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# Modification of a Pd-loaded electrode with a carbon nanotubes–polypyrrole interlayer and its dechlorination performance for 2,3-dichlorophenol†

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A novel Pd loaded Ti electrode was prepared with a carbon nanotubes and polypyrrole interlayer modification, referred to as a Pd/CNTs–PPy/Ti electrode. Pretreated CNTs were deposited on the substrate uniformly by electrophoretic deposition technique. Modification by the CNTs–PPy interlayer made the Pd particles smaller and well-distributed. Pulsed-current electrodeposition technique led to more compact Pd particles. The hydrogen adsorption current value increased 55.3% and the electrochemical active surface area (EASA) increased 58.7% compared to the electrode without modification. The prepared Pd/CNTs–PPy/Ti electrode was employed in dechlorination of 2,3-dichlorophenol (2,3-DCP) in aqueous solution. The potential impact factors on the electrocatalytic dechlorination were studied, including dechlorination current, initial pH value of catholyte, reaction temperature and common ions in aqueous solution. Complete dechlorination could be achieved within 70 min under the selected conditions *i.e.*, current of 5 mA, and an initial pH of 2.5 at ambient temperature. The common ions in aqueous solution, such as  $\text{NO}_3^-$ ,  $\text{CO}_3^{2-}$ ,  $\text{HCO}_3^-$ ,  $\text{Mg}^{2+}$ ,  $\text{K}^+$  and  $\text{Ca}^{2+}$ , had no obvious effect on the electrocatalytic dechlorination within the scope of this investigation. The prepared Pd/CNTs–PPy/Ti electrode exhibited promising dechlorination potential with higher electrochemical activity.

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## 1. Introduction

Chlorophenols (CPs), typical persistent organic pollutants (POPs), are widely used as raw materials in the production of wood preservatives and pesticides,<sup>1–3</sup> and have been listed as priority pollutants by the United States Environmental Protection Agency (USEPA).<sup>4,5</sup> They are highly toxic and resistant to biodegradation. They not only impact ecological security, but also threaten the life and health of human beings. Therefore, the degradation of CPs is of great significance.

There are several methods for CP removal, including biodegradation,<sup>6</sup> chemical reduction,<sup>7,8</sup> and electrocatalytic hydrogenolysis (ECH).<sup>9–11</sup> ECH is considered an environmentally-friendly and highly-efficient technology in the dechlorination of CPs. Active hydrogen ( $\text{H}^*$ ) is generated during the ECH process.  $\text{H}^*$  attacks CPs molecules chemisorbed on the electrode surface, leading to the dechlorination of CPs.<sup>12,13</sup>

Palladium (Pd) is usually used as catalyst in ECH process due to its strong ability in adsorption<sup>14,15</sup> and absorption of hydrogen to form hydride.<sup>16,17</sup> Many researchers attempted to develop methods for improving the electrocatalytic performance of the electrode, such as changing the substrate of the catalytic electrode or modifying the electrode by special materials. Polypyrrole (PPy), a typical conductive polymer with high electrical conductivity and chemical stability, can be used as support material to obtain highly dispersed metallic particles.<sup>18,19</sup> In our previous studies,<sup>20–23</sup> PPy modified composite Pd electrodes shown higher electrochemical activity for CPs dechlorination, indicating the PPy was conducive to Pd particles deposition dispersedly, and it can also play co-catalysis role.

Carbon nanotubes (CNTs) have received increasing attention since they were discovered in 1991.<sup>24</sup> CNTs exhibit good electric conductivity, large specific surface area and high energy density capacity, and have been regarded as an ideal catalyst support material.<sup>25,26</sup> CNTs have been used as Pd/Fe supporter for catalytic dechlorination of chlorinated organic compounds and complete removal of 2,4-dichlorophenol (2,4-DCP) could be achieved within 72 h.<sup>27</sup>

CNTs–PPy composite material might provide synergistic effects for the ECH of CPs. However, to the best of our knowledge, no attention has been paid to the investigation of ECH

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cathode modified with both PPy and CNTs. In the present work, Ti mesh was selected as the substrate due to its good stability. CNTs were introduced as Pd catalyst support material, and we designed the composite electrode which was modified commonly with CNTs and PPy interlayer.

Pd particles were deposited by pulsed-current electrodeposition technique, instead of direct current electrodeposition technique, aiming to further improve the morphology of the electrode surface. The prepared composite electrode, palladium/carbon nanotubes–polypyrrole/titanium electrode (Pd/CNTs–PPy/Ti electrode) was expected to have a better performance and was applied to the dechlorination of 2,3-dichlorophenol (2,3-DCP) in aqueous solution.

## 2. Experimental

### 2.1 Chemicals and materials

CNTs (>95%, length of 10–30  $\mu\text{m}$ , outer diameter of 20–30 nm) were obtained from Beijing Nachen Technology Co. Other chemicals and materials were similar to our previous research.<sup>28,29</sup>

### 2.2 Preparation of Pd/CNTs–PPy/Ti composite electrode

**2.2.1 Pretreatment of meshed Ti.** Ti mesh size and its pretreatment method were same as those in our previous research.

**2.2.2 Pretreatment of CNTs.** Chemical oxidation of CNTs was conducted in a mixture of concentrated acids ( $V_{\text{HNO}_3} : V_{\text{H}_2\text{SO}_4} = 1 : 3$ ) under the assistance of ultrasonication, followed by reflux at 333 K for further treatment. Different ultrasonication time and reflux time were discussed. The oxidized CNTs were washed to neutral with ultrapure water by centrifugation and filtration, and dried for later use.

**2.2.3 Ti mesh surface modification with PPy film and CNTs.** PPy was formed on the surface of pretreated meshed Ti mesh by electro-polymerization in a 50 mL mixed solution containing 0.04 mM Py solution with assistance of 1.0 mM SLS, and the applied current was 3 mA and polymerization time was 5 min. In the process of electro-polymerization, Ti mesh was set as anode, and Pt foil was used as cathode. Especially, the mixed solution was placed in ice bath.

CNTs–PPy/Ti supporting electrode was obtained by electrophoretic deposition of the pretreated CNTs with applied voltage of 10 V and time of 3 min.

**2.2.4 Deposition of Pd catalyst.** Pd/CNTs–PPy/Ti electrode was prepared by pulsed-current electrodeposition, CNTs–PPy/Ti was used as cathode and Pt foil was used as anode. Pulsed-current electrodeposition technique had better performance and Pd/CNTs–PPy/Ti electrode had higher hydrogen adsorption current value by comparison with current electrodeposition technique. Therefore, when it came to achieve the same hydrogen adsorption current value, pulsed-current electrodeposition technique could save more energy. The electrolyte solution consisted of 10 mM  $\text{PdCl}_2$  and 3.0 mM SDBS. Pd particles were loaded on CNTs–PPy/Ti surface by 1000 cycles pulsed-current electrodeposition, with the pulse current

density ( $j_p$ ) of 12.5  $\text{mA cm}^{-2}$  during the turn-on time, turn-on time ( $T_{\text{on}}$ ) of 1.25 s and turn-off time ( $T_{\text{off}}$ ) of 0.75 s. For comparison, Pd/Ti electrode without modification of CNTs–PPy interlayer was also prepared under the same experimental conditions.

### 2.3 Dechlorination experiments of 2,3-DCP

The dechlorination processes were carried out in a two-compartment electrolytic cell (Fig. 1), similar with our previous research.<sup>30</sup> 2,3-DCP initial concentration was 100  $\text{mg L}^{-1}$ , and 0.05 M  $\text{Na}_2\text{SO}_4$  solution was used as supporting electrolyte in both catholyte and anolyte. Potential impact factors, including different applied current, different initial pH values of catholyte adjusted with 0.05 M  $\text{H}_2\text{SO}_4$  solution and common ions in aqueous solution, were discussed.

### 2.4 Analytical methods

The analytical equipments and methods for cyclic voltammetry (CV), transmission microscope (TEM), scanning electron microscope (SEM), X-ray diffractometer (XRD), inductively coupled plasma-atomic emission spectrometry (ICP-AES), and ion chromatography (IC) followed our previous relevant research.<sup>31,32</sup> The level of CNTs surface charge was characterized by the zeta potential (Microtrac Zetatrac, USA), the concentration of CNTs dispersion was 0.5  $\text{mg mL}^{-1}$ , which pH value was adjusted by  $\text{H}_2\text{SO}_4$  solution and NaOH solution. The concentrations of 2,3-DCP and intermediate products were detected by high performance liquid chromatography (HPLC, waters 1525 binary pump and waters 2489 UV/visible detector, USA, C18 column, Kromasil) at wavelength of 280 nm. The mobile phase was 30% water and 70% methanol with flow rate of 0.8  $\text{mL min}^{-1}$ .

### 2.5 Calculation methods

The 2,3-DCP removal efficiency and the dechlorination efficiency were calculated following the methods reported in the literature.<sup>30</sup>

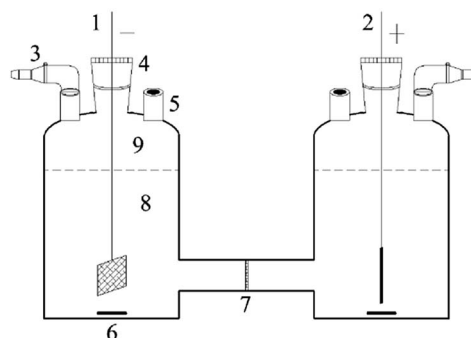


Fig. 1 A schematic of the electrolytic cell for batch experiments, (1) Pd/CNTs–PPy/Ti cathode; (2) Pt anode; (3) expandable syringe; (4) gas-tight adapter; (5) sampling port; (6) stir bar; (7) cation-exchange membrane; (8) electrolyte (30 mL); (9) headspace of cell (10 mL). All adapters were fixed with tape and clamps.



### 3. Results and discussion

#### 3.1 Preparation parameters of CNTs pretreatment and electrophoretic deposition

To obtain a stable and high-efficient electrode, fabrication of homogeneous CNTs film on the PPy/Ti supporting electrode was an essential step, and it was achieved by electrophoretic deposition method. Several factors affect the electrophoretic deposition process, especially the zeta potential of CNTs suspension and the electrophoretic deposition time.<sup>33,34</sup>

**3.1.1 Parameters of CNTs pretreatment.** Considering raw CNTs were inert and hydrophobic. It is vital to disperse CNTs homogeneously in a suitable solvent before electrophoretic deposition process.<sup>35</sup> Carboxyl groups could be introduced on the surface of CNTs by chemical oxidation treatment and provided negative surface charge,<sup>36–38</sup> so the oxidized CNTs had good hydrophilicity and could disperse in water uniformly, and this which was beneficial to the electrophoretic deposition process.

Fig. 2 shows the TEM images of CNTs under different ultrasonication time and reflux time. As shown in Fig. 2a, the raw CNTs had impurities and were poor dispersibility, which indicated the process of CNTs oxidation treatment was insufficient by ultrasonication for 1 h and reflux for 1 h (Fig. 2b). When extending ultrasonication time to 4 h, the treated CNTs behaved better dispersibility. Similar morphologies could be seen from Fig. 2c and d when the reflux time was further extended to 2 h. The oxidized CNTs appeared clean and ordered, and it indicated that the impurities were almost removed.

A higher absolute value of the zeta potential indicates stability of CNTs' suspensions, which is favorable to CNTs

deposition. If the absolute value of zeta potential is lower than 25 mV, the repulsive force between the particles is not strong enough to overcome the van der Waals force,<sup>34,39</sup> resulting in agglomeration and sedimentation of the particles.

Fig. 3 shows the zeta potentials of CNTs' suspension with the different pH values, H<sub>2</sub>SO<sub>4</sub> solution and NaOH solution were used to adjust the pH value of CNTs' suspension. The highest zeta potential absolute value, 53 mV, was achieved at pH value of 6 under condition of ultrasonication for 4 h and reflux for 2 h. The zeta potential absolute value became higher with the extension of reaction time, which indicated better stability of CNTs' suspension. This was because more active groups were introduced on the surface of CNTs with longer reaction time.

Therefore, ultrasonication for 4 h and reflux for 2 h at CNTs suspension pH value of 6 were selected as the appropriate CNTs pretreatment conditions.

**3.1.2 Parameters of CNTs electrophoretic deposition.** Electrophoretic deposition time of CNTs was investigated. When deposition time was only 1 min, a small amount of CNTs were loaded on the surface of electrode (Fig. 4a). When deposition time was extended to 3 min, a compact and uniform CNTs film was achieved (Fig. 4b). But when the deposition time continued to increase, we could clearly see the deposition layer of CNTs partly fall off as it was extensively thick. Consequently, 3 min was employed as the deposition time.

#### 3.2 Characterization of Pd/CNTs-PPy/Ti electrode

**3.2.1 Morphologies of electrode surface.** Fig. 5 shows the SEM images of the prepared electrodes. Pd particles on the surface of Pd/Ti electrode presented spherical agglomerates (Fig. 5a). Pd particles on the surface of Pd/CNTs-PPy/Ti electrode (Fig. 5b) were well-distributed with the modification of CNTs-PPy interlayer. It can be inferred that CNTs-PPy interlayer was favorable to the uniform dispersion of Pd particles, which would benefit the stability of the electrode.

In previous studies,<sup>23</sup> Pd particles deposited on the surface of Pd/PPy-SDBS/Ti electrode with conch-like shape, interestingly, the surface morphology of the Pd/CNTs-PPy/Ti electrode in this study appeared "pinpoint" shape, more compact and uniform, which could provide higher specific surface area and larger reaction contact sites for dechlorination process. This may be

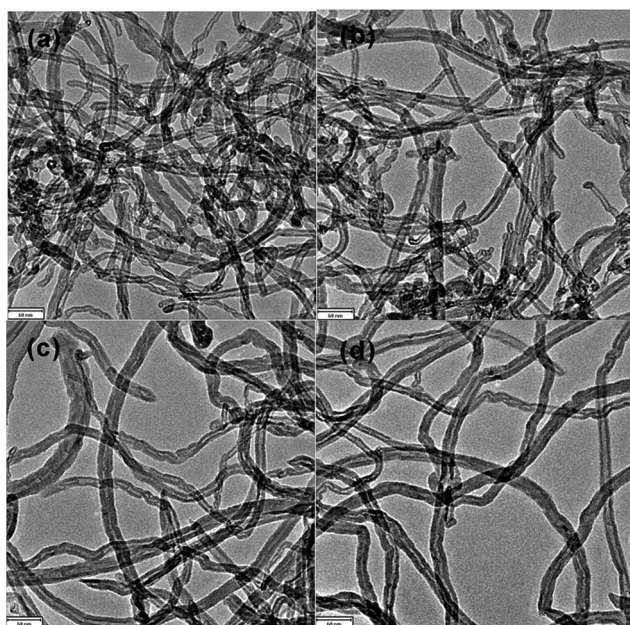


Fig. 2 TEM images of CNTs with different chemical oxidation conditions. Raw CNTs (a), ultrasonication for 1 h and reflux for 1 h (b), ultrasonication for 4 h and reflux for 1 h (c), ultrasonication for 4 h and reflux for 2 h (d).

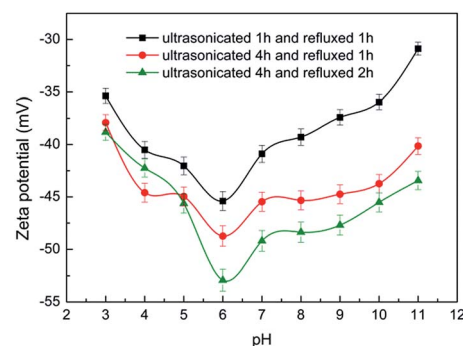


Fig. 3 Zeta potentials of CNTs suspension as a function of pH and pretreatment condition.



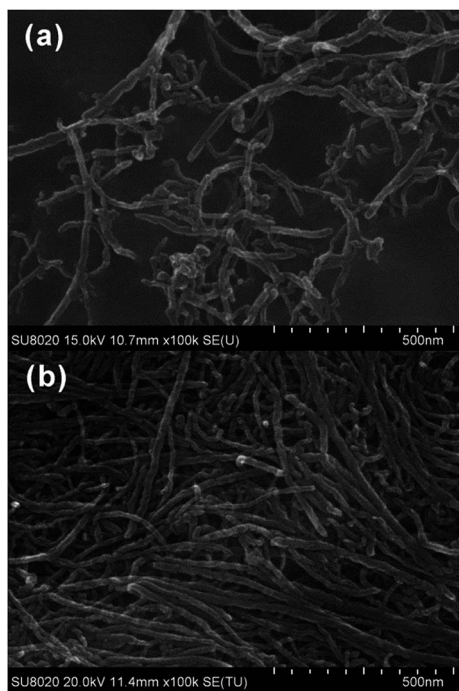


Fig. 4 SEM images of CNTs films under different electrophoretic deposition time. 1 min (a), 3 min (b).

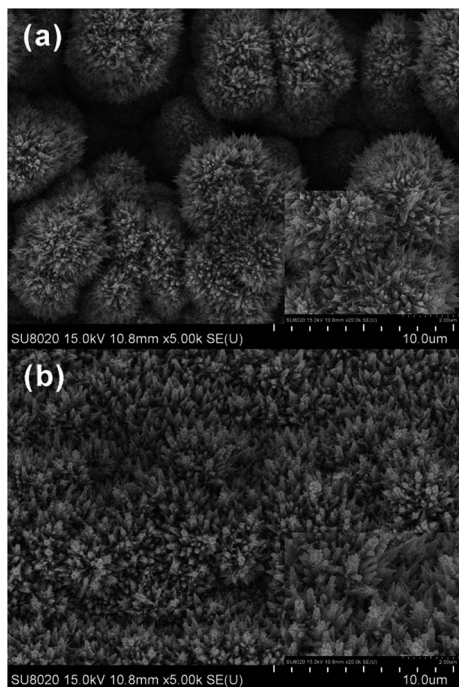


Fig. 5 SEM images (5000 $\times$ ) of Pd/Ti electrode (a) and Pd/CNTs-PPy/Ti electrode (b). The insets are the SEM images (20 000 $\times$ ).

due to the effect of pulsed-current electrodeposition. As reported in the literature,<sup>40</sup> dechlorination of CPs at Pd loaded electrode conformed to hydrogen spillover mechanism. H\* generated firstly on a polarized Pd particle and then diffused to a less polarized nearby Pd article by spillover. It was proposed that the

pinpoint like Pd on the Pd/CNTs-PPy/Ti could be polarized easily, which was conducive to H\* generation and ECH of CPs.

**3.2.2 Crystallite features.** Fig. 6 shows the XRD patterns of Pd/CNTs-PPy/Ti electrode and Pd/Ti electrode. The five strong diffraction peaks at  $2\theta$  of 40.2, 46.7, 68.2, 82.1 and 86.7 were attributed to Pd (111), Pd (200), Pd (220), Pd (311) and Pd (222) crystal lattice, respectively, representing the face-centered cubic (FCC) crystal structure. The crystallite size of Pd on Pd/CNTs-PPy/Ti electrode was 15.8 nm, which was smaller than that on Pd/Ti electrode (19.9 nm), which were calculated by XRD bundled software (EVA). The smaller crystallite size of the Pd/CNTs-PPy/Ti electrode might be due to the effect of CNTs-PPy interlayer.

**3.2.3 CV test.** Fig. 7 shows the CV curves of Pd/CNTs-PPy/Ti electrode and Pd/Ti electrode in 0.5 M H<sub>2</sub>SO<sub>4</sub> solution. In the potential range from  $-0.7$  to  $0.7$  V vs. Hg/Hg<sub>2</sub>SO<sub>4</sub>, two pairs of reduction and oxidation peaks were observed, the oxidation peaks existed from 100 mV to 600 mV and  $-500$  mV to  $-300$  mV corresponded to the oxidation of Pd and the desorption of hydrogen, respectively. It can be seen the curve existed from  $-300$  mV to 200 mV related to the reduction of Pd oxide. A sort of plateau around  $-600$  mV that could indeed be adsorption of hydrogen, and then massive hydrogen discharge around  $-800$  mV.<sup>41</sup> It can be apparently observed that Pd/CNTs-PPy/Ti electrode displayed a higher hydrogen adsorption peak current than Pd/Ti electrode, which suggested the Pd/CNTs-PPy/Ti electrode possessed faster reaction rate under the same conditions.<sup>42</sup> The CNTs-PPy/Ti electrode had a better charge-discharge capacity compared with PPy/Ti electrode and Ti electrode (shown in Fig. S1†). The CNTs-PPy interlayer may act as co-catalyst to ECH. Pd/CNTs-PPy/Ti electrode may offer more chances for the ECH and exhibit higher electrocatalytic activity in the dechlorination reaction.

As shown in Table 1, the hydrogen adsorption current value (at about  $-600$  mV) of Pd/CNTs-PPy/Ti electrode was improved to  $-160$  mA, which increased 55.3% compared with that of Pd/Ti electrode,  $-103$  mA. According to the mechanism of ECH,<sup>43,44</sup> the higher hydrogen adsorption current value in the voltammogram, the more hydrogen atoms were adsorbed on the electrode, which would contribute to a higher efficiency for electrochemical reductive dechlorination of CPs. Therefore, the electrocatalytic performance of electrodes was improved with the modification of CNTs-PPy interlayer. Pd/CNTs-PPy/Ti electrode had better potential for CPs dechlorination.

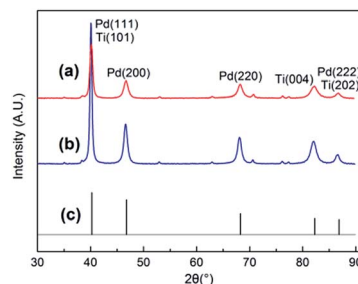


Fig. 6 XRD patterns of Pd/CNTs-PPy/Ti electrode (a), Pd/Ti electrode (b) and Pd (JCPDS-ICDD, 87-0638).



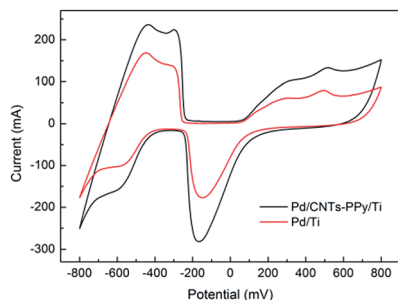


Fig. 7 CV curves of prepared electrodes in 0.5 M H<sub>2</sub>SO<sub>4</sub> solution. Scan rate: 50 mV s<sup>-1</sup>.

Table 1 Characterization parameters of prepared electrodes.  $j_p = 12.5$  mA cm<sup>-2</sup>,  $T_{on} = 1.25$  s,  $T_{off} = 0.75$  s, 1000 cycles,  $C_{Pd(i)} = 10$  mmol L<sup>-1</sup>,  $C_{SDBS} = 3.0$  mmol L<sup>-1</sup>

Electrode	Hydrogen adsorption current (mA)	Pd (mg cm <sup>-2</sup> )	EASA (m <sup>2</sup> g <sup>-1</sup> )
Pd/CNTs-PPy/Ti	160	3.55	16.35
Pd/Ti	103	3.55	10.30

The loading levels of Pd particles on the prepared electrodes were investigated by ICP-AES. The Pd content on Pd/CNTs-PPy/Ti electrode and Pd/Ti electrode were both 3.55 mg cm<sup>-2</sup>.

The electrochemical active surface areas (EASA) of the prepared electrodes were calculated in light of the literatures.<sup>45-47</sup>  $EASA \approx Q/Km$ , where charge  $Q$  is based on the charge of hydrogen atoms adsorption/desorption in the CV test,  $m$  is the mass of Pd of the prepared electrodes and  $K$  is the Pd conversion factor which is 210  $\mu\text{C cm}^{-2}$ . The calculated EASA values were shown in Table 1, the EASA of Pd/CNTs-PPy/Ti electrode was improved to 16.35 m<sup>2</sup> g<sup>-1</sup>, which increased 58.7% compared with that of the Pd/Ti electrode, 10.30 m<sup>2</sup> g<sup>-1</sup>. Pd/CNTs-PPy/Ti electrode with larger EASA was more beneficial to the electrocatalytic reductive dechlorination of chlorophenols. The utilization efficiency of Pd was improved with the modification of CNTs-PPy interlayer.

### 3.3 Parameters of 2,3-DCP dechlorination

**3.3.1 Effects of dechlorination currents.** The influence of dechlorination current was investigated with initial catholyte pH of 2.0 under different constant currents (3 mA, 5 mA, 8 mA and 10 mA, respectively). Fig. 8a plots the effect of constant current on the removal of 2,3-DCP. The removal was accelerated with the increase of current, reached 47% and 100% at 70 min with the applied current of 3 mA and 5 mA, respectively. When the current was increased to 8 mA and 10 mA, complete removal of 2,3-DCP could be achieved within 50 min.

Fig. 8b plots the dechlorination efficiencies of 2,3-DCP varying with the applied current. During the first 30 min, higher dechlorination current led to higher dechlorination efficiency, which consisted with the removal tendency of 2,3-DCP. Dechlorination efficiency decreased when current increased after 30 min. 2,3-DCP dechlorination efficiencies at 70 min were

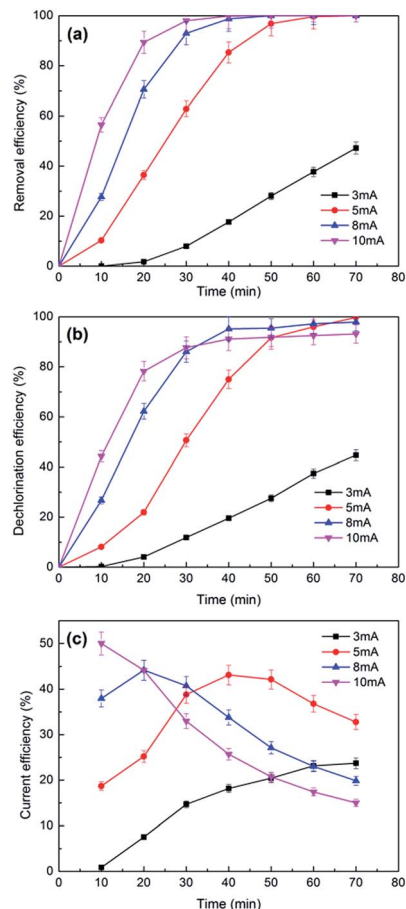


Fig. 8 The effects of currents on the removal (a), dechlorination efficiency (b) and current efficiency (c) of 2,3-DCP dechlorination. Initial pH, 2.0.

100%, 98% and 93% with the current of 5 mA, 8 mA and 10 mA, respectively. Dechlorination efficiency was lower than removal, possibly because the adsorption and dechlorination were carried out simultaneously, part of the adsorbed 2,3-DCP were not dechlorinated in time.

According to the previous studies,<sup>48,49</sup> hydrogen evolution reaction (HER) occurred during the ECH process. The side reaction HER would be speeded up with the current increasing. In addition, when the concentration of 2,3-DCP decreased gradually, the HER process became more strong, which inhibited the ECH process. Fig. 8c shows the current efficiencies of 2,3-DCP dechlorination under different currents. When current of 10 mA was employed, the current efficiency decreased as time went on, indicated that HER influenced dechlorination process seriously at 10 mA. When current of 3 mA was used, the current efficiency increased over time. When currents of 5 mA and 8 mA were used, current efficiencies increased at the early stage and decreased latterly. The current efficiencies at 70 min under 8 mA and 10 mA were 19% and 15%, respectively, much lower than 32% achieved under 5 mA. Considering the reaction efficiency and the cost, 5 mA was selected as the appropriate dechlorination current. The lower energy consumption was part of cleaner production.



**3.3.2 Effects of initial pH values of catholyte.** Under the constant current of 5 mA, the effect of initial pH of catholyte was investigated with different pH values of 2.0, 2.5, 2.8 and 3.2. Fig. 9 shows the influence of initial pH of catholyte on the 2,3-DCP dechlorination. 2,3-DCP could not be completely removed and dechlorinated within 70 min with initial pH value of 3.2, which might be due to the insufficient  $H^+$ . When initial pH value was 2.8, 100% removal and 95% dechlorination efficiency were obtained at 70 min. Complete removal and 100% dechlorination efficiency could be achieved at pH value of 2.5 and 2.0 at 70 min. More hydrogen ions accelerated the generation of  $H^+$ , which promoted the ECH process as well as HER process. The current efficiency under pH 2.0 was lower than that under pH 2.5 due to the stronger HER reaction. Table 2 lists the variation of pH values before and after the dechlorination. Initial pH value of 2.5 was suitable in consideration that neutral solution was beneficial to the subsequent treatment.

**3.3.3 Effects of temperature.** Under constant current of 5 mA and catholyte initial pH of 2.5, the dechlorination experiments were conducted with the temperature of 298 K, 308 K, 318 K and 328 K, respectively. As shown in Fig. 10, the removal and dechlorination efficiency were improved with temperature increased. At 298 K and 308 K, 100% removal and 100%

**Table 2** Variation of pH values of catholyte before and after dechlorination process.  $I = 5$  mA,  $t = 70$  min.  $C_{0,2,3-DCP} = 100$  mg  $L^{-1}$ ,  $C_{Na_2SO_4} = 0.05$  mol  $L^{-1}$

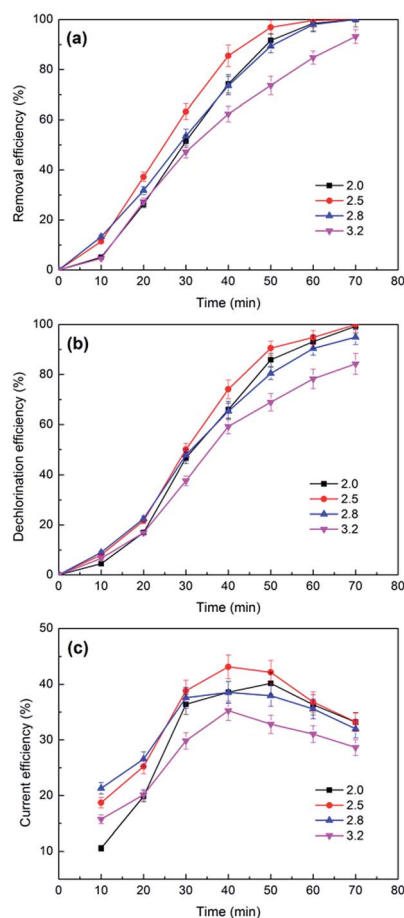
	pH value			
Initial	2.00	2.50	2.80	3.20
Terminal	2.38	6.08	9.35	11.19

dechlorination efficiency of 2,3-DCP dechlorination were achieved at 70 min. Removal efficiency reached 100% at 60 min and dechlorination efficiency reached 100% at 70 min with the temperature of 318 K. When temperature was increased to 328 K, 2,3-DCP could be completely removed at 50 min and thoroughly dechlorinated at 60 min. The high reaction rate was observed at high temperature.

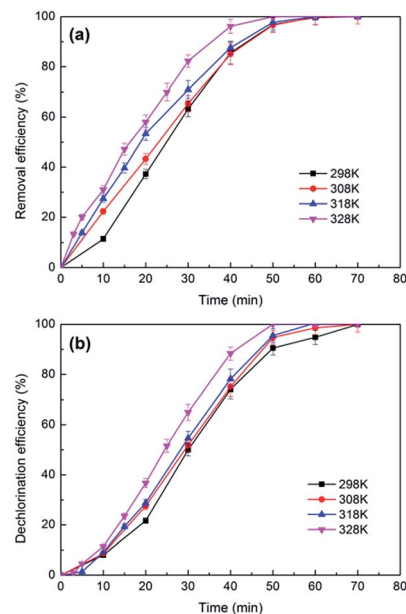
**3.3.4 Effects of common ions in aqueous solution.** The effects of common anions ( $NO_3^-$ ,  $CO_3^{2-}$  and  $HCO_3^-$ ) and cations ( $Mg^{2+}$ ,  $K^+$  and  $Ca^{2+}$ ) in aqueous solution were investigated. The same concentration level of 100 mg  $L^{-1}$  was selected for the discussed common ions. Experiments without the ions mentioned above were conducted for comparison.

The dechlorination experiments with addition of  $NO_3^-$ ,  $CO_3^{2-}$  and  $HCO_3^-$  were investigated under constant current of 5 mA and catholyte initial pH of 2.5. Results showed that the removal and dechlorination efficiency of 2,3-DCP could reach 100% within 70 min, almost similar to the control experiment, as shown in Fig. S2.† The addition of  $NO_3^-$ ,  $CO_3^{2-}$  and  $HCO_3^-$  had no obvious effect on the dechlorination.

Similarly, dechlorination experiments with addition of  $Mg^{2+}$ ,  $K^+$ ,  $Ca^{2+}$  were investigated under constant current of 5 mA and catholyte initial pH of 2.5. Both the removal and dechlorination efficiency of 2,3-DCP could reach 100%, as shown in Fig. S3.†



**Fig. 9** The effects of catholyte initial pH values on the removal (a) and dechlorination efficiency (b) and current efficiency (c) of 2,3-DCP dechlorination. Constant current, 5 mA.



**Fig. 10** The effects of temperature on the removal (a) and dechlorination efficiency (b) of 2,3-DCP dechlorination. Constant current, 5 mA; initial pH, 2.5.



distinct change was observed in comparison to the control experiment, indicating that  $Mg^{2+}$ ,  $K^+$ ,  $Ca^{2+}$  had no obvious effect on the dechlorination.

### 3.4 Dechlorination pathways of 2,3-DCP

Fig. 11 plots the concentrations of 2,3-DCP and intermediates during the dechlorination. As the dechlorination of 2,3-DCP progressing, the concentration of phenol increased gradually. The concentration of 2-chlorophenol (2-CP) increased in the first 30 min and then decreased. No 3-chlorophenol (3-CP) was detected during the whole dechlorination process. Therefore, it was inferred that the pathways of 2,3-DCP dechlorination might follow two routes (Fig. 12): *meta*-Cl was dechlorinated firstly, generating 2-CP, which was dechlorinated further to phenol (route a); *meta*-Cl and *ortho*-Cl were dechlorinated simultaneously (route b). Phenol was the main dechlorination product of 2,3-DCP. The *meta*-Cl was substituted easier due to the steric hindrance effect.<sup>50,51</sup>

### 3.5 Long term performance of Pd/CNTs-PPy/Ti electrode

Under constant current of 5 mA and catholyte initial pH of 2.5, eight times consecutive reduction experiments with 2,3-DCP initial concentration of  $100\text{ mg L}^{-1}$  were conducted. As shown in Fig. 13, complete removal and dechlorination could be achieved within 70 min in each experiment. This indicated that Pd/CNTs-PPy/Ti electrode had good stability.

### 3.6 The electrocatalytic comparison of Pd/CNTs-PPy/Ti electrode, Pd/PPy/Ti electrode, Pd/CNTs/Ti electrode and Pd/Ti electrode

The contrast experiments were used to further reveal the synergistic effect of Pd, CNTs, and polypyrrole for dechlorination of 2,3-dechlorophenol. In Fig. 14, it could be seen that the CV curve of Ti electrode nearly is a straight line, which meant the capacity of adsorbing hydrogen of Ti electrode could be negligible comparing with PPy/Ti electrode. Besides, compared with Ti electrode, it could be clearly seen that the increased hydrogen adsorption current of CNTs-PPy/Ti electrode (16.0 mA) was higher than the sum (6.1 mA) of PPy/Ti electrode and CNTs/Ti electrode. So the better electrocatalytic performance of

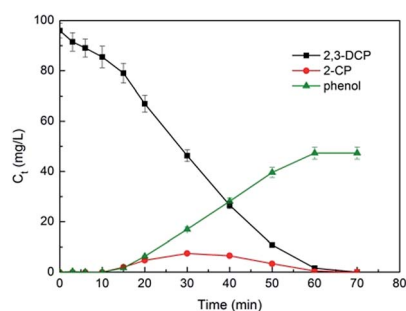


Fig. 11 Variations of 2,3-DCP concentration and intermediate products concentration during the dechlorination. Constant current, 5 mA; initial pH, 2.5.

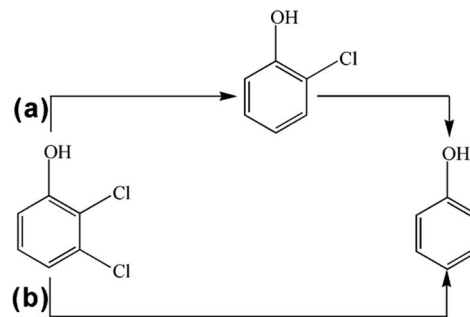


Fig. 12 Proposed pathways of 2,3-DCP dechlorination.

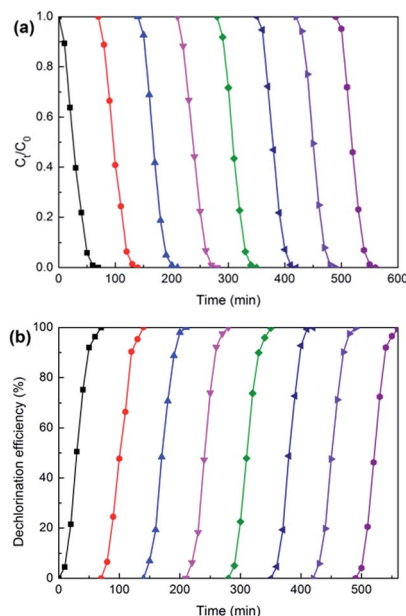


Fig. 13 Variations of 2,3-DCP concentration (a) and dechlorination efficiency (b) during the repeated dechlorinations of 2,3-DCP on Pd/CNTs-PPy/Ti electrode. Constant current, 5 mA; initial pH, 2.5.

CNTs-PPy/Ti electrode was likely to be the synergistic effect of Pd, CNTs, and polypyrrole.

In Fig. 15, it could be found that hydrogen adsorption current of CNTs-PPy/Ti electrode can be negligible comparing with Pd/CNTs-PPy/Ti electrode and Pd/Ti electrode. And

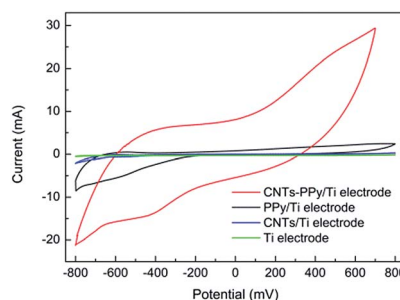


Fig. 14 CV curves of CNTs-PPy/Ti electrode, PPy/Ti electrode, CNTs/Ti electrode and Ti electrode in  $0.5\text{ M H}_2\text{SO}_4$  solution. Scan rate:  $50\text{ mV s}^{-1}$ .



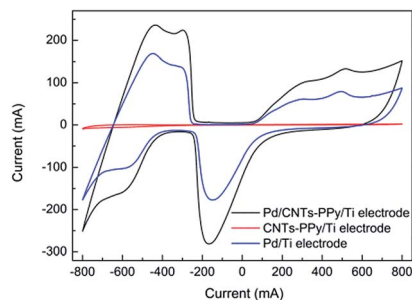


Fig. 15 CV curves of Pd/CNTs-PPy/Ti electrode, CNTs-PPy/Ti electrode and Pd/Ti electrode in 0.5 M  $\text{H}_2\text{SO}_4$  solution. Scan rate:  $50 \text{ mV s}^{-1}$ .

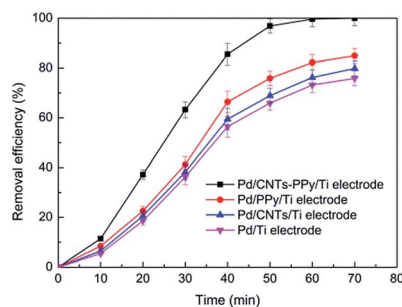


Fig. 16 The removal efficiency of Pd/CNTs-PPy/Ti electrode, Pd/PPy/Ti electrode, Pd/CNTs/Ti electrode and Pd/Ti electrode on 2,3-DCP dechlorination. Constant current, 5 mA; initial pH, 2.5.

hydrogen adsorption current of Pd/CNTs-PPy/Ti electrode was higher than the Pd/Ti electrode. It was likely to be the synergistic effect of Pd, CNTs, and polypyrrole to improve the electrocatalytic performance of the electrode. Besides, we further investigated the removal efficiency of Pd/CNTs-PPy/Ti electrode, Pd/PPy/Ti electrode, Pd/CNTs/Ti electrode and Pd/Ti electrode on 2,3-DCP dechlorination under the same conditions (constant current, 5 mA; initial pH, 2.5). Fig. 16 showed the removal efficiency of Pd/CNTs-PPy/Ti electrode was higher than the other three electrode, and the increased removal efficiency of the Pd/CNTs-PPy/Ti electrode (24.1%) was higher than the sum (13.0%) of Pd/PPy/Ti electrode and Pd/CNTs/Ti electrode compared with Pd/Ti electrode at 70 min, which meant the electrocatalytic performance of the electrode was promoted distinctively when the modified interlayer was CNTs-PPy. So the better electrocatalytic performance of Pd/CNTs-PPy/Ti electrode was likely to be the synergistic effect of Pd, CNTs, and polypyrrole.

## 4. Conclusions

Pd loaded electrode with the modification of CNTs-PPy interlayer showed higher electrocatalytic activity. CNTs-PPy interlayer made Pd particles uniformly deposit on the electrode surface. Pulsed-current electrodeposition technique made the electrode surface get a finer grain structure, and Pd particles were more compact and smaller. Pd/CNTs-PPy/Ti electrode had

higher hydrogen adsorption current value of  $-160 \text{ mA}$  and larger EASA of  $16.35 \text{ m}^2 \text{ g}^{-1}$ . 2,3-DCP could be completely removed and 100% dechlorinated under constant current of 5 mA and catholyte initial pH of 2.5 within 70 min on Pd/CNTs-PPy/Ti electrode. The common ions in aqueous solution including  $\text{NO}_3^-$ ,  $\text{CO}_3^{2-}$ ,  $\text{HCO}_3^-$ ,  $\text{Mg}^{2+}$ ,  $\text{K}^+$  and  $\text{Ca}^{2+}$ , had no obvious effect on the dechlorination within the scope of investigation. The intermediates of 2,3-DCP dechlorination were 2-CP and phenol. Pd/CNTs-PPy/Ti electrode presents the application potential due to its good stability and high electrocatalytic activity.

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