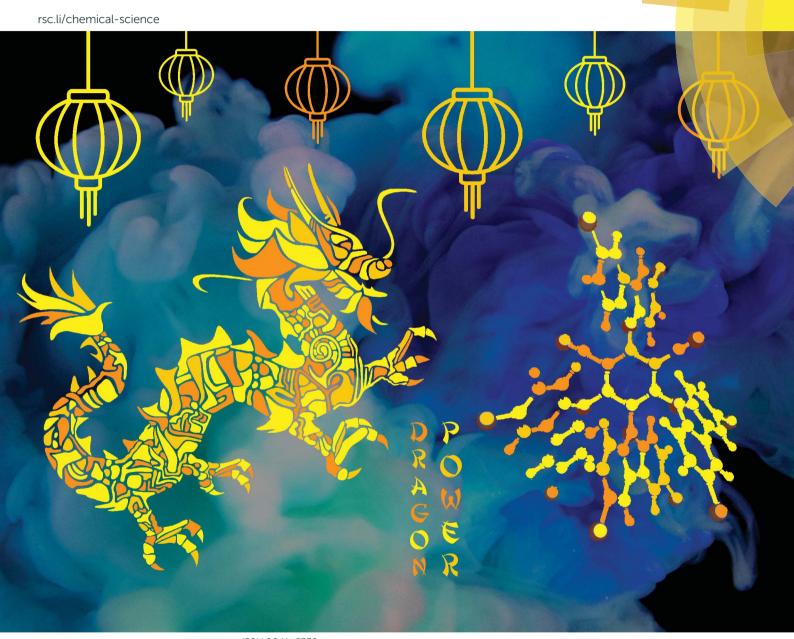
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## Fine-tuned organic photoredox catalysts for fragmentation-alkynylation cascades of cyclic oxime ethers†

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Fine-tuned organic photoredox catalysts are introduced for the metal-free alkynylation of alkylnitrile radicals generated *via* oxidative ring opening of cyclic alkylketone oxime ethers. The redox properties of the dyes were determined by both cyclic voltammetry and computation and covered an existing gap in the oxidation potential of photoredox organocatalysts.

#### 1. Introduction

Cascade reactions leading to several bond forming/bond breaking events are important tools to access efficiently the complex carbon skeleton of organic compounds.¹ Single electron chemistry is well suited for cascade transformations due to the high reactivity of radical intermediates combined with neutral reaction conditions.² However, many radical-based cascade reactions still require the stoichiometric use of toxic reagents, such as tin hydrides and/or harsh reaction conditions. In the last decade, visible light driven photoredox catalysis has emerged as a general method for radical generation under mild oxidative or reductive conditions without the need for toxic reagents.³ First successes were met with metal complexes as catalysts.⁴ More recently organic dyes have been introduced, as they are often cheap, non-toxic and easy to modify.⁵

In the case of reductive quenching cycles for the generation of radical by oxidation, the reduction potential of the photoactivated catalyst is an essential parameter (Fig. 1). The reduction potential of the most often used iridium catalyst (Ir[dF(CF<sub>3</sub>) ppy]<sub>2</sub>(dtbbpy))PF<sub>6</sub> (1) (+1.21 V  $\nu s$ . SCE in MeCN)<sup>6</sup> is lower than frequently used organic dyes, such as Fukuzumi dye (2) and dicyanoanthracene (3) (+2.06 V  $\nu s$ . SCE in MeCN).<sup>5,7</sup> To enrich the range of transformations available to organic catalysts, using dyes with lower oxidation values would be highly desirable.

In this context, 2,4,5,6-tetra(9*H*-carbazol-9-yl)isophthalonitrile (4CzIPN, **4a**) has emerged as an alternative organic dye with a lower reduction potential (+1.35 V  $\nu s$ . SCE in MeCN).8 Nevertheless, there is a need for further catalysts with reduction

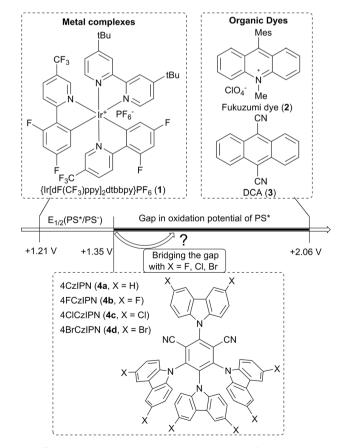


Fig. 1 Bridging the gap in the reduction potential of activated organic photocatalysts.

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potentials between +1.4 V and + 2.0 V to allow challenging transformations asking for efficient oxidation.

In this context, heteroatom centered radicals initiated transpositions have found widespread applications in synthetic chemistry.9 Zard and coworkers developed iminyl radical initiated cascades using tin reagents and electrophilic olefins to trap the generated intermediate (Scheme 1A).10 Transition metal catalysis (Pd, Ir, Fe) and/or heating at temperature higher than 90 °C were later reported to avoid the use of tin reagents. 11 In 2017, Shi and co-workers described one single example of a Heck-like coupling at room temperature under photoredox conditions (Scheme 1B).12 After Shi's pioneering example, Xiao and coworkers developed in 2018 a general method for C-C bond formation after reductive cleavage of cycloalkyloxime esters for the introduction of both alkanes and alkenes.13 Zhou and coworkers reported a multicomponent alkylation etherification of alkenes using a similar strategy.14 These breakthroughs greatly enhanced the synthetic potential of fragmentation cascades, but the exclusive use of reductive conditions limited the types of bond formations possible.

Recently, Studer and coworkers<sup>15</sup> and Leonori and coworkers<sup>16</sup> independently reported a new approach for the generation of iminyl radicals based on an oxidative quenching cycle starting from carboxylic acid derived oxime ethers. In 2018, this activation strategy was used by Leonori to develop the fluorination, chlorination and azidation of alkylnitrile radicals generated by fragmentation of iminyl radicals (Scheme 1B).<sup>17</sup>

A) C-C bond formation after ring opening via iminyl radical Bu<sub>3</sub>SnH/AIBN or TM, heat, µwave Zard, Uemura, Selander, bond Shi, Guo, Castle 10-11 formation

B) Iminyl radical generation and ring opening via photoredox catalysis

C) This work: oxidative photoredox catalysis for C-C bond formation

Scheme 1 Existing methods for C-C bonds formation after iminyl radical initiated ring opening (A and B) and our work based on oxidative photoredox catalysis (C) (PC = photoredox catalysis, EBX = ethynyl benziodoxolone)

up to 99% yield

However, no example of C-C bond formation was reported. In this respect, our group and others have shown that ethynylbenziodoxolones (EBX) reagents 5 allowed the alkynylation of radicals generated through oxidative photoredox cycles. 18 We therefore envisioned that EBX reagents would be ideally suited overcome these limitations and deliver remotelyfunctionalized alkynyl nitrile through an oxidative fragmentation C-C bond formation sequence. Such a transformation has never been reported under photoredox conditions.19 However, preliminary results showed that only low yields could be obtained in this reaction using established photoredox catalysts 1-3 (see Results and Discussions).

Herein, we report for the first time the use of the modified 4XCzIPN dyes 4b-d in photoredox catalysis. In particular, 4ClCzIPN dve 4c led to a highly efficient fragmentation alkynylation process. Both theory and experiment showed that 4c has an increased reduction potential compared to 4a, leading to the efficient fragmentation of both four- and five-membered cyclic oxime ethers to give compounds containing versatile alkyne and nitrile functionalities in good yield under mild conditions.

#### Results and Discussions

#### Dye synthesis and properties

The catalysts 4a-d were easily synthesized from commercially available 2,4,5,6-tetrafluoroisophthalonitrile (6) and the corresponding carbazoles 7a-d via nucleophilic aromatic substitution (Scheme 2).20 Good yields were obtained, except for fluorosubstituted dye 4b.

The redox properties of catalysts 4a-d were then determined (Table 1). Ground state reduction and oxidation potentials of -1.21 V and +1.52 V for catalyst 4a in acetonitrile have been reported based on cyclic voltammetry.84 These values combined with the difference in absorbance and emission signals of the dye allowed to estimate the potentials in the excited state: +1.35 V and -1.04 V for reduction and oxidation, respectively (entry 1). In our hand, slightly different values were obtained (entry 2). In particular, a higher reduction potential of +1.59 V

Scheme 2 Synthesis of 4XCzIPN dves 4a-d

Table 1 Literature (lit), measured (mes) and computed (comp) values for reduction potentials of dyes  $4a-d^a$ 

Entry	Catalyst	Solvent	$E_{1/2}\!\!\left(\mathbf{P}/\mathbf{P}^-\right)$	$E_{1/2}(P^+/P)$	$E_{0-0}$	$E_{1/2}\!\!\left(\mathbf{P^*}\!/\mathbf{P^-}\right)$	$E_{1/2}(\mathbf{P}^+/\mathbf{P}^*)$
1	<b>4a</b> (lit)	CH <sub>3</sub> CN	-1.21	+1.52	2.56	+1.35	-1.04
2	4a(mes)	CH <sub>3</sub> CN	-1.05	+1.68	2.64	+1.59	-0.96
3	<b>4c</b> (mes)	CH <sub>3</sub> CN	-0.97	+2.05	2.68	+1.71	-0.63
4	4a(comp)	CH <sub>3</sub> CN	-1.29	+1.56	$2.64^{b}$	+1.35	-1.08
5	<b>4b</b> (comp)	CH <sub>3</sub> CN	-1.18	+1.67	_	_	_
6	<b>4c</b> (comp)	CH <sub>3</sub> CN	-1.10	+1.76	$2.68^{b}$	+1.58	-0.92
7	<b>4d</b> (comp)	CH <sub>3</sub> CN	-0.83	+1.89	$2.65^{b}$	+1.82	-0.76
8	4a(comp)	$CH_2Cl_2$	-1.29	+1.67	$2.59^{b}$	+1.30	-0.92
9	<b>4b</b> (comp)	$CH_2Cl_2$	-1.18	+1.79	$2.60^{b}$	+1.42	-0.81
10	<b>4c</b> (comp)	$CH_2Cl_2$	-1.10	+1.87	$2.59^{b}$	+1.49	-0.72
11	<b>4d</b> (comp)	$\mathrm{CH_2Cl_2}$	-0.85	+1.98	$2.58^{b}$	+1.73	-0.60

<sup>&</sup>lt;sup>a</sup> Potentials in V vs. SCE. The excitation energy  $E_{0-0}$  was estimated by the point of intersection of the normalized absorbance and emission signals.  $E_{1/2}(P^+/P^*) = E_{1/2}(P^+/P) - E_{0-0}$  and  $E_{1/2}(P^*/P^-) = E_{0-0} + E_{1/2}(P/P^-)$ . See ESI for details. <sup>b</sup> Experimental values of  $E_{0-0}$  were used.

was observed in the excited state. We then turned to halogenated dyes. With catalyst 4c, both the cationic and anodic shifts were measured, with reduction and oxidation potentials of -0.97 V and +2.05 V (entry 3). This resulted in an increased reduction potential of +1.71 V of the photoexcited dye. Unfortunately, the cyclic voltammograms of dyes 4b and 4d could not be determined in acetonitrile due to their limited solubility. Only the absorption/emission spectra of dye 4d could be measured. We therefore turned to theory to have a more reproducible and expanded access to redox potential values. At the PCM-UAKS/PBE0-D3BJ/def2-SVP level, the ground state reduction potentials decrease and the oxidation potentials increase in the order of 4a, 4b, 4c, 4d (entries 4-7).21 This leads to a higher reduction potential in the excited state for catalyst 4c and 4d compared to 4a (+1.58 V and + 1.82 V respectively compared to + 1.35 V). Therefore, both measurement and computation confirmed the increased potential for 4c. The same trends but with lower reduction potentials were obtained in dichloromethane as solvent (entries 8-11). In this case, absorption and emission spectra could be measured for all dyes, but no good quality cyclic voltammograms could be acquired. Values in the ground state were therefore again obtained by computation.

The trend in reduction potentials increasing in the order of 4a < 4b < 4c < 4d (substituents on catalyst 4: H < F < Cl < Br) can be rationalized as follows. Upon reduction, an electron is added to the lowest unoccupied molecular orbital (LUMO), located mostly on the central isophthalonitrile ring (Fig. 2). However, the LUMO also involves the carbazole moieties in the 4 and 6 positions of the isophthalonitrile ring and is potentially stabilized by them to a greater extent in the case of Cl and Br substituents (resonance donors) compared to H and F. We also note that while molecules 4a and 4b are fairly symmetric, systems 4c and 4d feature noticeable distortion of the 4-carbazole (Fig. 2). This peculiar structure, observed both in the gasphase optimized geometries, obtained at several dispersioncorrected DFT levels, and experimental crystal structures,22 is due to strong Cl···Cl and Br···Br contacts in 4c and 4d, respectively.23

#### Optimization of the fragmentation cascade

Our investigations began with the gram scale synthesis of hydroxylamine **10** (Scheme 3, eqn (1)).<sup>24</sup> Nucleophilic substitution of commercially available *N*-hydroxyphthalimide (**8**) and ethyl 2-bromo-2-methylpropanoate (**9**), followed by acidic hydrolysis of the phthalimide, gave 5.8 g of **10** in 86% yield over 2 steps. Condensation of **10** with cyclobutanone (**11**) led to oxime ether **12a** in 90% yield as a crystalline solid (Scheme 3, eqn (2)).

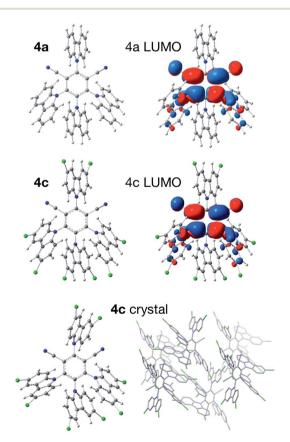


Fig. 2 Optimized structures and LUMO plots (isovalue 0.02) of the dyes 4a and 4c at the PBE0/def2-SVP level, as well as the crystal structure of 4c. See the SI for the plots of 4b and 4d.

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Scheme 3 Gram scale synthesis of 1-carboxy-1-methylethoxyammonium chloride (10), and condensation with cyclobutanone (11) to give model substrate 12a.

A cyclic voltammetry measurement on the corresponding potassium carboxylate 12a' showed a clear oxidation peak at +1.48 V vs. SCE in DMF. We decided then to first test the best catalysts reported by Studer: <sup>15</sup> (Ir[dF(CF<sub>3</sub>)ppy]<sub>2</sub>(dtbbpy))PF<sub>6</sub> (1) and Leonori:16,17 Fukuzumi's catalyst 2 (Table 2). Based on our previous studies on photoredox catalysis using EBX reagents, 18a,b we used DCE as solvent and an excess of Ph-EBX reagent 5a (2.0 equiv.). With 1 mol% of 1a, full conversion

Table 2 Optimization of the reaction conditions

Entry	Catalyst	Base	Time	Conversion <sup>a</sup>	Yield <sup>b</sup>
$1^{c,d}$	1	Cs <sub>2</sub> CO <sub>3</sub>	2 h	>95%	55%
$2^d$	2	Cs <sub>2</sub> CO <sub>3</sub>	2 h	50%	20%
$3^d$	3	Cs <sub>2</sub> CO <sub>3</sub>	2 h	90%	5%
4	4a	$Cs_2CO_3$	2 h-6 h	60%	50%
5	4c	$Cs_2CO_3$	2 h	>95%	70%
6	4c	K <sub>2</sub> CO <sub>3</sub>	2 h	>95%	75%
7	4c	K <sub>2</sub> CO <sub>3</sub>	1 h	>95%	80%
8	4d	$K_2CO_3$	1 h	>95%	75%
9	4b	K <sub>2</sub> CO <sub>3</sub>	1 h	>95%	75%
$10^e$	4c	K <sub>2</sub> CO <sub>3</sub>	1 h	>95%	70%
$11^f$	4c	K <sub>2</sub> CO <sub>3</sub>	1 h	>95%	80%
12	4c	K <sub>2</sub> CO <sub>3</sub>	0.5 h	>95%	80%
13	_	$K_2CO_3$	1 h	NR	NR
14	4c	_	1 h	<10%	5%
$15^g$	4c	$K_2CO_3$	1 h	NR	NR

<sup>&</sup>lt;sup>a</sup> Reaction conditions: using 0.1 mmol 12a (1 equiv.), 0.2 mmol 5a (2.0 equiv.), 5 mol% 1-4 (0.05 equiv.) in DCE (2.0 mL) for 2 h at RT. The conversion of 12a by NMR is given. NR = no reactivity. b NMR yield using dibromomethane as internal standard. <sup>c</sup> Using 1 mol% of 1. <sup>d</sup> Using 1.0 equiv. of base. <sup>e</sup> Using 1.5 equiv. of 5a. <sup>f</sup> Using 3.0 mol% of organic dye 4c. g Without irradiation.

and 55% NMR yield of alkyne 13a were reached after 2 h, whereas only 50% conversion and 20% NMR vield was obtained with 2 (Table 2, entries 1-2). In both cases, significant side products formation was observed. DiCyanoAnthracene (DCA, 3) gave only 5% NMR yield of 13a (entry 3). With 4CzIPN (4a) we were pleased to observe 60% conversion of 12a and 50% NMR yield of 13a after 2 h (entry 4).

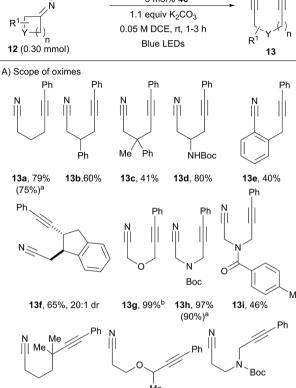
Increasing reaction time (6 h) and/or temperature (50 °C) was not successful to improve conversion. Speculating that stronger oxidizing dyes could lead to better results, catalyst 4c was then examined. Full conversion of 12a was observed and 70% of 13a could be isolated (entry 5). K<sub>2</sub>CO<sub>3</sub> was found to be slightly better than Cs<sub>2</sub>CO<sub>3</sub> as a base (entry 6), allowing us to reduce the reaction time to 1 h (entry 7). The other new photocatalysts 4b and 4d were then tested, and provided similar photocatalytic activity as 4c (entries 8-9). Due to its better solubility and ease of synthesis, 4c was selected as photocatalyst to continue the study. Decreasing the number of equivalents of 5a to 1.5 resulted in lower yield and more side reactions (entry 10). Interestingly, the catalyst loading could be decreased to 3 mol% (entry 11) or the reaction time reduced to only 30 min (entry 12) without any effect on the yield. Finally, control experiments showed that photocatalyst, base and light were all required for the reaction to proceed (entries 13-15).

#### Scope of the fragmentation cascade

We then investigated the scope of oximes at 0.30 mmol scale using 3 mol% of 4c (Scheme 4A). Oxime 12a afforded 13a in 79% isolated yield. Phenyl, alkyl and protected amines were tolerated at the β-position (13b-d, 41-80% yield). Benzocyclobutenone oxime 12e afforded propargylic benzocarbonitrile 13e in 40% yield. Bicyclic compound 13f was obtained in 65% yield and excellent diastereoselectivity. Oxetanone and azetidinones oximes ethers were also excellent substrates (13g-i). Cyclopentanones derivatives were also successful. Using a gemdimethyl substituent for the generation of a δ-tertiary alkylnitrile radical allowed the isolation of 13j in 77% yield. In the work of Leonori and coworkers, a benzylic position was used to promote ring opening.17 In our case, an α-heteroatom was exploited. With an oxygen linker, product 13k was obtained in 65% yield, whereas a nitrogen group led to a modest yield (30%) of 13l. In addition, compounds 13a and 13h were obtained in similar yields using sunlight instead of LEDs.

We then turned our attention to the scope of the reagents (Scheme 4B). The conversion of azetidinone oxime ether 12h into  $\alpha$ - $\alpha$ -cyanoalkynylamines was achieved in good yields, tolerating electron-withdrawing groups (13m-o) and electrondonating groups (13p) in para position of the benzene ring.<sup>25</sup> Good yields were also obtained with a fluorine in meta or a bromine in ortho positions (13q and 13r). Cyclobutanone oxime 12a could also be used with fluorinated arene substituents on the EBX (13s-t). Products 13u and 13v bearing a trifluoromethyl and bromine at the para-position were obtained in moderate yields from cyclopentanone oxime 12j. Alkynyl nitriles are important building blocks combining two highly useful

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13j, 77% 13k 65% 13I, 30% B) Scope of EBX reagents n-C<sub>5</sub>H<sub>11</sub>

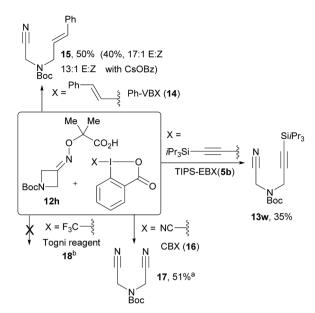
Scheme 4 Scope of the reaction. Using oximes 12 (1.0 equiv.), photocatalyst 4c (3 mol%), EBX 5 (2 equiv.) and K2CO3 (1.1 equiv.) in DCE (0.05 M) at rt, 1-3 h. aUsing sunlight instead of blue LEDs. NMR yield is given. bOn 0.10 mmol scale

13t, 76%

Me

13v. 34%

functional groups.<sup>26</sup> For example, compounds such as 13a have been used in cyclization reactions with alkynes, 26a-c vinyl iodonium salts,26d or benzyne equivalents.26e



Scheme 5 Extension to other C-C bond forming reactions. Reaction conditions: 5 mol% 4c, K2CO3 (1.1 equiv.), 0.05 M DCE, blue LEDs, rt, 1 h. <sup>a</sup>Reaction run of reagent 16 (2 equiv.) for 14 h. NMR vield is given (isolated yield: 31%) breaction run of reagent 18 (2 equiv.) for 24 h.

Preliminary investigations were then conducted for other types of C-C bond formation (Scheme 5). A silyl EBX reagent gave a lower yield of product 13w and alkyl EBX could not be used. Preliminary results for an alkenylation cascade were also obtained using Phenyl Vinyl BenziodoXolone (PhVBX, 14), recently introduced by Olofsson and co-workers.27 Under the optimized reaction conditions, alkene 15 was isolated in 50% yield with a 12:1 E/Z selectivity.<sup>28</sup> Changing the base from K<sub>2</sub>CO<sub>3</sub> to CsOBz afforded lower yield (40%), but with enhanced selectivity (17:1 E/Z). Using two equivalents of cyanobenziodoxolone (CBX, 16),18b bisnitrile 17 was obtained in 51% yield. With Togni reagent 18,29 no desired product was formed.

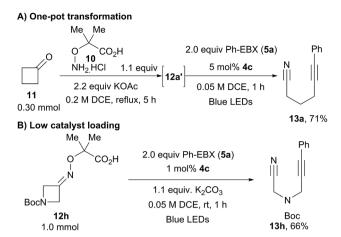
A one-pot protocol starting directly from the ketones was then developed (Scheme 6A). After condensation of cyclobutanone (11) with hydroxylamine 10, addition of two equivalents of Ph-EBX (5a) and 5 mol% of the organic dye followed by one hour of irradiation delivered alkyne 13a in 71% yield. A frequent drawback of organic dyes is their limited stability, leading often to low turnover numbers. Nevertheless, 66% of 13h could still be isolated when using only 1 mol% of 4c at a 1.0 mmol scale (Scheme 6B).

According to previous studies on EBX reagents, 18a,b and the redox potential of dye 4c, we assume that the reaction starts with the oxidation of the potassium carboxylate 12a' ( $E_{1/2}$  = +1.48 V vs. SCE in DMF) by the excited state PS\* of the organic dye to generate carboxyl radical I and the reduced state PS<sup>-</sup> (Scheme 7). Fast decarboxylation releases the  $\alpha$ -oxy radical II, which can either lead to iminyl radical III after acetone extrusion, or can be directly trapped by EBX reagents forming 19. Although 19 was not detected, its hydrated form 20 was observed by 1H NMR. The iminyl radical III then fragments into a nucleophilic alkyl nitrile radical IV, which reacts with EBX 5.

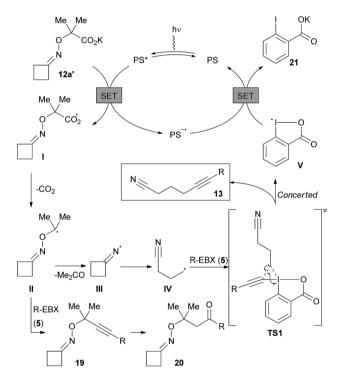
13r. 48%

13s, 71%

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Scheme 6 One-pot protocol and low catalyst loading



Scheme 7 Speculative reaction mechanism

For this step, a concerted mechanism via transition state TS1 can be envisioned according to our previous studies<sup>18b</sup> to afford the desired alkynylnitrile and radical V. A  $\beta$ -addition followed by a 1,2 shift of the EBX substituent could also be considered depending on the migratory aptitude.<sup>18b</sup> Final reduction of V allows the regeneration of the ground state photocatalyst PS and generates carboxylate 21.

#### Conclusions

In summary, we have described the fine-tuning of organic dyes for the synthesis of alkynyl nitriles based on the alkynylation of alkylnitrile radicals generated after a visible light-driven oxidative ring fragmentation reaction involving an iminyl radical as intermediate. Preliminary results for the corresponding alkenylation and cyanation were also achieved. High yields were obtained in a short reaction time with low catalyst loading. We envision a bright future for these new dyes and further applications of cascade approaches in challenging C–C bond forming reactions.

#### Conflicts of interest

There are no conflicts to declare.

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