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[View Journal](#) | [View Issue](#)Cite this: *J. Mater. Chem. A*, 2021, 9, 19865Controlling the Co–S coordination environment in Co-doped WS₂ nanosheets for electrochemical oxygen reduction†Wei Hong,  ‡ Erika Meza  ‡ and Christina W. Li  *

Cobalt sulfide nanomaterials are among the most active and stable catalysts for the electrocatalytic oxygen reduction reaction in pH 7 electrolyte. However, due to the complexity and dynamism of the catalytic surfaces in cobalt sulfide bulk materials, it is challenging to identify and tune the active site structure in order to achieve low overpotential oxygen reduction reactivity. In this work, we synthesize isolated Co sites supported on colloidal WS₂ nanosheets and develop a synthetic strategy to rationally control the first-shell coordination environment surrounding the adsorbed Co active sites. By studying Co–WS₂ materials with a range of Co–S coordination numbers, we are able to identify the optimal active site for pH 7 oxygen reduction catalysis, which comprises cobalt atoms bound to the WS₂ support with a Co–S coordination number of 3–4. The optimized Co–WS₂ material exhibits an oxygen reduction onset potential of 0.798 V vs. RHE, which is comparable to the most active bulk phases of cobalt sulfide in neutral electrolyte conditions.

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Introduction

Recent interest in coupling fuel cells and electrolyzers to biological systems has led to an effort to develop electrocatalysts that operate efficiently under biologically-compatible conditions: pH 7 aqueous buffer and ambient pressure and temperature.^{1–9} The oxygen reduction reaction (ORR) in neutral media presents catalytic challenges that are distinct from those present in acidic and alkaline electrolytes. The state-of-the-art catalysts in acidic and alkaline media, supported noble metal nanoparticles and metal oxide nanostructures, are susceptible to deactivation in neutral electrolyte due to surface poisoning and catalyst dissolution, respectively.^{10–12}

First-row transition metal sulfides have emerged as alternatives to Pt-based catalysts in ORR under a wide range of pH conditions due to their high activity, stability, and low cost.^{9,13–18} Literature studies have focused primarily on tuning the structure of nickel and cobalt sulfides to understand the role that metal sulfide phase and composition play in dictating ORR reactivity.^{19–25} The Co₉S₈ phase, in particular, has been identified as uniquely active amongst the first-row transition metal sulfides for alkaline and neutral ORR and has been studied in a wide range of composite nanostructures.^{26–33} Computational

studies have postulated that the M–S coordination environment at the catalytic surface influences oxygen adsorbate binding energies and thus ORR catalytic turnover.^{34,35} However, when experimentally altering the crystal structure or composition of a bulk metal sulfide phase, it is challenging to isolate the role that coordination environment plays amidst the multiple geometries and oxidation states that exist within any given phase as well as the dynamic nature of the surface under electrocatalytic conditions.^{35,36}

Supported single atom catalysts (SACs) are intriguing model systems for catalytic surface sites on bulk materials because their active site structure can be studied at the atomic level.^{37–39} Previous work on single metal atoms supported on metal oxide and heteroatom-doped carbon materials has shown the importance of local coordination environment and metal–support interaction in influencing catalyst selectivity and reactivity.^{40–44} Particular effort has been invested in active site characterization for metal single atoms supported on N-doped carbon.^{45–49} Recent work on Co SACs have postulated that the number and chemical nature of nitrogen atoms in the first coordination shell may have an impact on catalytic activity in the alkaline oxygen reduction reaction.^{50,51} While different single atom coordination environments have been observed in these examples, it remains challenging to synthetically control the coordination environment surrounding a single atom catalyst. Single atoms supported on MoS₂ have also been studied extensively, primarily for use as electrocatalysts in the hydrogen evolution reaction (HER).⁵² However, doped first-row transition metal single atoms in MoS₂ are not themselves catalytically active sites for HER but rather serve to tune the

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electronic and catalytic properties of the MoS₂ surface.^{53–55} In this work, we put forward a synthetic strategy to generate cobalt single atoms supported on WS₂ nanosheets, in which the Co–S coordination environment can be explicitly tuned through controlled introduction of excess sulfur. These Co–WS₂ nanosheets provide the basis for a systematic study on how the local coordination environment of surface cobalt sulfide active sites influence oxygen reduction reactivity in neutral electrolyte.

Experimental methods

Synthesis of colloidal Co-doped WS₂ nanosheets

Colloidal 1T-WS₂ nanosheets capped with oleylamine were synthesized based on a previously reported method.^{56,57} To the dried WS₂ nanosheets (26.2 mg, 0.007 mmol), a 0.1 M solution of *n*-BuLi in hexanes (10 mL) was added under inert atmosphere and stirred for 2 hours. The Li-treated nanosheets (Li-WS₂) were then rinsed with excess hexanes and redispersed in NMF (1 mL). To a 1 mL solution of Li-WS₂, a Co precursor solution (30 μ L or 100 μ L of 70 mM CoCl₂ in NMF) was then added and stirred for 24 hours at room temperature under nitrogen. After the Co functionalization step, the nanosheets were precipitated with ethanol (10 mL) and hexanes (20 mL) followed by centrifugation at 8700 rpm for 10 minutes. The supernatant was decanted and the cleaning step was repeated one time.

K₂S impregnation and annealing treatment for Co-WS₂

The colloidal Co-WS₂ nanosheets were first supported on carbon black by mixing 1 mL of a 7 mM solution of Co-WS₂ nanosheets with 11.6 mg carbon in 4 mL ethanol at room temperature. The supported Co-WS₂ samples were centrifuged at 8700 rpm for 10 min and resuspended in 2 mL of ethanol : IPA (v/v; 1 : 3) solution. Next, varying amounts of K₂S solution (70 mM K₂S in EtOH) were added, ranging from 0.5 eq. to 10 eq. with respect to Co. The mixture was left stirring in air at room temperature for 24–36 hours until the solvent fully evaporated. The dried powder was then annealed at 300 °C for 2 hours under N₂. After annealing, the sample was successively rinsed with ethanol, water, and IPA to remove unincorporated K₂S.

Scanning/transmission electron microscopy

Low-resolution TEM images were acquired using an FEI Tecnai T20 TEM equipped with a 200 kV LaB₆ filament. Low-resolution HAADF-STEM imaging and EDS mapping were obtained on an FEI Talos F200X S/TEM with a 200 kV X-FEG field-emission source and a super X-EDS system. High-resolution HAADF-STEM images were collected using a Thermo Scientific Themis Z, a spherical aberration corrected S/TEM with a 300 kV X-FEG field-emission source. Associated EDS spectra were obtained with a quad-silicon FEI Super X drift detector.

In order to obtain atomic-resolution HAADF-STEM images, it is critical to remove all organic ligand and solvent residue from the TEM sample. Colloidal Co-WS₂ samples for HR-STEM were cleaned four additional times with ethanol and hexanes prior to drop casting onto a Au grid coated with an ultrathin carbon

film. The Au grid was then submerged in a suspension of carbon black in hexanes, and N₂ gas was bubbled into the solution for 1 min to agitate the suspension.⁵⁸ The TEM grid was air-dried and stored under inert atmosphere prior to imaging.

X-ray absorption spectroscopy

X-ray absorption spectroscopy experiments were carried out at the 10-ID and 10-BM beamlines at the Advanced Photon Source at Argonne National Laboratory. Samples were pressed into a self-supporting pellet, and spectra were obtained in transmission mode at room temperature. The Demeter software package was used to analyze the collected data.⁵⁹ Data was collected at the Co K-edge (7.7089 keV) using metallic Co foil for energy calibration. EXAFS coordination parameters were obtained by a least-squares fit in *R*-space of the *k*²-weighted Fourier transform data from 2.5 to 8.5 Å^{−1} for Co-WS₂ samples and 2.5 to 9.0 Å^{−1} for control samples. A Co foil reference sample was first fit to its known crystallographic parameters to obtain an amplitude reduction factor (*S*₀²) for the Co K-edge. EXAFS fitting of the first coordination shell was carried out between 1.1 and 2.3 Å in *R*-space. Fittings were done by refining bond distances (*R*), coordination numbers (CN) and energy shift (*E*₀). The Debye–Waller factor (σ^2) was kept constant for each sample.

Electrochemical measurements

Electrochemical experiments were conducted on a Pine Wave-Driver 20 bipotentiostat. A catalyst ink was obtained by sonicating the catalyst powder in a solution containing 75.6% water, 24% IPA, and 0.4% Nafion (v/v) to obtain a nominal concentration of 7 mM based on Co-WS₂. The working electrode was prepared by drop casting 10 μ L of the catalyst ink onto a polished glassy carbon electrode with 5 mm diameter. The catalyst film was dried in air for 25 min at a rotation speed of 700 rpm. The counter electrode was a graphite rod. The electrolyte used for electrochemical experiments was 1.0 M sodium phosphate buffer solution (PBS, pH 7) or 0.1 M KOH (pH 13). Currents are reported with anodic current as positive and cathodic current as negative. Potentials were measured against a Ag/AgCl reference electrode (3 M NaCl) and converted to the RHE reference scale using:

$$E \text{ (vs. RHE)} = E \text{ (vs. Ag/AgCl)} + 0.210 \text{ V} + 0.0591 \text{ V} \times \text{pH}$$

Cyclic voltammetry (CV) scans for Co-WS₂ samples were performed in N₂-saturated 0.1 M KOH electrolyte with a scan rate of 50 mV s^{−1}. Oxygen reduction reaction voltammetry was carried out using a rotating ring-disk electrode in a single-compartment glass cell containing 150 mL of 1.0 M PBS electrolyte. The solution was purged with O₂ for at least 30 min prior to the start of the experiment. For all ORR experiments, a rotation rate of 1600 rpm was utilized. One cathodic linear sweep voltammetry scan was collected at 10 mV s^{−1} prior to the reported anodic LSV, obtained by scanning at 1 mV s^{−1} from 0.0 V to 0.9 V vs. RHE. The onset potential (*E*_{onset}) for ORR is defined

to be the potential at which the ORR current density reaches -0.1 mA cm^{-2} . All E_{onset} values are obtained from an average of three sample runs to account for variability in the electrode drop-drying process, and the error bar reflects the standard deviation (Fig. S11 and S12†). The Pt ring in the rotating-ring disk electrode was held at 1.26 V vs. RHE during ORR linear sweep voltammetry.

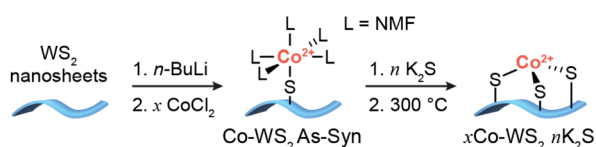
Uncompensated resistances were measured for a set of Co-WS₂ samples, which ranged from 20–36 Ω in 1.0 M PBS electrolyte. Catalytic measurements are reported without iR compensation. To assess the stability of Co-WS₂ catalysts in ORR, galvanostatic electrolysis was performed at -0.25 mA cm^{-2} in the standard electrolyte (1 M PBS) as well as with added methanol (1 M MeOH + 1 M PBS). Accelerated CV scanning was performed in a high-purity 1.0 M sodium phosphate electrolyte (pH 7, 99.9%). During the CV stability test, anodic LSV scans at 1 mV s^{-1} were obtained in the 1st, 5th, 20th, and 200th scan while all intervening CV scans were collected at 200 mV s^{-1} (Fig. S16†).

Results and discussion

Synthesis

The synthetic strategy to generate Co single atoms with tunable Co-S coordination environment, depicted in Scheme 1, is based on previous work from our group on the synthesis of single atom Ni-doped WS₂ nanosheets.⁵⁷ Colloidal WS₂ nanosheets were synthesized *via* a literature method and activated using *n*-butyllithium (*n*-BuLi) to generate dangling sulfide defects on the basal planes of WS₂ (Li-WS₂).⁵⁶ A dilute solution of CoCl₂ in *N*-methylformamide (NMF) was then introduced into the colloidal Li-WS₂ nanosheets under an inert atmosphere and permitted to passively adsorb over 24 hours. Any excess CoCl₂ remaining in solution was readily removed through precipitation and centrifugation of the Co-doped WS₂ nanosheets (colloidal Co-WS₂). Two loadings of Co on WS₂ were synthesized in order to vary the degree of dopant clustering and aggregation – a low loading (0.3 equiv.) to access more isolated Co atoms and a high loading (1.0 equiv.) to generate larger Co aggregates.

In our previous work, we showed that Ni dopants adsorbed relatively weakly and in a monodentate fashion to the WS₂ surface during the colloidal synthesis. In order to tune the Co-S coordination and more strongly anchor the Co dopants to the WS₂ surface, we develop a method herein for controlled sulfidation of the adsorbed Co atoms. The colloidal Co-WS₂ nanosheets are supported on carbon black (xCo-WS₂ As-Syn, where *x*



Scheme 1 Synthetic strategy for the deposition of Co single atoms onto colloidal WS₂ nanosheets followed by controlled sulfidation and annealing steps.

refers to the nominal Co loading) and impregnated with a variable amount of potassium sulfide, ranging from 0.5 to 10 equivalents with respect to the adsorbed Co atoms. The impregnated samples are then annealed at 300 °C under an inert atmosphere and rinsed repeatedly with solvent to remove unreacted sulfides ($\text{xCo-WS}_2 \text{ nK}_2\text{S}$).

Scanning transmission electron microscopy

We utilize high-angle annular dark-field scanning transmission electron microscopy (HAADF-STEM) coupled to energy-dispersive X-ray spectroscopy (EDS) to characterize the morphology and elemental distribution of Co and W on the Co-WS₂ nanosheets. Colloidal WS₂ comprises circular nanosheets with $\sim 100 \text{ nm}$ diameter and 10–14 layer thickness (Fig. S1a†).⁵⁷ After *n*-BuLi activation and Co adsorption, the nanosheet morphology is retained, and Co atoms are evenly distributed across the WS₂ based on the low-resolution HAADF-STEM image and EDS maps (Fig. S1b–f†). Based on EDS, Co loadings on the WS₂ nanosheets follow closely with the amount introduced in solution, ranging from 0.2 up to 1.0 equivalents with respect to W (Table S1†). At high Co loading, we anticipate that a significant fraction of Co species may be loosely intercalated or physisorbed within the WS₂ layers rather than coordinatively bound to the surface.

High-resolution HAADF-STEM imaging allows us to confirm the presence of Co single atoms in the low loading 0.3Co-WS₂ colloidal nanosheets in two surface adsorption sites (Fig. 1a). Due to the *n*-BuLi treatment, the underlying WS₂ nanosheets exhibit variable thickness, ranging from monolayer to multilayer, and regions where W vacancies and structural disorder are present (Fig. 1b and S2a†). In the monolayer regions, the ordered hexagonal lattice of W atoms for the WS₂ nanosheet is clearly visible, and deviations to the hexagonal array in terms of



Fig. 1 (a) Schematic depicting side and top views for S-top and W-top binding of Co atoms on WS₂, (b) high-resolution STEM image of colloidal 0.3Co-WS₂, (c) HR-STEM region containing S-top Co sites, and (d) HR-STEM region containing W-top Co sites.

both atomic position and contrast provide evidence for the presence of Co single atoms (Fig. 1c and d). Atoms interstitial to the hexagonal W array, highlighted in Fig. 1c, S2b and c,[†] represent Co atoms bound on-top of a S atom or in a S vacancy (S-top). In addition, analysis of annular-dark field (ADF) Z-contrast shows that Co atoms are also present on-top of W atoms in the hexagonal lattice, highlighted in Fig. 1d, S2d and e (W-top). Interestingly, Co atoms bound in W-top sites are frequently found adjacent to W vacancies in the lattice, suggesting that dangling sulfide moieties may be responsible for Co coordination in these positions.

In addition to ADF Z-contrast analysis, we also obtained EDS spectra in small regions throughout the high-resolution STEM image to ascertain that Co atoms can be spectroscopically detected on colloidal 0.3Co-WS₂. While the EDS signal is quite weak when focused on 3 × 3 nm square regions of the monolayer Co-WS₂ nanosheets, we can clearly distinguish peaks associated with the W M, S K α , and Co K α transitions (Fig. S2f and g[†]). In addition, when the EDS spectra is obtained over a slightly larger region, the atomic ratio of Co : W is $\sim 0.3 : 1$, similar to the elemental ratios obtained in the lower resolution EDS data.

Prior to sulfur addition, colloidal Co-WS₂ nanosheets are supported on carbon black, which causes the circular nanosheets to fold onto themselves rather than lie flat on the hydrophobic carbon surface (Fig. 2a and b). Impregnation of 0.5–10 equiv. of K₂S (with respect to Co) and annealing at 300 °C under N₂ does not alter the overall nanosheet morphology or the average loading of Co relative to W (Fig. 2c, d, S3, S4 and

Table S1[†]). However, we begin to see evidence of Co clustering in the STEM-EDS elemental maps after K₂S annealing. In 0.3Co-WS₂ samples treated with K₂S, Co atoms remain colocalized with W and S and evenly distributed across the WS₂ nanosheets in the lower-resolution STEM-EDS maps (Fig. 2c). Obtaining EDS spectra and quantifying the Co : W ratios at higher resolution reveals spatial variation in the Co distribution on WS₂ (Fig. S5[†]). In 8 × 8 nm regions across the 0.3Co-WS₂ 3K₂S sample, we observe Co : W (X : 1) ratios ranging from 0.14–0.29, suggesting that Co atoms are beginning to cluster but larger cobalt sulfide nanoparticles have not yet formed on the nanosheets. At the higher loading of Co (1.0Co-WS₂ nK₂S), cobalt sulfide nanoparticles are observed even in the low-resolution STEM-EDS maps (Fig. 2d). Aggregates of Co atoms (10–30 nm) that are not co-localized with the underlying W are clearly present.

Powder X-ray diffraction (XRD) patterns show the expected WS₂ phase transition from 1T to 2H after thermal annealing and no Co-containing crystalline phases (Fig. S6 and S7[†]). In the 1.0Co-WS₂ nK₂S samples, the CoO_x or CoS_x aggregates observed in the STEM-EDS maps must therefore be amorphous or comprise many small nanocrystalline domains.

Oxygen reduction reaction

We began by evaluating the oxygen reduction reactivity of the low and high loading Co-WS₂ catalysts as a function of the amount of K₂S incorporated during thermal treatment. The ORR catalytic activity of all samples was measured using linear scan voltammetry (LSV) on a rotating ring-disk electrode in O₂-saturated 1 M phosphate buffer solution (PBS, pH = 7). Negligible ring current is detected using any of the Co-WS₂ catalysts, which indicates quantitative selectivity toward the 4 e[−] reduction of O₂ to H₂O under these conditions (Fig. S13[†]). In order to compare the reactivity of various Co-WS₂ and bulk control samples to one another, we utilize the ORR onset potential (E_{onset}), the potential at which the catalyst attains -0.1 mA cm^{-2} of ORR current density, as the primary catalytic metric due to differences in sample mass-transport properties at higher current densities (Fig. S8 and S9[†]).

The annealed Li-WS₂ nanosheets alone show poor ORR activity with E_{onset} of 0.575 V vs. RHE (Fig. S8[†]). The addition of Co to the WS₂ sample at low loading (0.3Co-WS₂) in the absence of K₂S (0 K₂S) induces a modest increase in ORR onset potential to 0.614 V vs. RHE (Fig. 3a and c). Introduction of varying equivalents of K₂S to the 0.3Co-WS₂ sample has a dramatic effect on ORR catalysis. Upon mild sulfidation at 0.5 equiv. K₂S, the onset potential for ORR immediately increases by 54 mV to 0.668 V vs. RHE. From 0.5 to 3.0 equiv. of K₂S relative to Co, a steady increase in ORR onset potential is observed with the highest E_{onset} of 0.798 V vs. RHE occurring at 3.0 equiv. K₂S. Further increase in K₂S loading causes a decay in the ORR activity, dropping back down to 0.731 V vs. RHE at 10 equiv. K₂S. A modest improvement in catalytic activity is also observed when Li-WS₂ alone is annealed in the presence of excess sulfur (Fig. S8 and Table S2[†]). However, the onset potential of Li-WS₂ peaks at 0.639 V vs. RHE in the absence of doped Co atoms.

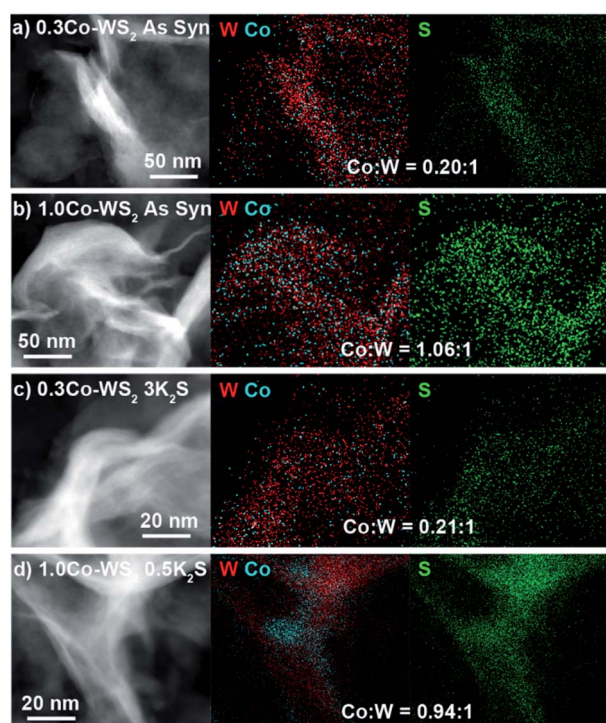


Fig. 2 STEM-EDS image and elemental maps of W, Co, and S for (a and b) as-synthesized 0.3Co and 1.0Co-WS₂ supported on C as well as (c and d) 0.3Co and 1.0Co-WS₂ after annealing with K₂S.



Fig. 3 Oxygen reduction reaction activity in O₂-saturated 1 M PBS (pH = 7) electrolyte. (a and b) ORR LSVs for the full series of 0.3Co-WS₂ samples treated with K₂S, (c) *E*_{onset} vs. equivalents of K₂S for Co-WS₂ catalysts compared to Co and WS₂ control samples.

At the higher Co loading (1.0Co-WS₂), the ORR activity is also sensitive to K₂S loading but with a somewhat different trend relative to the lower Co loading (Fig. 3b). In the absence of K₂S, 1.0Co-WS₂ exhibits an *E*_{onset} of 0.680 V vs. RHE. The *E*_{onset} rises rapidly with the addition of K₂S and reaches a peak at 1 equiv. K₂S, significantly earlier than on the 0.3Co-WS₂ sample (Fig. 3c). Intriguingly, the *E*_{onset} plateaus at this point and stays relatively constant at ~0.81 V vs. RHE even up to 10 equiv. K₂S. The fast rise and lack of a peak in the onset potential trend is consistent with the presence of aggregated CoS_x nanoparticles rather than isolated Co atoms or small Co clusters. The lower utilization of Co in the aggregated sample means a smaller amount of K₂S is required to fully sulfidize the catalytic surface, and excess K₂S serves only to convert more of the nanoparticle core into sulfided species without further influencing the surface.

Comparing the high and low loading Co-WS₂ at their most active sulfidation point, we find that the ORR onset potentials are quite similar between the two samples despite the very different amounts of Co present (Fig. 3c). These data further support that the low loading 0.3Co-WS₂ displays excellent catalytic utilization of Co atoms due to the relatively unaggregated structure of supported Co atoms on WS₂. For this reason, we believe that the 0.3Co-WS₂ treated with varying equivalents with K₂S serves as a useful structural model to understand the role of Co-S coordination in dictating ORR reactivity.

To evaluate the catalytic stability of Co-doped WS₂ catalysts, we undertook galvanostatic bulk electrolysis at -0.25 mA cm^{-2} on the 0.3Co-WS₂ 3K₂S sample and a Pt/C control (Fig. S17†). While Pt/C is more active, Co-WS₂ is somewhat more stable in 1 M PBS electrolyte. Over 60 min of electrolysis, the 0.3Co-WS₂ 3K₂S catalyst decays by 47 mV while Pt/C decays by 90 mV. Notably, when 1 M MeOH is added to the electrolyte, the overpotential required to drive ORR increases by 180 mV on Pt/C due to carbon monoxide poisoning of the Pt surface while remaining essentially unchanged on 0.3Co-WS₂.

X-ray absorption spectroscopy

To probe the local coordination environment and electronic state of adsorbed Co atoms in Co-WS₂, we obtain Co K-edge X-

ray absorption spectroscopy (XAS) for Co-WS₂ samples in three states: as-synthesized (As-Syn), K₂S-treated (K₂S), and post-catalysis (PBS). These spectra are compared to the precursor molecular complex, Co(NMF)_xCl₂, as well as a control sample of amorphous cobalt sulfide (CoS_x) with approximate composition of CoS_{1.2} (Fig. S14 and S15†).^{23,24} Fitting of the Fourier-transformed X-ray absorption fine structure (EXAFS) data was carried out on a subset of samples that showed sufficient Co K-edge absorption, and coordination numbers (CN) and bond distances (*R*) for both Co-O and Co-S scattering were obtained. Full fitting parameters and overlays of fitted and experimental data are provided in Table S4 and Fig. S20–S22.†

The X-ray absorption near edge spectroscopy (XANES) data for the precursor Co(NMF)_xCl₂ complex exhibits Co K-edge energy of 7721 eV and high white line intensity, consistent with a Co²⁺ ion primarily coordinated to oxygen atoms, as anticipated for a CoCl₂ precursor dissolved in the coordinating



Fig. 4 Co K-edge XANES and EXAFS spectra for (a) Co(NMF)_xCl₂, (b) As-synthesized 0.3Co-WS₂, (c) 0.3Co-WS₂ annealed with 0.5 equiv. of K₂S, (d) 0.3Co-WS₂ annealed with 3.0 equiv. of K₂S, (e) 0.3Co-WS₂ 3K₂S immersed in 1 M PBS electrolyte, and (f) amorphous CoS_x. Dotted lines indicate the edge energies or scattering peaks for Co-O and Co-S coordination.

Table 1 Co K-edge EXAFS fitting parameters for Co(NMF)_x complex and Co-WS₂ samples at varying synthetic stages including as-synthesized, K₂S-treated, and PBS-treated

| Sample | Pair | CN | <i>R</i> (Å) | <i>E</i> ₀ (eV) |
|--|------|-----------|--------------|----------------------------|
| Co(NMF) _x Cl ₂ | Co-O | 6.0 ± 0.5 | 2.08 ± 0.007 | -2.3 ± 0.9 |
| 0.3Co-WS ₂ 0.5K ₂ S | Co-O | 3.8 ± 0.2 | 2.10 ± 0.004 | 7.4 ± 0.3 |
| | Co-S | 2.5 ± 0.1 | 2.25 ± 0.004 | |
| 0.3Co-WS ₂ 3K ₂ S | Co-O | 1.7 ± 0.4 | 2.10 ± 0.022 | 3.3 ± 0.6 |
| | Co-S | 3.9 ± 0.3 | 2.25 ± 0.006 | |
| 1.0Co-WS ₂ As-Syn | Co-O | 4.3 ± 0.4 | 2.08 ± 0.009 | 0.30 ± 0.9 |
| 1.0Co-WS ₂ 0K ₂ S | Co-O | 4.3 ± 0.3 | 2.08 ± 0.007 | -0.29 ± 0.8 |
| 1.0Co-WS ₂ 1K ₂ S | Co-O | 5.5 ± 0.4 | 2.10 ± 0.006 | -0.52 ± 0.7 |
| 1.0Co-WS ₂ 3K ₂ S | Co-O | 4.9 ± 0.2 | 2.10 ± 0.004 | -1.2 ± 0.3 |
| | Co-S | 0.5 ± 0.1 | 2.25 ± 0.024 | |
| 1.0Co-WS ₂ 10K ₂ S | Co-S | 4.4 ± 0.2 | 2.25 ± 0.005 | -3.1 ± 0.5 |
| 1.0Co-WS ₂ 0K ₂ S PBS | Co-O | 4.1 ± 0.1 | 2.04 ± 0.003 | -1.0 ± 0.4 |
| 1.0Co-WS ₂ 1K ₂ S PBS | Co-O | 5.0 ± 0.3 | 2.10 ± 0.006 | -2.1 ± 0.6 |
| | Co-S | 0.8 ± 0.2 | 2.25 ± 0.025 | |
| 1.0Co-WS ₂ 3K ₂ S PBS | Co-S | 4.5 ± 0.2 | 2.25 ± 0.005 | -4.8 ± 0.5 |
| 1.0Co-WS ₂ 10K ₂ S PBS | Co-S | 4.6 ± 0.4 | 2.25 ± 0.007 | -2.8 ± 0.9 |

solvent NMF (Fig. 4a).⁶⁰ Likewise, the EXAFS spectrum shows a scattering peak at a short radial distance due to pure Co-O coordination. Fitting of the EXAFS data shows a Co-O bond length of 2.08 Å with CN ~6, consistent with an octahedral molecular complex (Table 1). On the opposite end of the spectrum, amorphous CoS_x exhibits a completely suppressed XANES white line and a significantly lower Co K-edge energy of 7714 eV with a strong pre-edge feature at 7710 eV, indicative of mixed tetrahedral and octahedral coordination (Fig. 4f).⁶¹ Compared to Co-O coordination, Co-S coordination generates a scattering pathway at higher radial distance due to the longer average bond length, which ranges from 2.18 to 2.45 in known bulk phases of cobalt sulfide (Table S5 and Fig. S25†). The XANES and EXAFS of these two control samples serve as bookends for the Co-WS₂ samples treated with varying amounts of sulfur because we expect the first-shell coordination environment to transform from oxygen-rich to sulfur-rich with increasing K₂S loading.

We begin our XAS analysis with the low loading 0.3Co-WS₂ *n*K₂S series of samples because their structural homogeneity and lack of aggregation lend themselves to clear structure-activity correlations. In the as-synthesized 0.3Co-WS₂ sample, we observe small perturbations to the adsorbed Co coordination environment relative to the free Co(NMF)_xCl₂ complex. A drop in white line intensity in the XANES and a small shift to longer radial distance for the first Co-X (X = O or S) scattering pathway relative to the free complex are both indicative of an increase in the number of sulfur atoms relative to oxygen atoms in the first coordination sphere (Fig. 4b). However, the persistence of the white line peak indicates that Co atoms remain primarily oxygen-bound in this form, when passively-adsorbed to the Li-WS₂ nanosheet in the solution phase. These data are consistent with the observation of Co atoms bound in S-top sites and W-top sites adjacent to vacancies in the high-resolution STEM images of colloidal 0.3Co-WS₂. We postulate that the complex likely binds in a mono- or bidentate fashion to

dangling sulfur sites on the lithiated WS₂ surface, and the remaining coordination sites remain occupied by NMF ligands.

Upon K₂S impregnation and annealing, the XAS data for 0.3Co-WS₂ shows a significant increase in the Co-S coordination to the WS₂ surface. Treatment with 0.5 equiv. or 3 equiv. of K₂S completely suppresses the sharp white line feature diagnostic of Co²⁺-O binding and shifts the Co K-edge energy downwards by 3 eV, approaching the edge energy of the CoS_x control sample (Fig. 4c and d). The EXAFS scattering peak also shifts to longer radial distance, and the sample treated with 3 equiv. K₂S essentially matches the peak position of CoS_x. Fitting of the EXAFS data indicates that the Co-S CN of 0.3Co-WS₂ increases steadily from 2.5 at 0.5 equiv. K₂S up to 3.9 at 3 equiv. K₂S while the Co-O CN drops correspondingly (Table 1). In all of these samples, no significant scattering density is observed beyond 3.0 Å, indicating that Co atoms remain relatively isolated on the WS₂ surface even after thermal annealing. No metallic Co-Co scattering or second coordination sphere Co-X-Co (X = O or S) scattering features are observed based on comparisons to Co foil and a variety of known CoO_x and CoS_x crystalline phases (Fig. S20, S25, Tables S4 and S5†). Together, these data show that Co atomic coordination to the WS₂ surface can be readily controlled in the 0.3Co-WS₂ system simply by varying the amount of impregnated K₂S.

Lastly, we characterized the 0.3Co-WS₂ samples after 1 hour immersion in phosphate buffer solution, catalytically equivalent to the material after 5–20 CV scans, to understand how the adsorbed Co sites evolve in aqueous electrolyte (Fig. S26 and Table S6†). XANES data on 0.3Co-WS₂ 3K₂S immersed in PBS shows a similar Co K-edge energy but a slight increase in white line intensity relative to the pre-catalysis material (Fig. 4e). The EXAFS spectrum also shows a small shift in the first-shell scattering peak to lower radial distance. These data indicate that the Co-S bonds that anchor the dopants to the WS₂ surface hydrolyze slightly in the presence of aqueous electrolyte, likely causing the decay in catalytic activity observed in the stability tests (Fig. S16 and S17†). However, Co surface sites remain relatively unaggregated on the WS₂ surface after the electrolyte treatment, as evidenced by the continued absence of a second coordination sphere Co-S-Co scattering feature between 3–4 Å

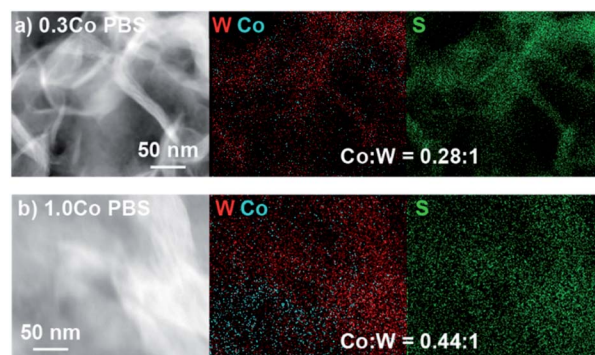


Fig. 5 STEM-EDS image and elemental maps of W, Co, and S for (a) 0.3Co-WS₂ 3K₂S and (b) 1.0Co-WS₂ 3K₂S after treatment with PBS electrolyte.

(Fig. 4e and S24c†). STEM-EDS mapping and elemental analysis after electrolyte treatment show that Co loading and spatial distribution remain unchanged relative to the pre-catalysis samples (Fig. 5a and S27†). Notably, Co atoms are fully retained in the 0.3Co-WS₂ *n*K₂S samples after immersion in aqueous electrolyte, suggesting that all Co atoms are coordinatively attached to the WS₂ surface (Table S7†).

We can now correlate XAS structural characterization on the 0.3Co-WS₂ samples to their initial ORR activity because, unlike in bulk or nanoparticle cobalt sulfide materials, the majority of Co atoms are available on the surface for catalysis. We find that 0.3Co-WS₂ treated with 3 equiv. K₂S, exhibiting three- to four-fold coordination of Co to sulfur atoms on the WS₂ surface, provides an optimal electronic and geometric environment for pH 7 ORR catalysis. We hypothesize that the activation of adsorbed Co atoms with K₂S up to 3 equiv. stems from the necessity for a more electron-rich metal center, enabled by Co-S coordination, in order to turn over adsorbed OH intermediates in the catalytic cycle.^{35,62} Further increasing K₂S, however, likely causes over-coordination with sulfur and therefore loss of available coordination sites for O₂ binding and activation (Fig. S29†). During catalysis, extended exposure to aqueous electrolyte induces partial hydrolysis of Co-S bonds, leading to a slight decay in ORR activity.

As a comparison to our Co-WS₂ materials, we also obtained ORR catalytic data on three Co-containing phases with very different bulk composition and local Co coordination environment: Co(OH)₂, amorphous CoS_x (~CoS_{1.2}), and CoS₂ (Fig. S14 and S15†). The Co(OH)₂ nanosheets with pure Co-O coordination show the poorest ORR onset potential of 0.514 V vs. RHE (Fig. 3c and Table S3†). Even the 0.3Co-WS₂ sample treated with 0 equiv. K₂S is substantially more active due to the interaction with and contribution of the underlying WS₂ nanosheets. The amorphous CoS_x sample with an intermediate S : Co ratio of ~1 : 1 shows the highest *E*_{onset} among the bulk phases at 0.767 V vs. RHE, nearly matching the onset potential of the optimal 0.3Co-WS₂ 3K₂S catalyst. The most sulfur-rich sample, comprising CoS₂ nanoparticles, exhibits lower ORR activity with an onset potential of 0.660 V vs. RHE, which mirrors the drop in activity of 0.3Co-WS₂ 10K₂S when it becomes over-saturated with sulfur. The similar trends in onset potential as a function of cobalt sulfidation between 0.3Co-WS₂ and bulk Co-containing phases provides an initial indication that the local coordination environment identified for Co atoms supported on WS₂ may also apply to active sites formed *in situ* on the surface of bulk cobalt sulfide species during ORR catalysis.

Because XAS spectra represent an average over all Co species in the sample, the high loading 1.0Co-WS₂ sample, comprising large Co aggregates and heterogeneous speciation, is much more difficult to directly correlate to catalytic reactivity. Nonetheless, the XAS data at each stage of the synthetic and catalytic process provide useful information about the speciation and evolution of adsorbed Co species at high loading.

In the as-synthesized 1.0Co-WS₂, both the XANES and EXAFS at the Co K-edge closely resemble the spectra for the free Co(NMF)_xCl₂ complex, suggesting that a significant fraction of Co complexes at this high loading are intercalated or weakly

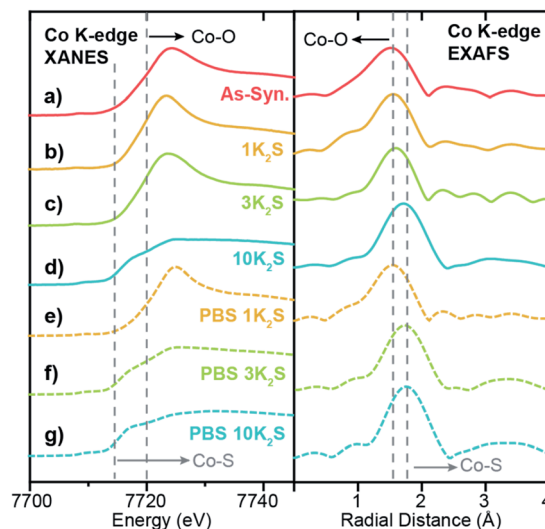


Fig. 6 Co K-edge XANES and EXAFS spectra for 1.0Co-WS₂ samples at varying stages of synthesis and catalysis. (a) As-synthesized, (b–d) thermally treated with varying equivalents of K₂S, and (e–g) immersed in 1 M PBS electrolyte. Dotted lines indicate the edge energies or scattering peaks for Co–O and Co–S coordination.

physisorbed due to the ionic interaction between the negatively charged WS₂ sheets and cationic Co complexes (Fig. 6a). After K₂S addition and thermal annealing in the presence of 1 and 3 equiv. K₂S, small increases in EXAFS first-shell scattering intensity are observed but with no significant increase in radial distance (Fig. 6b and c). Coupled to the persistently high white line intensity in the XANES, these data indicate that the bulk of the Co atoms in 1.0Co-WS₂ remain unsulfidized up to 3 equiv. K₂S (Table 1). At 10 equiv. K₂S, an abrupt increase in first-shell scattering distance and drop in white line intensity reveals that the majority of Co atoms have been converted to cobalt sulfide species (Fig. 6d). The relatively high equivalents of K₂S required to sulfidize the Co species in the 1.0Co-WS₂ is consistent with the presence of aggregated and intercalated Co species, clearly indicating that these average XAS spectra are not representative of catalytically active Co species.

After the PBS immersion, elemental analysis using XRF reveals that significant amounts of Co are leached out of the 1.0Co-WS₂ nanosheets, which is in stark contrast to the full Co retention observed in the 0.3Co-WS₂ samples (Table S7†). The Co spatial distribution also becomes more non-uniform and aggregated based on STEM-EDS elemental maps after PBS treatment (Fig. 5b, S28 and Table S8†). The quantity of leached Co is greatest at low K₂S loadings, likely because the loosely intercalated Co(NMF)_x molecular complexes observed in the as-synthesized sample are not yet converted to insoluble CoS_x aggregates and can be readily removed in strongly ionic solutions. As a result, the post-PBS XAS spectra are more representative of the Co species present during catalysis for the high loading 1.0Co-WS₂ sample (Fig. 6e–g). The difference in XAS spectra between the pre- and post-electrolyte treatment is clearest in the 1.0Co-WS₂ sample at intermediate K₂S loading (3 K₂S). While the spectrum prior to PBS immersion appears

minimally sulfidized, the spectrum after PBS immersion shows the longer radial distance and suppressed white line intensity characteristic of a cobalt sulfide species (Fig. 6c and f). This sample clearly demonstrates that unreacted Co complexes after K_2S annealing are dissolved out during the immersion in electrolyte solution, leaving behind only the sulfidized Co nanoparticles. Analysing the full series of post-PBS samples, we observe an increase in Co–S CN and decrease in Co–O CN as the K_2S loading is increased (Table 1). Full sulfidation is observed by 3 equiv. K_2S , which exhibits Co–S CN of 4.5 and no contribution from Co–O. Unlike in the 0.3Co–WS₂ samples, however, it is impossible to correlate these average Co–S coordination numbers directly to ORR catalysis because of the presence of large CoS_x aggregates and the fact that the majority of Co atoms are not available at the surface for catalysis. As anticipated, all 1.0Co–WS₂ samples after PBS treatment display a second coordination sphere Co–X–Co (X = O or S) scattering feature, which corroborates the presence of large aggregates and nanoparticles observed in the STEM-EDS images (Fig. S24†).

Conclusion

In conclusion, we find that relatively isolated Co atoms supported on WS₂ show high activity for the pH 7 ORR reaction when partially coordinated to sulfur ligands. Both undercoordination and overcoordination to sulfur prove detrimental to catalysis. The optimal Co active site comprises approximate 3–4-fold coordination to sulfur atoms on the surface of WS₂ and two labile coordination sites occupied by oxygen-based solvent ligands. These dilute Co–WS₂ catalysts show similar catalytic activity to the most active bulk phases of cobalt sulfide with an onset potential for ORR of 0.798 V vs. RHE in pH 7 phosphate buffer electrolyte. Unlike bulk CoS_x systems, however, the majority of Co atoms in dilute Co–WS₂ are available for catalysis, and the Co coordination environment obtained by XAS is representative of catalytically active sites. As a result, we postulate that the optimal coordination environment ascertained herein may be representative of the dynamic active sites present *in situ* on the surfaces of bulk cobalt sulfide catalysts. The activity of dilute Co–WS₂ clearly illustrates that the electronic and geometric environment created by Co–S bonds in the first coordination sphere provides a key descriptor for ORR catalysis in cobalt sulfide materials.

Author contributions

WH and EM contributed equally to this work.

Conflicts of interest

There are no conflicts to declare.

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References

- 1 C. Santoro, C. Arbizzani, B. Erable and I. Ieropoulos, *J. Power Sources*, 2017, **356**, 225–244.
- 2 B. E. Logan, B. Hamelers, R. A. Rozendal, U. Schröder, J. Keller, S. Freguia, P. Aelterman, W. Verstraete and K. Rabaey, *Environ. Sci. Technol.*, 2006, **40**, 5181–5192.
- 3 M. V. Kannan and G. G. Kumar, *Biosens. Bioelectron.*, 2016, **77**, 1208–1220.
- 4 Z. X. Hu, X. X. Zhou, Y. Lu, R. M. Jv, Y. Liu, N. Li and S. W. Chen, *Electrochim. Acta*, 2019, **296**, 214–223.
- 5 M. Kodali, S. Herrera, S. Kabir, A. Serov, C. Santoro, I. Ieropoulos and P. Atanassov, *Electrochim. Acta*, 2018, **265**, 56–64.
- 6 L. Z. Liu, G. Zeng, J. X. Chen, L. L. Bi, L. M. Dai and Z. H. Wen, *Nano Energy*, 2018, **49**, 393–402.
- 7 X. H. Tang and H. Y. Ng, *Electrochim. Acta*, 2017, **247**, 193–199.
- 8 O. Schaetzle, F. Barriere and U. Schroder, *Energy Environ. Sci.*, 2009, **2**, 96–99.
- 9 B. Yan, N. M. Concannon, J. D. Milshtein, F. R. Brushett and Y. Surendranath, *Angew. Chem., Int. Ed.*, 2017, **56**, 7496–7499.
- 10 J. M. Falkowski, N. M. Concannon, B. Yan and Y. Surendranath, *J. Am. Chem. Soc.*, 2015, **137**, 7978–7981.
- 11 P. W. Cai, J. H. Huang, J. X. Chen and Z. H. Wen, *Angew. Chem., Int. Ed.*, 2017, **56**, 4858–4861.
- 12 W. W. Zhao, P. Bothra, Z. Y. Lu, Y. B. Li, L. P. Mei, K. Liu, Z. H. Zhao, G. X. Chen, S. Back, S. Siahrostami, A. Kulkarni, J. K. Nørskov, M. Bajdich and Y. Cui, *ACS Appl. Energy Mater.*, 2019, **2**, 8605–8614.
- 13 H. Behret, H. Binder and G. Sandstedt, *Electrochim. Acta*, 1975, **20**, 111–117.
- 14 H. Behret, H. Binder, W. Clauberg and G. Sandstedt, *Electrochim. Acta*, 1978, **23**, 1023–1029.
- 15 Y. J. Feng, A. Gago, L. Timperman and N. Alonso-Vante, *Electrochim. Acta*, 2011, **56**, 1009–1022.
- 16 M. R. Gao, J. Jiang and S. H. Yu, *Small*, 2012, **8**, 13–27.
- 17 V. Vij, S. Sultan, A. M. Harzandi, A. Meena, J. N. Tiwari, W. G. Lee, T. Yoon and K. S. Kim, *ACS Catal.*, 2017, **7**, 7196–7225.
- 18 H. L. Wang, Y. Y. Liang, Y. G. Li and H. J. Dai, *Angew. Chem., Int. Ed.*, 2011, **50**, 10969–10972.
- 19 L. Zhu, D. Susac, M. Teo, K. C. Wong, P. C. Wong, R. R. Parsons, D. Bizzotto, K. A. R. Mitchell and S. A. Campbell, *J. Catal.*, 2008, **258**, 235–242.
- 20 D. Susac, L. Zhu, M. Teo, A. Sode, K. C. Wong, P. C. Wong, R. R. Parsons, D. Bizzotto, K. A. R. Mitchell and S. A. Campbell, *J. Phys. Chem. C*, 2007, **111**, 18715–18723.

- 21 A. Arunchander, S. G. Peera and A. K. Sahu, *ChemElectroChem*, 2017, **4**, 1544–1553.
- 22 J. M. Falkowski and Y. Surendranath, *ACS Catal.*, 2015, **5**, 3411–3416.
- 23 Y. F. Cao, Y. Y. Meng, S. C. Huang, S. M. He, X. H. Li, S. F. Tong and M. M. Wu, *ACS Sustainable Chem. Eng.*, 2018, **6**, 15582–15590.
- 24 P. Ganesan, M. Prabu, J. Sanetuntikul and S. Shanmugam, *ACS Catal.*, 2015, **5**, 3625–3637.
- 25 B. L. Chen, R. Li, G. P. Ma, X. L. Gou, Y. Q. Zhu and Y. D. Xia, *Nanoscale*, 2015, **7**, 20674–20684.
- 26 Q. L. Zhu, W. Xia, T. Akita, R. Q. Zou and Q. Xu, *Adv. Mater.*, 2016, **28**, 6391–6398.
- 27 Z. X. Wu, J. Wang, M. Song, G. M. Zhao, Y. Zhu, G. T. Fu and X. E. Liu, *ACS Appl. Mater. Interfaces*, 2018, **10**, 25415–25421.
- 28 Y. X. Zhou, H. B. Yao, Y. Wang, H. L. Liu, M. R. Gao, P. K. Shen and S. H. Yu, *Chem.–Eur. J.*, 2010, **16**, 12000–12007.
- 29 Y. P. Tang, F. Jing, Z. X. Xu, F. Zhang, Y. Y. Mai and D. Q. Wu, *ACS Appl. Mater. Interfaces*, 2017, **9**, 12340–12347.
- 30 S. F. Fu, C. Z. Zhu, J. H. Song, S. Feng, D. Du, M. H. Engelhard, D. D. Xiao, D. S. Li and Y. H. Lin, *ACS Appl. Mater. Interfaces*, 2017, **9**, 36755–36761.
- 31 Z. Xiao, G. Z. Xiao, M. H. Shi and Y. Zhu, *ACS Appl. Mater. Interfaces*, 2018, **10**, 16436–16448.
- 32 S. L. Zhang, D. Zhai, T. T. Sun, A. J. Han, Y. L. Zhai, W. C. Cheong, Y. Liu, C. L. Su, D. S. Wang and Y. D. Li, *Appl. Catal., B*, 2019, **254**, 186–193.
- 33 F. Bai, X. Qu, J. Wang, X. Chen and W. S. Yang, *ACS Appl. Mater. Interfaces*, 2020, **12**, 33740–33750.
- 34 R. A. Sidik and A. B. Anderson, *J. Phys. Chem. B*, 2006, **110**, 936–941.
- 35 B. Yan, D. Krishnamurthy, C. H. Hendon, S. Deshpande, Y. Surendranath and V. Viswanathan, *Joule*, 2017, **1**, 600–612.
- 36 D. Susac, A. Sode, L. Zhu, P. C. Wong, M. Teo, D. Bizzotto, K. A. R. Mitchell, R. R. Parsons and S. A. Campbell, *J. Phys. Chem. B*, 2006, **110**, 10762–10770.
- 37 X. F. Yang, A. Q. Wang, B. T. Qiao, J. Li, J. Y. Liu and T. Zhang, *Acc. Chem. Res.*, 2013, **46**, 1740–1748.
- 38 A. Q. Wang, J. Li and T. Zhang, *Nat. Rev. Chem.*, 2018, **2**, 65–81.
- 39 G. Giannakakis, M. Flytzani-Stephanopoulos and E. C. H. Sykes, *Acc. Chem. Res.*, 2019, **52**, 237–247.
- 40 Y. Xin, N. N. Zhang, Q. Li, Z. L. Zhang, X. M. Cao, L. R. Zheng, Y. W. Zeng and J. A. Anderson, *ACS Catal.*, 2018, **8**, 1399–1404.
- 41 G. Q. Gan, X. Y. Li, L. Wang, S. Y. Fan, J. C. Mu, P. L. Wang and G. H. Chen, *ACS Nano*, 2020, **14**, 9929–9937.
- 42 Y. Xu, M. Y. Chu, F. F. Liu, X. C. Wang, Y. Liu, M. H. Cao, J. Gong, J. Luo, H. P. Lin, Y. Y. Li and Q. Zhang, *Nano Lett.*, 2020, **20**, 6865–6872.
- 43 J. Q. Zhang, Y. F. Zhao, C. Chen, Y. C. Huang, C. L. Dong, C. J. Chen, R. S. Liu, C. Y. Wang, K. Yan, Y. D. Li and G. X. Wang, *J. Am. Chem. Soc.*, 2019, **141**, 20118–20126.
- 44 Q. Wang, X. Huang, Z. L. Zhao, M. Y. Wang, B. Xiang, J. Li, Z. X. Feng, H. Xu and M. Gu, *J. Am. Chem. Soc.*, 2020, **142**, 7425–7433.
- 45 A. Zitolo, V. Goellner, V. Armel, M. T. Sougrati, T. Mineva, L. Stievano, E. Fonda and F. Jaouen, *Nat. Mater.*, 2015, **14**, 937–942.
- 46 L. L. Zhang, A. Q. Wang, W. T. Wang, Y. Q. Huang, X. Y. Liu, S. Miao, J. Y. Liu and T. Zhang, *ACS Catal.*, 2015, **5**, 6563–6572.
- 47 M. Y. Hu, S. N. Li, S. S. Zheng, X. H. Liang, J. X. Zheng and F. Pan, *J. Phys. Chem. C*, 2020, **124**, 13168–13176.
- 48 W. G. Liu, L. L. Zhang, W. S. Yan, X. Y. Liu, X. F. Yang, S. Miao, W. T. Wang, A. Q. Wang and T. Zhang, *Chem. Sci.*, 2016, **7**, 5758–5764.
- 49 W. G. Liu, L. L. Zhang, X. Liu, X. Y. Liu, X. F. Yang, S. Miao, W. T. Wang, A. Q. Wang and T. Zhang, *J. Am. Chem. Soc.*, 2017, **139**, 10790–10798.
- 50 A. Zitolo, N. Ranjbar-Sahraie, T. Mineva, J. K. Li, Q. Y. Jia, S. Stamatin, G. F. Harrington, S. M. Lyth, P. Krttil, S. Mukerjee, E. Fonda and F. Jaouen, *Nat. Commun.*, 2017, **8**, 957.
- 51 P. Q. Yin, T. Yao, Y. Wu, L. R. Zheng, Y. Lin, W. Liu, H. X. Ju, J. F. Zhu, X. Hong, Z. X. Deng, G. Zhou, S. Q. Wei and Y. D. Li, *Angew. Chem., Int. Ed.*, 2016, **55**, 10800–10805.
- 52 P. Vancso, Z. I. Popov, J. Peto, T. Ollar, G. Dobrik, J. S. Pap, C. Y. Hwang, P. B. Sorokin and L. Tapasztó, *ACS Energy Lett.*, 2019, **4**, 1947–1953.
- 53 L. Ji, P. F. Yan, C. H. Zhu, C. Y. Ma, W. Z. Wu, C. Wei, Y. L. Shen, S. Q. Chu, J. O. Wang, Y. Du, J. Chen, X. A. Yang and Q. Xu, *Appl. Catal., B*, 2019, **251**, 87–93.
- 54 K. Qi, X. Q. Cui, L. Gu, S. S. Yu, X. F. Fan, M. C. Luo, S. Xu, N. B. Li, L. R. Zheng, Q. H. Zhang, J. Y. Ma, Y. Gong, F. Lv, K. Wang, H. H. Huang, W. Zhang, S. J. Guo, W. T. Zheng and P. Liu, *Nat. Commun.*, 2019, **10**, 5231.
- 55 Q. Wang, Z. L. Zhao, S. Dong, D. S. He, M. J. Lawrence, S. B. Han, C. Cai, S. H. Xiang, P. Rodriguez, B. Xiang, Z. G. Wang, Y. Y. Liang and M. Gu, *Nano Energy*, 2018, **53**, 458–467.
- 56 B. Mahler, V. Hoepfner, K. Liao and G. A. Ozin, *J. Am. Chem. Soc.*, 2014, **136**, 14121–14127.
- 57 E. Meza, R. E. Diaz and C. W. Li, *ACS Nano*, 2020, **14**, 2238–2247.
- 58 C. Li, A. P. Tardajos, D. Wang, D. Choukroun, K. Van Daele, T. Breugelmans and S. Bals, *Ultramicroscopy*, 2021, 221.
- 59 B. Ravel and M. Newville, *J. Synchrotron Radiat.*, 2005, **12**, 537–541.
- 60 M. Uchikoshi and K. Shinoda, *Struct. Chem.*, 2019, **30**, 945–954.
- 61 L. van Haandel, G. Smolentsev, J. A. van Bokhoven, E. J. M. Hensen and T. Weber, *ACS Catal.*, 2020, **10**, 10978–10988.
- 62 J. K. Norskov, J. Rossmeisl, A. Logadottir, L. Lindqvist, J. R. Kitchin, T. Bligaard and H. Jonsson, *J. Phys. Chem. B*, 2004, **108**, 17886–17892.