



Cite this: *RSC Adv.*, 2017, 7, 11030

Study on the synthesis and tribological properties of anti-corrosion benzotriazole ionic liquid†

S. Zhang,^a L. Ma,^a R. Dong,^a C. Y. Zhang,^a W. J. Sun,^a M. J. Fan,^{*a} D. S. Yang,^{*a} F. Zhou^b and W. M. Liu^b

Anti-corrosion benzotriazole ionic liquid (IL, BTAP₄₄₄₄) was synthesized by neutralizing 1H-benzotriazole (BTAH) with tetrabutylphosphonium hydroxide (P₄₄₄₄OH). The physicochemical and tribological properties of BTAP₄₄₄₄ were measured and the results indicated that it had good lubricating properties for steel/steel, steel/copper and steel/aluminum friction pairs in reducing the friction coefficient and wear volume, especially when it worked for the steel/copper friction pair both at room temperature (RT) and 100 °C. Good performances were also found in terms of its anti-corrosion ability, thermal and hydrolysis stability, and viscosity temperature characteristics compared with commercially available synthetic oil poly-alpha-olefin (PAO 10) and traditional IL 1-butyl-3-methyl imidazolium tetrafluoroborate (bmimBF₄). The designing purpose of BTAP₄₄₄₄ is to enhance the anti-corrosion ability of the IL lubricant attributed to the introduction of a nitrogen heterocyclic structure, 1H-benzotriazole, which is normally used as antirust and corrosion inhibitor in mechanical engineering. The corrosion test results confirm our assumption and BTAP₄₄₄₄ shows excellent anti-corrosion ability for all the used friction pairs including steel/steel, steel/copper and steel/aluminum contacts.

Received 26th November 2016

Accepted 6th February 2017

DOI: 10.1039/c6ra27376a

rsc.li/rsc-advances

1. Introduction

Friction is everywhere, whether it is a bicycle, a car, a ship or an aircraft, and there is a phenomenon of friction and wear in the moving parts of mechanical equipment. Therefore, a lubricating oil is needed to reduce friction and wear. However, during long-term mechanical friction, a lubricating oil may degrade and produce some pollutants, which can cause damage to the friction surface. Especially in the presence of oxygen, water, acid or under other harsh conditions, obvious oxidation or corrosion phenomena will occur on the metal friction surface, increasing the mechanical wear and reducing the lubricating performance of a lubricating oil.¹ Therefore, it is necessary to use lubricating oil with multi-functional properties, such as anti-oxidation, corrosion resistance and so on, to improve energy utilization and material durability.

As is known to all, ionic liquids (ILs) have been used as a kind of high performance lubricants and attracted great attention in recent years.^{2,3} So, what is an IL? It is a salt with melting point below 100 °C and consists of only cation and anion which are all capable of facile tuning.⁴⁻⁶ ILs have some

excellent physical properties, such as low volatility, nonflammability, low melting point, good viscosity temperature characteristic, large heating capacity, outstanding chemical stability and so on.⁷⁻¹⁰ Because of these excellent characteristics, ILs can be used as high-performance lubricants or lubricant additives for base oil, grease or water.¹¹⁻¹⁶ However, when traditional ILs are used as lubricants or lubricant additives, they are often accompanied by corrosion during the process of friction, because most of them contain halogens, such as BF₄⁻, PF₆⁻, X⁻ (F⁻, Cl⁻, Br⁻, etc.). These halogen containing anions are prone to hydrolyze and generate acid compounds (HX), which erodes the surface of friction pairs.¹⁷ Metal corrosion is a serious problem in industry fields.^{18,19} It has attracted particular attention, and the protection of metal and its alloys is of paramount important.¹⁸ There are two ways to solve the problem of corrosion. One is completely avoiding the using of halogen containing anion. For example, the use of amino acid, tricyanomethanide, dicyanamide, and chelated orthoborate anion-based ILs.²⁰⁻²⁵ Although these ILs have low corrosion ability, their thermal stability is found to be poor. Besides, their synthesis procedure is still complex, therefore, it hinders its potential application.²⁶ Another way to avoid corrosion is adding preservatives or introducing a functional group with corrosion resistance to original IL molecules. For example, reduce corrosion by using 1H-benzotriazole (BTAH) as a preserving additive. Since the first report of BTAH in 1947, it has attracted much attention as a corrosion inhibitor in different applications.²⁷⁻³¹ BTAH and its derivatives were found to be effective

^aShaanxi Key Laboratory of Phytochemistry, College of Chemistry & Chemical Engineering, Baoji University of Arts and Sciences, Baoji 721013, China. E-mail: mingjinfan@163.com; Fax: +86-9173566000; Tel: +86-9173565561

^bState Key Laboratory of Solid Lubrication, Lanzhou Institute of Chemical Physics, Chinese Academy of Sciences, Lanzhou 730000, China

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c6ra27376a



corrosion inhibitors to various metals, such as copper, mild steel, *etc.* because of its unique space structure.^{32–39} Furthermore, BTAH is found to be the most effective copper corrosion inhibitors among BTAH and other two benzotriazole derivatives, 1-hydroxybenzotriazole (BTAOH) and 3-amino-1,2,4-triazole (ATA).^{18,40,41} However, free BTAH is not stable enough at high temperature, and its thermal decomposition temperature is only about 150 °C. At the same time, it has limited sublimation temperature of 98–99 °C, so BTAH cannot be used as an effective anticorrosive additive to IL lubricants under high temperature condition.⁴² The work of designing new IL molecular which has benzotriazole as a substituted group, such as the synthesis of imidazolium ILs bearing benzotriazole group, were conducted by Cai *et al.*, but the synthetic process of these ILs are too complicated to be conducted in practical using.^{43,44}

As far as we known, the study of directly using BTAH as an anion donor to design new IL lubricants has not been reported. Therefore, in this paper, a new type of anti-corrosion IL BTAP₄₄₄₄ was synthesized by using readily available chemicals BTAH and terabutylphosphonium hydroxide (P₄₄₄₄OH) as raw materials, which was found to have good performances in terms of thermal and hydrolysis stability, viscosity temperature characteristic, tribological property and especially anti-corrosion ability.

2. Experimental

2.1 Chemicals and synthesis of the ILs

Synthetic oil poly-alpha-olefin (PAO 10) was purchased from Exxon Mobil Companies and conventional IL 1-butyl-3-methyl imidazolium tetrafluoroborate (bmimBF₄) was provided by Lanzhou Institute of Chemical Physics. Other used chemicals were commercially available: benzotriazole (99%, J&K), P₄₄₄₄OH solution (40% in water, Aladdin). IL BTAP₄₄₄₄ was synthesized through an economical and feasible method. The detail process is presented in the ESI.†

2.2 Characterization

The molecular structure of BTAP₄₄₄₄ was determined by proton nuclear magnetic resonance spectroscopy (¹H NMR, 400 MHz), carbon nuclear magnetic resonance spectroscopy (¹³C NMR, 100 MHz) which were carried on an Agilent 400 MHz nuclear magnetic resonance spectrometer (NMR) by using CDCl₃ as a solvent and tetramethylsilane (TMS) as an external standard. The high-resolution mass spectra (HRMS) were conducted on an AB SCIEX TOFTM 4600 MS equipped with an electrospray ionization (ESI) probe. The Fourier transform infrared (FT-IR) spectra was recorded on a Nicolet Nexus 670 spectrometer (Thermo Fisher, USA) with wave numbers ranging from 700 cm^{−1} to 4000 cm^{−1}.

2.3 Thermal stability

The thermal stabilities of BTAP₄₄₄₄ and the reference samples (PAO 10, bmimBF₄ and BTAH) were studied on a Netzsch synchronous thermal analyzer system (DSC/DTA-TG, STA 449 F3). The temperature was set up to increase from RT to 600 °C

with a heating rate of 10 °C min^{−1}. The analysis was conducted in nitrogen atmosphere with a flow rate of 50 mL min^{−1}.

2.4 Viscosity

The kinetic viscosities of BTAP₄₄₄₄ and the reference samples (PAO 10 and bmimBF₄) were measured at 40 °C and 100 °C respectively by using a SYP1003-III kinematic viscosity tester. Viscosity indexes of the samples were calculated to evaluate their viscosity temperature characteristics.

2.5 Hydrolysis stability analysis

The hydrolysis stabilities of BTAP₄₄₄₄ and the reference samples were measured according to the procedure described in our previous work.^{45–48} Each sample was mixed with equal molar amount of water and then stirred at RT for a certain time. After 30 min, 1 h, 2 h until 48 h, their pH values were measured to indicate their hydrolysis stability. All of the ILs were found to be miscible with water and the mixed solutions were used to measure the pH value changes. But for PAO 10, because it is not miscible with water, only the water layer of the mixture was used to measure its pH value changes.

2.6 Corrosion test

2.6.1 Copper strip corrosion test. The copper strip corrosion test of BTAP₄₄₄₄ and the reference samples was studied according to the GB-T5096-1985 (91) procedure. The length, width and thickness of the copper strips used in the experiment are 12.0 mm, 12.0 mm and 3.0 mm, respectively. Copper strips were polished with sandpaper, and the surfaces were cleaned with acetone, and dried in air naturally. They were then immersed into the samples, and heated at 100 °C for 24 h in an oven. At the end of the test, the copper strips were taken out and washed with acetone thoroughly. The photos of the copper strips were taken and the surface morphologies were analyzed by an FEI Quanta 250 scanning electron microscope (SEM).

2.6.2 Accelerated corrosion test. Configuration of the test solution: a certain amount of a sample was dissolved in 25 mL methanol, then 25 mL of saturated calcium hydroxide [Ca(OH)₂] solution was added to form a 0.01 mol L^{−1} solution (sample concentration). The well-polished steel, copper and aluminum discs were soaked into the prepared solutions which were maintained at RT for a certain time (steel discs: 3 months, copper discs: 10 days, and aluminum discs: 5 days). In addition, the corresponding discs soaked in saturated [Ca(OH)₂] solution were used as blank comparisons under the same conditions.

2.7 Friction and wear test

The tribological properties of BTAP₄₄₄₄ and the reference samples on different friction pairs were evaluated by an Opti-mol SRV-V oscillating reciprocating friction and wear tester through a reciprocating ball-on-disk configuration test style. The tests were performed both at RT and 100 °C. Before the tests, a drop of a sample was added onto the ball-disc contact area. A AISI 52100 bearing steel ball (diameter = 10 mm, hardness = 59–63 HRC, and mean roughness = 100 nm) was



made to slide on a lower stationary disk (ϕ 24–7.9 mm) with load of 100 N, oscillation frequency of 25 Hz, sliding amplitude of 1 mm and duration of 30 min. During the tests, environmental humidity was maintained between 30–50%. The lower stationary steel, copper and aluminum discs were made of AISI 52100 bearing steel (hardness = 59–63 HRC), ZQSn 663 copper (hardness = 140–160 HV) and 2024 aluminum alloy (hardness = 140–170 HV), respectively. Before the friction and wear tests, the steel, copper, and aluminum disks were polished in turn with 400, 800 and 1500 Cw SiC sandpaper until the mean roughness of the surfaces were about 20–50, 100–200, and 100–200 nm, respectively. The homologous friction curves were recorded automatically by the SRV test rig. After the tests, losses of the wear volumes of different lower discs were measured by using a non-contact surface mapping profiler (BRUKER-NPFLEX). Three times of repeating measurements were carried for each sample, and average values are reported. The element composition and morphology of the worn surfaces were analyzed by using X-ray photoelectron spectroscopy (XPS) which was conducted on an AXIS SUPRA X-ray photoelectric spectrometer manufactured by UK Kratos Company and monochromatic anode Al target (Al K α 1486.6 eV) was used as the exciting source. The test condition is: voltage = 15 kV, power = 225 W, exit angle 0° and resolution = ± 0.3 eV. Before scanning, the worn surfaces were ultrasonic cleaned three times with ethanol. The binding energy of contaminated carbon (C 1s: 284.8 eV) was used as the reference.

3. Results and discussion

3.1 Structure characterization and thermal stability

The chemical structures, names and codes of BTAP₄₄₄₄, bmimBF₄, and PAO 10 are given in Table 1. The structure and purity of BTAP₄₄₄₄ was finely confirmed by ¹H NMR, ¹³C NMR, FT-IR and HRMS spectroscopic data. The detail data are presented in the ESI.†

Table 1 The chemical structures, names and codes of the used samples

Samples (code)	Chemical structural
Poly-alpha-olefin (PAO 10)	
1-Butyl-3-methylimidazolium tetrafluoroborate (bmimBF ₄)	
Benzotriazole terabutylphosphonium (BTAP ₄₄₄₄)	

The thermal stabilities of the samples were probed over the temperature range of RT to 600 °C. Fig. 1 shows the thermogravimetry (TG) curves of all the samples. It is clearly seen that BTAP₄₄₄₄ has no rapid weight loss below 300 °C and the temperature for its total decomposition is higher than 400 °C. The onset decomposition temperature of BTAP₄₄₄₄ is found to be lower than bmimBF₄ but far higher than that of PAO 10.

For example, about 85% of PAO 10 was decomposed at 350 °C, but only 20% of BTAP₄₄₄₄ was decomposed at this temperature. Especially, as compared with raw material BTAH, the thermal decomposition temperature is greatly improved. Neat 1H-benzotriazole has very low thermal decomposition temperature (about 150 °C). So, in terms of thermal stability, BTAP₄₄₄₄ has a great advantage to be used as a lubricant.

3.2 Viscosity

The kinematic viscosities of all the samples were measured at 40 °C and 100 °C, respectively. According to the test results, viscosity indexes were calculated to evaluate the viscosity-temperature characteristics of the samples (Table 2). It is clearly seen that the viscosity of BTAP₄₄₄₄ is far higher than the comparing samples both at 40 °C and 100 °C. But its viscosity index is relatively low, which is attributed to the molecular structural characteristic of BTAP₄₄₄₄. On the one hand, viscosity index of IL with aromatic ion is lower than that with nonaromatic ion because of the delocalized negative charge on aromatic ion and thus the weaker cation-anion interaction.⁴⁹ On the other hand, viscosity index of phosphonium-based IL with a phosphorus atom in the cationic center are lower than that of ammonium-based IL with a nitrogen atom in the cationic center.^{50,51}

3.3 Hydrolysis stability analysis

To affirm the hydrolysis stability of BTAP₄₄₄₄, it was mixed with a same mass of water and then the solution was stirred at RT. The pH value of the solution was tested periodically to reflect the hydrolysis qualitatively. From Fig. 2, it can be seen that the

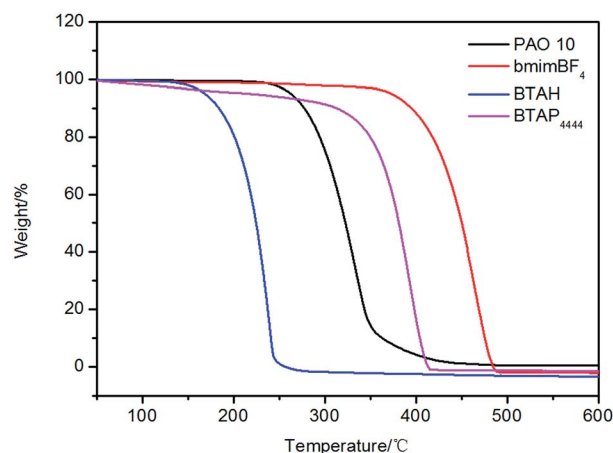


Fig. 1 Thermal decomposition temperatures of PAO 10, bmimBF₄, BTAH, and BTAP₄₄₄₄.



Table 2 Viscosities and viscosity indexes of the samples

Samples	Kinematic viscosity ($\text{mm}^2 \text{s}^{-1}$)		Viscosity index
	40 °C	100 °C	
PAO 10	68.8835	9.4659	123.32
bmimBF ₄	42.9653	7.1201	126.60
BTAP ₄₄₄₄	373.1334	20.7642	54.27

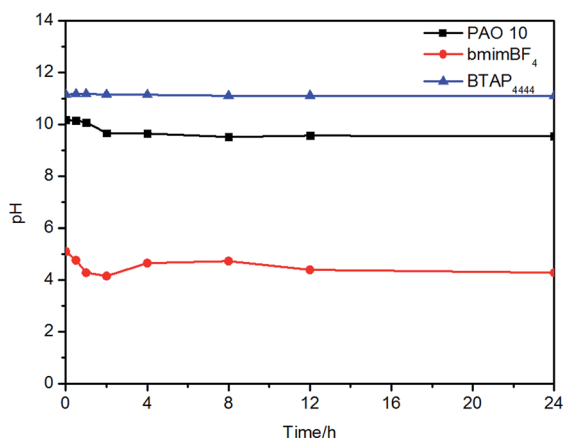


Fig. 2 The pH change curves of the samples.

hydrolytic stability of bmimBF₄ was relatively poor, whose pH decreased greatly at the beginning of the experiment (within the first 2 h).⁴⁵ Compared with bmimBF₄, the pH value of BTAP₄₄₄₄ had almost no change throughout the test. This means that BTAP₄₄₄₄ is stable and not easily hydrolyzed, which has the potential to be used as an efficient lubricant. Under the same condition, the pH value of PAO 10 did not change during the next 24 h. Above of all, these results indicate that BTAP₄₄₄₄ and PAO 10 have good hydrolysis stability to be used lubricants.

3.4 Corrosion test

3.4.1 Copper strip corrosion test. Copper strip corrosion test is adopted to test the corrosion ability of BTAP₄₄₄₄ and the reference samples. BTAP₄₄₄₄ was expected to have no corrosion or far less corrosion than traditional ILs.

The photographs of the adopted copper sheets are shown in Fig. 3. It is found that the copper strip immersed in BTAP₄₄₄₄ did not show any color change and remained bright compared with the fresh copper (Fig. 3). The color of the copper strip

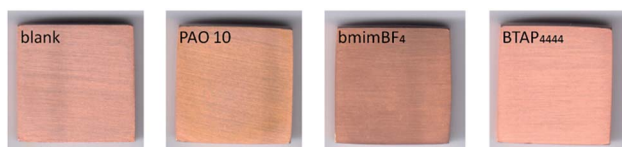


Fig. 3 The photographs of the copper strips adopted during the corrosion tests (blank: a new finished copper strip).

immersed in PAO 10 showed only a slight discoloration, while many black points were found on the copper strip immersed in bmimBF₄. The results indicated that bmimBF₄ contains halogen and is unstable, so hydrolysis occurred during the corrosion test resulting in the generation of corrosion to the copper strip. However, BTAP₄₄₄₄ has halogen-free characteristic, so there is not corrosion phenomenon occurring on the surface. According to the test results, it also can be seen that BTAP₄₄₄₄ may have anticorrosion property and it is worth to verify. In order to verify the anticorrosion function of BTAP₄₄₄₄, accelerated corrosion experiments were conducted.

3.4.2 Accelerated corrosion test. Accelerated corrosion test was carried in order to confirm the anticorrosion property of BTAP₄₄₄₄. The copper discs (would be used in the following tribological tests) were soaked in saturated Ca(OH)₂ solutions containing 0.01 mol L⁻¹ of the tested samples and maintained for 10 days. A copper disc immersed in saturated Ca(OH)₂ solution (without ILs) was used as a blank control experiment. During the test, BTAH was selected as another comparison to confirm the anticorrosion property of BTAP₄₄₄₄. Table 3 shows the pictures of the copper discs after the test. From Table 3, it can be clearly seen that severe corrosion occurred on the copper discs soaked in the blank and bmimBF₄ solutions. Although the copper disc in the solution containing BTAH was corroded very slightly and the copper surface showed little color change, BTAH is a solid and has low thermal stability, so it cannot be used as a lubricant. Table 3 also exhibits the SEM morphologies of the samples. According to the SEM morphologies, it can be seen that the skin layers of the copper discs soaked in the blank and 0.01 mol L⁻¹ bmimBF₄ solutions have been eroded away, and the surface structure of the discs changed obviously. By contrast, slight change occurred on the copper disc surfaces that were soaked in the solutions containing BTAH and BTAP₄₄₄₄. The texture structure of the surface which came from mechanical polishing before the test was clearly seen. These results undoubtedly confirm the anticorrosion property of BTAH and BTAP₄₄₄₄ to copper material.

The materials of steel and aluminum have been widely used in industrial production, and in the following experiments they will also be chosen to test the tribological properties of BTAP₄₄₄₄. So steel and aluminum discs were also selected to do the accelerated corrosion tests. The corrosion of steel is much slower than that of copper, so the test need a longer time and would be lasted for 3 months. However, due to the soft texture of aluminum, the accelerated corrosion test need a shorter time and would be lasted only for 5 days. The results of the tests are shown in Tables 4 and 5. From Table 4, it can be clearly seen that severe corrosion occurred on the steel discs soaked in the blank and bmimBF₄ solutions and the surface colors of the steel discs become dark black. The solutions also changed from colorless to deep brown. However, under the same condition, the steel discs immersed in BTAH and BTAP₄₄₄₄ solutions show no color change, and the texture structure of the surfaces which came from mechanical polishing before the test can be clearly seen. Also, no color change was found in the solutions. It means that BTAH and BTAP₄₄₄₄ have obvious anticorrosion characteristics for steel discs. The



Table 3 The photographs and SEM images of the copper discs used in the accelerated corrosion tests




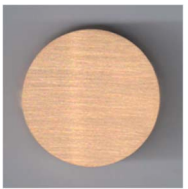

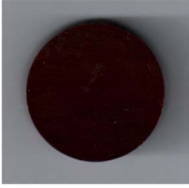

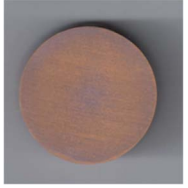
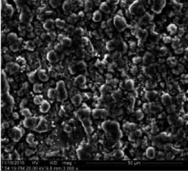

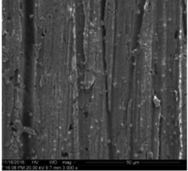
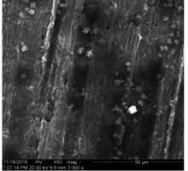
	Blank	0.01 mol L ⁻¹ bmimBF ₄	0.01 mol L ⁻¹ BTAH	0.01 mol L ⁻¹ BTAP ₄₄₄₄
Copper disc (before test)				
Copper disc (after test)				
SEM morphology (mag 3000×)				

Table 4 The photographs and SEM images of the steel discs used in the accelerated corrosion tests





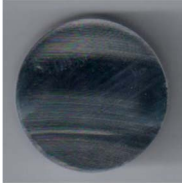
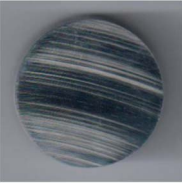
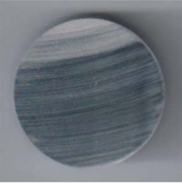



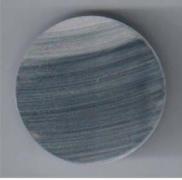
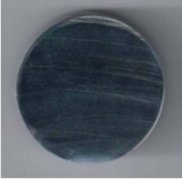
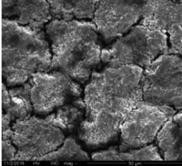
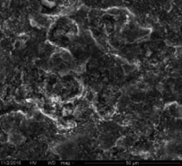
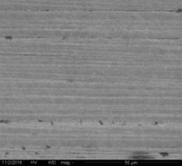
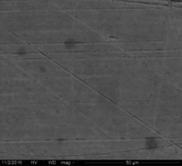
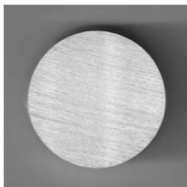
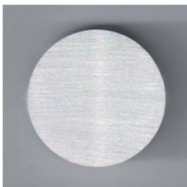
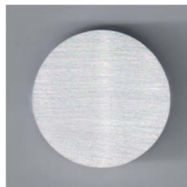
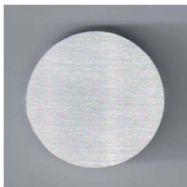


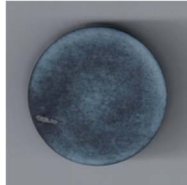

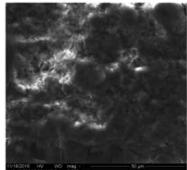
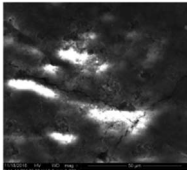
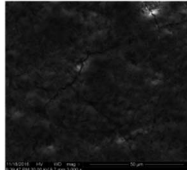
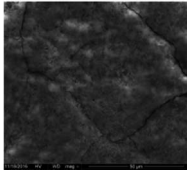
	Blank	0.01 mol L ⁻¹ bmimBF ₄	0.01 mol L ⁻¹ BTAH	0.01 mol L ⁻¹ BTAP ₄₄₄₄
Solution colour				
Steel disc (before test)				
Steel disc (after test)				
SEM morphology (mag 3000×)				

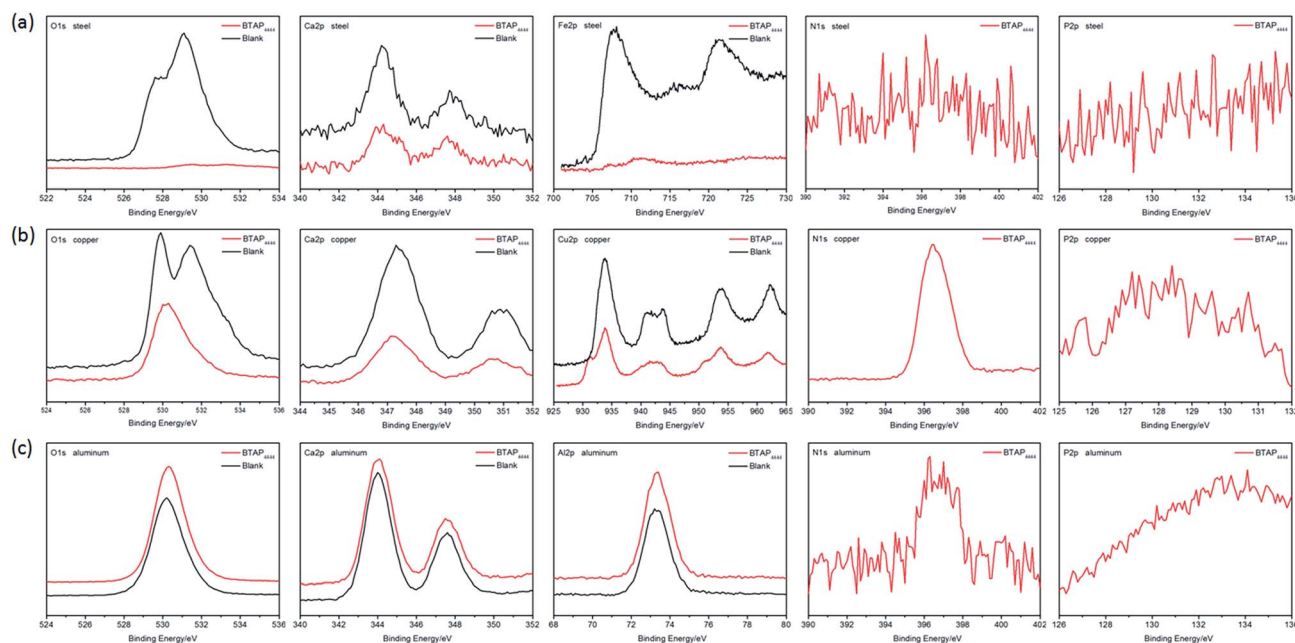


Table 5 The photographs and SEM images of the aluminum discs used in the accelerated corrosion tests

	Blank	0.01 mol L ⁻¹ bmimBF ₄	0.01 mol L ⁻¹ BTAH	0.01 mol L ⁻¹ BTAP ₄₄₄₄
Aluminium disc (before test)				
Aluminium disc (after test)				
SEM morphology (mag 3000×)				

morphologies of the aluminum discs in Table 5 also confirmed these characteristics. From Table 5, it can be seen that the aluminum disc surface soaked in the blank solution is completely broken, and there are some clear pits and fault exist. The aluminum disc soaked in bmimBF₄ solution is significantly blacker than those in BTAH and BTAP₄₄₄₄ solutions. The SEM morphology show that the surface has been seriously corroded, and the corrosion sediment attached to the

surface resulting in the surface height changing, so the brightness of the SEM morphologies is not consistent. However, far light corrosion was found on the discs surface soaked in BTAH and BTAP₄₄₄₄ solutions and protective films were formed on the surfaces. According to the above experimental data, it can be concluded that BTAP₄₄₄₄ has obvious anticorrosion capability for steel, copper and aluminum friction pairs when it was used as a lubricant.

**Fig. 4** The XPS spectra of the surfaces after accelerated corrosion tests (a: steel, b: copper, c: aluminum).

3.4.3 Mechanism analysis of anti-corrosion. In order to analyze the anti-corrosion mechanism of BTAP₄₄₄₄, XPS was used to analyze the surface composition of the corroded discs. The XPS spectra are shown in Fig. 4. From Fig. 4a and b, it can be clearly seen that no obvious N 1s and P 2p peaks were found on the corroded steel and aluminum discs. Only the peaks of O 1s, Ca 2p, and Fe 2p could be found on these disc surfaces and their intensities were weaker than those of the corresponding peaks appeared on the blank discs. So, during the corrosion process, there is no obvious chemical reactions other than oxidation occurred on the steel and aluminium discs. Therefore, it can be speculated that the main reason of the anti-corrosion property of BTAP₄₄₄₄ is the physical adsorption film formation between BTA[−] and the metal surface, which had been washed away by ultrasonic cleaning before the XPS test. Only sediment can be found on the steel and aluminium surfaces, such as CaO, CaCO₃, Fe₃O₄ and FeOOH.

However, obvious N 1s peaks at 396.38 eV and 396.50 eV were found on the corroded copper surface (Fig. 4c). Combining with the Cu 2p peak at 932.6 eV, surface protective films composed of copper complex compounds were formed on the surface. Other than these complex compounds, CuO, Cu₂O, Cu(OH)₂, CaO and CaCO₃ may be formed on the surface deducing from the characteristic peaks of O 1s, Cu 2p and Ca 2p. So there were not only physical adsorption but also chemical reaction happened on the copper surface and further formed protective films to achieve the purpose of anti-corrosion. From the above results, it can be concluded that strong physical and chemical interaction between BTA[−] and the surface of Cu alloy was proposed to account for the excellent anti-wear and anti-corrosion capability of BTAP₄₄₄₄.

3.5 Friction and wear test

The tribological properties of BTAP₄₄₄₄ and the reference samples PAO 10 and bmimBF₄ were studied on the above three friction pairs (steel/steel, steel/copper and steel/aluminum), which were frequently used in mechanical equipments. Fig. 5 shows the evolution of friction coefficients and wear volume losses of the sliding discs after the friction and wear test at RT. As shown in Fig. 5a, it can be seen that the friction coefficients of BTAP₄₄₄₄ is much lower than that of PAO 10 and a little bit higher than that of bmimBF₄ on the friction pair of steel/steel. The trend of the curve is very steady and the running-in time is short enough to be ignored. From Fig. 5b, it can be seen that the wear volume of BTAP₄₄₄₄ is also far smaller than that of PAO 10 and similar to that of bmimBF₄. These results confirm the good friction reducing and anti-wear properties of BTAP₄₄₄₄ for steel/steel friction pair.

Under the same conditions, the tribological properties of all the samples were tested on copper and aluminum friction pairs. It was found that BTAP₄₄₄₄ has good lubrication performance for steel/copper and steel/aluminum friction pairs, whether in terms of friction coefficient or wear volume (Fig. 6 and 7). Comparatively speaking, PAO 10 and bmimBF₄ were found to have longer running-in times and caused serious tribocorrosion during the sliding process, especially using bmimBF₄ as

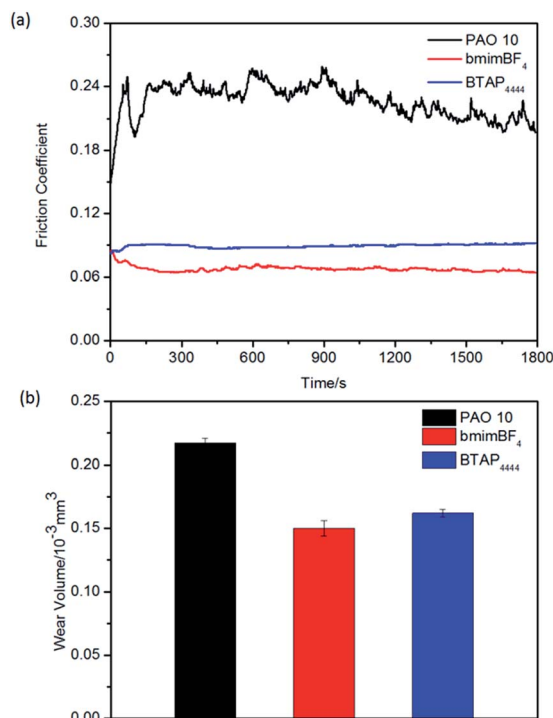


Fig. 5 Evolution of the friction coefficient (a) and wear volume losses (b) for different samples on steel/steel friction pairs at RT.

a lubricant for aluminum friction pair. On contrary, BTAP₄₄₄₄ has resistance to corrosion, and its friction coefficient and wear volume are far lower than those of others (Fig. 7).

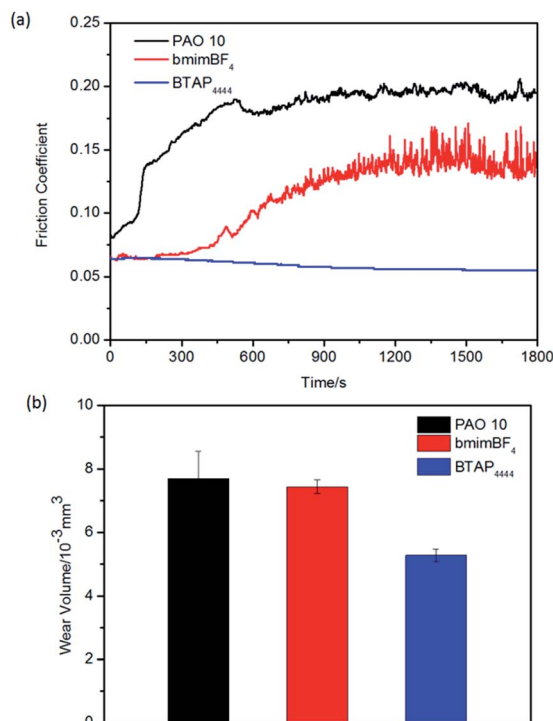


Fig. 6 Evolution of the friction coefficient (a) and wear volume losses (b) for samples on copper/steel friction pairs at RT.



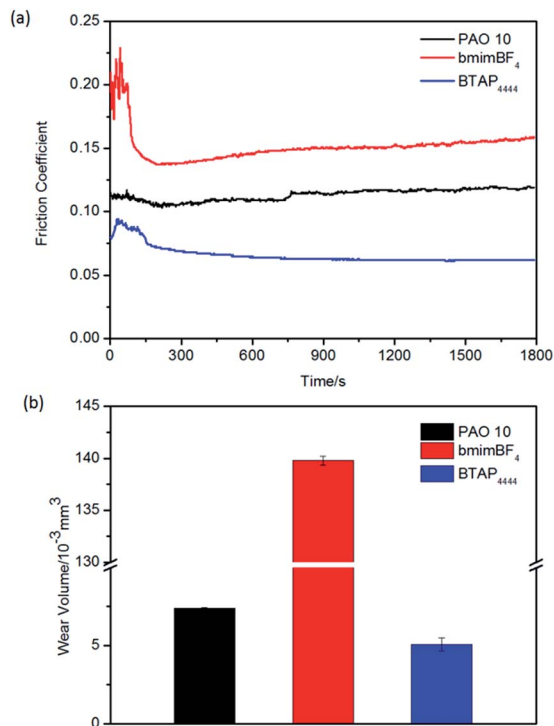


Fig. 7 Evolution of the friction coefficient (a) and wear volume losses (b) for different samples on aluminum/steel friction pairs at RT.

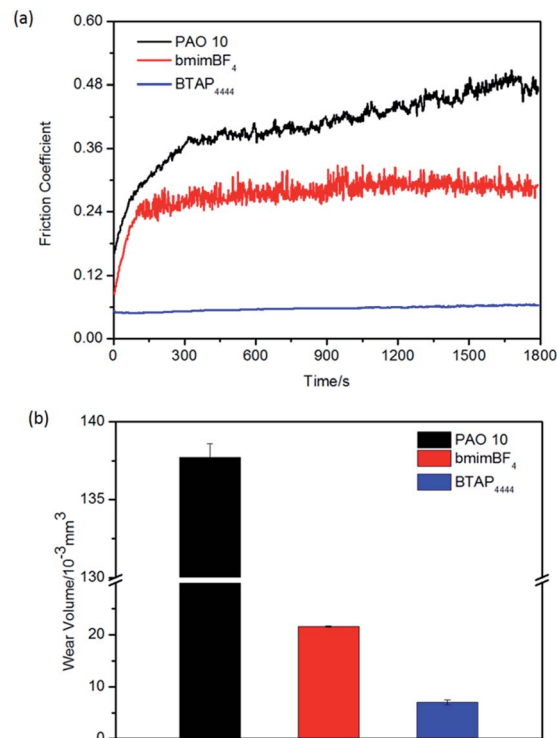


Fig. 9 Evolution of the friction coefficient (a) and wear volume losses (b) for different samples on copper/steel friction pairs at 100 °C.

At the same time, we also test the tribological properties of all the samples under high temperature conditions on the above three friction pairs. The results are shown in Fig. 8–10. From the

figures, it can be seen that BTAP₄₄₄₄ also has good lubrication property, especially on copper friction pairs. Both of the two comparisons have unstable and much larger friction

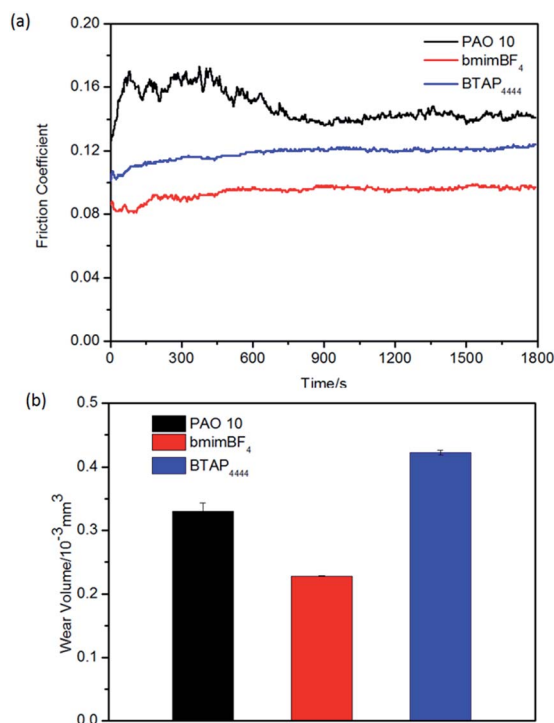


Fig. 8 Evolution of the friction coefficient (a) and wear volume losses (b) for different samples on steel/steel friction pairs at 100 °C.

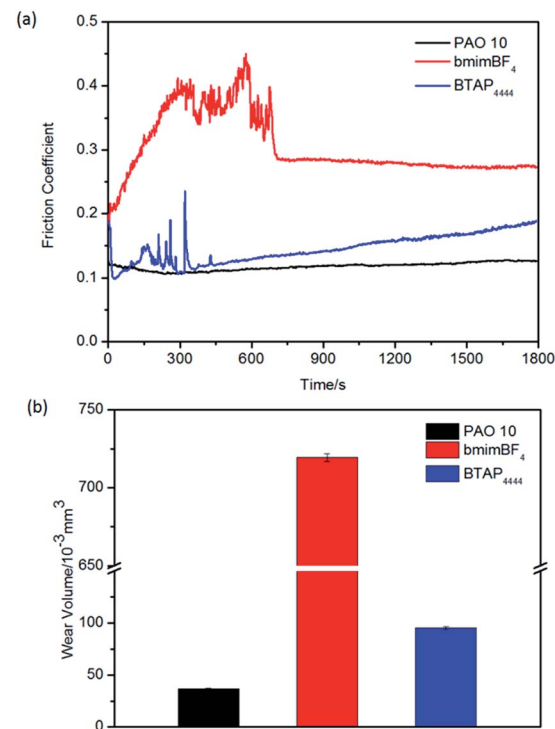


Fig. 10 Evolution of the friction coefficient (a) and wear volume losses (b) for different samples on aluminum/steel friction pairs at 100 °C.



coefficients and wear volumes. As a lubricant for copper friction pair, the friction coefficient of BTAP₄₄₄₄ is much lower and more stable than those of the two comparisons (Fig. 9a). From Fig. 9b, it can be seen that the wear volume of BTAP₄₄₄₄ is also very small, less than 1/15 of PAO 10, and just 1/3 of bmimBF₄. These indicate that BTAP₄₄₄₄ has remarkable wear resistance at high temperature for steel/copper friction pair. It can be concluded that the functional IL BTAP₄₄₄₄ has excellent friction-reducing and anti-wear properties than traditional IL bmimBF₄. So it is more suitable to be used as a lubricant, especially at high temperature conditions.

3.6 Mechanism analysis of friction reducing and anti-wear properties

In order to analyze the lubricating mechanism of BTAP₄₄₄₄, the element composition of the worn surfaces of different friction pairs were performed by using XPS analysis data. The XPS spectra are shown in Fig. 11. It can be seen that weak N 1s and P 2p peaks were found on the spectra of the worn steel surfaces both at RT and 100 °C (Fig. 11a). It reveals that new nitride and phosphorus compound such as FePO₄ and complex organic nitrogen compounds were generated on the worn surfaces during the sliding process. At the same time, combining with

the peaks of O 1s appearing at 530.5 eV and 531.5 eV, other oxides such as Fe₂O₃, Fe₃O₄, FeO, Fe(OH)₃ or Fe(OH)O were formed on the worn steel surfaces. According to these results, it can be inferred that the main reason for good tribological properties of BTAP₄₄₄₄ on steel/steel friction pair is probably attribute to the intrinsic polarity of BTAP₄₄₄₄ molecule to form effective physical adsorption film and further tribochemical reaction film on the sliding surfaces.^{52–54} These physical adsorption and tribochemical reaction films together act as effective protective films to reduce friction and wear on the steel surface.

From the Fig. 11b, it can be seen that obvious N 1s and P 2p peaks were found on the spectra of the worn copper surfaces both at RT and 100 °C. Similarly, combined with the O 1s peaks, CuO, Cu₂O or Cu(OH)₂ and other complex nitrogenous compounds may be formed on the worn copper surfaces. The N 1s peak at 400.4 eV can be identified as the appearing of copper complex compounds.⁵⁵ At the same time, combining with the mechanism of anti-corrosion, it can be concluded that BTA[−] can be easily absorbed on copper surface through strong N–Cu complexation and further react with the metal surface to form chemical adsorption and reaction films.⁵⁶ These protective films provides an effective barrier layer to prevent the metal–metal contact and achieve excellent anti-wear and friction-reducing property.^{57–61}

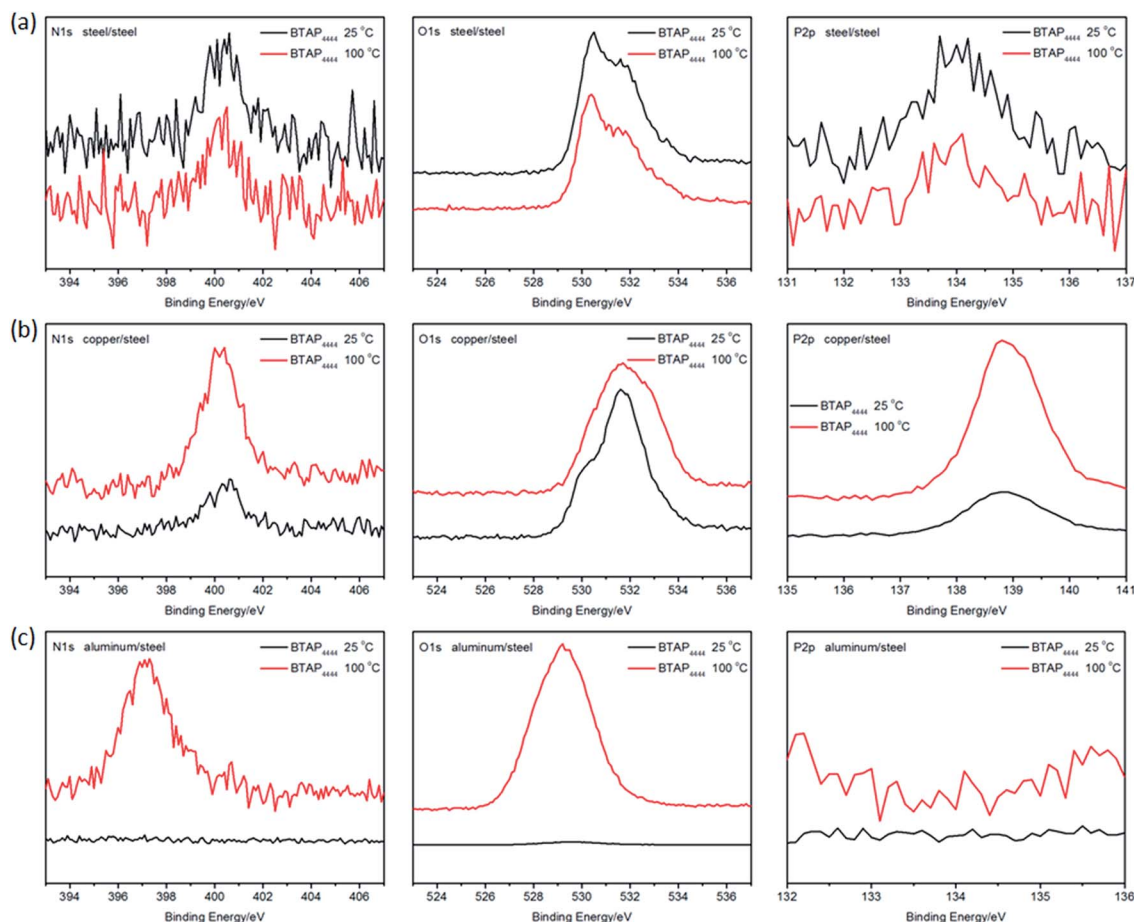


Fig. 11 The XPS spectra of the worn surfaces lubricated by BTAP₄₄₄₄ at RT and 100 °C on different friction pairs (a: steel discs, b: copper discs, c: aluminum discs).



Under the same conditions, however, no obvious N 1s and P 2p peaks could be found on the worn aluminum surfaces lubricated at RT and only weak N 1s peak can be found on the worn aluminum surface lubricated at 100 °C (Fig. 11c). Deducing from the O 1s peaks, there may be Al₂O₃ appearing on the worn aluminum surfaces, but the chemical absorption and reaction of BTAP₄₄₄₄ with the metal surface are very weak, so the anti-wear property of BTAP₄₄₄₄ on aluminum surface is not strong, especially at 100 °C (Fig. 10b).

4. Conclusions

In this work, anti-corrosion IL BTAP₄₄₄₄ was synthesized by using readily available 1H-benzotriazole as an anion donor to react with terabutylphosphonium hydroxide. The physical and chemical properties of the IL were systematically studied and its tribological performance was tested on steel/steel, copper/steel, aluminum/steel friction pairs both at RT and 100 °C. At the same time, its corrosion resistance was measured using both copper strip corrosion test and accelerated corrosion test. The results show that BTAP₄₄₄₄ has high thermal and hydrolytic stability, good tribological properties and high corrosion resistance to different materials. It is a new kind of anti-corrosion IL lubricant, which has good friction reducing and anti-wear properties.

Acknowledgements

We really thank the financial support from the National Natural Science Fund (51675006), an open fund of the State Key Laboratory of Solid Lubrication, Lanzhou Institute of Chemical Physics (LSL-1408), project of Science and Technology Department of Shaanxi Province (2016JZ017) and the local servicing research project of the Education Department of Shaanxi Province (15JF007). The authors would like to express our gratitude to the ILs provided by Professor Deng's group in Lanzhou Institute of Chemical Physics Chinese Academy of Sciences.

Notes and references

- 1 M. Rasberger, *Chemistry and Technology of Lubricants*, Springer, Germany, 1997.
- 2 C. F. Ye, W. M. Liu, Y. X. Chen and L. G. Yu, *Chem. Commun.*, 2001, **21**, 2244–2245.
- 3 F. Zhou, Y. M. Liang and W. M. Liu, *Chem. Soc. Rev.*, 2009, **38**, 2590–2599.
- 4 P. S. Bakshi, R. Gusain, M. Dhawaria, S. K. Suman and O. P. Khatri, *RSC Adv.*, 2016, **6**, 46567–46572.
- 5 N. V. Plechkova and K. R. Seddon, *Chem. Soc. Rev.*, 2008, **37**, 123–150.
- 6 T. Welton, *Chem. Rev.*, 1999, **99**, 2071–2083.
- 7 C. G. Venier and E. W. Casserly, *Lubr. Eng.*, 1991, **47**, 586–591.
- 8 M. J. Dube, D. Bollea, W. R. Jones, M. Marchetti and M. J. Jansen, *Tribol. Lett.*, 2003, **15**, 3–8.
- 9 J. Q. Ma, C. J. Pang, Y. F. Mo and M. W. Bai, *Wear*, 2007, **263**, 1000–1007.
- 10 Y. Wang, L. P. Wang, Y. F. Mo and Q. J. Xue, *Tribol. Lett.*, 2011, **41**, 163–170.
- 11 A. E. Jiménez, M. D. Bermúdez, F. J. Carrión and G. Martínez-Nicolás, *Wear*, 2006, **261**, 347–359.
- 12 M. R. Cai, Z. Zhao, R. M. Liang, F. Zhou and W. M. Liu, *Tribol. Lett.*, 2010, **40**, 215–224.
- 13 B. S. Phillips and J. S. Zabinski, *Tribol. Lett.*, 2004, **17**, 533–541.
- 14 W. M. Liu, C. F. Ye, Q. Y. Gong, H. Z. Wang and P. Wang, *Tribol. Lett.*, 2002, **13**, 81–85.
- 15 Y. Q. Xia, S. J. Wang, F. Zhou, H. Z. Wang, Y. M. Lin and T. Xu, *Tribol. Int.*, 2006, **39**, 635–640.
- 16 A. E. Jiménez, M. D. Bermúdez and P. Iglesias, *Tribol. Int.*, 2009, **42**, 1744–1751.
- 17 R. Gusain and O. P. Khatri, *RSC Adv.*, 2015, **5**, 25287–25294.
- 18 A. Kokalj, S. Peljhan, M. Finšgar and I. Milošev, *J. Am. Chem. Soc.*, 2010, **132**, 16657–16668.
- 19 Y. Ma, F. Han, Z. Li and C. G. Xia, *ACS Sustainable Chem. Eng.*, 2016, **4**, 633–639.
- 20 Z. H. Song, Y. M. Liang, M. J. Fan, F. Zhou and W. M. Liu, *RSC Adv.*, 2014, **4**, 19396–19402.
- 21 Y. Kondo, T. Koyama, R. Tsuboi, M. Nakano, K. Miyake and S. Sasaki, *Tribol. Lett.*, 2013, **51**, 243–249.
- 22 F. U. Shah, S. Glavatskih, D. R. MacFarlane, A. Somers, M. Forsyth and O. N. Antzutkin, *Phys. Chem. Chem. Phys.*, 2011, **13**, 12865–12873.
- 23 V. Totolin, I. Minami, C. Gabler and N. Dorr, *Tribol. Int.*, 2013, **67**, 191–198.
- 24 R. Gusain, R. Singh, K. L. N. Sivakumar and O. P. Khatri, *RSC Adv.*, 2014, **4**, 1293–1301.
- 25 R. Gusain, P. Gupta, S. Saran and O. P. Khatri, *ACS Appl. Mater. Interfaces*, 2014, **6**, 15318–15328.
- 26 R. Gusain, P. Gupta, S. Saran and O. P. Khatri, *Ind. Eng. Chem. Res.*, 2016, **55**, 856–865.
- 27 Procter and L. Gamble, *Br. Pat.*, 652339, 1947.
- 28 R. Walker, *J. Chem. Educ.*, 1980, **57**, 789–791.
- 29 H. Brinch Madsen, *Stud. Conserv.*, 1967, **12**, 163–167.
- 30 H. Brinch Madsen, *Stud. Conserv.*, 1971, **16**, 120–122.
- 31 J. J. Kim, S. K. Kim and J. U. Bae, *Thin Solid Films*, 2002, **415**, 101–107.
- 32 F. E. T. Heakal and S. Haruyama, *Corros. Sci.*, 1980, **20**, 887–898.
- 33 A. Ababneh, M. Sheban and M. Abu-Dalo, *J. Mater. Civ. Eng.*, 2012, **24**, 141–151.
- 34 D. Gopi, K. M. Govindaraju, V. Collins Arun Prakash, D. M. Angeline Sakila and L. Kavitha, *Corros. Sci.*, 2009, **51**, 2259–2265.
- 35 D. Gopi, K. M. Govindaraju, V. Collins Arun Prakash, V. Manivannan and L. Kavitha, *J. Appl. Electrochem.*, 2009, **39**, 269–276.
- 36 A. D. Modestov, G. D. Zhou, Y. P. Wu, T. Notoya and D. P. Schweinsberg, *Corros. Sci.*, 1994, **36**, 1931–1946.
- 37 C. Verma, P. Singh, I. Bahadur, E. E. Ebenso and M. A. Quraishi, *J. Mol. Liq.*, 2015, **209**, 767–778.
- 38 N. Soltani, M. Behpour, E. E. Oguzie, M. Mahluji and M. A. Ghasemzadeh, *RSC Adv.*, 2015, **5**, 11145–11162.



- 39 C. Verma, L. Olasunkanmi, I. B. Obot and M. A. Quraishi, *RSC Adv.*, 2016, **6**, 1–52.
- 40 M. M. Antonijevic and M. B. Petrovic, *Int. J. Electrochem. Sci.*, 2008, **3**, 1–28.
- 41 M. Finšgar and I. Milošev, *Corros. Sci.*, 2010, **52**, 2737–2749.
- 42 X. Liu, F. Zhou, Y. M. Liang and W. M. Liu, *Tribol. Lett.*, 2006, **23**, 191–196.
- 43 M. R. Cai, R. M. Liang, F. Zhou and W. M. Liu, *ACS Appl. Mater. Interfaces*, 2011, **3**, 4580–4592.
- 44 Z. H. Song, M. R. Cai, Y. M. Liang, M. J. Fan, F. Zhou and W. M. Liu, *RSC Adv.*, 2013, **3**, 21715–21721.
- 45 M. J. Fan, D. S. Yang, X. L. Wang, W. M. Liu and H. Z. Fu, *Ind. Eng. Chem. Res.*, 2014, **53**, 17952–17960.
- 46 M. J. Fan, X. L. Wang, D. S. Yang, D. M. Wang, Y. Y. Yan, c. Y. Zhang and X. C. Liu, *Tribol. Int.*, 2015, **92**, 344–352.
- 47 R. R. Sahoo and S. K. Biswas, *J. Colloid Interface Sci.*, 2009, **333**, 707–718.
- 48 C. J. Reeves, P. L. Menezes, T. C. Jen and M. R. Lovell, *Tribol. Int.*, 2015, **90**, 123–134.
- 49 G. R. Yu, D. C. Zhao, L. Wen, S. D. Yang and X. C. Chen, *AIChE J.*, 2012, **58**, 2885–2899.
- 50 S. Seki, T. Kobayashi, Y. Kobayashi, K. Takei, H. Miyashiro, K. Hayamizu, S. Tsuzuki, T. Mitsugi and Y. Umebayashi, *J. Mol. Liq.*, 2010, **152**, 9–13.
- 51 K. Tsunashima and M. Sugiya, *Electrochem. Commun.*, 2007, **9**, 2353–2358.
- 52 Z. B. Zhou, H. Matsumoto and K. Tatsumi, *Chem.–Eur. J.*, 2006, **12**, 2196–2212.
- 53 J. Qu, P. J. Blau, S. Dai, H. Luo and H. M. Meyer III, *Tribol. Lett.*, 2009, **35**, 181–189.
- 54 R. T. Haasch, T. Y. Lee, D. Gall, J. E. Greene and I. Petrov, *Surf. Sci. Spectra*, 2000, **7**, 204–212.
- 55 <http://srdata.nist.gov/xps/>.
- 56 Y. Jiang and J. B. Adams, *Surf. Sci.*, 2003, **529**, 428–442.
- 57 C. Verma, L. O. Olasunkanmi, I. B. Obot, E. E. Ebenso and M. A. Quraishi, *RSC Adv.*, 2016, **6**, 15639–15654.
- 58 D. G. Shchukin, S. V. Lamaka, K. A. Yasakau, M. L. Zheludkevich, M. G. S. Ferreira and H. Möhwald, *J. Phys. Chem. C*, 2008, **112**, 958–964.
- 59 V. Palanivel, Y. Huang and W. J. van Ooij, *Prog. Org. Coat.*, 2005, **53**, 153–168.
- 60 D. Raps, T. Hack, J. Wehr, M. L. Zheludkevich, A. C. Bastos, M. G. S. Ferreira and O. Nuyken, *Corros. Sci.*, 2009, **51**, 1012–1121.
- 61 M. Garcia-Heras, A. Jimenez-Morales, B. Casal, J. C. Galvan, S. Radzki and M. A. Villegas, *J. Alloy. Compd.*, 2004, **380**, 219–224.

