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Cite this: *Green Chem.*, 2024, **26**, 2599

Stabilization of Cu⁺ sites by amorphous Al₂O₃ to enhance electrochemical CO₂ reduction to C₂₊ products†

Hailian Cheng,^{a,b} Shuaiqiang Jia,^{*a,b} Jiapeng Jiao,^{a,b} Xiao Chen,^{a,b} Ting Deng,^{a,b} Cheng Xue,^{a,b} Mengke Dong,^{a,b} Jianrong Zeng,^d Chunjun Chen,^{a,b} Haihong Wu,^{†a,b} Mingyuan He^{a,b} and Buxing Han^{†a,b,c}

Catalytic conversion of CO₂ to produce fuels and chemicals is of great importance, and the electrocatalytic CO₂ reduction reaction (eCO₂RR) is considered one of the most attractive pathways. Multi-carbon (C₂₊) products are more desirable in many cases. To date, Cu-based catalysts, especially Cu⁺ sites, have been found to be the most efficient for the production of C₂₊ products. However, the retention of Cu⁺ sites at high cathodic potentials in the eCO₂RR remains a great challenge. In this study, we designed and synthesized CuAl-oxide-derived (Cu_xAl_y-OD, *x* and *y* are the mass percentages of Cu and Al) catalysts for the eCO₂RR to C₂₊ products. During the eCO₂RR process, the predominant Cu species changed to Cu⁺, and the resulting electrocatalyst showed a high faradaic efficiency (FE) of 81.6% for C₂₊ products in alkaline aqueous solutions. Experimental studies showed that the presence of a stable amorphous Al₂O₃ phase stabilized the Cu⁺ sites by oxidation state control, leading to high selectivity and activity for the production of C₂₊ products. This work provides a strategy for improving the stability of Cu⁺ in the catalyst to enhance the performance of the eCO₂RR.

Received 18th November 2023,
Accepted 12th January 2024

DOI: 10.1039/d3gc04492k

rsc.li/greenchem

Introduction

The catalytic conversion of CO₂ to produce fuels and chemicals is of great importance, as it ameliorates both environmental concerns and the energy crisis.^{1–4} The electrocatalytic CO₂ reduction reaction (eCO₂RR), in particular, is considered one of the most attractive pathways for carbon cycle rebalancing due to its mild operating conditions, ease of operation, product adjustment, and potential for synergy with renewable energy sources.^{5–8} To date, Cu-based catalysts have been found to be the most efficient for the production of multi-carbon (C₂₊) hydrocarbons and oxygenated products.^{9–11} However, Cu-

based catalysts still have certain limitations, including poor stability,¹² poor selectivity for specific products, and competition from the hydrogen evolution reaction (HER).¹³

To address these challenges, various strategies have been employed to improve the selectivity, activity, and stability of Cu-based catalysts for the eCO₂RR, including surface treatment of Cu electrodes,^{14–16} alloying,^{17–20} doping of other metal elements to modulate the electronic structure^{21–23} and formation of coordination compounds.^{24–29} It is noteworthy that in order to stabilize the active sites of Cu catalysts, researchers have attempted to use metal–organic frameworks (MOFs) with specially designed ligands to achieve the stabilization of the Cu²⁺ site,^{16,30} but unfortunately, the Cu-MOF catalysts reported so far did not continuously stabilize Cu²⁺ during electrolysis but were electrochemically reduced to Cu⁰.³¹ At the same time, it has also been reported that Cu²⁺ is incorporated into oxide matrices to form Cu–M–O solid solutions in order to stabilize Cu²⁺ during the eCO₂RR process.³² However, most of the target products obtained by the above strategy of stabilizing Cu²⁺ active sites were methane or other monocarbon products. Therefore, the development of catalysts with stabilized Cu⁺ sites for enhancing the dimerization of the reaction intermediate *CO and hence the selectivity of higher value-added C₂₊ products remains challenging.

Here, we prepared a catalyst capable of stabilizing the Cu⁺ site in the eCO₂RR process by introducing the metal Al to form

^aShanghai Key Laboratory of Green Chemistry and Chemical Processes, State Key Laboratory of Petroleum Molecular & Process Engineering, School of Chemistry and Molecular Engineering, East China Normal University, Shanghai, 200062, China. E-mail: sqjia@chem.ecnu.edu.cn, hhlwu@chem.ecnu.edu.cn, hanbx@iccas.ac.cn

^bInstitute of Eco-Chongming, 20 Cuinia Road, Chenjia Town, Chongming District, Shanghai, 202162, China

^cBeijing National Laboratory for Molecular Sciences, CAS Key Laboratory of Colloid and Interface and Thermodynamics, CAS Research/Education Center for Excellence in Molecular Sciences, Center for Carbon Neutral Chemistry, Institute of Chemistry, Chinese Academy of Sciences, Beijing, 100190, China

^dShanghai Synchrotron Radiation Facility, Shanghai Advanced Research Institute, Chinese Academy of Sciences, 201204, China

†Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d3gc04492k>

a $\text{Cu}_x\text{Al}_y\text{-OD}$ nanosheet catalyst. Remarkably, using the $\text{Cu}_{100}\text{Al}_{7.88}\text{-OD}$ catalyst as an electrode in a flow cell with a 1 M KOH electrolyte, we obtained a total current density of 692.6 mA cm^{-2} with a faradaic efficiency (FE) of 81.6% for C_{2+} products, which was twice as high as that of $\text{FE}_{\text{C}_{2+}}$ over a pure CuO-OD catalyst. *In situ* Raman spectroscopy measurements and X-ray absorption spectroscopy (XAS) experiments showed that the broadened orbital in interfacial Al_2O_3 offers a lower orbital for extra electrons than Cu^+ , which can effectively retain nearby Cu^+ , and the highly active Cu^+ sites significantly enhance C–C coupling, improving the selectivity and activity of the eCO_2RR to C_{2+} products.

Results and discussion

The synthesis of $\text{Cu}_x\text{Al}_y\text{-OD}$ electrodes involves a two-step process (Fig. 1a). First, the $\text{Cu}_x\text{Al}_y\text{-Ox}$ catalysts were prepared by the co-precipitation of $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ and $\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$ salts (the details of the synthesis method are provided in the ESI†), and then transferred to a hydrothermal reactor and placed in a 100°C oven for 18 h. Further $\text{Cu}_x\text{Al}_y\text{-OD}$ electrodes were obtained by the redox reaction cyclic voltammetry (CV) technique. The mass percentages of Cu and Al in the as-synthesized $\text{Cu}_x\text{Al}_y\text{-Ox}$ catalysts were determined by inductively coupled plasma optical emission spectroscopy (ICP-OES) (Table S1†). The $\text{Cu}_{100}\text{Al}_{7.88}\text{-Ox}$ catalyst was used as

a representative example. Scanning electron microscopy (SEM) and transmission electron microscopy (TEM) images showed that $\text{Cu}_{100}\text{Al}_{7.88}\text{-Ox}$ exhibited a two-dimensional nanosheet morphology along with a rough edge structure (Fig. 1b and c). According to the ESI Fig. S1† with the increase of Al doping, the morphology of the catalyst changed from a two-dimensional nanosheet structure to a flower-like structure. The high-resolution transmission electron microscopy (HRTEM) image (Fig. 1d) showed that the lattice fringes in the $\text{Cu}_{100}\text{Al}_{7.88}\text{-Ox}$ catalyst were 0.250 nm and 0.274 nm, corresponding to the distance of the (111) and (110) planes of CuO, respectively.^{33,34} It is noteworthy that no lattice fringes belonging to Al were found in HRTEM images, which may be because Al_2O_3 in the catalyst was an amorphous physical phase (Fig. S2 and S3†). Meanwhile, from the energy-dispersive X-ray spectroscopy (EDS) elemental mapping images (Fig. 1e), Cu, Al, and O were uniformly distributed in the $\text{Cu}_{100}\text{Al}_{7.88}\text{-Ox}$ catalyst.

The electronic and chemical structures of the as-synthesized catalysts were characterized by different methods. The X-ray diffraction (XRD) patterns of the Cu_xAl_y catalysts before ($\text{Cu}_x\text{Al}_y\text{-Ox}$) and after CV activation ($\text{Cu}_x\text{Al}_y\text{-OD}$) are shown in Fig. 2a and b. It can be seen that before CV activation, the Cu phase of all catalysts was only CuO (JCPDS#48-1548). However, after CV activation the pure CuO catalyst changed into Cu^0 ,



Fig. 1 Preparation and characterization of catalysts. (a) Schematic illustration of the process to prepare $\text{Cu}_x\text{Al}_y\text{-Ox}$ catalysts. (b) SEM, (c) TEM, (d) HRTEM, and (e) EDS elemental mapping images of the $\text{Cu}_{100}\text{Al}_{7.88}\text{-Ox}$ catalyst.

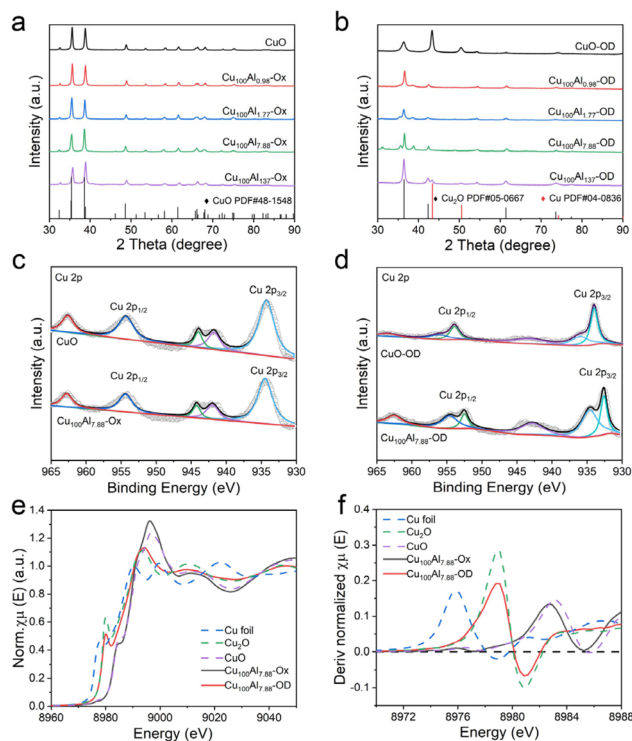


Fig. 2 XRD patterns of CuO and $\text{Cu}_x\text{Al}_y\text{-Ox}$ catalysts (a) before and (b) after CV activation in a CO_2 -saturated 1 M KOH electrolyte. Cu 2p XPS spectra of CuO and $\text{Cu}_{100}\text{Al}_{7.88}\text{-Ox}$ catalysts (c) before CV activation and (d) after CV activation in a CO_2 -saturated 1 M KOH electrolyte. (e) Normalized Cu K-edge XANES spectra and (f) derivative normalized spectra of different catalysts. For comparison, reference spectra from a Cu foil, Cu_2O , and CuO are also shown.

whereas the $\text{Cu}_x\text{Al}_y\text{-OD}$ catalyst was Cu_2O (JCPDS#05-0667). It is noteworthy that we did not find the diffraction peaks of Al in the XRD pattern, probably due to the insufficient crystallinity of Al_2O_3 in the catalysts prepared by this method, which existed as an amorphous phase.³⁵ X-ray photoelectron spectroscopy (XPS) experiments were used to gain insight into the electronic interaction between Cu and Al as well as the surface chemical states of each element in $\text{Cu}_{100}\text{Al}_{7.88}\text{-Ox}$ and $\text{Cu}_{100}\text{Al}_{7.88}\text{-OD}$ catalysts. Fig. 2c shows that the valence state of Cu in both CuO and $\text{Cu}_{100}\text{Al}_{7.88}\text{-Ox}$ catalysts is Cu^{2+} accompanied by strong satellite peaks. After CV activation (Fig. 2d), the CuO electrode was easily reduced to Cu^0 in the catalyst without Al doping, while $\text{Cu}_x\text{Al}_y\text{-OD}$ mostly retained Cu^+ . The XPS spectra of Al 2p showed that Al was present in the physical phase of Al_2O_3 both before and after the activation of the $\text{Cu}_{100}\text{Al}_{7.88}\text{-Ox}$ catalyst (Fig. S4†). Combined with the XRD and XPS results, it can be observed that the CuO catalyst was reduced to metallic Cu^0 , which proved that Cu^{2+} without oxidation state control protection can be easily reduced, further illustrating the importance of amorphous Al_2O_3 in stabilizing the Cu^+ site.

To further investigate the chemical state and electronic structure of Cu in the $\text{Cu}_{100}\text{Al}_{7.88}\text{-Ox}$ and $\text{Cu}_{100}\text{Al}_{7.88}\text{-OD}$ catalysts, synchrotron X-ray absorption spectroscopy (XAS) was performed on the catalysts prepared in this work, and the spectra of Cu foil, Cu_2O and CuO as references were obtained. As shown in Fig. 2e, normalized X-ray absorption near edge structure spectra (XANES) of the Cu K-edge showed that the Cu species in $\text{Cu}_{100}\text{Al}_{7.88}\text{-Ox}$ exhibited an oxidized state (+2), the same as the CuO reference sample, whereas the Cu species in $\text{Cu}_{100}\text{Al}_{7.88}\text{-OD}$ exhibited the same +1 valence state as the Cu_2O reference sample. Moreover, this was further confirmed by the corresponding derivative normalized spectra (Fig. 2f) and the k_3 -weighted Cu K edge Fourier-transform extended X-ray absorption fine structure (FT-EXAFS) spectra (Fig. S5†), which indicated that the presence of amorphous Al_2O_3 could stabilize Cu^+ to prevent its reduction to Cu^0 . This is consistent with the results of the XPS and XRD analyses.

The eCO_2RR catalytic performances of $\text{Cu}_x\text{Al}_y\text{-OD}$ catalysts were tested in a flow cell. 1 M KOH was used as the electrolyte, and the catalysts were loaded onto a gas diffusion layer (GDL). Gaseous products were probed and quantified during the eCO_2RR using gas chromatography (GC) and liquid products were examined after the completion of the electrochemical reaction *via* ^1H nuclear magnetic resonance (NMR). The linear sweep voltammetry (LSV) curves were measured in CO_2 - and Ar-saturated 1 M KOH electrolytes, the $\text{Cu}_{100}\text{Al}_{7.88}\text{-OD}$ electrode exhibited a higher current density in the CO_2 -saturated electrolyte compared to the results in the Ar-saturated electrolyte, indicating the occurrence of an eCO_2RR on the $\text{Cu}_{100}\text{Al}_{7.88}\text{-OD}$ electrode (Fig. 3a). We tested all the $\text{Cu}_x\text{Al}_y\text{-OD}$ and CuO catalysts, including CuO-OD , $\text{Cu}_{100}\text{Al}_{0.98}\text{-OD}$, $\text{Cu}_{100}\text{Al}_{1.77}\text{-OD}$, $\text{Cu}_{100}\text{Al}_{7.88}\text{-OD}$, and $\text{Cu}_{100}\text{Al}_{137}\text{-OD}$, and compared their eCO_2RR performances. As shown in Fig. 3b, gaseous products (H_2 , CO , CH_4 , and C_2H_4) and liquid products (HCOOH , CH_3COOH , $\text{CH}_3\text{CH}_2\text{OH}$ (EtOH), and $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$ (PrOH))

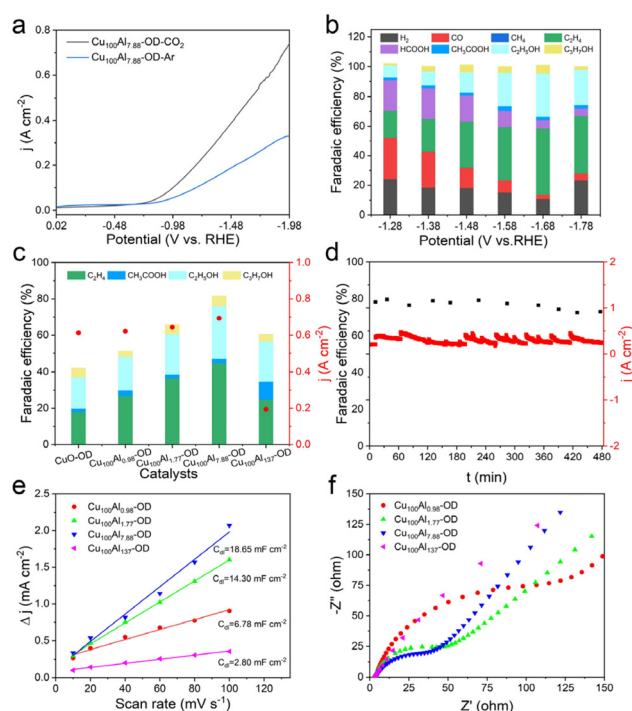


Fig. 3 Electrochemical performances of the $\text{Cu}_x\text{Al}_y\text{-OD}$ catalyst. (a) LSV curves for $\text{Cu}_{100}\text{Al}_{7.88}\text{-OD}$ in CO_2 -saturated and Ar-saturated 1 M KOH electrolytes. (b) FE of the products over $\text{Cu}_{100}\text{Al}_{7.88}\text{-OD}$ at different applied potentials. (c) FE_{C_2+} and total current densities over $\text{Cu}_x\text{Al}_y\text{-OD}$ at -1.68 V vs. RHE. (d) Electrochemical stability test of the $\text{Cu}_{100}\text{Al}_{7.88}\text{-OD}$ electrode at -1.68 V vs. RHE in a flow cell. (e) Comparison of double-layer capacitance in a 1 M KOH electrolyte with $\text{Cu}_x\text{Al}_y\text{-OD}$. (f) Nyquist plots of $\text{Cu}_x\text{Al}_y\text{-OD}$ in the 1 M KOH electrolyte. Data were obtained at ambient temperature and pressure with a CO_2 stream of 20 sccm.

were detected in the present catalytic system. According to the product distribution, the decreasing FE of H_2 and increasing FE of carbon products become obvious after Al doping into Cu catalysts. At a potential of -1.68 V vs. RHE, the FE of C_2+ products over the $\text{Cu}_{100}\text{Al}_{7.88}\text{-OD}$ electrode could reach 81.6% with a current density as high as 692.6 mA cm^{-2} , in which the partial current density of C_2+ products could reach 565.2 mA cm^{-2} (Fig. S6†). The catalyst maintained a high C_2+ yield as the applied potential increased. In contrast, pure CuO catalysts showed a maximum FE_{C_2+} of only 41.8% compared to a potential of -1.68 V vs. RHE, and produced more H_2 and C_1 products, respectively. Fig. 3c shows the FE_{C_2+} and j_{total} over different $\text{Cu}_x\text{Al}_y\text{-OD}$ electrodes at -1.68 V vs. RHE, and it was found that the amount of Al doping was crucial for the selectivity of the C_2+ product. A comparison of the electrocatalytic performances of the catalysts with different Cu/Al mass ratios revealed that the C_2+ product exhibited good FE and partial current density at the $\text{Cu}_{100}\text{Al}_{7.88}\text{-OD}$ electrode (Fig. S7–S10†). In addition to catalytic activity, stability is also one of the important criteria for evaluating catalyst performance. Subsequently, we conducted a stability test in a flow cell (Fig. 3d), in which the current density of the catalyst and FE_{C_2+}

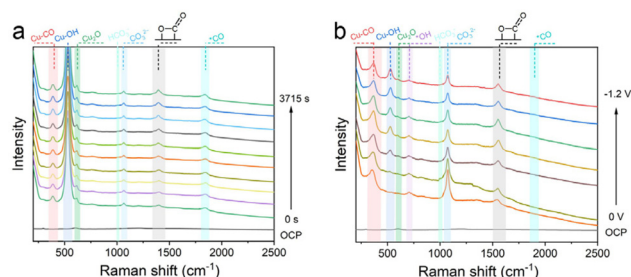


Fig. 4 *In situ* Raman spectra of (a) the Cu₁₀₀Al_{7.88}-OD electrode at -1.68 V vs. RHE with different reaction times during the eCO₂RR in a 0.5 M KOH electrolyte and (b) the Cu₁₀₀Al_{7.88}-OD electrode at various potentials during the eCO₂RR in a 0.5 M KOH electrolyte. All potentials are given vs. RHE.

remained unchanged after 8 h of electrolysis at -1.68 V vs. RHE, indicating that Cu₁₀₀Al_{7.88}-OD had a good stability for the eCO₂RR.

To explore the reasons for the excellent performance of the Cu₁₀₀Al_{7.88}-OD electrode, electrochemically active surface area (ECSA) and electrochemical impedance spectra (EIS) of Cu_xAl_y-OD electrodes were analyzed. The ECSA of the catalyst was then estimated through the electrochemical double-layer capacitance (*C*_{dl}) measurements. The CV curves with various scan rates in the non-faradaic region were recorded. As shown in Fig. 3e, the linear slopes show that the *C*_{dl} value for the Cu₁₀₀Al_{7.88}-OD electrode (18.65 mF cm⁻²) is larger than those of the other catalysts, which demonstrates that doping of a suitable amount of Al can generate more active sites than the pure Cu catalyst. The EIS was also carried out to probe the effect of the metal ratios of the catalysts on the charge transport kinetics at the open circuit potential (OCP). It showed that the charge transfer resistance of the Cu₁₀₀Al_{7.88}-OD electrode was lower than that of other catalysts, indicating a favorable kinetics toward the eCO₂RR (Fig. 3f).³⁶

To gain insight into characterizing both the structural evolution of catalyst surfaces and reaction intermediates at electrode/electrolyte interfaces during electrochemical reactions, *in situ* Raman spectroscopy was used to monitor the change of interfacial species and highly active sites of the Cu₁₀₀Al_{7.88}-OD electrode under different electrolysis times and potentials (Fig. 4). As shown in Fig. 4a, after prolonged electrolysis, the Cu₁₀₀Al_{7.88}-OD electrode showed a Raman peak at 620 cm⁻¹, which was attributed to Cu₂O.³⁷ This absorption peak corresponds to the peak at this location under OCP conditions and remains stable for more than 3000 s of continuous electrolysis, indicating that the Cu⁺ site is stable during the eCO₂RR. The *in situ* Raman results demonstrate that the broadened orbital in interfacial amorphous Al₂O₃ offers a lower orbital for extra electrons than Cu⁺, which can effectively retain nearby Cu⁺.^{38,39} It is worth noting that the catalyst was not exposed to ambient air during the Raman test, so the Cu⁺ site was generated after CV activation and remained stable during the eCO₂RR process. The retention of the band at around 530 cm⁻¹ for Cu-OH can be explained by the increase of the

local pH near the catalyst surface at high reaction potentials (-1.68 V vs. RHE).⁴⁰ Interestingly, the Raman peaks at 1330 cm⁻¹ belong to the CO₂⁻ in-stretching vibrations, which confirms the first intermediate during the eCO₂RR over the Cu₁₀₀Al_{7.88}-OD electrode.⁴¹ The peaks around 1880 cm⁻¹ are assigned to the C≡O stretching vibrations. Further investigation of the mechanism of the C₂₊ products from the Cu₁₀₀Al_{7.88}-OD electrode is very interesting. In particular, the peaks located in the range of 1800–2000 cm⁻¹ are associated with the bridge site of *CO absorption (*CO_{bridge}), which can favor *CO–CO coupling.⁴²

The *in situ* Raman spectra of the Cu₁₀₀Al_{7.88}-OD electrode at various potentials are shown in Fig. 4b; a Raman peak at 1012 cm⁻¹ (representing a small amount of bicarbonate) and a Raman peak at 1068 cm⁻¹ (representing carbonate) were detected in the Cu₁₀₀Al_{7.88}-OD electrode. According to previous studies, the density of localized proton donors near the catalyst surface can be estimated from the ratio of HCO₃⁻/CO₃²⁻.⁴³ As the applied potential decreases, the HCO₃⁻/CO₃²⁻ ratio increases, indicating a decrease in proton donors near the Cu₁₀₀Al_{7.88}-OD electrode surface. These alkaline conditions promote the formation of C₂₊ products, which is consistent with our electrochemical results (Fig. 3c).^{44,45}

Conclusions

In summary, we have proposed a method to stabilize the Cu⁺ oxidation state with a suitable amount of Al₂O₃, and the catalyst can promote CO₂ electroreduction to C₂₊ products with high catalytic activity, selectivity, and stability. Combined *in situ* spectroscopic and electrochemical measurements indicate that Cu⁺ stabilized by the amorphous Al₂O₃ enhances the adsorption stability of the *CO intermediate, which in turn promotes its further dimerization (C–C coupling) and thus significantly improves the selectivity for C₂₊ products. This discovery effectively solves the stability problem of Cu-based catalysts for the eCO₂RR and provides an effective way to design low-cost, energy-efficient catalysts.

Author contributions

H. L. C., S. Q. J., H. H. W. and B. X. H. proposed the project, designed the experiments, and wrote the manuscript. H. L. C. and S. Q. J. performed most of the experiments and analyzed the experimental data. J. P. J., X. C., T. D., C. X. M. K. D. and C. J. C. provided help in materials synthesis and characterization. H. L. C., S. Q. J., X. C. and J. R. Z. conducted XAS measurements and analyzed the results. H. H. W., M. Y. H. and B. X. H. co-supervised the whole project. All authors discussed the results and commented on the manuscript.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The work was supported by the National Key R&D Program of China (2023YFA1507901, 2020YFA0710201), the National Natural Science Foundation of China (22003070, 22293015, 22121002), the China Postdoctoral Science Foundation (2023M731096), and the Research Funds of Happiness Flower ECNU (2020ST2203). We thank the staff of beamline BL13SSW at Shanghai Synchrotron Radiation Facility for the support for the experiments of *in situ* XAS measurements.

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