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Behavior of interactions between hydrogen chalcogenides and an anthracene π -system elucidated by QTAIM dual functional analysis with QC calculations†‡

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The nature of $\text{EH}_2 \cdots \pi(\text{C}_{14}\text{H}_{10})$ interactions ($\text{E} = \text{O}, \text{S}, \text{Se}$ and Te) of an anthracene system was elucidated by applying QTAIM dual functional analysis (QTAIM-DFA) after clarification of the structural features with quantum chemical (QC) calculations. π -HB (hydrogen bond) interactions were detected for $\text{E} = \text{O}, \text{S}, \text{Se}$ and Te , whereas π -EB (chalcogen bond) interactions were observed for $\text{E} = \text{O}$ in $(\text{EH}_2) \cdots \pi(\text{C}_{14}\text{H}_{10})$, where the bond paths connected H in EH_2 to $\text{C}_{14}\text{H}_{10}$ in π -HB, and they connected E in EH_2 to C_{10}H_8 in π -EB. The QTAIM-DFA parameters of (R, θ) and (θ_p, κ_p) were evaluated for the interactions *via* analysing the plots of $H_b(r_c)$ versus $H_b(r_c) - V_b(r_c)/2$ for the interactions at the bond critical points. Data obtained from the perturbed structures around the fully optimized structures were employed for the plots, in addition to the fully optimized structures. Data obtained from the fully optimized structures were analysed using (R, θ) , which corresponded to the static nature, and those obtained from the perturbed structures were analysed using (θ_p, κ_p) , which represented the dynamic nature of the interactions, where θ_p corresponds to the tangent line of the plot and κ_p is the curvature. The θ and θ_p values are less than 90° for all the interactions examined, except for the ${}^1\text{H} \cdots {}^{11}\text{C}(\pi)$ interaction in $\text{TeH}_2 \cdots \pi(\text{C}_{14}\text{H}_{10})$ ($\text{C}_1: \text{IIB}_{\text{Atc}}$), where ${}^1\text{H}$ is located closer to the centre of $\text{C}_{14}\text{H}_{10}$. Therefore, the interactions examined were predicted to have vdW nature, appeared in the pure-CS (closed shell) interaction region, although ${}^1\text{H} \cdots {}^{11}\text{C}(\pi)$ was predicted to have the pure-CS/typical-HB nature without covalency. Additionally, the π -HB interaction seems to be slightly stronger than π -EB in $(\text{OH}_2) \cdots \pi(\text{C}_{14}\text{H}_{10})$.

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Introduction

Significant attention has been paid to the interactions between hydrogen chalcogenides (EH_2 ; $\text{E} = \text{O}, \text{S}, \text{Se}$ and Te) and aromatic π -systems, and some structures have been reported for benzene adducts formed through π -interactions.^{1–8} The interactions in benzene π -systems are mainly characterized by the hydrogen bonds formed between EH_2 and the benzene π -system, which act as a proton donor and proton acceptor, respectively, and have been called hydrogen bonds of the π -type (π -HBs) herein.^{1–8} The π -HB (hydrogen bond) interactions were also investigated for 1 : 1 and 2 : 1 adducts between EH_2 ($\text{E} = \text{O}$ and

S) and the naphthalene π -system.^{9–20} Indeed, there have been many investigations on π -adducts, which are mainly based on a theoretical background; however, these investigations seem to be rather limited to structural features. The nature of the π -HB interactions has seldom been reported, especially for the anthracene π -system, $\pi(\text{C}_{14}\text{H}_{10})$, to the best of our knowledge.

Very recently, we reported the nature of the $\text{EH}_2 \cdots \pi$ interactions for benzene and naphthalene π -systems, together with their structural features.^{1,2,10,11} The H atom(s) in EH_2 is (are) connected to the benzene π -system, $\text{EH}_2 \cdots \pi(\text{C}_6\text{H}_6)$, *via* bond paths (BPs). Through careful examination of the BPs in the adducts, another type of interaction was also detected in $\text{H}_2\text{E} \cdots \pi(\text{C}_6\text{H}_6)$, where E in EH_2 is joined to the benzene π -system *via* a BP. Such interaction was called a chalcogen π -type (π -EB) bond. EH_2 is connected to the C atoms or BCPs (bond critical points: r_c : *)²¹ on the $\text{C}=\text{C}$ bonds of C_6H_6 by BPs. In the case of the benzene π -system, π -EBs seem more important relative to π -HBs, although the predicted importance may change depending on the calculation system.²

What happens if EH_2 ($\text{E} = \text{O}, \text{S}, \text{Se}$ and Te) interacts with the anthracene π -system? The anthracene π -system contains two types of (three) benzene π -systems and one type of (two)

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† Dedicated to Professor Marian Mikołajczyk (Professor at the Centre of Molecular and Macromolecular Studies, Polish Academy of Sciences in Lodz, Poland) on the occasion of his 80th birthday.

‡ Electronic supplementary information (ESI) available: QTAIM-DFA approach, Cartesian coordinates for optimized structures of $(\text{EH}_2) \cdots \pi(\text{C}_{14}\text{H}_{10})$ ($\text{E} = \text{O}, \text{S}, \text{Se}$ and Te). See DOI: 10.1039/c7ra04224h



naphthalene π -system, in addition to the original anthracene π -system. What are the differences and similarities in the $\text{EH}_2 \cdots \pi$ interactions between the anthracene π -system and the π -systems of naphthalene and benzene? The nature of the interactions in $(\text{EH}_2) \cdots \pi(\text{C}_{14}\text{H}_{10})$ ($\text{E} = \text{O}, \text{S}, \text{Se}$ and Te) is elucidated together with its structural feature. Scheme 1 illustrates some of the structures expected for $(\text{EH}_2) \cdots \pi(\text{C}_{14}\text{H}_{10})$ of the 1 : 1 adducts, which are inferred from the structures of $(\text{EH}_2) \cdots \pi(\text{C}_{10}\text{H}_8)$ and $(\text{EH}_2) \cdots \pi(\text{C}_6\text{H}_6)$. Type IIA_{Atc} and type IIC_{Atc} in $(\text{EH}_2) \cdots \pi(\text{C}_{14}\text{H}_{10})$ are defined for the adduct, where EH_2 locates on the central and outer benzene rings of anthracene, respectively, whereas it is above the $^{11}\text{C}-^{12}\text{C}$ (or $^{13}\text{C}-^{14}\text{C}$) bond in type IIB_{Atc} . The structures for the 1 : 1 adduct of $(\text{EH}_2) \cdots \pi(\text{C}_6\text{H}_6)$ and $(\text{EH}_2) \cdots \pi(\text{C}_{10}\text{H}_8)$ are also shown in Scheme 1 for convenience of discussion. The optimized structures are called type Ia_{Bzn} , Ib_{Bzn} and type II_{Bzn} for $(\text{EH}_2) \cdots \pi(\text{C}_6\text{H}_6)$ and type I_{Nap} and type II_{Nap} for $(\text{EH}_2) \cdots \pi(\text{C}_{10}\text{H}_8)$.

The QTAIM (quantum theory of atoms-in-molecules) approach, introduced by Bader^{22,23} enables the analysis of the nature of chemical bonds and interactions.^{24–30} Interactions seem to be defined by the corresponding BPs, but we must be careful to use the correct terminology with the concept.³¹ BCP is an important concept in QTAIM, which is a point along the BP at the interatomic surface, where charge density, $\rho(r)$, reaches a minimum.³² It is denoted by $\rho_b(\mathbf{r}_c)$, in addition to other QTAIM functions at BCPs, such as Laplacians of $\rho(r)$ ($\nabla^2 \rho_b(\mathbf{r}_c)$), total electron energy densities $H_b(\mathbf{r}_c)$, potential energy densities $V_b(\mathbf{r}_c)$ and kinetic energy densities $G_b(\mathbf{r}_c)$, together with $k_b(\mathbf{r}_c) (= V_b(\mathbf{r}_c)/G_b(\mathbf{r}_c))$.

Recently, we proposed QTAIM dual functional analysis (QTAIM-DFA),^{33–36} according to QTAIM.^{22–30,37} QTAIM-DFA enables experimental chemists to analyse their own results concerning chemical bonds and interactions with their own image.³⁸ QTAIM-DFA provides an excellent possibility for evaluating, classifying and understanding weak to strong interactions in a unified form.^{33–36} To elucidate the nature of the interactions in question with QTAIM-DFA, $H_b(\mathbf{r}_c)$ is plotted

versus $H_b(\mathbf{r}_c) - V_b(\mathbf{r}_c)/2 [= (\hbar^2/8m)\nabla^2 \rho_b(\mathbf{r}_c)]$, where both the x and y axes are given units of energy. In our treatment, data for perturbed structures around fully optimized structures are employed for the plots, in addition to the fully optimized structures.^{33–36} We propose the concept of “the dynamic nature of interactions” which originated from the data containing the perturbed structures.^{33a,34–36,39}

Data from the fully optimized structures correspond to the static nature of interactions. QTAIM-DFA is applied to typical chemical bonds and interactions and rough criteria are established. The rough criteria can distinguish the chemical bonds and interactions in question from others. QTAIM-DFA and the criteria are explained in the ESI with Schemes S1 and S2, Fig. S1 and eqn (S1)–(S7).[†] The basic concept of the QTAIM approach is also surveyed.

QTAIM-DFA is applied to elucidate the dynamic and static nature of the interactions in $(\text{EH}_2) \cdots \pi(\text{C}_{14}\text{H}_{10})$ ($\text{E} = \text{O}, \text{S}, \text{Se}$ and Te). The discussion is limited to the 1 : 1 adducts of $(\text{EH}_2) \cdots \pi(\text{C}_{14}\text{H}_{10})$ for simplicity. Herein, we present the results of the investigations on the nature of the interactions in question. The interactions are classified and characterized as a reference by employing the criteria.

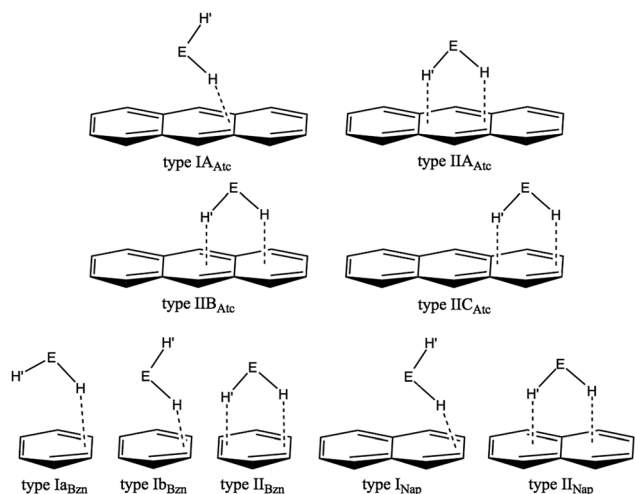
Methodological details in calculations

Structures were optimized using the Gaussian 09 programs.⁴⁰ The 6-311+G(3df) basis set⁴¹ was employed for O, S and Se and the basis set of the (7433111/743111/7411/2 + 1s1p1d1f) type from Sapporo Basis Set Factory⁴² was used for Te with the 6-311+G(d, p) basis set⁴¹ for C and H. The basis set system (BSS) is called BSS-F according to examinations of the BSSs in a previous study. The Møller–Plesset second order energy correlation (MP2) level⁴³ was applied to the calculations. Optimized structures were confirmed by the frequency analysis.

QTAIM functions were calculated by employing the wfn files using the Gaussian 09 program package⁴⁰ with the same method for optimizations, and the data were analysed with the AIM2000 program.⁴⁴ The normal coordinates of internal vibrations (NIV) obtained by the frequency analysis were employed to generate the perturbed structures,^{35,36} which is explained in eqn (1). The k -th perturbed structure in question (S_{kw}) was generated by the addition of the normal coordinates of the k -th internal vibration (N_k) to the standard orientation of a fully optimized structure (S_0) in the matrix representation.³⁵ The coefficient f_{kw} in eqn (1) controls the difference in the structures between S_{kw} and S_0 : f_{kw} is determined to satisfy eqn (2) for the interaction in question, where, r and r_0 stand for the distances in question in the perturbed and fully optimized structures, respectively, with a_0 of Bohr radius (0.52918 Å). The perturbed structures with NIV correspond to that with r being elongated or shortened by $0.05a_0$ or $0.1a_0$, relative to r_0 , as shown in eqn (2). N_k of five digits are used to predict S_{kw} . The selected vibration must contain the motion of the interaction in question most effectively among all the zero-point internal vibrations.

$$S_{kw} = S_0 + f_{kw} N_k \quad (1)$$

$$r = r_0 + wa_0 \quad (w = (0), \pm 0.05 \text{ and } \pm 0.1; a_0 = 0.52918 \text{ \AA}) \quad (2)$$



Scheme 1 Structures expected for $(\text{EH}_2) \cdots \pi(\text{C}_{14}\text{H}_{10})$ ($\text{E} = \text{O}, \text{S}, \text{Se}$ and Te), based on that optimized for $(\text{EH}_2) \cdots \pi(\text{C}_6\text{H}_6)$ and $(\text{EH}_2) \cdots \pi(\text{C}_{10}\text{H}_8)$.



$$y = c_0 + c_1x + c_2x^2 + c_3x^3 \quad (3)$$

(R_c^2 : square of correlation coefficient).

In the QTAIM-DFA treatment, $H_b(r_c)$ is plotted *versus* $H_b(r_c) - V_b(r_c)/2$ for data of five points of $w = 0, \pm 0.05$ and ± 0.1 , as shown in eqn (2). Each plot is analysed using a regression curve of the cubic function, as shown in eqn (3), where $(x, y) = (H_b(r_c) - V_b(r_c)/2, H_b(r_c))$ ($R_c^2 > 0.99999$ usually).⁴⁵

Results and discussion

Structural feature of $(EH_2) \cdots \pi(C_{14}H_{10})$ (E = O, S, Se and Te)

Before the final structural optimizations for $(EH_2) \cdots \pi(C_{14}H_{10})$ (E = O, S, Se and Te), the minima were searched for systematically with MP2/BSS-F. The search was started assuming typically type IIA_{Atc}, type IIB_{Atc} and type IIC_{Atc} structures of the C_1 symmetry. The type IIA_{Atc} and type IIB_{Atc} structures of $(EH_2) \cdots \pi(C_{14}H_{10})$ were optimized for E = O and S, whereas only type IIB_{Atc} was optimized for E = Se and Te. The processes of the convergence are summarized in Table S1 of the ESI.†

The type IIA_{Atc} structure optimized for $(OH_2) \cdots \pi(C_{14}H_{10})$ with MP2/BSS-F was apparently different from type IIB_{Atc}, as confirmed by the r_1 values. In the case of $(SH_2) \cdots \pi(C_{14}H_{10})$, one structure was optimized, which was close to the type IIA_{Atc} structure if initially assuming a typical type IIA_{Atc} structure. The type IIA_{Atc} structure optimized for $(SH_2) \cdots \pi(C_{14}H_{10})$ [$SH_2 \cdots \pi(C_{14}H_{10})$ (C_1 : IIA_{Atc})] seems somewhat distorted from the C_s symmetry.⁴⁶ The search converged to another one, if started assuming typical type IIB_{Atc} and type IIC_{Atc} structures, which are close to type IIA_{Atc}. Indeed, their r_1 values are very close to each other, but the ϕ_2 and ϕ_3 values seem meaningfully different for the two types (Fig. 1c and d, respectively). Therefore, the second structure is (tentatively) called type IIB_{Atc}, herein. No type I_{Atc}

structures were detected after similar treatment, even if optimizations started from those very close to type IA_{Atc}.

Table 1 collects the structural parameters selected for the optimized structures of $(EH_2) \cdots \pi(C_{14}H_{10})$ (E = O, S, Se and Te), $r_1, r_2, r_3, \theta_1, \theta_2, \theta_3, \phi_1, \phi_2$ and ϕ_3 , which are defined in Scheme 2. The optimized structures are not shown in the figures, but they can be found in molecular graphs, which are drawn on the optimized structures (see Fig. 1). What factors appear to control the optimized structures? We compared the H...H distance in EH_2 , $r(H, H: EH_2)$, with the $^{11}C \cdots ^{13}C$ distance in $C_{14}H_{10}$, $r(^{11}C, ^{13}C: C_{14}H_{10})$. Indeed, $r(H, H: EH_2)$ for E = O (1.522 Å), S (1.931 Å), Se (2.088 Å) and Te (2.361 Å) is shorter than $r(^{11}C, ^{13}C: C_{14}H_{10})$ (2.452 Å for the central benzene ring), but the differences in r ($\Delta r = r(C, C: C_{14}H_{10}) - r(H, H: EH_2)$) are larger than 0.5 Å for O ($\Delta r = 0.9303$ Å) and S (0.5215 Å), whereas they are smaller than 0.4 Å for Se (0.3644 Å) and Te (0.0913 Å) (see Table S2 of the ESI†). The type IIB_{Atc} structures of $(EH_2) \cdots \pi(C_{14}H_{10})$ (E = O, S, Se and Te) would form with no limitations of Δr , whereas it may be necessary for Δr larger than around 0.5 Å to give the type IIA_{Atc} structures (E = O and S). In the case of type IIB_{Atc}, one H in EH_2 seems to be slightly above the $^{11}C \cdots ^{13}C$ bond for E = S, Se and Te, but OH_2 seems to exist almost right above the $^{11}C \cdots ^{13}C$ bond, which may also be controlled by Δr . Indeed, the difference in Δr between E = S and Se seems small, at a first glance, but the small difference would play an important role in the appearance of the type IIA_{Atc} structure.

What are the stabilization energies in the formation of the adducts? Table 1 contains the ΔE (ΔE_{ES} and ΔE_{Ent}) values for $(EH_2) \cdots \pi(C_{14}H_{10})$, where, ΔE_{ES} and ΔE_{Ent} are the ΔE values on the energy surfaces and those with the thermal corrections to enthalpies, respectively [$\Delta E = (E((EH_2) \cdots \pi(C_{14}H_{10}))) - (E(EH_2) + E(C_{14}H_{10}))$]. An excellent correlation was obtained in the plot of ΔE_{Ent} *versus* ΔE_{ES} ($y = 5.262 + 1.047x$; $R_c^2 = 0.998$ ($n = 6$: number of data points)), although is not shown in the figure. Therefore,

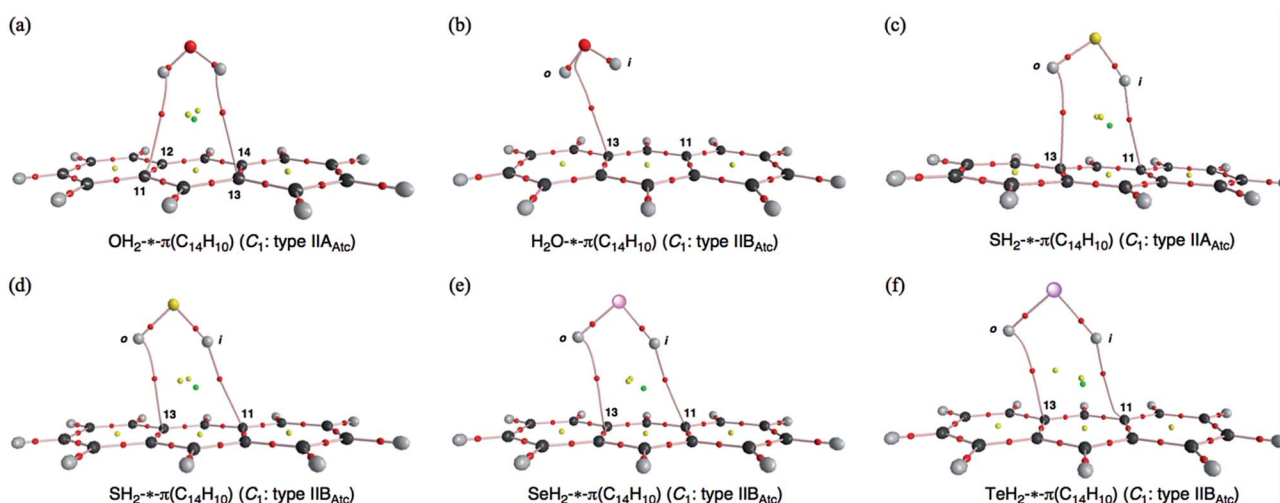


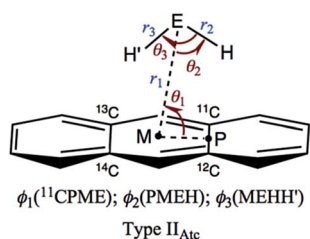
Fig. 1 Molecular graphs for $OH_2 \cdots \pi(C_{14}H_{10})$ (C_1 : type IIA_{Atc}) (a), $H_2O \cdots \pi(C_{14}H_{10})$ (C_1 : type IIB_{Atc}) (b), $SH_2 \cdots \pi(C_{14}H_{10})$ (C_1 : type IIA_{Atc}) (c), $SH_2 \cdots \pi(C_{14}H_{10})$ (C_1 : type IIB_{Atc}) (d), $SeH_2 \cdots \pi(C_{14}H_{10})$ (C_1 : type IIB_{Atc}) (e) and $TeH_2 \cdots \pi(C_{14}H_{10})$ (C_1 : type IIB_{Atc}) (f), calculated with MP2/BSS-F. The bond critical points (BCPs) are denoted by red dots, ring critical points (RCPs) by yellow dots and cage critical points (CCPs) by green dots, together with bond paths by pink lines. Carbon atoms are in black and hydrogen atoms are in grey, with oxygen, sulphur, selenium and tellurium atoms in red, yellow, pink and purple, respectively.



Table 1 Structural parameters for (EH₂)···π(C₁₄H₁₀) (E = O, S, Se and Te) optimized at the MP2 level with BSS-F^{abc}

Species (X–Y) (symmetry: type)	r ₁ (Å)	r ₂ (Å)	r ₃ (Å)	θ ₁ (°)	θ ₂ (°)	θ ₃ (°)	φ ₁ (°)	φ ₂ (°)	φ ₃ (°)	ΔE _{ES} ^d (kJ mol ⁻¹)	ΔE _{Ent} ^e (kJ mol ⁻¹)
OH ₂ ···π(C ₁₄ H ₁₀) (C ₁ : IIA _{Atc})	3.2865	0.9629	0.9629	89.9	51.8	103.5	-92.5	-1.7	2.8	-22.7	-18.1
H ₂ O···π(C ₁₄ H ₁₀) (C ₁ : IIB _{Atc})	3.6682	0.9631	0.9624	63.4	69.6	103.6	-93.8	-14.3	3.5	-21.9	-17.6
SH ₂ ···π(C ₁₄ H ₁₀) (C ₁ : IIA _{Atc}) ^g	3.6162	1.3389	1.3400	80.5	60.6	92.2	-92.1	-6.4	-3.3	-28.7	-25.0
SeH ₂ ···π(C ₁₄ H ₁₀) (C ₁ : IIB _{Atc}) ^g	3.6165	1.3389	1.3400	80.5	60.6	92.2	-91.8	-2.1	-1.9	-28.7	-25.0
TeH ₂ ···π(C ₁₄ H ₁₀) (C ₁ : IIB _{Atc})	3.6959	1.4597	1.4603	78.9	61.5	91.3	-91.5	3.0	-0.2	-31.7	-28.3
TeH ₂ ···π(C ₁₄ H ₁₀) (C ₁ : IIB _{Atc})	3.7703	1.6576	1.6578	78.1	64.2	91.5	-89.9	-11.4	-17.5	-40.5	-36.8

^a BSS-F: the 6-311+G(3df) basis set was employed for O, S and Se and the basis set of the (7433111/743111/7411/2 + 1s1p1d1f) type for Te with the 6-311+G(d,p) basis set for C and H. ^b See Scheme 2 for the structural parameters. ^c Optimized structures are not given in the figures but can be found in the molecular graphs drawn on the optimized structures (see Fig. 1). ^d ΔE_{ES} = E_{ES}((EH₂)*-π(C₁₄H₁₀)) - (E_{ES}(EH₂) + E_{ES}(C₁₄H₁₀)) on the energy surface. ^e ΔE_{Ent} = E_{Ent}((EH₂)*-π(C₁₄H₁₀)) - (E_{Ent}(EH₂) + E_{Ent}(C₁₄H₁₀)) with the thermal corrections to enthalpies. ^f Very close to C_s. ^g The structures are very close to each other; however, they are analysed as two different structures here since the differences in φ₂ and φ₃ seem meaningful. They are called IIA_{Atc} and IIB_{Atc}.

Scheme 2 Structural Parameters Illustrated for Type II_{Atc} of (EH₂)*-π(C₁₄H₁₀) (E = O, S, Se and Te).

the ΔE_{ES} values will be employed for the discussion of ΔE. A good correlation was also obtained in the plot of ΔE_{ES} versus r₁, if the data for H₂O···π(C₁₄H₁₀) (C₁: IIB_{Atc}) are neglected (y = -1410.42 + 820.87x - 121.30x²; R_c² = 0.975 (n = 5)), where the structure of H₂O···π(C₁₄H₁₀) (C₁: IIB_{Atc}) seems different from the others.

What are the relations between ΔE_{Atc}, ΔE_{Nap} and ΔE_{Bzn}? The ΔE_{Nap} values are linearly proportional to ΔE_{Bzn} in type II. The ratio of ΔE_{Nap}/ΔE_{Bzn} becomes larger in the order of E = O (ΔE_{Nap}/ΔE_{Bzn} = 1.34) < S (1.60) < Se (1.68) < Te (1.75), which shows a substantial chalcogen dependence. On the other hand, the ΔE_{Atc}/ΔE_{Nap} ratio seems almost constant for E = O (ΔE_{Atc}/ΔE_{Nap} = 1.01) < S (1.07) < Se (1.09) < Te (1.11). The different basis sets for Te may somewhat affect the evaluated values for E = Te, which are from the Sapporo Basis Set Factory, whereas the others are from the Gaussian09 program. What mechanisms operate to control the ratios? It is difficult to clarify this, however, we examined the ratios of r_{1 : Bzn}/r_{1 : Nap} and r_{1 : Atc}/r_{1 : Nap}. The r_{1 : Bzn}/r_{1 : Nap} ratios become larger in the order of 1.01 (E = O) < 1.13 (S) < 1.15 (Se) < 1.19 (Te) and the r_{1 : Atc}/r_{1 : Nap} ratios increase similarly in the order of 1.00 (E = O) < 1.09 (S) < 1.11 (Se) < 1.13 (Te). However, the chalcogen dependence of the ratios seems somewhat smaller for r_{1 : Atc}/r_{1 : Nap}, relative to the case for r_{1 : Bzn}/r_{1 : Nap}. These results would be responsible for the observed ΔE/ΔE ratios, although the vdW and/or covalent radii of chalcogens must also be carefully examined for the discussion.

Before the application of QTAIM-DFA to (EH₂)*-π(C₁₄H₁₀), the molecular graphs and contour plots are examined next.

Molecular graphs and contour plots of ρ(r) for (EH₂)*-π(C₁₄H₁₀)

Fig. 1 illustrates the molecular graphs for (EH₂)*-π(C₁₄H₁₀) (E = O, S, Se and Te), which were calculated with MP2/BSS-F. The BCPs are clearly detected, which contain that expected between EH₂ and C₁₄H₁₀. Each BP with a BCP connects each H in EH₂ and the carbon atom of ¹¹C or ¹³C in C₁₄H₁₀ for all optimized structures in Table 2, except for the type IIB_{Atc} of H₂O*-π(C₁₄H₁₀). The BP with BCP joins O in OH₂ and a carbon atom of ¹¹C or ¹³C in C₁₄H₁₀ of OH₂*-π(C₁₄H₁₀). Therefore, the interactions between EH₂ and C₁₄H₁₀ are classified by π-HBs for all the adducts in Table 2, except for the type IIB_{Atc} of H₂O*-π(C₁₄H₁₀), the interaction of which should be classified by the π-EB type. Ring critical points (RCPs) and cage critical points (CCPs) are also detected in Fig. 1. Fig. 2 shows the contour plots of ρ(r) for (EH₂)*-π(C₁₄H₁₀) (E = O, S, Se and Te) calculated with MP2/BSS-F. The contour plots of ρ(r) are drawn on a plane containing E, BCP (*) of the (EH₂)*-π type and an atom or BCP suitable for the contour plots of (EH₂)*-π(C₁₄H₁₀). The BCPs are well located at the (three dimensional) saddle points of ρ(r) in the species. Negative Laplacians and trajectory plots are drawn in Fig. S2 and S3 of the ESI.† It is well visualized how the BCPs are classified through ∇²ρ(r) and the space around the species is well divided by the atoms in it.

Survey of the interactions in (EH₂)*-π(C₁₄H₁₀) (E = O, S, Se and Te)

As shown in Fig. 1 and 2, some BPs are apparently curved. Therefore, the lengths of the BPs (r_{BP}) will be substantially different from the straight-line distances (R_{SL}) in some cases. The r_{BP} and R_{SL} values are collected in Table S2 of the ESI.† for (EH₂)*-π(C₁₄H₁₀) (E = O, S, Se and Te), which were evaluated with MP2/BSS-F, together with the differences between r_{BP} and R_{SL} (Δr_{BP} = r_{BP} - R_{SL}). There are two types of BPs in each adduct since each adduct has C₁ symmetry. Therefore, it is necessary to define the BPs to distinguish the two. Letting ¹H and ⁰H in EH₂ connect to π(C₁₄H₁₀) through BPs, which are placed more inside and more outside, respectively, in relation to the centre of C₁₄H₁₀, then ¹BP and ⁰BP can be defined as ¹BP (¹H*-π) and ⁰BP (⁰H*-π), respectively.



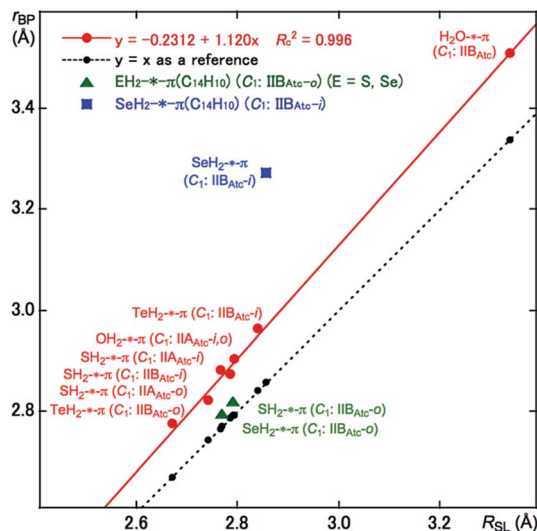


Fig. 3 Plots of r_{BP} versus R_{SL} for $(EH_2)^*-\pi(C_{14}H_{10})$ ($E = O, S, Se$ and Te), evaluated with BSS-F at the MP2 level.

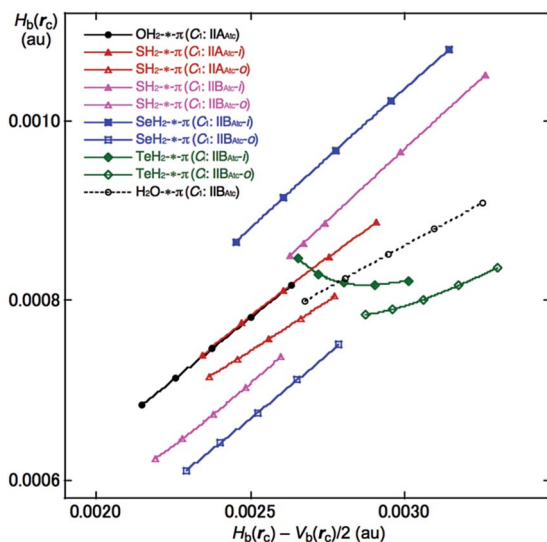


Fig. 4 Plots of $H_b(r_c)$ versus $H_b(r_c) - V_b(r_c)/2$ for $(EH_2)^*-\pi(C_{14}H_{10})$, evaluated with MP2/BSS-F. Compounds with marks are shown in the figure.

to eqn (S3)–(S6) of the ESI ‡ by applying QTAIM-DFA. These results are discussed next.

Nature of π -HBs and π -EBs in $(EH_2)^*-\pi(C_{14}H_{10})$ ($E = O, S, Se$ and Te)

Table 2 collects the QTAIM-DFA parameters of (R, θ) and (θ_p, κ_p) for the π -HB and π -EB interactions, together with the frequencies and the force constants (k_f) correlated to NIV, which are employed to generate the perturbed structures. The nature of π -HBs and π -EBs in $(EH_2)^*-\pi(C_{14}H_{10})$ is discussed based on the (R, θ, θ_p) values employing the standard values as a reference (see Scheme S2 of the ESI ‡). It should be instructive to survey the criteria related to that herein before a detailed discussion.

The criteria tell us that $45^\circ < \theta < 90^\circ$ for pure-CS interactions. The θ_p value predicts the character of the interactions. In the p-CS region, the character of the interactions will be the vdW type for $45^\circ < \theta_p < 90^\circ$, whereas it will be the t -HB (typical hydrogen bonds) type with no covalency (t -HB $_{nc}$) for $90^\circ < \theta_p \leq 125^\circ$, where $\theta_p = 125^\circ$ tentatively corresponds to $\theta = 90^\circ$.

The θ and θ_p values are less than 90° for all the π -HBs and π -EB in $(EH_2)^*-\pi(C_{14}H_{10})$ examined herein, except for the $Te^1H^0H^*-\pi(C_{14}H_{10})$. Therefore, the π -HBs and π -EB interactions in $(EH_2)^*-\pi(C_{14}H_{10})$ in Table 2 are all classified by the pure-CS interactions and characterized as the vdW type (p-CS/vdW), except for $Te^1H^0H^*-\pi(C_{14}H_{10})$. Although 0BP ($Te^1H^0H^*-\pi(C_{14}H_{10})$) is predicted to have the nature of (p-CS/vdW), 1BP ($Te^1H^0H^*-\pi(C_{14}H_{10})$) is predicted to have the nature of (p-CS/ t -HB $_{nc}$). The π -EB interaction in the $H_2O^*-\pi(C_{14}H_{10})$ type [$(\theta, \theta_p) = (73.9^\circ, 79.3^\circ)$] is predicted to be somewhat stronger than the π -HB interaction of the $OH^*-\pi(C_{14}H_{10})$ type [$(\theta, \theta_p) = (72.6^\circ, 74.5^\circ)$] for $OH^*-\pi(C_{14}H_{10})$. The π -HB interactions in the anthracene system are predicted to be very similar to that in the naphthalene system, which seem slightly stronger than that in the benzene system. The results are in accordance with that derived from the energies for the formation of the adducts.

The delocalization indexes and ellipticity are the important parameters to clarify and understand the nature of the $(EH_2) \cdots \pi$ interactions. 47 The nature of the $(EH_2) \cdots \pi$ interactions will be discussed elsewhere based on these parameters for the series of $(EH_2) \cdots \pi(C_6H_6)$, $(EH_2) \cdots \pi(C_{10}H_8)$, and $(EH_2) \cdots \pi(C_{14}H_{10})$ ($E = O, S, Se$ and Te).

Conclusions

The behaviour of the interactions for (EH_2) adducts with the anthracene π -system ($E = O, S, Se$ and Te) in a 1 : 1 ratio is elucidated by applying QTAIM-DFA. The structures were optimized with MP2/BSS-F. Two types of structures were optimized for $E = O$ and S , whereas one was optimized for $E = Se$ and Te . The BCPs are clearly detected in the molecular graphs. The interactions are the π -HB type for $(EH_2)^*-\pi(C_{14}H_{10})$ ($C_1: IIA_{Atc}$) ($E = O$ and S) and $(EH_2)^*-\pi(C_{14}H_{10})$ ($C_1: IIB_{Atc}$) ($E = S, Se$ and Te), whereas they are the π -EB type for $H_2O^*-\pi(C_{14}H_{10})$ ($C_1: IIB_{Atc}$). The QTAIM-DFA parameters of (R, θ) and (θ_p, κ_p) are calculated for the species, according to eqn (S3)–(S6). ‡ The θ values are less than 90° for all interactions examined in this work. The θ_p values are also less than 90° for all the interactions in question, except for ${}^1H^*-\pi(C_{14}H_{10})$ in $TeH_2^*-\pi(C_{14}H_{10})$ ($C_1: IIB_{Atc}$). Therefore, all the π -HB and π -EB interactions examined in this work are classified by pure-CS interactions and characterized as vdW in nature (p-CS/vdW), although ${}^1H^*-\pi(C_{14}H_{10})$ in $TeH_2^*-\pi(C_{14}H_{10})$ ($C_1: IIB_{Atc}$) is predicted to have the p-CS/ t -HB nature without covalency. The π -EB interaction is predicted to be somewhat stronger than π -HB in $(OH_2)^*-\pi(C_{14}H_{10})$. It is demonstrated that the predicted nature of $(EH_2)^*-\pi(C_{14}H_{10})$ is closer to that of $(EH_2)^*-\pi(C_{10}H_8)$ rather than that of $(EH_2)^*-\pi(C_6H_6)$. The π -HB and π -EB interactions in the anthracene π -system are well elucidated by applying QTAIM-DFA.



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- 31 Interactions would be easily imaged by means of QTAIM if they can be defined as the corresponding bond paths (BPs), especially for experimental chemists. However, it is demonstrated that the detection of the BPs between two atoms in a molecule emerging from natural alignment of the gradient vector field of the one-electron density of a molecule is neither necessary nor a sufficient condition for the presence of a chemical bond between those atoms.⁴⁸ In this connection, it is pointed out that the terms line paths (LPs) and line critical points (LCPs) should be used in place of BPs and BCPs, respectively.^{48b} Consequently, the dynamic and static nature in this work should be regarded as the investigation performed at LCPs on LPs corresponding to the EH₂–π interactions in the anthracene–π system. Nevertheless, BPs and BCPs are used in this paper for the discussion of the EH₂–π(C₁₄H₁₀) interactions, similarly to the case of the previous discussion for the EH₂–π(C₆H₆) and EH₂–π(C₁₀H₈) interactions. The interactions expected between EH₂ and



- $\pi(\text{C}_{14}\text{H}_{10})$ are clearly detected by BPs with BCPs which is another reason to use BPs and BCPs in this work.
- 32 Critical points (CPs) are characterized by the rank (ω) and the signature (σ). The CPs of the species in the three-dimensional space are classified by $\omega = 3$, which generally corresponds to all species. On the other hand, σ is defined by the simple algebraic sum of the signs of $\partial^2 \rho_b(r_c)/\partial r_i^2$ ($r_i = x, y$ and z for $i = 1, 2$ and 3 , respectively), where the + and - signs of $\partial^2 \rho_b(r_c)/\partial r_i^2$ are counted as +1 and -1, respectively. Therefore, $\sigma = -3, -1, 1$ and 3 correspond to attractors (nuclei), bond critical points (BCPs), ring critical points (RCPs) and cage critical points (CCPs), respectively. Namely, BCP is characterized by $(\omega, \sigma) = (3, -1)$.^{17,18}
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