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Electrochemical surface reconstruction of nickel cobalt pyrophosphate to Ni/Co-hydroxide-(oxy) hydroxide: an efficient and highly durable battery-type supercapacitor electrode material†

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Developing transition metal-based electrode materials with high charge storage ability and excellent cycle life is a substantial challenge for supercapacitor applications. This report demonstrates an improved specific capacitance and excellent cycling stability for a Ni/Co-hydroxide-(oxy)hydroxide electrode material obtained from the electrochemical reconstruction of a Ni/Co-pyrophosphate precursor. Herein, we have prepared a series of nickel/cobalt pyrophosphate ($\text{Ni}_x\text{Co}_{100-x}$ 600) ($x = 90, 80, 70, 60, 50, 0$) materials through a simple co-precipitation method followed by thermal annealing. Upon electrochemical treatment, the developed nickel/cobalt pyrophosphate materials were transformed into porous Ni/Co-hydroxide-(oxy)hydroxides. This is due to the etching of pyrophosphate anions with electrochemical cycling, confirmed by rigorous physical characterization. Among them, Ni-rich electrodes showed higher specific capacitance values due to the dominant intercalation redox behavior (battery-type storage), yet they suffer from poor cycle stability. In contrast, Co-rich electrodes display low specific capacitance and excellent capacitance retention as the charge storage mainly occurs through surface redox processes. Interestingly, the incorporation of an adequate amount of cobalt into the nickel system has dramatically improved the charge storage ability as well as cycling stability owing to the synergistic effect between Ni and Co redox centers and the balanced charge storage phenomenon contributed by both the surface and intercalation redox processes. Especially, the material developed from $\text{Ni}_{60}\text{Co}_{40}$ 600 has exhibited a maximum specific capacitance of 566 F g^{-1} at 1 A g^{-1} current density and achieved an excellent capacitance retention of 84% ($\sim 99\%$ coulombic efficiency) up to 10 000 cycles. Further, the corresponding hybrid device ($\text{Ni}_{60}\text{Co}_{40}$ 600//AC) has displayed an operational voltage window of 1.6 V and achieved a maximum energy density of 14.9 W h kg^{-1} at 794.5 W kg^{-1} power density and could retain $\sim 86\%$ of its initial capacity even after 25 000 cycles, thereby indicating the suitability of the electrode for practical applications.

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Introduction

The modern world demands the inevitable use of energy storage systems (ESSs) in devices ranging from small electronic gadgets to heavy electric vehicles.^{1–4} In this context, two potential candidates, *i.e.*, batteries and supercapacitors (SCs), are found to be promising owing to their high energy and power output, respectively.^{5,6} Although SCs possess high power density, they

still suffer from low specific energy compared to batteries.⁷ To boost the energy density of SCs, the development of high specific capacity electrode materials and the construction of hybrid devices with a wide operational voltage window (V) is important.^{8–10} In this context, battery-type electrodes are emerging as suitable alternatives due to their promising specific capacity compared to pseudo-capacitors and electrochemical double-layer capacitors (EDLCs).^{11,12} These materials store charges through intercalation and surface redox processes.^{13,14}

Recently, Ni-based materials have been vastly explored for SCs due to their relatively cheap, earth-abundant nature, and battery-like charge storage nature.¹⁵ However, the poor cycling stability is a major bottleneck for Ni systems due to aggregation in longer cycles. Especially, nickel based sulfides (Ni_xS_y),^{16,17} selenides (Ni_xSe_y),^{18,19} phosphides (Ni_xP_y),^{20,21} phosphates

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(Ni_x(PO₄)_y),²² *etc.* delivered promising capacitance but suffer from poor cycling stability. There are a few reports that tried to enhance the cycling stability of the aforementioned materials. For instance, Nath *et al.* improved the cycling stability of nickel sulfide–nickel hydroxide structures by composite formation with thiophene and pyrrole-based copolymers.²³ The co-polymer improves the stability by limiting the aggregation of the active material. Further, the stability can be improved by tuning the morphology of Ni-based materials. Recently, Zhu and coworkers boosted up the cycling stability of flower-type Ni(OH)₂ particles through composite formation with ZnO flakes through electrodeposition.²⁴ In the composite structure, the individual ZnO and Ni(OH)₂ mutually prop each other, demonstrating excellent stability and morphology retention during a long run. In another report, Chen *et al.* demonstrated that cobalt incorporation in nickel selenide aids in the improvement of the cycling stability.²⁵ In this context, different Co based materials are found to be promising candidates for supercapacitors.^{26,27} Especially, Co based LDHs^{28,29} and oxyhydroxides³⁰ show excellent specific capacities and cycling stabilities.

Further, to improve charge storage, the development of highly porous electrodes with a larger number of metal centers is highly desired. In this context, electrochemical reconstruction of the precursor material is one of the ingenious strategies to enhance the electrochemical activity and stability. The reconstruction process is generally associated with the etching of functional moieties from the precursor materials during the electrochemical preconditioning process. The etching of the chemical groups (anionic and cationic) usually results in the exposure of more metal centers towards electrolytes and leads to facile transformation of metal precursors to their corresponding hydroxide/oxyhydroxide species. For instance, Liu *et al.* reported the transformation of trimetallic vanadium–cobalt–iron nitrides into their (oxy)hydroxide species by a surface reconstruction process *via* V etching.³¹ Similarly, P etching assists in the conversion of a NiCoP pre-catalyst to Ni(Co)OOH active species by surface reconstruction during the water oxidation reaction.³² In another report, Han *et al.* observed that cobalt iron phosphide displayed good electrochemical activity upon a bulk surface reconstruction process *via* etching of phosphorus.³³ Recently, our group has reported enhanced electrochemical activity due to porous hydroxide-(oxy)hydroxide formation by a structural reconstruction process *via* etching of phosphate and vanadate ions.^{34,35} Yet, the etching-based reconstruction phenomenon is majorly limited for water oxidation catalysts and scarcely reported for supercapacitor applications. Our idea is to develop transition metal pyrophosphates and evaluate their potential for supercapacitor applications. Metal pyrophosphates (M_xP₂O₇) possess interesting structures and are evolving as promising candidates for SCs due to their cost effectiveness, wide operational voltage window, and high redox activity.³⁶ The presence of pyrophosphate units can make the materials more porous through the etching of anionic moieties during electrochemical cycling. Further, there are very few reports on pyrophosphate systems explored for energy storage applications. For instance, Senthilkumar *et al.* studied the aqueous hybrid device performance

of Ni₂P₂O₇ and graphitic carbon, which displays a good operating voltage window of 1.6 V.³⁷ Similarly, Wang and co-workers developed a porous Ni₂P₂O₇ nanowire, which shows a good specific capacity of 772.5 F g⁻¹.³⁸ Despite achieving high specific capacities, the above systems lack in good cycling stability. Our idea is to introduce an appropriate amount of cobalt into nickel pyrophosphate to improve the capacity retention and stability. We presume that cobalt incorporation enhances the surface storage kinetics, thereby improving the cycling stability.

In this work, we have discussed the synthesis of amorphous nickel cobalt pyrophosphate (P₂O₇⁴⁻) which upon electrochemical bulk structural reconstruction generates active nickel–cobalt hydroxide-(oxy)hydroxide *via* etching of pyrophosphate ions. A series of bimetallic pyrophosphates Ni_xCo_{100-x} 600 (*x* = 100, 90, 80, 70, 60, 50, 0) are prepared by the co-precipitation method using water and ethylene glycol as solvents followed by thermal annealing. Interestingly, the introduction of Co into the Ni system improved both the charge storage ability and retention rate by controlling both the intercalation and surface redox phenomena. Particularly, the material developed after the reconstruction of Ni₆₀Co₄₀ 600 manifests a maximum specific capacitance of 566 F g⁻¹ at 1 A g⁻¹ current density and a good capacitive retention of 84% up to 10 000 cycles. The transformation of Ni₆₀Co₄₀ 600 to Ni₆₀Co₄₀(OOH) *via* phosphate etching and synergistic effects between Ni and Co centers was found to be the primary reason for the observed result. Further, a hybrid device was assembled using Ni₆₀Co₄₀ 600 and activated charcoal (AC). The hybrid device (Ni₆₀Co₄₀ 600//AC) provides an operational voltage window of 1.6 V and delivers a maximum energy density of 14.9 W h kg⁻¹ at 794.5 W kg⁻¹ power density. Additionally, the Ni₆₀Co₄₀ 600//AC device displayed excellent capacitive retention of 86% even after 25 000 cycles, indicating the potential of the developed material for practical applications.

Results and discussion

All the pyrophosphate materials were initially synthesized by a co-precipitation process using Co(NO₃)₂·6H₂O, NiCl₂·6H₂O, and (NH₄)₂HPO₄ as precursors. The above chemicals were dissolved in a water and ethylene glycol mixture and converted into ammonium metal hydrate. Subsequently, calcination at 600 °C results in desired pyrophosphate materials (see the details in the ESI†). For simplicity, the sample codes along with the Ni and Co ratio and the reaction temperature are summarized in Table 1.

Characterization

The phase information of the synthesized materials was identified by PXRD analysis. The diffraction patterns of as-prepared Co₁₀₀ RT and Ni₁₀₀ RT closely matched with the standard patterns of ammonium cobalt phosphate hydrate ([NH₄]CoPO₄·H₂O) (JCPDS No. # 01-089-6598) and ammonium nickel phosphate hydrate ([NH₄]NiPO₄·6H₂O) (JCPDS No. # 00-021-0034), respectively (Fig. S1a†). Upon calcination, Co₁₀₀ RT and Ni₁₀₀ RT at 600 °C were converted into the corresponding pyrophosphate and denoted as Co₁₀₀ 600 and Ni₁₀₀ 600,

Table 1 Sample codes for the as-prepared pyrophosphate materials with different Ni/Co ratios and reaction conditions

Sample codes	Ni/Co	Reaction temperature
Co ₁₀₀ RT	0/100 (only Co)	Room temperature (RT)
Ni ₁₀₀ RT	100/0 (only Ni)	RT
Ni ₆₀ Co ₄₀ RT	60/40	RT
Ni ₁₀₀ 600	100/0	600 °C
Co ₁₀₀ 600	0/100	600 °C
Ni ₉₀ Co ₁₀ 600	90/10	600 °C
Ni ₈₀ Co ₂₀ 600	80/20	600 °C
Ni ₇₀ Co ₃₀ 600	70/30	600 °C
Ni ₆₀ Co ₄₀ 600	60/40	600 °C
Ni ₅₀ Co ₅₀ 600	50/50	600 °C

respectively. The PXRD pattern of Co₁₀₀ 600 matched well with the monoclinic phase of Co₂P₂O₇ (JCPDS No. # 01-079-0825, *P*₂/*c* space group), whereas Ni₁₀₀ 600 displayed amorphous nature (Fig. 1a). This observation has been reported previously.^{37–40} Additionally, Ni₁₀₀ RT was calcined at 800 °C, and the PXRD pattern of the resulting sample (Ni₁₀₀ 800) matched with that of a pure phase of Ni₂P₂O₇ (JCPDS No. # 01-074-1604). The above results confirm that Ni₁₀₀ 600 is very likely the amorphous phase of Ni₂P₂O₇. The amorphous nature of Ni₆₀Co₄₀ 600 might be caused by the presence of higher Ni content in the system. Similar diffraction patterns are observed for the samples prepared with varying Ni/Co ratios except for Ni₅₀Co₅₀ 600 which resembles the Co₁₀₀ 600 pattern (Fig. S1b†). Further, thermogravimetric analysis (TGA) of Ni₁₀₀ RT and Ni₆₀Co₄₀ RT gives information about calcination temperatures required for the pyrophosphate formation. The calculated mass loss up to

600 °C is around 26% and 26.54% for Ni₁₀₀ RT and Ni₆₀Co₄₀ RT, respectively. This is maintained up to 800 °C in a static air atmosphere (Fig. S2†). The weight loss is associated with the elimination of ammonia and structural water molecules and results in the transformation of [NH₄]NiPO₄·6H₂O to Ni₂P₂O₇.^{38,41,42} Furthermore, the PXRD pattern of the post TGA-analyzed Ni₁₀₀ RT sample, denoted as Ni₁₀₀ (post TGA), displays a small peak at ~29°, confirming the formation of Ni₂P₂O₇ (Fig. 1a). Subsequently, the chemical functional groups present in the materials were studied by Fourier-transformed infrared (FTIR) analysis. The FTIR spectra of Ni₁₀₀ (post TGA) and all the annealed samples show two characteristic vibrational modes for PO₃ and P–O–P units (Fig. 1b and S3†). The band in the 1061–1084 cm⁻¹ and 568–585 cm⁻¹ regions can be attributed to PO₃ stretching and bending modes, respectively. In addition, a weak band noted in the 944–983 cm⁻¹ region and a strong band appearing around the 737–752 cm⁻¹ region are identified to be P–O–P bridge vibrations, confirming the presence of P₂O₇⁴⁻.^{43–46} To understand the porosity and surface area of the as-prepared materials, nitrogen adsorption desorption experiments were performed. Co₁₀₀ 600, Ni₁₀₀ 600, and Ni₆₀Co₄₀ 600 show a type-IV isotherm and H3 hysteresis loop, manifesting mesoporous nature and slit-like pore configuration in the materials (Fig. S4†).⁴⁷ The surface area values are found to be 32, 12, and 14 m² g⁻¹ for Ni₁₀₀ 600, Co₁₀₀ 600, and Ni₆₀Co₄₀ 600, respectively. In addition, the pore size distribution (PSD) was analyzed from the BJH distribution (Fig. S5†). The measured pore diameters for Ni₁₀₀ 600 and Co₁₀₀ 600 are 7.6 nm and 2.1 nm, respectively. The above values further corroborate the feasibility of enhanced intercalation or diffusion of electrolyte ions in Ni₁₀₀ 600 compared to Co₁₀₀ 600. Notably, an adequate

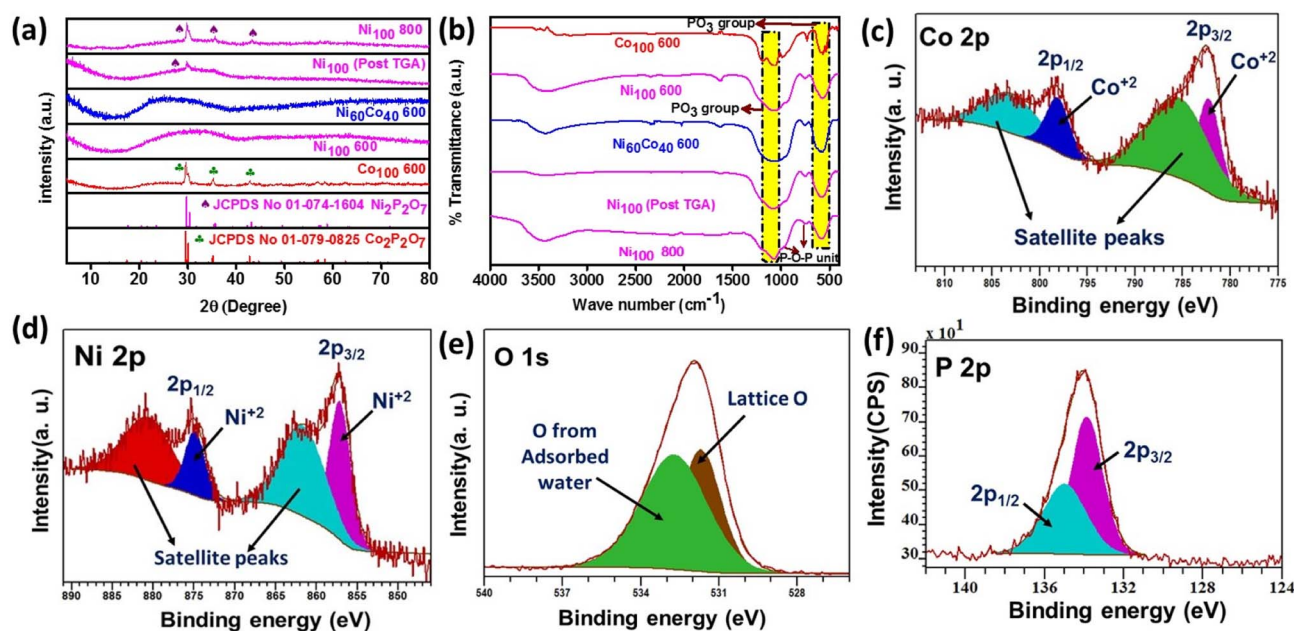


Fig. 1 (a) PXRD pattern and (b) FTIR spectra of Ni₁₀₀ 600, Co₁₀₀ 600, Ni₆₀Co₄₀ 600, Ni₁₀₀ 800, and Ni₁₀₀ (post TGA). The standard XRD patterns of Ni₂P₂O₇ (JCPDS No. # 01-074-1604) and Co₂P₂O₇ (JCPDS No. # 01-079-0825) are shown in (a). High-resolution XPS analysis of (c) Co 2p, (d) Ni 2p, (e) O 1s, and (f) P 2p for as-prepared Ni₆₀Co₄₀ 600.

quantity of Co introduction into the Ni system modulates the porosity in Ni₆₀Co₄₀ 600 (Table S1†). This helps regulate the charge storage mechanism by keeping the surface storage and intercalation phenomenon at an optimum level.

Chemical state analysis

The chemical state information of the elements present in Ni₆₀Co₄₀ 600 was obtained from X-ray photoelectron spectroscopic (XPS) analysis. The XPS survey scan of Ni₆₀Co₄₀ 600 confirms the presence of Ni, Co, O, and P elements in the material (Fig. S6†). Further, peak fitting of the Co 2p spectrum shows two main peaks positioned at 782.3 (2p_{3/2}) and 798.2 eV (2p_{1/2}). The binding energy (BE) difference of 15.9 eV between the peaks confirms the Co²⁺ state (Fig. 1c).^{48,49} This is further supported by the characteristic satellite signals of Co²⁺ positioned at 785.6 and 803.4 eV. In the Ni 2p spectrum, the 857.2 and 874.9 eV peaks correspond to Ni 2p_{3/2} and 2p_{1/2}, respectively. The noted peak separation of 17.7 eV suggests the presence of nickel in a +2 valence state (Fig. 1d).^{50,51} Additionally, the characteristic satellite peaks for Ni 2p noted at 862 and 880.7 eV confirm that nickel is present as Ni²⁺. The O 1s scan gives two peaks centered at 531.7 and 532.7 eV ascribed to lattice oxygen and adsorbed water molecules, respectively (Fig. 1e).³⁴ The peak fitted P 2p spectrum gives two peaks (due to spin-orbit splitting) at 134 and 134.9 eV for 2p_{3/2} and 2p_{1/2}, respectively. The observed peak separation of about 0.9 eV indeed authenticates the presence of a pyrophosphate ligand in the precursor framework (Fig. 1f).⁵² These aforementioned studies confirm the formation of bimetallic nickel cobalt pyrophosphate.

Morphological analyses

The morphology of the as-prepared materials was studied by field emission scanning electron microscopy (FESEM) and transmission electron microscopy (TEM). As shown in Fig. S7(a and b)†, Co₁₀₀ RT displays a hierarchical floret-like assembly,

presumably formed through the self-assembly of micro-petals. Fig. S7(c and d)† display random floppy-kind architectures (size ~30–60 nm) for Ni₁₀₀ RT. However, the observed morphology of Ni₆₀Co₄₀ RT is similar to that of Ni₁₀₀ RT (Fig. S7(e and f)†). Further, the calcined samples show isotropic morphological conversion along with some cracks throughout the surface. These pores could be attributed to the elimination of structural water and ammonia molecules (supported by TGA analysis) (Fig. 2a and b). Further, the Ni-rich system retained its initial particle-like texture even after calcination at 600 °C (Fig. S8(a–d)†). Interestingly, upon calcination of Ni₆₀Co₄₀ RT, the initial particles are rearranged and transformed into interconnected network-like morphology with extended connectivity in Ni₆₀Co₄₀ 600 (Fig. 2c). Similar observation was made in the Ni₇₀Co₃₀ 600 material (Fig. S8(e and f)†). Meanwhile the Ni₅₀Co₅₀ 600 resembles Co₁₀₀ 600 (Fig. S8(g and h)†). Further, the TEM analysis of Ni₆₀Co₄₀ 600 confirms the network-like morphology with long-range connectivity among the particles (Fig. 2(d–g)). EDS analysis corroborates the presence of Ni, Co, O, and P elements in Ni₆₀Co₄₀ 600 (Fig. 2h). In addition, the ICP-AES analysis shows that the Ni : Co ratio is 0.52 : 0.48 in the as-prepared Ni₆₀Co₄₀ 600 (Table S2†).

Electrochemical activation

The precursors were drop-cast on carbon paper and activated through electrochemical preconditioning using cyclic voltammetry (CV) at a 50 mV s⁻¹ scan rate in the voltage window 0–0.65 V (vs. Ag/AgCl) in 1 M KOH (see the details in the ESI†). Interestingly, CV results show a continuous increase in the current density up to ~200 cycles for Ni₁₀₀ 600, Co₁₀₀ 600, and Ni₆₀Co₄₀ 600 (Fig. 3(a–c)). Especially an observable improvement in the current density is noticed in the 200th cycle compared to the 1st cycle (Fig. 3(d–f)). This can be attributed to the occurrence of major structural reconstruction during electrochemical preconditioning. For instance, Allam and coworkers have studied a bimetallic Ni–Co oxy phosphide

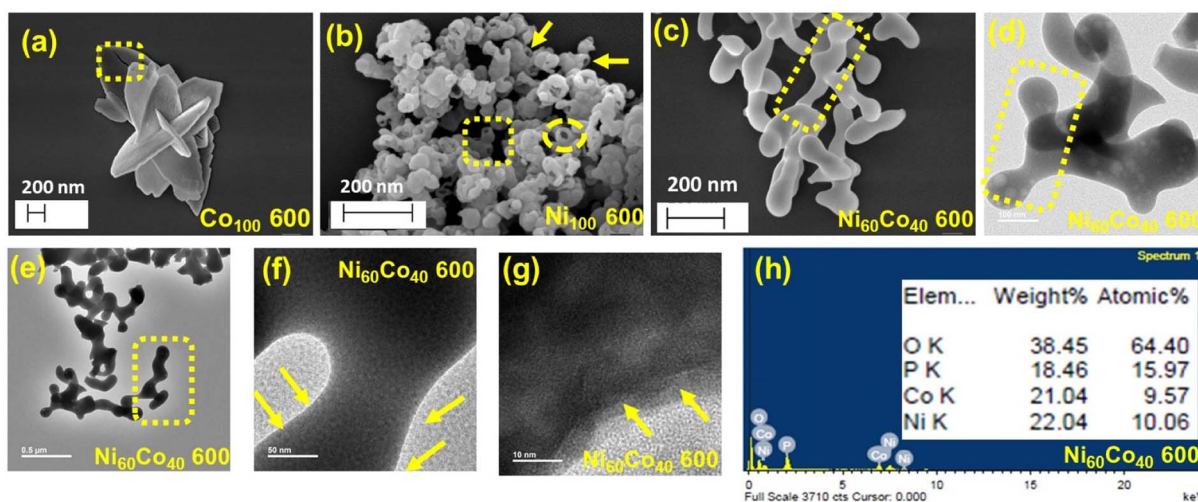


Fig. 2 FESEM images of as-prepared (a) Co₁₀₀ 600, (b) Ni₁₀₀ 600, and (c) Ni₆₀Co₄₀ 600, (d–g) TEM images and (h) EDS (Energy-dispersive X-ray spectroscopy) analysis for Ni₆₀Co₄₀ 600.

supercapacitor and emphasized the *in situ* generation of (oxy) hydroxide through structural reconstruction.⁵³ Further, during preconditioning, well-distinguished redox peaks characteristic of $\text{Co}^{2+}/\text{Co}^{3+}$, $\text{Co}^{3+}/\text{Co}^{4+}$, and $\text{Ni}^{2+}/\text{Ni}^{3+}$ redox couples are observed with stepwise increment in the current density. The CV curves illustrate the faradaic nature of the charge storage and the exposure of more Ni and Co centers to the electrolyte during activation.^{54,55} Similar results are observed for different compositions of Ni/Co optimized samples (Fig. S9†). The aforementioned results strongly suggest that metal pyrophosphates are likely converted into electrochemically active hydroxide-(oxy) hydroxide ($\text{M}(\text{OH})_2/\text{MOOH}$) ($\text{M} = \text{Ni}, \text{Co}$) species mostly through surface reconstruction (*vide infra*). This phenomenon has been earlier reported during electrochemical processes in some electrocatalysts.⁵⁶ Among all the prepared materials, $\text{Ni}_{60}\text{Co}_{40}$ 600 manifests the maximum area under CV curves during electrochemical preconditioning. To optimize the electrolyte concentration, we have performed the CV and GCD tests of $\text{Ni}_{60}\text{Co}_{40}$ 600 in 0.5 M, 1 M and 2 M KOH electrolytes. We observed the highest redox current and maximum specific capacitance in 1 M KOH (Fig. S10†) and hence 1 M KOH was chosen for further electrochemical studies. To get more information, detailed structural analyses were performed on the electrochemically activated $\text{Ni}_{60}\text{Co}_{40}$ 600 sample.

Post preconditioning characterization

To examine the chemical state, morphology, and elemental content of $\text{Ni}_{60}\text{Co}_{40}$ 600 after preconditioning, *ex situ* studies like XPS, FESEM, HRTEM, and ICP-AES were performed. The XPS survey scan of the electrochemically activated $\text{Ni}_{60}\text{Co}_{40}$ 600 electrode material confirms the presence of Ni, Co, and O, and a trace amount of P after preconditioning (Fig. S11†). Further,

the deconvoluted Co 2p spectrum shows peaks at 780.5 eV ($2p_{3/2}$) and 795.6 eV ($2p_{1/2}$). The noted BE difference of 15.1 eV between the two peaks corroborates Co^{3+} formation.⁵⁷ Additionally, the two peaks observed at 781.8 eV ($2p_{3/2}$) and 797.4 eV ($2p_{1/2}$) with a peak separation of ~ 16 eV suggest the presence of Co^{2+} (Fig. 4a).^{47,58,59} Further, after activation, all Ni^{2+} centers transformed to Ni^{3+} , which was confirmed by the presence of characteristic Ni^{3+} peaks at 855.9 eV ($2p_{3/2}$) and 873.7 eV ($2p_{1/2}$) (Fig. 4b).⁶⁰ The peak fitted O 1s spectra display three peaks at 529.6, 531.3, and 532.6 eV, which are assigned to lattice oxygen, O from the hydroxyl group in an alkaline environment and adsorbed water molecules, respectively (Fig. 4c).^{48,61} These observations further validate the formation of MOOH ($\text{M} = \text{Ni}, \text{Co}$) during the preconditioning process. Interestingly, the P 2p spectrum shows an observable decrease in phosphorus content compared to the as-prepared material (Fig. 4d). This implies the etching of P (likely as pyrophosphate) during preconditioning. Further ICP-AES analysis was performed with the electrolyte sequentially collected after 0, 25, 50, 100, 150, and 200 CV cycles of preconditioning for the $\text{Ni}_{60}\text{Co}_{40}$ 600 sample. Interestingly, the amount of P in the aliquots gradually increased from 0 to 150 cycles in the electrolyte. This observation confirms that P etches out likely as $\text{P}_2\text{O}_7^{4-}$ during the electrochemical preconditioning of the working electrode (Fig. S12 and Table S3†). Interestingly, electrochemical preconditioning results in a noticeable change in the morphology of the material. The FESEM study (on the carbon paper loaded material) on the electrochemically preconditioned $\text{Ni}_{60}\text{Co}_{40}$ 600 indicates that the initial network structure transformed into an agglomerated texture (Fig. S13†). Further, close inspection of TEM images revealed the formation of a porous flake-like morphology, which could be beneficial for better charge storage (Fig. 4e and f). The HRTEM image and its corresponding FFT pattern additionally

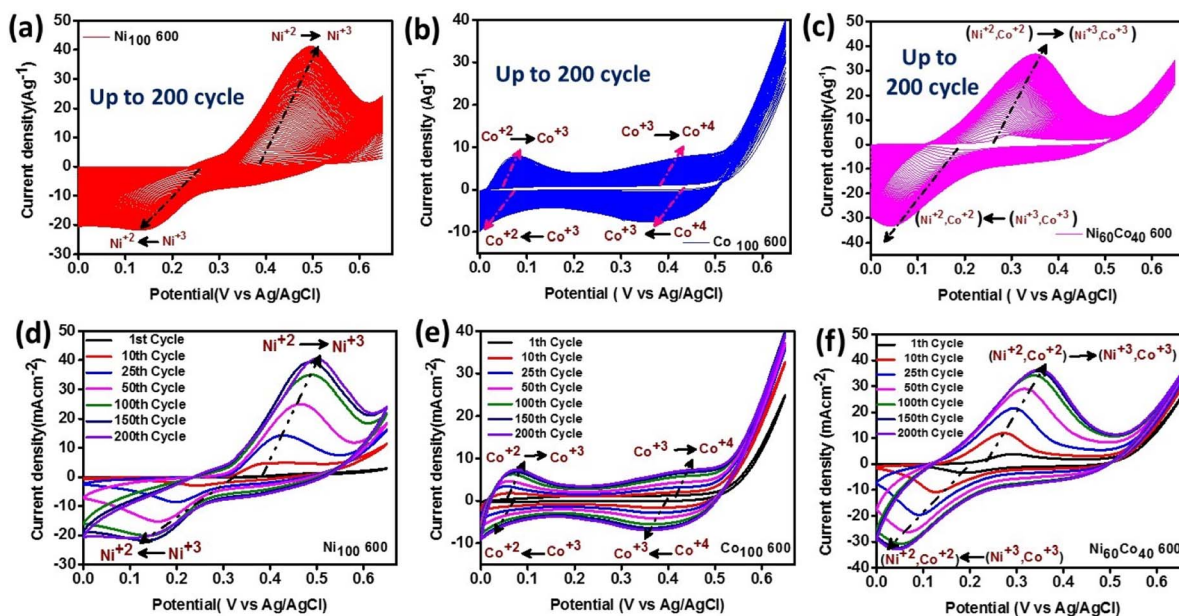


Fig. 3 CV curves indicate current density improvement during preconditioning (1 to 200 cycles) for (a and d) Ni_{100} 600, (b and e) Co_{100} 600 and (c and f) $\text{Ni}_{60}\text{Co}_{40}$ 600.

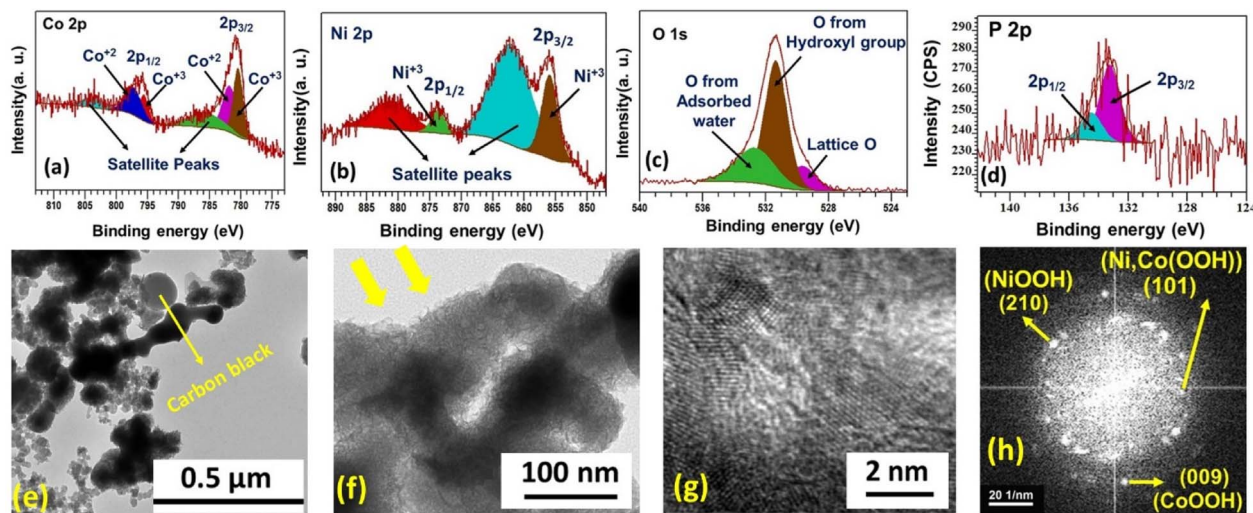


Fig. 4 XPS spectrum of (a) Co 2p, (b) Ni 2p, (c) O 1s, and (d) P 2p after preconditioning of Ni₆₀Co₄₀ 600, and (e and f) TEM images, (g) HRTEM image and (h) corresponding FFT pattern of Ni₆₀Co₄₀ 600 after electrochemical preconditioning. In the inset of figure (e), the sphere represents the carbon black used for the preparation of the working electrode.

support the existence of (210) and (009) planes characteristic of NiOOH (JCPDS No. # 00-027-0956) and CoOOH (JCPDS No. # 07-0169), respectively (Fig. 4g).^{62,63} Moreover, the appearance of the (101) plane in the FFT pattern confirms the formation Ni₂Co(OOH) in the activated material (Fig. 4h).^{64,65} Based on our understanding and literature survey, we believe that the formation of Ni,Co hydroxide/(oxy)hydroxide from Ni,Co pyrophosphates is facilitated by the phosphate etching during the electrochemical process (detailed chemical mechanism is given in the ESI, page no. S24†).

Electrochemical performance of the activated material. The reconstructed materials (active materials) after electrochemical preconditioning were evaluated for supercapacitor performance. Initially, CV measurements of all the activated materials were recorded at a 5 mV s⁻¹ scan rate (within 0–0.65 V vs. Ag/AgCl). The CV results reveal that the area under the CV curve for the sample Ni₁₀₀ 600 is higher compared to the Co₁₀₀ 600 sample demonstrating higher charge storage ability in the former (Fig. 5a). Upon Co incorporation the charge storage ability was altered. Especially, the material obtained after preconditioning of Ni₆₀Co₄₀ 600 displays maximum charge storage ability compared to Co₁₀₀ 600, Ni₁₀₀ 600, and all Ni/Co materials with different compositions (Fig. 5a and S14†). Further, the CVs of all activated materials were recorded at varying scan rates (5 to 50 mV s⁻¹). A significant increase in the redox current is noted for the activated Ni₆₀Co₄₀ 600 with increasing scan rate and the cathodic and anodic peaks shifted toward more negative and positive potential regions, respectively. The notable peak-to-peak gap at a high scan rate indicates that charge storage is principally diffusion-controlled (Fig. S15†). A similar behaviour was observed for all other materials with different Ni/Co compositions (Fig. S16†). Subsequently, the specific capacity (*C_s*) of all activated materials has been measured from their CV curves (see eqn (1) and (2), ESI†). Among all the materials, electrochemically activated Ni₆₀Co₄₀ 600 displays the highest *C_s*

value. Further, with increasing scan rates, the *C_s* value decreases, which is attributed to insufficient time for the charge cumulation towards the electrode (Fig. 5b and S17†).⁶⁶ Meanwhile, the Co incorporation into the nickel system increased the capacitive retention rate. For example, the activated Co₁₀₀ 600 and Ni₁₀₀ 600 retained 72% and 44% of their initial specific capacity, respectively, at a 50 mV s⁻¹ scan rate. Interestingly, after cobalt incorporation, the retention rate increased to 52% in Ni₆₀Co₄₀ 600 at the same scan rate (Fig. 5b).

Mechanism. Post Co incorporation into Ni systems, the observed increase in the area under the CV curve and redox current density is ascribed to the synergistic interaction of Ni and Co centers in the bimetallic system. This has enthused us to look into the mechanistic details of charge storage trajectory occurring in different electrochemically activated materials. This is done by calculating the *b* value, which is estimated from the slope of the log(peak current) vs. log(scan rate) plot using the power law (eqn (1) and (2)).⁶⁷

$$i = av^b \quad (1)$$

$$\log(i) = \log(a) + b \log(v) \quad (2)$$

where *v* denotes the scan rate (mV s⁻¹), *i* is the cathodic (*i_c*) or anodic (*i_a*) peak current (mA), and *a* and *b* are variable parameters. The *b* value specifies the behaviour of charge storage. For example, when the *b* value is ~1, the *i* (= *i_c* or *i_a*) is linearly proportional to the scan rate (*i* ∝ *v*), which signifies that the material stores the charge by a surface redox process (pseudocapacitor). Meanwhile, a *b* value of ~0.5 indicates intercalation redox or diffusion control behaviour (battery-type capacitor).⁵ The estimated *b* values for Ni₁₀₀ 600 and Co₁₀₀ 600 are 0.5 and 0.98, indicating that charge storage mainly occurs through intercalation (battery-type) and surface redox (pseudocapacitive nature) processes, respectively. Interestingly, gradual incorporation of Co into the Ni system increases the *b* value successively

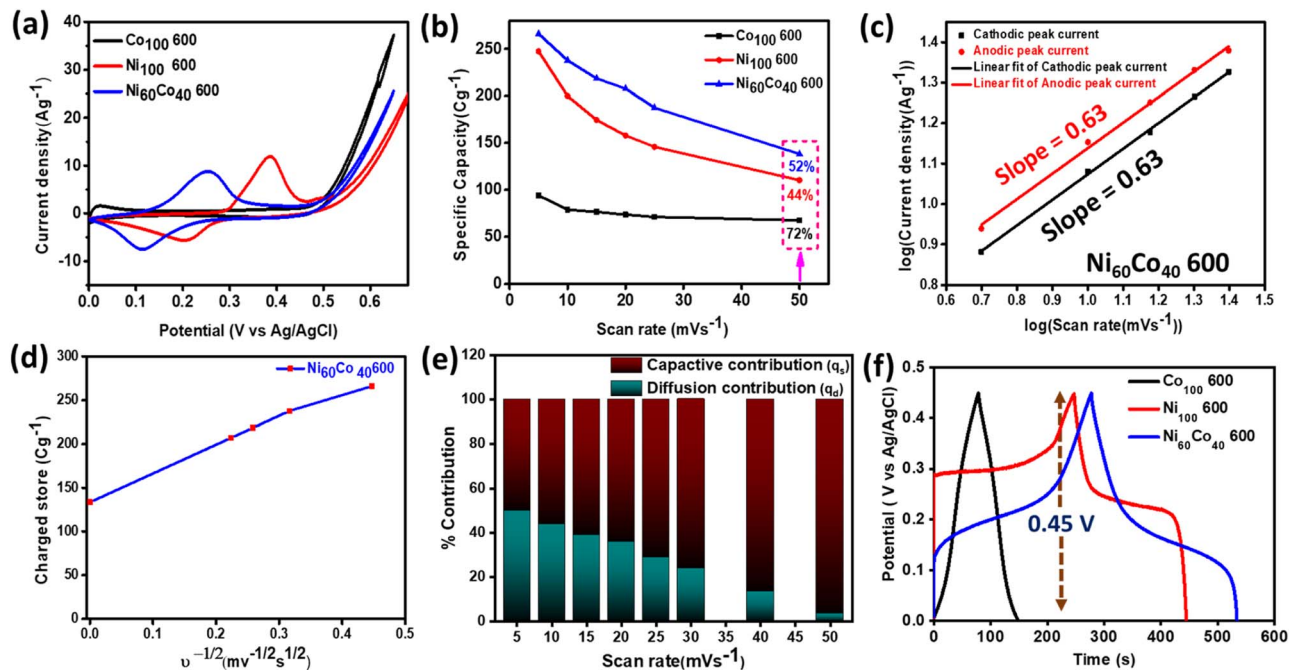


Fig. 5 (a) CV curves, (b) plot of C_s vs. scan rate for activated Ni_{100} 600, Co_{100} 600, and $Ni_{60}Co_{40}$ 600 samples, (c) b value estimation (within the 5 to 25 mV s^{-1} scan rate range), (d) C_s vs. $\nu^{-1/2}$ plot for activated $Ni_{60}Co_{40}$ 600, (e) the bar diagram showing % contribution from diffusion and capacitive processes at 5 to 50 mV s^{-1} scan rates and (f) GCD curves for Ni_{100} 600, Co_{100} 600, and $Ni_{60}Co_{40}$ 600 recorded at 1 A g^{-1} .

(Fig. S18).† This indicates that Co modulates the charge storage phenomenon by increasing the surface contribution to the Ni system.

Especially, the b value for the activated $Ni_{60}Co_{40}$ 600 is 0.63, suggesting that the material stores the charge through both intercalation and surface redox processes (Fig. 5c). Further, Trasatti's method was adopted to quantify both diffusion and capacitive contributions towards the charge storage of the materials. Generally, Trasatti's method correlates the diffusion (q_d) and surface (q_s) stored charges with varying scan rates. The total charge (q_t) accumulated under the CV curve is a combination of the capacitive component (q_s) and diffusive component (q_d) (eqn (3) and (4)).⁶⁸

$$q_t = q_s + q_d \quad (3)$$

$$q_t = q_s + \text{const } \nu^{-1/2} \quad (4)$$

Normally, the q_s linearly varies with the scan rate (ν) due to quicker kinetics at the electrode surface compared to the bulk. On the other hand, the q_d component varies as the inverse square root of scan rate ($\nu^{-1/2}$). Generally, at a low scan rate due to sufficient time for diffusion, intercalation kinetics are dominant. Further, at a high scan rate ($\nu \rightarrow \infty$) the entire contribution predominately arises from the surface of the electrode (capacitive contribution).⁶⁹ Finally, the y -axis intercept from the q_t vs. $\nu^{-1/2}$ plot (at 5 to 20 mV s^{-1}) gives the information of q_s (Fig. 5d). Subsequently, the bulk contribution (q_d) is obtained by subtracting q_s from q_t . The contributions of q_s and q_d for Ni_{100} 600 and Co_{100} 600 are shown as a bar diagram in Fig. S19.† At a low scan rate (5 mV s^{-1}), Ni_{100} 600 and Co_{100} 600 show

79.4% and 42.6% diffusion contributions, respectively (Fig. S20†). These outcomes further support the inherent diffusion behaviour of the Ni system over the Co system and correlate with the estimated b value (*vide supra*). Interestingly, after the incorporation of an optimum quantity of Co into the Ni system, the activated $Ni_{60}Co_{40}$ 600 shows a 49.8% diffusion contribution (at a 5 mV s^{-1} scan rate). Further, at a high scan rates $Ni_{60}Co_{40}$ 600 displays 28.9% (25 mV s^{-1}) and only 3.5% (50 mV s^{-1}) diffusion contribution indicating predominant surface contribution towards the charge storage at a very high scan rate (Fig. 5e and S21).†

The galvanostatic charge–discharge (GCD) curves of all electrochemically activated materials were recorded at 1 A g^{-1} current density within the voltage window of 0–0.45 V. In the case of Ni_{100} 600, the GCD curve shows a long plateau, indicating battery-type charge storage or intercalation redox characteristics.⁷⁰ In contrast, Co_{100} 600 shows a quasi-triangular GCD profile suggesting surface redox or intrinsic pseudocapacitive behaviour. Interestingly, the GCD curve for $Ni_{60}Co_{40}$ 600 contains both a plateau and distorted quasi-triangular shape (Fig. 5f). Similar results are noticed for materials with different Ni/Co ratios (Fig. 6a). These results further support that Co incorporation into Ni systems results in both intercalation and surface contributions towards charge storage, *i.e.*, intercalation type pseudocapacitance (Fig. S22†).

Besides, the GCD profiles display a small iR drop associated with internal solution resistance and electrical contacts between the electrode and current outlets. Among the different activated materials, $Ni_{60}Co_{40}$ 600 delivered the maximum discharge time and highest specific capacitance (C_{sp}) value. The calculated C_{sp}

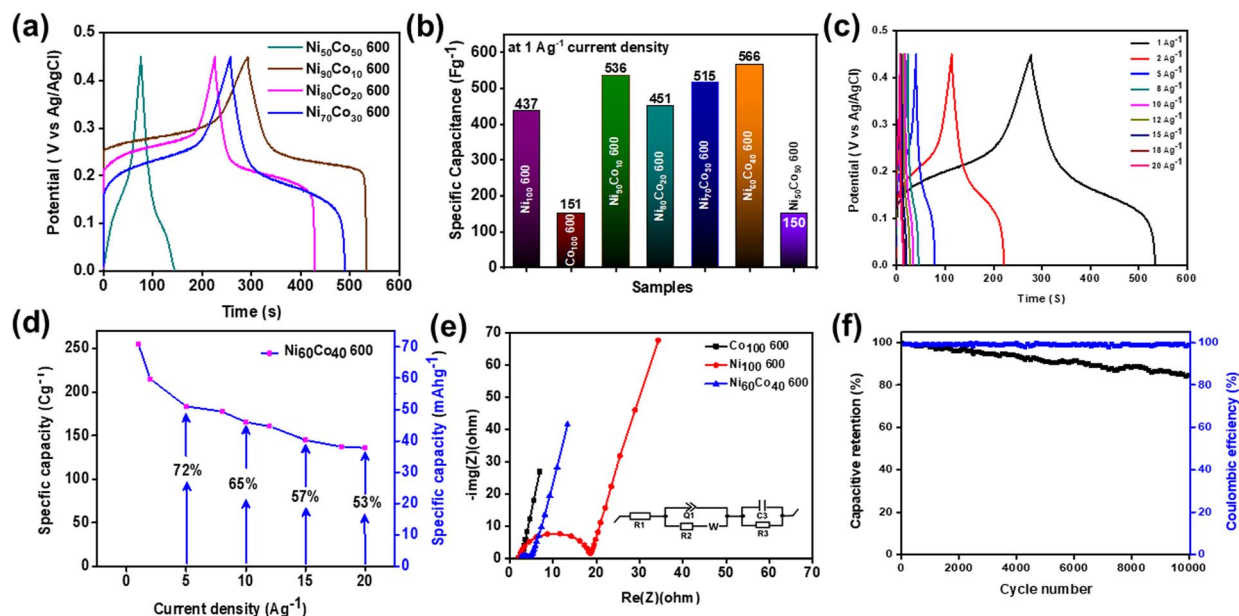


Fig. 6 (a) GCD curves (recorded within 0–0.45 V and 1 A g^{-1}), (b) bar diagram showing the C_{sp} for the as-prepared material with different Ni/Co ratios, (c) GCD curves of activated $\text{Ni}_{60}\text{Co}_{40} 600$ at various current densities ($1\text{--}20 \text{ A g}^{-1}$), (d) C_s vs. current density plot, (e) Nyquist plot for activated $\text{Ni}_{100} 600$, $\text{Co}_{100} 600$ and $\text{Ni}_{60}\text{Co}_{40} 600$ (inset shows the Randles circuit for impedance data fitting, where R , W , C , and Q parameters are the resistor, Warburg impedance, capacitor, constant phase element, respectively), and (f) durability test of activated $\text{Ni}_{60}\text{Co}_{40} 600$, performed at constant 10 A g^{-1} current density (up to 10 000 cycles).

values (recorded at 1 A g^{-1} current density) of the respective materials are shown as a bar diagram in Fig. 6b. Subsequently, the GCDs of all activated materials were measured at different current densities (1 to 20 A g^{-1}) (Fig. 6c and S23).[†] The calculated specific capacity (C_s) value of electrochemically activated $\text{Ni}_{60}\text{Co}_{40} 600$ is 254.79 C g^{-1} ($70.77 \text{ mA h g}^{-1}$) at 1 A g^{-1} current density. Further, with increasing current densities, the C_s values decreased (Fig. 6d and S24).[†] This is due to limiting diffusion time in a faster charge–discharge process.⁷¹ On the other hand, the activated $\text{Ni}_{60}\text{Co}_{40} 600$ could retain 72% of its initial capacity at 5 A g^{-1} current density and 53% at 20 A g^{-1} indicating the good rate capability of $\text{Ni}_{60}\text{Co}_{40} 600$ (Fig. 6d). Further the diffusion and interfacial charge transfer resistance (R_{ct}) of the developed materials were monitored using electrochemical impedance spectroscopy (EIS) at the open circuit potential (OCP). The Nyquist plot of electrochemically activated $\text{Co}_{100} 600$, $\text{Ni}_{100} 600$, and $\text{Ni}_{60}\text{Co}_{40} 600$ displays a semicircle and a vertical line in the high and low frequency zones, respectively (Fig. 6e). At high frequencies, the real axis intersection denotes equivalent series resistance (ESR) due to interfacial connections between the electrode material/current outlet and the inherent resistance of the electrolyte ions. At low frequency, the semicircle diameter and vertical line slope represent R_{ct} and Warburg impedance (Z_w) associated with diffusion kinetics, respectively.^{68,72} Interestingly, after introduction of the requisite amount of Co, the semicircle diameter decreased in activated $\text{Ni}_{60}\text{Co}_{40} 600$ compared to $\text{Ni}_{100} 600$. This observation indicates that cobalt incorporation into Ni systems helps in improving the charge transfer kinetics. Further, the observed $\sim 70^\circ$ slope of the vertical line for $\text{Ni}_{60}\text{Co}_{40} 600$ manifests facile diffusion kinetics of electrolyte ions during charge storage.

Cycling stability is another important parameter to evaluate the suitability of the materials for supercapacitors. All the activated Ni/Co materials were subjected to continuous 5000 GCD cycles at 10 A g^{-1} current density. The activated $\text{Ni}_{100} 600$ shows only a 39% retention rate which suggests poor electrochemical robustness of the Ni system. Meanwhile $\text{Co}_{100} 600$ shows 96% capacitive retention indicating good stability (Fig. S25a[†]).

Interestingly, after gradual Co incorporation into the Ni system, cycling stability increases effectively. Especially, the activated $\text{Ni}_{60}\text{Co}_{40} 600$ displays an excellent retention rate of 94% (up to 5000 cycles). The retention rate for other materials with different Ni/Co compositions is shown in Fig. S25b.[†] From the above results, a 3D scatter plot is drawn showing the b value (both cathodic and anodic) and % capacitive retention to understand the relationship between the charge storage procedure and stability (Fig. S26).[†]

The 3D plot further proves that gradual Co incorporation increases the b value and surface contribution to the overall charge storage process. In addition, the plot shows an increase in the stability of the Ni system with Co incorporation. Further, among all the activated materials, $\text{Ni}_{60}\text{Co}_{40} 600$ offers the best performance (Table S4).[†] This has encouraged us to extend the cycling stability test for longer cycles. The activated $\text{Ni}_{60}\text{Co}_{40} 600$ shows a retention rate of 84% up to 10 000 cycles at 10 A g^{-1} current density and $\sim 99\%$ coulombic efficiency indicating excellent cycling stability and reversibility, respectively (Fig. 6f). The observed high stability value for $\text{Ni}_{60}\text{Co}_{40} 600$ is found to be better compared to other reported pyrophosphate materials (see Table S7).[†] Further, the stability test indicates that the final

GCD curve shows only a little decrease in the plateau region as well as in the discharge time compared to the first cycle (Fig. S27 and S28).[†] This could be assigned to mild suppression of electrolyte migration channels present into the material during long GCD cycles.

Characterization after the electrochemical stability test

Meticulous physical characterization has been performed on the material after the stability studies. The XPS spectrum of the deconvoluted P 2p spectrum displays no characteristic signal indicating that the material's surface is devoid of any P (Fig. S29[†]). In addition, the P 2p XPS spectra of the as-prepared Ni₆₀Co₄₀ 600, after preconditioning treatment and post-stability measurement, are compared in Fig. 7a. The results clearly indicate the etching of P (likely as P₂O₇⁴⁻) during preconditioning. In addition, the XPS study confirms the existence of Co²⁺ and Co³⁺ along with Ni³⁺ even after

a long-term durability experiment (Fig. S30).[†] The peak position values are shown in Table S5.[†] The post-stability Co, Ni 2p, and O 1s XPS spectra closely match the XPS results obtained after preconditioning, revealing the structural robustness of the activated material. The above results further indicate the possible formation of M(OH)₂/MOOH species (M = Ni, Co) in the post-stability material. Duan *et al.* demonstrated similar P etching through an electrochemical transformation from FePO₄ in a nickel rich system to NiFe-OOH.⁷³ Recently, Shreenivasa *et al.* noticed the absence of P in the post-XPS spectra owing to the surface reconstruction of BiPO₄ to Bi₂O₃.⁷⁴ Further, ICP-AES analysis of as-prepared Ni₆₀Co₄₀ 600, after preconditioning and post stability testing, further proves that the majority of P etching occurs from the as-prepared material during the preconditioning and hardly any change is noted after stability (Table S6).[†] The mechanism of the etching phenomenon is summarised and mentioned in

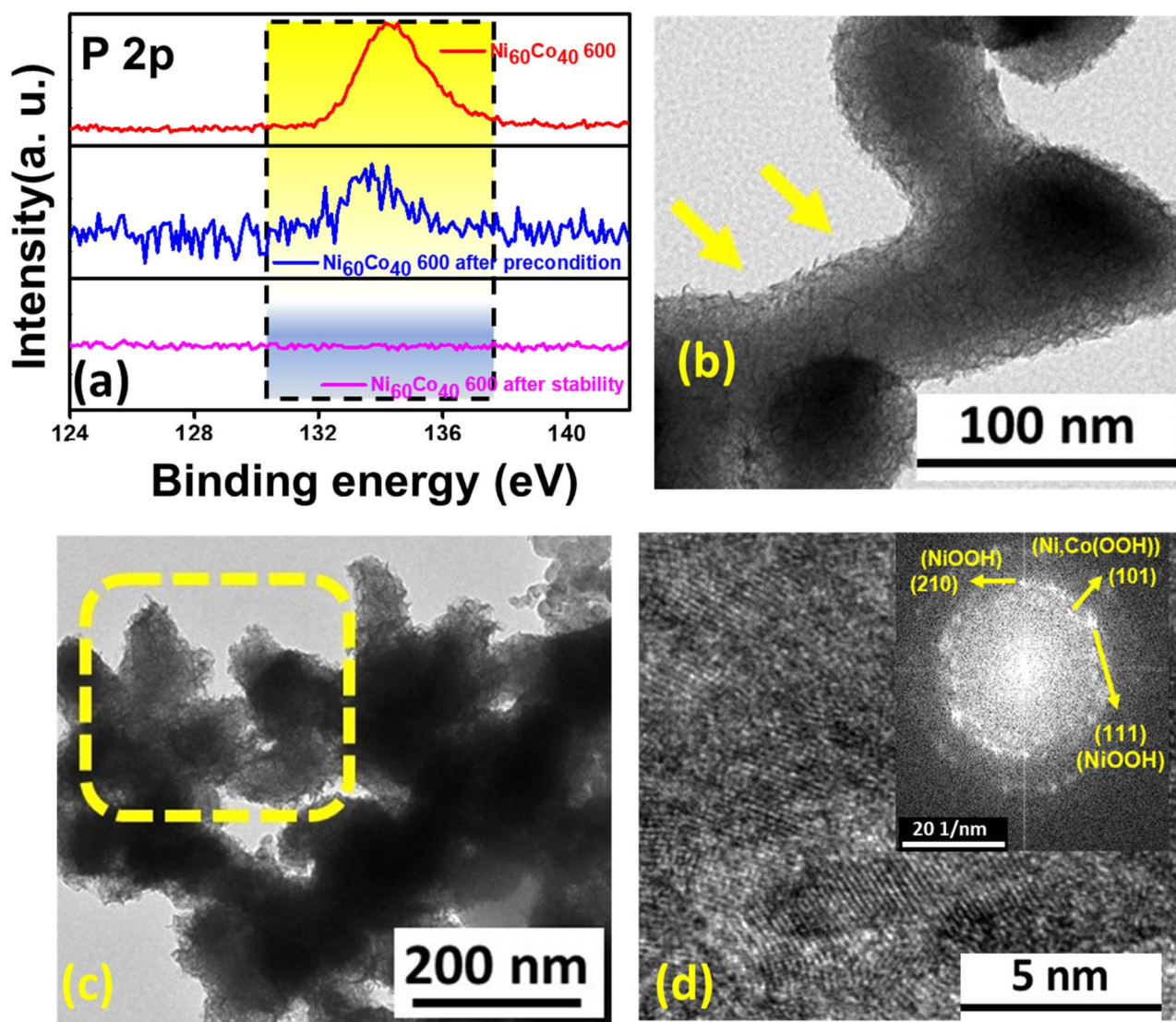


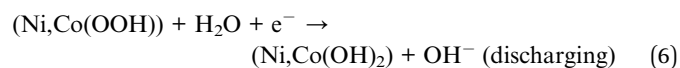
Fig. 7 (a) Comparison of total P intensity in the as-prepared material, after preconditioning and after stability testing, (b and c) TEM images, and (d) HRTEM image and corresponding FFT pattern (inset).

detail in the ESI.† The TEM images of the post stability $\text{Ni}_{60}\text{Co}_{40}$ 600 sample display retention of the porous flake-like morphology as acquired after preconditioning (Fig. 7(b and c)). Moreover, the HRTEM image and the corresponding FFT pattern show (210), (111), and (101) planes, which corroborates the formation of NiOOH and $\text{Ni}_2\text{Co}(\text{OOH})$, respectively (Fig. 7d).^{62,64,65,75} Further, the EDS study and elemental mapping analyses confirmed the presence of Co, Ni, O, and C (from the carbon paper substrate) and significant loss of P in the material after the extended durability test (Fig. S31).† The results above strongly support the formation of redox-active $\text{M}(\text{OH})_2/\text{MOOH}$ ($\text{M} = \text{Ni}, \text{Co}$) species through structural reconstruction of NiCo pyrophosphate.

Further, to understand the importance of Ni_2Co hydroxides/oxyhydroxides derived through electrochemical reconstruction of the corresponding pyrophosphates, Ni_2Co hydroxides were directly prepared (synthesis and characterisation are given in ESI page no. S22, Fig. S32 and S33†) and their activity was evaluated. The electrochemically activated $\text{Ni}_{60}\text{Co}_{40}$ 600 displays better reconstruction (Fig. S34(a–d)†) and a higher area under the CV curve, compared to the as-synthesized hydroxides, manifesting better charge storage ability of the reconstructed material (Fig. S34e).† In addition, the GCD curves additionally support that activated $\text{Ni}_{60}\text{Co}_{40}$ 600 has a higher discharge time than the as-prepared Ni_2Co hydroxide (Fig. S34f).† The calculated C_{sp} for Ni_2Co hydroxide is 220 F g^{-1} (at 1 A g^{-1}), which is much lower than that of $\text{Ni}_{60}\text{Co}_{40}$ 600, indicating the suitability of pyrophosphate as an efficient precursor material for supercapacitor applications (Fig. S34g).†

Further, the PXRD analysis was performed on the carbon paper-supported material to unveil the phase formation at charged and discharged states. The broad diffraction peaks

between ~ 36 and 39° and ~ 65 and 66.8° for the charged electrode indicate the formation of $(\text{Ni},\text{Co}(\text{OOH}))$ (NiOOH JCPDS No. # 00-027-0956 and CoOOH JCPDS No. # 07-0169) (Fig. S35a).†^{76,77} In the case of the discharged electrode, the broad peaks observed between ~ 37 and 38.6° and ~ 61.4 and 63° suggest the formation of nickel cobalt hydroxide $(\text{Ni},\text{Co}(\text{OH})_2)$ ($\text{Ni}(\text{OH})_2$ JCPDS No. # 01-073-1520 and $\text{Co}(\text{OH})_2$ JCPDS No. # 00-045-0031) (Fig. S35b).† Therefore, the proposed charged and discharged reaction could be attributed as follows (eqn (5) and (6)).



Hybrid device performance

The noted high cycling stability and good charge storage capability of $\text{Ni}_{60}\text{Co}_{40}$ 600 motivated us to construct an asymmetric hybrid device. The hybrid device was fabricated in a two-electrode setup using $\text{Ni}_{60}\text{Co}_{40}$ 600 as the positive electrode and activated charcoal (AC) as the negative electrode (Fig. S36).† Initially, the CVs of positive and negative electrodes were separately recorded in a three-electrode setup. AC shows an EDLC nature, whereas activated $\text{Ni}_{60}\text{Co}_{40}$ 600 shows battery-type behaviour within their operational voltage window of -1 to 0 V and 0 – 0.6 V , respectively (Fig. 8a). Before constructing the device, the mass balance between positive and negative electrodes was maintained by the charge/mass balance equation (ESI†). Subsequently, a $\text{Ni}_{60}\text{Co}_{40}$ 600//AC hybrid device was

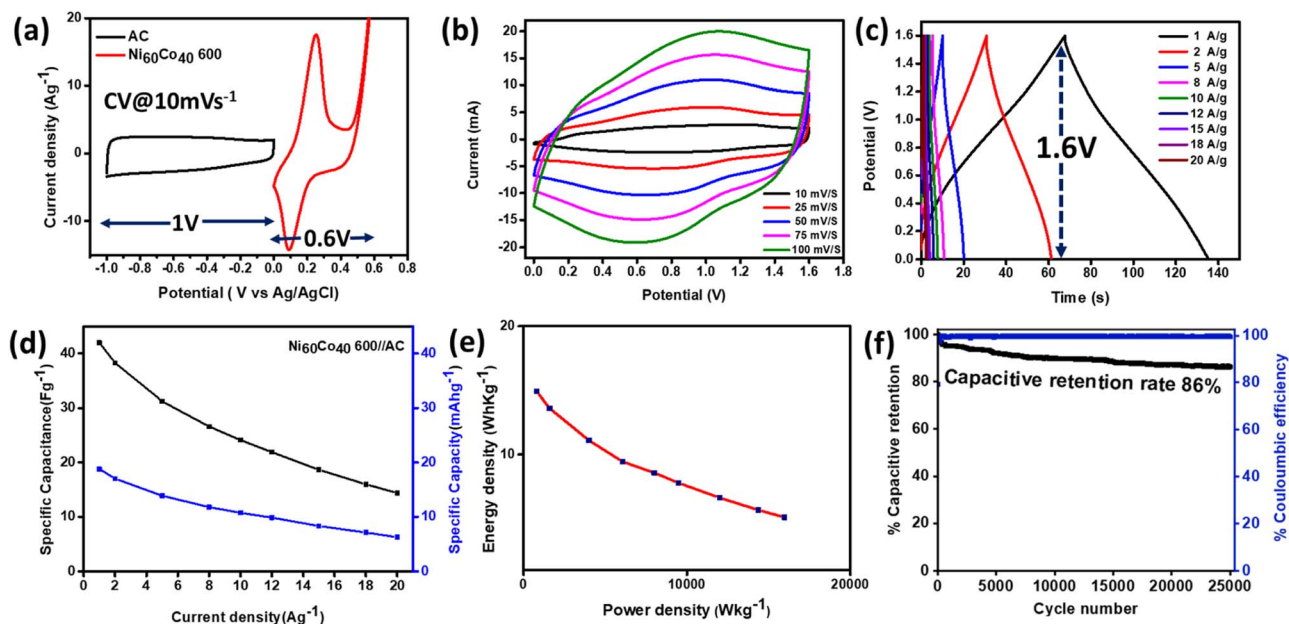


Fig. 8 CVs of AC and $\text{Ni}_{60}\text{Co}_{40}$ 600 measured in a 3-electrode setup, (b) CVs of the $\text{Ni}_{60}\text{Co}_{40}$ 600//AC hybrid device measured at varying scan rates, (c) GCD curves for $\text{Ni}_{60}\text{Co}_{40}$ 600//AC recorded at individual current density (within 0–1.6 V), (d) variation of C_{sp} and C_{s} vs. current density, (e) Ragone plot and (f) cycling stability test (at 10 A g^{-1} current density) of the hybrid device.

constructed with a series connection between Ni₆₀Co₄₀ 600 and AC. The voltage window of the hybrid device was determined by recording the CVs (5 mV s⁻¹ scan rate) at 1.2 to 1.8 V. The CV result suggests that Ni₆₀Co₄₀ 600//AC has achieved a maximum operational voltage window of 1.6 V without any polarization (Fig. S37a).† The CVs of the hybrid device were recorded by varying the scan rates (10 to 100 mV s⁻¹ within 0 to 1.6 V), which display both EDLC and redox behaviour (Fig. 8b).⁷⁸ The 1.6 V operational voltage window for the Ni₆₀Co₄₀ 600//AC hybrid device is further confirmed by recording the GCD in separate voltage windows (Fig. S37b).† Subsequently, the GCDs of the Ni₆₀Co₄₀ 600//AC hybrid device were collected at different current densities (1 to 20 A g⁻¹) at 1.6 V (Fig. 8c). The noted specific capacitance (specific capacity) value for the Ni₆₀Co₄₀ 600//AC device is 42 F g⁻¹ (18.79 mA h g⁻¹) at 1 A g⁻¹, which could retain about 14.37 F g⁻¹ (6.22 mA h g⁻¹) at 20 A g⁻¹ current density, indicating a reasonable retention rate (Fig. 8d). The power and energy densities of the hybrid device were calculated from different GCD values (Fig. S38).† It is clear from the Ragone plot shown in Fig. 8e that the hybrid device displays a maximum of 14.9 W h kg⁻¹ energy density at a power density of 794.5 W kg⁻¹. The value obtained for the carbon paper deposited substrate is quite comparable with values reported in the literature for other pyrophosphate-based materials (Table S8).† In addition, the hybrid device could retain 5.1 W h kg⁻¹ energy density at a high-power density of 15 965.2 W kg⁻¹. Further, the hybrid device shows 86% capacitive retention and ~99% coulombic efficiency up to 25 000 GCD cycles, implying excellent cycling stability and good chemical reversibility (Fig. 8f). The comparison of the electrochemical activity, including long term stability performance of the developed Ni₆₀Co₄₀ 600//AC device with the reported pyrophosphate-based hybrid devices is illustrated in Table S8.† The data support the impressive cycling stability and capacitive retention of the Ni₆₀Co₄₀ 600//AC hybrid device compared to other systems, mainly using a cost-effective carbon substrate.

Conclusions

This work demonstrated an electrochemically induced surface reconstruction process for developing an efficient charge storage material. A simple co-precipitation process using different proportions of nickel and cobalt salts, followed by annealing, produced a series of amorphous bimetallic NiCo pyrophosphate precursors. Electrochemically induced reconstruction of the prepared bimetallic pyrophosphates resulted in the formation of highly porous Ni/Co hydroxide-(oxy)hydroxide species. Among the different pyrophosphate precursors, Ni-rich materials showed strong intercalation or battery-type charge storage behaviour and high specific capacitance compared to Co-rich systems. In contrast, the Co-rich system provides good cycling stability over the Ni system owing to predominant surface redox charge storage. Interestingly, introducing the requisite amount of Co into Ni systems improves the specific capacitance (due to the synergistic effect) and the cycling stability by modulating the intercalation and surface redox contribution towards the charge storage process. Particularly,

the material obtained from the reconstruction of Ni₆₀Co₄₀ 600 showed a maximum specific capacitance of 566 F g⁻¹ at 1 A g⁻¹ current density through both intercalation and surface redox processes. Further, the material displayed 84% retention of its initial capacity after 10 000 cycles, suggesting good cycling stability. In addition, the Ni₆₀Co₄₀ 600//AC hybrid device showed a maximum operational voltage window of 1.6 V with an energy density of 14.9 W h kg⁻¹ at 794.5 W kg⁻¹ power density. Further, the hybrid device retained 86% of its initial capacity after 25 000 cycles, suggesting excellent stability. These results indicate that the Ni₆₀Co₄₀ 600//AC hybrid device is a potential candidate for practical applications. To our knowledge, this is the first detailed report on an electrochemically induced structurally reconstructed pyrophosphate material for supercapacitor applications. We strongly envision that the work unlocks the potential of the reconstruction strategy using transition metal pyrophosphates, which can be extended to other earth-abundant transitional metal-based pyrophosphate materials for energy storage applications.

Author contributions

A. R., H. R. Inta, and V. M. conceptualized the scheme of this work. A. R. carried out all the synthesis, total physical characterization, and entire electrochemical measurements. H. R. Inta, S. Ghosh, and V. M. constantly assisted A. R. during manuscript writing. A. M., H. V. S. R. M. Koppiseti, and B. R. Verma constantly aided A. R. in performing the HRTEM, FESEM, XPS, and ICP-AES studies. S. Bag, S. Ghosh, and A. R. performed the BET study. The whole work was supervised by V. M. together with valuable suggestions. Before submission, all authors have given assent to the final version of the manuscript.

Conflicts of interest

The authors reveal no conflict of interest.

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