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Self-assembling characteristics of amphiphilic zwitterionic brush random copolymers at the air–water interface

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Self-assembling characteristics of a series of amphiphilic zwitterionic brush random copolymers, poly(oxy(11-(3-sulfonylpropyltrimethylglycyl)undecylesterthiomethyl)ethylene-co-oxy(*n*-dodecylthiomethyl)ethylene) (PECH-DMAPS_{*m*}, where *m* is the mol% of DMAPS end group), were investigated at an air–water interface by using surface pressure–area isotherms, infrared spectroscopy, and X-ray reflectivity analysis. Interestingly the random polymers (*m*: 20–60 mol%) always formed Langmuir monolayer structures only rather than any other structures, regardless of the surface pressures. The Langmuir monolayers possessed enhanced lateral ordering together with conformational changes of the backbone and bristles through increasing the surface pressure. The monolayer structures were basically composed of a hydrophobic bristle phase in the air side and a hydrophilic backbone and bristle phase in the water side. The unique, highly ordered Langmuir monolayer structures could be realized by positive, cooperative efforts of several factors such as the compositional balance of hydrophobic and hydrophilic zwitterionic bristles, the air- and water-induced segregations of the bristles and backbone, the lateral ordering capabilities of the alkyl groups and the alkylenyl linkers in the bristles under the assistance of surface pressure, and the relatively strong anchoring power of DMAPS end groups in the water side using their high water affinity due to the hydrophilic zwitterionic nature. Overall, PECH-DMAPS_{*m*} revealed a very unique self-assembling behavior at the air–water interface, always producing the Langmuir monolayer structure.

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Introduction

Biomimicking polymers have gained great attention from academia and industry because of their potential applications as biomaterials in various fields of biomedical science, tissue culture, pharmaceuticals, and drug delivery.^{1,2} Therefore, much effort has been given to develop biomimicking polymers; as a result, a number of biomimicking polymers were reported.^{1,3–15} They can be classified into two major families: biomolecule-containing polymer brushes or self-assembled monolayers (SAMs) tethered to substrate surfaces^{1,3–7} and biomolecule-containing graft, comb-like, random and block

copolymers.^{13–15} However, the first family has some serious drawbacks, as follows. Much care is needed for the selection of substrates in the aspect of biocompatibility and reliability. The substrates employed require surface functionalization or activation before tethering of polymers or surface-induced polymerization and SAM formation. The tethering approach of polymers to substrate surfaces is limited by steric hindrance, causing incomplete surface coverage. The SAM approach possesses structural defects due to the non-uniformity and the chemical instability. Therefore, the first family of materials is generally restricted to *in vitro* applications. In the case of the second family, there are also some disadvantages. A number of reaction steps are generally necessary to synthesize biomolecule-containing polymers. Thus, the development of synthetic schemes with high yields is required. And, it is not assured that the biomolecule groups are exposed to the surface, because detailed structural information has not been given. Their structures and functionalities were mainly investigated in bulk state or fiber form. In applications they are highly demanded to use as a thin film rather than a bulk state or fiber form. Moreover, they very often elicit inflammatory responses, including foreign body responses and fibrous encapsulation,^{1,2} indicating that their biocompatibility is needed to improve

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more. Therefore, the development of polymeric biomaterials with high performances still remains in the exploration stage.

Recently we have developed a new strategy to develop high performance polymeric biomaterials with significantly enhanced biocompatibility. Our strategy stands on the molecular design and synthesis of chemically well-defined brush (*i.e.*, comb-like) polymers bearing proper biomolecules at the bristle ends which can enable to favorably self-assemble as an organized structure, providing biomolecule-enriched surface. We have successfully demonstrated this strategy and, indeed, newly introduced several polymeric biomaterial systems.^{16–27} Their thin film morphologies were well investigated, confirming that they self-assembled favorably as horizontally-oriented multibilayer structures with fully extended chain conformation, always providing biomolecule-rich film surface. In fact, these biomaterials can be used in aqueous solutions including buffers or in contact with water or buffers, in addition to the film and bulk state. Thus, it is necessary to investigate their self-assembly behaviors at an air–water interface or in water and buffer solutions.

In this study, we investigated Langmuir film formation characteristics of amphiphilic zwitterionic brush random copolymers, poly(oxy(11-(3-sulfonylpropyltrimethylglycyl)-undecylesterthiomethyl)ethylene-co-oxy(*n*-dodecylthiomethyl)ethylene)s (PECH-DMAPS_{*m*}, where *m* is the mol% of DMAPS (*i.e.*, sulfobetaine) end group) at an air–water interface using surface pressure–area isotherm, infrared spectroscopy, and X-ray reflectivity analysis. PECH-DMAPS₁₀₀ was found to dissolve in water, failed to form Langmuir film. In contrast, the other brush random polymers (*m*: 20–60 mol%) nicely demonstrated to always form only Langmuir monolayer structure, regardless of the surface pressures. The Langmuir monolayer films were further improved their structural ordering by increasing the surface pressure. These interesting monolayer formation behaviors could be understood with taking into account all possible chemical, physical, and thermodynamic factors.

Experimental

Materials

A series of well-defined PECH-DMAPS_{*m*} polymers were prepared, according to the synthetic method described in the literature.²³ Poly(epichlorohydrin) (PECH) was synthesized as a base polymer from epichlorohydrin by cationic ring-opening polymerization. The obtained PECH polymer was determined to have a weight-average molecular weight \bar{M}_w of 38 500 and \bar{M}_n of 22 700 by using gel permeation chromatography calibrated with polystyrene standards. A portion of the PECH polymer was reacted with sodium 11-hydroxyundecylthiolate and further reacted with 1,3-propanesultone, producing PECH-DMAPS_{*m*} (Fig. 1). The PECH-DMAPS_{*m*} polymers were identified by proton and carbon nuclear magnetic resonance spectroscopy analysis. The PECH-DMAPS_{*m*} polymers were thermally stable up to 190–215 °C (= T_d , degradation temperature) in nitrogen atmosphere, depending on the compositions. A glass transition temperature T_g was observed over the temperature range –33 °C to –25 °C, whereas a melting temperature T_m was detected over the

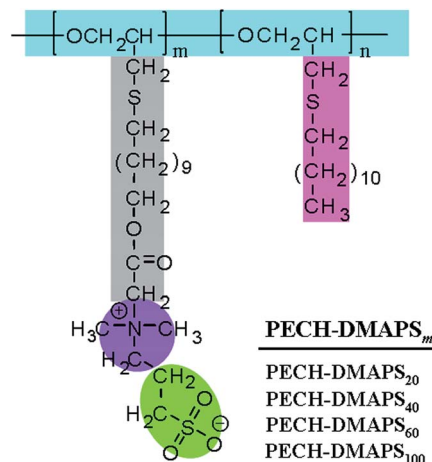


Fig. 1 Chemical structures of the PECH-DMAPS_{*m*} polymers used in this study.

temperature range 0 °C to 30 °C. Higher DMAPS content in the brush polymer lowered T_d and T_g but increased T_m .

Langmuir film formation

Each brush polymer sample was dissolved in chloroform (analytical grade, Sigma-Aldrich, St. Louis, MO, USA), giving polymer solution with a concentration of 0.001 wt%. The obtained polymer solutions were filtered through polytetrafluoroethylene (PTFE) membrane microfilters with a pore size of 0.2 μm. A Langmuir–Blodgett (LB) trough system (model 302M, Nima, Coventry, England) was used for the preparation of Langmuir films at air–water interface. The trough was filled with deionized water and then maintained at room temperature. Each brush polymer solution was spread onto the air–water interface by using a Hamilton microsyringe. The compression was started around 30 min after spreading in order to allow the evaporation of the used chloroform solvent. The surface pressure–area (π -*A*) isotherm was recorded at a compression rate of 10 cm² min^{–1} in a symmetric mode with an initial area of 245 cm².

XR measurement

Taking into account the measured π -*A* isotherms, *in situ* specular XR measurements were carried out at chosen surface pressures during Langmuir film formation process, using a Bruker diffractometer (model D8, Bruker AXS, Germany) with a vertical goniometer, which allows the study of a liquid surface without disturbing a LB trough during measurements.^{28,29} The LB trough was enclosed in a plexiglass box with Kapton windows, and the incoming and reflected beams were detected through the Kapton-covered windows. The LB trough setup was placed on an active vibration table platform (NANO 30, Halcyonics, Goettingen, Germany). An anode X-ray generator was used to produce CuK α radiation ($\lambda = 1.542 \text{ \AA}$), followed by paralleling the incident beam with a Göbel mirror (model GM III, Incoatec, Geesthacht, Germany). Two slits prior to the sample defined the incident beam size and reduced the vertical



divergence. The measured XR data were quantitatively analyzed with a recursive formula given by the Parratt dynamic theory.^{30,31}

IR measurement

Attenuated total reflection Fourier transform infrared (ATR-FTIR) spectroscopy measurements were conducted on LB films that were prepared at various surface pressures. All IR measurements were obtained by recording 256 scans with a resolution of 2 cm^{-1} using FTIR spectrometers (model Cary 640, Agilent Technologies, Santa Clara, CA, USA; model UATR, Perkin-Elmer, Waltham, MA, USA) with a Peltier-cooled DLATGS detector. A ZnSe ATR crystal was purchased from the Pike Technologies (Madison, WI, USA) and used for ATR-FTIR measurements. The ZnSe crystal was dipped in water on the LB trough and then Langmuir film was prepared on water surface. The barrier in the LB trough was moved with a compression rate of $10\text{ cm}^2\text{ min}^{-1}$. When surface pressure was reached at a target pressure, the ZnSe crystal was pulled up with a rate of 2 mm min^{-1} while the barrier was compressed to maintain the target pressure. The LB film deposited onto the ZnSe crystal was loaded onto the ATR-FTIR cell after evaporation of the residual water and then IR spectroscopic measurements were conducted. In addition, the polymers in bulk states were examined by IR spectroscopy.

Results and discussions

π - A isotherms were successfully measured at the air-water interface for PECH-DMAPS₂₀, PECH-DMAPS₄₀, and PECH-DMAPS₆₀. However, π - A isotherm could not be measured for PECH-DMAPS₁₀₀ because of its good solubility in water. Representatives of the measured π - A isotherm curves are shown in Fig. 2. For the PECH-DMAPS₂₀ molecules spread on the water surface, the surface pressure π begins to rise up when the surface area reaches to 9148 \AA^2 per molecule. Then, with decreasing area, the π value initially increases very slowly until around 0.5 mN m^{-1} and changes to a rapidly rising mode. The π value continues such rapid rise, reaching to around 30 mN m^{-1} . Thereafter, the π value turns to increase slowly with decreasing surface area until 33 mN m^{-1} and again changes to a rising

mode, ultimately reaching to 36 mN m^{-1} . A similar mode of π - A isotherm curve is observed for PECH-DMAPS₆₀. However, the π value starts to rise at the surface area of 9647 \AA^2 per molecule, which is slightly larger than that of PECH-DMAPS₂₀. The rising mode of π is changed at 1, 23, and 33 mN m^{-1} . PECH-DMAPS₄₀ shows somewhat different π - A isotherm curve from those of PECH-DMAPS₂₀ and PECH-DMAPS₆₀. The π value begins to rise at the surface area of 9723 \AA^2 per molecule, which is slightly larger than that of PECH-DMAPS₆₀ but much larger than that of PECH-DMAPS₂₀. The rising mode of π is changed at 1 and 28 mN m^{-1} .

With the π - A isotherms above, Langmuir films were sampled out at chosen π values and transformed onto ZnSe prisms for IR spectroscopy analysis. In addition, IR spectroscopy analysis was extended for the brush polymers in solid states. Representatives of the measured IR spectra are shown in Fig. 3. In general, the methyl (CH_3) group of alkyl-containing compounds is known to reveal asymmetric stretching vibration peak around 2955 cm^{-1}

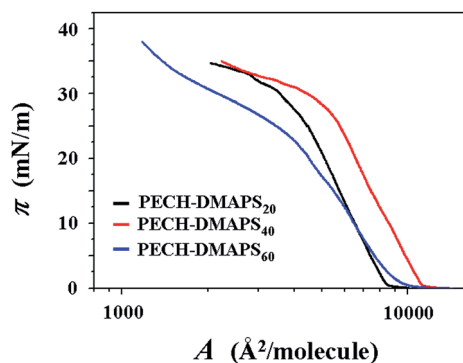


Fig. 2 Representative surface pressure–area (π - A) isotherms of the amphiphilic zwitterionic brush random polymers at the air–water interface.

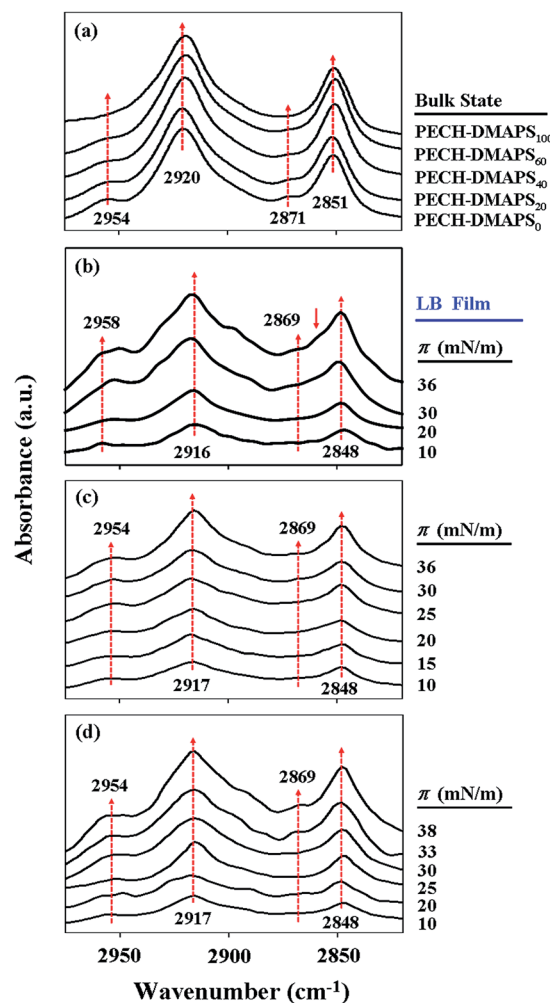


Fig. 3 Representative IR spectra measured for the brush polymers in bulk states as well as the Langmuir–Blodgett (LB) films prepared at various surface pressures: (a) brush polymers in bulk states; (b) PECH-DMAPS₂₀; (c) PECH-DMAPS₄₀; (d) PECH-DMAPS₆₀. The LB films were prepared on ZnSe prisms.



and symmetric vibration peak around 2875 cm^{-1} .^{32–37} And their methylenyl (CH_2) units are also known to exhibit asymmetric stretching vibration peak around 2920 cm^{-1} and symmetric vibration peak around 2850 cm^{-1} .^{32–37} For both cases, the asymmetric vibration peak is stronger in intensity than the symmetric vibration peak. The CH_2 stretching peaks are much stronger in intensity than the CH_3 stretching bands. On the other hand, the methyl (N-CH_3) group of zwitterionic sulfobetaine-containing compounds has been reported to show asymmetric stretching vibration peak around 3020 cm^{-1} and symmetric vibration peak around 2854 cm^{-1} ; in general both the peaks are known to be very weak and broad, perhaps due to their low absorptivities.^{34,35} Taking these facts into account, the characteristic C–H stretching vibration peaks observed for the PECH-DMAPS_{*m*} polymers in Langmuir films or bulk states of this study could be assigned.

As shown in Fig. 3a, PECH-DMAPS₀ in bulk state reveals CH_3 stretching peaks at 2954 and 2871 cm^{-1} and CH_2 stretching peaks at 2920 and 2851 cm^{-1} . PECH-DMAPS₁₀₀ in bulk state also exhibits CH_2 stretching peaks at 2920 and 2851 cm^{-1} , which are very similar to those of PECH-DMAPS₀; but the symmetric vibration peak of the methyl groups (N-CH_3) is not discernible clearly because of its very and broad nature. All brush random copolymers in bulk states show IR spectra which resemble that of PECH-DMAPS₀.

For the PECH-DMAPS₂₀ film formed at $\pi = 10\text{ mN m}^{-1}$, the CH_2 asymmetric and symmetric stretching vibrations are observed in slightly lower frequencies (2916 and 2848 cm^{-1} respectively) (Fig. 3b), suggesting that the bristles are in ordered states. However, the CH_3 asymmetric vibration peak of *n*-dodecylthiomethyl bristles weakly appears in slightly higher frequency (2958 cm^{-1}). Furthermore, the symmetric CH_3 vibration peak could not be discernible. These observations inform that the *n*-dodecylthiomethyl bristles are in a less ordered state rather than in a well-ordered state. The results collectively indicate that in the Langmuir film the bristles are present in a relatively low degree of ordered state.

As the surface pressure π is increased, the CH_2 asymmetric and symmetric stretching frequencies are varied slightly (Fig. 3b). In fact, the CH_2 stretching bands of alkyl compounds are known to vary very little in frequency; for example, they shift only 1 cm^{-1} as the alkyl packing density changes by a factor of 1000.³⁴ And both the hydrophobic and the hydrophilic bristles possess CH_2 units. The CH_2 units in the hydrophobic bristle may be less associated with water molecules, leading to a red shift in the stretching frequencies. This red shift effect may compete with a blue shift caused by the CH_2 units in the hydrophilic bristle that may have more opportunity to associate with water molecules. Due to these characteristics, in this study it is hard to see significant variations in the CH_2 stretching frequencies with increasing surface pressure.

In contrast, the CH_3 asymmetric stretching band shows red shifts as the π value is increased; in addition, the CH_3 symmetric stretching band becomes discernible in the films formed with higher surface pressures. These results indicate that the packing order of the *n*-dodecylthiomethyl bristles is enhanced with increasing surface pressure. Moreover,

a shoulder peak is additionally observed at 2860 cm^{-1} in the films formed with $\pi = 30$ and 36 mN m^{-1} , suggesting that the disordered and ordered DMAPS bristles associated with water molecules coexist in the films.

Overall, these results inform that both the hydrophobic and the hydrophilic bristles of PECH-DMAPS₂₀ tend to pack together laterally under increasing surface pressure, forming ordered states. Such ordered states result from segregations of the hydrophobic and the hydrophilic bristles at the air–water interface which are induced by surface pressure in the film formation process.

Similar trends are observed in the Langmuir film formation processes of PECH-DMAPS₄₀ and PECH-DMAPS₆₀ (Fig. 3c and d). It is noted that for the film formed with $\pi = 10\text{ mN m}^{-1}$ the CH_3 asymmetric vibration peak of *n*-dodecylthiomethyl bristles appears at 2954 cm^{-1} rather than at 2958 cm^{-1} , indicating that the hydrophobic bristles have already formed an ordered state at the air–water interface even under such low surface pressure. The symmetric CH_3 vibration peak is discernible in the films formed with high surface pressures. But, in these cases a shoulder peak at 2860 cm^{-1} could not be discernible clearly because of its weak and broad nature.

Overall, the IR analysis results inform that the hydrophobic *n*-dodecylthiomethyl bristles, as well as the hydrophilic DMAPS bristles undergo lateral packing in the Langmuir film formation process of PECH-DMAPS_{*m*}; their packing orders are enhanced with increasing surface pressure. Such packing orderings are associated with segregations of the hydrophobic and the hydrophilic bristles at the air–water interface under surface pressure.

With the π -*A* isotherms and IR spectroscopy analysis results discussed above, XR measurements were conducted during Langmuir layer formation processes of the PECH-DMAPS_{*m*} polymers. Fig. 4 shows representatives of the measured XR data. To quantitatively analyze the measured XR data, all possible

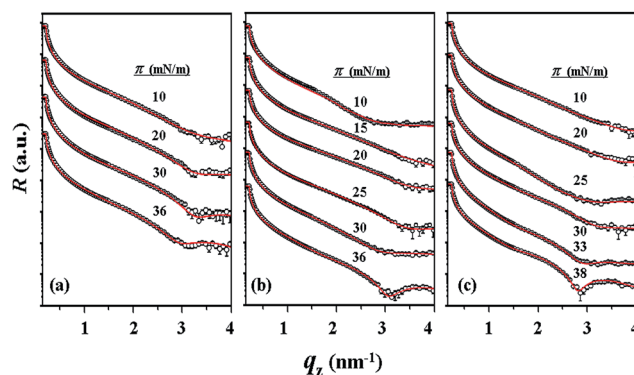


Fig. 4 Representative specular X-ray reflectivity *R* profiles measured for the Langmuir films formed on the air–water interface under various surface pressures: (a) PECH-DMAPS₂₀; (b) PECH-DMAPS₄₀; (c) PECH-DMAPS₆₀. q_z is the magnitude of the scattering vector along the direction of the film thickness, which is defined by $q_z = (4\pi/\lambda)\sin\theta_i$, where λ is the wavelength of the X-ray beam and θ_i is the grazing incidence angle. The symbols are the measured data and the solid lines represent the best fit curves.



structural models were considered and tested. As a result, structural models consisted of two to four sublayers were found to be most suitable for analyzing the XR data, depending on the surface pressure and the chemical composition of the random copolymer. All XR data were successfully analyzed using a recursive formula based on the Parratt dynamic theory.^{30,31} The analysis results are displayed in Fig. 5 and 6. From the analysis results, morphological structures are further proposed for the Langmuir layers formed at various conditions, as shown in Fig. 7–9.

For PECH-DMAPS₂₀, the Langmuir film is formed with a thickness t_f of 2.05 nm and an electron density $\rho_{e,f}$ of 359.4 nm⁻³ under $\pi = 10$ mN m⁻¹. The film consists of three sublayers: the first sublayer ($t_1 = 0.60$ nm and $\rho_{e,1} = 375.9$ nm⁻³), the second sublayer ($t_2 = 0.52$ nm and $\rho_{e,2} = 411.6$ nm⁻³), and the third sublayer ($t_3 = 0.93$ nm and $\rho_{e,3} = 319.4$ nm⁻³). The overall film thickness t_f is slightly larger than the length (1.80 nm) of the fully extended hydrophobic bristle but much smaller than that (2.84 nm) of the fully extended hydrophilic bristle; here the bristle lengths were estimated by performing molecular simulations using the Cerius² software package. And the $\rho_{e,1}$ and $\rho_{e,2}$ values are larger than that (336.2 nm⁻³) of water layer; but the $\rho_{e,3}$ value is lower than that of water layer. These results collectively indicate that a Langmuir monolayer is formed at $\pi = 10$ mN m⁻¹, where the bristles might be tilted

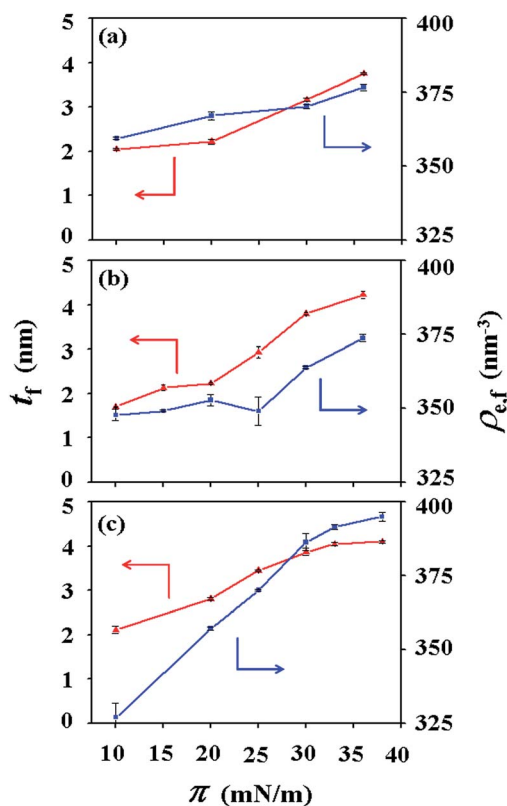


Fig. 5 Thickness t_f and electron density $\rho_{e,f}$ variations of the Langmuir film prepared on the air–water interface as a function of surface pressure: (a) PECH-DMAPS₂₀; (b) PECH-DMAPS₄₀; (c) PECH-DMAPS₆₀.

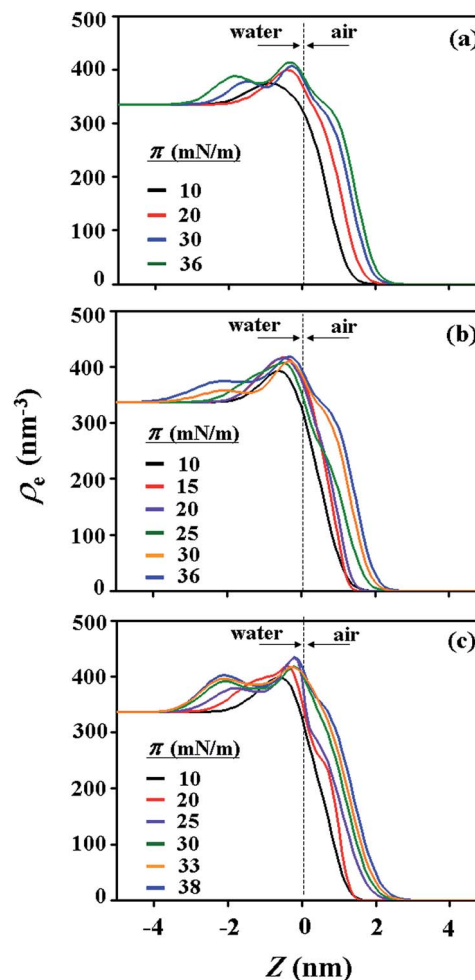


Fig. 6 Electron density ρ_e profiles as a function of the film thickness, for the films prepared at the air–water interface under various surface pressures: (a) PECH-DMAPS₂₀; (b) PECH-DMAPS₄₀; (c) PECH-DMAPS₆₀.

with certain angles with respect to the backbone and/or be in less extended conformation; the first and second sublayers are immersed in the water side while the third sublayer is present in the air side. Considering the polymer chemical structure, the first sublayer is composed mainly with the hydrophilic, zwitterionic DMAPS moieties while the second sublayer consists of the hydrophilic backbones and the linker parts of the bristles. The third sublayer is made of the hydrophobic *n*-dodecylthiomethyl bristles in the air side.

The Langmuir film is found to undergo structural changes including thickening and densification by increasing surface pressure π . The film is slightly thickened and densified at $\pi = 20$ mN m⁻¹. The film is further thickened and densified at $\pi = 30$ mN m⁻¹, consequently causing structural change that the three-sublayer structure is transformed to a four-sublayer structure. The four-sublayer structure continues thickening and densification by increasing π up to 36 mN m⁻¹. The ultimately obtained film is characterized by a set of structural parameters: $t_f = 3.75$ nm ($t_1 = 0.80$ nm, $t_2 = 0.90$ nm, $t_3 = 0.60$ nm, and $t_4 = 1.45$ nm) and $\rho_{e,f} = 376.7$ nm⁻³. The t_f value is



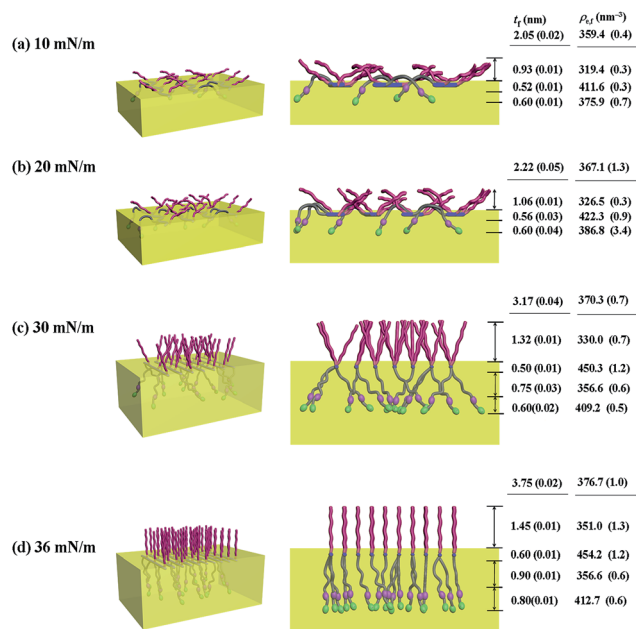


Fig. 7 Molecular models of the Langmuir films of PECH-DMAPS₂₀ prepared at the air–water interface under various surface pressures: (left) view with a tilt angle from the top; (right) front view along the polymer backbone.

somewhat smaller than the sum (4.64 nm) of the lengths of the hydrophobic bristle fully extended upward and the hydrophilic, zwitterionic DMAPS bristle fully extended downward. Moreover, the t_4 value is slightly smaller than the length (1.80 nm) of the fully extended hydrophobic bristle; in fact, the sum of t_4 and a half of t_3 (this part is made of the inner portion of the hydrophobic bristle) is almost same with the length of the fully extended hydrophobic bristle. And the fourth sublayer reveals the lowest ρ_e value, compared to those of the other sublayers. This result is an indication that no or less water molecules are associated with the fourth sublayer. In comparison, the sum of t_1 , t_2 , and a half of t_3 (this part is made of the inner portion of the hydrophilic bristle), which is originated from the hydrophilic bristle, is much smaller than the length (2.84 nm) of the fully extended hydrophilic bristle. These results collectively inform that in the Langmuir film formed at $\pi = 36 \text{ mN m}^{-1}$, the fourth sublayer is composed of the hydrophobic bristles in nearly extended conformation and positioned in the air side while the first and second sublayers including roughly a half of the third sublayer are consisted of the DMAPS bristles in less extended conformation and present in the water side. The lateral packing of such the nearly-extended hydrophobic bristles is further supported by the lateral packing of the hydrophilic polymer backbones being extended under surface pressure as high as 36 mN m^{-1} .

Similar Langmuir monolayer formation behaviors are observed for PECH-DMAPS₄₀ and PECH-DMAPS₆₀. However, the ultimately formed Langmuir films are relatively thicker, compared to the PECH-DMAPS₂₀ film. The films formed with the highest surface pressures are found to have a thickness slightly smaller than the sum of the lengths of the fully extended

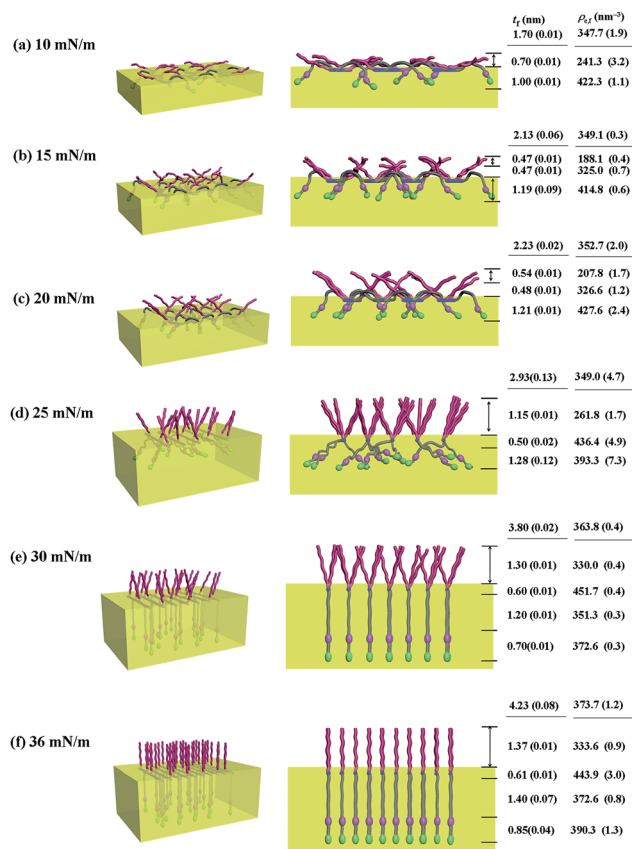


Fig. 8 Molecular models of the Langmuir films of PECH-DMAPS₄₀ prepared at the air–water interface under various surface pressures: (left) view with a tilt angle from the top; (right) front view along the polymer backbone.

hydrophobic bristle upward and the fully extended hydrophilic bristle downward. These results are evidenced that the hydrophobic and hydrophilic bristles, which are segregated into the air and the water side respectively, are in nearly-extended conformations. And the Langmuir films are identified to be composed of four sublayers rather than three sublayers.

The above π -A isotherm, IR spectroscopy and XR analysis results collectively provide the following insights into the correlations of chemical structure, self-assembly and Langmuir film morphology in PECH-DMAPS_m, an amphiphilic, zwitterionic random copolymer system, at the air–water interface.

Firstly, the PECH-DMAPS_m polymer bearing higher mol% of DMAPS tends to cover a larger area of water surface. Such the water surface coverage reaches to a maximum as the DMAPS content approaches to around 40 mol%. These suggest that the hydrophilic, zwitterionic DMAPS bristles, as well as the hydrophilic backbones make important role in the spread of PECH-DMAPS_m at the water surface *via* their favorable interactions with or solvations in water while the hydrophobic *n*-dodecylthiomethyl bristles contribute to the Langmuir monolayer formation and its stability against the complete solvation (*i.e.*, dissolution) of the polymer in water. However, Langmuir films could not be made from PECH-DMAPS₁₀₀ because it is completely soluble in water. This fact indicates that the



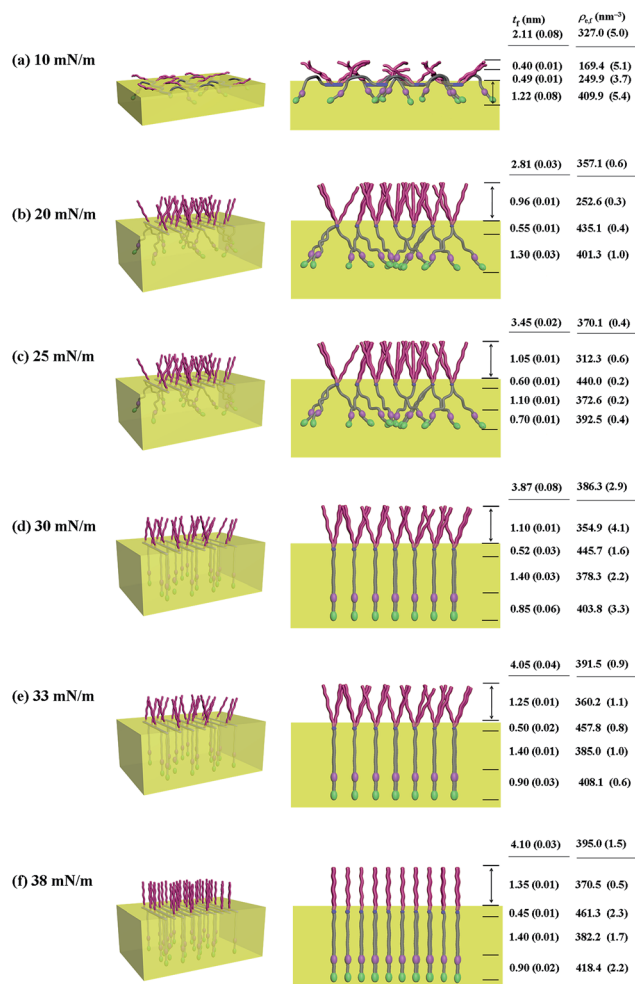


Fig. 9 Molecular models of the Langmuir films of PECH-DMAPS₆₀ prepared at the air–water interface under various surface pressures: (left) view with a tilt angle from the top; (right) front view along the polymer backbone.

solvation power in water of the hydrophilic, zwitterionic DMAPS bristle with the aid of the hydrophilic backbone is relatively stronger than the stabilization power of the *n*-dodecylthiomethyl bristle for Langmuir monolayer. Therefore, a Langmuir film could be formed at the air–water interface from PECH-DMAPS_{*m*} containing <100 mol% DMAPS in which the structural stabilization power of the hydrophobic bristle is balanced with the water-solvation power of the hydrophilic, zwitterionic bristle in a certain level.

Secondly, once a Langmuir film of PECH-DMAPS_{*m*} is formed, it could undergo structural changes by increasing surface pressure, revealing highly ordered monolayer structure. The highly ordered monolayer structure could be achieved by thickening and densification through the conformational changes of the backbone and bristles to nearly or fully extended forms and their lateral packing under increasing surface pressure. Such conformational changes and lateral packing could be driven by the hydrophobic bristles, the alkylenyl linkers in the hydrophilic bristles, and the backbone chain. In particular, the alkylenyl linkers of the hydrophilic bristles present in the water

side could get a certain degree of stabilization (*i.e.*, enthalpy gain) by the lateral packing with the aid of their hydrophobicity against the solvation with water molecules. This could make a significant contribution to enhance the ordering and stability of the Langmuir film. Such enthalpy gain could be more feasible for the hydrophobic alkyl groups in the air side.

Thirdly, at the water surface all PECH-DMAPS_{*m*} polymers (DMAPS: 20–60 mol%) formed only monolayers rather than bilayer or multilayers. These interesting, unique monolayer formation behaviors might be attributed to the relatively high solvation power of DMAPS moieties in the random copolymers under the presence of the hydrophobic bristles.

Lastly, the Langmuir monolayer structure formations are the evidences that the bristles of PECH-DMAPS_{*m*} favorably underwent segregation, forming a hydrophobic bristle phase and a hydrophilic bristle phase. The segregation behavior is remarkable in regard to the random copolymeric nature. In general, such kind of segregation is very rare or not easily happened for random copolymers in bulk or film state. Even though the PECH-DMAPS_{*m*} polymers are random copolymers, such the segregation could be driven at the air–water interface by the selective interactions of the bristles with water molecules and air, leading to Langmuir monolayer formation.

Conclusions

A series of amphiphilic, zwitterionic brush random copolymers, PECH-DMAPS_{*m*}, was investigated at the air–water interface in the aspect of their self-assembling behavior by using π -A isotherm, IR spectroscopy, and XR analyses. These analyses collectively provide insights about the self-assembling behavior of the random copolymer system as follows. PECH-DMAPS₂₀, PECH-DMAPS₄₀, and PECH-DMAPS₆₀ successfully demonstrated the formation of Langmuir layer at the air–water interface, whereas PECH-DMAPS₁₀₀ failed to make Langmuir film because of its high solubility in water. The observed Langmuir films might result from the dual functionalities of PECH-DMAPS_{*m*}, namely its favorable and unfavorable interactions with water molecules in a balanced manner. The favorable interaction with water could be driven by both the hydrophilic backbone and the hydrophilic, zwitterionic DMAPS groups at the bristle ends while the unfavorable interaction with water could be attributed to the hydrophobic *n*-dodecylthiomethyl bristles. Therefore, the composition of the bristles plays a key role in the Langmuir film formation. Interestingly, PECH-DMAPS_{*m*} always forms only Langmuir monolayer film, regardless of the surface pressure in the film formation process. Instead, the Langmuir monolayer undergoes to enhance the lateral packing order in both the hydrophobic bristle phase and the hydrophilic bristle and backbone phase. In particular, the hydrophobic linkers of the bristles could make a significant contribution to the overall lateral packing of the bristles although they are present in the water side. That is, the hydrophobic linkers could maximize enthalpy gains thermodynamically through the lateral packing with the aid of increasing surface pressure, consequently making positive contribution to enhance the packing order of Langmuir film



structure. Furthermore, the zwitterionic DMAPS moieties may have made significant contribution to hold the hydrophilic bristle and backbone phase in the waterside by using their high water affinity (or water salvation power), therefore keeping only Langmuir monolayer structure. Overall, PECH-DMAPS_m revealed very unique self-assembling behavior at the air–water interface, always producing Langmuir monolayer structure.

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Notes and references

- 1 J. E. Raynor, J. R. Capadona, D. M. Collard, T. A. Petrie and A. J. García, *Biointerphases*, 2009, **4**, FA3.
- 2 J. M. Anderson, *Annu. Rev. Mater. Res.*, 2001, **31**, 81.
- 3 B. Zhao and W. J. Brittain, *Prog. Polym. Sci.*, 2000, **25**, 677.
- 4 D. Bozukova, C. Pagnoulle, M.-C. De Pauw-Gillet, S. Desbief, R. Lazzaroni, N. Ruth, R. Jerome and C. Jerome, *Biomacromolecules*, 2007, **8**, 2379.
- 5 H. Ma, D. Li, X. Sheng, B. Zhao and A. Chilkoti, *Langmuir*, 2006, **22**, 3751.
- 6 Z. Zhang, T. Chao, S. Chen and S. Jiang, *Langmuir*, 2006, **22**, 10072.
- 7 E. Ostuni, R. G. Chapman, M. N. Liang, G. Meluleni, G. Pier, D. E. Ingber and G. M. Whitesides, *Langmuir*, 2001, **17**, 6336.
- 8 R. G. Chapman, E. Ostuni, M. N. Liang, G. Meluleni, E. Kim, L. Yan, G. Pier, H. S. Warren and G. M. Whitesides, *Langmuir*, 2001, **17**, 1225.
- 9 S. Chen, J. Zheng, L. Li and S. Jiang, *J. Am. Chem. Soc.*, 2005, **127**, 14473.
- 10 Z. Zhang, S. Chen, Y. Chang and S. Jiang, *J. Phys. Chem. B*, 2006, **110**, 10799.
- 11 W. Feng, J. L. Brash and S. Zhu, *Biomaterials*, 2006, **27**, 847.
- 12 J. Zheng, L. Li, S. Chen and S. Jiang, *Langmuir*, 2004, **20**, 8931.
- 13 H. Kitano, T. Mori, Y. Takeuchi, S. Tada, M. Gemmei-Ide, Y. Yokoyama and M. Tanaka, *Macromol. Biosci.*, 2005, **5**, 314.
- 14 Y. Chang, S. Chen, Z. Zhang and S. Jiang, *Langmuir*, 2006, **22**, 2222.
- 15 A. B. Lowe, M. Vamvakaki, M. A. Wassall, L. Wong, N. C. Billingham, S. P. Armes and A. W. Lloyd, *J. Biomed. Mater. Res.*, 2000, **52**, 88.
- 16 G. Kim, S. Park, J. Jung, K. Heo, J. Yoon, H. Kim, I. J. Kim, J. R. Kim, J. I. Lee and M. Ree, *Adv. Funct. Mater.*, 2009, **19**, 1631.
- 17 G. Kim, L. Y. Hong, J. Jung, D.-P. Kim, H. Kim, I. J. Kim, J. R. Kim and M. Ree, *Biomaterials*, 2010, **31**, 2517.
- 18 G. Kim, Y. Rho, S. Park, H. Kim, S. Son, H. Kim, I. J. Kim, J. R. Kim, W. J. Kim and M. Ree, *Biomaterials*, 2010, **31**, 3816.
- 19 J. Jung, J. C. Kim, Y. Rho, M. Kim, H. Kim and M. Ree, *ACS Appl. Mater. Interfaces*, 2011, **3**, 2655.
- 20 J. C. Kim, J. Jung, Y. Rho, M. Kim, W. Kwon, H. Kim, I. J. Kim, J. R. Kim and M. Ree, *Biomacromolecules*, 2011, **12**, 2822.
- 21 Y. Rho, J. Min, J. Yoon, B. Ahn, S. Jung, K. Kim, P. Shah, J.-S. Lee and M. Ree, *NPG Asia Mater.*, 2012, **4**, e29.
- 22 J. C. Kim, M. Kim, J. Jung, H. Kim, I. J. Kim, J. R. Kim and M. Ree, *Macromol. Res.*, 2012, **20**, 746.
- 23 M. Kim, J. C. Kim, Y. Rho, J. Jung, W. Kwon, H. Kim and M. Ree, *J. Mater. Chem.*, 2012, **22**, 19418.
- 24 J. C. Kim, Y. Rho, G. Kim, M. Kim, H. Kim, I. J. Kim, J. R. Kim and M. Ree, *Polym. Chem.*, 2013, **4**, 2260.
- 25 J. Jung, H. Kim and M. Ree, *Soft Matter*, 2014, **10**, 701.
- 26 J. C. Kim, M. Kim, J. Jung, J.-S. Lee, B. J. Ree, H. Kim, I. J. Kim, J. R. Kim and M. Ree, *J. Polym. Sci., Part A: Polym. Chem.*, 2015, **53**, 1151.
- 27 K. Kwon, C. Kim, J. Lee, H. Kim and M. Ree, *Biomacromolecules*, 2016, **17**, 974.
- 28 I. Kuzmenko, H. Rapaport, K. Kjaer, J. Als-Nielsen, I. Weissbuch, M. Lahav and L. Leiserowitz, *Chem. Rev.*, 2001, **101**, 1659.
- 29 M. Tolan, *X-ray Scattering from Soft Matter Thin Films*, Springer-Verlag Berlin, Heidelberg, 1999.
- 30 L. G. Parratt, *Phys. Rev.*, 1954, **95**, 359.
- 31 J. Bolze, M. Ree, H. S. Yoon, S. H. Chu and K. Char, *Langmuir*, 2001, **17**, 6683.
- 32 M. D. Porter, T. B. Bright, D. L. Allara and C. E. D. Chidsey, *J. Am. Chem. Soc.*, 1987, **109**, 3559.
- 33 M. R. Linford and C. E. D. Chidsey, *J. Am. Chem. Soc.*, 1993, **115**, 12631.
- 34 R. B. Viana, A. B. F. da Silva and A. S. Pimentel, *Adv. Phys. Chem.*, 2012, **2012**, 903272.
- 35 K.-S. S. Kung and K. F. Hayes, *Langmuir*, 1993, **9**, 263.
- 36 Y.-L. Su, J. Wang and H.-Z. Liu, *J. Phys. Chem. B*, 2002, **106**, 11823.
- 37 E. E. Dormidontova, *Macromolecules*, 2002, **35**, 987.

