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Iron(II)-catalyzed asymmetric intramolecular olefin aminochlorination using chloride ion[†]

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An iron-catalyzed enantioselective and diastereoselective intramolecular olefin aminochlorination reaction is reported (ee up to 92%, dr up to 15 : 1). In this reaction, a functionalized hydroxylamine and chloride ion are utilized as nitrogen and chlorine sources, respectively. This new method tolerates a range of synthetically valuable internal olefins that are all incompatible with existing asymmetric olefin aminochlorination methods.

Introduction

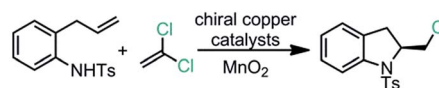
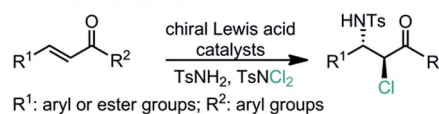
Enantioselective olefin halo-functionalization reactions constitute a range of synthetically valuable yet challenging transformations.¹ Although a variety of excellent asymmetric olefin halo-oxygenation reactions have been discovered,² there are much fewer asymmetric olefin aminohalogenation methods available.³ In particular, there have been just a few reported catalytic asymmetric olefin aminochlorination reactions.⁴ In one instance, Feng discovered the chiral Lewis acid-catalyzed aminochlorination of chalconic and other α,β -unsaturated olefins.^{4a,c} Also, Chemler reported copper-catalyzed aminochlorination of terminal olefins with chlorine radical donors in the presence of MnO_2 (Scheme 1A).^{4b} Despite these and other important discoveries, catalytic asymmetric aminochlorination methods for internal, non-chalconic olefins have yet to be developed. These methods would be synthetically valuable because they would readily provide vicinal amino chlorides, a class of important chiral building blocks. Moreover, asymmetric olefin aminochlorination that proceeds through an iron-nitrenoid intermediate has not yet been reported.⁵

We previously discovered $\text{Fe}(\text{BF}_4)_2$ -based catalysts for both diastereoselective and enantioselective intramolecular olefin aminofluorination reactions.⁶ Our initial attempts to apply these catalysts to olefin aminochlorination reactions led to either low diastereoselectivity or low yield, presumably due to the reason that chlorine and fluorine atom-transfer may proceed through

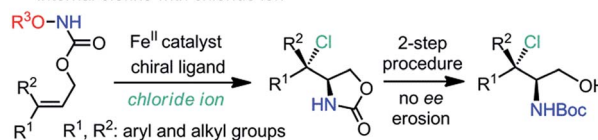
distinct mechanisms. Therefore, we explored a range of activating group–ligand combinations and discovered entirely new catalytic conditions for asymmetric olefin aminochlorination. Herein, we describe iron-catalyzed enantioselective and diastereoselective intramolecular aminochlorination for a range of internal, non-chalconic olefins (ee up to 92%, dr up to 15 : 1). In these reactions, a functionalized hydroxylamine and chloride ion were utilized as nitrogen and chlorine sources, respectively. This method tolerates a range of synthetically valuable internal olefins that are all incompatible with existing asymmetric olefin aminochlorination approaches; it also provides a new approach that is complementary to known methods for the asymmetric synthesis of amino chlorides with contiguous stereogenic centers.

Prior to this research, Bach reported an FeCl_2 -catalyzed racemic intramolecular olefin aminochlorination method using acyl azides, TMSCl , and EtOH under ligand-free conditions.⁷

A) previous work: asymmetric aminochlorination of chalconic and terminal olefins



B) current work: iron-catalyzed asymmetric aminochlorination of internal olefins with chloride ion



Scheme 1 Catalytic asymmetric olefin aminochlorination: summary of this work and other existing asymmetric methods.

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Table 1 Catalyst discovery for the iron-catalyzed diastereoselective olefin aminochlorination reaction

Entry ^a	Fe(X) ₂	Ligand (mol%)	Conversion ^b	Yield ^c	dr ^b (anti : syn)
1	FeCl ₂	None	62%	45%	2 : 1
2	FeCl ₂	L1 (20)	>95%	80%	>20 : 1
3	Fe(NTf ₂) ₂	L1 (20)	>95%	86%	>20 : 1
4	Fe(NTf ₂) ₂	L2 (10)	>95%	82%	0.83 : 1
5	Fe(NTf ₂) ₂	L3 (10)	61%	34%	0.25 : 1
6	Fe(NTf ₂) ₂	L4 (20)	>95%	75%	1.8 : 1

^a Unless stated otherwise, the reactions were carried out under a nitrogen atmosphere. TBAC: tetra-*n*-butylammonium chloride. ^b Conversion and dr were determined by ¹H NMR. ^c Isolated yield.

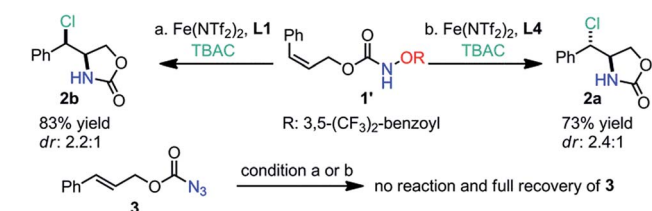
Excellent syn-selectivity was observed with styrenyl olefins (dr up to > 20 : 1). However, poor diastereoselectivity was recorded with non-styrenyl acyclic olefins (dr: 1 : 1). The new method presented here has a few unique features which complement the existing iron-catalyzed olefin aminochlorination method. First, excellent anti-selectivity has been observed across a wide range of styrenyl and non-styrenyl olefins. Second, good to excellent enantioselectivity has been achieved with a variety of internal, non-chalconic olefins (ee up to 92%). Finally, acyl azides are non-reactive under the described reaction conditions (*vide infra*), which suggests that iron-nitrenoid generation may proceed *via* different pathways compared with the known azide activation pathway.

Results and discussion

A cinnamyl alcohol-derived acyloxy carbamate **1** was selected as the model substrate for catalyst discovery (Table 1).⁸ In the presence of tetra-*n*-butylammonium chloride (TBAC), we observed that FeCl₂ alone catalyzed a sluggish reaction under ligand-free conditions (entry 1, 45% yield, dr: 2 : 1).⁹ However, the FeCl₂-phenanthroline (L1) complex catalyzed the anti-

Table 2 Substrate scope of the iron-catalyzed diastereoselective olefin aminochlorination reaction

1	2	3
86% yield, dr >20:1 from <i>E</i> olefin 83% yield, dr: 0.46:1 from <i>Z</i> olefin	86% yield, dr >20:1 ^a	70% yield, dr: 7:1
4	5	6
67% yield, dr: 10:1	76% yield, dr: 10:1	76% yield, dr: 12:1
7	8	9
61% yield, dr >20:1 ^a	59%, dr >20:1	93% yield, dr: 4.7:1 from <i>E</i> olefin ^a 84% yield, dr: 7:1 from <i>Z</i> olefin ^a
10	11	12
50% yield, dr >20:1	76% yield	69% yield, dr >20:1 ^b
13	14	15
77% yield ^c	88% yield, dr: 1.7:1	64% yield, dr >20:1



Scheme 2 Iron-catalyzed aminochlorination with a *cis* olefin and an acyl azide. ^aReaction conditions: Fe(NTf₂)₂ (10 mol%), L1 (20 mol%), TBAC (2.5 equiv.), CH₂Cl₂, 0 °C, 2 h. ^bReaction conditions: Fe(NTf₂)₂ (10 mol%), L4 (20 mol%), TBAC (2.5 equiv.), CH₂Cl₂, 0 °C, 2 h.

^a Reaction conditions: -15 °C, 2 h. ^b Reaction conditions: 0 °C, 5 h. ^c Reaction conditions: 0 °C, 12 h.



aminochlorination with significantly improved yield and dr (entry 2, 80% yield, dr > 20 : 1). We also noted that the Fe(NTf₂)₂-L1 complex provided essentially the same reactivity and diastereoselectivity (entry 3, 86% yield, dr > 20 : 1). Interestingly, the Fe(NTf₂)₂-bisoxazoline (L2) complex resulted in a loss of diastereoselectivity (entry 4, 82% yield, dr: 0.83 : 1). Furthermore, the Fe(NTf₂)₂-L3 complex promoted the syn-aminochlorination with moderate yield and dr (entry 5, 34% yield, dr: 0.25 : 1). We also observed that the Fe(NTf₂)₂-L4 complex catalyzed the anti-aminochlorination with a modest dr (entry 6, 75% yield, dr: 1.8 : 1). Notably, an iron-L4 complex resulted in high dr and reaction rate in the previously reported olefin aminofluorination reaction.⁶ These observations suggest that ligands are involved in the diastereoselectivity-determining step and provide excellent opportunities for diastereo-control.

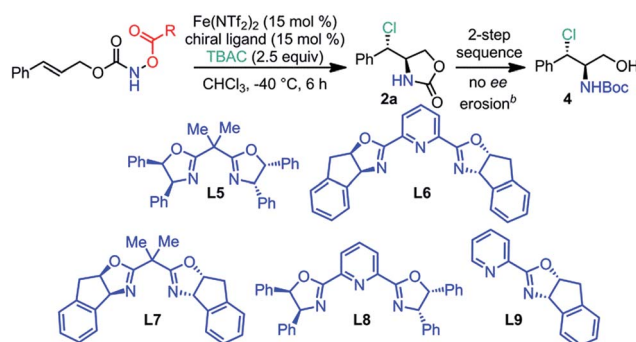
The observed ligand-enabled diastereo-control with *trans*-olefin **1** prompted us to evaluate *cis*-olefin **1'** (Scheme 2). To our surprise, the Fe(NTf₂)₂-L1 complex catalyzed syn-aminochlorination, while the Fe(NTf₂)₂-L4 complex promoted anti-aminochlorination with essentially the same dr (Scheme 2). The different reaction profiles for isomeric olefins **1** and **1'** suggest that the aminochlorination reaction is neither stereospecific nor fully stereo-convergent, which is significantly different from the iron-catalyzed olefin aminofluorination reaction.⁶

Furthermore, an acyl azide **3** was evaluated under the reaction conditions as a control experiment. Interestingly, the acyl azide **3** was fully recovered and no aminochlorination product

was detected. These results suggest that the activation of acyloxy carbamates (**1** and **1'**) may proceed *via* different pathways compared with the known azide activation pathway.⁷

We subsequently explored a range of olefins under the optimized conditions to evaluate the scope and limitations of this anti-aminochlorination method (Table 2). We discovered that di-substituted styrenyl olefins are generally good substrates; both electron-donating and electron-withdrawing substituents are compatible with this method (entries 1–4). Importantly, *ortho*-substituents and pyridyl groups are both tolerated (entries 5–6). Furthermore, extended aromatics, including naphthyl olefins, are reasonable substrates (entries 7–8). Moreover, isomeric ene-yne are both excellent substrates for the stereo-convergent and anti-selective method (entry 9). Additionally, we observed that both styrenyl and non-styrenyl tri-substituted olefins undergo aminochlorination smoothly with excellent dr (entries 10–11).¹⁰ We also discovered that a cyclohexyl-substituted olefin was an excellent substrate (entry 12, dr > 20 : 1). Further exploration revealed that both 1,1-disubstituted olefins and dienes are viable substrates with excellent regioselectivity (entries 13–14). Most notably, a cyclic olefin could also undergo highly diastereoselective anti-aminochlorination (entry 15, dr > 20 : 1), yielding a product which is difficult to obtain with known methods.¹¹ Since the FeCl₂-L1 complex provides essentially the same dr and yield in these diastereoselective reactions, FeCl₂ can be a convenient substitute for Fe(NTf₂)₂ in racemic reactions.

Table 3 Catalyst discovery for the iron-catalyzed asymmetric olefin aminochlorination reaction



Entry ^a	R	Ligand	Conversion ^c	Yield ^d	dr ^c (anti : syn)	ee ^e (anti)	ee ^e (syn)
1	3,5-(CF ₃) ₂ -Ph	L5	>95%	53%	9.9 : 1	84%	<5%
2	3,5-(CF ₃) ₂ -Ph	L6	>95%	68%	0.5 : 1	24%	79%
3	3,5-(CF ₃) ₂ -Ph	L7	88%	61%	1.7 : 1	<5%	<5%
4	3,5-(CF ₃) ₂ -Ph	L8	>95%	32%	2.5 : 1	47%	30%
5	3,5-(CF ₃) ₂ -Ph	L9	>95%	82%	0.5 : 1	8%	24%
6 ^f	3,5-(CF ₃) ₂ -Ph	L5	>95%	51%	11.0 : 1	90%	<5%
7 ^f	CH ₃	L5	>95%	42%	1.1 : 1	97%	<5%
8 ^f	CH ₂ Cl	L5	>95%	67%	9.6 : 1	89%	<5%
9 ^{f,g}	CH ₂ Cl	L5	>95%	58%	9.0 : 1	83%	<5%

^a Unless stated otherwise, the reactions were carried out under a nitrogen atmosphere with 4 Å molecular sieves. ^b Reaction conditions: Boc₂O, Et₃N, DMAP; then Cs₂CO₃, MeOH, 85% over two steps; see ESI for details. ^c Conversion and dr were determined by ¹H NMR. ^d Isolated yield. ^e Enantiomeric excess (ee) was measured by HPLC with chiral columns; the absolute stereochemistry was determined by X-ray crystallographic analysis of an analog of **2a**. ^f The reaction was carried out at -60 °C for 12 h. ^g The FeCl₂-L5 complex was used.



In order to fulfil the need for catalytic asymmetric olefin aminochlorination, we further explored asymmetric induction for internal, non-chalconcic olefins with a variety of iron–chiral ligand complexes (Table 3).¹² First, we discovered that the iron–L5 complex induced diastereoselective and enantioselective anti-aminochlorination, albeit with a low yield, mostly due to the competing aminohydroxylation reaction (entry 1, 53% yield, dr: 9.9 : 1). Interestingly, the anti-addition product **2a** was obtained with excellent ee (84% ee), while the syn-addition product **2b** was obtained essentially as a racemate (<5% ee).¹³ Additionally, a two-step procedure can convert **2a** to a chlorinated amino alcohol triad **4** without ee erosion.¹⁴ Next, we observed that the iron–L6 complex induced moderately diastereoselective syn-aminochlorination (entry 2, 68% yield, dr: 0.48 : 1). To our surprise, the anti-addition product **2a** was obtained with moderate ee (24% ee), while the syn-addition product **2b** was isolated with significant ee (79% ee). Furthermore, we evaluated chiral ligands L7 and L8 and determined that they are less effective for asymmetric induction (entries 3–4). Additionally, chiral ligand L9 induced fast yet non-selective aminochlorination with a high overall yield (entry 5).¹⁵ With the iron–L5 complex in hand, we subsequently explored other reaction parameters. First, a decreased reaction temperature was found to benefit both dr and ee (entry 6, dr: 11 : 1 and 90% ee for **2a** at –60 °C). Next, replacing the 3,5-bis(trifluoromethyl) benzoyl activating group with a smaller acetyl group further enhanced the ee (entry 7, 97% ee for **2a**); however, much lower dr and yield were obtained (entry 7, dr: 1.1 : 1, 42% yield). Finally, a chloroacetyl activating group induced an effective balance between overall yield and stereoselectivity (entry 8, 67% yield, dr: 9.6 : 1 and 89% ee for **2a**). We also observed that the FeCl₂–L5 complex induced a slightly less selective reaction with a lower yield (entry 9, 58% yield, dr: 9.0 : 1 and 83% ee for **2a**).

In order to evaluate the scope of this asymmetric method, we explored the asymmetric induction with a range of internal olefins (Table 4). The chiral catalyst provides excellent asymmetric induction with styrenyl olefins. A range of *para*-substituted styrenyl olefins with different electronic properties were converted to the corresponding aminochlorination products with high dr and ee (entries 1–6, dr: 9.6–15 : 1, ee: 86–91%). Additionally, *meta*-substituted styrenyl olefins are also good substrates but with slightly decreased ee (entries 7–9, dr: 10–15 : 1, ee: 80–87%). However, we discovered that *ortho*-substitution of styrenes has a deleterious effect on ee (entries 10–11, dr: 4.5–12 : 1, ee: 77–79%). Interestingly, both α - and β -naphthyl olefins are excellent substrates (entries 12–13, dr: 4.5–10 : 1, ee: 89–92%). To our delight, a 3-pyridyl olefin with a basic nitrogen atom is a reasonable substrate for the asymmetric aminochlorination (entry 14, dr: 1.8 : 1, ee: 70% for the anti-diastereomer). Moreover, we observed that the iron–L5 complex can induce significant ee in the aminochlorination with non-styrenyl olefins (entry 15, dr: 2 : 1, ee: 54% for the anti-diastereomer). To our surprise, the iron–L6 complex proved to be uniquely effective for the asymmetric induction with tri-substituted olefins, while the iron–L5 complex was less effective (entry 16, dr: 2.3 : 1, ee: 86% for the anti-diastereomer).¹⁶

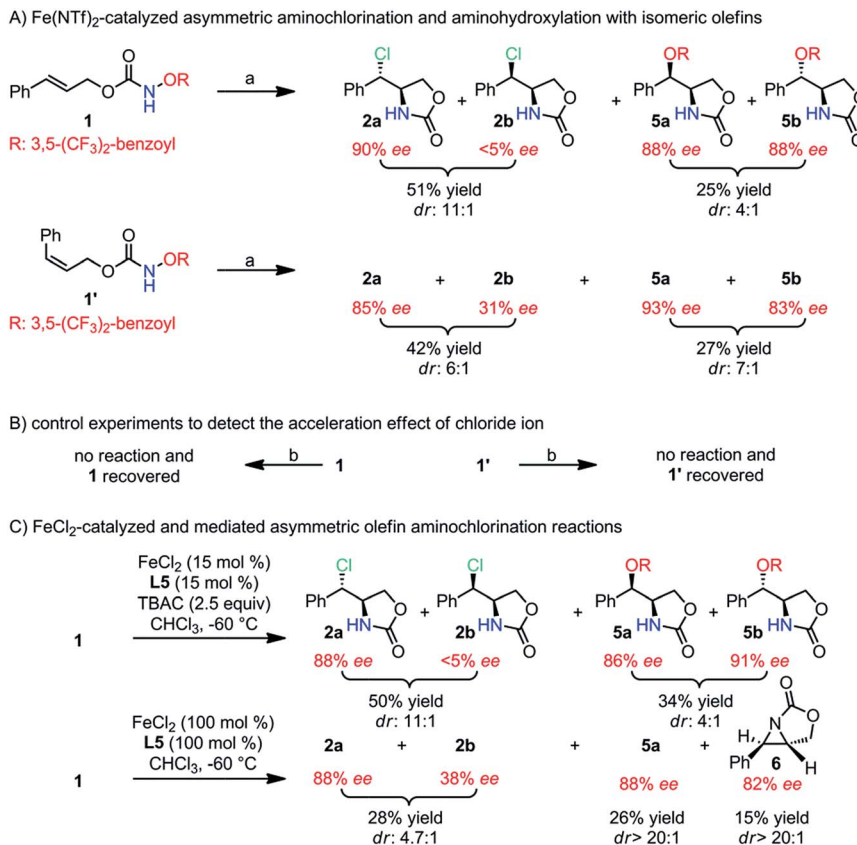
Table 4 Substrate scope for the iron-catalyzed asymmetric olefin aminochlorination reaction

olefinic substrates ^a		$\text{Fe}(\text{NTf}_2)_2$ (15 mol %) L5 (15 mol %) TBAC (2.5 equiv)	olefin aminochlorination products ^a	
		$\xrightarrow{\text{CHCl}_3, -60^\circ\text{C}, 12\text{ h}}$		
1			67% yield, dr: 9.6:1 89% ee	
2			65% yield, dr: 15:1 91% ee	
3			69% yield, dr: 5.2:1 87% ee	
4			84% yield, dr: 12:1 90% ee	
5			62% yield, dr: 11:1 88% ee	
6			71% yield, dr: 11:1 86% ee	
7			75% yield, dr: 12:1 87% ee	
8			63% yield, dr: 10:1 80% ee	
9			71% yield, dr: 15:1 80% ee	
10			78% yield, dr: 4.5:1 77% ee	
11			55% yield, dr: 12:1 79% ee	
12			63% yield, dr: 10:1 92% ee ^b	
13			53% yield, dr: 4.5:1 89% ee ^b	
14			51% yield, dr: 1.8:1 70% ee	
15			66% yield, dr: 2:1 54% ee ^{b,c}	
16			45% yield, dr: 2.3:1 86% ee ^{b,d}	

^a Unless stated otherwise, mono-chloroacetyl was selected as the activating group for asymmetric catalysis; the ee for all syn-aminochlorination products was less than 5%. ^b Bis(trifluoromethyl)-benzoyl was selected as the activating group. ^c The ee for the syn-addition product was 12%. ^d L6 was used as the ligand for asymmetric induction; the ee for the syn-addition product was 50%.

During the exploration of substrate scope, it was surprising to observe completely different ee values for anti- and syn-diastereomers (e.g. **2a** and **2b**). In contrast, exactly the same ee for both diastereomeric products was observed in the iron-catalyzed aminofluorination of **1**.⁶ In order to obtain greater mechanistic insights, we carried out ee analysis for all isolable products using several control experiments (Scheme 3). First, in an Fe(NTf₂)₂-catalyzed reaction with *trans*-olefin **1**, two aminochlorination products were obtained (Scheme 3A, 90% ee for **2a**, <5% ee for **2b**, dr: 11 : 1).¹⁷ Simultaneously, diastereomers **5a** and **5b** were also isolated with the same ee as two competing olefin aminohydroxylation products (Scheme 3A, 88% ee for **5a** and **5b**, dr: 4 : 1). However, completely different selectivity (both dr and ee) was observed in an Fe(NTf₂)₂-catalyzed reaction with *cis*-olefin **1'** (Scheme 3A, 85% ee for **2a** and 31% ee for **2b**, dr:





Scheme 3 Control experiments to probe the mechanism. ^aReaction conditions: $\text{Fe}(\text{NTf}_2)_2$ (15 mol%), L1 (15 mol%), TBAC (2.5 equiv.), CHCl_3 , -60 °C, 12 h. ^bReaction conditions: $\text{Fe}(\text{NTf}_2)_2$ (15 mol%), L1 (15 mol%), CHCl_3 , -60 °C, 12 h.

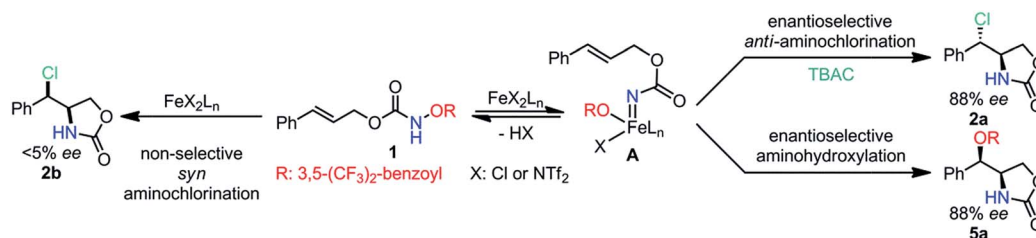
6 : 1; 93% ee for **5a** and 83% ee for **5b**, dr: 7 : 1). In both cases, **5a** and **5b** cannot be converted to **2a** under the reaction conditions.

These observations provide several important mechanistic insights. First, the non-stereospecificity observed in the iron-catalyzed olefin aminochlorination suggests that the formation of C–N and C–Cl bonds occurs in a stepwise fashion.¹⁸ Second, the lack of complete stereo-convergence between the reaction profiles of isomeric olefins (**1** and **1'**) suggests that C–N bond formation may be the rate- and ee-determining step.¹⁸ Furthermore, since essentially the same ee was observed for **2a**, **5a**, and **5b** from the reaction with *trans*-olefin **1**, it is likely that these products are derived from the same intermediate after the ee-determining step. Additionally, the fact that the syn-aminochlorination product **2b** was isolated as a racemate suggests that **2b** may be derived from non-stereoselective pathways

which are distinct from the one leading to the formation of **2a**, **5a**, and **5b**.

The product divergence (**2a** vs. **5a/b**) after the ee-determining step is mechanistically interesting. Therefore, we studied the effect of external chloride ion. To our surprise, in the absence of TBAC, the $\text{Fe}(\text{NTf}_2)_2$ -L5 complex alone was ineffective for the nitrogen atom-transfer at -60 °C; **1** and **1'** were both fully recovered (Scheme 3B). However, aminochlorination occurred as soon as a stoichiometric amount of TBAC was introduced. This observation suggests that the $\text{Fe}(\text{NTf}_2)_2$ -L5 complex may serve as a pre-catalyst and it may be activated by chloride ion *in situ*.

In order to test this hypothesis, we further carried out the FeCl_2 -catalyzed reaction in the presence of TBAC (Scheme 3C). Notably, **2a** was isolated with essentially the same ee as that obtained under the standard conditions (88% ee for **2a** and <5%



Scheme 4 Proposed mechanistic working hypothesis for the iron-catalyzed asymmetric aminochlorination of *trans*-olefin **1**.



ee for **2b**). This result suggests that the catalytically relevant species may also be generated from the FeCl₂-L5 complex.

To probe for more mechanistic details, we subsequently carried out the FeCl₂-promoted olefin aminochlorination in the absence of TBAC (100 mol% FeCl₂, 100 mol% L5, Scheme 3C). Under these conditions, FeCl₂ is the only available chlorine source. Surprisingly, we discovered that **2a** was obtained with essentially the same ee compared with the two previous control experiments (88% ee for **2a**). Furthermore, a syn-amino-hydroxylation product **5a** was isolated with excellent dr and ee (dr > 20 : 1, 88% ee). These observations suggest that Fe-Cl bond cleavage may be relevant for the chlorine atom-transfer step during the enantioselective anti-aminochlorination.¹⁹ In addition, we also identified a small amount of aziridine **6** (15% yield, 82% ee), and further discovered that it could not be converted to either **2a** or **5a** under the reaction conditions.

With the accumulated mechanistic evidence, we propose a plausible mechanistic working hypothesis for the iron-catalyzed asymmetric aminochlorination of *trans*-olefin **1** (Scheme 4). First, the iron catalyst reversibly cleaves the N-O bond in the acyloxy carbamate **1**, generating iron-nitrenoid **A** with chloride as a counter ion. From there, **A** may participate in enantioselective and diastereoselective aminochlorination and amino-hydroxylation to afford **2a** and **5a**, respectively. Since the aminochlorination-aminohydroxylation competition occurs after the ee-determining step, **2a** is obtained with essentially the same ee as **5a**. At the same time, **1** may be converted to **2b** via a non-stereoselective pathway which is distinct from the one leading to the formation of **2a** and **5a**. Further mechanistic studies are required to elucidate the details.

Conclusions

In conclusion, we have described an iron-catalyzed enantioselective and diastereoselective aminochlorination method for internal, non-chalconic olefins. This method tolerates a range of synthetically valuable olefins that are all incompatible with existing asymmetric olefin aminochlorination methods. It also provides a complementary approach for the asymmetric synthesis of amino chlorides with contiguous stereogenic centers. Our preliminary mechanistic studies revealed that an FeCl₂-derived nitrenoid may be a feasible reactive intermediate and that Fe-Cl bond cleavage may be relevant for stereoselective chlorine atom-transfer. Our current efforts are focused on the mechanistic investigation of this new reaction and method development for the enantioselective intermolecular olefin aminochlorination.

Acknowledgements

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- 9 The relative stereochemistry of **2a** was determined by comparison of the experimental NMR data with those reported in ref. 7. It was further corroborated by ¹H NMR and X-ray crystallographic analysis of a structural analog of **2a**. See ESI† for details.
- 10 The relative stereochemistry was assigned based on the ¹H NMR and X-ray crystallographic analysis of a structural analog described in ref. 6; see ESI† for details.
- 11 Complementary stereochemistry was achieved (in entry 15 of Table 2) compared with the known method reported in ref. 7, where the syn-aminochlorination product was isolated. This substrate did not undergo kinetic resolution with a chiral catalyst, the complex Fe(NTf₂)₂-L5. Both the starting material and product were isolated as racemates.
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- 15 For the synthesis of **L9**, see ref. 6.
- 16 The iron-L5 complex catalyzed the reaction favoring the syn-addition product (dr (anti/syn): 0.47 : 1); ee for the anti-addition product was 60% and ee for the syn-addition product was <5%. The relative stereochemistry was assigned based on the ¹H NMR and X-ray crystallographic analysis of a structural analog described in ref. 6; see ESI† for details.
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