

Cite this: *RSC Adv.*, 2016, 6, 19531

Synthesis, structure and coordination properties of novel bifunctional carboxylic derivatives of 1,3-alternate tetrathiacalix[4]arene†

Sergey N. Podyachev,^{*a} Gulnaz Sh. Gimazetdinova,^b Aidar T. Gubaidullin,^a Victor V. Syakaev,^a Svetlana N. Sudakova,^a Bulat M. Gabidullin,^a Vladimir T. Ivanov,^a Edward L. Gogolashvili^a and Alexander I. Konovalov^a

New bifunctional derivatives of 1,3-alternate tetrathiacalix[4]arene decorated with carboxylic, ester, hydrazide and/or hydrazone groups have been synthesized with good yields using the tetrathiacalix[4]arene derivatives with incorporated pairs of carboxylic and ester groups as versatile building blocks. The structural peculiarities of the obtained bifunctional compounds have been investigated by means of X-ray analysis, IR and NMR spectroscopy. The recognition ability of the synthesized macrocycles towards some alkali, alkali-earth and transition metal ions has been investigated applying a solvent extraction method. The results showed that the structure of a calix[4]arene platform as well as the nature of functional substitutes located on opposite sides of the macrocycle are critical for the coordination properties of the synthesized compounds.

Received 20th January 2016

Accepted 9th February 2016

DOI: 10.1039/c6ra01730d

www.rsc.org/advances

1. Introduction

The design and synthesis of artificial receptors for cation and anion guests is an exciting topic of current chemistry because of the extreme importance of these compounds for coordination and supramolecular chemistry as well as for biological processes involving molecular recognition of cationic, anionic and neutral species.^{1–7} The sensing properties of the receptors are usually based on their ability to form multiple-point binding sites during interaction with a variety of substrates. It is well-known that macrocyclic compounds with several donor centers can reveal a unique selectivity towards some substrates. However, there are some restrictive factors on the way of preparing such type of receptors and they are connected not only with an accurate simulation of the structure or prediction the recognition properties of these compounds but obviously deal with the problem of their further synthesis. In many cases, the solving of this problem opens new possibilities to the successful construction of artificial receptors. At the same time, current industrial needs require more and more advanced compounds possessing unique properties and enabling the

progress in the creation of novel functional materials. Therefore, a search for rational approaches to the synthesis of ligands with a desired set of properties is an actual task.

One of the strategies for the molecular design of ligands can be founded on the molecular LEGO approach. It assumes that even a trivial combination of the basic molecular platforms and functional building blocks makes possible to obtain a great number of various sophisticated structures with desired and predictable properties, applying, as a rule, relatively simple chemical protocols. The calix[*n*]arenes are known as a synthetically and commercially accessible class of compounds which can be successfully applied for these purposes due to their suitable molecular scaffolds.^{8,9} A distinctive feature of this class of macrocyclic compounds is their three-dimensional spatially preorganized structure that can be fixed in various isomeric forms. Moreover, such compounds give an opportunity for incorporating of a variety of substitutes both at the upper and lower rims. These peculiarities of calix[4]arenes provide their application as keystones for the creation of bifunctional receptors having unequal binding centers.^{10–12} A possibility of coordinating at least two types of substrates simultaneously extends an application area of these compounds. An interest to such type of calix[4]arene derivatives can be also explained by the fact that they can reveal the biomimetic properties.¹³ Additionally, a presence of spatially divided binding and reaction sites in the structure of the compounds gives an excellent opportunity for the preparation on their base the polymeric as well as covalently fixed on a hard substrate sensing materials.

Among calix[*n*]arene family compounds, the tetrathiacalix[4]arenes attract a special attention. In opposite to their classical

^aA. E. Arbuzov Institute of Organic and Physical Chemistry, Kazan Scientific Center of Russian Academy of Sciences, Arbuzov str., 8, 420088, Kazan, Russia. E-mail: spodyachev@iopc.ru; Fax: +7-843-273-1872

^bKazan National Research Technological University, K. Marks str., 68, 420015, Kazan, Russia

† Electronic supplementary information (ESI) available. CCDC 1437363–1437367. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c6ra01730d

analogues, they can be easily fixed in any of *cone*, *partial cone* or *1,3-alternate* isomer forms by using a template effect of metal cations just on the initial stage of their functionalizing. Although the *cone*-isomers usually reveal substantially higher binding efficiency towards different ions in comparison with *1,3-alternate* isomers, the latter are interesting due to their enhanced binding selectivity.^{14,15} Furthermore, the *1,3-alternate* tetrathiacalix[4]arenes can be considered as very suitable tectons for the construction of artificial heteroditopic receptors capable of exhibiting a mimic allosteric binding that plays a major role in biological systems.^{11,12,16,17}

In present work we demonstrate a synthetic strategy for the preparing of novel *1,3-alternate* tetrathiacalix[4]arene derivatives functionalized by pairs of carboxylic, ester, hydrazide, chloroanhydride as well as hydrazone groups and report the details of the synthesis and spectroscopic characterizations of these compounds. Their structural peculiarities have been investigated by means of X-ray analysis. The receptor properties of the synthesized bifunctional derivatives against alkali, alkali-earth and transition metals were studied by liquid-liquid extraction method and are discussed below.

2. Results and discussion

2.1. Synthesis

The compounds with ester and carboxylic groups often serve as starting compounds for a further synthesis of a wide range of organic compounds. Therefore, the calix[4]arene scaffolds

functionalized by both these groups can be potentially used as promising building blocks for the obtaining of bifunctional macrocyclic derivatives due to the different reactivity of ester and carboxylic groups.^{18,19} The synthetic route utilized by us for the preparation of bifunctional compounds is presented in Fig. 1. The synthesis starts from the parent tetrathiacalix[4]arenes **1a** and **1b** which are converted into the corresponding tetraethyl esters **2a** and **2b** after their alkylation with ethyl bromoacetate in the refluxing acetone and using Cs_2CO_3 as a base.^{20,21}

1,3-Alternate tetrathiacalix[4]arenes are known to bind Cs^+ ions only by “one side” of the macrocycle. This selectivity is obviously caused by the allosteric effect.^{22,23} Taking this fact into account, we have developed an effective protocol for the synthesis of dicarboxylic tetrathiacalix[4]arene derivatives **3a** and **3b**. The reflux of the compounds **2a** and **2b** in the THF solution containing 5 vol% of water and 12 equivalents of Cs_2CO_3 resulted in the formation of cesium salts of **3a** and **3b**, respectively. After treatment of the precipitates with HCl, the calix[4]arenes **3a** and **3b** were obtained with good yields. The details of the synthetic procedure and the supposed mechanism of the reaction as well as the structures of the compounds **3a**, **3b** and dicesium salt of **3a** determined by the X-ray analysis have been reported recently.²⁴

The obtained compounds **3a** and **3b** can undergo a further selective transformation on both sides of the tetrathiacalix[4]arene platform. We were especially interested in the introduction of hydrazide groups into the macrocyclic framework. According

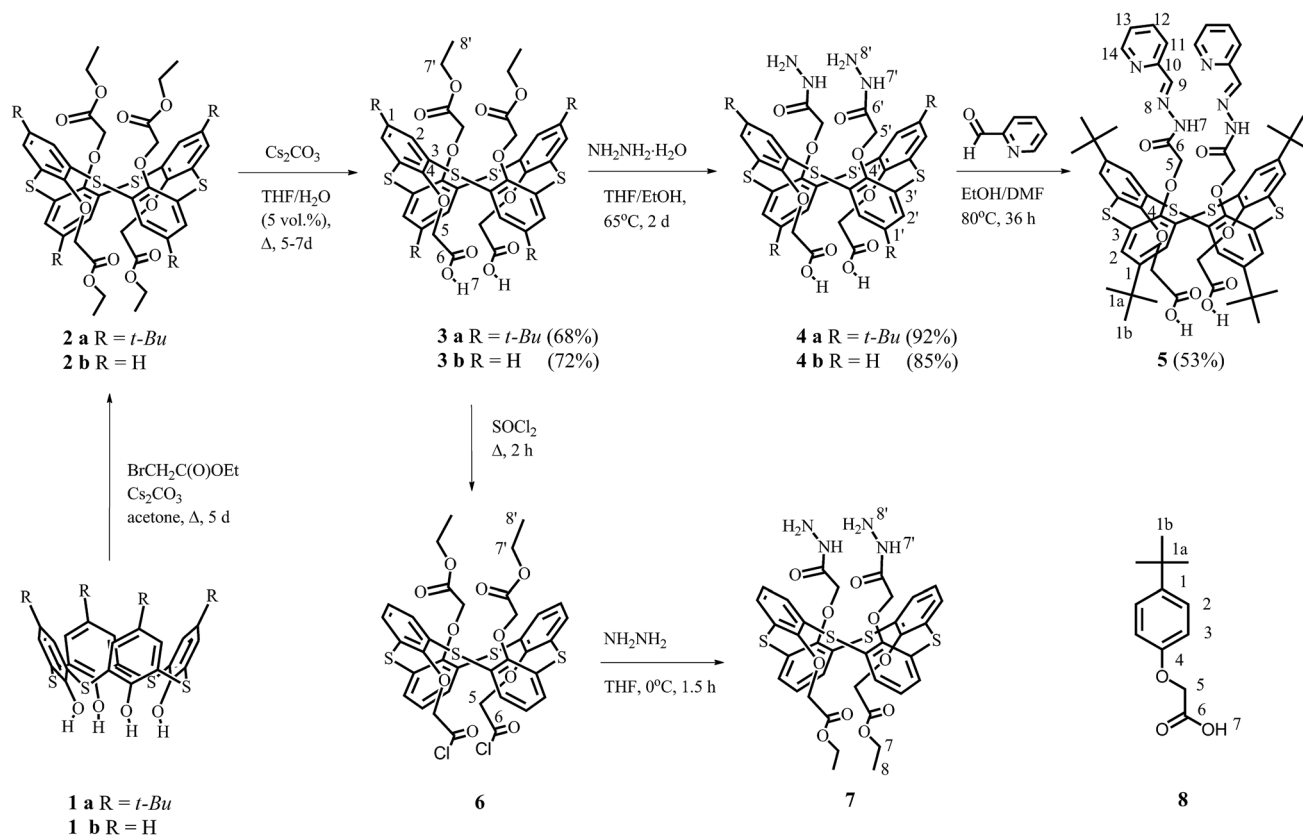


Fig. 1 Synthetic routes and structural formulae of the investigated compounds. Numbering system of atoms used in the experimental section.

Table 1 The frequencies of characteristic vibrations (ν , cm^{-1})^a in the IR spectra of compounds

		Vibration assignment							
Compound		$\nu_{\text{as}}\text{NH}_2\text{NH}$	$\nu_{\text{s}}\text{NH}_2\text{NH}$	δNH	δNH_2	$\nu\text{C}=\text{O}$	νPh	$\nu\text{CH}_{2(3)}$	νCSC
3a	KBr					1767	1450	2959	1050 br
						1738		2907	1090
						1692		2872	
	Nujol					1766	1460 br	— ^b	1053
						1738			1068
3b	KBr					1690 br			1089
						1762	1419 vbr	2981	1054
						1723 br		2903	1078
	Nujol					1761	1462 br	— ^b	1058
						1724 br	1437 sh		1078
4a	KBr	3423 br		1541 br	1626 sh	1663 br	1417 sh		
		3333				1758	1478 sh	2963	1041
						1682 br	1450	2909	
	Nujol	3422 br		— ^c	1627	1758	1430	2872	
		3330 br				1681 vbr	1457 br	— ^b	1041 br
4b	KBr	3416 br	3210 sh	1522 br	1620 sh	1718 sh	1418	2912	1039 br
		3319 br				1683 vbr			
		3415 br	3210 br	1522	1620 sh	1720 sh	1461	— ^b	1041 br
	Nujol	3318				1685 vbr	1436		
							1418		
5	KBr	3445 vbr ^d		1534		1757 sh	1467	2962	1044 br
		3329 ^d				1707 br	1448	2907	
		3234 ^d					1429	2872	
	Nujol	3459 br ^d		1537		1758	1463 br	— ^b	1046
		3293 ^d				1717 vbr			
6	KBr					1815	1474	2989	1060
						1802	1422	2903	
						1766	1408		
	Nujol					1815	1462	— ^b	1060
						1802	1422		
7	KBr	3429	3210 sh	1505	1632	1766	1408		
		3328 vbr				1765	1473	2981	1038 br
						1729	1420 br	2906	
	Nujol	3428	3210	1502	1632 sh	1677 br			
		3336 br				1728	1462 br	— ^b	1039
						1675 br	1422 sh		

^a Additional characteristics of absorption bands: sh – shoulder; br – broad; vbr – very broad. The maxima of rather intensive bands of signals are solid underlined. The maxima of weak intensive bands of signals are dashed underlined. ^b Characteristic vibrations are not due to overlap with Nujol bands. ^c Absorption band no detected. ^d Characteristic vibrations of νNH .

to the investigations accomplished in the last decade, the hydrazide derivatives of calix[4]arene can act as selective receptors for anions and cations.^{14,15,25–27} Furthermore, these compounds can be successfully applied as key reagents for the synthesis of various nitrogen containing derivatives, including acylhydrazones, acylsemicarbazones and heterocycles.

The tetrathiacalix[4]arenes **4a** and **4b** functionalized by acetylhydrazide groups have been prepared in 92% and 85% yields by the reflux of the corresponding tetrathiacalixarenyloxyacetic acid diesters **3a** and **3b** with an excess of $\text{NH}_2\text{—NH}_2\cdot\text{H}_2\text{O}$ in the THF-EtOH solution for 2 days (Fig. 1). The crude products isolated from the reaction mixture can include the hydrazine molecules captured due to the interactions with carboxylic groups of the tetrathiacalix[4]arene molecules. Therefore, for obtaining the hydrazine free products **4a** and **4b**, the solid

remainders were suspended in water and thereafter treated carefully with HCl to reach pH ~5.

The condensation of dihydrazide **4a** with picolinaldehyde has been accomplished in DMF-EtOH solution by heating the mixture at 80 °C over 36 h. After an appropriate treatment, the target dihydrazone **5** was obtained with the yield of 53%.

The synthesis of diester-dihydrazide derivative **7** started from the diester-dicarboxylic acid **3b**. At the first stage, **3b** was converted into dichloroanhydride derivative **6** under treatment with thionyl chloride and reflux the mixture for 2 h. A subsequent condensation of the crude product **6** with an anhydrous hydrazine dispersed in the absolute THF at 0 °C led to the formation of bifunctional derivative **7** with the yield 72%. The synthesis should be carried out with a special accuracy, since a variety of by-products can be formed during the reaction. It



should be also noted that we have attempted to obtain the compound **7** by the etherification of **4b** in the presence of EtOH and *p*-toluenesulfonic acid, as a catalyst. However, the yield of a target product did not exceed 30%.

In order to explore a complexation behavior of novel synthesized compounds, the monomeric carboxylic counterpart **8**, as a structural block of tetrathiacalix[4]arenes **3**–**5**, has been obtained from *p*-*tert*-butyl-phenoxyacetic acid ester.²⁸

2.2. IR characterization in solid

IR spectra of *p*-*tert*-butyltetrathiacalix[4]arenes **3a**, **4a**, **5** differ from the spectra of tetrathiacalix[4]arenes **3b**, **4b**, **6** and **7** due to the presence in former of the intensive bands $\nu(\text{CH}_3)$ at $\sim 2960\text{ cm}^{-1}$ assigned to the *tert*-butyl groups enclosed in the structure of the compounds (see Table 1 and Fig. 1, 2, 5, 8, 11, 14 and 17 in ESI†). In IR spectra of the compounds **3a** and **3b**, the absorption bands $\nu\text{C}=\text{O}$ belonging to the ester groups have almost the same maxima (1767 cm^{-1} and 1762 cm^{-1}). At the same time, the absorption bands $\nu\text{C}=\text{O}$ of carboxylic groups for **3a** appear as two peaks at 1738 cm^{-1} and 1692 cm^{-1} . In the case of **3b**, we observe a single but rather intensive peak at 1723 cm^{-1} in the spectrum obtained in KBr pellets. The second peak having a weak intensity and low-frequency was detected at 1663 cm^{-1} only in Nujol. The presence of both peaks in the spectra of **3a** and **3b** indicates obviously an occurrence as free carboxylic groups in the structure of the compounds and the groups participating in the hydrogen bonding as well.

The replacement of ester groups in the compounds on the hydrazide (for **4a** and **4b**) or the hydrazone (for **5**) ones leads to the appearance ($>3000\text{ cm}^{-1}$) of complicated absorption bands in their spectra caused by νNH vibrations. The similar peaks are observed in the spectrum of dihydrazide derivative **7** as well. The presence of amide-II absorption band ($\delta\text{NH} \sim 1520\text{--}1540\text{ cm}^{-1}$) in the spectra of **4a**, **4b**, **5** and **7** indicates the *trans*-conformation of amide groups in the structure of these compounds.^{29–31} At the same time, an absorption band $\nu\text{NH} \sim 3234\text{ cm}^{-1}$ which is also observed in the spectrum of the dihydrazone **5** proves a contribution of *cis*-form to the conformational composition of the compound.

The presence of low-frequency absorption bands $\nu_{\text{as}}\text{NHNH}_2$ along side with high-frequency peaks (3423 cm^{-1} and 3333 cm^{-1} for **4a**, 3416 cm^{-1} and 3318 cm^{-1} for **4b** in KBr) in the spectra of the compounds points to the fact that at least one of the hydrazide groups in these molecules participates in the hydrogen bonding. A similar picture is observed in the case of dihydrazone derivative **5** (3445 cm^{-1} and 3329 cm^{-1}) as well.

An occurrence of hydrazide and carboxylic groups in the structure of the compounds **4a** and **4b** leads to the appearance in their spectra of a broad absorption peak (at 1680 cm^{-1}) arising due to the overlap of $\nu\text{C}=\text{O}$ bands. In the case of **4a**, a low-intensity peak at 1758 cm^{-1} caused probably by the presence of free carboxylic groups is also detected. When going to the dihydrazone **5**, a broad absorption band $\nu\text{C}=\text{O}$ with a maximum at 1707 cm^{-1} was observed in the spectrum of this compound as well. The underfrequency for this vibration is obviously connected with the participation of carboxylic and

hydrazone groups in the hydrogen bonding. However, the hydrogen bonds, in this case, are not so strong, as those formed in the compounds **4a** and **4b**. In the case of spectra **6**, the doublet $\nu(\text{C}=\text{O})$ at 1815 cm^{-1} and 1802 cm^{-1} for chloroanhydride groups and a sharp single peak for ester groups at 1766 cm^{-1} are observed in KBr. An absorbance $\nu(\text{C}=\text{O})$ of ester groups for the compound **7** is detected exactly at the same frequency (1765 cm^{-1}) but appears as a quiet low intensive peak. At the same time, an absorbance of $\nu\text{C}=\text{O}$ of hydrazide groups is revealed as two peaks having an equal intensity at 1729 cm^{-1} and 1677 cm^{-1} , indicating the involving only one of the carbonyl oxygen atoms of these groups in the hydrogen bonding.

The hydroxyl groups appear as a broad dome-shaped $\nu(\text{OH})$ band in IR spectra of the compounds **3a**, **3b**, **4a** and **4b**. This band is underneath the $\nu(\text{=CH})$, $\nu(\text{–CH})$ and $\nu(\text{NH})$ absorptions in the region $\sim 3500\text{--}3100\text{ cm}^{-1}$ with a center being around $\sim 3400\text{ cm}^{-1}$ for **3a**, **3b** and $\sim 3250\text{ cm}^{-1}$ for **4a** and **4b**. These values of frequency are lower than ones observed for the stretching vibrations of free hydroxyl groups ($\sim 3500\text{ cm}^{-1}$), that proves their participation in the formation of H-bonds. In the case of dihydrazone derivative **5**, the absorbance $\nu(\text{OH})$ has more broadened shape with an uncertain maximum.

2.3. X-ray analysis

Suitable for X-ray analysis the colorless crystals of dihydrazides **4a** and **4b** were obtained by the recrystallization of their crude products from MeOH–CH₂Cl₂. We have also succeeded in the preparing the crystals of **4a** in DMSO when treated a crude product with HCl (see Experimental part). In former case, **4a** and **4b** calix[4]arenes were crystallized with the capturing two

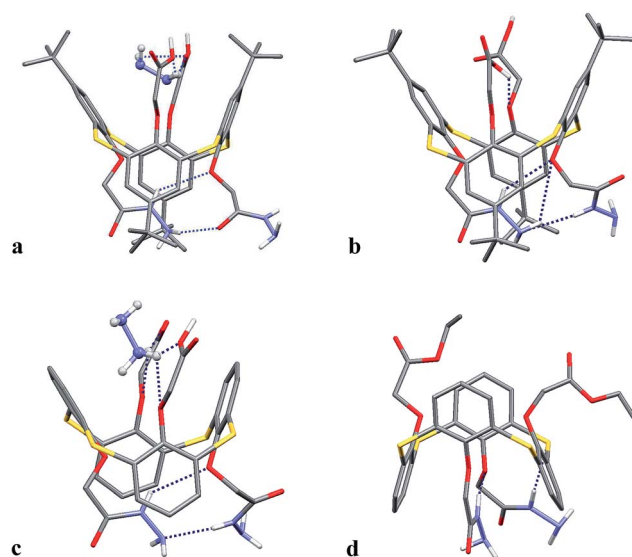


Fig. 2 The structures of compounds (a) **4a**· $2\text{NH}_2\text{NH}_2$; (b) **4a**; (c) **4b**· NH_2NH_2 and (d) **7** in the crystals. Only hydrogen atoms of OH and NH groups are shown. The disordered fragments are shown in positions with the highest occupancy. In the cases of (a) – the hydrazine molecule, (b) – the DMSO molecule and (c) – the water molecules outside of the thiacalix[4]arene cavity are not shown.



and one hydrazine molecules, affording consequently the structures **4a**·2NH₂NH₂ (Fig. 2a) and **4b**·NH₂NH₂ (Fig. 2c). In the second case, the crystals of **4a** were obtained without inclusion the hydrazine in their structure (Fig. 2b). The recrystallization of dihydrazide **7** from MeOH–CH₂Cl₂ has also brought the crystals do not containing hydrazine molecules (Fig. 2d). In the last case, the product was treated by water without adding HCl. The presence of carboxylic groups in **4a** and **4b** in contrast to **7** is obviously a main reason for the hydrazine binding by these compounds. The acidification of **4a** and **4b** solutions prevents the amine coordination. The crystals of dihydrazone **5** were obtained from MeOH–CH₂Cl₂ as well. All considered bifunctional derivatives **4a**, **4b**, **5** and **7** possess the 1,3-alternate isomer form established for their parent **3a** and **3b** tetrathiacalix[4]arenes previously.

The compound **4a**·2NH₂NH₂ crystallizes in *C2/c* space group with one hydrazine molecule being in thiacalix[4]arene's cavity and the another one being outside (Fig. 2a). The highly disordered solvent molecules can be localized between thiacalix[4]arene molecules in the crystal and occupy up to 1307 Å³ per unit cell. However, it should be noticed that these solvate molecules could not be detected properly using Fourier electron density synthesis under normal X-ray experiment conditions. A coordination of hydrazine molecules by the thiacalix[4]arene **4a** leads to an increase only one from four of dihedral angles formed by the aromatic moieties and a reference plane passing through sulfur atoms (for compare in **4a**: 119.1(2)°, 113.1(2)°, 116.0(2)° and 116.0(2)°; in **4a**·2NH₂NH₂: 127.4(2)°, 109.1(2)°, 105.7(2)° and 108.3(2)°).

The hydrazine free thiacalix[4]arene **4a** crystallizes in *P2₁/n* space group with one DMSO molecule per macrocycle directed outside of the cavity (Fig. 2b). All *tert*-butyl groups are disordered over two positions with relative occupancies: 0.77 : 0.23, 0.60 : 0.40, 0.65 : 0.35 and 0.65 : 0.35. The carboxylic group is also disordered over two positions with a relative occupancy being 0.57 : 0.43. Inequality of bond lengths in carboxylic groups indicate that both –COOH substituents in the molecule are not deprotonated.

The realization of classical hydrogen bonds in the molecules of thiacalix[4]arene **4a** results in the formation of H-dimers in the crystals due to the interaction between hydrazide groups of the centro-symmetrical pairs of molecules, whereas the carboxylic groups do not form intra- and intermolecular hydrogen bonds between each other. It is obviously caused by the fact that one of the carboxylic groups is involved in the hydrogen bonding with DMSO molecule located in the macrocycle cavity, and the another one participates in the intra-molecular hydrogen bonding with ester oxygen of an opposite phenoxylic fragment of the thiacalix[4]arene framework.

The compound **4b**·NH₂NH₂ crystallizes in *P2₁/c* space group. The asymmetric part of the unit cell contains one thiacalix[4]arene molecule together with one hydrazine molecule being inside (Fig. 2c), as in the case of **4a**·2NH₂NH₂, and four co-crystallized water molecules located outside the cavity. Dihedral angles between aromatic moieties and a reference plane passing through sulfur atoms (109.3(1)°, 104.1(1)°, 100.7(1)° and 123.6(1)°) are slightly less than the corresponding values for

4a. An absence of bulky *tert*-butyl groups at the upper rim of thiacalix[4]arene favours the tube-shaped structure of **4b** which is clear observed for this compound even under encapsulation.

Insufficiently high quality of **4a**·2NH₂NH₂ and **4b**·NH₂NH₂ crystals prevents a deep analysis of the bond distances and angles in these molecules as well as makes difficult a precise determination of the hydrogen atoms in OH and NH₂ groups. Nevertheless, it is clearly seen that the hydrazine molecule is coordinated due to the participation of both carboxylic groups. It should be also noticed that dicarboxylic tetrathiacalix[4]arene derivatives can generate a formation of salt structures with an involvement of deprotonated carboxylic group of macrocycle and the hydrazinium ion.

The dihydrazide **7** crystallizes in *P2₁/c* space group. The asymmetric part of the unit cell comprises just one thiacalix[4]arene molecule and does not enclose solvate molecules (Fig. 2d). The dihedral angles between the aromatic moieties and the plane defined by four sulfur atoms were determined as 111.5(1)°, 99.1(1)°, 105.2(1)° and 115.8(1)°. One of the hydrazide and one of the ester groups are disordered over two positions with relative occupancy ratios of 0.45 : 0.55 and 0.42 : 0.58, respectively. The hydrazide substitutes are involved in intra-molecular as well as the intermolecular hydrogen bonds simultaneously, which results in 3D supramolecular structure formed by **7** in the crystal.

The dihydrazide derivatives **4a**, **4b** and **7** have *trans*-amide conformation of the acetylhydrazide fragments and the non-hydrogen atoms of these groups are located practically at the same plain. One of the hydrazide substituents closes a pseudo-cavity of the thiacalix[4]arene and prevents a penetration of solvent molecules (Fig. 2a–d). Moreover, this spatial structure of the molecules is additionally stabilized by a system of intra-molecular and intermolecular hydrogen bonds.

Dihydrazone derivative **5** crystallizes in *P1* space group. The asymmetric part of the unit cell contains one thiacalix[4]arene molecule and one methanol molecule which acts as a bridge between two hydrazone fragments (Fig. 3). Dihedral angles between aromatic moieties and a reference plane passing through sulfur atoms are 116.9(5)°, 103.7(5)°, 105.3(5)° and 110.3(4)°. One of the *tert*-butyl groups is disordered over two positions with a relative occupancy ratio of 0.54 : 0.46.

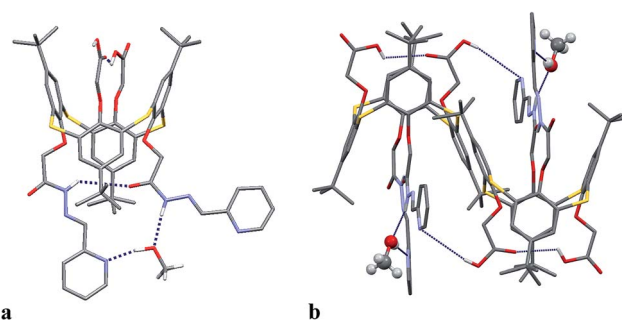


Fig. 3 The geometry of the molecule (a) and H-dimer (b) in the crystal of **5** in the crystal. Hydrogen atoms are not shown, except for the hydrazone and carboxyl groups of **5** and MeOH.





Table 2 ^1H chemical shifts^a (ppm) and observed spin–spin coupling (Hz) of 1,3-alternate tetrathiacalix[4]arenes **3–7** and compound **8**

Compound	4a	4b	5	6	7	8
Solvent	DMSO- <i>d</i> ₆	DMSO- <i>d</i> ₆	DMSO- <i>d</i> ₆	CDCl ₃	DMSO- <i>d</i> ₆	DMSO- <i>d</i> ₆
Fragment ^b	–COOH	–COOH	–COOH	–C(O)Cl	–COOEt	–COOH
H1b	1.212	1.200	1.122 1.152 1.216 1.229	–NHN=C–Py 1.244, (1.033; 1.104) ^d		1.249
H1						
H2	7.389	7.386	7.382 7.481 7.498	7.456 7.481 (7.616; 7.625) ^d	7.026, t, ³ <i>J</i> = 7.7 7.4 7.38, d, ³ <i>J</i> = 7.4 7.5 7.50, d, ³ <i>J</i> = 7.5	6.984, t, ³ <i>J</i> = 7.8 7.447 d, ³ <i>J</i> = 7.8 7.285, d, ³ <i>J</i> = 8.76
H3						6.816, d, ³ <i>J</i> = 8.76 4.616
H5	4.316	4.283	4.356 4.385 4.428 4.452 4.505	3.870; 4.318 (5.038; 5.203) ^d	4.487	4.437
H7	— ^e	7.59	12.37	9.865; 11.104 (11.62; 11.799) ^d	4.23, q, ³ <i>J</i> = 6.9 1.28, t, ³ <i>J</i> = 6.9	7.231 — ^e
H8						
H9				8.087; 8.459, (7.959; 7.978) ^d	4.088, q, ³ <i>J</i> = 7.1 1.164, t, ³ <i>J</i> = 7.1	12.917
H11				7.908 (7.716) ^d		
H12				7.786 (7.860) ^d		
H13				7.346 (7.432) ^d		
H14				8.625 (8.541) ^d		

^a Numbering according to Fig. 1. ^b Assignment for structural fragment of tetrathiacalix[4]arene containing the indicated functional group. ^c Registered only *trans*-form. ^d Assignment for *trans*–(*cis*)-form of hydrazine fragments. ^e No detected signal.

In spite of the fact that strong donor and acceptor groups are present in the molecule **5**, only pair interactions between the hydroxyl groups of carboxylic substitutes of one molecule and the pyridine nitrogen atoms of a neighboring molecule are realized in the crystal. These interactions result in the formation of centrosymmetric H-dimers (or pseudo-tetramers, if two solvate MeOH molecules are taking into account). Such type of the molecules' arrangement, however, does not lead to the closest packing. Indeed, the voids revealed in the crystals occupy up to 246 Å³ per unit cell and can be filled up by the highly disordered solvent molecules.

The ^1H and ^{13}C spectra of bifunctional derivatives **4a**, **4b**, **6** and **7** as well as the spectra of starting compounds **3a** and **3b** described by us earlier²⁴ show two sets of signals assigned to tetrathiacalix[4]arene scaffold (Table 2 and 3), although each of $-\text{OCH}_2\text{R}$ pairs of functional groups produces only one combination of signals. Such spectral picture for the synthesized compounds is in a full agreement with their expected structure and is typical for the distal location of the identical substitutes in molecule.

The determination of the isomer form for the “classical” calix[4]arenes containing methylene bridges in their structure can be easily performed by using a simple “de Mendoza rule”.^{8,33} In the case of tetrathiacalix[4]arenes, a more sophisticated analysis of the chemical shifts for other groups of atoms in comparison with the model compounds as well as 2D NMR experiments are usually required. However, in our case, the situation is simplified due to the fact that *n*-butyl groups are just bulky enough to prevent rotation of the alkylated phenolic rings through the main annulus of thiacalix[4]arene, at least at temperatures up to 413 K.³⁴ Since the synthetic pathway to the desired compounds started from the conformationally immobilized *1,3-alternate* precursors **3a** and **3b** with well documented structures,²⁴ we have assumed that their derivatives **4a**, **4b**, **6** and **7** would also adopt the *1,3-alternate* isomer form. The 2D NOESY experiments for the compounds **4–7** have been nevertheless accomplished to confirm this fact. Really, the cross-peaks between aromatic and *tert*-butyl protons and the protons

Table 3 ^{13}C chemical shifts^a (ppm) of 1,3-alternate tetrathiacalix[4]arenes **3–7** and compound **8**

Compound	4a	4b	5	6	7	8
Solvent	DMSO- <i>d</i> ₆	DMSO- <i>d</i> ₆	DMSO- <i>d</i> ₆	CDCl ₃	DMSO- <i>d</i> ₆	DMSO- <i>d</i> ₆
Fragment ^b	-COOH	-COOH	-COOH	-C(O)Cl	-COOEt	-COOH
C1a	31.0	30.7	33.88	-NHN=C-Py		33.7
C1b	34.1	33.9	30.70	33.7 (34.1) ^d		31.3
C1	147.0	147.2	146.2	146.6 (146.7) ^d	124.4	143.1
C2	128.8	129.8	130.5, 131.7	132.5 (134.0) ^d	137.2	125.9
C3	127.1	126.9	127.0	127.4 (127.3) ^d	128.5	113.8
C4	154.5	155.3	155.3, 155.8	157.4 (156.7) ^d	159.4	155.4
C5	65.6	68.3	67.3, 68.3, 68.6	69.6; 68.7, (67.3) ^d	68.8	64.4
C6	168.6	165.5	168.8, 169.0	163.1; 164.0, (168.6; 168.9) ^d	168.3	170.2
C7					60.9	
C8					14.3	
C9						
C10				148.4 (143.2) ^d		
C11				152.8 (153.1) ^d		
C12				120.2 (119.9) ^d		
C13				136.7 (136.8) ^d		
C14				124.4 (124.0) ^d		
				149.4 (149.3) ^d		

^a Numbering according to Fig. 1. ^b Assignment for structural fragment of tetrathiaalix[4]arene containing the indicated functional group. ^c Registered only *trans*-form. ^d Assignment for *trans*-(*cis*)-form of hydrazine fragment.

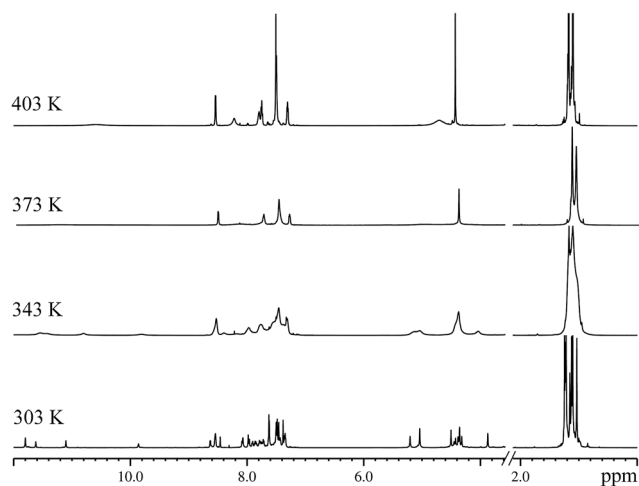


Fig. 4 The influence of the temperature change 303 K \rightarrow 403 K on ^1H NMR spectra of the dihydrazone *tert*-butyltetrahydrazone 5.

of OCH_2 groups belonging to the acid or hydrazide fragments of the adjacent structural blocks of tetrathiacalix[4]arene were detected in the spectrum of **4a**. It is quite clear that only in the case of the *1,3-alternate* isomer, these cross-peaks could be observed. The similar picture was characteristic for all synthesized compounds **4**–**7**.

It could be expected that an occurrence of amide fragments in the structures of **4a**, **4b**, **5** and **7** should result in the complication of their spectra due to the realization of *cis*- and *trans*-amide conformers for these compounds. In the case of compound **5**, a formation of additional spatial forms is also possible due to *E/Z* isomerization relative to the $\text{C}=\text{N}$ double bond. Indeed, a great number of peaks as well as the broadening and overlapping signals are observed in ^1H NMR spectrum of dihydrazone **5** at 303 K (Fig. 4). The various spatial forms for hydrazones have been analyzed carefully and discussed in detail in our previous publications.^{35,36} In accordance with the spectral criteria suggested in these articles, we have performed a conformational assignment for the investigated compounds and represented the results in the Tables 2 and 3.

It was established that only *trans*-amide conformer is realized for dihydrazides **4a**, **4b** and **7**. The hydrazone fragments in **5** were found to exist in $E_{\text{C}=\text{N}}$ isomer form. It is worth noting that the *trans*-form content for the compound **5** equals 42% and is practically similar to the conformational composition determined for the corresponding 4-*tert*-butylphenoxyacetylhydrazone (40%)³⁵ as well as for *1,3-alternate* tetrahydrazone *p*-*tert*-butyltetrahydrazone 5 (44%).³²

It was shown in our previous investigations that UV irradiation of 4-*tert*-butylphenol and resorcinol acetylhydrazone derivatives resulted in a partly conversion of the $E_{\text{C}=\text{N}}$ into the $Z_{\text{C}=\text{N}}$ isomer.^{31,35} According to NMR data, the stabilization of $Z_{\text{C}=\text{N}}$ isomer in these compounds was provided due to the occurrence of intramolecular hydrogen bonds between NH protons and the nitrogen atoms of 2-pyridinyl substitutes in the hydrazone fragments. However, the irradiation of CDCl_3 solution of the compound **5** by UV light for 2 h did not lead to the

$E_{\text{C}=\text{N}}/Z_{\text{C}=\text{N}}$ isomerization. Such situation was observed by us earlier for the calix[4]resorcinol and tetrathiacalix[4]arene (*cone* and *1,3-alternate*) acetylhydrazone derivatives also.^{31,36} The presence of neighboring acetylhydrazone groups capable of participating in intramolecular interactions as well as the sterical hindrances realized in tetrathiacalix[4]arene and calix[4]resorcinol macrocycles prevent evidently the UV-induced $E_{\text{C}=\text{N}}/Z_{\text{C}=\text{N}}$ isomerization.

An occurrence of a large number of polar groups possessing acceptor and donor properties makes possible for the investigated compounds a formation of dimer and even substantially greater molecular aggregates in low-polarity solvents. It could be expected that the compounds **4a** and **4b** bearing the carboxylic and hydrazide groups capable of formation of a various type of intermolecular hydrogen bonds are greatly predisposed to the generation of such aggregates. To check this assumption, we have applied a diffusion NMR method which is commonly used for studying an aggregation in host–guest systems³⁷ and investigating of the hydrogen-bonded assemblies in the solution-phase.^{38,39}

NMR DOESY experiments performed for the compounds **4a** and **4b** in the mixture of CDCl_3 – CD_3OD (1 : 1) have demonstrated that the investigated bifunctional derivatives does not form even dimers at the concentrations up to 30 mM. The analysis of a change of self-diffusion coefficients under increase of the ligand's concentration showed that the amount of the dimeric molecules for these compounds does not exceed 5%. It should be noticed that we have previously investigated the aggregation properties of dicarbonic acid **3b** possessing a dimeric structure in the crystal.²⁴ It appeared to our surprise that in CDCl_3 solution the percentage of **3b** molecules adopting the dimer form was not above 25%.

To summarize NMR experiments, we can conclude that the complication of NMR spectra for the compounds **4**–**7** may be caused, first of all, by the intramolecular interactions of functional groups and, in the case of the compound **5**, by *cis/trans* amide conformational isomerism. The magnetic anisotropy of aromatic rings in *1,3-alternate* isomer may additionally complicate the spectra of these compounds. This influence is especially noticeable for the compound **5**. Thus, the signals of methylene groups (H5) of acetylhydrazone fragments having *trans*-conformation differ on 0.45 ppm (Table 2 and Fig. 4). Under heating the samples up to 373 K, the proton peaks of carboxylic groups as well as the peaks of hydrazone group protons for different conformers coalesce and converge into sharp lines at 403 K. Under cyclic temperature changes 303 K \rightarrow 403 K \rightarrow 303 K, the spectral picture does not change, which testifies to the maintenance of the structural content for the compound **5**.

2.5. Extraction studies

The liquid–liquid extraction experiments were performed to examine the efficiency and selectivity of tetrathiacalix[4]arene derivatives **3a**, **3b**, **4a**, **4b**, **5**, **7** and their monomeric counterpart **8** in transferring of s- (Na^+ , Ca^{2+}), p- (Pb^{2+}) and d- (Co^{2+} , Ni^{2+} , Cu^{2+} , Zn^{2+} , Ag^+ and Cd^{2+}) metal ions from aqueous phase into



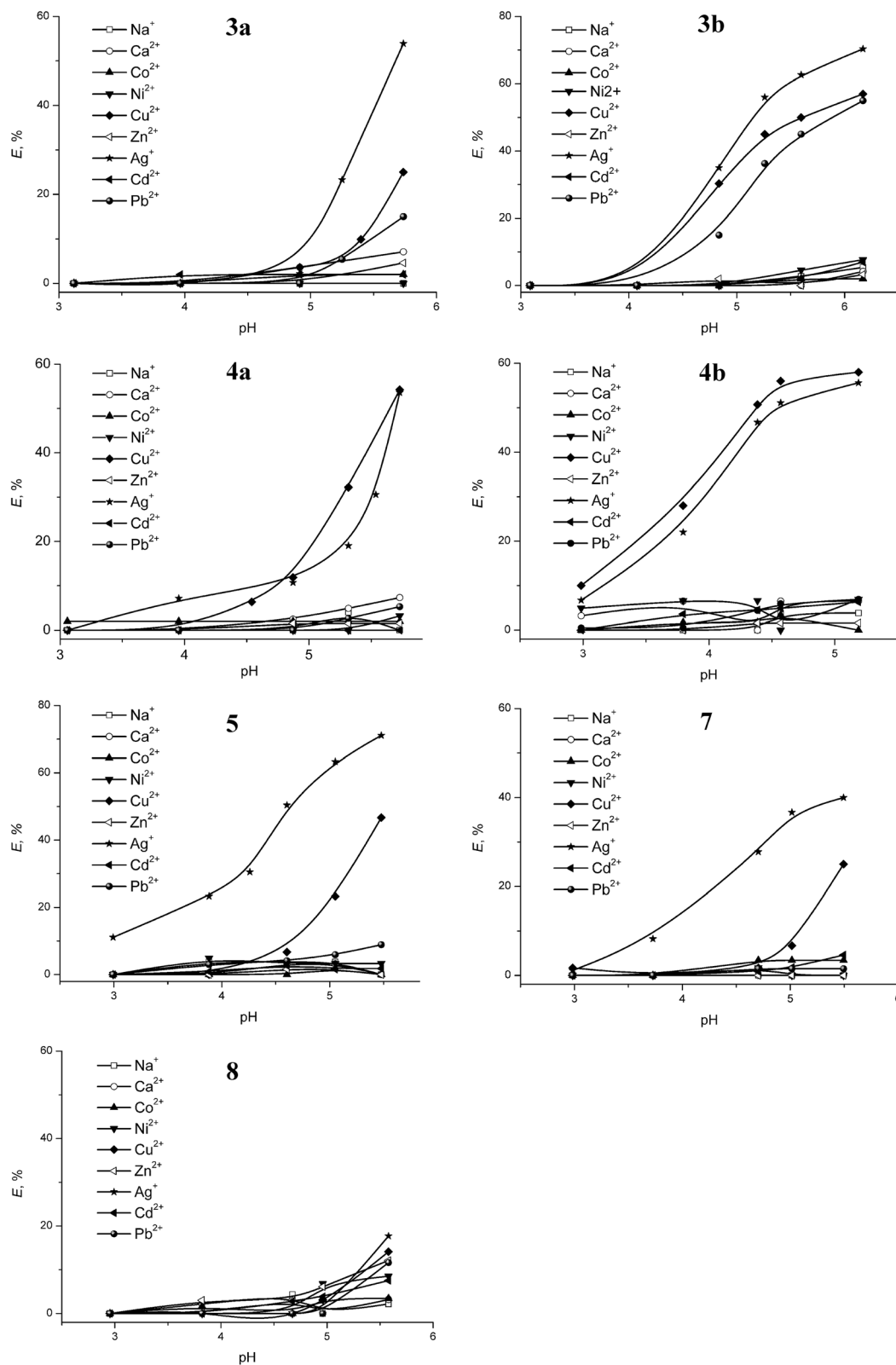


Fig. 5 Effect of pH on the extraction percentage ($E\%$) for different metal ions recovered by the ligands 3–7. $[L] = 0.2 \text{ mM}$, $[\text{metal ion}] = 0.1 \text{ mM}$, shaking time = 2 h at 298 K, phase ratio 1 : 1 (v/v), pH was adjusted with HNO_3 and Tris buffer.



chloroform at various pH (Fig. 5). The concentrations of ligands in chloroform and metal cations in aqueous phase were identical in all experiments.

It was established that an increase of pH of the aqueous phase up to pH \sim 5–6 leads to the enhancement of extraction ability of the investigated compounds towards metal ions. This phenomenon obviously may be explained by the ionization of carboxylic groups of the compounds at high pH, except the compound 7. From the other hand, a presence of lone electron pairs at the nitrogen atoms in dihydrazides **4a**, **4b**, **7** and dihydrazone **5** provides their basic properties. The protonation of hydrazide and hydrazone groups in these compounds, proceeding at the interface in acid conditions, prevents their complex formation with metal cations. This fact seems to be a main reason of the pH depending extraction revealed by the compound **7** having an incapable of deprotonation ester groups instead of the carboxylic ones.

All tetrathiacalix[4]arene derivatives **3–7** and the compound **8** have demonstrated a poor extraction towards s-elements ($E < 5\%$ for Na^+ and $E < 5\%$ for Ca^{2+}). At the same time, the tetrathiacalix[4]arene derivatives recovered selectively certain d- and p- ions at pH \sim 5–6. The selectivity of the ligands **3a** and **3b** towards the metal ions is reduced in the row $\text{Ag}^+ > \text{Cu}^{2+} > \text{Pb}^{2+}$.

The extraction efficiency goes up after removing *tert*-butyl groups from the upper rim of 1,3-*alternate* tetrathiacalix[4]arene on going from **3a** to **3b**: 54% and 70% for Ag^+ , 25% and 57% for Cu^{2+} , 15% and 45% for Pb^{2+} . In spite of decreasing the lipophilicity of the molecules, the weakening of steric difficulties obviously facilitates a process of metals recovery by the compound **3b**. It should be noted that the monomeric carboxylic acid **8** can also recover the Ag^+ (18%), Cu^{2+} (14%) and Pb^{2+} (11%) metal ions from aqua solutions, but the efficiency and selectivity are substantially lower than for tetrathiacalix[4]arene derivatives. It is worth noting that in the case of carboxylic derivatives of calix[4]-, [5]- and [6]arenes, a more effective binding of Pb^{2+} relatively Cu^{2+} has been established earlier.^{40–44}

A replacement of ester groups by the hydrazide ones for **4a** and **4b** leads to the dramatic decrease of the extraction efficiency towards Pb^{2+} ion (up to 5–7%). At the same time, the efficiency of the Ag^+ and Cu^{2+} ions recovery by these compounds remains the same. It is known that calix[*n*]arenes may exhibit an allosteric effect,^{22,23,45} which was previously detected for the binding of Pb^{2+} ions by the calix[5,6]arene carboxylic derivatives.^{43,46} In the case of tetrathiacalix[4]arenes **4a** and **4b**, the replacement of ester groups on the hydrazide ones can influence on the preorganization of carboxylic binding center similar to the allosteric effect. The hydrazide groups in the calix[4]arenes **4a**, **4b** and **7**, in contrast to the ester ones in the compounds **3a** and **3b**, are connected by the intramolecular hydrogen bonds (Fig. 2). Such circumstance obviously influences not only on the geometrical parameters, but on the rigidity of the molecules as well. For the dihydrazide derivatives **4a** and **4b**, these factors lead to a negative effect in the binding of Pb^{2+} ion which has greater ionic radius (1.19 Å) than Ag^+ (0.67 Å) and Cu^{2+} (0.73 Å) ions.⁴⁷

An enhanced efficiency of Ag^+ (71%) and Cu^{2+} (47%) ions recovery is observed for the hydrazone **5** in comparison with **3a** and **4a**. This fact is apparently caused by the simultaneous

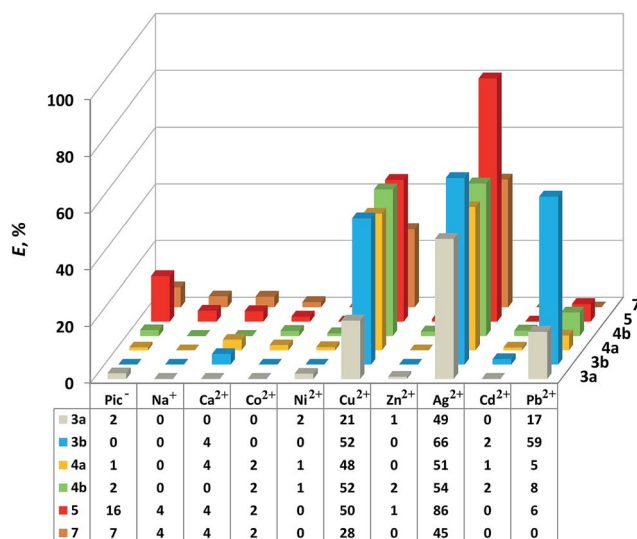


Fig. 6 Extraction percentage ($E\%$) for cations and Pic^- as a function of the nature of ligands **3–7**. $[\text{L}] = 0.2 \text{ mM}$, $[\text{Metal ion}] = 0.1 \text{ mM}$, $[\text{HPic}] = 0.4 \text{ mM}$, shaking time = 2 h at 298 K, phase ratio = 1 : 1 (v/v), pH = 6 ± 0.1 was adjusted with HNO_3 and Tris buffer.

cooperative participation of the carboxylic and hydrazone binding centers in the coordination of these cations.

In the case of dihydrazide **7** which does not have the carboxylic groups in its structure, a noticeable extraction of Ag^+ (40%) and Cu^{2+} (25%) cations is also observed. It was previously shown by the picrate liquid extraction method that 1,3-*alternate* tetrahydrazide of *p-tert*-butyl-tetrathiacalix[4]arene has revealed a high extraction efficiency exactly for these metal ions, whereas its *de-tert*-butyl analogue extracted more effectively in the row $\text{Ag}^+ > \text{Ni}^{2+} > \text{Cu}^{2+} > \text{Cd}^{2+}$.¹⁵ Generally, the data obtained for the selectivity of dihydrazide **7** are in accordance with the previously obtained results.

It should be noticed that the recovery of metal cations by the compound **7** can proceed only with participation of a counterion, but in the case of the compounds **3a**, **3b**, **4a**, **4b**, **5** and **8**, the extraction is probably realized *via* the ion-exchange process. To support this assumption we have performed an extraction experiment in the presence of Pic^- anion in aqua phase at pH = 6 ± 0.1 (Fig. 6).

Picrate is known to form preferentially second sphere complexes and can easier than “hard anions” be transferred from aqua to organic phase due to its enhanced lipophilic properties.⁴⁸ A transfer of this anion can be readily detected by UV-Vis spectrophotometric method. The analysis of the obtained UV data testified that the extraction of metal cations by the dihydrazide **7** proceeds with the participation of picrate anion (Fig. 6). In the case of dihydrazone **5**, a transfer of metal ions is also accompanied by the transfer of Pic^- anion. A mixed-mode extraction involving both the ion-exchange process as well as the metal salt transfer is evidently realized for **5**. A total extraction percentage of the Ag^+ and Cu^{2+} ions recovery by the compound **5** amounts to 136%. These facts indicate apparently that compound **5** coordinates the metal ions during extraction by both sides of tetrathiacalix[4]arene platform.



Thereby, it can be concluded that the synthesized bi-functional derivatives 3–7 are capable of an effective recovering of the Ag^+ and Cu^{2+} ions from aqua solutions as well as the Pb^{2+} ions in the case of the compounds 3a and 3b. The extraction of metal cations by the compounds 3–5 is mainly provided due to the binding of the metals by means of carboxylic groups. The preorganization of binding centers on the macrocyclic platform dramatically affects the efficiency and selectivity of the extraction. Moreover, the influence of the substitutes located on opposite side of the macrocycle in 1,3-*alternate* isomer and non-participated in the binding of the metal cations on the extraction properties has been also revealed. This phenomenon is obviously connected with the allosteric effect realized for this class of macrocyclic compounds.^{22,23,49}

3. Conclusions

In summary, we have reported the synthetic strategy for obtaining of novel bifunctional derivatives of 1,3-*alternate* tetrathiacalix[4]arene functionalized by carboxylic, ester, hydrazide and hydrazone groups immobilized on opposite sides of the macrocyclic platform. The detailed spectroscopic and structural characteristics as well as receptor properties towards metal ions for the synthesized compounds were established by using X-ray analysis, NMR, IR spectroscopy and liquid–liquid extraction method.

The structure of hydrazide fragments in derivatives 4a, 4b and 7 in solution as well as in solid state is characterized by *trans*-amide conformation which obviously promotes an effective chelate binding of metal ions. The acetylhydrazone fragments in tetrathiacalix[4]arene 5 adopt only the $E_{\text{C}=\text{N}}$ configuration. The $E_{\text{C}=\text{N}}/Z_{\text{C}=\text{N}}$ isomerization was not observed for this compound even at UV-light irradiation of its CDCl_3 solutions. At the same time the acetylhydrazone substitutes are in *cis*- and *trans*-amide conformations, that leads to the realization of a large number of spatial forms of dihydrazone 5 and as a consequence to the significant complication of NMR and IR spectra. The presence of carboxylic and hydrazide groups in the compounds 4a and 4b predisposed to the hydrogen bond formation does not produce a noticeable amount of associated molecules ($\sim 5\%$) in CDCl_3 – CD_3OD solutions.

The solvent extraction experiments have demonstrated that an increase of pH of the aqueous phase up to pH ~ 5 –6 leads to the rise of the extraction yields of the metals recovery by the compounds 3a, 3b, 4a, 4b, 5 due to the deprotonation of carboxylic groups. The hydrazide derivative 7 also demonstrates the pH depending extraction properties caused by the protonation of the nitrogen atoms in acid conditions.

The synthesized bifunctional derivatives 3–7 are capable of an effective recognition at pH ~ 5 –6 of the Ag^+ and Cu^{2+} ions from aqua solutions as well as the Pb^{2+} ions in the case of the compounds 3a and 3b. A nature of the binding centers and their preorganization on macrocyclic platform dramatically affect the efficiency and selectivity of the extraction. The extraction ability of dicarbonic acids 3a and 3b is significantly higher than for monomeric counterpart 8. The extraction efficiency goes up after removing *tert*-butyl groups from the upper rim of 1,3-

alternate tetrathiacalix[4]arene platform on going from 3a to 3b. A replacement of ester groups by the hydrazide and pyridinyl hydrazone ones (compounds 4a,b and 5) leads to the more selective recovery of Ag^+ and Cu^{2+} ions. In the case of dihydrazone 5, a mixed-mode extraction is realized which involves both the ion-exchange process as well as the metal salt transfer.

Thus, we can conclude that the developed synthetic strategy can be successfully applied for the preparing a wide range of bifunctional tetrathiacalix[4]arene derivatives capable of exhibiting excellent receptor properties.

4. Experimental section

4.1. General remarks

All chemicals were used as commercially received without further purification. CHCl_3 and DMF were distilled over P_2O_5 . THF was purified by distillation over KOH and sodium. CDCl_3 (99.8% isotopic purity) and DMSO-d_6 (99.5% isotopic purity) from Aldrich were used for NMR spectroscopy.

Microanalyses of C, H, and N were carried out with use of the CHN-3 analyser. Melting points of compounds were measured using a Boetius hotstage apparatus. The purity of the compounds was monitored by TLC. IR absorption spectra were recorded on a Vector-22 Bruker FT-IR spectrophotometer with the resolution of 4 cm^{-1} as Nujol emulsions and KBr pellets of compounds. Mass spectra (MALDI) were detected on a Finnigan MALDI-TOF Dynamo mass spectrometer. NMR experiments were performed on a Bruker AVANCE-600 spectrometer at 303 K equipped with a 5 mm broadband probehead working at 600.13 MHz in ^1H and 150.864 MHz in ^{13}C experiments. Chemical shifts in ^1H and ^{13}C spectra (Table 2) were reported relative to the solvent as internal standard (CDCl_3 $\delta(^1\text{H}) = 7.27\text{ ppm}$, $\delta(^{13}\text{C}) = 77.2\text{ ppm}$; DMSO $\delta(^1\text{H}) = 2.50\text{ ppm}$, $\delta(^{13}\text{C}) = 39.5\text{ ppm}$). Assignment was accomplished by means of 2D ^1H – ^{13}C HSQC and 2D ^1H – ^{13}C HMBC methods. The pulse programs of the HSQC, HMBC, FT-PGSE and NOESY experiments were taken from Bruker software library. The Fourier transform pulsed-gradient spin-echo (FT-PGSE) experiments⁵⁰ were performed by BPP-STE-LED (bipolar pulse pair-stimulated echo-longitudinal eddy current delay) sequence.⁵¹

4.2. Synthesis

The synthetic routes and the structural formulae of the investigated compounds are shown in Fig. 1. The parent tetrathiacalix[4]arenes 1a⁵² and 1b⁵³ as well as the tetrathiacalix[4]arenyloxyacetic acid ethyl esters in 1,3-*alternate* conformation 2a²¹ and 2b²² were prepared according to the literature methods. The dicarbonic acids 3a and 3b were obtained by the selective hydrolysis of tetrathiacalix[4]arene tetraesters 2a and 2b in the presence of an excess of Cs_2CO_3 as described by us earlier.²⁴ ^1H , ^{13}C NMR and IR data of the synthesized compounds are presented in Tables 1–3

5,11,17,23-Tetra-*tert*-butyl-25,27-di[(hydrazinocarbonyl)methoxy]-26,28-di[(hydroxycarbonyl)methoxy]-2,8,14,20-tetrathiacalix[4]arene (1,3-*alternate* 4a). To the solution of 1,3-*alternate* isomer of *p-tert*-butyltetrathiacalix[4]arene 3a (1.0 g; 1 mmol) in



the mixture of THF (15 ml) and MeOH (15 ml), an excess of $\text{NH}_2\text{-NH}_2\cdot\text{H}_2\text{O}$ (98%, 0.98 ml; 32 mmol) was added under stirring. The reaction mixture was heated at 65 °C for 2 days. Then the solvent was removed with using of vacuum distillation at 80 °C and solid remainder was suspended in H_2O (20 ml). The acidity of the solution was kept by the concentrated HCl (0.3–0.6 ml) to reach pH ~5. The matured precipitate was filtrated and washed with water to a neutral pH. After drying the precipitate under reduced pressure at 90 °C, the target product **4a** was obtained as a white powder. Yield: 0.90 g (92%). Mp 272 °C (decomp.). Anal. Calcd for $\text{C}_{48}\text{H}_{60}\text{N}_4\text{O}_{10}\text{S}_4$ (981.27): C, 58.75; H, 6.16; S, 13.07; N, 5.71. Found: C, 58.44; H, 6.10; S, 13.36; N, 5.47. Mass spectrum (MALDI-TOF): m/z : = 981 $[\text{M} + \text{H}]^+$; 1003 $[\text{M} + \text{Na}]^+$; 1019 $[\text{M} + \text{K}]^+$.

25,27-Di[(hydrazinocarbonyl)methoxy]-26,28-di[(hydroxycarbonyl)methoxy]-2,8,14,20-tetrathiacalix[4]arene (1,3-alternate 4b). Prepared similar to the procedure for **4a** using tetrathiacalix[4]arene **3b** (1.0 g; 1.3 mmol) and an excess of $\text{NH}_2\text{-NH}_2\cdot\text{H}_2\text{O}$ (98%; 1.3 ml; 42 mmol). The target product **4b** was obtained as a white powder. Yield: 0.84 g (85%). Mp 287 °C (decomp.). Anal. Calcd for $\text{C}_{32}\text{H}_{28}\text{N}_4\text{O}_{10}\text{S}_4$ (756.85): C, 50.78; H, 3.73; S, 16.95; N, 7.4. Found: C, 50.45; H, 3.48; S, 16.57; N, 7.75. Mass spectrum (MALDI-TOF): m/z : = 757 $[\text{M} + \text{H}]^+$; 779 $[\text{M} + \text{Na}]^+$; 795 $[\text{M} + \text{K}]^+$.

5,11,17,23-Tetra-tert-butyl-25,27-di[(2-pyridinylmethylidene)-hydrazinocarbonyl]-methoxy-[(hydrazinocarbonyl)methoxy]-26,28-di[(hydroxycarbonyl)methoxy]-2,8,14,20-tetrathiacalix[4]arene (1,3-alternate 5). To the solution of 1,3-alternate isomer **4a** (0.49 g; 0.5 mmol) in the mixture of EtOH (12 ml) and DMF (12 ml), an excess of the picolinaldehyde (0.28 ml, 3 mmol) was added under stirring. The reaction mixture was heated at 80 °C for 36 h. The solid formed was filtered off and washed by EtOH. The product **5** was obtained as a white powder. Yield: 0.31 g (53%). Mp 198 °C. Anal. Calcd for $\text{C}_{60}\text{H}_{66}\text{N}_6\text{O}_{10}\text{S}_4$ (1159.46): C, 62.15; H, 5.74; S, 11.06; N, 7.25. Found: C, 62.25; H, 5.48; S, 11.37; N, 7.45. Mass spectrum (MALDI-TOF): m/z : = 1159 $[\text{M} + \text{H}]^+$; 1181 $[\text{M} + \text{Na}]^+$.

25,27-Di[(chlorocarbonyl)methoxy]-26,28-di[(hydroxycarbonyl)methoxy]-2,8,14,20-tetrathiacalix[4]arene (1,3-alternate 6). To the diester-dicarboxylic acid **3b** (1.25 g, 1.2 mmol), a fresh distilled SOCl_2 (10 ml, 138 mmol) was added. The reaction mixture was refluxed for 2 h under argon atmosphere. Then the solvent was removed from the reaction mixture by distillation. The residue was drying under vacuum at 80 °C. The target product **6** was obtained as a yellowish powder. Yield: 1.2 g (96%). Mp 148–152 °C.

25,27-Di[(hydrazinocarbonyl)methoxy]-26,28-di[(ethoxycarbonyl)methoxy]-2,8,14,20-tetrathiacalix[4]arene (1,3-alternate 7). A mixture of absolute THF (40 ml) and $\text{NH}_2\text{-NH}_2$ (0.75 ml, 24 mmol) previously distilled on KOH was cooled down to 0 °C under vigorous stirring. Then dichloroanhydride **6** (0.82 g, 1 mmol) in 20 ml of THF was added dropwise during 0.5 h and the reaction mixture was vigorously stirred at 0 °C during 1.5 h. Thereafter, a solvent was removed under reduced pressure at 40 °C. The obtained residue has been treated by H_2O (50 ml) and extracted with dichloromethane (50 ml). The organic layer was separated, washed twice with water and dried with MgSO_4 . After filtration, the solvent was removed by distillation. The

residue was drying under vacuum at 80 °C. Yield: 0.59 g (72%). Mp 169–171 °C. Anal. Calcd for $\text{C}_{36}\text{H}_{36}\text{N}_4\text{O}_{10}\text{S}_4$ (812.95): C, 53.19; H, 4.46; S, 15.78; N, 6.89. Found: C, 53.33; H, 4.27; S, 15.34; N, 6.77. Mass spectrum (MALDI-TOF): m/z : = 813 $[\text{M} + \text{H}]^+$; 835 $[\text{M} + \text{Na}]^+$; 851 $[\text{M} + \text{K}]^+$.

4-tert-Butyl-phenoxyacetic acid (8). To 4-tert-butyl-phenoxyacetic acid ethyl ester⁵⁴ (2.13 g, 9 mmol) the solution of KOH (0.76 g, 13.5 mmol) in water (6 ml) was added. The reaction mixture was heated at 70 °C for 0.5 h. After cooling it down to 5–10 °C, the concentrated 37% HCl (2 ml) was added under stirring. The resulting mixture was extracted with dichloromethane (3 × 5 ml). The organic layer was separated and washed twice with 10% HCl (10 ml) and water. After drying it over MgSO_4 , the solvent was removed by distillation. The residue was dried under vacuum at 100 °C. The crude product was recrystallized from hexane (12 ml). Yield: 1.64 g (87%). Mp 86–88 °C. Anal. calcd for $\text{C}_{12}\text{H}_{16}\text{O}_3$ (208.25): C, 69.21; H, 7.74. Found: C, 69.33; H, 7.67.

4.3. Crystal structure

The X-ray diffraction data for the crystals of **4a**, **4a**·2 NH_2NH_2 , **4b**· NH_2NH_2 , **5** and **7** were collected on a Bruker Smart Apex II CCD diffractometer in the ω and ϕ -scan modes using graphite monochromated Mo K_α (λ = 0.71073 Å) radiation at 150(2) K (**4a**) and at 296(2) K for other samples. Data were corrected for the absorption effect using SADABS program.⁵⁵ The structures were solved by direct method and refined by the full matrix least-squares using SHELXTL⁵⁶ and WinGX⁵⁷ programs. All non-hydrogen atoms were refined anisotropically. The hydrogen atoms were inserted at calculated positions and refined as riding atoms except the hydrogen atoms on solvents molecules, hydroxyl groups and amino groups, which were located from difference maps and refined using a riding model. Data collections: images were indexed, integrates, and scaled using the APEX2 (ref. 58) data reduction package. Analysis of the intermolecular interactions was performed using the program PLATON.⁵⁹ Mercury program package⁶⁰ was used for figures preparation.

Crystallographic data (excluding structure factors) for the structure **4a**, **4a**·2 NH_2NH_2 , **4b**· NH_2NH_2 , **5** and **7** have been deposited in the Cambridge Crystallographic Data Centre as supplementary publication numbers CCDC 1437363–1437367 respectively.

Crystallographic data for 4a. $\text{C}_{48}\text{H}_{60}\text{N}_4\text{O}_{10}\text{S}_4\cdot\text{C}_2\text{H}_6\text{OS}$, colorless plate, $0.25 \times 0.14 \times 0.07 \text{ mm}^3$, M = 1059.36, monoclinic, a = 13.163(5) Å, b = 29.199(11) Å, c = 15.137(5) Å, β = 91.473(5)°, V = 5816(4) Å³, T = 150(2) K, space group $P2_1/n$, Z = 4, $\mu(\text{Mo K}_\alpha)$ = 0.255 mm^{−1}, ρ_{calc} = 1.210 g cm^{−3}, $F(000)$ = 2248, 67 104 reflections measured, 15 532 independent reflections (R_{int} = 0.1253), 797 parameters, 447 restraints. Final indices: R_1 = 0.1075, wR_2 = 0.2694 (4474 reflections with $I > 2\sigma_1$), R_1 = 0.2629 (all data), wR_2 = 0.3323 (all data), GoF = 0.875, largest difference in peak and hole (0.545 and −0.533 e Å^{−3}).

Crystallographic data for 4a·2 NH_2NH_2 . $\text{C}_{48}\text{H}_{60}\text{N}_4\text{O}_{10}\text{S}_4\cdot 2\text{N}_2\text{H}_4$, colorless block, $0.41 \times 0.32 \times 0.13 \text{ mm}^3$, M = 1045.34, monoclinic, T = 296(2) K, a = 35.723(8) Å, b = 13.121(3) Å, c =



25.422(6) Å, $\beta = 108.104(3)^\circ$, $V = 11\,326(4)\text{ Å}^3$, space group $C2/c$, $Z = 8$, $\mu(\text{Mo K}\alpha) = 0.226\text{ mm}^{-1}$, $\rho_{\text{calc}} = 1.226\text{ g cm}^{-3}$, $F(000) = 4448$, 13 275 independent reflections, 3458 reflections with $I > 2\sigma_I$, 650 parameters, 121 restraints. Final indices: $R_1 = 0.0952$ ($I > 2\sigma_I$), $wR_2 = 0.2080$ ($I > 2\sigma_I$), $R_1 = 0.2839$ (all data), $wR_2 = 0.2684$ (all data), GoF (goodness-of-fit on F^2) = 0.763, largest difference in peak and hole (0.562 and -0.481 e Å^{-3}).

Crystallographic data for 4b·NH₂NH₂. C₃₂H₂₈N₄O₁₀S₄·N₂·H₄·4H₂O, colorless block, $0.439 \times 0.387 \times 0.246\text{ mm}^3$, $M = 860.94$, monoclinic, $T = 296(2)\text{ K}$, $a = 10.3996(8)\text{ Å}$, $b = 16.7027(13)\text{ Å}$, $c = 21.9306(18)\text{ Å}$, $\beta = 90.2490(10)^\circ$, $V = 3809.3(5)\text{ Å}^3$, space group $P2_1/c$, $Z = 4$, $\mu(\text{Mo K}\alpha) = 0.325\text{ mm}^{-1}$, $\rho_{\text{calc}} = 1.501\text{ g cm}^{-3}$, $F(000) = 1800$, 41 221 reflections measured, 9032 independent reflections ($R_{\text{int}} = 0.0261$), 7845 reflections with $I > 2\sigma_I$, 522 parameters, 38 restraints. Final indices: $R_1 = 0.0589$ ($I > 2\sigma_I$), $wR_2 = 0.1455$ ($I > 2\sigma_I$), $R_1 = 0.0669$ (all data), $wR_2 = 0.1503$ (all data), GoF = 1.110, largest difference in peak and hole (0.944 and -0.620 e Å^{-3}).

Crystallographic data for 5. C₆₀H₆₆N₆O₁₀S₄·CH₃OH, colorless block, $0.109 \times 0.083 \times 0.037\text{ mm}^3$, $M = 1191.47$, triclinic, $T = 296(2)\text{ K}$, $a = 13.783(13)\text{ Å}$, $b = 15.091(14)\text{ Å}$, $c = 17.68(2)\text{ Å}$, $\alpha = 71.891(12)^\circ$, $\beta = 85.063(18)^\circ$, $\gamma = 69.735(12)^\circ$, $V = 3278(6)\text{ Å}^3$, space group $P\bar{1}$, $Z = 2$, $\mu(\text{Mo K}\alpha) = 0.204\text{ mm}^{-1}$, $\rho_{\text{calc}} = 1.207\text{ g cm}^{-3}$, $F(000) = 1260$, 39 008 reflections measured, 16 924 independent reflections ($R_{\text{int}} = 0.2261$), 3408 reflections with $I > 2\sigma_I$, 752 parameters, 690 restraints. Final indices: $R_1 = 0.1530$ ($I > 2\sigma_I$), $wR_2 = 0.4178$ ($I > 2\sigma_I$), $R_1 = 0.3498$ (all data), $wR_2 = 0.4690$ (all data), GoF = 0.897, largest difference in peak and hole (0.692 and -0.483 e Å^{-3}).

Crystallographic data for 7. C₃₆H₃₆N₄O₁₀S₄, colorless plate, $0.43 \times 0.25 \times 0.08\text{ mm}^3$, $M = 812.93$, monoclinic, $T = 296(2)\text{ K}$, $a = 9.9936(9)\text{ Å}$, $b = 18.4615(15)\text{ Å}$, $c = 20.6801(17)\text{ Å}$, $\beta = 94.046(5)^\circ$, $V = 3805.9(6)\text{ Å}^3$, space group $P2_1/c$, $Z = 4$, $\mu(\text{Mo K}\alpha) = 0.312\text{ mm}^{-1}$, $\rho_{\text{calc}} = 1.419\text{ g cm}^{-3}$, $F(000) = 1696$, 47 387 reflections measured, 9433 independent reflections ($R_{\text{int}} = 0.0689$), 4234 reflections with $I > 2\sigma_I$, 618 parameters, 354 restraints. Final indices: $R_1 = 0.0606$ ($I > 2\sigma_I$), $wR_2 = 0.1482$ ($I > 2\sigma_I$), $R_1 = 0.1632$ (all data), $wR_2 = 0.1905$ (all data), GoF = 1.010, largest difference in peak and hole (0.622 and -0.342 e Å^{-3}).

4.4. Competitive extraction of metal cations

The CHCl₃ was saturated with H₂O to prevent volume changes during the extraction. Aqueous solutions (10 ml) containing metal salts (NaNO₃; Ca(NO₃)₂·4H₂O, Co(NO₃)₂·6H₂O, Ni(NO₃)₂·6H₂O, Cu₂(CH₃COO)₄·2H₂O, Zn(NO₃)₂·4H₂O, Pb(NO₃)₂, AgNO₃ and Cd(CH₃COO)₂·2H₂O) were prepared with the concentration for each metal ion $1 \times 10^{-4}\text{ M}$. To adjust pH of the solutions a tris(hydroxymethyl)aminomethane (Tris) and HNO₃ were added arbitrarily. The equal volumes of the aqueous mixture of metal salts and the CHCl₃ solution of extractant (C_{L3a,L3b,L4a,L4b,L5,L7,L8} = $2 \times 10^{-4}\text{ M}$) were magnetically stirred in a flask. The extraction equilibrium was reached after vigorous stirring for 1.5 h at 25 °C. Then two phases were allowed to settle for 1 h and afterwards separated by centrifugation.

The pH and concentration of metal ions in aqueous phase before and after extraction were measured. The relative

concentrations of the cations in the aqueous phase were determined by the applying of atomic absorbance spectrometer AAS 1 N (Carl Zeiss Jena) with the use of oxidative air-acetylene flame. Quantification was made by referring on a standard solution containing a mixture of salts. Blank experiments without added hosts were carried out under the same experimental conditions. The percentage of extraction was calculated as a ratio $E\% = 100 \times (C_0 - C_1)/C_0$, where C_0 and C_1 are the initial and equilibrium concentrations of metal ion in the aqueous solution determined before and after extraction, respectively. $E\%$ uncertainties are generally $\leq 2\%$.

Notes and references

- 1 J. W. Steed and J. L. Atwood, *Supramolecular Chemistry*, Wiley, Chichester, UK, 2000.
- 2 H. W. Lin, *Angew. Chem.*, 2007, **119**, 911–914.
- 3 L. Prodi, F. Bolletta, M. Montaliti and N. Zaccheroni, *Coord. Chem. Rev.*, 2000, **205**, 59–83.
- 4 V. M. Mirsky and A. K. Yatsimirsky, *Artificial Receptors for Chemical Sensors*, Wiley-VCH, Weinheim, 2010.
- 5 G. Yu, C. Han, Z. Zhang, J. Chen, X. Yan, B. Zheng, S. Liu and F. Huang, *J. Am. Chem. Soc.*, 2012, **134**, 8711–8717.
- 6 G. Yu, M. Xue, Z. Zhang, J. Li, C. Han and F. Huang, *J. Am. Chem. Soc.*, 2012, **134**, 3248–3251.
- 7 G. Yu, X. Zhou, Z. Zhang, C. Han, Z. Mao, C. Gao and F. Huang, *J. Am. Chem. Soc.*, 2012, **134**, 19489–19497.
- 8 C. D. Gutsche, *Calixarenes: An Introduction*, The Royal Society of Chemistry, Cambridge, 2nd edn, 2008.
- 9 Z. Asfary, V. Böhmer, J. Harrowfield and J. Vicens, *Calixarenes 2001*, Kluwer Academic Publishers, Dordrecht, The Netherlands, 2001.
- 10 O. K. Sung and C. N. Kye, *Bull. Korean Chem. Soc.*, 2002, **23**, 640–642.
- 11 K. Manoj, D. Abhimanew and B. Vandana, *Tetrahedron*, 2009, **65**, 7510–7515.
- 12 X.-I. Ni, H. Tomiyasu, T. Shimizu, C. Pérez-Casas, Z. Xi and T. Yamato, *J. Inclusion Phenom. Macrocyclic Chem.*, 2010, **68**, 99–108.
- 13 J. Mendez-Arroyo, J. Barroso-Flores, A. M. Lifschitz, A. A. Sarjeant, C. L. Stern and C. A. Mirkin, *J. Am. Chem. Soc.*, 2014, **136**, 10340–10348.
- 14 S. N. Podyachev, S. N. Sudakova, V. V. Syakaev, A. K. Galiev, R. R. Shagidullin and A. I. Kononov, *Supramol. Chem.*, 2008, **20**, 479–486.
- 15 S. N. Podyachev, N. E. Burmakina, S. N. Sudakova, V. V. Syakaev and A. I. Kononov, *Supramol. Chem.*, 2010, **22**, 339–346.
- 16 F. Botha, J. Budka, V. Eigner, O. Hudeček, L. Vrzal, I. Císařová and P. Lhoták, *Tetrahedron*, 2014, **70**, 477–483.
- 17 L. Kovbasyuk and R. Krämer, *Chem. Rev.*, 2004, **104**, 3161–3187.
- 18 S. Patai, *The Chemistry of Carboxylic Acids and Esters*, London, 1969.
- 19 S. Patai, *The Chemistry of Acid Derivatives*, New York, 1979.
- 20 N. Iki, F. Narumi, T. Fujimoto, N. Morohashi and S. Miyano, *J. Chem. Soc., Perkin Trans. 2*, 1998, 2745–2750.



- 21 P. Lhotak and J. Sykora, *Collect. Czech. Chem. Commun.*, 2000, **65**, 757–771.
- 22 A. Ikeda, Y. Suzuki, M. Yoshimura and S. Shinkai, *Tetrahedron*, 1998, **54**, 2497–2508.
- 23 J. Budka, P. Lhotak, V. Michlova and I. Stibor, *Tetrahedron Lett.*, 2001, **42**, 1583–1586.
- 24 S. N. Podyachev, S. N. Sudakova, B. M. Gabidullin, V. V. Syakaev, A. T. Gubaidullin, W. Dehaen and A. I. Konovalov, *Tetrahedron Lett.*, 2012, **53**, 3135–3139.
- 25 S. N. Podyachev, V. V. Syakaev, S. N. Sudakova, R. R. Shagidullin, D. V. Osyana, L. V. Avvakumova, B. I. Buzykin, Sh. K. Latypov, V. D. Habicher and A. I. Konovalov, *J. Inclusion Phenom.*, 2007, **58**, 55–61.
- 26 E. Quinlan, S. E. Matthews and T. Gunnlaugsson, *J. Org. Chem.*, 2007, **72**, 7497–7503.
- 27 M. A. Qazi, Ü. Ocak, M. Ocak, S. Memon and I. B. Solangi, *J. Fluoresc.*, 2013, **23**, 575–590.
- 28 X. Li, S.-L. Gong, W.-P. Yang, Y.-Y. Chen and X.-G. Meng, *Tetrahedron*, 2008, **64**, 6230–6237.
- 29 N. B. Colthup, L. H. Daly and S. E. Wiberley, *Introduction to infrared and Raman spectroscopy*, London, Academic Press, 1964.
- 30 L. J. Bellamy, *The IR Spectra of Complex Organic Molecules*, London: Methuen, N.-Y., Wiley, 2nd edn, 1958.
- 31 S. N. Podyachev, N. E. Burmakina, V. V. Syakaev, S. N. Sudakova, R. R. Shagidullin and A. I. Konovalov, *Tetrahedron*, 2009, **65**, 408–417.
- 32 S. N. Podyachev, B. M. Gabidullin, V. V. Syakaev, S. N. Sudakova, A. T. Gubaidullin, W. D. Habicher and A. I. Konovalov, *J. Mol. Struct.*, 2011, **1001**, 125–133.
- 33 C. Jaime, J. de Mendoza, P. Prados, P. Nieto and C. Sanchez, *J. Org. Chem.*, 1991, **56**, 3372–3376.
- 34 P. Lhotak, *Eur. J. Org. Chem.*, 2004, 1675–1692.
- 35 V. V. Syakaev, S. N. Podyachev, B. I. Buzykin, S. K. Latypov, V. D. Habicher and A. I. Konovalov, *J. Mol. Struct.*, 2006, **788**, 55–62.
- 36 V. V. Syakaev, S. N. Podyachev, A. T. Gubaidullin, S. N. Sudakova and A. I. Konovalov, *J. Mol. Struct.*, 2008, **885**, 111–121.
- 37 P. Stilbs, *Prog. Nucl. Magn. Reson. Spectrosc.*, 1987, **19**, 1–45.
- 38 Y. Cohen, L. Avram and L. Frish, *Angew. Chem., Int. Ed. Engl.*, 2005, **44**, 520–554.
- 39 T. Brand, E. J. Cabrita and S. Berger, *Prog. Nucl. Magn. Reson. Spectrosc.*, 2005, **46**, 159–196.
- 40 N. T. K. Dung and R. Ludwig, *New J. Chem.*, 1999, **23**, 603–607.
- 41 K. Ohto, Y. Fujimoto and K. Inoue, *Anal. Chim. Acta*, 1999, **387**, 61–69.
- 42 B. B. Adhikari, M. Gurung, H. Kawakita and K. Ohto, *Analyst*, 2011, **136**, 3758–3769.
- 43 B. B. Adhikari, M. Gurung, A. B. Chetry, H. Kawakita and K. Ohto, *RSC Adv.*, 2013, **3**, 25950–25959.
- 44 B. B. Adhikari, M. Gurung, H. Kawakita and K. Ohto, *Solvent Extr. Ion Exch.*, 2013, **31**, 483–498.
- 45 Z. Asfari, V. Lamare, J. F. Dozol and J. Vicens, *Tetrahedron Lett.*, 1999, **40**, 691.
- 46 B. B. Adhikari, K. Ohto, M. Gurung and H. Kawakita, *Tetrahedron Lett.*, 2010, **51**, 3481–3485.
- 47 R. D. Shannon, *Acta Crystallogr., Sect. A: Cryst. Phys., Diffraction, Theor. Gen. Crystallogr.*, 1976, **32**, 751–767.
- 48 Y. Takeda, F. Vögtle and E. Weber, *Topics of curr. Chem.*, Springer, Berlin, 1984, vol. 121, pp. 1–38.
- 49 Z. Asfari, V. Lamare, J. F. Dozol and J. Vicens, *Tetrahedron Lett.*, 1999, **40**, 691–694.
- 50 W. S. Price, *NMR Studies of Translational Motion*, University Press, Cambridge, 2009.
- 51 D. Wu, A. Chen and C. S. Johnson, *J. Magn. Reson.*, 1995, **115**, 260–264.
- 52 H. Kumagai, M. Hasegawa, S. Miyanari, Y. Sugawa, Y. Sato, T. Hori, S. Ueda, H. Kamiyama and S. Miyano, *Tetrahedron Lett.*, 1997, **38**, 3971–3972.
- 53 Y. Higuchi, M. Narita, T. Niimi, N. Ogawa, F. Hamada, H. Kumagai, N. Iki, S. Miyano and C. Kabuto, *Tetrahedron*, 2000, **56**, 4659–4666.
- 54 A. Yamada, T. Murase, K. Kikukawa, T. Arimura and S. Shinkai, *J. Chem. Soc., Perkin Trans. 2*, 1991, **5**, 793–797.
- 55 G. M. Sheldrick, *SADABS, Program for empirical X-ray absorption correction*, Bruker-Nonius, 2004.
- 56 G. M. Sheldrick, *SHELXTL v.6.12, Structure Determination Software Suite*, Bruker AXS, Madison, Wisconsin, USA, 2000.
- 57 L. J. Farrugia, *J. Appl. Crystallogr.*, 1999, **32**, 837–838.
- 58 APEX2 (Version 2.1), SAINTPlus, Data Reduction and Correction Program (Version 7.31A), BrukerAXS Inc., Madison, Wisconsin, USA, 2006.
- 59 A. L. Spek, *J. Appl. Crystallogr.*, 2003, **36**, 7–13.
- 60 I. J. Bruno, J. C. Cole, P. R. Edgington, M. K. Kessler, C. F. Macrae, P. McCabe, J. Pearson and R. Taylor, *Acta Cryst.*, 2002, **58**, 389–397.

