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Incorporating doped carbon nanodots and metal ions as an excellent artificial peroxidase for H₂O₂ detection†

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Novel nitrogen-doped carbon nanodots (NCdots) were successfully synthesized by a hydrothermal method using 1-(3-dimethylaminopropyl)-3-ethylcarbodiimide hydrochloride (EDC) as a carbon–nitrogen source. The NCdots were characterized by methods including TEM, EDX, SAED, XPS, and UV-vis. The NCdots were found to possess intrinsic chromogenic activity, which could be used as an excellent artificial peroxidase to oxidize H₂O₂ catalytically by metal ions to produce a chromogenic reaction. The reason for the chromogenic mechanism might be attributed to the structural alteration of NCdots by free radical scavenging similar to Fenton's reagent in the presence of H₂O₂. This offers a simple and selective colorimetric method for the determination of some metal ions in environmental systems.

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Introduction

Carbon-based nanodots consisting of graphene quantum dots (GQDs) and carbon quantum nanodots (CQDs, Cdots or CDs) are a new species of carbon nanomaterial with sizes less than 10 nm.¹ As a consequence of their special surface state and unique physical and chemical properties, carbon-based nanodots have gradually become an important new nanocarbon member due to their benign, abundant and inexpensive nature.^{2,3} Nevertheless, to date, nearly all applications of pure carbon-based nanodots have low surface activity limiting their application.^{4,5} Compared with pure carbon nanodots (Cdots), doping Cdots with heteroatoms is more attractive for analytical applications due to their outstanding intrinsic properties including higher photostability and fluorescent quantum yield, low photobleaching, more surface-passivated groups, and other special physicochemical properties.^{6–8} Moreover, doped Cdots are commonly biocompatible and have low toxicity, which makes them superior to quantum dots.² Recently, applications

of doped Cdots in metal ion and biomolecule detection based on their optical characteristics have been demonstrated by several groups.^{9–12} These applications are important and attractive. Up to now, the surface activity of the doped Cdots as an artificial peroxidase has rarely been reported although there are some studies on intrinsic enzymatic activities of Cdots.¹³ Herein, we demonstrate the intrinsic chromogenic activity of doped Cdots in the oxidation of H₂O₂ catalytically in the presence of some metal ions to produce a brown color.

Hydrogen peroxide (H₂O₂), as a principal member of the reactive oxygen class and a by-product of reactions catalyzed by many oxidase enzymes, plays important roles in biological systems, and industrial, pharmaceutical, and other fields.¹⁴ Therefore, the development of effective analytical methods for the selective detection of trace amounts of H₂O₂ is an especially important subject in current chemical research. To date, various traditional techniques and methods, including electrochemical strategies, spectroscopic approaches, and enzymatic methods, have been developed for H₂O₂ detection.^{15–17} Among them, enzymatic methods of detection have proved to be very powerful owing to their high substrate specificity and efficiency. However, limited natural sources, high-cost purification processes, time-consuming preparation, easy inactivation and inherent instability restrict their applications to some extent. To

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overcome these drawbacks, increasing attention has focused on enzyme mimetics. Compared with natural enzymes, the efficiency of most non-enzymes is still lower, especially enzyme activity, and which can be dramatically decreased by the addition of coatings and bioconjugation. Thus, developing a convenient and simple method for the preparation of nanomaterials with intrinsic enzymatic activities remains highly desirable before using enzymatic techniques for H₂O₂ detection.

In this work, a one-pot simple green method was developed to synthesize nitrogen-doped Cdots (NCdots) using 1-(3-dimethylaminopropyl)-3-ethylcarbodiimide hydrochloride (EDC) as a carbon–nitrogen precursor, without any other additives or further chemical modifications. According to EDX and XPS analysis, the compositions of NCdots were determined, and the results showed that the resultant NCdots possess intrinsic organic redox dye (ORD)-like chromogenic activity, and could oxidize H₂O₂ catalytically in the presence of some metal ions to produce a chromogenic reaction. The NCdots were successfully used for the analysis of some metal ions and H₂O₂ based on the NCdots and the metal ion reaction.

Experimental

Reagents and chemicals

1-(3-Dimethylaminopropyl)-3-ethylcarbodiimide hydrochloride (EDC) was purchased from Sigma-Aldrich. Hydrogen peroxide (H₂O₂), copper sulfate pentahydrate (CuSO₄·5H₂O), cobalt chloride hexahydrate (CoCl₂·6H₂O), manganese chloride tetrahydrate (MnCl₂·4H₂O), iron sulfate heptahydrate (FeSO₄·7H₂O), iron chloride (FeCl₃), chromium chloride hexahydrate (CrCl₃·6H₂O) and tris(hydroxymethyl)aminoethane (Tris) were obtained from Shanghai Sinopharm Chemical Reagent Co. Ltd. (Shanghai, China). Other chemicals and solvents were of analytical grade and were used without further purification. Double distilled water was used throughout.

Apparatus

Fluorescence (FL) spectra were recorded on a Fluoromax-4 fluorescence spectrofluorometer (Horiba, USA). UV-vis absorption spectra were recorded on a Shimadzu UV-2550 spectrophotometer (Tokyo, Japan). Transmission electron microscopy

(TEM) and selected-area electron diffraction (SAED) patterns were conducted on a JEM-2100 transmission electron microscope (JEOL Ltd.). Energy dispersive X-ray spectroscopy (EDX) was carried out using a FEI Sirion 200 scanning electron microscope (FEI). X-ray photoelectron spectroscopy (XPS) analysis was performed on a Thermo ESCALAB 250XI X-ray photoelectron spectrometer (USA).

Synthesis of NCdots

Fluorescent NCdots were prepared by traditional hydrothermal treatment of EDC. Briefly, 0.55 g of EDC was dissolved in 12.5 mL of deionized water and 12.5 mL of ethanol to form a homogeneous solution under sonication. The 25 mL as-prepared solution was then put into a poly(tetrafluoroethylene) (Teflon)-lined autoclave and heated at 180 °C for 8.0 h, and then cooled to room temperature naturally. The resulting solution was centrifuged at 10 000 rpm for 30 min to remove the solid residue, and a brown-yellow aqueous solution of NCdots was obtained.

Fluorescent detection procedure for H₂O₂

For the H₂O₂ sensing experiments, solutions of 50 μL Tris–HNO₃ buffer (50 mM, pH 6.0), 10 μL NCdots (1 mg mL⁻¹), 50 μL Cu²⁺ ion (1.0 mM), and different amounts of H₂O₂ solution were added to a centrifuge tube (1.5 mL). Then, the solution was completely mixed with a vortex mixer at room temperature for a few seconds to accelerate the reaction. The final volume of the mixture was adjusted to 500 μL with double distilled water. The mixture was equilibrated at room temperature for 30 min before FL spectra measurements were recorded. The resulting solutions were studied by FL spectra at room temperature with excitation at 330 nm, and both the excitation and emission slit widths were 5 nm.

Results and discussion

Synthesis and characterization of fluorescent NCdots

TEM was performed to observe the morphology of the prepared NCdots. A TEM image of the NCdots supported on carbon nanodots synthesized using EDC as the carbon–nitrogen source in the precursor solution is shown in Fig. 1A. The NCdots thus formed are uniform, monodisperse, and the corresponding particle size distribution histogram (inset in Fig. 1A) indicates that these carbon nanodots have diameters ranging from 5.0 to 7.0 nm (50 NCdots have been used to calculate the size distribution, RSD 5.17%). The HRTEM image in the inset of Fig. 1A clearly reveals that the diffraction contrast of the NCdots is very high and with obvious lattice fringes (2.17 Å), which is consistent with the (102) diffraction planes of sp² graphitic carbon,

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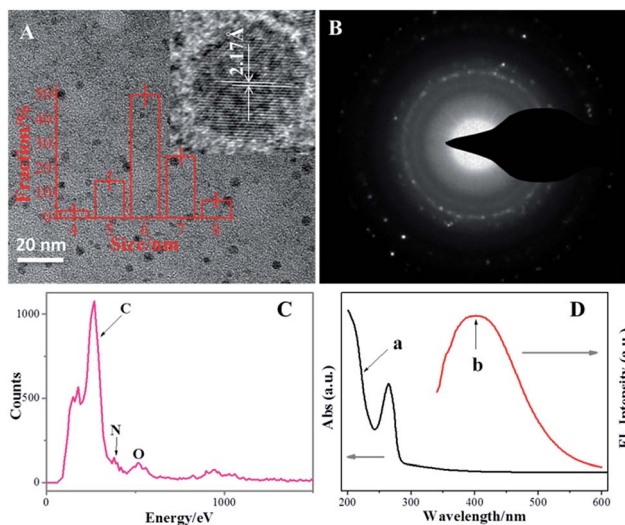


Fig. 1 TEM image (A), SAED image (B), EDX (C), and UV-vis/FL spectra (D) of the NCdots. Inset in (A) shows the diameter distribution and HRTEM of the NCdots.

implying that the NCdots maintained similar crystallinity to graphene.¹⁸ So the prepared NCdots are uniform and of high crystallinity. Meanwhile, Fig. 1B shows the SAED pattern of the NCdots. SAED reveals the crystalline nature of the NCdots, as also confirmed in the inset in Fig. 1A. The formation of NCdots was confirmed by EDX as shown in Fig. 1C. EDX measurement of the NCdots revealed the presence of mainly carbon, nitrogen, and oxygen, confirming that nitrogen had been successfully doped in the NCdots (9.069 at%) by the present one-pot *in situ* approach.

XPS is often used to determine the surface composition and elemental analysis of carbon nanomaterials. The XPS survey spectrum (Fig. 2A) of the NCdots shows three typical peaks at 285.7, 399.8, and 531.5 eV, which are attributed to C_{1s} , N_{1s} , and O_{1s} , respectively. The XPS spectrum results indicate that these NCdots are mainly composed of C (83.162%), as well as a limited

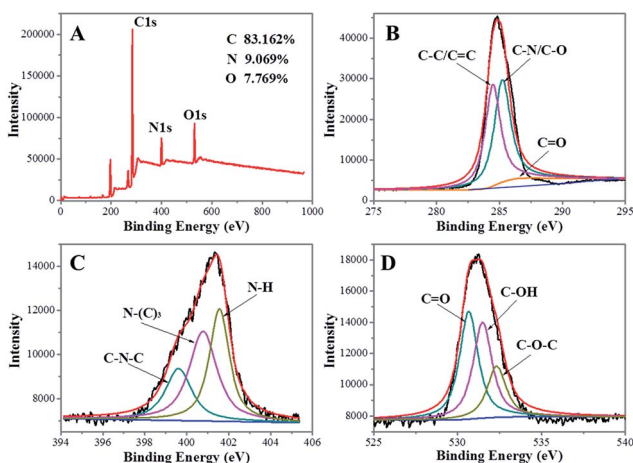


Fig. 2 (A) XPS survey spectra of the as-produced NCdots. High-resolution spectra of C_{1s} (B), N_{1s} (C) and O_{1s} (D) in NCdots.

amount of N (9.069%) and O (7.769%) elements (the atomic ratio of C/N/O is about 11 : 1 : 1). The high resolution spectrum of C_{1s} exhibits three main peaks (Fig. 2B). The binding energy peak at 284.5 eV confirms the graphitic structure (sp^2 , C=C) of the NCdots. The peaks at about 285.3 and 286.5 eV suggest the presence of C-N/C-O and C=O groups, respectively.³ The high-resolution spectrum of N_{1s} (Fig. 2C) shows three peaks at 399.6, 400.7 and 401.5 eV, which are attributed to the C-N-C, N-(C)₃, and N-H bands, respectively.¹⁹ The O_{1s} spectrum (Fig. 2D) shows three peaks at 530.6, 531.5, and 532.4 eV, which are attributed to C=O, C-OH and C-O-C groups, respectively.⁶ Such observations also suggest that the NCdots have a graphitic carbon structure. Therefore, we could conclude that the NCdots obtained are mainly composed of polycyclic heterocyclic aromatic C=C species derived from the dehydration and polymerization of EDC and ethanol, as well as possessing abundant hydroxyl, amino and carbonyl moieties on their surface.

Fig. 1D shows that the suspension of NCdots has a typical UV-vis absorption peak at 263 nm with a weak shoulder at \sim 330 nm. The absorption peak at \sim 263 nm was assigned to the σ - π and π - π^* transitions originating from the carbene-like triplet state of aromatic sp^2 domains of carbon nanodots, which leads to virtually no observed FL signal,²⁰ while the other transition centered at \sim 330 nm is due to the trapping of excited-state energy by the surface states and results in strong emission.²¹ An excitation-dependent emission was observed with the red-shifted emission peaks from long excitation (Fig. S1, ESI[†]), showing the multicolor

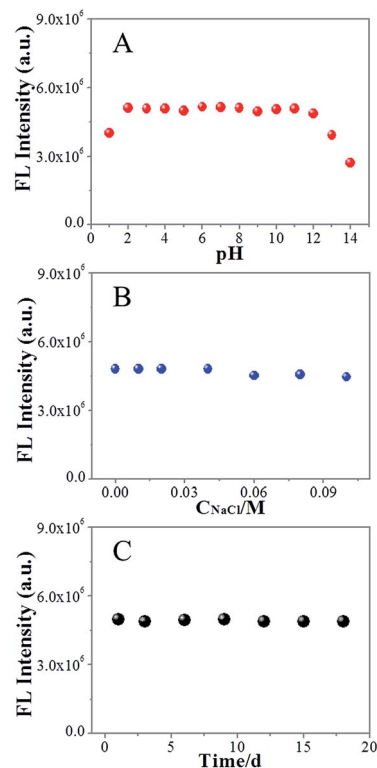


Fig. 3 FL intensity at 390 nm (excitation at 330 nm) of the NCdots at different pH values (A), different concentrations of NaCl solution (B), and the stability of the NCdots as a function of storage time (C). RSD ($n = 5$) is 4.23% (A), 4.57% (B) and 3.95% (C), respectively.



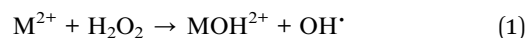
properties of the NCdots and characteristics of carbon materials.²² The emission intensity increased in the excitation range of 270–330 nm and then decreased gradually. The maximum emission peak was observed at 390 nm for an excitation of 330 nm (Fig. 1D, curve b).

To study the photostability of NCdots under different conditions, the FL intensity of NCdots was investigated at different pH values, ionic strength environments and storage times (Fig. 3). Fig. 3A shows the fluorescence intensities of NCdots at different pH values. It is seen that the FL of the NCdots is strong and stable over a wide range of pH values (2–12, Fig. 3A). When the pH value is lower than 2 or higher than 12, the FL intensities decrease gradually. Therefore, the effect of pH can be understood in terms of the change in surface charge owing to intermolecular and intramolecular protonation/deprotonation, which could be attributed to the presence of CN units and O-containing groups on the surface and edge.⁶ To confirm the FL stability of NCdots under high ionic strength environments, we measured the FL intensities in the presence of different concentrations of NaCl (up to 0.1 M). As shown in Fig. 3B, only a slight change in FL intensity was observed, indicating the high stability of the NCdots even under an environment of high ionic strength. Meanwhile, it is worth mentioning that these NCdots in suspension can be very stable for 18 days, without observation of any floating or precipitated nanodots (Fig. 3C). These results indicate the excellent photostability of the NCdots, probably due to the electrostatic repulsions between the negatively charged nanodots resulting in electrosteric stabilization. These findings indicate that NCdots have great potential for sensing applications under physiological conditions.

Investigation of the chromogenic mechanism

The organic redox dye (ORD)-like activity of the NCdots was evaluated in the catalysis of H₂O₂ in the presence of metal ions.

NCdots could catalyze H₂O₂ in the presence of some metal ions, and produced a typical chromogenic reaction. As shown in Fig. 4A, six stronger chromogenic reactions were observed for NCdots upon addition of Cu²⁺, Mn²⁺, Co²⁺, Fe²⁺, Fe³⁺ or Cr³⁺ ions, while other metal ions were not able to exert a significant chromogenic reaction. Based on the above results, it can be speculated that the chromogenic mechanism might be attributed to the structural alteration of NCdots by the excellent OH[•] radical scavenging properties in the presence of H₂O₂, like Fenton's reagent, resulting in the mixed solution showing a chromogenic reaction.^{23,24} It is well-known that OH[•] radicals can be generated, as with Fenton's reagent, in which these six metal ions react with H₂O₂, producing OH[•] radicals and metal hydroxides (eqn (1)).²⁵ Meanwhile, the FL intensity of these mixed solutions can be easily observed visually under UV light (365 nm), and the solution of NCdots in the presence of Cu²⁺ ions exhibits weak FL intensity (Fig. 4B). In contrast, the FL intensity of NCdots shows no obvious change in the presence of the other metal ions which perhaps produce the chromogenic reaction, simultaneously, accompanied by specific chelation between Cu²⁺ ion and NCdots resulting in FL quenching. Inspired by the fact that the fluorescence intensity of NCdots depends on the presence of H₂O₂, a straightforward strategy for H₂O₂ sensing has been explored based on NCdots and Cu²⁺ ion (Fig. 4C and D).



Condition optimization

In order to obtain optimal H₂O₂ sensing, the Cu²⁺ ion concentration was investigated. Fig. 5 shows the alteration of fluorescence of NCdots at various concentrations of Cu²⁺ ions. The relative FL intensity (FL₀/FL) at 390 nm increases in the concentration range of 0.01–0.1 mM and then decreases

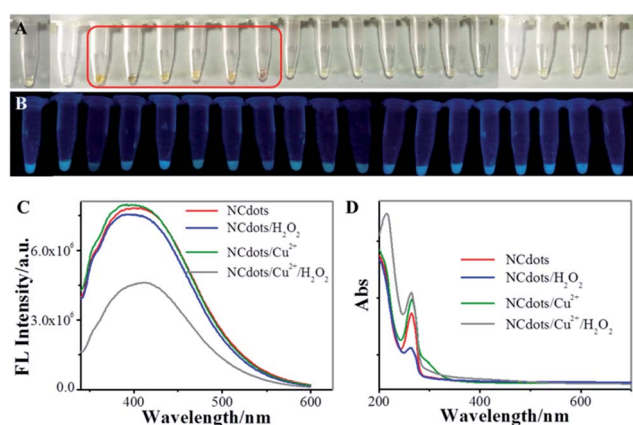


Fig. 4 (A) Images of production of colored products for blank, H₂O₂, and Cu²⁺, Mn²⁺, Co²⁺, Fe²⁺, Fe³⁺, Cr³⁺, Hg²⁺, Mg²⁺, Ba²⁺, K⁺, Zn²⁺, Cd²⁺, Sn²⁺, Ni²⁺, Pb²⁺, Ag⁺ or Na⁺ in the presence of H₂O₂, respectively. (B) Photographs under UV light (365 nm) corresponding to (A). (C) FL spectra of free NCdots and in the presence of H₂O₂, Cu²⁺ ion, and both H₂O₂ and Cu²⁺ ions, respectively. (D) UV-vis spectra of free NCdots and in the presence of H₂O₂, Cu²⁺ ion, and both H₂O₂ and Cu²⁺ ions, respectively.

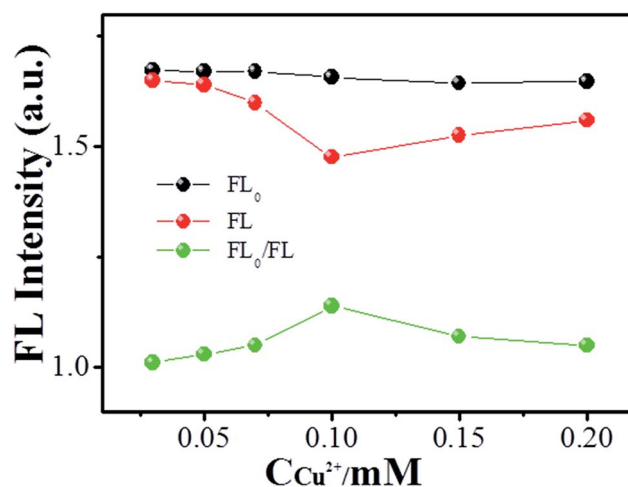


Fig. 5 Influence of Cu²⁺ ions on FL quenching in the concentration range 0.01–0.2 mM Cu²⁺. The black profile is for NCdots alone, the red profile is for NCdots/Cu²⁺ in the presence of H₂O₂, and the green profile shows the ratio of NCdots alone to NCdots/Cu²⁺. Concentration: NCdots: 20 μg mL⁻¹, H₂O₂: 500 mM, Tris-HNO₃ buffer, pH 6.0.



gradually. The reason for this phenomenon may be because there is chelation between OH[•] radicals and Cu²⁺ ions at high concentrations of metal ions and, therefore, the resulting reduced chemical reaction between NCdots and OH[•] radicals will make them soluble in aqueous media with high concentrations of metal ions. That is, Cu²⁺ ions could complex with OH[•] radicals to form the insoluble hydrated oxide Cu(OH)₂ under high concentration conditions, preventing coordination of OH[•] radicals with the NCdots, leading to incomplete FL quenching due to suppression of the alteration of the molecular structure of NCdots. With respect to this sensitivity, a concentration of 0.1 mM of Cu²⁺ ions was selected as an intermediate mediator for H₂O₂ detection.

Sensitive and selective sensing of H₂O₂

On the basis of the above discussion, we explored the feasibility of using NCdots as a FL probe for the detection of H₂O₂ under optimal conditions (Tris-HNO₃ buffer of pH 6.0, Cu²⁺ ion: 0.1 mM), as shown in Fig. 6. Fig. 6A shows the FL responses of NCdots/Cu²⁺ complexes when different concentrations of H₂O₂ were added into the system, and sequential decreases of the FL emission at 390 nm were observed with increasing concentrations of H₂O₂. The H₂O₂ dependence plot (Fig. 6B) of [(FL₀ - FL)/FL₀] (FL₀ and FL are the highest fluorescence intensities of NCdots/Cu²⁺ complexes excited at 330 nm in the absence and presence of H₂O₂, respectively) showed good linearity with H₂O₂ concentration in the range 0.05–300 mM. The limit of detection (LOD) for H₂O₂, at a signal-to-noise ratio of three, was estimated to be 15.89 μM (3δ). Such a high LOD might be attributed to indirect detection rather than direct detection through FL of NCdots.

Considering the promise of the NCdots/Cu²⁺ complex sensor system for application in the biological and environmental fields, the selectivity of the FL sensor for H₂O₂ was evaluated. Under optimal conditions, we repeated the FL intensity changes of the NCdots/Cu²⁺ complexes in the presence of competitive biomolecules five times under the same conditions, respectively, including uric acid (UA), dopamine (DA), ascorbic acid (AA), and glucose (Glu), as shown in Fig. S2† (RSD, 3.85%, *n* = 5). A higher FL quenching was observed for NCdots upon addition of H₂O₂ in the presence of Cu²⁺ ions (0.1 mM), while

the other substances were not able to exert a significant quenching effect. Therefore, these observations suggest that the proposed method is capable of discriminating between H₂O₂ and the interfering biomolecules.

Conclusions

In this study, we have successfully synthesized one type of NCdots through a simple, one-pot, hydrothermal approach, in which EDC was used as the carbon–nitrogen source. We also found that, in the chromogenic reaction, six metal ions play an important role in determining the chromogenic reagent properties of the prepared NCdots. Based on the results, for the first time, NCdots were successfully applied as an artificial peroxidase to detect multi-metal ions and H₂O₂. The green synthesis method of NCdots using EDC was feasible and could have economic benefits.

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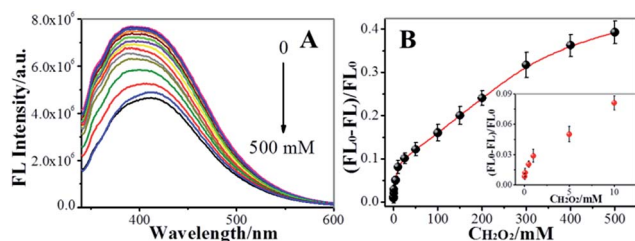


Fig. 6 (A) FL emission spectra of NCdots/Cu²⁺ in the presence of different concentrations of H₂O₂ (0, 0.05, 0.1, 0.5, 1.0, 5.0, 10.0, 25.0, 50.0, 100.0, 150.0, 200.0, 300.0, 400.0, 500.0 mM, top to bottom, excitation at 330 nm). (B) Plot of the enhanced FL signals [(FL₀ - FL)/FL₀] versus H₂O₂ concentration. Concentration: NCdots: 20 μg mL⁻¹, Cu²⁺: 0.1 mM, Tris-HNO₃ buffer, pH 6.0.



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