    | 103  |
| Pt <sub>1</sub> /Mn-NiO                     | −129        | −100                       | NA       | NA       | 0.5 M KOH + seawater             | 30 h (100 mA cm <sup>−2</sup> )   | 96   |
| Co-N,P-HCS                                  | −164        | −10                        | NA       | NA       | 1 M KOH + seawater               | 100 h                             | 50   |
| Cu <sub>2</sub> S@NiS@Ni-NiMo               | −177        | −500                       | 2.16 Ω   | NA       | 1 M NaOH + 0.5 M NaCl            | 2000 h (500 mA cm <sup>−2</sup> ) | 104  |
| Cr-Co <sub>x</sub> P                        | −194        | −100                       | NA       | NA       | 1 M KOH + seawater               | 140 h                             | 105  |
| c-NiFe/a-NiFe LDH                           | −200        | −100                       | ≈5 Ω     | NA       | 1 M KOH + 0.5 M NaCl             | 100 h (500 mA cm <sup>−2</sup> )  | 106  |
| CoFe-Ni <sub>2</sub> P                      | −241        | −100                       | NA       | NA       | 1 M KOH + seawater               | 100 h (500 mA cm <sup>−2</sup> )  | 107  |
| Fe-Mo-Co-O                                  | −250        | −50                        | NA       | NA       | 1 M KOH + seawater               | 72 h (10 mA cm <sup>−2</sup> )    | 94   |
| Co <sub>x</sub> P <sub>v</sub> @NC          | −260        | −1000                      | NA       | NA       | 1 M KOH + seawater               | 100 h (800 mA cm <sup>−2</sup> )  | 15   |
| Cu-MoO <sub>2</sub>                         | −330.2      | −10                        | NA       | NA       | Artificial seawater (neutral pH) | 10 h                              | 93   |
| Mn-Ni <sub>2</sub> P-Fe <sub>2</sub> P      | −470        | −1000                      | NA       | NA       | 1 M KOH + seawater               | 100 h (500 mA cm <sup>−2</sup> )  | 108  |

Ref.: reference;  $R_{ct}$ : charge transfer resistance;  $R_p$ : polarization resistance; HER: hydrogen evolution reaction;  $\eta$ : overpotential; NA: not applicable;  $j$  (mA cm<sup>−2</sup>): current density (mA cm<sup>−2</sup>); NRAHM-NiO: nanorod array based hierarchical NiO microspheres; c: crystalline; a: amorphous.

**Table 7** Overpotential ( $\eta$ ) to achieve 10, 50, 100, 500, or 1000 mA cm<sup>-2</sup> for the oxygen evolution reaction (OER) of various electrocatalysts in seawater based electrolyte, where  $R_{ct}$  and  $R_p$  ( $R_{ct}$ : charge transfer resistance;  $R_p$ : polarization resistance) are acquired from electrochemical impedance spectroscopy (EIS)

| OER electrocatalysts                                 | $\eta$ (mV) | $j$ (mA cm <sup>-2</sup> ) | $R_{ct}$      | $R_p$           | Electrolyte                      | Stability                         | Ref. |
|------------------------------------------------------|-------------|----------------------------|---------------|-----------------|----------------------------------|-----------------------------------|------|
| NRAHM-NiO                                            | 70          | 10                         | 3.4 $\Omega$  | 0.5611 $\Omega$ | 1 M KOH + 0.5 M NaCl             | 100 h (500 mA cm <sup>-2</sup> )  | 97   |
| CoFe-Ni <sub>2</sub> P                               | 224         | 10                         | NA            | NA              | 1 M KOH + seawater               | 130 h (500 mA cm <sup>-2</sup> )  | 107  |
| Ni@CNTs-Mo <sub>x</sub> C/Ni <sub>2</sub> P          | 230         | 10                         | NA            | NA              | 1 M KOH + seawater               | 50 h (50 mA cm <sup>-2</sup> )    | 103  |
| s-NiFeCrMo-OH                                        | 242         | 100                        | NA            | NA              | 1 M KOH + 0.5 M NaCl             | 400 h (500 mA cm <sup>-2</sup> )  | 101  |
| Ni(Fe)(OH) <sub>2</sub>                              | 257         | 10                         | NA            | NA              | 1 M KOH + seawater               | 100 h (250 mA cm <sup>-2</sup> )  | 98   |
| CoFePBA-Co <sub>2</sub> P                            | 257         | 10                         | 0.57 $\Omega$ | NA              | 1 M NaOH + 0.5 M NaCl            | 100 h (2000 mA cm <sup>-2</sup> ) | 110  |
| CoCr <sub>0.7</sub> Rh <sub>1.3</sub> O <sub>4</sub> | 273         | 10                         | NA            | NA              | 1 M KOH + 0.5 M NaCl             | 12 h (20 mA cm <sup>-2</sup> )    | 95   |
| Fe <sub>1</sub> /Mn-NiO                              | 290         | 100                        | NA            | NA              | 0.5 M KOH + seawater             | 30 h (100 mA cm <sup>-2</sup> )   | 96   |
| RuMoNi                                               | 291         | 100                        | NA            | NA              | 1 M KOH + seawater               | 3000 h (500 mA cm <sup>-2</sup> ) | 111  |
| c-NiFe/a-NiFe LDH                                    | 300         | 100                        | NA            | NA              | 1 M KOH + 0.5 M NaCl             | 100 h (500 mA cm <sup>-2</sup> )  | 106  |
| Co,Cl-Cu(OH) <sub>2</sub>                            | 310         | 100                        | NA            | NA              | 1 M KOH + seawater               | 10 h (500 mA cm <sup>-2</sup> )   | 99   |
| (Ni, Fe)OOH@Ni <sub>x</sub> P                        | 318         | 500                        | 0.85 $\Omega$ | NA              | 1 M KOH + seawater               | 100 h (500 mA cm <sup>-2</sup> )  | 92   |
| Co-N,P-HCS                                           | 320         | 10                         | NA            | NA              | 1 M KOH + seawater               | 100 h                             | 50   |
| Ru/NiFeOOH                                           | 330         | 500                        | NA            | NA              | 1 M KOH + seawater               | 400 h (100 mA cm <sup>-2</sup> )  | 102  |
| Cr-Co <sub>x</sub> P                                 | 334         | 100                        | NA            | NA              | 1 M KOH + seawater               | 140 h                             | 105  |
| Mn-Ni <sub>2</sub> P-Fe <sub>2</sub> P               | 358         | 1000                       | NA            | NA              | 1 M KOH + seawater               | 100 h (500 mA cm <sup>-2</sup> )  | 108  |
| Fe-Mo-Co-O                                           | 365         | 50                         | NA            | NA              | 1 M KOH + seawater               | 72 h (10 mA cm <sup>-2</sup> )    | 94   |
| Co <sub>x</sub> P <sub>y</sub> @NC                   | 386         | 1000                       | NA            | NA              | 1 M KOH + seawater               | 100 h (800 mA cm <sup>-2</sup> )  | 15   |
| Co-(NiFe)N@NiS <sub>x</sub>                          | 425         | 100                        | 0.14 $\Omega$ | NA              | Artificial seawater (neutral pH) | 50 h (100 mA cm <sup>-2</sup> )   | 112  |
| NiS <sub>2</sub> pS <sub>x</sub>                     | 460         | 10                         | 33.7 $\Omega$ | NA              | 0.5 M NaCl (neutral pH)          | 90 h                              | 113  |

Ref.: reference;  $R_{ct}$ : charge transfer resistance;  $R_p$ : polarization resistance; OER: oxygen evolution reaction;  $\eta$ : overpotential; NA: not applicable;  $j$  (mA cm<sup>-2</sup>): current density (mA cm<sup>-2</sup>); NRAHM-NiO: nanorod array based hierarchical NiO microspheres; c: crystalline; a: amorphous.

**Table 8** Potential to achieve 10, 100, 200, or 1000 mA cm<sup>-2</sup> for overall water splitting for several kinds of bifunctional electrocatalysts in seawater based electrolyte, where non-bifunctional electrocatalysts are indicated by "nb"

| Bifunctional electrocatalysts for HER and OER                                         | Overall water splitting (V) | $j$ (mA cm <sup>-2</sup> ) | Electrolyte          | Stability                         | Ref. |
|---------------------------------------------------------------------------------------|-----------------------------|----------------------------|----------------------|-----------------------------------|------|
| nbFe <sub>1</sub> /Mn-NiO <sup>OER</sup> //Pt <sub>1</sub> /Mn-NiO <sup>HER</sup>     | 1.44                        | 10                         | 0.5 M KOH + seawater | 30 h (100 mA cm <sup>-2</sup> )   | 96   |
| Cr-Co <sub>x</sub> P                                                                  | 1.54                        | 10                         | 1 M KOH + seawater   | 160 h                             | 105  |
| nbCo,Cl-Cu(OH) <sub>2</sub> <sup>OER</sup> //Ru,Cl-Cu(OH) <sub>2</sub> <sup>HER</sup> | 1.55                        | 10                         | 1 M KOH + seawater   | 100 h (200 mA cm <sup>-2</sup> )  | 99   |
| Ni@CNTs-Mo <sub>x</sub> C/Ni <sub>2</sub> P                                           | 1.56                        | 10                         | 1 M KOH + seawater   | 50 h (50 mA cm <sup>-2</sup> )    | 103  |
| Fe-Mo-Co-O                                                                            | 1.59                        | 10                         | 1 M KOH + seawater   | 72 h (10 mA cm <sup>-2</sup> )    | 94   |
| Mn-Ni <sub>2</sub> P-Fe <sub>2</sub> P                                                | 1.64                        | 10                         | 1 M KOH + seawater   | 120 h (500 mA cm <sup>-2</sup> )  | 108  |
| NRAHM-NiO                                                                             | 1.66                        | 100                        | 1 M KOH + 0.5 M NaCl | 168 h (240 mA cm <sup>-2</sup> )  | 97   |
| CoFe-Ni <sub>2</sub> P                                                                | 1.738                       | 100                        | 1 M KOH + seawater   | 350 h (1000 mA cm <sup>-2</sup> ) | 107  |
| nb(Ni, Fe)OOH@Ni <sub>x</sub> P <sup>OER</sup> //Ni <sub>x</sub> P <sup>HER</sup>     | 1.81                        | 500                        | 1 M KOH + seawater   | 50 h (500 mA cm <sup>-2</sup> )   | 92   |
| Co <sub>x</sub> P <sub>y</sub> @NC                                                    | 1.884                       | 1000                       | 1 M KOH + seawater   | 100 h (200 mA cm <sup>-2</sup> )  | 15   |
| Co-N,P-HCS                                                                            | ≈2.12                       | 200                        | 1 M KOH + seawater   | 1000 h (10 mA cm <sup>-2</sup> )  | 50   |
| Ni-B-P                                                                                | 2.76                        | 500                        | 1 M KOH + seawater   | 8 h (500 mA cm <sup>-2</sup> )    | 114  |

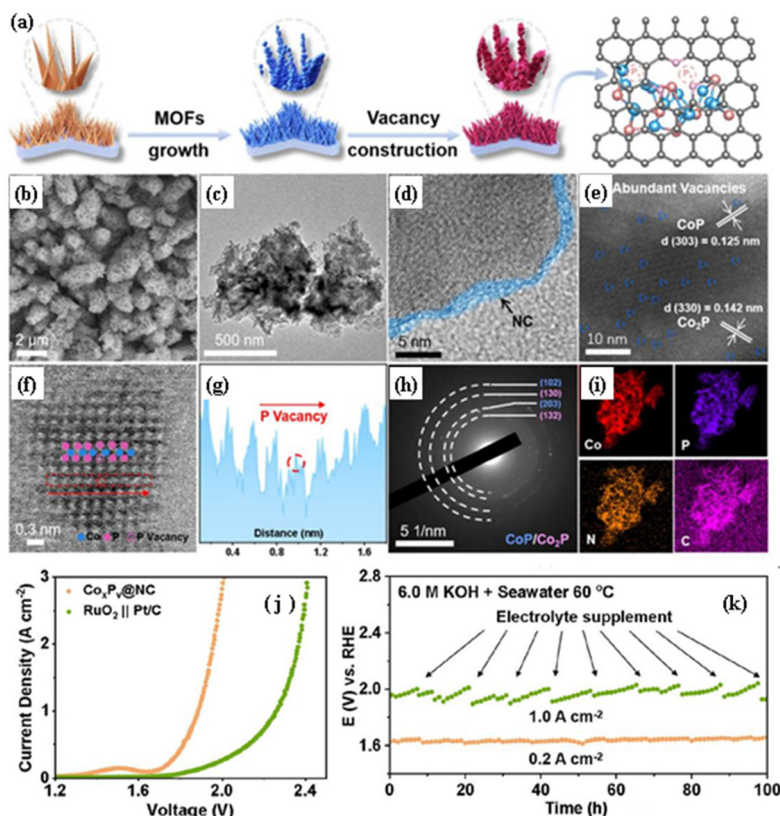
nb: non-bifunctional electrocatalyst; Ref.: reference;  $j$  (mA cm<sup>-2</sup>): current density (mA cm<sup>-2</sup>); NRAHM-NiO: nanorod array based hierarchical NiO microspheres.

Fig. 8(b)). The ultrathin nanosheets further wrapped the outer layer of nanowires (TEM image; Fig. 8(c)). The surface of Co<sub>x</sub>P is wrapped in a NC shell with ~0.5 nm thickness (HRTEM image; Fig. 8(d)). The lattice fringes of Co<sub>2</sub>P and CoP inside the carbon layers reveal  $d$ -spacings of 0.142 nm and 0.125 nm, which correspond to the (330) and (303) lattice planes of Co<sub>2</sub>P and CoP, respectively (HRTEM image; Fig. 8(e)), while abundant vacancies and lattice distortions are observed. The marked red circles show the anionic P vacancies (STEM-BF image; Fig. 8(f)), while the corresponding intensity profile depicts the anionic P vacancies (Fig. 8(g)). Co<sub>x</sub>P<sub>y</sub>@NC is polycrystalline, where the diffraction rings correspond to the facets of CoP and Co<sub>2</sub>P (SAED pattern; Fig. 8(h)), while it contains Co, P, N, and C, which are uniformly distributed (Fig. 8(i)),

40.4 wt% Co<sub>x</sub>P<sub>y</sub>. DFT calculations and *in situ* Raman spectra reveal that the enhancement of the catalytic performance is attributed to the quick reconstruction of active species for the OER, and P vacancies provide an optimal adsorption energy with intermediates for the HER.

The fabrication of asymmetric Co-N<sub>3</sub>P<sub>1</sub> implanted into carbon can provide symmetry-breaking electronic structures, enhance the conductivity, enable the affinity of strong oxygen-containing intermediates with weak Cl<sup>-</sup> adsorption and moderate H adsorption, and that could enhance the performance for HER and OER. Wang *et al.*<sup>50</sup> observed that Co-N<sub>3</sub>P-HCS (HCS: hollow carbon spheres) exhibited higher activity and lower charge transfer resistance than those of HCS, Co-HCS, and Co-N-HCS for the HER and OER in seawater with 1 M KOH, while it





**Fig. 8** (a) Schematic diagram illustrating the synthesis process for  $\text{Co}_x\text{P}_y\text{@NC}$ ; (b) SEM image and (c) TEM image of  $\text{Co}_x\text{P}_y\text{@NC}$ ; (d and e) HRTEM images and (f) STEM-BF image of  $\text{Co}_x\text{P}_y\text{@NC}$ ; (g) intensity profile obtained from the corresponding areas in part (f); (h) SAED pattern and (i) EDS mapping of  $\text{Co}_x\text{P}_y\text{@NC}$  for Co, P, N and C; (j) LSV curves for  $\text{Co}_x\text{P}_y\text{@NC}$ // $\text{Co}_x\text{P}_y\text{@NC}$  in comparison with those of the  $\text{RuO}_2\text{||Pt/C}$  couple in 1 M KOH with seawater; (k) chronopotentiometry curve of  $\text{Co}_x\text{P}_y\text{@NC}$ // $\text{Co}_x\text{P}_y\text{@NC}$  in 6 M KOH with seawater at 60 °C (reproduced with permission from ref. 15. Copyright 2023, Elsevier Ltd).

exhibited high stability for 1000 h. Co–N,P–HCS is composed of an asymmetric Co– $\text{N}_3\text{P}_1$  structure, which is implanted on hollow carbon spheres. The DFT calculations disclose that the asymmetric Co– $\text{N}_3\text{P}_1$  sites with symmetry-breaking electronic structures can be formed due to P-doping into Co–N–C, enabling the affinity of strong oxygen-containing intermediates with weak  $\text{Cl}^-$  adsorption and moderate H adsorption.

Bahuguna *et al.*<sup>113</sup> reported that  $\text{NiS}_{2p}\text{S}_x$  achieved an overpotential of 460 mV at  $10 \text{ mA cm}^{-2}$  for the oxygen evolution reaction (OER) in 0.5 M NaCl (neutral pH), highlighting its exceptional activity. It also maintains high stability over 90 h and presents a low charge transfer resistance of  $33.7 \Omega$ . 3D  $\text{NiS}_{2p}\text{S}_x$  is a binder-free, catenated sulphur type Ni polysulfide-based chlorine-phobic electrocatalyst.  $\text{NiS}_{2p}\text{S}_x$  produces pure oxygen without the formation of any chlorine in 0.5 M NaCl. This high selectivity towards oxygen evolution is attributed to the limited accessibility of the electrostatically and sterically hindered Ni metal centers to the larger  $\text{Cl}^-$  ions. Furthermore, the generation of catenated sulphur type polysulfide species ( $\text{pSn}^{2-}/\text{S}^{2-} = 2.1$ ) due to the rapid electrochemical evolution of the  $\text{NiS}_{2p}\text{S}_x$  surface can further facilitate the intrinsic chlorine-phobicity of the electrocatalyst,

which can lead to pure oxygen evolution at current densities up to  $300 \text{ mA cm}^{-2}$ .

Additionally, Badreldin *et al.*<sup>112</sup> reported that the Co–(NiFe) N@ $\text{NiS}_x$  catalyst exhibited a remarkable performance for the OER in artificial seawater at neutral pH. It achieves an  $\eta$  of 425 mV at  $100 \text{ mA cm}^{-2}$ , indicating its exceptionally high activity. Furthermore, it maintains a current retention of 84.3% to deliver approximately  $200 \text{ mA cm}^{-2}$  after 168 h, highlighting its remarkable stability. The catalyst also exhibits a low charge transfer resistance ( $0.14 \Omega$ ) and a high faradaic efficiency of 87.4% at 1.75 V (vs. RHE), further confirming its outstanding performance for oxygen evolution reactions in seawater. Co–(NiFe)N@ $\text{NiS}_x$  is composed of Co–(NiFe) oxide/nitride with  $\text{NiS}_x$  as an interlayer. The formation of stable anionic  $[\text{O–N}]^{\delta-}$  and cationic  $[\text{Co–(NiFe)}]^{\delta+}$  surface species coupled with the  $\text{NiS}_x$  interlayer can facilitate the dissociation of  $\text{H}_2\text{O}$ , which could improve the activity and kinetics for the OER.

Xu *et al.*<sup>104</sup> demonstrated that  $\text{Cu}_2\text{S@NiS@Ni–NiMo}$  achieved an impressive  $-177 \text{ mV}$  overpotential for the HER at  $-500 \text{ mA cm}^{-2}$  in 1 M NaOH + 0.5 M NaCl, showcasing excellent activity. It maintains a stable performance without decay for 2500 h at the same current density and exhibits a low

charge transfer resistance of 2.16  $\Omega$ , underscoring its exceptional durability and conductive efficiency. The synthesis route to  $\text{Cu}_2\text{S@NiS@Ni-NiMo}$  on Cu foam can be scalable to over 100  $\text{cm}^2$ . The enhanced performance for the HER of  $\text{Cu}_2\text{S@NiS@Ni-NiMo}$  is attributed to the following factors: (1) the unique boundary interface can enhance the high intrinsic activity, (2) the superaerophobic nanoarray architecture can facilitate the evolution of gas and improve mass transport, (3) the polyanion-rich passivating layers can improve the corrosion resistance.

Huang *et al.*<sup>107</sup> reported that  $\text{CoFe-Ni}_2\text{P}$  achieved an overpotential of 224 mV at 10  $\text{mA cm}^{-2}$  for the OER in 1 M KOH + seawater, indicating exceptional activity. Remarkably, it maintains its stability, with no decay observed over 500 h at a current density of 500  $\text{mA cm}^{-2}$ , highlighting its superior stability and minimal charge transfer resistance. For the HER,  $\text{CoFe-Ni}_2\text{P}$  reaches an overpotential of -241 mV at -100  $\text{mA cm}^{-2}$ , further confirming its high efficiency. In the context of overall water splitting in the same medium,  $\text{CoFe-Ni}_2\text{P}$  achieves a potential of 1.738 V at 100  $\text{mA cm}^{-2}$ , showcasing its outstanding activity. Moreover, it demonstrates enduring stability over 100 h at a current density of 500  $\text{mA cm}^{-2}$ , underscoring its durability and efficient performance.  $\text{CoFe-Ni}_2\text{P}$  is composed of nickel phosphide, which is modified by Co and Fe, where the conductivity of  $\text{Ni}_2\text{P}$  can be enhanced by Fe atoms to facilitate electron transfer, while the self-reconstruction process can be enhanced by Co atoms to form  $\text{CoFe-NiOOH}$  species *in situ* on the surface of the electrode. Furthermore, the selective adsorption of  $\text{OH}^-$  ions, which significantly hinders the adsorption of  $\text{Cl}^-$  ions, can be achieved by the *in situ*-generated  $\text{CoFe-NiOOH}$  species, and this could enhance the performance for seawater splitting even at a huge current density.

Besides, Liu *et al.*<sup>110</sup> found that  $\text{CoFePBA-Co}_2\text{P}$  demonstrated an overpotential of 257 mV at 10  $\text{mA cm}^{-2}$  for the OER in 1 M NaOH + 0.5 M NaCl, indicating exceptional activity. Remarkably, this catalyst maintains its performance, generating a current density of 1000  $\text{mA cm}^{-2}$  over 1000 h in 20 wt% NaOH + seawater and sustaining a current density of 2000  $\text{mA cm}^{-2}$  for 100 h in 20 wt% NaOH + saturated NaCl, highlighting its outstanding stability. Additionally, it shows a low charge transfer resistance of 0.57  $\Omega$ , further confirming its efficient electrochemical behavior.  $\text{CoFePBA-Co}_2\text{P}$  is composed of cobalt ferricyanide/cobalt phosphide with a cap/pin structure. Molecular dynamics simulations and experimental investigations reveal that  $\text{PO}_4^{3-}$  and  $\text{Fe(CN)}_6^{3-}$  formed by the electrode can effectively hinder  $\text{Cl}^-$  through dense coverage and electrostatic repulsion, leading to about a 5-fold decrease in  $\text{Cl}^-$  adsorption, thereby enhancing the corrosion resistance.

Notably, Song *et al.*<sup>105</sup> discovered that  $\text{Cr-Co}_x\text{P}$  delivered an overpotential of 334 mV at 100  $\text{mA cm}^{-2}$  for the OER and -194 mV for the HER at -100  $\text{mA cm}^{-2}$  in 1 M KOH + seawater, indicating outstanding activity for both processes. This catalyst maintains its performance with remarkable stability over 140 h for both the OER and HER, alongside demonstrating a low charge transfer resistance. In the context of overall

water splitting in the same medium,  $\text{Cr-Co}_x\text{P}$  achieves a potential of 1.54 V at 10  $\text{mA cm}^{-2}$ , maintaining exceptional activity and extending its stability to 160 h, highlighting its efficiency and durability under challenging conditions.  $\text{Cr-Co}_x\text{P}$  is composed of  $\text{Co}_x\text{P}$ , which is doped with vein-like Cr. DFT calculations and experimental studies reveal that the doping of  $\text{Co}_x\text{P}$  with Cr can regulate the electronic structure of  $\text{Co}_x\text{P}$  and effectively obviate the work function of  $\text{Co}_x\text{P}$ , which can facilitate electron transfer between absorbates and the catalyst surface, and decrease the energy barriers for water dissociation and the rate-determining steps for the OER and HER. Furthermore, the doping of  $\text{Co}_x\text{P}$  with Cr can also prevent Co sites with robust valence states to retain their high performance during the OER process, which can enhance the performance for seawater electrolysis.

Luo *et al.*<sup>108</sup> demonstrated that the  $\text{Mn-Ni}_2\text{P-Fe}_2\text{P}$  catalyst, synthesized *via* a hydrothermal-phosphorization method, exhibited an outstanding electrocatalytic performance for both oxygen and hydrogen evolution reactions (OER and HER) in 1 M KOH + seawater. Specifically, it achieves an overpotential of 358 mV at a current density of 1000  $\text{mA cm}^{-2}$  for the OER and -470 mV at the same density for the HER, highlighting its exceptional activity. Remarkably, this catalyst shows high stability, maintaining around 500  $\text{mA cm}^{-2}$  for 200 h in the OER and for over 100 h in the HER, with minimal decay observed. Furthermore, in overall water splitting,  $\text{Mn-Ni}_2\text{P-Fe}_2\text{P}$  achieves a potential of 1.64 V at 10  $\text{mA cm}^{-2}$ , sustaining nearly 500  $\text{mA cm}^{-2}$  for 120 h, indicating its durability and low charge transfer resistance. This combination of high activity and stability underlines the effectiveness of the  $\text{Mn-Ni}_2\text{P-Fe}_2\text{P}$  system for sustainable water electrolysis in seawater. The enhanced performance of  $\text{Mn-Ni}_2\text{P-Fe}_2\text{P}$  for seawater electrolysis is attributed to the following factors: (1) Mn doping and  $\text{Ni}_2\text{P/Fe}_2\text{P}$  heterointerfaces can provide optimal electronic structures that can provide an optimal adsorption energy for intermediates to enhance the intrinsic HER and OER activities. (2) Mn doping and heterointerfaces can generate abundant active sites, while the 3D nanoflower structure can expose abundant active sites. (3) Nickel-iron phosphides can provide high corrosion resistance.

In addition, Ahasan Habib *et al.*<sup>114</sup> observed that  $\text{Ni-B-P//Ni-B-P}$  afforded a potential of 2.76 V at a huge current density of 500  $\text{mA cm}^{-2}$ , suggesting substantial activity and stability for overall water splitting in 1 M KOH + seawater. It was obtained through a hydrothermal approach.  $\text{Ni-B-P}$  is composed of nickel-boride-phosphide microspheres, which can modify the electronic structure, enhance the conductivity, afford/expose abundant active sites, and enhance the corrosion resistance, which could enhance the performance for seawater electrolysis.

Wang *et al.*<sup>103</sup> found that  $\text{Ni@CNTs-Mo}_x\text{C/Ni}_2\text{P}$  delivered an overpotential of 230 mV at 10  $\text{mA cm}^{-2}$  for the OER and -65.9 mV for the HER at -10  $\text{mA cm}^{-2}$  in 1 M KOH + seawater, indicating its outstanding activity for both reactions. This catalyst also maintains its performance with high stability over 50 h and demonstrates a low charge transfer resistance. In

terms of overall water splitting in the same medium, Ni@CNTs-Mo<sub>x</sub>C/Ni<sub>2</sub>P achieves a potential of 1.56 V at 10 mA cm<sup>-2</sup>, maintaining its excellent activity and stability for 50 h, showcasing its efficiency and durability in electrolysis applications. Ni@CNTs-Mo<sub>x</sub>C/Ni<sub>2</sub>P is composed of hierarchically structured molybdenum carbide/nickel phosphide micro-pillar heterostructures, which are bridged with nickel nanoparticles encapsulated in carbon nanotubes that can enhance the conductivity, provide abundant multiple electroactive sites and an optimal electronic structure to elevate the density of states near the Fermi level, afford an optimal adsorption free energy, and enhance the performance for seawater electrolysis.

Moreover, Kang *et al.*<sup>111</sup> reported that RuMoNi achieved an overpotential of 291 mV at 100 mA cm<sup>-2</sup> for the OER in 1 M KOH + seawater, highlighting its exceptional activity. Impressively, it maintains strong stability over 3000 h at 500 mA cm<sup>-2</sup> and demonstrates a low charge transfer resistance, underscoring its efficiency and endurance under electrolysis conditions. RuMoNi exhibits the *in situ* formation of molybdate ions on its surface, which can repel chloride ions, and enhance the corrosion resistance and performance for seawater electrolysis. An anion exchange membrane electrolyzer with the RuMoNi catalyst delivers a huge current density of 1000 mA cm<sup>-2</sup> at 1.72 V with an energy conversion efficiency of 77.9%, exhibiting a GGE of \$0.85, which is far lower than the GGE of \$2.0 (\$2.0 GGE: 2026 technical target of the United

Stated Department of Energy; GGE: price per gallon of gasoline equivalent for H<sub>2</sub> generation).

## 5. Photocatalytic water splitting

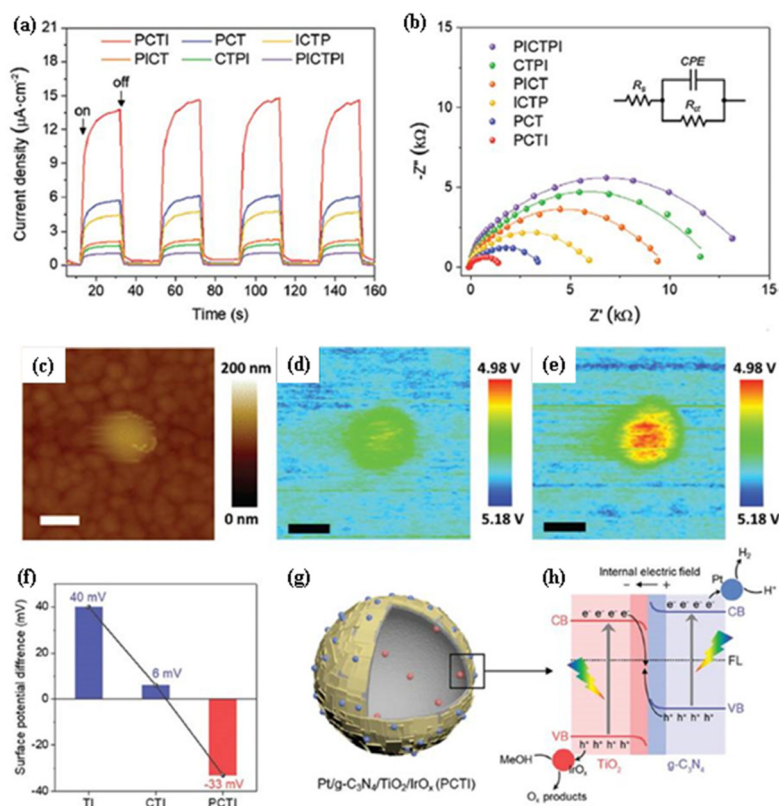
### 5.1. Nano-photocatalysts with a heterojunction/cocatalyst loading/defect engineering/homojunction

Spatial separation of suitable cocatalysts on a Z-scheme heterostructure can enhance the surface reaction rate and make charge separation and transfer of the photogenerated electrons and holes efficient, which can enhance photocatalytic H<sub>2</sub> evolution. Moon *et al.*<sup>7</sup> prepared a Z-scheme Pt/g-C<sub>3</sub>N<sub>4</sub>/TiO<sub>2</sub>/IrO<sub>x</sub> (PCTI) hollow sphere photocatalyst through electrostatic interactions. PCTI with efficient light absorption and interfacial charge separation abilities shows a much higher H<sub>2</sub> evolution rate (8.15 mmol h<sup>-1</sup> g<sup>-1</sup>; Table 9) and a considerably higher apparent quantum yield (24.3% at 330 nm) with 0.5 wt% Pt cocatalyst on g-C<sub>3</sub>N<sub>4</sub> and 1.2 wt% IrO<sub>x</sub> cocatalyst on TiO<sub>2</sub>. Moreover, to further facilitate the reaction kinetics and surface charge separation, Pt and IrO<sub>x</sub>, as two cocatalysts, have been spatially divided along the Z-scheme charge-transfer pathway. The photoelectrochemical measurements demonstrate the role of Pt and IrO<sub>x</sub> cocatalysts on g-C<sub>3</sub>N<sub>4</sub>/TiO<sub>2</sub>. The Pt and IrO<sub>x</sub> cocatalysts loaded on g-C<sub>3</sub>N<sub>4</sub>/TiO<sub>2</sub> (PCTI) exhibit significantly higher transient photocurrent density when compared to that

**Table 9** Photocatalytic H<sub>2</sub> evolution rate and apparent quantum yield (AQY) of various kinds of photocatalysts with/without cocatalysts using pure water/sacrificial agent as a reaction medium

| Photocatalysts with/without cocatalysts                                                 | Light source                                 | Reaction medium                                     | H <sub>2</sub> evolution rate (mmol h <sup>-1</sup> g <sup>-1</sup> ) | AQY           | Ref. |
|-----------------------------------------------------------------------------------------|----------------------------------------------|-----------------------------------------------------|-----------------------------------------------------------------------|---------------|------|
| N-NaTaO <sub>3</sub> @Ta <sub>3</sub> N <sub>5</sub> /Rh@Cr <sub>2</sub> O <sub>3</sub> | 300 W Xe lamp                                | Pure water                                          | 25.12                                                                 | 9.32%@350 nm  | 11   |
| Ru-In SA/TiO <sub>2</sub>                                                               | 300 W Xe lamp ( $\lambda \geq 300$ nm)       | Pure water                                          | ~4.8362                                                               | NA            | 8    |
| PCN-Ni-CoP                                                                              | 300 W Xe lamp                                | Pure water                                          | 0.143                                                                 | 2.7%@380 nm   | 116  |
| Ni@NiS <sub>x</sub> -CdS                                                                | 300 W Xe lamp ( $\lambda > 420$ nm)          | Na <sub>2</sub> S + Na <sub>2</sub> SO <sub>3</sub> | 78.7                                                                  | 36.74%@420 nm | 117  |
| NiSeS/ZnSe                                                                              | Visible light ( $\lambda > 400$ nm)          | Ascorbic acid                                       | 73.54                                                                 | 57.68%@400 nm | 118  |
| TSP                                                                                     | 300 W Xe lamp (200 nm < $\lambda$ < 1000 nm) | Methanol                                            | 62.37                                                                 | 45.9%@365 nm  | 115  |
| Pt-CdS <sub>x</sub> -T                                                                  | 300 W Xe lamp ( $\lambda \geq 420$ nm)       | Lactic acid                                         | 42.7                                                                  | 43.6%@470 nm  | 119  |
| RuNi/g-C <sub>3</sub> N <sub>4</sub>                                                    | 300 W Xe lamp (>400 nm)                      | Triethanolamine                                     | 35.1                                                                  | 1.27%@380 nm  | 120  |
| ZrC@ZnIn <sub>2</sub> S <sub>4</sub>                                                    | 300 W Xe lamp (>800 nm)                      | Triethanolamine                                     | 32.87c                                                                | NA            | 121  |
| NiCo <sub>2</sub> S <sub>4</sub> /ZnIn <sub>2</sub> S <sub>4</sub>                      | 300 W Xe lamp ( $\lambda > 420$ nm)          | Ascorbic acid                                       | 19.654d                                                               | 56.7%@400 nm  | 17   |
| 15% CoS <sub>2</sub> /CdS-L                                                             | 5 W LED                                      | Lactic acid                                         | 19.22                                                                 | 22.6%@420 nm  | 55   |
| Au@MoS <sub>2</sub> -ZnO <sup>b</sup>                                                   | AM 1.5 G filter                              | Na <sub>2</sub> SO <sub>3</sub>                     | 13.09                                                                 | 21.3a%@420 nm | 122  |
| Ru@Nb <sub>2</sub> O <sub>5</sub> /Nb <sub>2</sub> C                                    | 300 W Xe lamp                                | Glycerol                                            | 10.11                                                                 | 41.25%@313 nm | 123  |
| Zn-Ni <sub>3</sub> P <sub>4</sub> /CdS                                                  | 300 W Xe lamp                                | Lactic acid                                         | 8.969                                                                 | NA            | 124  |
| CDs/CCN                                                                                 | 300 W Xe lamp                                | Triethanolamine                                     | 8.682                                                                 | 1.41%@420 nm  | 125  |
| Pt/g-C <sub>3</sub> N <sub>4</sub> /TiO <sub>2</sub> /IrO <sub>x</sub>                  | 300 W Xe lamp                                | Methanol                                            | 8.153                                                                 | 24.3%@330 nm  | 7    |
| Protonated NH <sub>2</sub> -MIL-125                                                     | 300 W Xe lamp                                | Acetonitrile + triethanolamine                      | 6.7628                                                                | NA            | 126  |
| CN/CNQDs                                                                                | 300 W Xe lamp                                | Triethanolamine                                     | 4.0877                                                                | 6.74%@420 nm  | 127  |
| ClZS/MoS <sub>2</sub> /CDs                                                              | LED light ( $\lambda \geq 420$ nm)           | L-Ascorbic acid                                     | 3.706                                                                 | 34.23%@450 nm | 128  |
| TiO <sub>2</sub> -Pt/FeHAp                                                              | 300 W Xe lamp                                | Triethanolamine                                     | 3.026                                                                 | 4.72%@370 nm  | 129  |
| PtO@Ti <sub>3</sub> C <sub>2</sub> /TiO <sub>2</sub>                                    | 300 W Xe lamp                                | Methanol                                            | 2.54                                                                  | 4.2%@365 nm   | 130  |
| Amorphous-NiB/TiO <sub>2</sub>                                                          | 4 LED lamps (3 W; 365 nm)                    | Lactic acid                                         | 2.334                                                                 | 16.30%        | 131  |
| Bi <sub>20</sub> TiO <sub>32</sub> /Bi <sub>4</sub> Ti <sub>3</sub> O <sub>12</sub>     | 250 W Xe lamp ( $\lambda > 400$ nm)          | Methanol                                            | 1.89                                                                  | NA            | 132  |
| BOC-CNF@CNT                                                                             | 300 W Xe lamp                                | Triethanolamine                                     | 0.02542                                                               | NA            | 133  |

Ref.: reference; NA: not applicable; AQY: apparent quantum yield; TSP: twin S-scheme photocatalyst comprising of 2D graphitic carbon nitride nanosheets, which is inserted between anatase TiO<sub>2</sub> nanoparticles and H-doped rutile TiO<sub>2</sub> nanorods, HAp: hydroxyapatite; ClZS/MoS<sub>2</sub>/CDs: Cu-In-Zn-S quantum dots/MoS<sub>2</sub>/carbon dots; QDs: quantum dots; CDs/CCN: carbon dots/C-doping g-C<sub>3</sub>N<sub>4</sub>; 21.3a%@420 nm for Au@MoS<sub>2</sub>; Au@MoS<sub>2</sub>-ZnOb: piezo-photocatalyst (H<sub>2</sub> evolution was obtained under light and ultrasonic waves); 32.87c: photothermal-assisted photocatalytic H<sub>2</sub> evolution rate of 32.87 mmol h<sup>-1</sup> g<sup>-1</sup>; 19.654d: photothermal-assisted photocatalytic H<sub>2</sub> evolution rate of 19.654 mmol h<sup>-1</sup> g<sup>-1</sup>.



**Fig. 9** (a) Chopped photocurrent–time response of PCTI in comparison with those of other photocatalysts recorded at 0.0 V versus Ag/AgCl in 0.5 M Na<sub>2</sub>SO<sub>4</sub>, where sequential illumination of 20 s on/off time intervals were used. (b) Nyquist plot of PCTI in comparison with other photocatalysts obtained at 0.6 V versus Ag/AgCl in 0.5 M Na<sub>2</sub>SO<sub>4</sub> under light irradiation (inset depicts the corresponding equivalent circuit). (c) Kelvin probe force microscopy (KPFM) topography mapping of PCTI, (d) corresponding surface potential (SP) mapping of PCTI in darkness, and (e) corresponding SP mapping of PCTI under light irradiation. (f) ΔSP of PCTI in comparison with TI and CTI, which was calculated based on evolution of the structure. (g) Schematic diagram illustrating the physical structure of PCTI. (h) Schematic diagram illustrating the charge-transfer process within PCTI under light irradiation (reproduced with permission from ref. 7. Copyright 2022, Wiley-VCH GmbH).

of other cocatalysts loaded on g-C<sub>3</sub>N<sub>4</sub>/TiO<sub>2</sub> as shown in Fig. 9(a), where the chopped photocurrent–time response of the photocatalysts was recorded at 0.0 V versus Ag/AgCl in 0.5 M Na<sub>2</sub>SO<sub>4</sub> (sequential illumination: 20 s on/off time interval). Enhanced surface charge separation of PCTI is observed, where the photogenerated electrons and photoexcited holes can be trapped at Pt on the external surface of CT and IrO<sub>x</sub> on the internal surface of CT, respectively, which could enhance the performance for the water redox reaction. Moreover, the Pt and IrO<sub>x</sub> cocatalysts loaded on g-C<sub>3</sub>N<sub>4</sub>/TiO<sub>2</sub> (PCTI) exhibit lower charge transfer resistance ( $R_{ct} = 1.52$  kΩ) when compared to that of other cocatalysts loaded on g-C<sub>3</sub>N<sub>4</sub>/TiO<sub>2</sub> as shown in Fig. 9(b), for Nyquist plots obtained at 0.6 V versus Ag/AgCl in 0.5 M Na<sub>2</sub>SO<sub>4</sub> under light irradiation. This suggests a faster charge transfer process in PCTI. These results further reveal efficient charge separation and transfer for spatially separated cocatalysts. The spatial charge separation in PCTI under irradiation was further demonstrated by photo-KPFM analysis (KPFM: Kelvin probe force microscopy). Fig. 9(c) depicts the KPFM topography mapping of PCTI, while the corresponding surface potential (SP) mapping of PCTI in darkness and under light irradiation is depicted in Fig. 9(d) and (e), respectively.

Moreover, Fig. 9(f) depicts ΔSP (difference in SP) of PCTI in comparison with TI and CTI, which is calculated based on structural evolution to analyze the photoresponse and photo-generated electron–hole pairs of the photocatalyst. In general, if light strikes a semiconductor, the photoexcited electrons can accumulate at the material's surface, leading to the negative shift in SP, whereas a positive shift in SP can be observed if the electron density decreases. The ΔSP of 40 mV (highly increased SP) for TI after light illumination denotes a decrease in the electron density at the surface, suggesting the transfer of electrons to the internal IrO<sub>x</sub> particles from the external TiO<sub>2</sub> shell due to the large work function of IrO<sub>x</sub>. Nevertheless, the ΔSP of 6 mV (slightly increased SP) for CTI is much lower than that of the ΔSP of 40 mV (highly increased SP) for CTI after light illumination due to the formation of a heterojunction of TI with g-C<sub>3</sub>N<sub>4</sub>, where the photoexcited electrons can be accumulated on the outermost g-C<sub>3</sub>N<sub>4</sub> layer and the inward migration of electrons can be decreased because of the Z-scheme charge transfer. In contrast, the ΔSP of −33 mV (highly decreased SP) for PCTI after light illumination denotes that photoexcited electrons can migrate to g-C<sub>3</sub>N<sub>4</sub> and the external PCTI surface, where the migrated electrons can be



particularly trapped by the Pt cocatalyst. The efficiency of the  $\text{H}_2$  evolution rate can be primarily determined by the surface reaction rate and the efficiency of photogenerated carrier separation and migration. Thus, the significantly higher  $\text{H}_2$  evolution rate ( $8.15 \text{ mmol h}^{-1} \text{ g}^{-1}$ ) of PCTI can be attributed to the following factors, as shown in Fig. 9(g) and (h): (1) enhanced charge separation and migration efficiency at the bulk phase and surface because of the synergy of the Z-scheme CT heterostructure; (2) separately loaded Pt and  $\text{IrO}_x$  cocatalysts. A built-in electric field can be generated adjacent to the PCTI heterojunction because of the mismatch in Fermi energy levels in the case of Z-scheme charge transfer within PCTI, leading to the accumulation of electrons and holes at the interfaces, which can cause band bending. The efficient separation and transportation of photogenerated electrons and holes can occur due to interfacial band bending and the internal electric field, which can retain a strong reduction and oxidation ability, where the electrons are in the conduction band of  $\text{g-C}_3\text{N}_4$  and holes are in the valence band of  $\text{TiO}_2$ . Moreover, the separately loaded Pt and  $\text{IrO}_x$  cocatalysts on the surfaces of the hollow spheres can enhance the surface reaction kinetics, where the cocatalysts can decrease surface recombination and reduce the activation energy for oxidation and reduction reactions in water splitting. If Pt and  $\text{IrO}_x$  cocatalysts are loaded in reverse, then this could reverse the reactions to retard water splitting. Hence, the loading of Pt on  $\text{g-C}_3\text{N}_4$  and  $\text{IrO}_x$  on the  $\text{TiO}_2$  surface could be essential to the enhancement of the photocatalytic  $\text{H}_2$  evolution activity.

Using only pure water without any sacrificial agent, Lu *et al.*<sup>11</sup> observed that the  $\text{N-NaTaO}_3@\text{Ta}_3\text{N}_5/\text{Rh}@\text{Cr}_2\text{O}_3$  photocatalyst demonstrated a remarkably high photocatalytic  $\text{H}_2$  evolution rate of  $25.12 \text{ mmol h}^{-1} \text{ g}^{-1}$  when used for overall photocatalytic water splitting, achieving a high apparent quantum yield of 9.32% at 350 nm. Furthermore, this photocatalyst exhibits a low charge transfer resistance. The ultrathin  $\text{Ta}_3\text{N}_5$  shell (thickness:  $\approx 5 \text{ nm}$ ) can enhance the lifetimes of the photogenerated charges, while the core-shell heterojunction with type-I band alignment can enable smooth charge flow from  $\text{N-NaTaO}_3$  to  $\text{Ta}_3\text{N}_5$ , especially with the aid of the shared communal Ta atoms at the interface, and the efficiency can be further enhanced by the *in situ* deposition of a core-shell structured  $\text{Rh}@\text{Cr}_2\text{O}_3$  cocatalyst.

Moreover, Ruan *et al.*<sup>115</sup> synthesized a twin S-scheme (step scheme) photocatalyst (TSP) comprised of 2D graphitic carbon nitride nanosheets, which were inserted between anatase  $\text{TiO}_2$  nanoparticles and H-doped rutile  $\text{TiO}_2$  nanorods. The TSP photocatalyst shows an exceptionally high  $\text{H}_2$  evolution rate of  $62.37 \text{ mmol h}^{-1} \text{ g}^{-1}$  and an impressive quantum yield of 45.9% at 365 nm, while maintaining a low charge transfer resistance. The enhanced photocatalytic performance of TSP is ascribed to optimal charge separation and transfer due to the unique twin S-scheme structure. Time-resolved characterization studies reveal that rapid electron trapping and suppressed recombination of photogenerated carriers contribute to TSP enhanced carrier separation and transfer efficiency. Moreover, the S-scheme pathway of charge transfer enables

electrons and holes to remain in the CB of  $\text{C}_3\text{N}_4$  and the VB of  $\text{TiO}_2$ , respectively, ensuring a robust redox ability.

In addition, Wang *et al.*<sup>55</sup> observed that the 15%  $\text{CoS}_2/\text{CdS-L}$  photocatalyst exhibited a much higher photocatalytic  $\text{H}_2$  evolution rate ( $19.22 \text{ mmol h}^{-1} \text{ g}^{-1}$ ) with a much higher apparent quantum yield of 22.6% at 420 nm, while it exhibited a low charge transfer resistance. Theoretical and experimental studies reveal that the fully exposed (002) surface of  $\text{CdS-L}$  (leaf-like  $\text{CdS}$ ) possesses superior activity for  $\text{H}_2$  evolution, while the electron flow between the  $\text{CdS-L}$  and  $\text{CoS}_2$  S-scheme heterojunction can be efficiently enhanced by the utilization of the interface heterojunction strategy, which can tune the potential barrier difference. Thus, the design of a photocatalyst with more exposed crystal planes on the basis of fabricating S-scheme heterojunctions can afford abundant active sites leading to enhanced reactivity.

Das *et al.*<sup>132</sup> discovered that the  $\text{Bi}_{20}\text{TiO}_{32}/\text{Bi}_4\text{Ti}_3\text{O}_{12}$  photocatalyst achieved a remarkable photocatalytic hydrogen evolution rate of  $1.89 \text{ mmol h}^{-1} \text{ g}^{-1}$ , accompanied by a low charge transfer resistance. This enhanced performance is attributed to a synergistic type-I bridged coupled Z-scheme electron migration process, facilitating prolonged excited state lifetimes, increased interfacial contact, rapid electron channeling, and improved optical absorption. Similarly, Shu *et al.*<sup>129</sup> found that the  $\text{TiO}_2\text{-Pt/FeHAp}$  (hydroxyapatite) photocatalyst reached a high hydrogen evolution rate of  $3.026 \text{ mmol h}^{-1} \text{ g}^{-1}$ , with an impressive apparent quantum yield of 4.72% at 370 nm, while also displaying a low charge transfer resistance. This indicates its efficiency for photocatalytic hydrogen production.  $\text{TiO}_2\text{-Pt/FeHAp}$  is obtained by the following steps: first,  $\text{TiO}_2\text{-Pt}$  was prepared by photoreduction of 1 wt% Pt clusters onto sub-15 nm  $\text{TiO}_2$  nanoparticles using 20 vol% methanol. Finally,  $\text{TiO}_2\text{-Pt/FeHAp}$  was obtained by an electrostatic assembly method. Thus, a built in electric field is formed in the  $\text{TiO}_2\text{-Pt/FeHAp}$  heterojunction, which drives the photogenerated electrons to flow from FeHAp on  $\text{TiO}_2$  nanoparticles to the Pt cocatalyst *via* a type-II path. The improved photocatalytic performance of  $\text{TiO}_2\text{-Pt/FeHAp}$  is ascribed to enhanced interfacial charge migration and separation due to the heterostructure, a large specific surface area stemming from sub-15 nm  $\text{TiO}_2$ , and facile  $\text{H}_2$  production kinetics associated with the Pt cocatalyst.

Moreover, Chen *et al.*<sup>128</sup> used Cu-In-Zn-S quantum dots (CIZS QDs),  $\text{MoS}_2$  and carbon dots (CDs) to construct the CIZS/ $\text{MoS}_2$ /CDs photocatalyst, which showed an impressive photocatalytic  $\text{H}_2$  evolution rate ( $3.706 \text{ mmol h}^{-1} \text{ g}^{-1}$ ) with a high apparent quantum yield of 34.23% at 450 nm and a low charge transfer resistance. The transient photovoltage measurements reveal that  $\text{MoS}_2$  has a negligible effect on the charge extraction rate, whereas the CDs significantly elevate the attenuation constant of the charge recombination process from 0.178 ms to 0.260 ms because of its electron sink effect. In addition, various studies along with an electrocatalytic investigation reveal that  $\text{MoS}_2$  plays a vital role in the hydrogen evolution reaction, whereas CIZS facilitates a light harvesting process and CDs enhance the charge sink effect.

Xu *et al.*<sup>123</sup> demonstrated that the Ru@Nb<sub>2</sub>O<sub>5</sub>/Nb<sub>2</sub>C photocatalyst achieved a remarkable hydrogen evolution rate of 10.11 mmol h<sup>-1</sup> g<sup>-1</sup>, coupled with an impressive apparent quantum yield of 41.25% at 313 nm, and maintained a low charge transfer resistance, highlighting its superior photocatalytic efficiency. The Ru@Nb<sub>2</sub>O<sub>5</sub>/Nb<sub>2</sub>C catalyst was synthesized through a one-pot hydrothermal method, during which the Nb<sub>2</sub>CT<sub>x</sub> MXene's reducibility was crucial to the oxidation process, actively producing H<sub>2</sub> from water splitting. This process reduces Ru<sup>3+</sup> to Ru nanoparticles *in situ*, which are then deposited on Nb<sub>2</sub>O<sub>5</sub> nanowires formed through the partial hydrothermal oxidation of Nb<sub>2</sub>CT<sub>x</sub>, illustrating the seamless integration of catalytic materials for enhanced photocatalytic hydrogen production. The valence state of the noble metal Ru/Pt obtained through this one-pot hydrothermal synthesis method was closer to the metallic state when compared to that of a traditional photodeposition method, and this decreased the charge transfer resistance by 82.5%.

Han *et al.*<sup>120</sup> demonstrated the RuNi/g-C<sub>3</sub>N<sub>4</sub> photocatalyst's outstanding performance, delivering a hydrogen evolution rate of 35.1 mmol h<sup>-1</sup> g<sup>-1</sup> and an apparent quantum yield of 1.27% at 380 nm. This catalyst also features a low charge transfer resistance, highlighting its efficacy for photocatalytic hydrogen generation. RuNi/g-C<sub>3</sub>N<sub>4</sub> with about 2.3% Ru content exhibits a higher photocatalytic H<sub>2</sub> evolution rate (>twice) than that of the Pt/g-C<sub>3</sub>N<sub>4</sub> photocatalyst with about 2.3% Pt content under simulated sunlight irradiation. It is composed of a 2D structured bimetallic RuNi alloy as a co-catalyst of g-C<sub>3</sub>N<sub>4</sub> nanosheets. The integration of RuNi with g-C<sub>3</sub>N<sub>4</sub> effectively reduces the transfer barrier for photogenerated electrons and provides ample reaction sites for photocatalytic H<sub>2</sub> generation. Additionally, the synergistic interaction between the Ni and Ru metals in the RuNi alloy enhances its performance as a co-catalyst, outperforming Pt in this role.

Besides, Yang *et al.*<sup>130</sup> found that the PtO@Ti<sub>3</sub>C<sub>2</sub>/TiO<sub>2</sub> photocatalyst achieved a notable hydrogen evolution rate of 2.54 mmol h<sup>-1</sup> g<sup>-1</sup> and an apparent quantum yield of 4.2% at 365 nm, showcasing its effectiveness at photocatalytic hydrogen production alongside featuring a low charge transfer resistance. The *in situ* grown TiO<sub>2</sub> nanosheets on Ti<sub>3</sub>C<sub>2</sub> MXene can enhance charge separation, while the deposited PtO nanodots with a higher oxidation state can decrease the hydrogen back oxidation reaction, and lead to improved photocatalytic H<sub>2</sub> generation. In the PtO@Ti<sub>3</sub>C<sub>2</sub>/TiO<sub>2</sub> photocatalyst, photogenerated electrons and holes on TiO<sub>2</sub> are transferred to PtO and Ti<sub>3</sub>C<sub>2</sub>, respectively. Simultaneously, photogenerated holes from PtO migrate to Ti<sub>3</sub>C<sub>2</sub>, while electrons remain on PtO, facilitating efficient charge separation. Then, the holes on Ti<sub>3</sub>C<sub>2</sub> can be consumed by reacting with the sacrificial agent methanol, while the electrons on PtO can react with protons to form H<sub>2</sub>, where the hydrogen back oxidation reaction can be suppressed.

Moreover, Zhang *et al.*<sup>117</sup> reported that the Ni@NiS<sub>x</sub>-CdS photocatalyst achieved an impressive hydrogen evolution rate of 78.7 mmol h<sup>-1</sup> g<sup>-1</sup> and an apparent quantum yield of

36.74% at 420 nm, demonstrating efficient charge separation with a low charge transfer resistance. The crystalline CdS nanorod was prepared by a solvothermal method, while the amorphous Ni@NiS<sub>x</sub> co-catalyst was prepared and assembled onto CdS through a partial reduction strategy. The Ni@NiS<sub>x</sub>-CdS photocatalyst exhibits a unique combination of amorphous and crystalline phases, resulting from the encapsulation of Ni clusters within an amorphous NiS<sub>x</sub> layer. Its gradient work function variation enhances the charge extraction efficiency, effectively suppressing carrier recombination. Furthermore, it regulates the S p-band center away from the Fermi level, optimizing the Gibbs free energy for atomic H<sub>2</sub> adsorption.

In addition, research by Ma *et al.*<sup>127</sup> shows that the CN/CNQD (quantum dot) photocatalyst delivers a notable hydrogen evolution rate of 4.0877 mmol h<sup>-1</sup> g<sup>-1</sup> and achieves an apparent quantum yield of 6.74% at 420 nm, showcasing effective charge separation, as indicated by its low charge transfer resistance. The CN/CNQD composite is constructed from a planar carbon nitride (CN) structure, which is seamlessly integrated with carbon nitride quantum dots (CNQDs). This unique composite is achieved through a novel phototriggered self-assembly strategy, where covalent bonds are forged at the atomic junctions between CN and CNQDs. This approach differs from conventional interfacial modification methods that rely solely on hydrogen bonding or van der Waals forces. Consequently, the CN/CNQD hybrid boasts superior interfacial interactions, enabling efficient carrier migration between its components. Additionally, the presence of a continuous  $\pi$ -conjugated structure within CN/CNQDs provides an ideal milieu for carrier transport. Furthermore, an inherent in-plane electric field acts as a driving force for the directional transfer of electrons and holes, thereby enhancing the separation and transportation of photogenerated carriers.

Besides, Long *et al.*<sup>131</sup> observed that the amorphous-NiB/TiO<sub>2</sub> photocatalyst exhibited a much higher photocatalytic H<sub>2</sub> evolution rate (2.334 mmol h<sup>-1</sup> g<sup>-1</sup>) with a much higher apparent quantum yield of 16.30%, while it exhibited a low charge transfer resistance. Amorphous-NiB/TiO<sub>2</sub> is composed of amorphous NiB nanodots (0.5 to 1 nm), which are anchored onto the TiO<sub>2</sub> surface through a novel light-induced route. Theoretical and experimental investigations reveal that the amorphization of NiB can induce reverse electron transfer from B to Ni, which contributes to an optimal electronic structure on Ni active sites and facilitates the hydrogen desorption process on the Ni active sites, leading to enhanced efficiency for H<sub>2</sub> generation.

Moreover, Zeng *et al.*<sup>118</sup> observed an exceptionally high photocatalytic H<sub>2</sub> evolution rate of 73.54 mmol h<sup>-1</sup> g<sup>-1</sup> for the NiSeS/ZnSe photocatalyst, coupled with a noteworthy apparent quantum yield of 57.68% at 400 nm, while maintaining a low charge transfer resistance. The NiSeS cocatalyst with dual Ni-Se vacancies is integrated with ZnSe to obtain NiSeS/ZnSe heterojunctions, where a local polarization electric field can be induced by the dual Ni-Se vacancies, greatly enhancing

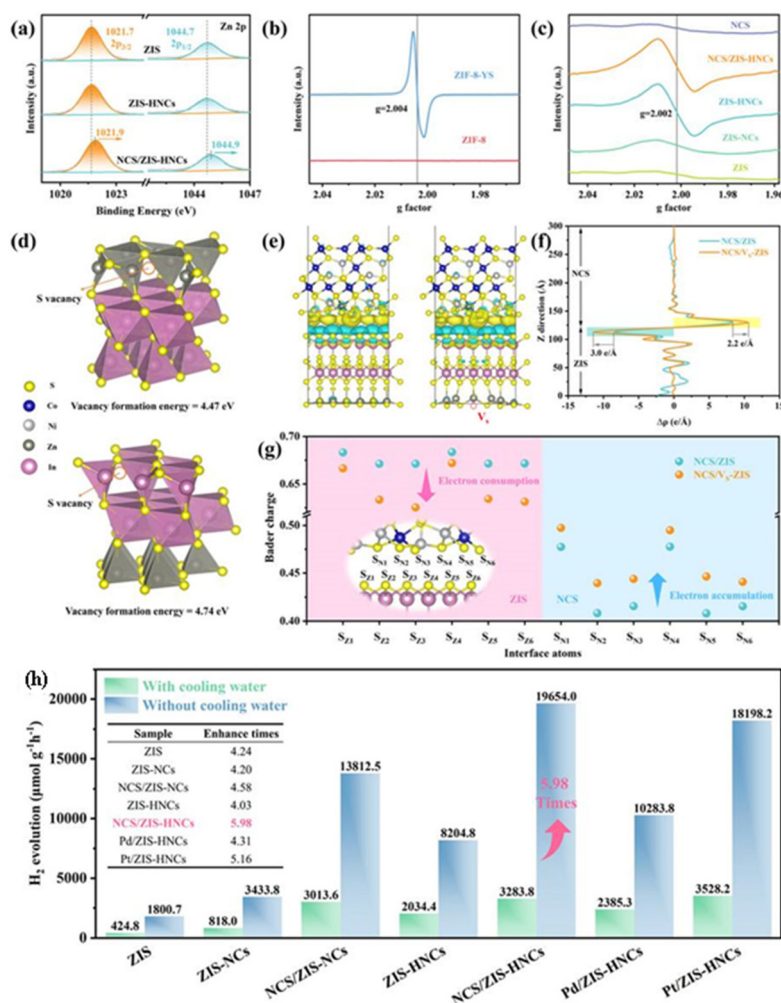
surface charge transfer and effectively reducing the reaction barrier within the photocatalytic process.

Tang *et al.*<sup>119</sup> precisely anchored Pt sub-nanoclusters (SNCs) onto the tandem CdS homojunction (CdS<sub>x</sub>-T) through an *in situ* photoreduction route. The CdS<sub>x</sub>-T photocatalyst, enriched with coordination-unsaturated S atoms, serves as an ideal anchor for Pt single nanoclusters (SNCs), promoting the formation of interfacial Pt-S bonds. These bonds enhance the directional transfer of photocarriers at the metal-semiconductor interface. Additionally, the CdS<sub>x</sub>-T homojunction, characterized by a tandem potential barrier, promotes the efficient separation of photocarriers through an internal electrostatic field. The size of Pt SNCs plays a pivotal role in the hydrogen ion (H<sup>+</sup>) adsorption capacity and electron transfer kinetics at the metal-semiconductor interface. Consequently, the Pt-CdS<sub>x</sub>-T photocatalyst demonstrates an impressive photocatalytic hydrogen evolution rate of 42.7 mmol h<sup>-1</sup> g<sup>-1</sup>, along with a remarkable apparent quantum yield of 43.6% at 470 nm. Furthermore, it exhibits a low charge transfer resistance, highlighting its superior photocatalytic performance.

The sluggish kinetics of photocatalytic H<sub>2</sub> evolution can be enhanced by efficient replenishment through the photothermal effect, where hot electrons can produce a photothermal effect by means of phonon-phonon scattering. Nevertheless, the hot electrons have short lifetimes due to fast relaxation kinetics, leading to a low energy utilization efficiency. The fabrication of hollow-structured ZnIn<sub>2</sub>S<sub>4</sub> can enhance the efficiency of interfacial electron transfer at the NiCo<sub>2</sub>S<sub>4</sub>/ZnIn<sub>2</sub>S<sub>4</sub> photocatalyst junction, extending the hot electron lifetimes for efficient solar energy utilization and significantly improving the hydrogen evolution rate. Guo *et al.*<sup>17</sup> observed that the NiCo<sub>2</sub>S<sub>4</sub>/ZnIn<sub>2</sub>S<sub>4</sub> hollow nanocage (NCS/ZIS-HNC) photocatalyst demonstrated a remarkable hydrogen evolution rate of 19.654 mmol g<sup>-1</sup> h<sup>-1</sup> and an impressive apparent quantum efficiency of 56.7% at 400 nm, when exposed to visible and near-infrared (NIR) light. Hollow-structured ZnIn<sub>2</sub>S<sub>4</sub> with surface S vacancies obtained by a soft template method can control the electron accumulation/consumption ability of interfacial atoms to enhance the interfacial interaction with NiCo<sub>2</sub>S<sub>4</sub>. Hollow-structured ZnIn<sub>2</sub>S<sub>4</sub> enhances the number of photogenerated electrons through multiple light scattering, facilitating their efficient transfer at the interface with NiCo<sub>2</sub>S<sub>4</sub>. This transfer accumulates photogenerated electrons on the surface of NiCo<sub>2</sub>S<sub>4</sub>, which are further excited to hot electrons by the localized surface plasmon resonance (LSPR) effect. This process continuously generates hot electrons by intrinsic excitation, prolonging their lifetimes for efficient solar energy utilization. Fig. 10(h) depicts the photocatalytic hydrogen evolution rates of NiCo<sub>2</sub>S<sub>4</sub>/ZnIn<sub>2</sub>S<sub>4</sub> in comparison with its counterparts in the presence and absence of cooling water at an indoor temperature of 25 °C. NiCo<sub>2</sub>S<sub>4</sub>/ZnIn<sub>2</sub>S<sub>4</sub> in the absence of cooling water exhibits a H<sub>2</sub> evolution rate of 19.654 mmol g<sup>-1</sup> h<sup>-1</sup>, which is 5.98 times higher than that in the presence of cooling water (3.2838 mmol g<sup>-1</sup> h<sup>-1</sup>). This further demonstrates that the photothermal effect plays a

pivotal role in the substantial enhancement of the H<sub>2</sub> evolution rate for the NiCo<sub>2</sub>S<sub>4</sub>/ZnIn<sub>2</sub>S<sub>4</sub> photocatalyst. Moreover, compared to ZIS-HNCs, lower charge transfer resistance and higher transient photocurrent density were obtained for NCS/ZIS-HNCs, suggesting effective migration of photogenerated electrons at the interface. The high resolution Zn 2p XPS spectrum obtained for NCS/ZIS-HNCs (Fig. 10(a)) shows higher binding energies compared to those of ZIS-HNCs, indicating the existence of strong interfacial interactions between ZIS-HNCs and NCS. ZIS-HNCs were derived from ZIF-8-YS, and therefore ZIF-8-YS was characterized. Fig. 10(b) depicts the ESR spectrum of ZIF-8-YS in comparison with that of ZIF-8. A noticeable ESR response at a *g*-value of 2.004 is observed for ZIF-8-YS, which can be ascribed to the generation of abundant defects during the hollowing process. In this process, the organic frameworks coordinated with Zn<sup>2+</sup> can be destroyed by free H<sup>+</sup> ions released from tannic acid to generate abundant defects, leading to a disordered crystal structure for ZIF-8-YS. Moreover, even after releasing H<sup>+</sup> ions, tannic acid possesses numerous electrons, enabling it to form a new coordination bond with Zn<sup>2+</sup>. This, in turn, partially mitigates the subsequent sulphuration reaction of thioacetamide. When compared to ZIS, ZIS-NCs, and NCS, both NCS/ZIS-HNCs and ZIS-HNCs exhibit a pronounced ESR response at a *g*-value of 2.002 (Fig. 10(c)), which is attributed to sulfur vacancies resulting from unpaired electrons with low spin states in the S 3p orbital, suggesting the formation of sulfur vacancies around Zn atoms during the preparation of hollow-structured ZIS. In addition, DFT calculations were conducted to investigate the role of sulfur vacancies in enhancing strong interfacial interactions. Fig. 10(d) depicts the ZIS crystal structure with the S vacancy formation energy at different sites, which reveals that the S vacancies around Zn atoms have a lower formation energy of 4.47 eV, compared to that around In atoms (4.74 eV). Fig. 10(e) compares the electron density differences of NCS/VIS and NCS/ZIS. Both NCS/VIS and NCS/ZIS afford electron consumption on ZIS (cyan region) and electron accumulation on NCS (yellow region). Notably, NCS/VIS with S vacancies shows higher interfacial electron accumulation/consumption as shown in Fig. 10(f), suggesting a stronger interfacial interaction caused by S vacancies. Moreover, the Bader charge of interfacial S atoms was calculated to understand this phenomenon. As shown in Fig. 10(g), a decreasing trend is observed for the Bader charges of all interfacial S atoms (labeled as S<sub>Z1</sub>-S<sub>Z6</sub>) at the ZIS part, whereas an increasing trend is observed for all the S atoms (labeled as S<sub>N1</sub>-S<sub>N6</sub>) at the NCS part. This suggests that the interfacial S atoms with S vacancies can improve the electron accumulation/consumption ability and enhance efficient transfer of more photogenerated carriers. Subsequently, the transferred electrons are excited to hot electrons through the LSPR effect to constantly enable the intrinsic excitation of NCS, promoting the photothermal effect and ultimately enhancing photocatalytic H<sub>2</sub> evolution.

Shi *et al.*<sup>121</sup> prepared ZrC@ZnIn<sub>2</sub>S<sub>4</sub> through a facile water bath method. The formation of a Schottky junction in the



**Fig. 10** (a) High resolution XPS spectra of Zn 2p obtained from NCS/ZIS-HNCs in comparison with its counterparts; (b) ESR spectrum of ZIF-8-YS in comparison with that of ZIF-8; (c) ESR spectrum of NCS/ZIS-HNCs in comparison with those of its counterparts; (d) ZIS crystal structure depicting the S vacancy formation energy at different sites; (e) electron density difference of NCS/Vs-ZIS (right) in comparison with that of NCS/ZIS (left); (f) planar-averaged electron density difference  $\Delta\rho(z)$  of NCS/Vs-ZIS in comparison with that of NCS/ZIS; (g) Bader charge of interfacial atoms of NCS/Vs-ZIS in comparison with that of NCS/ZIS; (h) photocatalytic hydrogen evolution rates of NCS/ZIS-HNCs in comparison with those of its counterparts in the presence and absence of cooling water (reproduced with permission from ref. 17. Copyright 2023, The Royal Society of Chemistry).

ZrC@ZnIn<sub>2</sub>S<sub>4</sub> photocatalyst can kinetically accelerate photo-generated carrier transfer and separation, enhancing photocatalytic water splitting for H<sub>2</sub> formation. Moreover, the photothermal effect induced by the LSPR effect (LSPR: local surface plasmon resonance) of metallic ZrC particles can thermodynamically decrease the activation energy of the reaction and enable the forward reaction in the ZrC@ZnIn<sub>2</sub>S<sub>4</sub> heterostructure system. Finally, a desirable H<sub>2</sub> evolution rate (32.87 mmol h<sup>-1</sup> g<sup>-1</sup>) with a much higher photothermal conversion efficiency of 43.54% using NIR irradiation ( $\lambda > 800$  nm) was obtained. Alternatively, Guo *et al.*<sup>122</sup> observed that the Au@MoS<sub>2</sub>-ZnO piezo-photocatalyst exhibited a much higher piezo-enhanced photocatalytic H<sub>2</sub> evolution rate (13.09 mmol h<sup>-1</sup> g<sup>-1</sup>) under light and ultrasonic waves, but it exhibited a low charge transfer resistance. It was obtained by a hydrothermal method. DFT

calculations and photoelectric characterization reveal that the ultrasonic field's spatial compression and cavitation effect can enhance the dynamics of mass transfer. Specifically, the cavitation effect enhances the chemical bond breakage-reformation process, thereby accelerating the surface redox reaction, and the ultrasonic field can tune the built-in electric field, thereby facilitating transfer of the photoinduced carriers.

## 5.2. Nano-photocatalyst with elemental doping/metal free/carbon based materials/MOFs

A metal-free photocatalyst can be prepared by integrating melon with highly reduced graphene oxide, which could enhance the visible light utilization, narrow the bandgap, enable charge redistribution at the interface, and improve the performance for photoelectrochemical hydrogen evolution.



Ashraf *et al.*<sup>54</sup> demonstrated that photoelectrochemical hydrogen evolution could be enhanced by bandgap engineering of melon with highly reduced graphene oxide (HRG), where melon is the uncondensed form of polymeric carbon nitrides. The 1% HRG@melon nanocomposite exhibits a higher photocurrent density ( $71 \mu\text{A cm}^{-2}$ ) than that of melon ( $24 \mu\text{A cm}^{-2}$ ) under alkaline conditions, while the photocurrent density is further increased to  $630 \mu\text{A cm}^{-2}$  with the addition of triethanolamine as a hole scavenger, and it exhibits a low charge transfer resistance. The addition of HRG can reduce the bandgap of melon from 2.6 eV to 2.4 eV, enhance visible light utilization and charge redistribution at the interface between melon and HRG, and finally lead to improved performance for photoelectrochemical hydrogen evolution.

The decoration of  $\text{TiO}_2$  with Ru/In dual-single atoms can produce oxygen vacancies and elevate the  $\text{Ti}^{3+}$  concentration, accelerating charge separation and ultimately promoting photocatalytic  $\text{H}_2$  evolution. Peng *et al.*<sup>8</sup> observed that Ru-In SA/ $\text{TiO}_2$  ( $\text{TiO}_2$  doped with Ru/In single atoms) could alter the separation of photogenerated carriers during photocatalytic  $\text{H}_2$  evolution using pure water. Ru-In SA/ $\text{TiO}_2$  exhibits a much higher photocatalytic  $\text{H}_2$  evolution rate of  $174.1 \mu\text{mol h}^{-1}$  and a lower charge transfer resistance than those of In SA/ $\text{TiO}_2$ , Ru SA/ $\text{TiO}_2$ , and  $\text{TiO}_2$ . The decoration of Ru/In dual-single atoms on  $\text{TiO}_2$  increases the oxygen vacancies and elevates  $\text{Ti}^{3+}$ , leading to accelerated charge separation. Femtosecond transient absorption spectroscopy reveals that the distribution separation of photogenerated electrons and holes of  $\text{TiO}_2$  are accelerated by doping with Ru and In single atoms, respectively.

Besides, a carbon doped and carbon dot loaded graphitic carbon nitride (CDs/CCN) photocatalyst was obtained by Yuan *et al.* using a direct thermal polymerization route.<sup>125</sup> CDs/CCN shows a boosted photocatalytic  $\text{H}_2$  evolution rate ( $8.682 \text{ mmol h}^{-1} \text{ g}^{-1}$ ) with a high apparent quantum yield of 1.41% at 420 nm. Transient absorption spectroscopy investigations reveal that C-doping of g- $\text{C}_3\text{N}_4$  induced shallow trap states that could confine photoinduced electrons to alleviate direct recombination and deep trapping of photogenerated carriers. The integration of CDs with CCN can create a fast hole transfer path and provide additional reaction sites for the oxidation of triethanolamine (serving as a sacrificial agent) *via* extracted holes, giving rise to efficient spatial separation of electron-hole pairs.

Moreover, Luo *et al.*<sup>124</sup> found that doping  $\text{Ni}_5\text{P}_4$  with Zn could lower its capacitance and enhance the release of photoexcited electrons. Besides, Zn-doped  $\text{Ni}_5\text{P}_4$  exhibits a low overpotential and high specific surface area, which can provide abundant active sites and facilitate the kinetics for  $\text{H}_2$  production. The constructed Zn- $\text{Ni}_5\text{P}_4$ /CdS photocatalyst demonstrates an exceptionally high rate of photocatalytic  $\text{H}_2$  evolution ( $8.969 \text{ mmol h}^{-1} \text{ g}^{-1}$ ), which is 2.4 times higher than that of  $\text{Ni}_5\text{P}_4$ /CdS ( $3.668 \text{ mmol h}^{-1} \text{ g}^{-1}$ ). Specifically, Zn doping decreases the capacitance of  $\text{Ni}_5\text{P}_4$  from 680 to 540  $\text{mF g}^{-1}$  while enhancing electron transport, which in turn mitigates the negative effects of capacitance and promotes the release of photoexcited electrons.

Zhang *et al.*<sup>116</sup> precisely anchored Ni-doped cobalt phosphide nanoparticles ( $\sim 4 \text{ nm}$ ) on phosphatized carbon nitride nanosheets through an absorption-phosphidation strategy to obtain the PCN-Ni-CoP photocatalyst. It achieves a remarkably high photocatalytic  $\text{H}_2$  evolution rate of  $0.143 \text{ mmol h}^{-1} \text{ g}^{-1}$  and impressive apparent quantum yield of 2.7% at 380 nm through overall water splitting, and maintains a low charge transfer resistance. PCN-Ni-CoP is composed of Ni-doped cobalt phosphide nanoparticles ( $\sim 4 \text{ nm}$ ), which are precisely anchored on phosphatized carbon nitride nanosheets through an absorption-phosphidation strategy. In the PCN-Ni-CoP photocatalyst, the CoP nanoparticles with atomic-scale doped Ni element can effectively facilitate the surface catalytic reaction kinetics and act as the active center for the redox reaction to generate  $\text{H}_2$ . The PCN nanosheets with higher electrical conductivity and a lower band gap energy can efficiently enhance photogenerated charge separation and transfer. Moreover, the Co-N coordination effect between Ni-CoP and PCN can also contribute to the acceleration of interfacial electron transfer.

In addition, Zhou *et al.*<sup>133</sup> utilized BiOCl (BOC) semiconductor and cellulose nanofibers@carbon nanotubes (CNF@CNT) to construct a 2D lamellar membrane, BOC-CNF@CNT, which showed a satisfactory photocatalytic  $\text{H}_2$  evolution rate ( $0.02542 \text{ mmol h}^{-1} \text{ g}^{-1}$ ) and a low charge transfer resistance. The CNTs effectively function as transfer channels for photogenerated carriers, optimizing the separation of electrons and holes while minimizing interfacial charge transfer resistance and mitigating carrier recombination. Consequently, this enhances the photocatalytic  $\text{H}_2$  production performance of BOC-CNF@CNT.

Moreover, Luo *et al.*<sup>126</sup> synthesized the  $\text{NH}_2$ -MIL-125 photocatalyst using HCl vapor as a protonating reagent through natural volatilization of hydrochloric acid. The  $-\text{NH}_3^+\text{Cl}^-$  group (a moiety with charge and ample hydrogen) is introduced into  $\text{NH}_2$ -MIL-125, which leads to efficient charge carrier separation and a high photocatalytic  $\text{H}_2$  evolution rate ( $6.7628 \text{ mmol h}^{-1} \text{ g}^{-1}$ ). Characterization studies including PL, FT-IR, and *in situ* DRIFTS demonstrate that the protonation effect not only enhances the electron migration rates but also effectively accelerates the separation of photogenerated electron-hole pairs during the photocatalytic reaction.

## 6. Photocatalytic HX splitting

### 6.1. Lead free perovskite based photocatalysts

Photocatalytic  $\text{H}_2$  production using hydrohalic acid (HX) is a promising strategy, where value-added chemicals (such as  $\text{X}_2/\text{X}_3^-$ ) can also be generated along with clean, renewable  $\text{H}_2$  fuel. However, developing environmentally friendly lead-free photocatalysts with strong acid resistance and a visible light response is a challenging task. The *in situ* formation of a bismuth-based perovskite heterojunction has recently been proved as an effective way to attain better photocatalytic HI splitting performance. Tang *et al.*<sup>5</sup> successfully fabricated a lead-free bismuth-based hybrid perovskite by utilizing a

solvent engineering technique, which enabled the *in situ* growth of heterojunctions at the interface of MA<sub>3</sub>Bi<sub>2</sub>I<sub>9</sub> (methylammonium bismuth iodide) and DMA<sub>3</sub>BiI<sub>6</sub> (tri(dimethylammonium)hexa-iodobismuthate). The well-matched type-II heterostructure of BBP-5 gives rise to a much higher H<sub>2</sub> evolution rate of 198.2 μmol h<sup>-1</sup> g<sup>-1</sup> (Table 10; Fig. 11(b)) and apparent quantum efficiency, compared to that of its counterparts, without using noble metal co-catalysts under the illumination of 100 mW cm<sup>-2</sup> visible light ( $\lambda \geq 420$  nm) in an aqueous hydrogen iodide solution. Besides, there are no side products from the photocatalytic HI splitting (Fig. 11(c)). Moreover, MA<sub>3</sub>Bi<sub>2</sub>I<sub>9</sub>/DMA<sub>3</sub>BiI<sub>6</sub> (BBP-5) can maintain a high photocatalytic H<sub>2</sub> evolution activity for 100 h (Fig. 11(d)). Photoelectrochemical characterization studies illustrate that BBP-5 possesses a lower charge transfer resistance ( $R_{ct}$ : 1.15 kΩ cm<sup>2</sup>) and photoluminescence (PL) peak intensity, compared to that of its counterparts (other Bi-based perovskites), suggesting suppression of photogenerated charge recombination and a facilitated charge diffusion rate at the heterojunction interface. BBP-5 exhibits higher PL lifetimes (short-lived lifetime,  $\tau_1$ : 14.63 ns; long-lived lifetime,  $\tau_2$ : 115.67 ns; average lifetimes,  $\tau$ : 37.98 ns) compared to that of its counterparts, suggesting an extended exciton diffusion lifetime because of the reduction of nonradiative recombination in the MA<sub>3</sub>Bi<sub>2</sub>I<sub>9</sub>/DMA<sub>3</sub>BiI<sub>6</sub> heterojunctions. This reduction can facilitate the H<sub>2</sub> evolution activity due to the efficient charge separation of BBP-5 (Fig. 11(a)).

The fabrication of a mixed-halide perovskite with a bandgap funnel structure is another viable strategy to improve the photocatalytic H<sub>2</sub> evolution performance. Tang *et al.*<sup>140</sup> prepared an environmentally friendly lead-free MA<sub>3</sub>Bi<sub>2</sub>Cl<sub>9-x</sub>I<sub>x</sub> photocatalyst with a bandgap funnel structure. Because of the gradient neutral formation energy within the mixed halide perovskite, the content of iodide element diminishes from the surface to the interior across the MA<sub>3</sub>Bi<sub>2</sub>Cl<sub>9-x</sub>I<sub>x</sub> perovskite,

resulting in the formation of a graded bandgap funnel structure, which can facilitate photoinduced charge transfer from the interior to the surface. Consequently, an enhanced photocatalytic H<sub>2</sub> evolution rate of  $341 \pm 61.7$  μmol h<sup>-1</sup> with a Pt co-catalyst was achieved under the illumination of visible light, which was higher when compared to that of MA<sub>3</sub>Bi<sub>2</sub>Cl<sub>9</sub>. The content of iodide element diminishes from the surface to the interior of the MA<sub>3</sub>Bi<sub>2</sub>Cl<sub>9-x</sub>I<sub>x</sub> perovskite because of the gradient neutral formation energy within the mixed halide perovskite. Thus, a graded bandgap funnel structure has been formed due to the aligned energy levels of the iodide/chloride-mixed perovskite (MA<sub>3</sub>Bi<sub>2</sub>Cl<sub>9-x</sub>I<sub>x</sub>), which facilitate photoinduced charge transfer to the surface from the interior, leading to an enhanced photocatalytic redox reaction.

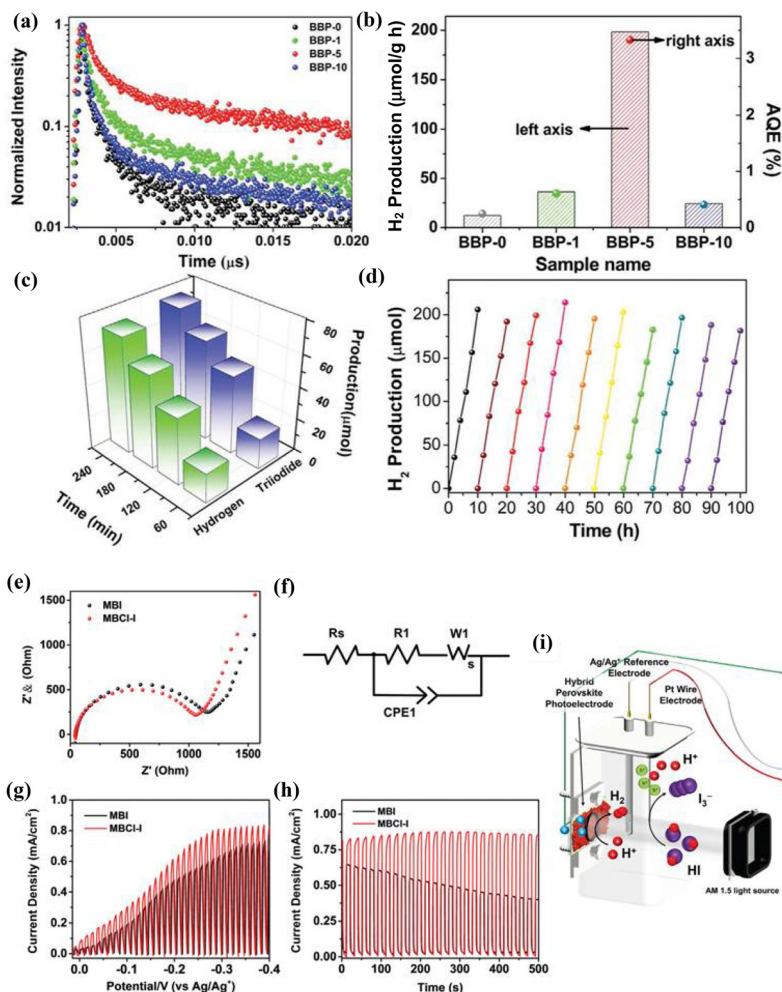
Besides, understanding the dynamic interfacial interactions of bandgap funneling in Bi-based mixed-halide perovskites in the photoelectrochemical system are crucial. Hence, Tang *et al.*<sup>9</sup> employed photophysical and photoelectrochemical characterization studies to reveal that MA<sub>3</sub>Bi<sub>2</sub>Cl<sub>9-y</sub>I<sub>y</sub> with a bandgap funnel structure exhibited a higher photocurrent density (Fig. 11(g and h)) and lower charge transfer resistance (Fig. 11(e)) compared to those of MA<sub>3</sub>Bi<sub>2</sub>I<sub>9</sub>. The photophysical and (photo-)electrochemical phenomena of solid-solid and solid-liquid interfaces for MA<sub>3</sub>Bi<sub>2</sub>Cl<sub>9-y</sub>I<sub>y</sub> and MA<sub>3</sub>Bi<sub>2</sub>I<sub>9</sub> are confirmed using electrochemical voltammetric and temperature dependent transient photoluminescence techniques. The stronger electronic coupling of MA<sub>3</sub>Bi<sub>2</sub>Cl<sub>9-y</sub>I<sub>y</sub>, stemming from the facilitated overlap of electronic wavefunctions, results in a higher diffusion coefficient and electron-transfer rate, ultimately enhancing photoelectrochemical H<sub>2</sub> production.

Zhou *et al.*<sup>18</sup> integrated Pt single atoms into Cs<sub>2</sub>SnI<sub>6</sub> to improve the photocatalytic activity. The PtSA/Cs<sub>2</sub>SnI<sub>6</sub> catalyst prepared exhibits enhanced tolerance towards HI aqueous solutions and achieves outstanding photocatalytic H<sub>2</sub> production, which is about 176.5 times higher when compared to

**Table 10** Photocatalytic H<sub>2</sub> evolution rate and apparent quantum yield (AQY) of various kinds of Pb-free/Pb-based perovskite photocatalysts with/without cocatalysts using HX and H<sub>3</sub>PO<sub>2</sub> as a reaction medium

| Pb-free/Pb-based perovskite photocatalysts with/without cocatalysts               | Light source                                         | Reaction solution                          | H <sub>2</sub> evolution rate (mmol h <sup>-1</sup> g <sup>-1</sup> ) | AQY             | Ref. |
|-----------------------------------------------------------------------------------|------------------------------------------------------|--------------------------------------------|-----------------------------------------------------------------------|-----------------|------|
| MAPbI <sub>3</sub> /MoS <sub>2</sub>                                              | 280 W Xe lamp ( $\lambda \geq 420$ nm)               | HI and H <sub>3</sub> PO <sub>2</sub>      | 29.389                                                                | 22.1% @ 500 nm  | 134  |
| MoSe <sub>2</sub> /MAPbBr <sub>3-x</sub> I <sub>x</sub>                           | 300 W Xe lamp ( $\lambda > 420$ nm)                  | HBr, HI and H <sub>3</sub> PO <sub>2</sub> | 22.935                                                                | 37.67% @ 530 nm | 10   |
| PtSA/FAPbBr <sub>3-x</sub> I <sub>x</sub>                                         | 300 W Xe lamp (AM 1.5G)                              | HBr, HI and H <sub>3</sub> PO <sub>2</sub> | 6.826                                                                 | 33.4% @ 530 nm  | 32   |
| MAPbI <sub>3</sub> /BP                                                            | 300 W Xe lamp ( $\lambda \geq 420$ nm)               | HI and H <sub>3</sub> PO <sub>2</sub>      | 3.742                                                                 | 23.2% @ 420 nm  | 135  |
| Ni <sub>3</sub> C/MAPbI <sub>3</sub>                                              | 300 W Xe lamp ( $\lambda \geq 420$ nm)               | HI and H <sub>3</sub> PO <sub>2</sub>      | 2.362                                                                 | 16.6% @ 420 nm  | 136  |
| MAPbI <sub>3</sub> /CoP                                                           | 150 W Xe lamp ( $\lambda \geq 420$ nm)               | HI and H <sub>3</sub> PO <sub>2</sub>      | 2.0875                                                                | NA              | 137  |
| MAPbI <sub>3</sub> /MoS <sub>2</sub>                                              | 9 LED lamp (10 W; 380 nm $\leq \lambda \leq 780$ nm) | HI and H <sub>3</sub> PO <sub>2</sub>      | ~2.061                                                                | NA              | 138  |
| aMAPb(I <sub>1-x</sub> Br <sub>x</sub> ) <sub>3</sub>                             | 300 W Xe lamp ( $\lambda \geq 420$ nm)               | HI, HBr and H <sub>3</sub> PO <sub>2</sub> | 1.471                                                                 | NA              | 2    |
| CH <sub>3</sub> NH <sub>3</sub> PbI <sub>3</sub> /rGO                             | 300 W Xe lamp ( $\lambda > 420$ nm)                  | HI and H <sub>3</sub> PO <sub>2</sub>      | 0.9389                                                                | 1.4% @ 450 nm   | 139  |
| MA <sub>3</sub> Bi <sub>2</sub> Cl <sub>9-x</sub> I <sub>x</sub> /Pt              | 300 W Xe lamp ( $\lambda \geq 420$ nm)               | HCl, HI and H <sub>3</sub> PO <sub>2</sub> | ~0.682                                                                | 1.25% @ 435 nm  | 140  |
| PtSA/Cs <sub>2</sub> SnI <sub>6</sub>                                             | 300 W Xe lamp ( $\lambda \geq 420$ nm)               | HI and H <sub>3</sub> PO <sub>2</sub>      | 0.43                                                                  | NA              | 18   |
| Cs <sub>2</sub> AgBiBr <sub>6</sub> /N-C                                          | 300 W Xe lamp ( $\lambda \geq 420$ nm)               | HBr and H <sub>3</sub> PO <sub>2</sub>     | 0.38                                                                  | 0.59% @ 420 nm  | 141  |
| MA <sub>3</sub> Bi <sub>2</sub> I <sub>9</sub> /DMA <sub>3</sub> BiI <sub>6</sub> | 300 W Xe lamp ( $\lambda \geq 420$ nm)               | HI and H <sub>3</sub> PO <sub>2</sub>      | 0.1982                                                                | >3%             | 5    |
| MoS <sub>2</sub> /Cs <sub>2</sub> AgBiBr <sub>6</sub>                             | 300 W Xe lamp ( $\lambda \geq 420$ nm)               | HBr and H <sub>3</sub> PO <sub>2</sub>     | 0.0875                                                                | 0.20% @ 450 nm  | 142  |

Ref.: reference; NA: not applicable; AQY: apparent quantum yield; SA: single atom; aMAPb(I<sub>1-x</sub>Br<sub>x</sub>)<sub>3</sub>: MAPb(I<sub>1-x</sub>Br<sub>x</sub>)<sub>3</sub> with  $x = 0.10$ ; BP: black phosphorus; rGO: reduced graphene oxide.

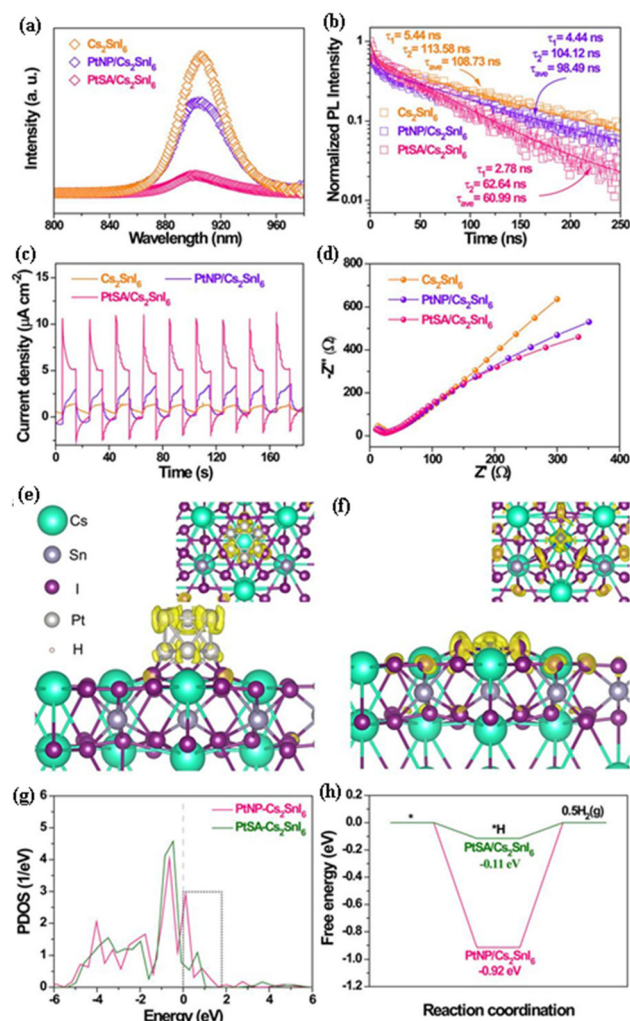


**Fig. 11** (a) Time-resolved photoluminescence (TRPL) spectrum of BBP-5 in comparison with those of its counterparts obtained at 650 nm; (b) photocatalytic  $\text{H}_2$  production along with the apparent quantum efficiency (AQE) of BBP-5 in comparison with those of its counterparts in HI saturated solution; (c) comparison of the amount of  $\text{H}_2$  (green) and  $\text{I}_3^-$  (purple) generated over time using the BBP-5 photocatalyst; (d) stable photocatalytic  $\text{H}_2$  evolution over 100 h using the BBP-5 photocatalyst in HI saturated solution, where  $\text{H}_3\text{PO}_2$  was added to the HI solution (reproduced with permission from ref. 5. Copyright 2020, Wiley-VCH GmbH). (e) Nyquist plot of MBI-I in comparison with that of MBI recorded at 0.14 V versus  $\text{Ag}/\text{Ag}^+$  electrode in  $\text{CH}_2\text{Cl}_2$  solution containing 0.1 M tetrabutylammonium hexafluorophosphate ( $\text{TBAPF}_6$ ) under visible-light illumination, where (f) depicts the corresponding equivalent electrical circuit; (g) chopped photocurrent–voltage ( $J$ – $V$ ) curve of MBI-I in comparison with that of MBI measured under visible-light irradiation; (h) chopped photocurrent response of MBI-I in comparison with that of MBI recorded at  $-0.3$  V vs.  $\text{Ag}/\text{Ag}^+$  under visible-light irradiation; (i) schematic diagram illustrating the working principle of the photoelectrochemical HI splitting cell using MBI-I/FTO (or  $\text{MBI}/\text{FTO}$ ) as a photoelectrode (reproduced with permission from ref. 9. Copyright 2022, Wiley-VCH GmbH).

that of Pt nanoparticles supported on the  $\text{Cs}_2\text{SnI}_6$  perovskite. As a result,  $\text{PtSA}/\text{Cs}_2\text{SnI}_6$  exhibits a significantly higher photocatalytic activity for  $\text{H}_2$  production (turnover frequency of 70.6 per h per Pt), which is about 176.5 times higher when compared to that of Pt nanoparticles supported on the  $\text{Cs}_2\text{SnI}_6$  perovskite, while it exhibits a significantly higher cycling durability. The PL technique was used to determine the role of Pt single atoms on the kinetics of charge carrier transfer and recombination on  $\text{Cs}_2\text{SnI}_6$ . Fig. 12(a) depicts the steady-state photoluminescence spectrum of the  $\text{PtSA}/\text{Cs}_2\text{SnI}_6$  photocatalyst in comparison with  $\text{Cs}_2\text{SnI}_6$  and  $\text{PtNP}/\text{Cs}_2\text{SnI}_6$  (NP: nanoparticles). Characterization studies including PL, TRPL, photocurrent, and EIS were conducted to investigate the intrinsic

properties of  $\text{PtSA}/\text{Cs}_2\text{SnI}_6$ . The results indicated that  $\text{PtSA}/\text{Cs}_2\text{SnI}_6$  exhibited a lower PL intensity, faster average decay lifetime, higher photocurrent density, and reduced electrochemical impedance, demonstrating that the evenly distributed atomic Pt species could significantly promote the separation and transfer of charge carriers. DFT calculations illustrated maps of the charge density differences before and after the photoexcitation of  $\text{PtSA}/\text{Cs}_2\text{SnI}_6$  (Fig. 12(f)) in comparison with  $\text{PtNP}/\text{Cs}_2\text{SnI}_6$  (Fig. 12(e)). The added photogenerated electrons tend to be distributed over the entire Pt nanoparticles in  $\text{PtNP}/\text{Cs}_2\text{SnI}_6$  (Fig. 12(e)), which can decrease the electron density per Pt atom in Pt nanoparticles, whereas the electrons in  $\text{PtSA}/\text{Cs}_2\text{SnI}_6$  are only located between the Pt single atoms





**Fig. 12** (a) Steady-state photoluminescence spectrum, (b) time-resolved transient photoluminescence decay, (c) photocurrent response spectrum, and (d) Nyquist plot of the PtSA/Cs<sub>2</sub>SnI<sub>6</sub> photocatalyst in comparison with those of Cs<sub>2</sub>SnI<sub>6</sub> and PtNP/Cs<sub>2</sub>SnI<sub>6</sub>; the maps depict the charge density differences before and after photoexcitation of (f) PtSA/Cs<sub>2</sub>SnI<sub>6</sub> and (e) PtNP/Cs<sub>2</sub>SnI<sub>6</sub>, where the isosurface of charge density is 0.001 e Å<sup>-3</sup>, while the insets depict the top view, and the additional electron distribution is indicated by the yellow region; the model adds an excess electron to describe the photogenerated electron; (g) the PDOS (5d states) of PtSA/Cs<sub>2</sub>SnI<sub>6</sub> in comparison with that of PtNP/Cs<sub>2</sub>SnI<sub>6</sub>, where the dashed line indicates the Fermi level; (h) the calculated energy profile for hydrogen generation on PtSA/Cs<sub>2</sub>SnI<sub>6</sub> in comparison with that of PtNP/Cs<sub>2</sub>SnI<sub>6</sub> (reproduced with permission from ref. 18. Copyright 2021, Springer Nature).

and three neighboring I atoms (Fig. 12(f)), suggesting the existence of high electron density on the Pt-I<sub>3</sub> sites, which can be ascribed to a strong metal-support interaction (SMSI) effect, leading to higher HER activity on PtSA/Cs<sub>2</sub>SnI<sub>6</sub>. The SMSI effect, which is conducive to photogenerated charge transfer between the photocatalyst and cocatalyst, was also supported by the calculated PDOS of PtSA/Cs<sub>2</sub>SnI<sub>6</sub> in comparison with PtNP/Cs<sub>2</sub>SnI<sub>6</sub> (Fig. 12(g)), where most 5d states of PtSA are observed below the Fermi level, suggesting an elec-

tron-rich state in PtSA. The integrated PDOS areas of uncaptured Pt 5d states above the Fermi level in PtSA are calculated to be 0.71, which is lower than that of PtNP (1.19), further suggesting an electron-rich state in PtSA. This implies that the PtSA species has a superior ability to capture electrons from Cs<sub>2</sub>SnI<sub>6</sub>, resulting in higher H<sub>2</sub> production of PtSA/Cs<sub>2</sub>SnI<sub>6</sub>. PtSA exhibits a significantly lower energy barrier (0.11 eV) than that of PtNP (0.92 eV; Fig. 12(h)), and the enhanced electronic properties of PtSA on Cs<sub>2</sub>SnI<sub>6</sub> lead to elevated photocatalytic H<sub>2</sub> production on PtSA/Cs<sub>2</sub>SnI<sub>6</sub>. Thus, the electronic property and unique coordination structure of Pt-I<sub>3</sub> sites as well as strong metal-support interactions collectively reduce the reaction energy barrier and contribute to an excellent photocatalytic H<sub>2</sub> evolution performance.

Moreover, Chen *et al.*<sup>56</sup> prepared lead-free Cs<sub>3</sub>Bi<sub>0.6</sub>Sb<sub>1.4</sub>I<sub>9</sub> perovskites with a high photocatalytic H<sub>2</sub> evolution rate and a low charge transfer resistance. Cs<sub>3</sub>Bi<sub>2x</sub>Sb<sub>2-2x</sub>I<sub>9</sub> doped with Sb can effectively diminish the contribution of Bi<sup>3+</sup> to the conduction band, which can suppress the impact of Bi vacancies on the band structure. Therefore, Cs<sub>3</sub>Bi<sub>2x</sub>Sb<sub>2-2x</sub>I<sub>9</sub> exhibits an optimal bandgap and higher optical absorption when compared to those of Cs<sub>3</sub>Bi<sub>2</sub>I<sub>9</sub> and Cs<sub>3</sub>Sb<sub>2</sub>I<sub>9</sub>, leading to enhanced performance for photocatalytic H<sub>2</sub> evolution. Besides, Zhang *et al.*<sup>142</sup> observed that the MoS<sub>2</sub>/Cs<sub>2</sub>AgBiBr<sub>6</sub> photocatalyst exhibited a high photocatalytic H<sub>2</sub> evolution rate through HX splitting, while it exhibited a low charge transfer resistance. The enhanced performance of MoS<sub>2</sub>/Cs<sub>2</sub>AgBiBr<sub>6</sub> can be attributed to the kinetics-facilitated heterostructure comprised of stable MoS<sub>2</sub> and Cs<sub>2</sub>AgBiBr<sub>6</sub>. In addition, Jiang *et al.*<sup>141</sup> found that the addition of N-C with a high surface area could inhibit the aggregation of Cs<sub>2</sub>AgBiBr<sub>6</sub> nanoparticles and afford further pathways for the migration of photoinduced carriers. Doping with nitrogen can enhance photoinduced electron transfer from Cs<sub>2</sub>AgBiBr<sub>6</sub> to N-C, leading to spatially separated electrons and holes with a prolonged electron life time, and boost the photocatalytic performance.

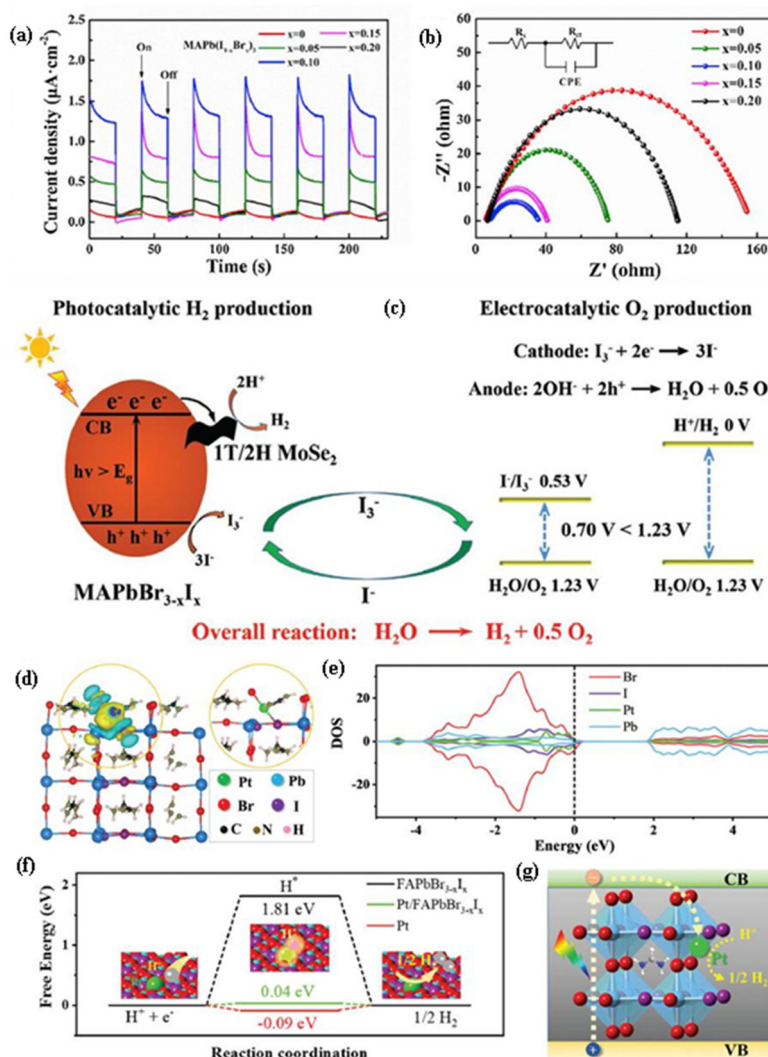
## 6.2. Lead based perovskites as photocatalysts

The introduction of Br into MAPbI<sub>3</sub> can enhance charge separation and transfer, which can enhance photocatalytic H<sub>2</sub> evolution. Zhao *et al.*<sup>2</sup> reported that MAPb(I<sub>1-x</sub>Br<sub>x</sub>)<sub>3</sub> ( $x = 0.10$ ) exhibited the highest H<sub>2</sub> evolution rate of 1471 μmol h<sup>-1</sup> g<sup>-1</sup> under visible light illumination ( $\lambda \geq 420$  nm) in HI/HBr aqueous solutions without using Pt cocatalyst; this was about 40 times larger than that of pure MAPbI<sub>3</sub>. And it also maintains high stability for photocatalytic H<sub>2</sub> production for 252 h (more than 60 runs) without changes to the perovskite structure. MAPb(I<sub>1-x</sub>Br<sub>x</sub>)<sub>3</sub> exhibits an obvious current increase under visible light irradiation, as indicated by the photocurrent responses and decreased semicircle in the Nyquist plot, suggesting superior separation and migration of charge carriers. The enhancement in the photocatalytic H<sub>2</sub> evolution of MAPb(I<sub>1-x</sub>Br<sub>x</sub>)<sub>3</sub> ( $x = 0.10$ ) is ascribed to the Br-incorporated perovskite with a lower Pb-H energy and lead at the defect site being exposed to amine cations.



Liu *et al.*<sup>10</sup> innovatively integrated photocatalytic HI splitting with electrocatalysis, creating an efficient closed-loop system for pure water splitting. This system is bifurcated into two parts: the first one involves the generation of H<sub>2</sub> using the MoSe<sub>2</sub>/MAPbBr<sub>3-x</sub>I<sub>x</sub> photocatalyst *via* photocatalytic HI splitting. Subsequently, the I<sub>3</sub><sup>-</sup> ions obtained from this photocatalytic process are subjected to electrolysis for electrocatalytic I<sub>3</sub><sup>-</sup> reduction, resulting in the production of O<sub>2</sub>. This integrated approach is schematically depicted in Fig. 13(c), highlighting the coherence and efficiency of the pure water splitting process. The 3.0 wt% MoSe<sub>2</sub>/MAPbBr<sub>3-x</sub>I<sub>x</sub> catalyst demonstrates an outstanding photocatalytic hydrogen production rate of 22.935 mmol h<sup>-1</sup> g<sup>-1</sup> under visible light illu-

mination, coupled with reduced charge transfer resistance, indicating its superior charge separation and transfer capabilities. The enhanced photocatalytic H<sub>2</sub> production activity is attributed to efficient charge separation, abundant active sites for H<sub>2</sub> generation, and a small energy barrier for HI splitting. Hence, the electrolyzer can operate efficiently at a low cell voltage of 1.22 V and a current density of 10 mA cm<sup>-2</sup>, facilitating the electrocatalytic reduction of I<sub>3</sub><sup>-</sup> and the production of O<sub>2</sub>. This voltage is notably lower than the theoretical requirement of 1.23 V for electrocatalytic pure water splitting. Simultaneously, this approach significantly reduces the dependence on H<sub>3</sub>PO<sub>2</sub> as a sacrificial agent in the photocatalytic HI splitting system.



**Fig. 13** (a) Chopped photocurrent responses and (b) Nyquist plots of MAPb(I<sub>1-x</sub>Br<sub>x</sub>)<sub>3</sub> photocatalysts in saturated HI/HBr solution, where  $x = 0, 0.05, 0.10, 0.15$ , and  $0.20$  (reproduced with permission from ref. 2. Copyright 2019, Elsevier B.V.). (c) Schematic diagram depicting the mechanism of pure water splitting using tandem photocatalytic H<sub>2</sub> generation and electrocatalytic O<sub>2</sub> generation with I<sub>3</sub><sup>-</sup>/I<sup>-</sup> as a redox shuttle mediator (reproduced with permission from ref. 10. Copyright 2023, Wiley-VCH GmbH). (d) The difference in the charge densities of Pt/FAPbBr<sub>3-x</sub>I<sub>x</sub>, where the inset depicts its local structure, while the isosurface value is set at 0.001 e Å<sup>-3</sup> (positive charge: yellow color; negative charge: cyan color); (e) projected density of states (PDOS) of the photocatalyst (Pt/FAPbBr<sub>3-x</sub>I<sub>x</sub>); (f) calculated free energy of H<sup>+</sup> adsorption for Pt/FAPbBr<sub>3-x</sub>I<sub>x</sub> in comparison with those of Pt and FAPbBr<sub>3-x</sub>I<sub>x</sub>, where the inset depicts the corresponding schematic mechanism; (g) schematic diagram illustrating the photocatalytic process for Pt/FAPbBr<sub>3-x</sub>I<sub>x</sub> (reproduced with permission from ref. 32. Copyright 2022, The Royal Society of Chemistry).

The integration of single Pt atoms into  $\text{FAPbBr}_{3-x}\text{I}_x$  can enhance the charge density around the Pt atom, afford optimal charge separation and transfer, and decrease the energy barrier for  $\text{H}_2$  production, leading to enhanced photocatalytic  $\text{H}_2$  production. Wu *et al.*<sup>32</sup> observed that  $\text{PtSA}/\text{FAPbBr}_{3-x}\text{I}_x$  (SA: single-atoms; FA =  $\text{CH}(\text{NH}_2)_2$ ) exhibited a much higher photocatalytic  $\text{H}_2$  evolution rate of  $682.6 \text{ mmol h}^{-1}$  under simulated sunlight irradiation with a high STH of 4.50%. Using a self-adsorption and photoreduction method, Pt single atoms are immobilized individually and coordinated with surface halide anions of  $\text{FAPbBr}_{3-x}\text{I}_x$  through replacing formamidinium groups. Higher photocurrent responses and the smaller semicircle in the Nyquist plot demonstrate efficient charge separation and transfer in  $\text{PtSA}/\text{FAPbBr}_{3-x}\text{I}_x$ . Time-/space-resolved photoelectron formation and migration characteristics were investigated with an advanced PL microscope at the single-particle level. This study reveals that photogenerated electrons from the perovskite can be effectively captured and trapped by individual Pt atoms, and then be preserved for the reduction of protons in the photocatalysis process. The difference in the charge densities of  $\text{Pt}/\text{FAPbBr}_{3-x}\text{I}_x$  shown in Fig. 13(d) reveals that the loading of single Pt atoms can induce increased charge density. The projected density of states (PDOS) (Fig. 13(e)) shows that orbital hybridization can occur between the coordinated Br/I atoms and the Pt introduced due to the loading of single Pt atoms, which can facilitate charge transfer to the Pt site from the perovskite, leading to enhancement of the proton reduction process. Fig. 13(f) depicts the calculated free energy of  $\text{H}^+$  adsorption ( $\Delta G_{\text{H}^+}$ ) for  $\text{PtSA}/\text{FAPbBr}_{3-x}\text{I}_x$  in comparison with those of Pt and  $\text{FAPbBr}_{3-x}\text{I}_x$ .  $\text{PtSA}/\text{FAPbBr}_{3-x}\text{I}_x$  exhibits the lowest adsorption energy of 0.04 eV (almost close to zero), when compared to those of  $\text{FAPbBr}_{3-x}\text{I}_x$  (1.18 eV) and bare Pt (−0.09 eV), suggesting the promising proton reduction ability of  $\text{PtSA}/\text{FAPbBr}_{3-x}\text{I}_x$ .

Li *et al.*<sup>135</sup> constructed the  $\text{MAPbI}_3/\text{BP}$  (BP: black phosphorus) type I heterojunction, where BP could function as an electron promoter to catch electrons from  $\text{MAPbI}_3$  and facilitate charge carrier separation, leading to an impressive photocatalytic  $\text{H}_2$  evolution rate ( $3.742 \text{ mmol h}^{-1} \text{ g}^{-1}$ ) and an apparent quantum yield of 23.2% at 420 nm. Besides, Zhao *et al.*<sup>136</sup> observed that the  $\text{Ni}_3\text{C}/\text{MAPbI}_3$  photocatalyst exhibited a high photocatalytic  $\text{H}_2$  evolution rate ( $2.362 \text{ mmol h}^{-1} \text{ g}^{-1}$ ) through HX splitting with a high apparent quantum yield of 16.6% @420 nm, while it exhibited a low charge transfer resistance. The enhanced performance of  $\text{Ni}_3\text{C}/\text{MAPbI}_3$  is ascribed to the enhanced capabilities of charge carrier transfer and separation along with the enormous reactive centers on the surface of  $\text{MAPbI}_3$  due to the decoration of  $\text{Ni}_3\text{C}$ . In addition, Cai *et al.*<sup>137</sup> observed that the  $\text{MAPbI}_3/\text{CoP}$  photocatalyst exhibited a high photocatalytic  $\text{H}_2$  evolution rate ( $2.0875 \text{ mmol h}^{-1} \text{ g}^{-1}$ ) through HX splitting, while it exhibited a low charge transfer resistance. The CoP nanoparticle co-catalyst can effectively extract photogenerated electrons from  $\text{MAPbI}_3$  to facilitate photoinduced charge separation, and enhance the interfacial catalytic reaction. Moreover, Wang *et al.*<sup>138</sup> observed that the

$\text{MAPbI}_3/\text{MoS}_2$  photocatalyst exhibited a high photocatalytic  $\text{H}_2$  evolution rate ( $\sim 2.061 \text{ mmol h}^{-1} \text{ g}^{-1}$ ) through HX splitting, while it exhibited a low charge transfer resistance. The integration of the  $\text{MoS}_2$  co-catalyst with  $\text{MAPbI}_3$  can efficiently transfer photogenerated electrons from  $\text{MAPbI}_3$  to  $\text{MoS}_2$ , causing efficient separation of photogenerated carriers, while photogenerated electrons on  $\text{MoS}_2$  can migrate to the surface active sites, leading to an enhancement in the reduction of protons to  $\text{H}_2$ . Besides, Guan *et al.*<sup>134</sup> observed that the  $\text{MAPbI}_3/\text{MoS}_2$  photocatalyst exhibited a significantly higher photocatalytic  $\text{H}_2$  evolution rate ( $29.389 \text{ mmol h}^{-1} \text{ g}^{-1}$ ) through HX splitting with a considerably higher solar HI splitting efficiency of 7.35% and much higher apparent quantum yield of 22.1% @500 nm, while it exhibited a low charge transfer resistance. The integration of  $\text{MoS}_2$  nanoflowers, with abundant active sites, an optimal band gap and unsaturated species, with  $\text{MAPbI}_3$  can efficiently facilitate charge separation and this can enhance the reduction process for  $\text{H}_2$  production. Wu *et al.*<sup>139</sup> announced that the  $\text{CH}_3\text{NH}_3\text{PbI}_3/\text{rGO}$  photocatalyst exhibited a high photocatalytic  $\text{H}_2$  evolution rate of  $0.9389 \text{ mmol h}^{-1} \text{ g}^{-1}$  through HX splitting, demonstrating remarkable stability over 200 h. The integration of rGO with  $\text{CH}_3\text{NH}_3\text{PbI}_3$  enables the efficient transfer of photogenerated electrons from  $\text{MAPbI}_3$  to rGO sites, significantly enhancing the reduction process for protons on rGO's surface. This results in improved  $\text{H}_2$  production rates, further highlighting the photocatalyst's performance and stability.

## 7. Summary and perspectives

Boosting the production of clean hydrogen fuel has been regarded as one of the most promising strategies to address the increasing energy crisis and environmental pollution. Electrocatalysis and photocatalysis are highly anticipated to achieve efficient  $\text{H}_2$  production by using catalysts with desirable properties. With the aim of prompting further research in this field, this review comprehensively articulated the recent advances in various nanostructured electrocatalysts and photocatalysts. In detail, electrocatalysts including oxide/hydroxide/layered double hydroxide, sulfide/selenide, phosphide/nitride, and carbon/alloy/MOF/organic based materials applied in different solution environments (*e.g.* acidic, alkaline, and seawater electrolyte) were discussed. And modification strategies aimed at an enhanced HER and OER performance have been evaluated. Subsequently, photocatalysts explored for water and HX splitting are also reviewed by elaborating modification approaches like heterojunction/cocatalyst loading/defect engineering/homojunction/elemental doping. More importantly, the enhancement mechanisms of  $\text{H}_2$  evolution were investigated from the catalyst's design to characterization. Accordingly, we propose the following issues that may be conducive to a deeper understanding of complex mechanisms and better development of potential nano-electro- or -photocatalysts.

(1) Adopting appropriate design and modification strategies for electrocatalysts according to specific reaction conditions is

important. By tailoring the physicochemical properties of electrocatalysts, such as their composition, morphology, and surface structure, researchers can significantly improve the catalysts' activity, stability, and durability under operating conditions.

(2) The photocatalytic  $\text{H}_2$  evolution process involves initial light absorption, charge carrier generation, and subsequent conversion of  $\text{H}_2\text{O}/\text{H}^+$  into  $\text{H}_2$ . Therefore, fabricating photocatalysts with efficient exciton separation and highly active surface catalytic sites is essential. Additional investigation of the complex relationship between the light exciton process and hydrogen evolution process on photocatalysts is also necessary. Understanding their interactions at a fundamental level will enable the development of photocatalysts with tailored properties, such as optimized band structures, enhanced charge carrier mobility, and increased surface reactivity, all of which are essential for improving the efficiency and scalability of  $\text{H}_2$  production from solar energy.

(3) It is encouraged to gain deeper insights into the underlying reaction mechanisms through the employment of characterization studies and theoretical calculations, to design and synthesize desired electrocatalysts and photocatalysts. With the aid of powerful characterization methods, detailed insights into the morphology, crystallinity, composition, surface states, and interfacial charge carrier dynamics of catalysts are readily obtained. Combined with predictive insights into the energetics, kinetics, and mechanisms of catalytic reactions that are offered by DFT, we can foster a feedback loop where experimental data inform theoretical models, which in turn predict new materials and modifications that can be tested experimentally.

## Author contributions

M.-I. J. and H.-Y. H. conceptualized the manuscript. M.-I. J. and H. T. wrote the original draft. M.-I. J., H. T., S. P. S., W. N., J.-J. K., C.-W. H., K.-C. C., F.-F. L., B. H., J. C. C. and H.-Y. H. contributed to the revision of the manuscript. H.-Y. H. supervised the writing and revision of the manuscript and provided the resources.

## Data availability

Data availability does not apply to this article because no new data are generated or analyzed in this study.

## Conflicts of interest

The authors declare no competing interests.

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