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The conversion of donor to acceptor and rational design for diketopyrrolopyrrole-containing small molecule acceptors by introducing nitrogen-atoms for organic solar cells†

Jie Pan,^a Hang Yin,^a Yu-Zhong Xie,^a Guang-Yan Sun^{ID}*^a and Zhong-Min Su^{ID}*^b

For the purpose of improving the power conversion efficiencies (PCEs) of bulk heterojunction (BHJ) devices, non-fullerene small acceptors have been presented by modifying desirable donors, such as a diketopyrrolopyrrole (DPP)-based acceptor with two thiazoles, namely poly{3,6-bis(5-hexyldecyl-2thenyl)-2,5-dihydro-2,5-di(alkyl)pyrrolo[3,4]pyrrolo-1,4-dionethiazole} (PDPP2TzT), synthesized by introducing two nitrogens onto the copolymer based on DPP and terthiophene (poly[{2,5-bis(2-hexyldecyl)-2,3,5,6-tetrahydro-3,6-dioxopyrrolo[3,4-c]pyrrole-1,4-diy]}-alt-{{2,2':5',2''-terthiophene-5,5''-diy}}]) (PDPP3T). In order to clarify the transfer mechanism, we put forward a comparative theoretical study of PDPP3T and PDPP2TzT using density functional theory/time-dependent density functional theory (DFT/TD-DFT) calculations. Subsequently, we determined that the mechanism of the conversion from a donor to an acceptor was that the insertion of nitrogen could reduce the LUMO energy by changing the electron density and intramolecular interactions. Based on this mechanism, a series of acceptors were designed with the advantages of PDPP2TzT. The results illustrate that **b4** showed the best performance when blended with PDPP5T as the donor for BHJ devices. Finally, we hope our investigations will provide guidelines for further rational design of acceptor materials for BHJ devices.

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1. Introduction

Bulk heterojunction (BHJ) devices, which exhibit clear phase separation, are famous for converting solar energy into electrical energy.¹ Ordinarily, a classical BHJ device contains both donors and acceptors to enable the entire photoelectric conversion process to occur at the heterojunction interface.² However, in contrast to the great deal of experimental and theoretical studies on electron donor materials, the electron acceptor materials have fallen far behind.^{3–5} For the purpose of improving the power conversion efficiencies (PCEs), more researchers have begun to pay attention to the polymer acceptors and they have determined that it is possible to secure prominent acceptors by introducing electron-withdrawing groups into low level donors. For instance, Barbarella and

co-workers reported a dendritic oligothiophene-*s,s*-dioxide acceptor by adding oxygen to one thienyl ring of oligothiophene, and when it was combined with poly(3-hexylthiophene-2,5-diy) (P3HT), a PCE of 0.3% could be achieved.⁶ Anthony and co-workers synthesized cyano-pentacene acceptors *via* cyano functionalization of pentacenes, finding that the energy levels would drop by 0.14 eV with the reduction of each cyano group.⁷ Sonar and co-workers designed a series of diketopyrrolopyrrole (DPP)-based acceptors by inserting fluoride groups, resulting in an excellent PCE as high as 2.05%.⁸ On this basis, Janssen and co-workers synthesized a new n-type acceptor (poly{3,6-bis(5-hexyldecyl-2thenyl)-2,5-dihydro-2,5-di(alkyl)pyrrolo[3,4]pyrrolo-1,4-dionethiazole}) (PDPP2TzT) (Fig. 1) by replacing adjacent thiophene groups in the copolymer based on DPP and terthiophene (poly[{2,5-bis(2-hexyldecyl)-2,3,5,6-tetrahydro-3,6-dioxopyrrolo[3,4-c]pyrrole-1,4-diy]}-alt-{{2,2':5',2''-terthiophene-5,5''-diy}}]) (PDPP3T) (Fig. 1) with thiazole units and transforming the donor to an acceptor.^{9–14} Here, the introduction of nitrogen is performed with the aim of improving the open circuit voltage (V_{oc}) and short-circuit current density (J_{sc}) by reducing the energy levels and band gap.^{15,16} PDPP2TzT is a promising acceptor material candidate because of its relatively low energy levels (−5.63 eV and −4.00 eV), high electron mobility ($0.13 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$), and broad infrared absorption (300–900 nm).¹⁷

^aDepartment of Chemistry, Faculty of Science, Yanbian University, Yanji, 133002, People's Republic of China. E-mail: gysun@ybu.edu.cn

^bInstitute of Functional Material Chemistry, Faculty of Chemistry, Northeast Normal University, Changchun 130024, People's Republic of China. E-mail: zmsu@nenu.edu.cn

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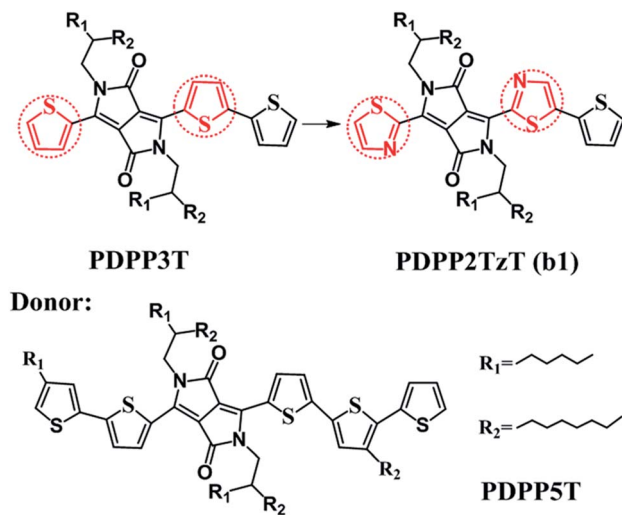


Fig. 1 Molecular structures of PDPP3T, PDPP2TzT, and PDPP5T.

To the best of our knowledge, small molecule acceptor (SMA) materials, which can be acquired from polymer units, have better-defined molecular structures, higher purity without batch-to-batch variations and higher V_{oc} than polymers.^{18,19} Therefore, in this context, we deeply studied the monomer of **PDPP3T** and **PDPP2TzT (b1)** as a SMA material to investigate the reason why the introduction of nitrogen made the donor convert into an acceptor using density functional theory (DFT) and time-dependent density functional theory (TD-DFT) calculations. Moreover, for the purpose of achieving a superior PCE, we designed a series of SMA materials according to the transfer mechanism of donors to acceptors. Also, we chose (poly[*2,5*-bis-(2-hexyldecyl)-2,3,5,6-tetrahydro-3,6-dioxopyrrolo[3,4-*c*]pyrrole-1,4-diyl]-alt-*[[2,2':5',2''*pentathiophene]-5,5''-diyl]) (**PDPP5T**), which is based on DPP and penta-thiophene (Fig. 1), as the complementary donor to construct interface models, **PDPP5T/b1-10**.¹⁷ Some parameters such as the Frontier molecular orbitals (FMOs), energy levels, absorption spectra and charge transport abilities related to V_{oc} , J_{sc} , and PCE were estimated to evaluate the performance of the SMAs in organic solar cells (OSCs).

2. Computational method

2.1. Models

According to the computed results, the alkyl-branched chains (R_1 and R_2 , shown in Fig. 1) were proved to make no difference to the electronic structures and optical properties of the materials (shown in Table S1[†]), therefore these chains were replaced with hydrogen during the whole calculation for the purpose of reducing the computational cost.^{5,19} We tested four hybrid functions, B3LYP, PBE0, B3P86 and B3PW91, to obtain the equilibrium geometries with the 6-31G(d) basis set for the monomer models **PDPP5T** and **PDPP2TzT**.^{19,20} It turned out that the B3P86/6-31G(d) level is in good accordance with the experimental values; therefore, all of the ground-state geometries were optimized at this level (shown in Table S2[†]). As for the

absorption spectra and excited-state energies, the TD-B3P86/6-31G(d) level was utilized. In addition, the fragment orbital correlation maps and charge density difference (CDD) maps, which were used to account for the conversion of the donors and acceptors, were implemented in Multiwfn 3.3.²¹

For the interfaces, the intermolecular distance between the isolated donor and acceptor (monomer, $n = 1$) is a vital factor that affects the performance of the active layers.²² Therefore, we screened plenty of distances. Firstly, we made the left donor unit, the acceptor unit and the right donor unit of the donor correspond to the right donor unit, the acceptor unit and the left donor unit, respectively. This way we obtained nine distances and determined the most stable configuration when the acceptor units of the donor and acceptor corresponded, namely, $X = 0 \text{ \AA}$ (shown in Table S3[†]). Secondly, we made 0 \AA the basis and 1 \AA the interval, and filtered four distances in the positive and negative direction, respectively, along the Y axis. The results are summarized in Table S4[†] and indicate that $Y = 2 \text{ \AA}$ gives the most stable configuration. A number of researchers have proved that 3.5 \AA is the optimal distance for the Z axis of D/A interfaces, which is accordance with the similar system of Bredas and co-workers.²³ Therefore, we thirdly made 3.5 \AA the basis and 0.1 \AA the interval, and filtered nine distances along the Z axis. The results summarized in Table S5[†] show that $Z = 3.4 \text{ \AA}$ is the most suitable distance. Therefore, we formed face-to-face π - π stacking between the donor units, as well as the acceptor units, as the final structure (as an example, the structural model of **PDPP5T/b1** is shown in Fig. S1[†]). Then the ground-state geometries of the **PDPP5T/b1-10** interface models were optimized using the dispersion-corrected B3LYP-D3(BJ) functional with the 6-31G(d) basis set, which has been demonstrated to afford a reasonable interpretation of the interface model of heterojunction structures by Troisi and co-workers.²⁴ To further study the relevant parameters at the D/A interface, all calculations related to excited states were evaluated by TD-DFT theory at the CAM-B3LYP/6-31G(d) level, taking long-rang correction into account, which has been proved to be suitable for evaluating the excitation energies for D/A systems.²⁵⁻²⁷ All of the above calculations were completed in the Gaussian 09 program package.²⁸

2.2. Marcus rate expression

For the weak coupling between D and A, the non-adiabatic electron transfer theory was employed to describe the inter-CT and inter-CR. The inter-CT ($k_{\text{inter-CT}}$) and inter-CR ($k_{\text{inter-CR}}$) rates at the interface formed by **PDPP5T** and **b1-b10** can be expressed by the Marcus semi-classical model:^{29,30}

$$k = \sqrt{\frac{4\pi^3}{h^2 \lambda k_B T}} |V_{DA}|^2 \exp\left(-\frac{(\Delta G + \lambda)^2}{4\lambda k_B T}\right) \quad (1)$$

where λ and V_{DA} represents the total reorganization energy and the electronic coupling at the D/A interface, respectively. ΔG is the Gibbs free energy change of the reaction, h is the Planck constant, k_B is the Boltzmann constant, and T is the temperature defined as 300 K in our calculations. The computational details of λ and ΔG are given in Section S1 of the ESI[†] In



addition, the transfer integrals (V_{DA}) were calculated using the site-energy corrected method with the PW91 functional through the ADF package.^{5,31,32}

3. Results and discussion

3.1. Comparative study of PDPP3T and **b1**

In order to investigate the mechanism *via* which the donor converts to an acceptor, we calculated some parameters, such as the FMOs, energy levels, CDD maps, and fragment orbital correlation maps. Then we defined **PDPP5T** as the donor and **b1** as the acceptor to construct a BHJ device and computed some parameters concerned with the PCE.

3.1.1. FMOs and energy levels. As is well-known, the stimulation, separation, and transfer of excitons is closely related to the distribution of the FMOs, especially the HOMO and LUMO energy levels. Meanwhile, parameters such as V_{oc} and ΔE_{L-L} , which affect the photoelectric conversion efficiency, are correlated with the energy levels. Therefore, choosing an efficient calculation method to obtain the energy levels is a high priority. For the FMOs displayed in Fig. 2, the consummate distribution could be seen for the HOMOs and LUMOs of both **PDPP3T** and **b1**, indicating that the energy of photoelectric conversion is mainly created by the transition of $\pi \rightarrow \pi^*$ and the introduction of nitrogen is unable to change the π -electron delocalization of the molecule through which efficient charge separation could easily be achieved. In addition, we can directly observe the phase changes of the LUMO energy level between **PDPP3T** and **b1**. As the delocalization changed slightly, the LUMO energy level decreased a little as shown in Fig. 2. It is feasible for **b1** to be an acceptor compared with **PDPP3T**. It is also credible to some extent that the LUMO level (-3.67 eV) is close to that of poly[2,7-(5,5-bis-(3,7-dimethyloctyl)-5H-dithieno[3,2-*b*:20,30-*d*]pyran)-alt-4,7-(5,6-difluoro-2,1,3-benzothiadiazole)]/[6,6] phenyl-C71-butyric acid methyl ester ([70]PCBM), which is the most

common acceptor material.¹⁷ Thus we deduce that the introduction of nitrogen may affect the electron density and the intramolecular interactions, and then reduce the LUMO energy level which makes the donors convert to acceptors.

3.1.2. The fragment orbital correlation maps and CDD maps. In order to probe into the internal construction of the LUMO energy level, the fragment orbital correlation is well used, as stated in Fig. 3. We partitioned **PDPP3T** into thiophene, DPP, and bithiophene. Other than substituting bithiophene for thiophene and thiazole, **b1** is divided similarly. It is obviously seen that the low-lying LUMO in **PDPP3T** is mainly attributed to the contribution from the LUMO of thiophene (about 12%), the LUMO+1 of DPP (about 49%) and the LUMO of bithiophene (about 18%). Meanwhile, when bithiophene is replaced with thiophene and thiazole, the contribution rose to 24% leading to an increase in the total electron density for the low-lying LUMO energy level. To this end, the LUMO energy level is greatly decreased and this improves the probability of exciton transition and the ΔE_{L-L} of exciton separation. Meanwhile, the intramolecular and intermolecular interactions, including the

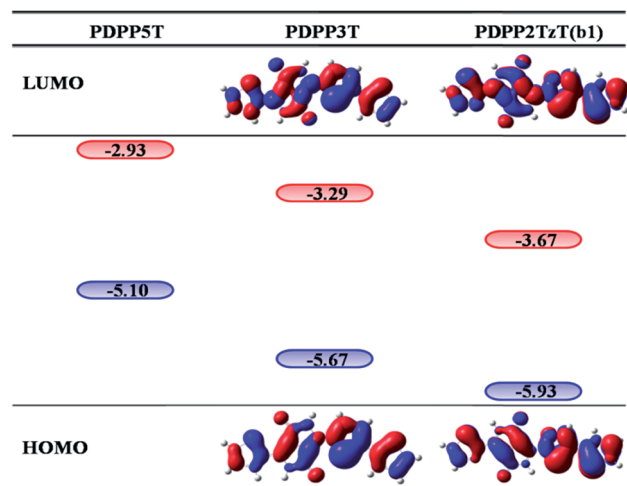


Fig. 2 The Frontier molecular orbitals and energy levels for **PDPP3T** and **b1** (**PDPP2TzT**) at the B3P86/6-31G(d) level, where the navy and red colors represent the HOMO energy levels (eV) and LUMO energy levels (eV), respectively.

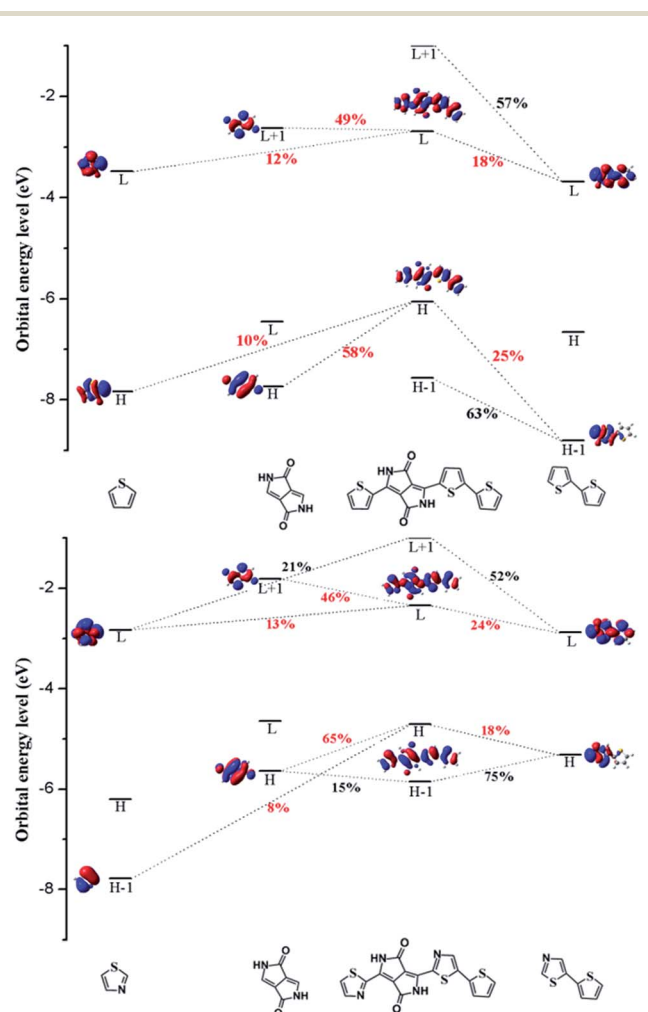


Fig. 3 The molecular orbital correlation diagram for **PDPP3T** (top) and **b1** (bottom) calculated at the B3P86/6-31G(d) level of theory, where H and L stand for the HOMO energy levels and LUMO energy levels, respectively.



A–A and D–A interactions, may increase as a result of adjusting the device configuration and the charge transfer.

CDD is the charge density difference between the ground and excited states and can be used to analyze the charge transfer in conjugated backbones.³¹ Hence, we obtained CDD maps of **PDPP3T** and **b1** to prove the deduction that the charge transfer increases with the addition of nitrogen, as shown in Fig. 4. From the plot, we can intuitively see that the predominant charge transfers for the $S_0 \rightarrow S_1$ transitions of **PDPP3T** and **b1** are practically identical. Therefore, we are left with nothing better than the second choice of S_3 for **PDPP3T** and S_4 for **b1**. It is obvious that the electron density of thiazole is elevated compared with thiophene due to the existence of electron-withdrawing nitrogen, which may enhance the charge transfer and the electron mobility of **b1**. This result is highly in agreement with the analysis of the FMOs and the fragment orbital correlation maps.

In short, we can conclude that the introduction of nitrogen can change the electron density and intramolecular interactions, which then causes the conversion of donors to acceptors.

3.1.3. The BHJ device of PDPP5T/b1. The general rule is that the energy levels of acceptors must be inferior to the energy levels of donors. Furthermore, only when the difference in LUMO energy between the donors and acceptors, namely ΔE_{L-L} , is larger than 0.3 eV can charge separation occur at the interface.¹⁹ According to the details above, we use **b1** as the acceptor together with **PDPP5T** as the donor to construct a BHJ device and characterize some parameters as summarized Table S6.† As you can see, the ΔE_{L-L} is large enough to split the charge, pointing to the existence of the device. Furthermore, V_{oc} is in accordance with the experimental value, suggesting the high performance of the BHJ device (the formulation of V_{oc} is shown in Section S2 of the ESI†).

In addition, we also figured out the energy gap, which directly determines the absorption spectra. As a consequence of the low-lying energy gap, the absorption spectrum exhibits

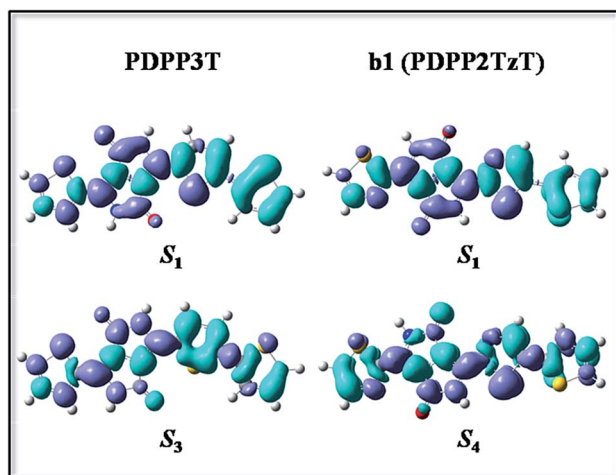


Fig. 4 The CDD map for **PDPP3T** and **b1** (**PDPP2TzT**) at the B3P86/6-31G(d) level, where the violet and turquoise colors represent an increase and a decrease in electron density, respectively.

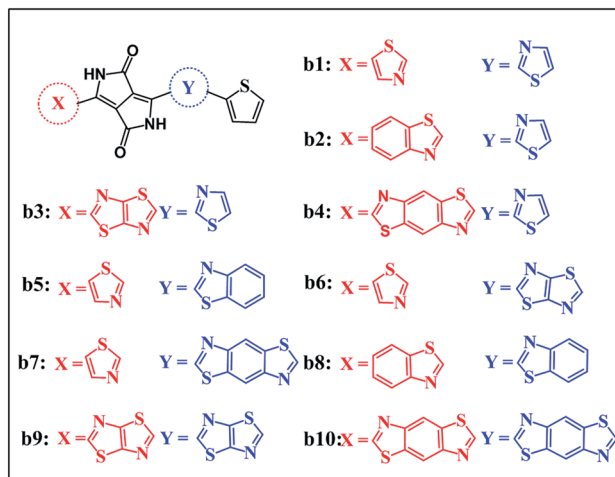


Fig. 5 Molecular structures of all the designed molecules **b1**–**b10**.

a broad absorption in the visible region as shown in Fig. 6. Hence, the device could capture a great deal of photons and increase J_{sc} . Eventually we acquired the $k_{inter-CT}$ and $k_{inter-CR}$ via the Marcus rate expression to evaluate the charge separation and recombination. The result is satisfactory as the $k_{inter-CT}$ is large enough to cause the charge to separate and the $k_{inter-CR}$ is low enough to avoid charge recombination. As all of the results have indicated the electron accepting ability of **b1**, we infer that the introduction of nitrogen may be a helpful mechanism to design novel acceptors.

3.2. Rational design of acceptors based on **b1**

According to the mechanism above, we simultaneously designed acceptors **b1**–**b10** with electron-withdrawing groups such as benzothiazole, thiadiazole, and benzodithiazole, which are shown in Fig. 5, and hoped to make a contribution to improve the performance of BHJ devices.

All of the designed acceptors were calculated with the same methods as **b1**. The calculated FMOs and energy levels are displayed in Fig. S1 and S2.† As seen in the FMOs, the delocalization of **b2**–**b10** is in accordance with that of **b1**, with both the HOMOs and LUMOs distributed over the whole molecule, except for in **b7**. In **b7**, there is obvious charge transfer, which means that the HOMO and LUMO are mainly delocalized at the right and left side, respectively. This may be due to the electron-donating benzodithiazole compared to the thiophene-capped diketopyrrolopyrrole (TDPP) units, which can form “push-pull” electronic structures to facilitate charge separation and tune the energy levels.^{33–35} However, **b4**, whose connection position is different from that of **b7**, cannot form the same peculiar structure and possesses plenty of electrons. Therefore, it is more delocalized and shows a relatively high LUMO energy level, which may enhance the V_{oc} of the BHJ device, as seen from Table 1. In addition, we can see that the ΔE_{L-L} of all molecules is larger than 0.3 eV, which means that efficient charge separation is guaranteed. Hence, all of the designed acceptors could be electron acceptors relative to **PDPP5T**.



The energy gap is the difference between the LUMO and HOMO, and may have an impact on the absorption properties of the molecules. Therefore, we should regard it as a reference. As shown in Fig. S2,[†] we can sort the energy gaps into the following order: **b4** > **b5** > **b8** > **b10** > **b7** > **b6** > **b1** > **b2** > **b3** > **b9**. As you can

Table 1 Calculated energy level values (eV), E_g (eV), ΔE_{L-L} (eV) and V_{oc} (V), for PDPP5T and **b1**–**b10** at the B3P86/6-31G(d) level

	HOMO	LUMO	E_g	ΔE_{L-L}	V_{oc}
PDPP5T	-5.10	-2.93			
b1	-5.93	-3.67	2.26	0.74	1.13
b2	-6.01	-3.76	2.25	0.83	1.04
b3	-6.05	-3.83	2.22	0.90	0.97
b4	-6.07	-3.68	2.39	0.75	1.12
b5	-6.12	-3.74	2.38	0.81	1.06
b6	-6.04	-3.77	2.27	0.84	1.03
b7	-6.04	-3.76	2.28	0.83	1.04
b8	-6.19	-3.83	2.36	0.90	0.97
b9	-6.05	-3.91	2.14	0.98	0.89
b10	-6.29	-3.95	2.34	1.02	0.85

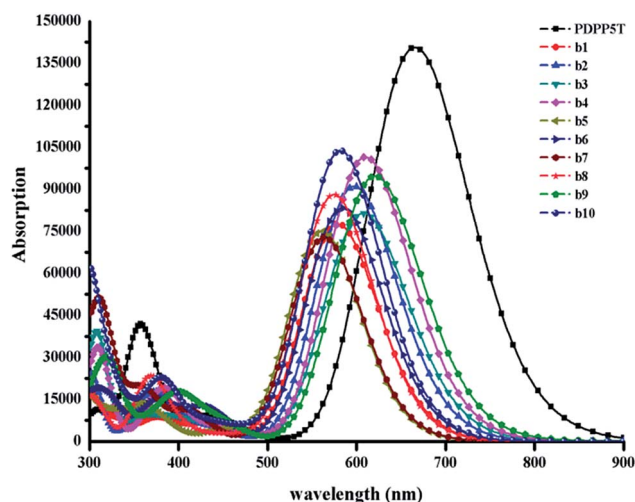


Fig. 6 The absorption spectra of all the designed molecules **b1**–**b10** and PDPP5T calculated at the B3P86/6-31G(d) level in a chloroform solvent. The value of the FWHM is 3000 cm^{-1} .

Table 2 Calculated internal reorganization energy λ_{int} (eV), external reorganization energy λ_{ext} (eV), total reorganization energy λ (eV), $\Delta G_{inter-CR}$ (eV) and $\Delta G_{inter-CT}$ (eV), V_{DA} (eV), $k_{inter-CR}$ (s^{-1}) and $k_{inter-CT}$ (s^{-1}) of PDPP5T/**b1**–**b10**

	λ_{int}	λ_{ext}	λ	$\Delta G_{inter-CR}$	$\Delta G_{inter-CT}$	V_{DA}	$k_{inter-CR}$	$k_{inter-CT}$
PDPP5T/ b1	0.33	0.19	0.52	-1.43	-0.83	0.059	1.64×10^7	1.34×10^{13}
PDPP5T/ b2	0.33	0.18	0.51	-1.34	-0.91	-0.187	1.74×10^9	3.94×10^{13}
PDPP5T/ b3	0.34	0.18	0.52	-1.27	-0.95	0.067	2.98×10^9	3.34×10^{12}
PDPP5T/ b4	0.34	0.16	0.50	-1.42	-0.97	0.013	2.93×10^5	5.26×10^{10}
PDPP5T/ b5	0.33	0.18	0.51	1.36	-1.02	0.028	2.07×10^7	1.33×10^{11}
PDPP5T/ b6	0.33	0.18	0.51	-1.33	-0.94	0.049	1.65×10^8	1.71×10^{12}
PDPP5T/ b7	0.34	0.17	0.51	-1.34	-0.94	0.075	2.83×10^8	3.40×10^{12}
PDPP5T/ b8	0.24	0.16	0.50	-1.27	-1.09	0.001	3.63×10^6	4.13×10^7
PDPP5T/ b9	0.34	0.17	0.51	-1.19	-0.95	0.075	2.08×10^{10}	3.40×10^{12}
PDPP5T/ b10	0.34	0.15	0.49	-1.15	-1.19	-0.160	1.13×10^{11}	3.87×10^{10}

see, **b9** exhibits the lowest energy gap, probably because of the strongest electron-withdrawing ability of thiadiazole. Therefore, **b9** may show the best conjugation and satisfactory absorption properties, which will be further discussed with reference to the absorption spectra below.

The simulated absorption spectra are provided in Fig. 6; all of the designed molecules exhibit wide absorption from 300 nm to 900 nm, which matches well with the solar spectrum. The maximum absorption peak with the largest oscillator strength mainly ranges from $S_0 \rightarrow S_1$ generated by the transition from the HOMO to the LUMO, as seen in ESI Table S7.[†] Furthermore, the range of the maximum absorption highly overlaps with that of PDPP5T, which has been proved to be helpful for achieving high photovoltaic efficiency.³⁶ It is worth nothing that **b9** with the lowest energy gap is bathochromically shifted compared to the other molecules. As the red shift of the absorption spectrum could result in the absorption of more sunlight in the visible spectrum and increase the J_{sc} of the corresponding OSCs, we deduce that **b9** will give a preminent charge separation at the D/A interface, though the maximum oscillator strength is slightly lower than that of **b4** and **b10**.⁵ Certainly, **b4** and **b10** are also capable of achieving outstanding charge separation due to their superior inter-chain π - π^* stacking, which will be discussed in the following section.

It is widely known that not only charge separation but also charge recombination occurs during photovoltaic conversions. Hence, both the inter-CT and inter-CR should be considered to probe the charge properties. The parameters related to inter-CT and inter-CR for $S_1 \rightarrow S_2$ are summarized in Table 2. Upon inspection of Table 2, we find that **b2**–**b10** together with PDPP5T possess comparatively high $k_{inter-CT}$ at about 10^{12} s^{-1} indicating the electron accepting abilities of the designed molecules. Unfortunately, we find that the $k_{inter-CR}$ of **b10** is 1 order of magnitude larger than the $k_{inter-CT}$, so the photovoltaic property of PDPP5T/**b10** is doubtful. Subsequently, the $k_{inter-CR}$ of **b9** is relatively large, which is harmful to the charge separation even though the $k_{inter-CT}$ is quite large. As far as we are concerned, the bigger $k_{inter-CR}$ originates from the inferior $|\Delta G|$ of **b9** and **b10**, as shown in Table 2. It is widely accepted that both the intermolecular CT and CR processes are regarded as exothermic reactions ($\Delta G < 0$).³⁷ And the electron transfer rate will reach a maximum where the values of $|\Delta G|$ and λ are equal. When



$|\Delta G| > \lambda$, the electron transfer rate will increase with a decrease in the $|\Delta G|$.¹⁹ For **b4**, the $k_{\text{inter-CT}}$ is large enough to promote charge separation and the $k_{\text{inter-CR}}$ is the lowest of all. This suggests that the BHJ device constructed from **PDPP5T** and **b4** will exhibit the best photovoltaic properties. Overall, when we define **PDPP5T** as the donor, **b4** is considered to be the most promising acceptor material because of its high LUMO energy level, broad absorption, large $k_{\text{inter-CT}}$ and small $k_{\text{inter-CR}}$.

4. Conclusion

In summary, we have studied the mechanism of the conversion of donors to acceptors by taking the difference of **PDPP3T** and **PDPP2TzT** as an example through DFT and TD-DFT methods. The mechanism of the conversion from a donor to an acceptor is that the insertion of nitrogen reduces the LUMO energy by changing the electron density and intramolecular interactions, which can be observed from the distribution of the FMOs, the fragment orbital correlation maps and CDD maps. Furthermore, on the basis of this mechanism, we designed a series of SMAs and calculated some parameters for the **PDPP5T/b1-b10** models to preliminarily characterize their performance in BHJ devices. The device not only supplies sufficient $\Delta E_{\text{L-L}}$ to ensure exciton separation but also has a high V_{oc} , which is beneficial to the performance of the devices. In addition, the absorption spectra are satisfactory for light absorption, which facilitates the J_{sc} . Sequentially, we calculated the $k_{\text{inter-CT}}$ and $k_{\text{inter-CR}}$ to clarify the electron accepting abilities of the SMAs, but unfortunately we found that the $k_{\text{inter-CR}}$ of **b10** is larger than the $k_{\text{inter-CT}}$, putting the photovoltaic properties of **b10** in doubt. By integrating all of the calculated parameters we found that all of the designed molecules with electron-withdrawing groups exhibited outstanding acceptor abilities and **b4** may be optimal due to its high LUMO energy level, broad absorption, and small $k_{\text{inter-CT}}$. Finally, we hope that our work will provide a novel direction for designing electron acceptors and then accelerate the development of BHJ devices.

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References

- J. You, L. Dou, K. Yoshimura, T. Kato, K. Ohya, T. Moriarty, K. Emery, C. C. Chen, J. Gao and G. Li, *Nat. Commun.*, 2013, **4**, 1446–1455.
- G. Yu, J. Gao, J. C. Hummelen, F. Wudl and A. J. Heeger, *Science*, 1995, **270**, 1789–1791.
- J.-I. Park, J. W. Chung, J.-Y. Kim, J. Lee, J. Y. Jung, B. Koo, *et al.*, *J. Am. Chem. Soc.*, 2015, **137**, 12175–12178.
- K. H. Hendriks, G. H. L. Heintges, V. S. Gevaerts, M. M. Wienk and R. A. J. Janssen, *Angew. Chem., Int. Ed.*, 2013, **52**, 8341–8344.
- Q. Q. Pan, S. B. Li, Y. Wu, G. Sun, Y. Geng and Z. M. Su, *RSC Adv.*, 2016, **6**, 81164–81173.
- G. Barbarella, L. Favaretto, A. Zanelli, G. Gigli, M. Mazzeo, M. Anni and A. Bongini, *Adv. Funct. Mater.*, 2005, **15**, 664–670.
- Y. F. Lim, Y. Shu, S. R. Parkin, J. E. Anthony and G. G. Malliaras, *J. Mater. Chem.*, 2009, **19**, 3049–3056.
- P. Sonar, G. M. Ng, T. T. Lin, A. Dodabalapur and Z. K. Chen, *J. Mater. Chem.*, 2010, **20**, 3626–3636.
- H. Bronstein, Z. Chen, R. Ashraf, W. Zhang, J. Du, J. Durrant, P. Tuladhar, K. Song, S. Watkins, Y. Geerts, M. Wienk, R. Janssen, T. Anthopoulos, H. Sirringhaus, M. Heeney and I. McCulloch, *J. Am. Chem. Soc.*, 2011, **133**, 3272–3275.
- L. Dou, G. Jing, E. Richard, J. You, C. C. Chen, K. C. Cha, Y. He, L. Gang and Y. Yang, *J. Am. Chem. Soc.*, 2012, **134**, 10071–10079.
- S. Zhang, L. Ye, Q. Wang, Z. Li, X. Guo, L. Huo, H. Fan and J. Hou, *J. Phys. Chem. C*, 2013, **117**, 9550–9557.
- I. Meager, R. S. Ashraf, S. Mollinger, B. C. Schroeder, H. Bronstein, D. Beatrup, M. S. Vezie, T. Kirchartz, A. Salleo and J. Nelson, *J. Am. Chem. Soc.*, 2013, **135**, 11537–11540.
- L. Huo, S. Zhang, X. Guo, F. Xu, Y. Li and J. Hou, *Angew. Chem., Int. Ed.*, 2011, **50**, 9697–9702.
- B. Tiede, A. R. Rabindranath, K. Zhang and Y. Zhu, *Beilstein J. Org. Chem.*, 2010, **6**, 830–845.
- Y. Cao, T. Lei, J. Yuan, J.-Y. Wang and J. Pei, *J. Polym.*, 2013, **4**, 5228–5236.
- B. Sun, W. Hong, Z. Yan, H. Aziz and A. Li, *Adv. Mater.*, 2014, **26**, 2636–2642.
- W. Li, W. S. C. Roelofs, M. Turbiez, M. M. Wienk and R. A. J. Janssen, *Adv. Mater.*, 2014, **26**, 3304–3309.
- J. Ren and M. L. Sun, *Chin. J. Org. Chem.*, 2016, **36**, 2284–2300.
- S. B. Li, Y. A. Duan, Y. Geng, H. Z. Gao, Y. Q. Qiu and Z. M. Su, *RSC Adv.*, 2015, **5**, 29401–29411.
- Y. A. Duan, Y. Geng, H. B. Li, X. D. Tang, J. L. Jin and Z. M. Su, *Org. Electron.*, 2012, **13**, 1213–1222.
- T. Lu and F. Chen, *J. Comput. Chem.*, 2012, **33**, 580–592.
- K. R. Graham, C. Cabanetos, J. P. Jahnke, M. N. Idso, L. A. El, G. O. Ngongang Ndjawa, T. Heumueller, K. Vandewal, A. Salleo and B. F. Chmelka, *J. Am. Chem. Soc.*, 2014, **136**, 9608–9618.
- M. K. Ravva, T. H. Wang and J. L. Bredas, *Chem. Mater.*, 2016, **28**, 22–25.
- T. Liu and A. Troisi, *J. Phys. Chem. C*, 2011, **115**, 2406–2415.
- Y. Li, T. Pullerits, M. Zhao and M. Sun, *J. Phys. Chem. C*, 2011, **115**, 21865–21873.
- T. Yanai, D. P. Tew and N. C. Handy, *Chem. Phys. Lett.*, 2004, **393**, 51–57.
- E. K. U. Gross and W. Kohn, *Phys. Rev. Lett.*, 1985, **55**, 2850–2852.
- M. J. Frisch GWT, H. B. Schlegel, G. E. S. Mar, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson HN, X. L. M. Caricato, A. F. H. P. Hratchian, G. Z. Izmaylov JB, J. L. Sonnenberg, M. E. M. Hada, K. Toyota, R. Fukuda JH, M. Ishida, T. Nakajima,



- Y. Honda, H. N. O. Kitao, J. T. Vreven JAM, J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd EB, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari AR, S. S. I. J. C. Burant, M. C. J. Tomasi, N. Rega JMM, M. Klene, J. E. Knox, J. B. Cross, V. Ba kken, C. Adamo, J. Jaramillo RG, R. E. Stratmann, A. J. A. O. Yazyev, R. Cammi, C. Pomelli JWO, R. L. Martin, K. Morokuma, G. V. G. Z akrzewski, A. Voth PS, J. J. Dannenberg, S. Dapprich, A. D. Daniels, Ö. Farkas, J. B. Foresman JVO, J. Cioslowski and D. J. Fox, *R. A. Gaussian 09, Revision A.02*, Gaussian, Inc, Wallingford CT, 2009.
- 29 R. A. Marcus, *Angew. Chem., Int. Ed. Engl.*, 1993, **32**, 1111–1121.
- 30 P. F. Barbara, T. J. Meyer and M. A. Ratner, *J. Phys. Chem.*, 1996, **100**, 13148–13168.
- 31 M. C. Scharber, D. Mühlbacher, M. Koppe, P. Denk, C. Waldauf, A. J. Heeger, *et al.*, *Adv. Mater.*, 2006, **18**, 789–794.
- 32 T. Huser, M. Yan and L. J. Rothberg, *Proc. Natl. Acad. Sci. U. S. A.*, 2000, **97**, 11187–11191.
- 33 Y. A. Duan, Y. Geng, H. B. Li, J. L. Jin, Y. Wu and Z. M. Su, *J. Comput. Chem.*, 2013, **34**, 1611–1619.
- 34 M. Y. Sui, S. B. Li, Q. Q. Pan, G. Y. Sun and Y. Geng, *J. Mol. Model.*, 2017, **23**, 281–288.
- 35 H. Yin, Y. Geng, G. Y. Sun and Z. M. Su, *J. Phys. Chem. C*, 2017, **121**, 2125–2134.
- 36 H. Yao, R. Yu, T. J. Shin, H. Zhang, S. Zhang, B. Jang, M. A. Uddin, H. Y. Woo and J. Hou, *Adv. Energy Mater.*, 2016, **6**, 742–757.
- 37 Y. Yi and V. Coropceanu, *J. Mater. Chem.*, 2011, **21**, 1479–1486.

