



Cite this: *Dalton Trans.*, 2015, **44**, 6638

Received 21st February 2015,

Accepted 2nd March 2015

DOI: 10.1039/c5dt00761e

www.rsc.org/dalton

Reduction of hydroxy-functionalised carbaboranyl carboxylic acids and ketones by organolithium reagents†

W. Neumann, M. Hiller, M. B. Sárosi, P. Lönnecke and E. Hey-Hawkins*

While the reaction of carbaboranyl carboxylic acids and ketones with organolithium reagents generally leads to cleavage of the *exo*-polyhedral C–C bond, introduction of a hydroxyl group at the second carbon atom of the cluster enables the reduction of the carbonyl compounds to tertiary alcohols. The proposed mechanism involving the formation of dimeric contact ion pairs was supported by X-ray crystallography and theoretical calculations.

Introduction

Carbonyl-functionalised carbaboranes are versatile building blocks in carbaborane chemistry.¹ However, one drawback of these compounds is the high lability of their *exo*-polyhedral C–C bond. Reaction of *ortho*-carbaboranyl ketones, carboxylic acids and esters with organolithium reagents mainly results in cleavage of the carbon–carbon bond rather than reduction of the carbonyl group, as the lithiated intermediates are probably destabilised due to the strong electron-withdrawing effect and steric hindrance of the cluster.² Less cleavage was observed with Grignard reagents, as the less polar O–Mg bonds apparently better stabilise the intermediate alkoxides.^{2e} In contrast, *ortho*-carbaboranyl aldehydes can be reduced to secondary alcohols by reaction with organolithium reagents with only little C–C bond cleavage.^{2e,f,3} We have recently reported that carbaboranyl carboxylic acids can also be reduced to the respective tertiary alcohols under such conditions with less *exo*-polyhedral C–C bond cleavage if a hydroxyl group is present at the second carbon atom of the cluster.^{2g} As the reduction led to a 1 : 1 mixture of tertiary alcohol and cleavage product, a reaction mechanism involving the formation of dimeric contact ion pairs was proposed. These dimers may enable intermolecular (intra-aggregate) transfer of an oxygen atom, thereby allowing the direct conversion of the acid to an alcohol instead of the expected ketone. Herein, further investigations into this unique reaction behaviour are presented.

The reduction of carboxylic acids by organolithium reagents typically results in the formation of geminal diolates, which spontaneously rearrange to the corresponding ketones upon acidic work-up. Reduction of the ketones to tertiary alcohols can then be achieved in a second step. In contrast, reduction of 1-hydroxy-1,2-dicarba-*closo*-dodecaboranyl-2-carboxylic acid (salborin) directly yields tertiary alcohols (reaction II, Table 1).^{2g} The formation of an intermediate ketone after intermolecular oxygen transfer has been proposed. To further investigate the formation of such an intermediate as well as the influence of the hydroxyl group on *exo*-polyhedral C–C bond cleavage, ketones of unsubstituted and hydroxy-functionalised *ortho*-carbaborane were synthesised and treated with *n*-BuLi.

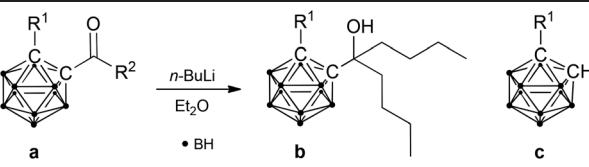
Results and discussion

1-Acetyl-1,2-dicarba-*closo*-dodecaborane was synthesised by treating lithiated *ortho*-carbaborane with acetyl chloride and separated from the disubstituted derivative by column chromatography (Scheme 1).^{2d} Accordingly, 1-acetoxy-2-acetyl-1,2-dicarba-*closo*-dodecaborane was prepared from lithiated 1-hydroxy-1,2-dicarba-*closo*-dodecaborane and acetyl chloride (Scheme 1). For subsequent deprotection of the hydroxyl group, a specific protocol had to be developed. Acetyl-protected alcohols are generally deprotected under alkaline conditions. However, carbaboranes tend to undergo deboronation of the cluster (decapping) under such conditions, and their lability is noticeably enhanced by carbonyl groups.⁴ The carboxylic acid congeners of the carbaboranyl ketone, salborin and asborin (1-acetoxy-1,2-dicarba-*closo*-dodecaboranyl-2-carboxylic acid), were shown to be already decapped under neutral conditions in aqueous solution.⁵ Furthermore, also the *exo*-polyhedral C–C bond of carbonyl-substituted carbaboranes

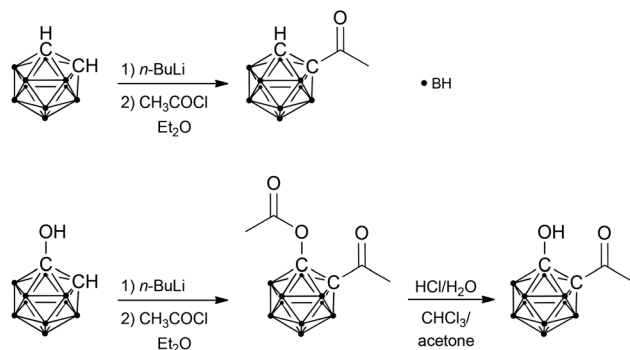
Universität Leipzig, Institut für Anorganische Chemie, Johannisallee 29, 04103 Leipzig, Germany. E-mail: hey@uni-leipzig.de

†Electronic supplementary information (ESI) available: Further crystal data. CCDC 1039317. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c5dt00761e



Table 1 Reaction of carbaboranyl carboxylic acids and ketones with *n*-BuLi


Reaction	R ¹	R ²	Products
I ^{2g}	H	OH	c
II ^{2g}	OH	OH	b + c (1 : 1)
III	H	CH ₃	c
IV	OH	CH ₃	b + c (1 : 4)
V	NHBoc ^a	OH	c
VI ^{2g}	OCH ₃	OH	c

^a Boc: *tert*-butoxy carbonyl.**Scheme 1** Syntheses of the studied carbaboranyl ketones.

is prone to cleavage under alkaline conditions.^{2b,d,6} Therefore, acidic conditions were chosen for hydrolysis of the ester. Although asborin was shown to readily be deacylated already under neutral conditions,⁵ the acidic environment was also expected to minimise the concurrent decapping of the cluster. However, even the use of acetone, which has been reported to suppress deboronation of carbonyl-functionalised carbaboranes,⁴ unexpectedly led to rapid decapping of the carbaboranyl ketone. Thus, chloroform was chosen as reaction medium and acetone was only added at the end of the reaction to accelerate deacylation (Scheme 1).

Both ketones were then treated with *n*-BuLi. As expected and reported for analogous derivatives,^{2b,ef} the reaction of 1-acetyl-1,2-dicarba-*closo*-dodecaborane resulted in cleavage of the *exo*-polyhedral C–C bond yielding unsubstituted *ortho*-carbaborane (reaction III, Table 1). In contrast, reaction of the hydroxy-functionalised carbaboranyl ketone gave the respective tertiary alcohol in a mixture with the cleavage product, 1-hydroxy-1,2-dicarba-*closo*-dodecaborane (reaction IV, Table 1). This confirms the previous findings that a hydroxyl group decreases C–C bond cleavage during reduction with organolithium reagents^{2g} and also supports the proposed formation

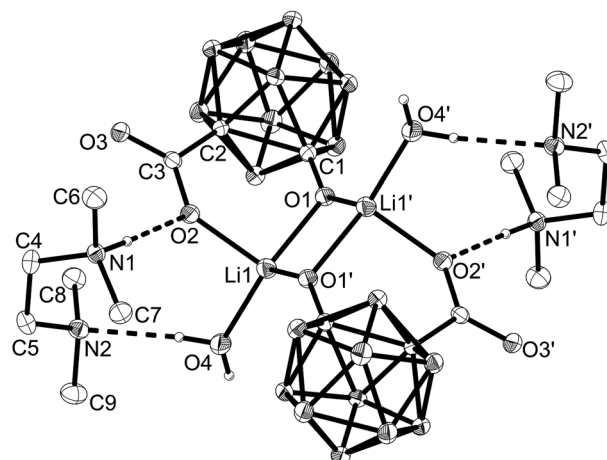


Fig. 1 ORTEP representation of the crystal structure of the lithium salt of 1-hydroxy-1,2-dicarba-*closo*-dodecaboranyl-2-carboxylic acid. Thermal ellipsoids are drawn at 50% probability; only selected hydrogen atoms are shown to indicate the direction of hydrogen bonds; all other hydrogen atoms are omitted for clarity. Selected bond lengths (Å): C1–O1 1.277(2), C1–C2 1.867(9), C1–C3 1.527(1), C3–O2 1.263(5), C3–O3 1.239(1). Crystallographic data: C₁₈H₅₈B₂₀Li₂N₄O₈, *M* = 688.76, triclinic, space group *P* $\bar{1}$, *a* = 7.2752(3), *b* = 11.5518(4), *c* = 12.3958(5) Å, *V* = 974.44(7) Å³, *Z* = 1, *T* = 130(2) K, 19 988 reflections measured, 5958 unique (*R*_{int} = 0.0250). Final *R* indices (all data): *R*₁ = 0.0452, *wR*₂ = 0.0922. Chains along the *a* axis are formed *via* intermolecular OH...O donor–acceptor bonds.

of a ketone-like intermediate during reduction of the corresponding carboxylic acid (salborin).

The previously proposed mechanism involved the formation of dimeric contact ion pairs.^{2g} This dimer formation is now supported by a crystal structure of the lithium salt of salborin (Fig. 1). In the solid state, a centrosymmetric alkoxide-bridged Li₂O₂ rhomboid is formed in which each Li ion is coordinated by both deprotonated hydroxyl groups and an oxygen atom of a carboxylate group. The tetrahedral coordination sphere of each Li ion is completed by a coordinating water molecule and thus formed exclusively by O atoms. Although generally known as an excellent Lewis base for the complexation of Li ions, the co-crystallised tetramethylethylenediamine (tmeda) molecules are not coordinated at the metal ions but protonated, and the resulting two cations compensate the charge of the dianionic dimeric contact ion pair. Apart from dimer formation, the carbaborane clusters do not interact with other clusters. In contrast, interaction of hydroxyl groups of adjacent clusters of salborin leads to formation of one-dimensional chains of dimers;⁵ the latter are formed *via* hydrogen bonds of the carboxylate groups, similar to other carbaboranyl carboxylic acids.^{5,7} Upon deprotonation, the hydroxyl groups seem to become the pivot of dimer formation.

Deprotonation also leads to a remarkable change in the geometry of the cluster caused by an increase in the C–O *exo*-donative π bonding of the hydroxyl group.⁸ Thus, the *exo*-polyhedral C–O bond (C(1)–O(1) 1.277(2) Å) is significantly shortened



compared to salborin (1.374(2) Å;⁵ Table S1 in the ESI†), with an associated increase in the bond order (C=O *ca.* 1.23 Å; C–O *ca.* 1.43 Å).^{8a,e} Delocalisation of the anionic charge of the oxygen atom into the carbaborane cage results in lengthening of the cluster C–C bond (C(1)–C(2)) concomitant with an opening of the *closo* geometry. While the C–C bond of salborin (1.660(1) Å) is only slightly longer than that of unsubstituted *ortho*-carbaborane (1.62 Å),⁹ the C(1)–C(2) bond in the lithium salt is significantly longer (1.867(9) Å). The B–B bonds remain mainly unaffected by deprotonation (Table S1 in the ESI†), whereas the similar C–O bond lengths of the carboxylato group (C(3)–O(2), C(3)–O(3)) indicate C–O double-bond character.

The arrangement of the clusters in the crystal structure of the lithium salt supports the importance of the hydroxyl groups for the unique reaction behaviour. In the reaction mixture, solvent molecules (Et₂O) may complete the coordination sphere of the lithium ions in place of the water molecules, and as structures in solution are dynamic, the bridging lithium ions may enable an orientation of both clusters appropriate for the proposed oxygen transfer between the clusters during reduction of the carbaboranyl acid (Scheme 2). Theoretical calculations support the rearrangement of the dimer in solution by revealing similar Gibbs free energies (ΔG) for different dimeric structures (Fig. 2): a centrosymmetric dimer, as found in the crystal structure (A), a *C*₂-symmetric dimer (B), a dimer with a bridging monodentate coordination (μ) of one carboxylato group (C), and a dimer with a bridging bidentate coordination ($\mu, \kappa^1: \kappa^1$) of one carboxylato group (D). Further dimeric structures with chelating carboxylato groups and non-bridging hydroxyl groups would be theoretically possible.

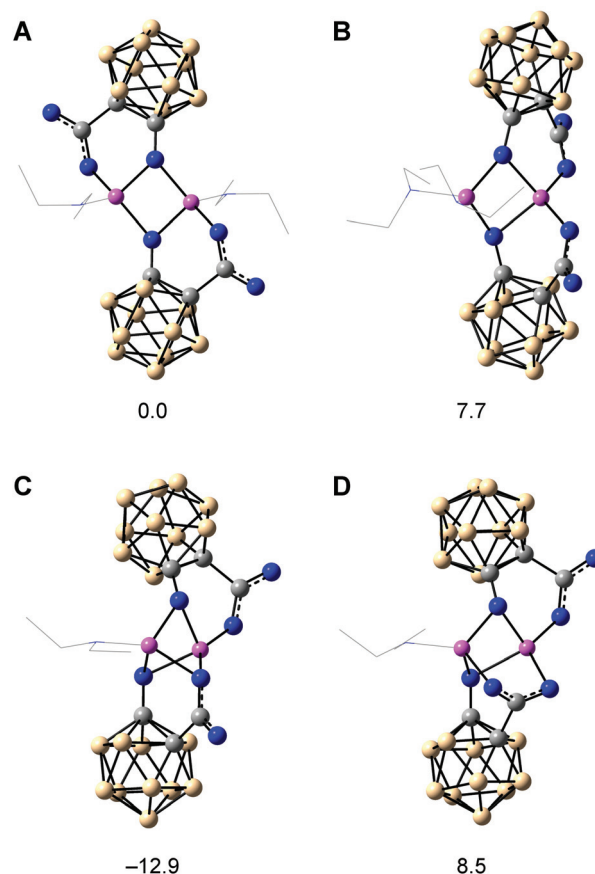
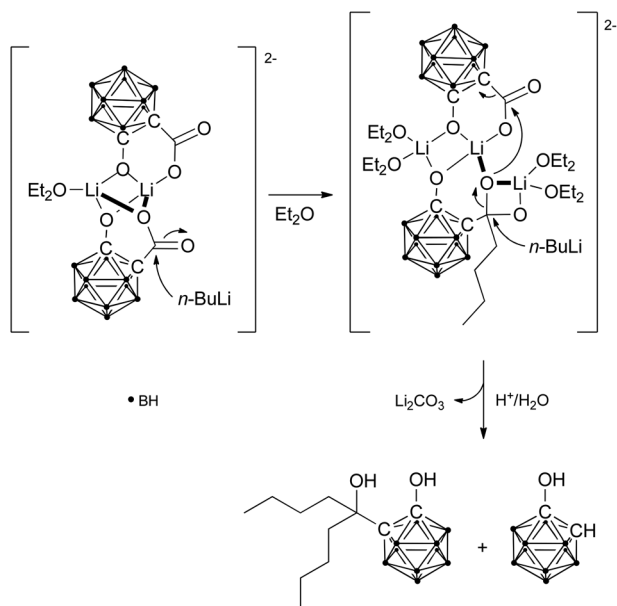


Fig. 2 Optimised geometries of the dimers of lithiated 1-hydroxy-1,2-dicarba-*closo*-dodecaboranyl-2-carboxylic acid. Diethyl ether molecules are shown as a wireframe representation. All geometries are minima, and the relative Gibbs free energies (ΔG) are given below each dimer in kJ mol⁻¹. The calculated free energy of an additional diethyl ether molecule has been included in the ΔG values of C and D.



Scheme 2 Proposed mechanism for the reduction of 1-hydroxy-1,2-dicarba-*closo*-dodecaboranyl-2-carboxylic acid by *n*-BuLi involving an intra-aggregate oxygen transfer.

However, as only bridging hydroxyl groups were found in the crystal structure, dimers with chelating carboxylato groups were not considered in the theoretical calculations. The modelled dimer C is thermodynamically favoured over the other dimeric structures, and the orientation of the carboxylato groups in this dimer may facilitate the proposed intra-aggregate oxygen transfer.

Nucleophilic attack of *n*-BuLi may first lead to a ketone-like intermediate (Scheme 2). The attack could either take place at the non-bridging or the bridging carboxylato group of C. An attack at the first could be sterically favoured as a result of less hindrance. However, nucleophilic attack at the bridging carboxylato group could be more favoured due to the higher electrophilicity of the carbon centre. Based on this attack a possible reaction intermediate was calculated (Fig. 3). Furthermore, the carboxylato bridge can be cleaved by the incoming lithium ion, and resulting free coordination sites could be filled by solvent molecules. As geminal diolates are generally unstable, oxygen transfer to the carboxylato group of the other cluster may then occur rapidly with simultaneous attack of the



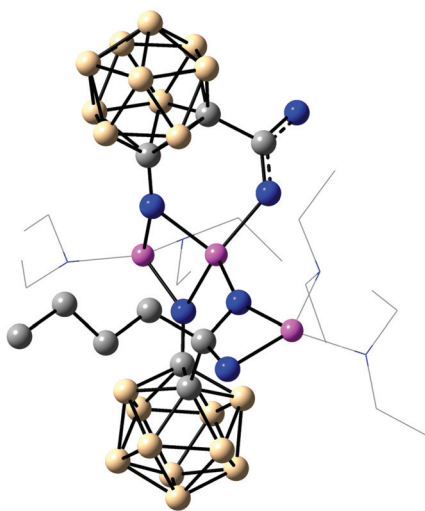


Fig. 3 Optimised geometry of a plausible intermediate formed after nucleophilic attack of *n*-BuLi at lithiated 1-hydroxy-1,2-dicarba-*closo*-dodecaboranyl-2-carboxylic acid (dimer C). Diethyl ether molecules are shown as a wireframe representation.

second *n*-BuLi molecule at this intermediate to generate the tertiary alkoxide and lithium carbonate. The bridging Li coordination of one of the carboxylate oxygen atoms may facilitate its intra-aggregate transfer resulting in the elimination of lithium carbonate.

Besides facilitating oxygen transfer from the geminal diolate, Li coordination by the alkoxide and the deprotonated hydroxyl group may also stabilise the intermediates and thereby suppress cleavage of the *exo*-polyhedral C–C bond during reduction of the acid (reaction II) or the ketone (reaction IV) and explain the different ratios between tertiary alcohol and cleavage product (Table 1). Furthermore, the *exo*-dative π bonding of the deprotonated hydroxyl group increases the electron density within the cluster and may thus also contribute to stabilisation of the intermediate alkoxides.

The importance of the hydroxyl group for the formation of tertiary alcohols was further corroborated by replacing the hydroxyl with an amino group. Similar to oxygen, nitrogen bases strongly coordinate at lithium cations (HSAB principle)¹⁰ and also show *exo*-dative π bonding.^{8e} Thus, an amino group at the carbaborane cluster may enable an analogous reaction sequence. However, as primary amines readily react with carbonyl groups and attack the carbaborane cluster resulting in deboronation,¹¹ a protecting group had to be used. The *tert*-butoxycarbonyl (Boc) group was chosen as a rather small group which is stable under alkaline conditions, also in the presence of *n*-BuLi, and can be directly introduced during the synthesis of the carbaboranyl amine from an intermediate isocyanate.¹² Carboxylation of the Boc-protected carbaboranyl amine has already been reported.¹² However, reaction with *n*-BuLi did not yield the tertiary alcohol but only led to cleavage of the *exo*-polyhedral C–C bond (reaction V, Table 1). This be-

haviour resembles the previously reported decarboxylation of 1-methoxy-1,2-dicarba-*closo*-dodecaboranyl-2-carboxylic acid (reaction VI, Table 1)^{2g} and indicates that the Boc group probably impedes coordination of lithium ions in a bridging manner.

Conclusions

Similar to 1-hydroxy-1,2-dicarba-*closo*-dodecaboranyl-2-carboxylic acid, reaction of the corresponding ketone with *n*-BuLi led to formation of a tertiary alcohol. This supports the proposed mechanism of reduction of the carboxylic acid proceeding *via* a ketone-like intermediate, thereby directly yielding the tertiary alcohol upon work-up. Formation of the proposed dimeric contact ion pairs was supported by a crystal structure of the lithium salt of salborin. The observed centrosymmetric Li₂O₂ arrangement with bridging deprotonated hydroxyl groups may enable intra-aggregate oxygen transfer during reduction of the acid. The proposed mechanism was also supported by theoretical calculations. Furthermore, the observed intramolecular coordination of lithium by the deprotonated hydroxyl and carbonyl group of the cluster may stabilise the intermediate alkoxides and reduce *exo*-polyhedral C–C bond cleavage during reduction of the acid and ketone. In addition, the strong *exo*-dative π -bonding of the deprotonated hydroxyl group may further stabilise the intermediate alkoxides. The importance of the hydroxyl group for this unique reactivity was further confirmed by reactions of derivatives without a second functional group (R = H) or with a bulky functional group (R = OMe, NHBoc) hindering coordination at the lithium cation, thus resulting in quantitative cleavage of the *exo*-polyhedral C–C bond upon treatment with *n*-BuLi.

Experimental

General

Reactions were carried out under a nitrogen atmosphere in anhydrous solvents, which were purified with an MBRAUN Solvent Purification System MB SPS-800 Series. Chemicals were used as purchased. Thin-layer chromatography (TLC) was performed on pre-coated glass plates (0.25 mm, silica gel 60 F₂₅₄); visualisation of carbaborane compounds on TLC plates was achieved by treatment with a solution of PdCl₂ (1% in MeOH) and gentle heating. Column chromatography was carried out with silica gel (0.035–0.070 mm, 60 Å). NMR spectra were recorded on Bruker AVANCE DRX 400 and AVANCE III HD 400 spectrometers. The chemical shifts of ¹H, ¹¹B and ¹³C NMR spectra are reported in parts per million at 400.13, 128.38, and 100.63 MHz, respectively, with tetramethylsilane as internal standard and referencing to the unified scale.¹³ Mass spectra were recorded on an ESQUIRE 3000 Plus Bruker-Daltonics ESI spectrometer.



Synthetic procedures

1-Acetyl-1,2-dicarba-closo-dodecaborane(12) (cf.¹⁴). 1,2-Dicarba-closo-dodecaborane(12) (0.50 g, 3.5 mmol) was dissolved in Et₂O (50 mL) and the solution was cooled to 0 °C. *n*-BuLi (2.2 mL, 1.5 M in *n*-hexane, 3.3 mmol, 0.95 equiv.) was added and the suspension was stirred for 90 min at 0 °C. The formed white precipitate was filtered off, washed with Et₂O and re-suspended in Et₂O (50 mL). Acetyl chloride (0.6 mL, 8.7 mmol, 2.5 eq.) was added at 0 °C and the suspension was stirred for 2 h at room temperature. Aqueous HCl (30 mL, 1 M) was added and the phases were separated. The aqueous phase was extracted with *n*-hexane, the combined organic layers were dried over MgSO₄, filtered and the solvent was removed under reduced pressure. Unconsumed starting material and the disubstituted analogue were separated from the product by column chromatography (*n*-hexane–ethyl acetate 10 : 1) to yield a pale yellow oil (0.05 g, 8%): ¹H NMR (CDCl₃): δ = 1.40–3.10 (10H, br, C₂B₁₀H₁₀), 2.43 (3H, s, CH₃), 4.16 (1H, s, CH). ¹¹B {¹H} NMR (CDCl₃): δ = –13.2 (2B, s), –12.8 (2B, s), –11.8 (2B, s), –8.7 (2B, s), –2.5 (2B, s). MS (ESI(–), acetone): *m/z* (%) 175.2 (100) [M – B][–]; the observed isotopic pattern was in agreement with the calculated one.

Reduction of 1-acetyl-1,2-dicarba-closo-dodecaborane(12). 1-Acetyl-1,2-dicarba-closo-dodecaborane(12) (0.045 g, 0.24 mmol) was dissolved in Et₂O (8 mL) and the solution was cooled to 0 °C. *n*-BuLi (0.40 mL, 1.5 M in *n*-hexane, 2.5 equiv.) was added and the solution was stirred for 1 h at room temperature. The solvent was removed under reduced pressure, aqueous HCl (15 mL, 1 M) was added and the solution was stirred for 1 h at room temperature. The solution was extracted with *n*-hexane, the combined organic phases were dried over MgSO₄, filtered and all volatile material was removed under reduced pressure. The product crystallised from the obtained oil as colourless crystals, which were identified as 1,2-dicarba-closo-dodecaborane(12) by ¹H and ¹¹B NMR. ¹H NMR (CDCl₃): δ = 1.00–3.10 (10H, br, C₂B₁₀H₁₀), 3.55 (2H, s, CH). ¹¹B {¹H} NMR (CDCl₃): δ = –14.8 (2B, s), –13.6 (4B, s), –9.2 (2B, s), –2.3 (2B, s).

1-Hydroxy-1,2-dicarba-closo-dodecaborane(12). The compound was synthesised according to the literature.¹⁵ ¹H NMR (CDCl₃): δ = 1.30–3.30 (10H, br, C₂B₁₀H₁₀), 3.97 (1H, s, CH), 4.63 (1H, br s, OH); the chemical shift of the signal of the acidic OH proton was found to be concentration-dependent. ¹¹B {¹H} NMR (CDCl₃): δ = –14.5 (2B, s), –12.1 (7B, s), –3.7 (1B, s).

1-Acetoxy-2-acetyl-1,2-dicarba-closo-dodecaborane(12). 1-Hydroxy-1,2-dicarba-closo-dodecaborane(12) (0.10 g, 0.6 mmol) was dissolved in Et₂O (15 mL) and the solution was cooled to 0 °C. *n*-BuLi (1.1 mL, 1.5 M in *n*-hexane, 1.6 mmol, 2.5 equiv.) was added and the suspension was stirred for 90 min at 0 °C. The formed white precipitate was filtered off, washed with Et₂O and re-suspended in Et₂O (15 mL). Acetyl chloride (0.14 mL, 1.9 mmol, 3.0 equiv.) was added at 0 °C and the suspension was stirred for 3 h at room temperature. Aqueous HCl (15 mL, 1 M) was added and the phases were

separated. The aqueous phase was extracted with *n*-hexane; the combined organic phases were dried over MgSO₄, filtered and the solvent was removed under reduced pressure. Unconsumed starting material and 1-acetoxy-1,2-dicarba-closo-dodecaborane(12) were separated from the product by column chromatography (*n*-hexane–ethyl acetate 10 : 1) to yield a colourless oil (0.02 g, 13%): ¹H NMR (CDCl₃): δ = 1.30–3.60 (10H, br, C₂B₁₀H₁₀), 2.10 (3H, s, CH₃), 2.47 (3H, s, CH₃). ¹¹B NMR (CDCl₃): δ = –13.5 (2B, d, ¹J_{B,H} = 184 Hz), –12.0 (2B, d, ¹J_{B,H} = 159 Hz), –10.8 (4B, d, ¹J_{B,H} = 168 Hz), –7.8 (1B, d, ¹J_{B,H} = 150 Hz), –3.2 (1B, d, ¹J_{B,H} = 154 Hz). ¹³C {¹H} NMR (CDCl₃): δ = 20.9 (CH₃), 28.8 (CH₃), 79.8 (C_{cluster}), 96.3 (C_{cluster}), 165.2 (CO), 189.7 (CO). MS (ESI(–), acetone): *m/z* (%) 234.1 (100) [M – B][–], 191.1 (30) [M – B – COCH₃][–]; the observed isotopic patterns were in agreement with the calculated ones.

1-Hydroxy-2-acetyl-1,2-dicarba-closo-dodecaborane(12). Deprotection of 1-acetoxy-2-acetyl-1,2-dicarba-closo-dodecaborane(12) was performed under acidic conditions (not under nitrogen atmosphere). Acetone as solvent led to a fast decapping of the cluster. Therefore, the reaction was carried out in CHCl₃, in which, however, deacylation proceeded very slowly, so that acetone was added for the final hours of the reaction. Deprotection in CHCl₃ was also tested under microwave conditions at 70 °C; however, even after 8 h the reaction mixture mainly contained the starting material.

1-Acetoxy-2-acetyl-1,2-dicarba-closo-dodecaborane(12) (0.25 mg) was dissolved in CHCl₃ (35 mL). HCl (37% *w/w*, 3.5 mL) was added and the reaction mixture was stirred for 2.5 d at room temperature. Acetone (10 mL) was added and the reaction mixture was stirred for another 5 h. The phases were separated and the aqueous phase was extracted with *n*-hexane. The combined organic phases were dried over MgSO₄, filtered and the solvent was removed under reduced pressure. Unconsumed starting material was separated from the product by column chromatography (*n*-hexane–ethyl acetate 10 : 1 → 0 : 1) to yield an orange oil (0.05 g, 24%): ¹H NMR (CDCl₃): δ = 1.30–3.00 (10H, br, C₂B₁₀H₁₀), 2.51 (3H, s, COCH₃), 2.65 (1H, br s, OH). ¹¹B {¹H} NMR (CDCl₃): δ = –18.1 (1B, s), –14.5 (2B, s), –11.8 (5B, br), –9.5 (1B, s), –3.8 (1B, s). MS (ESI(–), CH₃CN): *m/z* (%) 201.1 (100) [M – H][–]; the observed isotopic pattern was in agreement with the calculated one.

Reduction of 1-hydroxy-2-acetyl-1,2-dicarba-closo-dodecaborane(12). The reaction was carried out as described for the reduction of 1-acetyl-1,2-dicarba-closo-dodecaborane(12). The obtained yellow oil was analysed by ¹H and ¹¹B NMR and was found to contain 1-hydroxy-1,2-dicarba-closo-dodecaborane(12) (ca. 80%) and 5-[1-hydroxy-1,2-dicarba-closo-dodecaboranyl-(12)]-nonan-5-ol (ca. 20%): ¹¹B {¹H} NMR (CDCl₃): δ = –13.0 (4B, s), –11.2 (4B, s), –4.8 (2B, s).^{2g}

1-[(*tert*-Butyloxycarbonyl)amino]-1,2-dicarba-closo-dodecaborane(12). The compound was synthesised according to the literature.⁴ ¹H NMR (CDCl₃): δ = 1.30–3.50 (10H, br, C₂B₁₀H₁₀), 1.44 (9H, s, CH₃), 4.61 (1H, s, CH), 5.65 (1H, s, NH). ¹¹B NMR (CDCl₃): δ = –14.0 (2B, d, ¹J_{B,H} = 156 Hz), –11.7 (2B, d, ¹J_{B,H} = 210 Hz), –11.0 (4B, d, ¹J_{B,H} = 151 Hz), –7.7 (1B, d, ¹J_{B,H} = 152 Hz), –3.8 (1B, d, ¹J_{B,H} = 143 Hz). MS (ESI(–), CHCl₃–CH₃OH):



m/z (%) 258.1 (100) $[M - H]^-$, 249.1 (48) $[M - B]^-$; the observed isotopic patterns were in agreement with the calculated ones.

In situ formation of 1-[(*tert*-butyloxycarbonyl)amino]-1,2-dicarba-*closo*-dodecaboranyl(12)-2-carboxylic acid and reaction with *n*-BuLi. The carbaboranyl carboxylic acid was prepared *in situ* by deprotonation of the cluster C–H group with *n*-BuLi followed by reaction with gaseous CO₂; reduction of the acid was performed by a second addition of *n*-BuLi.^{2g}

1-[(*tert*-Butyloxycarbonyl)amino]-1,2-dicarba-*closo*-dodecaborane(12) (0.10 g, 0.4 mmol) was dissolved in Et₂O (15 mL) and the solution was cooled to 0 °C. *n*-BuLi (0.64 mL, 1.5 M in *n*-hexane, 2.5 equiv.) was added and the solution was stirred for 90 min at 0 °C. The solvent was removed, and the residue was washed with *n*-hexane (3 × 15 mL) and redissolved in Et₂O (40 mL). CO₂ was passed through the solution for 10 min at 0 °C. N₂ was passed through the solution for 10 min at 0 °C to remove any dissolved CO₂; then *n*-BuLi (0.64 mL, 1.5 M in *n*-hexane, 2.5 equiv.) was added. The solution was stirred for 1 h at room temperature. The solvent was removed under reduced pressure; aqueous HCl (30 mL, 1 M) was added and the solution was stirred for 1 h at room temperature. The solution was extracted with *n*-hexane and the combined organic phases were washed with water, dried over MgSO₄, filtered and all volatile material was removed under reduced pressure. The product crystallised as colourless, needle-like crystals, which were analysed by ¹H, ¹¹B NMR spectroscopy and MS and were identified as 1-[(*tert*-butyloxycarbonyl)amino]-1,2-dicarba-*closo*-dodecaborane(12).

Crystallography

Crystallisation of the lithium salt of salborin (cf.^{2g}). 1-Hydroxy-1,2-dicarba-*closo*-dodecaborane(12) was dissolved in Et₂O and the solution was cooled to 0 °C. *n*-BuLi (1.58 M in *n*-hexane, 2.5 equiv.) was added and the suspension was stirred for 90 min at 0 °C. The formed white precipitate was filtered off, washed with Et₂O and re-suspended in Et₂O. CO₂ was passed through the solution for 10 min at 0 °C while the precipitate dissolved. The solution containing the lithium salt of salborin was transferred to an NMR tube (non-inert conditions). After 1 d a few drops of tmeda were added to the solution, resulting in the formation of a precipitate. The precipitate was dissolved by the addition of *n*-hexane. The lithium salt crystallised from the solution within 8 weeks at room temperature as colourless, needle-like crystals.

X-ray crystallography. Data for the lithium salt of salborin were collected on an Oxford Diffraction CCD Gemini-S diffractometer (Agilent Technologies) with MoK_α radiation (λ = 0.71073 Å) and ω -scan rotation. Data reduction was performed with CrysAlis Pro¹⁶ including the programme SCALE3 ABSPACK for empirical absorption correction. The structure was solved with direct methods (SIR-92)¹⁷ and the refinement of all non-hydrogen atoms was performed with SHELXL-2013.¹⁸ All H atoms were located on difference Fourier maps calculated at the final stage of the structure refinement. CCDC 1039317 contains the supplementary crystallographic data for this paper.

Theoretical calculations

All calculations have been carried out with the Gaussian 09 programme package¹⁹ at the M06-2X/6-31+G(d,p) level of theory²⁰ using a pruned (99 590) grid. Explicit solvent molecules (diethyl ether) have been included wherever appropriate and the SMD solvent model (diethyl ether)²¹ has been used for all calculations.

Acknowledgements

Financial support from the Fonds der Chemischen Industrie (FCI, doctoral grant to W.N.) and the Graduate School “Leipzig School of Natural Sciences – Building with Molecules and Nano-objects (BuildMoNa)” is gratefully acknowledged.

Notes and references

- 1 R. N. Grimes, *Carboranes*, Elsevier, Amsterdam, 2nd edn, 2011.
- 2 (a) L. I. Zakharkin, *Dokl. Akad. Nauk SSSR*, 1965, **162**, 817; (b) V. I. Stanko, A. I. Klimova, Y. A. Chapovskii and T. P. Klimova, *Zh. Obshch. Khim.*, 1966, **36**, 1779; (c) L. I. Zakharkin and A. I. L'vov, *J. Organomet. Chem.*, 1966, **5**, 313; (d) L. I. Zakharkin and A. I. L'vov, *Zh. Obshch. Khim.*, 1967, **37**, 1217; (e) A. I. L'vov and L. I. Zakharkin, *Izv. Akad. Nauk SSSR, Ser. Khim.*, 1967, 2653; (f) R. N. Grimes, *Carboranes*, Elsevier, Amsterdam, 2nd edn, 2011, ch. 9.9, pp. 427–430 and ch. 9.10, pp. 436–437; (g) W. Neumann, M. Hiller, P. Lönnecke and E. Hey-Hawkins, *Dalton Trans.*, 2014, **43**, 4935.
- 3 (a) L. I. Zakharkin and A. I. L'vov, *Zh. Obshch. Khim.*, 1967, **37**, 742; (b) V. I. Stanko, V. A. Brattsev, N. E. Al'perovich and N. S. Titova, *Zh. Obshch. Khim.*, 1968, **38**, 1056; (c) J. L. Maurer, F. Berchier, A. J. Serino, C. B. Knobler and M. F. Hawthorne, *J. Org. Chem.*, 1990, **55**, 838.
- 4 J. J. Schaeck and S. B. Kahl, *Inorg. Chem.*, 1999, **38**, 204.
- 5 M. Scholz, G. N. Kaluderović, H. Kommera, R. Paschke, J. Will, W. S. Sheldrick and E. Hey-Hawkins, *Eur. J. Med. Chem.*, 2011, **46**, 1131.
- 6 (a) V. I. Stanko, T. V. Klimova and I. P. Beletskaya, *Dokl. Akad. Nauk SSSR, Ser. Khim.*, 1974, **216**, 329; (b) V. I. Stanko and T. V. Klimova, *Zh. Obshch. Khim.*, 1977, **47**, 2017; (c) D. A. Brown, H. M. Colquhoun, J. A. Daniels, J. A. H. MacBride, I. R. Stephenson and K. Wade, *J. Mater. Chem.*, 1992, **2**, 793; (d) H. Nakamura, K. Aoyagi and Y. Yamamoto, *J. Org. Chem.*, 1997, **62**, 780; (e) H. Nakamura, K. Aoyagi and Y. Yamamoto, *J. Organomet. Chem.*, 1999, **574**, 107.
- 7 (a) A. J. Welch, U. Venkatasubramanian, G. M. Rosair, D. Ellis and D. J. Donohoe, *Acta Crystallogr., Sect. C: Cryst. Struct. Commun.*, 2001, **57**, 1295; (b) U. Venkatasubramanian, D. Ellis, G. M. Rosair and A. J. Welch, *Acta Crystallogr., Sect. C: Cryst. Struct. Commun.*, 2003, **59**, o559; (c) U. Venkatasubramanian,



- D. Ellis, G. M. Rosair and A. J. Welch, *Acta Crystallogr., Sect. C: Cryst. Struct. Commun.*, 2003, **59**, o586; (d) U. Venkatasubramanian, D. J. Donohoe, D. Ellis, B. T. Giles, S. A. Macgregor, S. Robertson, G. M. Rosair, A. J. Welch, A. S. Batsanov, L. A. Boyd, R. C. B. Copley, M. A. Fox, J. A. K. Howard and K. Wade, *Polyhedron*, 2004, **23**, 629.
- 8 (a) D. A. Brown, W. Clegg, H. M. Colquhoun, J. A. Daniels, I. R. Stephenson and K. Wade, *J. Chem. Soc., Chem. Commun.*, 1987, 889; (b) R. Coult, M. A. Fox, W. R. Gill and K. Wade, *Polyhedron*, 1992, **11**, 2717; (c) I. Zharov, A. Saxena, J. Michl and R. D. Miller, *Inorg. Chem.*, 1997, **36**, 6033; (d) J. Llop, C. Viñas, J. M. Oliva, F. Teixidor, M. A. Flores, R. Kivekas and R. Sillanpää, *J. Organomet. Chem.*, 2002, **657**, 232; (e) L. A. Boyd, W. Clegg, R. C. B. Copley, M. G. Davidson, M. A. Fox, T. G. Hibbert, J. A. K. Howard, A. Mackinnon, R. J. Peace and K. Wade, *Dalton Trans.*, 2004, 2786.
- 9 (a) M. G. Davidson, T. G. Hibbert, J. A. K. Howard, A. Mackinnon and K. Wade, *Chem. Commun.*, 1996, 2285; (b) R. J. Blanch, M. Williams, G. D. Fallon, M. G. Gardiner, R. Kaddour and C. L. Raston, *Angew. Chem., Int. Ed.*, 1997, **36**, 504; (c) M. J. Hardie and C. L. Raston, *CrystEngComm*, 2001, **3**, 162; (d) M. A. Fox and A. K. Hughes, *Coord. Chem. Rev.*, 2004, **248**, 457; (e) A. R. Turner, H. E. Robertson, K. B. Borisenko, D. W. H. Rankin and M. A. Fox, *Dalton Trans.*, 2005, 1310.
- 10 R. G. Pearson, *J. Am. Chem. Soc.*, 1963, **85**, 3533.
- 11 L. I. Zakharkin and V. N. Kalinin, *Tetrahedron Lett.*, 1965, **6**, 407.
- 12 R. A. Kasar, G. M. Knudsen and S. B. Kahl, *Inorg. Chem.*, 1999, **38**, 2936.
- 13 R. K. Harris, E. D. Becker, S. M. Cabral de Menezes, R. Goodfellow and P. Granger, *Pure Appl. Chem.*, 2001, **73**, 1795.
- 14 E. Hey-Hawkins and M. Scholz, *Ger. Appl.*, DE 102007026701 A1, 2008; E. Hey-Hawkins and M. Scholz, *PCT Int. Appl.*, WO 2008145733 A2, 2008.
- 15 K. Ohta, T. Goto, H. Yamazaki, F. Pichierri and Y. Endo, *Inorg. Chem.*, 2007, **46**, 3966.
- 16 *CrysAlis Pro*, Agilent Technologies, Version 1.171.36.28.
- 17 SIR92: A. Altomare, G. Casciaro, C. Giacovazzo and A. Guagliardi, *J. Appl. Crystallogr.*, 1993, **26**, 343.
- 18 SHELXL-2013: G. M. Sheldrick, *Acta Crystallogr., Sect. A: Fundam. Crystallogr.*, 2008, **64**, 112.
- 19 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery Jr., J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, O. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski and D. J. Fox, *Gaussian 09, Revision A.02*, Gaussian, Inc., Wallingford CT, 2009.
- 20 (a) Y. Zhao and D. Truhlar, *Theor. Chem. Acc.*, 2008, **120**, 215; (b) Y. Zhao and D. Truhlar, *Theor. Chem. Acc.*, 2008, **119**, 525; (c) Y. Zhao and D. G. Truhlar, *Acc. Chem. Res.*, 2008, **41**, 157; (d) W. J. Hehre, R. Ditchfield and J. A. Pople, *J. Chem. Phys.*, 1972, **56**, 2257; (e) R. Ditchfield, W. J. Hehre and J. A. Pople, *J. Chem. Phys.*, 1971, **54**, 724.
- 21 A. V. Marenich, C. J. Cramer and D. G. Truhlar, *J. Phys. Chem. B*, 2009, **113**, 6378.

