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## Development of a functionally separated D- $\pi$ -A fluorescent dye with a pyrazyl group as an electron-accepting group for dye-sensitized solar cells†

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A functionally separated D- $\pi$ -A dye **OUK-3** with a pyrazyl group as an electron-withdrawing anchoring group and a carboxyl group as an additional anchoring group has been newly developed as a photosensitizer for dye-sensitized solar cells. The optical and electrochemical properties, adsorption states on TiO<sub>2</sub> nanoparticles, and photovoltaic performances in dye-sensitized solar cells (DSSCs) were investigated. It was found that the maximum adsorption amount of dye adsorbed on the TiO<sub>2</sub> electrode for **OUK-3** is 3 times as much as that of the D- $\pi$ -A dye sensitizer **OUK-1** with a pyrazyl group as an electron-withdrawing anchoring group. Moreover, this work revealed that the DSSC based on the dye **OUK-3** exhibits a higher photovoltage ( $V_{oc}$ ) value than the DSSC based on the dye **OUK-1**. On the basis of the FTIR spectra of the dyes adsorbed on TiO<sub>2</sub> nanoparticles and the electrochemical impedance spectroscopy (EIS) analysis of DSSCs based on the two dyes **OUK-1** and **OUK-3**, the differences of photovoltaic performances between the two dyes are discussed by taking into account the adsorption states of the dyes adsorbed on the TiO<sub>2</sub> surface. This work demonstrates that functionally separated D- $\pi$ -A dye sensitizers can achieve effective surface coverage of the TiO<sub>2</sub> electrode due to their high adsorption ability onto the TiO<sub>2</sub> electrode, leading to not only the improvement of light-harvesting efficiency (LHE), but also an increase in the number of injected electrons in the CB of TiO<sub>2</sub>, which is responsible for the higher  $V_{oc}$  value of functionally separated D- $\pi$ -A dye sensitizers.

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## Introduction

Dye-sensitized solar cells (DSSCs) employing dye-adsorbed TiO<sub>2</sub> electrodes have attracted much attention from chemists, physicists, and engineers because of enormous scientific interest in their construction and operational principles, and low cost of production since Grätzel and co-workers produced high-performance DSSCs based on a Ru-complex dye in 1991.<sup>1-10</sup> Much effort in molecular design and development of efficient dye sensitizers have been made to further improve the photovoltaic performances of DSSCs so far. In particular, donor-acceptor  $\pi$ -conjugated (D- $\pi$ -A) dye sensitizers having both electron-donating (D) and electron-accepting (A) groups linked by  $\pi$ -conjugated bridges would be expected to be one of the most promising classes of organic dye sensitizers, because

the wide variety of structures and facile modification provides potential for molecular design, with the introduction of substituents onto the chromophore skeletons allowing easy control not only of their photophysical and electrochemical properties (HOMO and LUMO levels), but also of their stereochemical structures.<sup>4-10</sup> Consequently, a new type of D- $\pi$ -A dye sensitizers with 2-(1,1-dicyanomethylene)rhodanine,<sup>11</sup> pyridine,<sup>12,13</sup> and 8-hydroxyquinoline<sup>14</sup> as an electron-withdrawing anchoring group, as well as conventional D- $\pi$ -A dye sensitizers with a carboxyl group have been designed and developed so far. Many research groups demonstrated that the conventional D- $\pi$ -A dye sensitizers with a carboxyl group are adsorbed on the TiO<sub>2</sub> film through the bidentate bridging linkage between the carboxyl group of the dye and Brønsted acid sites (surface-bound hydroxyl groups, Ti-OH) on the TiO<sub>2</sub> surface.<sup>4-10</sup> On the other hand, we have demonstrated that dye sensitizers with a pyridyl group are predominantly adsorbed on the TiO<sub>2</sub> through coordinate bonding between the pyridyl group of the dye and the Lewis acid site (exposed Ti<sup>n+</sup> cations) on the TiO<sub>2</sub> surface.<sup>12</sup> It was found that D- $\pi$ -A dye sensitizers with a pyridyl group can inject electrons efficiently to the conduction band (CB) of the TiO<sub>2</sub> electrode through coordinate bonding,

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rather than conventional D- $\pi$ -A dye sensitizers with carboxyl groups. Recently, DSSCs based on porphyrin dyes and D- $\pi$ -A dyes with pyridyl groups have been reported by some research groups.<sup>13</sup>

Moreover, to seek an epoch-making molecular design of D- $\pi$ -A dye sensitizers for high-performance DSSCs, some research groups designed and synthesized the functionally separated D- $\pi$ -A dye sensitizers with a pyridyl group,<sup>12a,c</sup> a cyano group,<sup>15,16</sup> a nitro group<sup>17</sup> or an aldehyde<sup>18</sup> as an electron-accepting group and a carboxyl group as an additional anchoring group. They investigated the effects of the interaction between the electron-accepting group of the dye and the TiO<sub>2</sub> surface on the photovoltaic performances of DSSCs. It was revealed that the biggest advantage of the functionally separated D- $\pi$ -A dye sensitizers is that one can introduce both strong electron-withdrawing groups into the electron acceptor moiety and an additional anchoring group, leading to not only high adsorption ability onto the TiO<sub>2</sub> film, but also the bathochromic shift and broadening of the photoabsorption property and efficient electron injection from the dye to the CB of the TiO<sub>2</sub> electrode. For example, Sun and co-workers designed and synthesized efficient near infrared D- $\pi$ -A dye sensitizers, with strong electron-withdrawing groups as the electron acceptors, in which the anchoring group (carboxyl group) in these dyes was separated from the electron acceptor unit (dicyanomethylene).<sup>15</sup> They also developed the functionally separated D- $\pi$ -A dye sensitizer with a nitro group as the electron-accepting group and a carboxyl group as the anchoring group. They demonstrated the effects of the interaction between the nitro group of the dye and the TiO<sub>2</sub> surface on the photovoltaic performances and UV-Vis absorption properties.<sup>17</sup> Thus, the functionally separated D- $\pi$ -A dye sensitizers would also be expected to be one of the most promising classes of organic dye sensitizers for DSSCs.

In this work, a functionally separated D- $\pi$ -A dye sensitizer **OUK-3** with a pyrazyl group as an electron-withdrawing anchoring group and a carboxyl group as an additional anchoring group has been newly developed. The optical and electrochemical properties, adsorption states on TiO<sub>2</sub> nanoparticles, and photovoltaic performances in dye-sensitized solar cells (DSSCs) were investigated (Scheme 1). It was found that the maximum adsorption amount of dyes adsorbed on the TiO<sub>2</sub> electrode for **OUK-3** is 3 times as much as that of the D- $\pi$ -A dye sensitizer **OUK-1**<sup>19</sup> with a pyrazyl group as an elec-

tron-withdrawing anchoring group. Moreover, this work revealed that the DSSC based on the dye **OUK-3** exhibits a higher photovoltage ( $V_{oc}$ ) value than the DSSC based on the dye **OUK-1**. On the basis of the FTIR spectra of the dyes adsorbed on TiO<sub>2</sub> nanoparticles and the electrochemical impedance spectroscopy (EIS) analysis of DSSCs based on the two dyes **OUK-1** and **OUK-3**, the differences of photovoltaic performances between the two dyes are discussed by taking into account the adsorption states of the dyes adsorbed on the TiO<sub>2</sub> surface.

## Results and discussion

### Synthesis

The synthesis of **OUK-1** has been reported elsewhere.<sup>19</sup> The synthetic pathway of **OUK-3** is shown in Scheme 2. The starting material **1**<sup>19</sup> was converted to stannyl compound **2** by treatment with *n*BuLi and then Me<sub>3</sub>SnCl. Compound **3** was prepared by the Stille coupling of **2** with 2-iodopyrazine. The reaction of **3** with ethyl 7-bromoheptanoate by using sodium hydride yielded the compound **4**, and then the hydrolysis of **4** with a base gave the functionally separated D- $\pi$ -A dye sensitizer **OUK-3**.

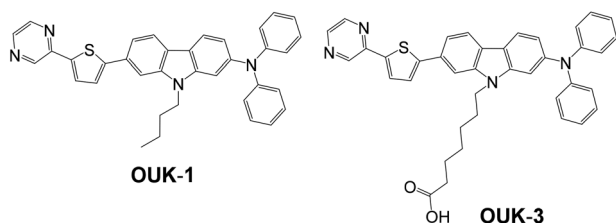
### Optical properties

The absorption and fluorescence spectra of **OUK-3** in 1,4-dioxane are shown in Fig. 1a (see Fig. S1a in ESI† for **OUK-1**). The two dyes show the absorption maximum ( $\lambda_{max}^{abs}$ ) at 402 nm, which is assigned to the intramolecular charge-transfer (ICT) excitation from the electron donor moiety (diphenylamino group) to the electron acceptor moiety (pyrazyl group). The molar extinction coefficient ( $\epsilon$ ) for the ICT band is 45 400 M<sup>-1</sup> cm<sup>-1</sup> for **OUK-1** and 48 300 M<sup>-1</sup> cm<sup>-1</sup> for **OUK-3**, respectively. The corresponding fluorescence bands for the two dyes occur at around 480 nm, and the fluorescence quantum yield ( $\Phi_f$ ) is 0.46 for **OUK-1** and 0.48 for **OUK-3**, respectively (Table 1).

The absorption spectrum of the dye **OUK-3** adsorbed on the TiO<sub>2</sub> film is shown in Fig. 1b (see Fig. S1b in ESI† for **OUK-1**). The absorption bands of **OUK-1** and **OUK-3** adsorbed on the TiO<sub>2</sub> film are broadened compared with those in 1,4-dioxane. However, the absorption peak wavelengths of **OUK-1** and **OUK-3** adsorbed on TiO<sub>2</sub> are similar to those in 1,4-dioxane. Thus, this result indicates that the two dyes form weak  $\pi$ -stacked aggregates on the TiO<sub>2</sub> surface.

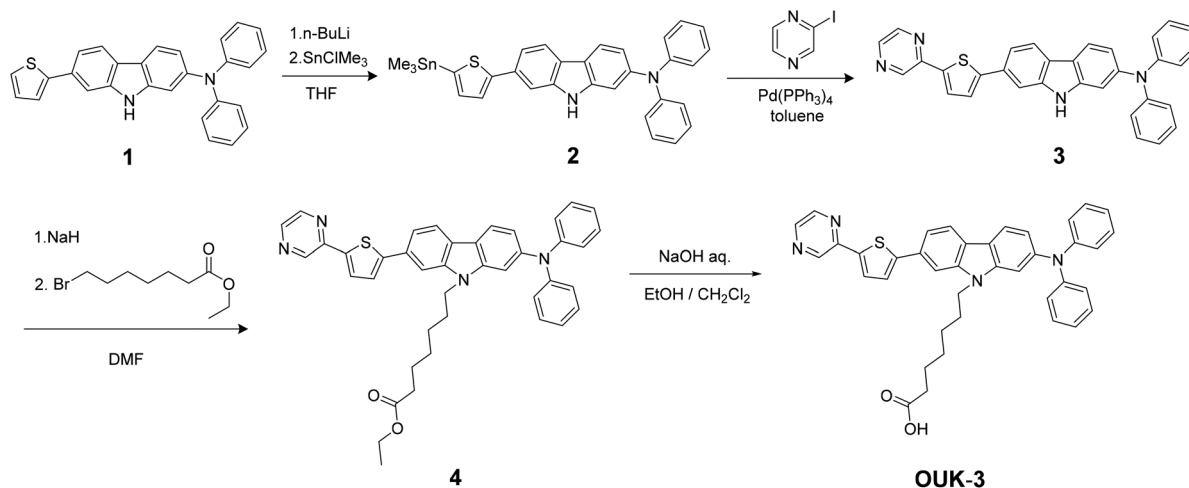
### Electrochemical properties

The electrochemical properties of **OUK-1** and **OUK-3** were determined by cyclic voltammetry (Fig. 2 and Table 1, see Fig. S2 in ESI† for **OUK-1**). The reversible oxidation waves for the two dyes were observed at 0.42 V vs. ferrocene/ferrocenium (Fc/Fc<sup>+</sup>). The corresponding reduction waves for the two dyes appeared at 0.35 V, thus showing that the oxidized states of the two dyes are stable. The HOMO energy level was evaluated from the half-wave potential for oxidation ( $E_{1/2}^{ox} = 0.39$  V for the two dyes). The HOMO energy level was 1.11 V vs. the normal

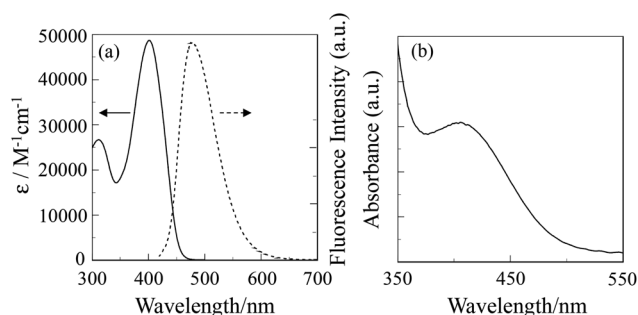


**Scheme 1** Chemical structures of the D- $\pi$ -A dye sensitizer **OUK-1** and the functionally separated D- $\pi$ -A dye sensitizer **OUK-3**.





**Scheme 2** Synthetic pathway of the functionally separated D- $\pi$ -A dye sensitizer **OUK-3**.



**Fig. 1** Absorption (–) and fluorescence (– –) spectra of **OUK-3** in 1,4-dioxane and (b) absorption spectrum of **OUK-3** adsorbed on the  $\text{TiO}_2$  film.

hydrogen electrode (NHE), thus indicating that the HOMO energy level is more positive than the  $\text{I}_3^-/\text{I}^-$  redox potential (0.4 V). This assures efficient regeneration of the oxidized dyes by electron transfer from the  $\text{I}_3^-/\text{I}^-$  in the electrolyte. The LUMO energy level was estimated from the  $E_{1/2}^{\text{ox}}$  and an intersection of the absorption and fluorescence spectra (445 nm; 2.79 eV for **OUK-1** and 443 nm; 2.80 eV for **OUK-3**). The LUMO energy levels of **OUK-1** and **OUK-3** were  $-1.68$  and  $-1.69$  V,

respectively. Evidently, the LUMO energy levels of **OUK-1** and **OUK-3** are higher than the energy level ( $E_{\text{cb}}$ ) of the CB of  $\text{TiO}_2$  ( $-0.5$  V), suggesting that an electron injection to the CB of  $\text{TiO}_2$  is thermodynamically feasible. Accordingly, the optical and electrochemical properties of **OUK-1** and **OUK-3** resemble each other very closely, showing that the effect of *N*-alkylation of the carbazole ring on the photophysical and electrochemical properties is negligible.

### Theoretical calculations

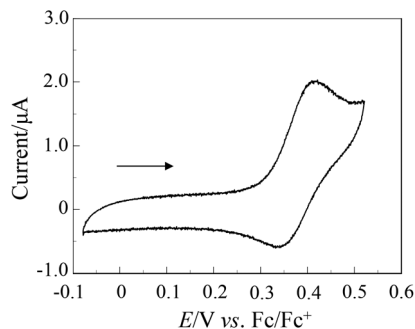
To examine the electronic structures of **OUK-1** and **OUK-3**, the molecular orbitals of the two dyes were calculated using density functional theory (DFT) at the B3LYP/6-31G(d, p) level. The DFT calculations indicate that for the two dyes the HOMOs were mostly localized on the diphenylamine-carbazole moiety containing a thiophene ring (Fig. 3, see Fig. S3 in ESI† for **OUK-1**). On the other hand, the LUMOs were mostly localized on the thienylpyridazine moiety. Accordingly, the DFT calculations reveal that dye excitations upon light irradiation induce a strong ICT from the diphenylamine-carbazole moiety to the pyridazine moiety.

**Table 1** Optical and electrochemical data, HOMO and LUMO energy levels, and DSSC performance parameters of **OUK-1** and **OUK-3**

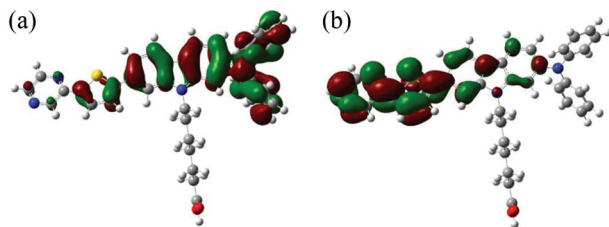
Dye	$\lambda_{\text{max}}^{\text{abs}}/\text{nm}$ ( $\epsilon \text{ M}^{-1} \text{ cm}^{-1}$ )	$\lambda_{\text{max}}^{\text{fl}}/\text{nm}$ ( $\Phi_{\text{f}}$ )	$E_{1/2}^{\text{ox}}/V$	HOMO <sup>c</sup> /V	LUMO <sup>c</sup> /V	Molecules <sup>d</sup> /cm <sup>-2</sup>	$J_{\text{sc}}^e/\text{mA cm}^{-2}$	$V_{\text{oc}}^e/\text{mV}$	FF <sup>e</sup>	$\eta^e$ (%)
<b>OUK-1</b>	402 (45400)	478 (0.46)	0.39	1.11	$-1.68$	$3.0 \times 10^{16}$	2.99	448	0.67	0.89
<b>OUK-3</b>	402 (48300)	477 (0.48)	0.39	1.11	$-1.69$	$9.6 \times 10^{16}$	4.48	550	0.64	1.58

<sup>a</sup> In 1,4-dioxane. Fluorescence quantum yields ( $\Phi_{\text{f}}$ ) were determined by using a calibrated integrating sphere system ( $\lambda_{\text{ex}} = 402$  for **OUK-1** and **OUK-3**). <sup>b</sup> Half-wave potentials for oxidation ( $E_{1/2}^{\text{ox}}$ ) vs.  $\text{Fc}/\text{Fc}^+$  were recorded in the DMF- $\text{Bu}_4\text{NClO}_4$  (0.1 M) solution. <sup>c</sup> vs. Normal hydrogen electrode (NHE). <sup>d</sup> The 9  $\mu\text{m}$  thick  $\text{TiO}_2$  electrode was immersed into a 1 mM dye solution in THF for 15 hours. The dye-coated film was immersed in a mixed solvent of THF-DMSO-NaOH aq. 1 M (5 : 4 : 1), which was used to determine the amount of dye molecules adsorbed onto the film by measuring the absorbance. <sup>e</sup> Under a simulated solar light (AM 1.5, 100  $\text{mW cm}^{-2}$ ).





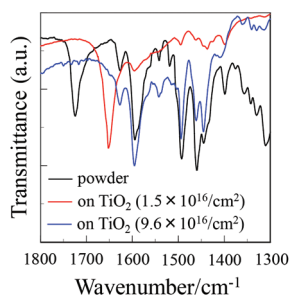
**Fig. 2** Cyclic voltammogram of **OUK-1** (0.1 mM) in DMF containing 0.1 M  $\text{Bu}_4\text{NClO}_4$  at a scan rate of  $10 \text{ mV s}^{-1}$ . The arrow denotes the direction of the potential scan.



**Fig. 3** (a) HOMO and (b) LUMO of **OUK-3** by the density functional theory (DFT) calculations at the B3LYP/6-31G(d, p) level.

### FTIR spectra

To elucidate the adsorption states of **OUK-1** and **OUK-3** on  $\text{TiO}_2$  nanoparticles, we measured the FTIR spectra of the dye powders and the dyes adsorbed on  $\text{TiO}_2$  nanoparticles (Fig. 4, see Fig. S4 in ESI† for **OUK-1**). For the dye powders of **OUK-3**, the  $\text{C}=\text{O}$  stretching band of the carboxyl group was observed at  $1724 \text{ cm}^{-1}$ . When the dye **OUK-3** was adsorbed on the  $\text{TiO}_2$  surface, the  $\text{C}=\text{O}$  stretching bands of the carboxyl group disappeared completely; this indicates the formation of a bidentate bridging linkage between the carboxyl group of the dye and the Brønsted acid site on the  $\text{TiO}_2$  surface. In addition, for the powders of the two dyes the  $\text{C}=\text{N}$  stretching band of the pyrazine ring was clearly observed at around  $1490 \text{ cm}^{-1}$ . Inter-



**Fig. 4** FTIR spectra of the dye powders and the dye ( $1.5 \times 10^{16}$  and  $9.6 \times 10^{16}$  molecules per  $\text{cm}^2$ ) adsorbed on  $\text{TiO}_2$  nanoparticles for **OUK-3**.

estingly, when the two dyes were adsorbed on the  $\text{TiO}_2$  surface, where the adsorption amount of the dye adsorbed on the  $\text{TiO}_2$  electrode is low ( $0.8 \times 10^{16}$  and  $1.5 \times 10^{16}$  molecules per  $\text{cm}^2$  for **OUK-1** and **OUK-3** respectively), the band at  $1490 \text{ cm}^{-1}$  disappeared completely and a new band appeared at around  $1650 \text{ cm}^{-1}$ , which indicates the formation of a pyrazinium ion with Brønsted acid sites on the  $\text{TiO}_2$  surface.<sup>19,20</sup> More interestingly, when the adsorption amount of the dye adsorbed on the  $\text{TiO}_2$  electrode is high ( $3.0 \times 10^{16}$  and  $9.6 \times 10^{16}$  molecules per  $\text{cm}^2$  for **OUK-1** and **OUK-3** respectively), the  $\text{C}=\text{N}$  stretching band at  $1490 \text{ cm}^{-1}$  is shifted by 6 and  $1 \text{ cm}^{-1}$  for **OUK-1** and **OUK-3**, respectively, to a higher wavenumber, that is, the band can be assigned to the hydrogen-bonded pyrazyl group to Brønsted acid sites on the  $\text{TiO}_2$  surface. These observations indicate that in the initial stage of dye adsorption both the dyes **OUK-1** and **OUK-3** are predominantly adsorbed on the  $\text{TiO}_2$  surface through the formation of a pyrazinium ion with Brønsted acid sites. However, both the dyes are adsorbed on the  $\text{TiO}_2$  surface through the hydrogen bonding at Brønsted acid sites as the dye adsorption progresses, although there still remain dye molecules adsorbed on the  $\text{TiO}_2$  surface through the formation of a pyrazinium ion with Brønsted acid sites. Consequently, the dye **OUK-1** is adsorbed on the  $\text{TiO}_2$  surface through formations of both hydrogen bonding of the pyrazyl group and the pyrazinium ion at Brønsted acid sites on the  $\text{TiO}_2$  surface. On the other hand, the dye **OUK-3** is adsorbed on the  $\text{TiO}_2$  surface through not only the bidentate bridging linkage of the carboxyl group but also the formations of hydrogen bonding of the pyrazyl group and the pyrazinium ion at Brønsted acid sites on the  $\text{TiO}_2$  surface. Thus, these results indicate that effective surface coverage of the  $\text{TiO}_2$  electrode is achieved successfully by employing the functionally separated D- $\pi$ -A dyes with two functional groups (cyclic azine and carboxyl group) possessing bonding ability to two points on Brønsted acid sites on the  $\text{TiO}_2$  surface. Moreover, it is worth mentioning here that the functionally separated D- $\pi$ -A dye sensitizers with a pyridyl group and a carboxyl group as an additional anchoring group was adsorbed on the  $\text{TiO}_2$  surface through not only the formation of a bidentate bridging linkage between the carboxyl group of the dye and the Brønsted acid site on the  $\text{TiO}_2$  surface, but also the coordinate bonding between the pyridyl group of the dye and the Lewis acid site on the  $\text{TiO}_2$  surface.<sup>12c</sup> Our work suggest that the binding modes of D- $\pi$ -A dye sensitizers with cyclic azine on the  $\text{TiO}_2$  surface can be changed by controlling the basicity and the electron density of cyclic azine.

### Dye-sensitized solar cells

The DSSC was prepared using the dye-adsorbed  $\text{TiO}_2$  electrode ( $9 \mu\text{m}$ ), Pt-coated glass as a counter electrode, and an acetonitrile solution with iodine (0.05 M), lithium iodide (0.1 M), and 1,2-dimethyl-3-propylimidazolium iodide (0.6 M) as an electrolyte. The photocurrent-voltage ( $I$ - $V$ ) characteristics were measured under simulated solar light (AM 1.5,  $100 \text{ mW cm}^{-2}$ ). The  $I$ - $V$  curve and the incident photon-to-current conversion efficiency (IPCE) spectrum of DSSC based on the dye **OUK-3**



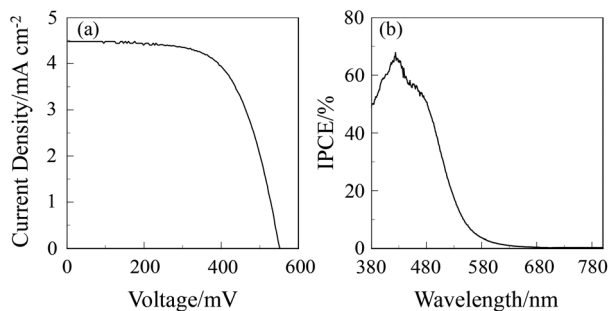


Fig. 5 (a)  $I$ - $V$  curve and (b) IPCE spectrum of DSSC based on **OUK-3**.

are shown in Fig. 5 (see Fig. S5 in ESI† for **OUK-1**). The photovoltaic performance parameters are collected in Table 1. The maximum adsorption amount of dyes adsorbed on  $\text{TiO}_2$  for **OUK-3** is 3 times as much as that of **OUK-1** ( $3.0 \times 10^{17}$  and  $9.6 \times 10^{16}$  molecules per  $\text{cm}^2$  for **OUK-1** and **OUK-3** respectively). The high adsorption ability of **OUK-3** relative to **OUK-1** is attributed to the formation of a bidentate bridging linkage between the carboxyl group of **OUK-3** and Brønsted acid sites on the  $\text{TiO}_2$  surface. The  $I$ - $V$  curves show that the short-circuit photocurrent density ( $J_{sc}$ ) and solar energy-to-electricity conversion yield ( $\eta$ ) of **OUK-3** ( $4.48 \text{ mA cm}^{-2}$  and 1.58%) are higher than that of **OUK-1** ( $2.99 \text{ mA cm}^{-2}$  and 0.89%) (Fig. 5a). The maximum IPCE value of **OUK-3** is 65% at 420 nm, which is higher than that (38% at 420 nm) of **OUK-1** (Fig. 5b). Moreover, it is worth mentioning here that the  $V_{oc}$  value of **OUK-3** (550 mV) is higher than that of **OUK-1** (448 mV). Thus, electrochemical impedance spectroscopy (EIS) analysis was performed to study the electron recombination process in DSSCs based on **OUK-1** and **OUK-3** in the dark under a forward bias of  $-0.60 \text{ V}$  with a frequency range of 10 mHz to 100 kHz. The large semicircle in the Nyquist plot (Fig. 6a, see Fig. S6a in ESI† for **OUK-1**), which corresponds to the midfrequency peaks in the Bode phase plots, represents the charge recombination between the injected electrons in  $\text{TiO}_2$  and  $\text{I}_3^-$  ions in the electrolyte, that is, the charge-transfer resistances at the  $\text{TiO}_2/\text{dye}/\text{electrolyte}$  interface. The Nyquist plots show that the resistance value for the large semicircle for **OUK-3** ( $50 \Omega$ ) is

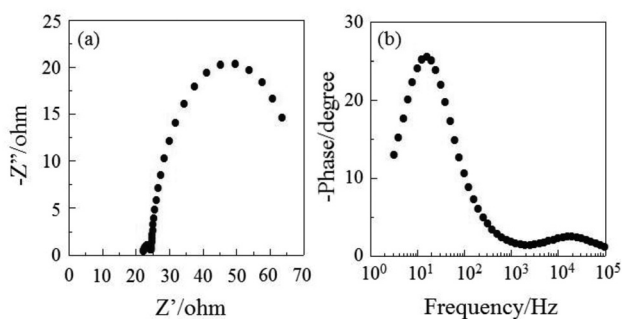


Fig. 6 (a) Nyquist plot and (b) Bode phase plot of DSSCs based on **OUK-3**.

larger than that of **OUK-1** ( $26 \Omega$ ), indicating that the electron recombination resistance of **OUK-3** is higher than that of **OUK-1**. The electron recombination lifetimes ( $\tau_e$ ) expressing the electron recombination between the injected electrons in  $\text{TiO}_2$  and  $\text{I}_3^-$  ions in the electrolyte, extracted from the angular frequency ( $\omega_{rec}$ ) at the midfrequency peak in the Bode phase plot (Fig. 6b, see Fig. S6b in ESI† for **OUK-1**) using  $\tau_e = 1/\omega_{rec}$ , are 11 ms for DSSC based on **OUK-3**, which is slightly larger than 10 ms for DSSC based on **OUK-1**. This result revealed that the charge recombination between the injected electrons in  $\text{TiO}_2$  and  $\text{I}_3^-$  ions in the electrolyte is not the major reason for the difference in the  $V_{oc}$  value between the two dyes, but the increase in the number of injected electrons in the CB of  $\text{TiO}_2$  and/or a negative shift of the  $E_{cb}$  of  $\text{TiO}_2$  by the formation of hydrogen bonding of the pyrazyl group at Brønsted acid sites on the  $\text{TiO}_2$  surface with the increasing dye loading of **OUK-3** on the  $\text{TiO}_2$  electrode may result in a higher  $V_{oc}$  value for **OUK-3**.<sup>4,6</sup> Consequently, the higher photovoltaic performance of **OUK-3** is attributed to the improvement of light-harvesting efficiency (LHE) and the increase in the number of injected electrons in the CB of  $\text{TiO}_2$  due to a high adsorption ability of the functionally separated D- $\pi$ -A dye onto the  $\text{TiO}_2$  electrode.

## Conclusions

We have designed and synthesized a functionally separated D- $\pi$ -A dye sensitizer **OUK-3** with a pyrazyl group as an electron-withdrawing anchoring group and a carboxyl group as an additional anchoring group for dye-sensitized solar cells. The dye **OUK-3** is adsorbed on the  $\text{TiO}_2$  surface through not only the bidentate bridging linkage of the carboxyl group but also the formation of hydrogen bonding of the pyrazyl group and the pyrazinium ion at Brønsted acid sites on the  $\text{TiO}_2$  surface. This work revealed that functionally separated D- $\pi$ -A dye sensitizers can achieve effective surface coverage of the  $\text{TiO}_2$  electrode due to its high adsorption ability onto the  $\text{TiO}_2$  electrode, leading to not only the improvement of light-harvesting efficiency, but also the increase in the number of injected electrons in the CB of  $\text{TiO}_2$ , which is responsible for the higher  $V_{oc}$  value of functionally separated D- $\pi$ -A dye sensitizers. This work also shows that the functionally separated D- $\pi$ -A dyes with two functional groups (cyclic azine and carboxyl group) possessing bonding ability to two points on the acid sites (Brønsted acid site and Lewis acid site) on the  $\text{TiO}_2$  surface would be expected to be one of the most promising classes of organic dye sensitizers for dye-sensitized solar cells. However, to further improve the photovoltaic performances of dye-sensitized solar cells based on the functionally separated D- $\pi$ -A dye sensitizers, the extension of the absorption spectrum of functionally separated D- $\pi$ -A dye sensitizers is necessary; for example, the destabilization of the HOMO level by the introduction of a stronger electron-donating group such as starburst triarylamine and the expansion of the  $\pi$ -conjugated system by the introduction of a long  $\pi$ -bridge such as terthio-



phene is one of the most effective strategies to improve the light-harvesting efficiency.

## Experimental

### General

Melting points were measured with a Yanaco micro melting point apparatus MP model. TG-DTA was carried out on a SII EXTRA6000 thermal analyser. IR spectra were recorded on a Perkin Elmer Spectrum One FT-IR spectrometer by the ATR method. High-resolution mass spectral data were acquired on a Thermo Fisher Scientific LTQ Orbitrap XL.  $^1\text{H}$  NMR spectra were recorded on a Varian-400 (400 MHz) FT NMR spectrometer with tetramethylsilane as an internal standard. Absorption spectra were observed with a Shimadzu UV-3150 spectrophotometer and fluorescence spectra were measured with a HORIBA FluoroMax-4 spectrofluorometer. The fluorescence quantum yields in solution were determined by a HORIBA FluoroMax-4 spectrofluorometer by using a calibrated integrating sphere system ( $\lambda_{\text{ex}} = 402$  nm for **OUK-1** and **OUK-3**). Cyclic voltammetry (CV) curves were recorded in the DMF/ $\text{Bu}_4\text{NClO}_4$  (0.1 M) solution with a three-electrode system consisting of  $\text{Ag}/\text{Ag}^+$  as the reference electrode, Pt plate as a working electrode, and Pt wire as a counter electrode by using a AMETEK Versa STAT 4 potentiostat. The highest occupied molecular orbital (HOMO) and the lowest unoccupied molecular orbital (LUMO) energy levels of **OUK-1** and **OUK-3** were evaluated from the spectral analyses and the CV data. The HOMO energy level was evaluated from the  $E_{1/2}^{\text{ox}}$ . The LUMO energy level was estimated from the  $E_{1/2}^{\text{ox}}$  and an intersection of absorption and fluorescence spectra (445 nm; 2.79 eV for **OUK-1** and 443 nm; 2.80 eV for **OUK-3**), which correspond to the energy gap between the HOMO and the LUMO. Electrochemical impedance spectroscopy (EIS) for DSSCs in the dark under a forward bias of  $-0.60$  V with a frequency range of 10 mHz to 100 kHz was measured with a AMETEK Versa STAT 3.

### Synthesis

***N,N*-Diphenyl-7-(5-(trimethylstannyl)thiophen-2-yl)-9*H*-carbazol-2-amine (2)**. To a THF solution (1.2 mL) of **1**<sup>19</sup> (0.1 g, 0.24 mmol) under an argon atmosphere was added 1.6 M hexane solution of *n*BuLi (0.3 ml) at  $-65$  °C. After stirring for 1 h,  $\text{Me}_3\text{SnCl}$  (0.075 g, 0.38 mmol) was added and the solution was stirred at  $-65$  °C for 2 h. After that, the solution was stirred for 8 h at room temperature. The reaction was quenched with water, and then the solution was extracted with dichloromethane. The dichloromethane extract was evaporated under reduced pressure. The resulting residue was subjected to reprecipitation from dichloromethane–hexane to give **2** (0.11 g, yield 78%) as a green powder; IR (ATR):  $\tilde{\nu} = 3417, 1583, 1486$   $\text{cm}^{-1}$ ;  $^1\text{H}$  NMR (400 MHz, acetone- $d_6$ , TMS)  $\delta = 0.40$  (s, 9H), 6.93 (dd,  $J = 2.0$  and 8.4 Hz, 1H), 7.02–7.05 (m, 2H), 7.09–7.11 (m, 4H), 7.15 (d,  $J = 2.0$  Hz, 1H), 7.24 (d,  $J = 3.4$  Hz, 1H), 7.28–7.32 (m, 4H), 7.64 (dd,  $J = 1.6$  and 8.2 Hz, 1H), 7.57

(d,  $J = 3.4$  Hz, 1H), 7.75 (d,  $J = 1.0$  Hz, 1H), 8.01 (d,  $J = 8.4$  Hz, 1H), 8.05 (d,  $J = 8.2$  Hz, 1H), 10.21 (s, –NH); HRMS (ESI):  $m/z$  (%): $[\text{M}^+]$  calcd for  $\text{C}_{31}\text{H}_{28}\text{N}_2\text{SSn}$ , 580.09897; found 580.09906.

***N,N*-Diphenyl-7-(5-(pyrazin-2-yl)thiophen-2-yl)-9*H*-carbazol-2-amine (3)**. A solution of **2** (0.1 g, 0.18 mmol), 2-iodopyrazine 0.028 ml, 0.28 mmol), and  $\text{Pd}(\text{PPh}_3)_4$  (0.014 g, 0.001 mmol) in toluene (2 mL) was stirred for 10 h at 110 °C under an argon atmosphere. The reaction mixture was diluted with water, and then the solution was extracted with dichloromethane. The dichloromethane extract was evaporated under reduced pressure. The residue was chromatographed on silica gel (dichloromethane as the eluent) to give **3** (0.07 g, yield 79%) as a yellow solid; m.p. 298–299 °C; IR (ATR):  $\tilde{\nu} = 3168, 1584, 1489$   $\text{cm}^{-1}$ ;  $^1\text{H}$  NMR (400 MHz, acetone- $d_6$ , TMS)  $\delta = 6.96$  (dd,  $J = 2.0$  and 8.4 Hz, 1H), 7.03–7.08 (m, 2H), 7.11–7.13 (m, 4H), 7.17 (d,  $J = 2.0$  Hz, 1H), 7.29–7.34 (m, 4H), 7.60–7.63 (m, 2H), 7.86 (d,  $J = 1.2$  Hz, 1H), 7.95 (d,  $J = 4.0$  Hz, 1H), 8.04 (d,  $J = 8.8$  Hz, 1H), 8.12 (d,  $J = 8.0$  Hz, 1H), 8.47 (d,  $J = 2.4$  Hz, 1H), 8.56 (dd,  $J = 1.6$  and 2.4 Hz, 1H), 9.18 (d,  $J = 1.6$  Hz, 1H), 10.31 (s, –NH); HRMS (APCI):  $m/z$  (%): $[\text{M} + \text{H}^+]$  calcd for  $\text{C}_{32}\text{H}_{23}\text{N}_4\text{S}$ , 495.16379; found 495.16351.

**Ethyl 7-(2-(diphenylamino)-7-(5-(pyrazin-2-yl)thiophen-2-yl)-9*H*-carbazol-9-yl)heptanoate (4)**. A solution of **3** (0.057 g, 0.12 mmol) in DMF (1.5 ml) was treated with sodium hydride (60%, 0.014 g, 0.58 mmol) and stirred for 1 h at room temperature. Ethyl 7-bromoheptanoate (0.0067 ml, 0.35 mmol) was added dropwise and the solution was stirred at room temperature for 3 h. The reaction was quenched with water, and then the solution was extracted with dichloromethane. The dichloromethane extract was evaporated under reduced pressure. The residue was chromatographed on silica gel (dichloromethane as the eluent) to give **4** (0.06 g, yield 79%) as a yellow viscous solid; IR (ATR):  $\tilde{\nu} = 1726, 1593, 1459$   $\text{cm}^{-1}$ ;  $^1\text{H}$  NMR (400 MHz, acetone- $d_6$ , TMS)  $\delta = 1.17$  (t,  $J = 7.2$  Hz, 3H), 1.31–1.34 (m, 4H), 1.49–1.54 (m, 2H), 1.82–1.86 (m, 2H), 2.21 (t,  $J = 7.2$  Hz, 2H), 4.01–4.06 (q, 2H), 4.39 (t,  $J = 6.8$  Hz, 2H), 6.94 (dd,  $J = 1.8$  and 8.5 Hz, 1H), 7.03–7.07 (m, 2H), 7.11–7.14 (m, 4H), 7.25 (d,  $J = 1.8$  Hz, 1H), 7.29–7.34 (m, 4H), 7.62 (dd,  $J = 1.6$  and 8.1 Hz, 1H), 7.67 (d,  $J = 3.9$  Hz, 1H), 7.92 (d,  $J = 1.1$  Hz, 1H), 7.95 (d,  $J = 3.9$  Hz, 1H), 8.05 (d,  $J = 8.5$  Hz, 1H), 8.13 (d,  $J = 8.1$  Hz, 1H), 8.47 (d,  $J = 2.5$  Hz, 1H), 8.56 (dd,  $J = 1.6$  and 2.5 Hz, 1H), 9.18 (d,  $J = 1.6$  Hz, 1H); HRMS (APCI):  $m/z$  (%): $[\text{M} + \text{H}^+]$  calcd for  $\text{C}_{41}\text{H}_{39}\text{O}_2\text{N}_4\text{S}$ , 651.27882; found 651.27863.

**7-(2-(Diphenylamino)-7-(5-(pyrazin-2-yl)thiophen-2-yl)-9*H*-carbazol-9-yl)heptanoic acid (OUK-3)**. To a solution of **4** (0.068 g, 0.1 mmol) in the mixed solvent of ethanol (27 ml) and dichloromethane (10 ml) was added dropwise aqueous NaOH (10 ml, 0.05 M) with stirring at 75 °C. After further stirring for 17 h, the solution was concentrated under reduced pressure. The residue was dissolved in dichloromethane, and washed with water, in which the aqueous layer was acidified to pH 3 with 2 M HCl. The dichloromethane extract was evaporated under reduced pressure. The resulting residue was subjected to reprecipitation from dichloromethane–hexane to give **OUK-3** (0.05 g, yield 77%) as a yellow powder; m.p. 129–131 °C, d.p. 376 °C; IR (ATR):  $\tilde{\nu} = 1724, 1594, 1459$   $\text{cm}^{-1}$ ;  $^1\text{H}$  NMR



(400 MHz, acetone- $d_6$ , TMS)  $\delta$  = 1.32–1.35 (m, 4H), 1.50–1.54 (m, 2H), 1.82–1.86 (m, 2H), 2.22 (t,  $J$  = 7.4 Hz, 2H), 4.40 (t,  $J$  = 6.8 Hz, 2H), 6.94 (dd,  $J$  = 1.7 and 8.4 Hz, 1H), 7.03–7.07 (m, 2H), 7.11–7.14 (m, 4H), 7.25 (s, 1H), 7.29–7.34 (m, 4H), 7.61 (d,  $J$  = 6.8 Hz, 1H), 7.67 (d,  $J$  = 3.9 Hz, 1H), 7.93–7.95 (m, 2H), 8.05 (d,  $J$  = 8.4 Hz, 1H), 8.12 (d,  $J$  = 8.2 Hz, 1H), 8.46 (s, 1H), 8.56 (s, 1H), 9.18 (s, 1H); HRMS (APCI):  $m/z$  (%): $[M + H^+]$  calcd for  $C_{39}H_{35}O_2N_4S$ , 623.24752; found 623.24744.

### Preparation of DSSCs based on dyes OUK-1 and OUK-3

The  $TiO_2$  paste (JGC Catalysts and Chemicals Ltd, PST-18NR) was deposited on a fluorine-doped-tin-oxide (FTO) substrate by doctor-blading, and sintered for 50 min at 450 °C. The 9  $\mu$ m thick  $TiO_2$  electrode was immersed into a 1 mM dye solution in THF for 15 hours sufficient to adsorb the photosensitizer. The DSSCs were fabricated by using the  $TiO_2$  electrode (0.5  $\times$  0.5  $cm^2$  in the photoactive area) thus prepared, with Pt-coated glass as a counter electrode, and a solution of 0.05 M iodine, 0.1 M lithium iodide, and 0.6 M 1,2-dimethyl-3-propylimidazolium iodide in acetonitrile as the electrolyte. The photocurrent–voltage characteristics were measured using a potentiostat under a simulated solar light (AM 1.5, 100  $mW\ cm^{-2}$ ). IPCE spectra were measured under monochromatic irradiation with a tungsten-halogen lamp and a monochromator. The dye-coated film was immersed in a mixed solvent of THF–DMSO–NaOH aq. 1 M (5 : 4 : 1), which was used to determine the amount of dye molecules adsorbed onto the film by measuring the absorbance. The quantification of the dye was made based on the  $\lambda_{max}$  (405 nm for OUK-1 and 402 nm for OUK-3) and the molar extinction coefficient of the dye in the above solution. Absorption spectra of the dyes adsorbed on  $TiO_2$  nanoparticles were recorded on the dye-adsorbed  $TiO_2$  film in the transmission mode with a calibrated integrating sphere system.

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### Notes and references

- B. O'Regan and M. Grätzel, *Nature*, 1991, **353**, 737.
- A. Hagfeldt, G. Boschloo, L. Sun, L. Kloo and H. Pettersson, *Chem. Rev.*, 2010, **110**, 6595.
- (a) T. Bessho, S. M. Zakeeruddin, C.-Y. Yeh, E. W.-G. Diau and M. Grätzel, *Angew. Chem., Int. Ed.*, 2010, **49**, 6646; (b) A. Yella, H.-W. Lee, H. N. Tsao, C. Yi, A. K. Chandiran, M. K. Nazeeruddin, E. W.-G. Diau, C.-Y. Yeh, S. M. Zakeeruddin and M. Grätzel, *Science*, 2011, **334**, 629; (c) L.-L. Li and E. W.-G. Diau, *Chem. Soc. Rev.*, 2013, **42**, 291; (d) S. Mathew, A. Yella, P. Gao, R. Humphry-Baker, B. F. E. Curchod, N. Ashari-Astani, I. Tavernelli, U. Rothlisberger, M. K. Nazeeruddin and M. Grätzel, *Nat. Chem.*, 2014, **6**, 242; (e) A. Yella, C.-L. Mai, S. M. Zakeeruddin, S.-N. Chang, C.-H. Hsieh, C.-Y. Yeh and M. Grätzel, *Angew. Chem., Int. Ed.*, 2014, **53**, 2973.
- (a) Z. Ning and H. Tian, *Chem. Commun.*, 2009, 5483; (b) Z. Ning, Y. Fu and H. Tian, *Energy Environ. Sci.*, 2010, **3**, 1170.
- A. Mishra, M. K. R. Fischer and P. Bäuerle, *Angew. Chem., Int. Ed.*, 2009, **48**, 2474.x
- (a) Y. Ooyama and Y. Harima, *Eur. J. Org. Chem.*, 2009, **18**, 2903; (b) Y. Ooyama and Y. Harima, *ChemPhysChem*, 2012, **13**, 4032.
- K. Ladomenou, T. N. Kitsopoulos, G. D. Sharma and A. G. Coutsolelos, *RSC Adv.*, 2014, **4**, 21379.
- N. Manfredi, B. Cecconi and A. Abbotto, *Eur. J. Org. Chem.*, 2014, 7069.
- (a) X. Wang, J. Yang, H. Yu, F. Li, L. Fan, W. Sun, Y. Liu, Z. Y. Koh, J. Pan, W.-L. Yim, L. yan and Q. Wang, *Chem. Commun.*, 2014, **50**, 3965; (b) S.-G. Li, K.-J. Jiang, J.-H. Huang, L.-M. Yang and Y.-L. Song, *Chem. Commun.*, 2014, **50**, 4309; (c) D. K. Panda, F. S. Goodson, S. Ray and S. Saha, *Chem. Commun.*, 2014, **50**, 5358; (d) K. Kakiage, Y. Aoyama, T. Yano, T. Otsuka, T. Kyomen, M. Unno and M. Hanaya, *Chem. Commun.*, 2014, **50**, 6379; (e) A. Amacher, C. Yi, J. yang, M. P. Bircher, Y. Fu, M. Cascella, M. Grätzel, S. Decurtins and S.-X. Liu, *Chem. Commun.*, 2014, **50**, 6540; (f) J. Yang, P. Ganesan, J. Teuscher, T. Moehl, Y. J. Kim, C. Yi, P. Comte, K. Pei, T. W. Holcombe, M. K. Nazeeruddin, J. Hua, S. K. Zakeeruddin, H. Tian and M. Grätzel, *J. Am. Chem. Soc.*, 2014, **136**, 5772.
- (a) R. B. Ambre, G.-F. Chang and C.-H. Hung, *Chem. Commun.*, 2014, **50**, 725; (b) T. Ikeuchi, H. Nomoto, N. Masaki, M. J. Griffith, S. Mori and M. Kimura, *Chem. Commun.*, 2014, **50**, 1941; (c) R. Agosta, R. Grisorio, L. De Marco, G. Romanazzi, G. P. Suranna, G. Gigli and M. Manca, *Chem. Commun.*, 2014, **50**, 9451; (d) A. Dessi, M. Calamante, A. Mordini, M. Peruzzin, A. Sinicropi, R. Basosi, F. F. de Biani, M. Taddei, D. Colonna, A. D. Carlo, G. Reginato and L. Zani, *Chem. Commun.*, 2014, **50**, 13952; (e) X. Sun, Y. Wang, X. Li, H. Ågren, W. Zhu, H. Tian and Y. Xie, *Chem. Commun.*, 2014, **50**, 15609.
- J. Mao, N. He, Z. Ning, Q. Zhang, F. Guo, L. Chen, W. Wu, J. Hua and H. Tian, *Angew. Chem., Int. Ed.*, 2012, **51**, 9873.
- (a) Y. Ooyama, S. Inoue, R. Asada, G. Ito, K. Kushimoto, K. Komaguchi, I. Imae and Y. Harima, *Eur. J. Org. Chem.*, 2010, 92; (b) Y. Ooyama, S. Inoue, T. Nagano, K. Kushimoto, J. Ohshita, I. Imae, K. Komaguchi and Y. Harima, *Angew. Chem., Int. Ed.*, 2011, **50**, 7429; (c) Y. Ooyama, T. Nagano, S. Inoue, I. Imae, K. Komaguchi, J. Ohshita and Y. Harima, *Chem. – Eur. J.*, 2011, **17**, 14837; (d) Y. Ooyama, N. Yamaguchi, I. Imae, K. Komaguchi, J. Ohshita and Y. Harima, *Chem. Commun.*, 2013, **49**, 2548; (e) Y. Ooyama, Y. Hagiwara, T. Mizumo, Y. Harima and



- J. Ohshita, *New J. Chem.*, 2013, **37**, 2479; (f) Y. Ooyama, T. Sato, Y. Harima and J. Ohshita, *J. Mater. Chem. A*, 2014, **2**, 3293.
- 13 (a) D. Daphnomili, G. D. Sharma, S. Biswas, T. K. R. Justin and A. G. Goutsolelos, *J. Photochem. Photobiol., A*, 2013, **253**, 88; (b) J. Lu, X. Xu, Z. Li, k. Cao, J. Cui, Y. Zhang, Y. Shen, Y. Li, J. Zhu, S. Dai, W. Chjen, Y. Cheng and M. Wang, *Chem. – Asian J.*, 2013, **8**, 956; (c) M.-D. Zhang, H.-X. Xie, X.-H. Ju, L. Qin, Q.-X. Yang, H.-G. Zheng and X.-F. Zhou, *Phys. Chem. Chem. Phys.*, 2013, **15**, 634; (d) D. Daphnomili, G. Landrou, P. Singh, A. Thomas, K. Yesudas, B. K. G. D. Sharma and A. G. Goutsolelos, *RSC Adv.*, 2012, **2**, 12899; (e) L. Wang, X. yang, S. Li, M. Cheng and L. Sun, *RSC Adv.*, 2014, **4**, 13677; (f) T. Sakurada, Y. Arai and H. Segawa, *RSC Adv.*, 2014, **4**, 13201; (g) J. Mao, D. Wang, S.-H. Liu, Y. Hang, Y. Xu, Q. Zhang, W. Wu, P.-T. Chou and J. Hua, *Asian J. Org. Chem.*, 2014, **3**, 153.
- 14 H. He, A. Gurung and L. Si, *Chem. Commun.*, 2012, **48**, 5910.
- 15 (a) Y. Ooyama, Y. Shimada, Y. Kagawa, I. Imae and Y. Harima, *Org. Biomol. Chem.*, 2007, **5**, 2046; (b) Y. Ooyama, Y. Shimada, Y. Kagawa, Y. Yamada, I. Imae, K. Komaguchi and Y. Harima, *Tetrahedron Lett.*, 2007, **48**, 9167; (c) Y. Ooyama, Y. Shimada, S. Inoue, T. Nagano, Y. Fujikawa, K. Komaguchi, I. Imae and Y. Harima, *New J. Chem.*, 2011, **35**, 111.
- 16 Y. Hao, X. Yang, J. Cong, H. Tian, A. Hagfeldt and L. Sun, *Chem. Commun.*, 2009, 4031.
- 17 J. Cong, X. Yang, J. Liu, J. Zhao, Y. Hao, Y. Wang and L. Sun, *Chem. Commun.*, 2012, **48**, 6663.
- 18 Y. Ooyama, Y. Hagiwara, Y. Oda, T. Mizumo, Y. Harima and J. Ohshita, *New J. Chem.*, 2013, **37**, 2336.
- 19 Y. Ooyama, K. Uenaka, Y. Harima and J. Ohshita, *RSC Adv.*, 2014, **4**, 30225.
- 20 (a) T. J. Dines, L. D. MacGregor and C. H. Rochester, *Phys. Chem. Chem. Phys.*, 2001, **3**, 2676; (b) M. I. Zaki, M. A. Hasan, F. A. Al-Sagheer and L. Pasupulety, *Colloids Surf., A*, 2001, **190**, 261; (c) O. Kasende and T. Zeegers-Huyskens, *J. Phys. Chem.*, 1984, **88**, 2132; (d) H. Takahashi, K. Mamola and E. K. Plyler, *J. Mol. Spectrosc.*, 1966, **21**, 217; (e) M. A. Montañez, I. L. Tocón, J. C. Otero and J. I. Marcos, *J. Mol. Struct.*, 1999, **482–483**, 201.

