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A ternary composite with manganese dioxide nanorods and graphene nanoribbons embedded in a polyaniline matrix for high-performance supercapacitors†

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A new ternary composite of manganese dioxide (MnO₂), polyaniline (PANI) and graphene nanoribbons (GNRs) has been fabricated through a two-step polymerization process. With manganese dioxide nanorods and graphene nanoribbons embedded in the PANI matrix, the MnO₂/GNRs/PANI ternary composites with 3-dimensional porous structures exhibited higher specific capacitance and better charge–discharge cycling performance than that of MnO₂ and MnO₂/GNRs binary composites in supercapacitors. The specific capacitance of the MnO₂/PANI/GNRs ternary composite can be achieved at 472 F g⁻¹. After 5000 cycles, the specific capacitance with 79.7% of the initial capacitance can be preserved. The enhanced specific capacitance and cycle life implies the structural stability of MnO₂/GNRs/PANI ternary composites and the synergistic effect between MnO₂, GNRs and PANI.

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1. Introduction

Supercapacitors, also called ultracapacitors, have attracted considerable attention, owing to their high power density, long-cycle performance, optimal operational safety and rapid charge–discharge rates.¹ In general, supercapacitors can be classified into two types depending on the charge storage mechanism and the corresponding active materials were used. Electrical double-layer capacitances (EDLCs), which store charges by reversible ion absorption between the electrode and electrolyte interface, mostly use carbon materials with high accessible surface area.^{2,3} Conventional carbon materials such as activated carbon, carbon nanotubes and carbon fiber usually exhibit good stability due to their special structures and large surface area.^{4–7} Nevertheless, their specific capacitance is limited due to the low theoretical capacity derived from EDLCs.^{8–10} The other type, known as pseudocapacitors, is based on redox reactions that occur largely on material surfaces to store charges.^{11,12} Typical pseudocapacitive materials include transition metal oxides, such as RuO₂, MnO₂, NiO,^{13,14} MoS₂ (ref.

15) and conducting polymers like polyaniline (PANI), polypyrrole, and polythiophene.^{16–18} However, they usually display poor cycle performance with structural deformation during the charge and discharge processes.¹⁹ Undoubtedly, the capacitive performance of supercapacitors depend strongly on the performance of the electrode materials including their structures, surface area, conductivity and electrochemical stability.²⁰ Therefore, increasing interest has been focused on the development of novel composites to enhance its capacitive performance.²¹

Graphene nanoribbons (GNRs), thin elongated strips of graphene with a high length-to-width ratio and straight edges, can be successfully synthesized *via* longitudinal unzipping of MWCNTs by a chemical oxidation route. In recent years, GNRs-based materials have been extensively studied due to the high available surface area, extraordinary mechanical strength and outstanding electrical conductivity.^{22–24} The excellent electronic properties make them attractive especially as supporting materials for the fabrication of composites, and as so far, it has been studied in a wide range of applications such as batteries, supercapacitors and fuel cells.^{25–28}

Nowadays, great interests are shown by researches on pseudocapacitance materials with excellent electrochemical characteristics such as metal oxides and conducting polymers. Among the metal oxides electrode materials, MnO₂ has been studied extensively because of its advantages of low cost, natural abundance, environmental safety and high theoretical specific capacitance. However, the capacitance of MnO₂ is limited by its poor electrical conductivity, which hinders the

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practical application in supercapacitors. On the other hand, among different kind of conducting polymers, PANI have been widely studied due to their controllable electrical conductivity and environmental stability. However, the drawbacks of the PANI electrode is also obvious, that is, a serious problem of typical volumetric swelling and shrinking happens during the charge–discharge process, which severely influence the cycle performance, limiting its practical usage in supercapacitors.

In recent years, binary or ternary composites of carbon materials, conducting polymers and metal oxides have been investigated to make full use of their advantages and overcome their disadvantages of each component. For example, multi-walled carbon nanotube/PANI/MnO₂ ternary coaxial nanostructures,²⁹ graphene/MnO₂/PANI ternary composite,³⁰ manganese ferrite/graphene/PANI nanostructure³¹ have been prepared to further enhance its capacitive performance. These researches have shown that it is an effective way by immobilizing carbon materials with pseudocapacitive conducting polymers and metal oxides. Because of the unique synergistic effects between carbon materials and pseudocapacitance materials may greatly improve the electrochemical properties of the hybrid composite.

In this paper, we report unique MnO₂/GNRs/PANI ternary composites with GNRs and MnO₂ nanorods embedded in PANI matrix. The introduction of PANI can not only prohibit the stacking of GNRs, but also fastens all of them directly interconnected with MnO₂ nanorods to form ideal conductive three-dimensional networks. Therefore, the electron transfers between GNRs and MnO₂ has been improved due to the electrical conductivity of PANI. On the other hand, GNRs provide mechanical structures which significantly reduce the volumetric change during the charge–discharge process. The MnO₂/GNRs/PANI ternary composites exhibit outstanding capacitive performance and excellent cycle performance with a highest specific capacitance of 472 F g⁻¹ at a current density of 1 A g⁻¹ and the capacitive retention of as high as 85% observed after 1000 cycles at a current density of 10 A g⁻¹.

2. Experimental

2.1 Materials

MWCNTs (purity, >95%; length, ~10 μm; diameter, >30 nm) were purchased from Chengdu Organic Chemistry Co., Ltd (China). Potassium permanganate (KMnO₄), manganese(II) chloride tetrahydrate (MnCl₂·4H₂O), ammonium persulfate ((NH₄)₂S₂O₈), aniline, concentrated sulfuric acid (H₂SO₄), and hydrogen peroxide (30% water solution, H₂O₂) were purchased from Shanghai Aladdin Bio-chem Technology Co., Ltd. All chemicals were used as received without further purification. All solutions were prepared with deionized (DI) water.

2.2 Preparation of MnO₂ nanorods³²

In a typical procedure, 0.32 g of KMnO₄ and 0.6 g of MnCl₂·4H₂O were added to 40 mL of deionized water with vigorously stirring for 30 min until formed the homogeneous solution. And then, the precursor mixture solution was further transferred

into a 50 mL of Teflon-lined autoclave and subsequently heated at 160 °C for 12 h. Then, it was cooled to room temperature naturally. The precipitates were obtained by centrifugation and sequentially washing with deionized water for several times, and then dried in vacuum at 80 °C for 6 h. Finally, the brown-colored powder of the MnO₂ nanorods was obtained.

2.3 Preparation of GNRs

GNRs were synthesized according to the method described by Tour *et al.*²² Briefly, 150 mg of MWCNTs was first dispersed in 200 mL of H₂SO₄ by stirring for 1 h. Then, 2.0 g of KMnO₄ was then slowly added into the MWCNTs/H₂SO₄ mixture under a bath of mixture of ice water. Afterward, the mixture was stirred at room temperature for 1 h and then heated to 60 °C for 1 h. After that, the reacted products were pour into 1000 mL of deionized water, and then adding 10 mL of H₂O₂ under ice water bath. The products of GNRs were harvested by centrifugation and washed with 10% (v/v) hydrochloric acid and deionized water for several times, and finally dried at 50 °C.

2.4 Preparation of MnO₂/PANI/GNRs ternary composites

The MnO₂/PANI/GNRs ternary composites were prepared by two-step *in situ* polymerization. 0.1 g of as-prepared GNRs were first added into 100 mL of water and sonicated for 1 h to form a yellow-brown GNRs/water dispersion. By then, 0.12 g of aniline was slowly added into the GNRs/water dispersion under vigorously stirred to obtain the stable GNRs/aniline suspension. After, 10 mL of mixture solution of hydrochloric acid and APS was slowly added into GNRs/aniline dispersion with magnetically stirring. The molar ratio of aniline, hydrochloric acid and APS was 1 : 1 : 1. When the reaction was conducted in ice bath for 10 minutes, 0.6 g of MnO₂ nanorods dispersed in 50 mL of 1 M HCl water solution was added to the above reacting dispersion. Almost simultaneously, 0.12 g of aniline were added again and stirred for 4 h. Finally, the deep-green MnO₂/PANI/GNRs ternary composites were obtained by filtration and washed with deionized water and ethanol in sequence and dried at 60 °C under vacuum.

2.5 Instruments and characterization

The morphologies of the products were investigated by a scanning electron microscope (SEM, JSM-7800F JEOL Co., Japan) and a transmission electron microscope (TEM, JEM-2100 JEOL Co., Japan), respectively. Infrared spectra were recorded on a Varian 3100 FT-IR spectrometer by using pressed KBr pellets. X-ray diffraction (XRD) patterns of the composites were measured on a XRD, Cu Kα radiation, Bruker D8 Advance, Germany. X-ray photoelectron spectroscopy (XPS) analysis was carried out on an ECSALAN 250 electron spectrometer using a monochromatized Al KR X-ray source (1486.71 eV photons).

2.5.1 Electrochemical measurement. The electrochemical properties of the composite were investigated on a CHI660E electrochemical workstation (Shanghai, China) with conventional three-electrode system at room temperature. To prepare the working electrodes, 80 wt% of active materials, 10 wt% acetylene black and 10 wt% PVDF were mixed to form



a homogenous slurry. Then the slurry was pressed onto nickel foam current collectors with the surface area around 1 cm^2 . Finally, the nickel foam was dried at $70 \text{ }^\circ\text{C}$ for 8 h under vacuum. Electrochemical properties of the samples were investigated *via* cyclic voltammetry (CV), galvanostatic charge-discharge (GCD) and electrochemical impedance spectroscopy (EIS) measurements. And the platinum electrode and saturated calomel electrode was used as counter and reference electrodes, respectively. And 1 M Na_2SO_4 solution was used as electrolyte. The potential range for CV tests was 0–0.8 V and the scan rate was $10\text{--}100 \text{ mV s}^{-1}$. The potential range for charge-discharge tests was 0–0.8 V. EIS measurements were carried out in the frequency range of 10^5 to 0.01 Hz at 5 mV amplitude referring to open circuit potential. Galvanostatic charge-discharge tests were conducted on a LAND CT2001A program testing system.

3. Results and discussion

As shown Fig. 1, $\text{MnO}_2/\text{PANI}/\text{GNRs}$ composites were fabricated by two-step polymerization. By step-oxidation, there are lots of oxygen-containing groups, such as epoxy, hydroxyl and carboxyl groups in the structure of GNRs.²⁶ These functional groups, acting as anchor sites, enable the aniline units subsequent *in situ* formation of PANI attaching on the surfaces and edges of GNRs.²⁴ After first step polymerization, since the MnO_2 nanorods can be acted as oxidant reactive template for oxidative polymerization of aniline, the aniline monomers adsorbed on the surfaces of GNRs can be oxidized to PANI polymer around the surfaces of the MnO_2 nanorods, accompanied by the

reduction of MnO_2 to soluble Mn^{2+} ions. Therefore, the $\text{MnO}_2/\text{GNRs}/\text{PANI}$ ternary composites with MnO_2 and GNRs dispersed in the PANI matrix were obtained after second-step polymerization.

FT-IR spectra were first used to characterize the formation of $\text{MnO}_2/\text{PANI}/\text{GNRs}$ composites. From Fig. 2a, regarding FT-IR spectrum of GNRs, there are characteristic peaks located at 1721, 1238 and 1040 cm^{-1} which are attributed to the vibration of carbonyl, epoxy and hydroxyl groups, respectively. The peak of 1610 cm^{-1} is assigned to the stretching vibration of $\text{C}=\text{C}$ group.^{24,26} Towards to PANI, the absorption peaks located at 1570 and 1492 cm^{-1} are ascribed to the $\text{C}=\text{C}$ stretching deformation of the quinoid ring and benzenoid rings in the emeraldine salt. The peaks corresponding to the $\text{C}-\text{N}$ stretching of secondary aromatic amine and $\text{C}=\text{N}$ stretching appear at 1302 and 1152 cm^{-1} , respectively.^{33–35} It is worth noting that the FTIR spectra of $\text{MnO}_2/\text{PANI}/\text{GNRs}$ almost shows similar peaks as PANI, which indicates the PANI were successfully introduced into the $\text{MnO}_2/\text{PANI}/\text{GNRs}$.

The XRD patterns were also demonstrated the formation of $\text{MnO}_2/\text{PANI}/\text{GNRs}$. For MnO_2 nanorods, significant XRD peaks were recorded at $2\theta = 12.98^\circ, 18.24^\circ, 25.72^\circ, 28.86^\circ, 37.66^\circ, 41.90^\circ, 49.80^\circ, 56.28^\circ, 60.32^\circ, 65.60^\circ$ and 69.24° , which could be well-assigned to the (100), (200), (220), (310), (211), (301), (411), (600), (521), (002) and (541) crystal planes of a pure tetragonal the $\alpha\text{-MnO}_2$ phase.^{32,36} For pure polymer of PANI, the diffraction peaks were located at $2\theta = 8.62^\circ, 16.34^\circ, 20.41^\circ, 25.32^\circ$ and 26.92° , illustrating that the *in situ* polymerized PANI are in highly doped emeraldine salt form.^{33,34} As for $\text{MnO}_2/\text{PANI}/\text{GNRs}$, a board peak is centered at $2\theta = 26.92^\circ$, showing similar peaks as PANI except for that the characteristic peaks are becoming obscure, which reveals that the degree of crystallinity of PANI in $\text{MnO}_2/\text{PANI}/\text{GNRs}$ become weak owing to the addition of GNRs and MnO_2 nanorods. Moreover, no obvious diffraction peaks of GNRs were observed in $\text{MnO}_2/\text{PANI}/\text{GNRs}$ indicating that the regular stacks of GNRs are separated by the incorporated MnO_2 and PANI, which is beneficial to its final electrochemical property.

Further, SEM and TEM images provided the structure and morphologies of $\text{MnO}_2/\text{PANI}/\text{GNRs}$. As shown in Fig. 3a, there is no tubulous structure and internal and external walls for GNRs, instead, ribbon-like thin graphene layer with crimped edges.

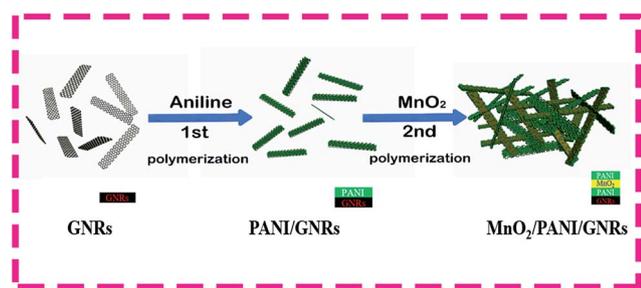


Fig. 1 Schematic illustration of fabrication of $\text{MnO}_2/\text{PANI}/\text{GNRs}$.

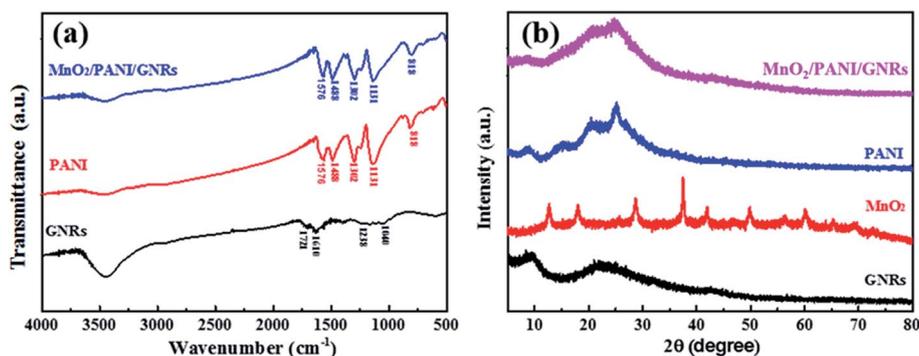


Fig. 2 (a) FT-IR spectra of GNRs, PANI and $\text{MnO}_2/\text{PANI}/\text{GNR}$, (b) XRD of GNRs, MnO_2 , PANI and $\text{MnO}_2/\text{PANI}/\text{GNR}$.



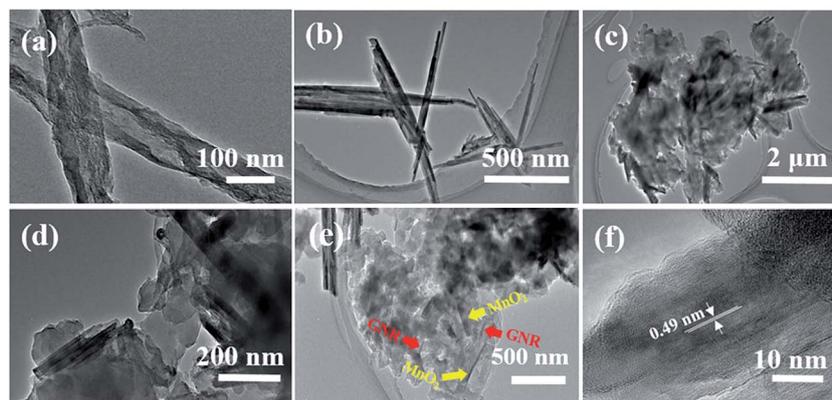


Fig. 3 TEM images of (a) GNRs, (b) MnO₂ nanorods, (c–e) MnO₂/PANI/GNRs composites, (f) HRTEM image of MnO₂/PANI/GNRs composites.

From Fig. 3b, the as-prepared MnO₂ nanoparticles exhibit rod-like structure. The mean width and length for the MnO₂ nanorods are 103.74 nm and 1.95 μm, respectively (Fig. S1 and S2, ESI[†]). After combination of PANI with two-step polymerization, the GNRs and MnO₂ nanorods were filled into the PANI matrix. From Fig. 3c–e, there are distinct MnO₂ nanorods embed in PANI. Due to the ultra-thin ribbon-like structure of GNRs, there is almost no GNRs presented in the SEM images of MnO₂/PANI/GNRs composites. As pointed in Fig. 3e, there are some curved GNRs inserted into the PANI matrix. As shown in Fig. 3f, the MnO₂ nanorods exhibits well crystallized with interplanar distance is 0.49 nm, corresponding to the (200) crystal plane. Brunauer–Emmett–Teller (BET) measurement was also conducted to characterize the structure of MnO₂/PANI/GNRs ternary composite. As shown in Fig. S3a (ESI[†]), the nitrogen adsorption–desorption isotherms of MnO₂/GNRs/PANI ternary composite showed a typical type IV with a hysteresis loop at relative pressure (P/P_0) of 0.4–1, demonstrating the presence of porous structure of MnO₂/GNRs/PANI ternary composite. The Brunauer–Emmett–Teller (BET) specific surface area of the NDPC was estimated as 326.7 m² g⁻¹. The pore size distribution curve of MnO₂/GNRs/PANI ternary composite estimated by Barrett–Joyner–Halenda (BJH), is illustrated in Fig. S3b (ESI[†]). The MnO₂/GNRs/PANI ternary composite is meso-pores with an intensive pore size distribution of ~4.07 nm. The three-dimensional porous MnO₂/GNRs/PANI

ternary composite with high specific surface can accommodate the volume change and facilitate penetration of electrolyte and fast diffusion of ions.^{37–40}

Obviously, as illustrated in Fig. 4a, utilizing the interaction between the GNRs and aniline monomers, the GNRs were first wrapped with PANI polymer chains after first polymerization. Afterward, due to the introduction of MnO₂ nanorods, the structures of composites were turned out to be a 3D porous structure with the GNRs interconnected with MnO₂, together embedding into PANI matrix, which we can believe this unique structure provides an efficient path for electron and ion transport.³⁵ As observed in Fig. 4b–d, the MnO₂/PANI/GNRs show the loose and porous 3D structure with embedding with MnO₂ nanorods and GNRs. The elements mapping images in Fig. 4e–h shows the uniform distributions of MnO₂ and GNRs in PANI matrix.

To further understanding the composition of MnO₂/PANI/GNRs, XPS was carried out and the results are shown in Fig. 5. As for GNRs, the carbon and oxygen signals were observed at 284.6 and 531.9 eV.²⁴ Obviously, after combination, the existences of C, N, O, Mn elements in MnO₂/PANI/GNRs can be confirmed from the wide range as shown in Fig. 5b. The XPS spectra of N 1s can be decomposed into three distinct curves, related to 398.6 eV (–N=), 399.7 eV (–NH–) and 400.1 eV (–N⁺–). More specifically, the peak located at 400.1 eV is assigned to the quinoid amine and nitrogen cationic radical (–N=) and the one

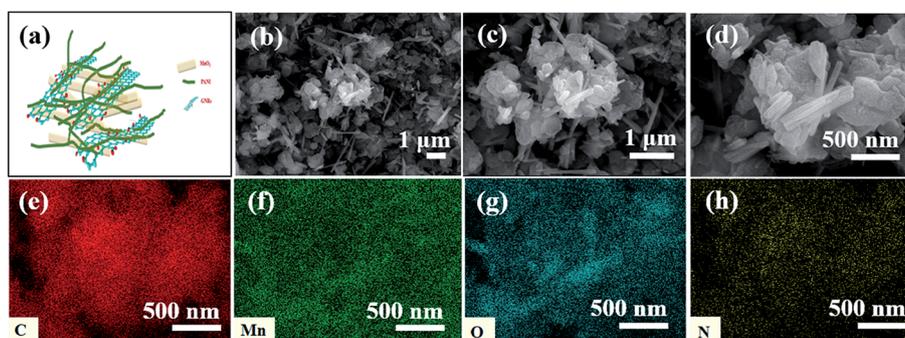


Fig. 4 (a) Structure schematic drawing for MnO₂/PANI/GNRs composites, (b and c) SEM images of MnO₂/PANI/GNRs composites with different magnifications, (d–h) elements mapping for C, Mn, O and N elements of MnO₂/PANI/GNRs composites.



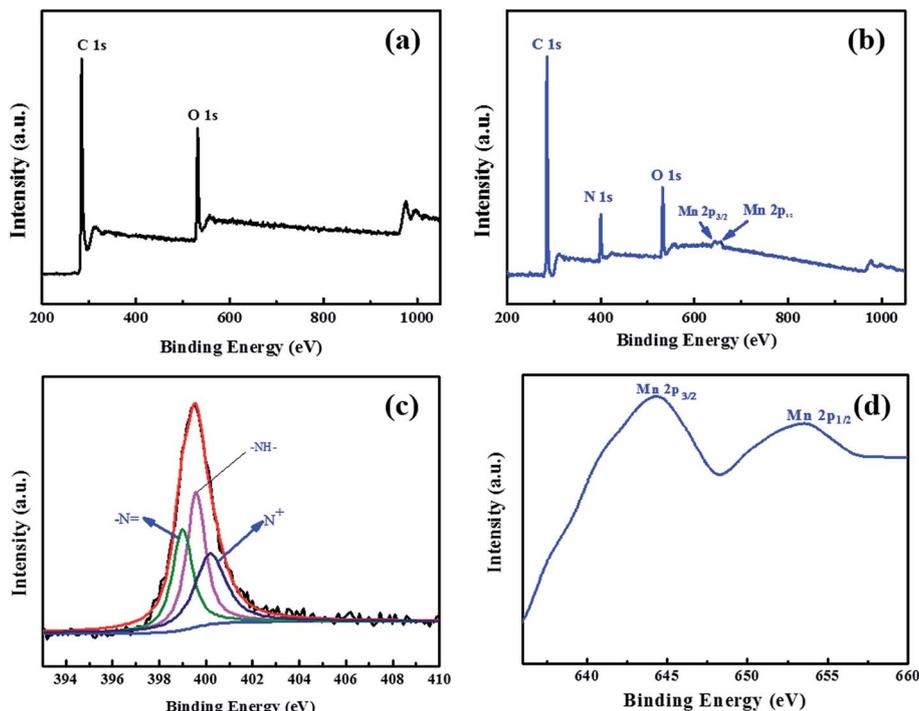


Fig. 5 XPS surveys of (a) GNRs and (b) MnO₂/PANI/GNRs, high-resolution N 1s (c) and Mn 2p (d) core-level XPS spectra of MnO₂/PANI/GNRs.

at 398.6 eV ($-N=$) is due to benzenoid amine.³⁵ The Mn 2p XPS spectrum exhibits two characteristic peaks, Mn 2p_{3/2} at 643.2 eV and Mn 2p_{1/2} peak at 654.8 eV, which are attributed to the spin-orbit peaks of α -MnO₂. The peak-to-peak separation between Mn 2p_{1/2} and Mn 2p_{3/2} level is 11.6 eV, which is approximately the same value as that literature for MnO₂,^{28,32,35} confirming the presence of MnO₂ in composites. From the XPS data of MnO₂/PANI/GNRs composite, the atomic ratios of C, N, O, Mn elements are 68.86%, 8.31%, 10.51% and 12.32%, respectively.

The wrapped PANI with MnO₂ and GNRs can not only offer significant capacitance because of its pseudocapacitance character but can also enhance the contact between GNRs and MnO₂ nanorods, thus, improving the capacitive performance. For understanding the electrochemical property and qualify the specific capacitance of the MnO₂/PANI/GNRs composites, the cyclic voltammograms (CV) measurements of electrode materials for MnO₂, MnO₂/GNRs and MnO₂/PANI/GNRs composites were performed in a 1 M Na₂SO₄ aqueous solution using a three-electrode system. As shown in Fig. 6a, under the electrochemical window from 0 to 0.8 V with a low scan rate of 5 mV s⁻¹, the MnO₂ electrode exhibits rectangular and symmetric CV curves, indicating the ideal pseudocapacitance feature of MnO₂. Incorporating with GNR, the increased area of CV curves is probably due to the combined electric double-layer and pseudocapacitive contribution to the total capacitance.²⁵ After introduction of PANI, the CV curve for MnO₂/PANI/GNRs ternary composites is different from that of MnO₂ and MnO₂/GNRs. There are two pairs redox peaks appeared for MnO₂/PANI/GNRs. It is attributed to the leucoemeraldine-emeraldine transition and the emeraldine-pernigraniline transition. Consequently, the capacitance of MnO₂/PANI/GNRs composites

partially comes from faradaic reactions of PANI at the electrode/electrolyte surface, which is different from that of the electric double-layer capacitance of GNRs and pseudocapacitance of MnO₂.^{35,36} Fig. 6c shows the CV curves of MnO₂/PANI/GNRs at different scan rates. The current density of the MnO₂/PANI/GNRs composites increased with the scan rate and the curve shape is stable. The results demonstrated that the MnO₂/PANI/GNRs has good electrochemical stability. Galvanostatic charge/discharge plots (1 A g⁻¹) of MnO₂, MnO₂/GNRs and MnO₂/PANI/GNRs composites were presented in Fig. 6b. During the charging and discharging steps, all the curves exhibit an equilateral triangle shape, indicating high reversibility of the hybrid materials during charge/discharge process. The specific capacitance can be calculated according to $C_m = I \times \Delta t / (\Delta V \times m)$ from the discharge curves, where C_m is the specific capacitance (F g⁻¹), I is the constant discharge current (A), Δt is the discharge time (s), ΔV is the potential window (V), m is the mass of active material in the working electrode (g).^{32,35} The gravimetric specific capacitances of MnO₂, GNRs/MnO₂ and MnO₂/PANI/GNRs were 136, 245 and 472 F g⁻¹ at the current density of 1 A g⁻¹, respectively. Obviously, the specific capacitance of MnO₂/PANI/GNRs ternary composite is higher than that of MnO₂ and GNRs/MnO₂, indicating the synergistic effect among GNRs, MnO₂ and PANI. From Fig. 6d, the specific capacitance for MnO₂/PANI/GNRs ternary composite at 1, 2, 5 and 10 A g⁻¹ were 472, 385, 347 and 306 F g⁻¹. From Table 1, the specific capacitance for MnO₂/PANI/GNRs ternary composite was higher than those of other reported previous reported GNR-MnO₂ (18%),⁴¹ GNR/MnO₂,⁴² PANI-MnO₂ nanowires,⁴³ PANI/MnO₂,⁴⁴ and PANI-GNRs-40.⁴⁵ From Fig. 7a, generally, the specific capacitance decreases with the increase of current density



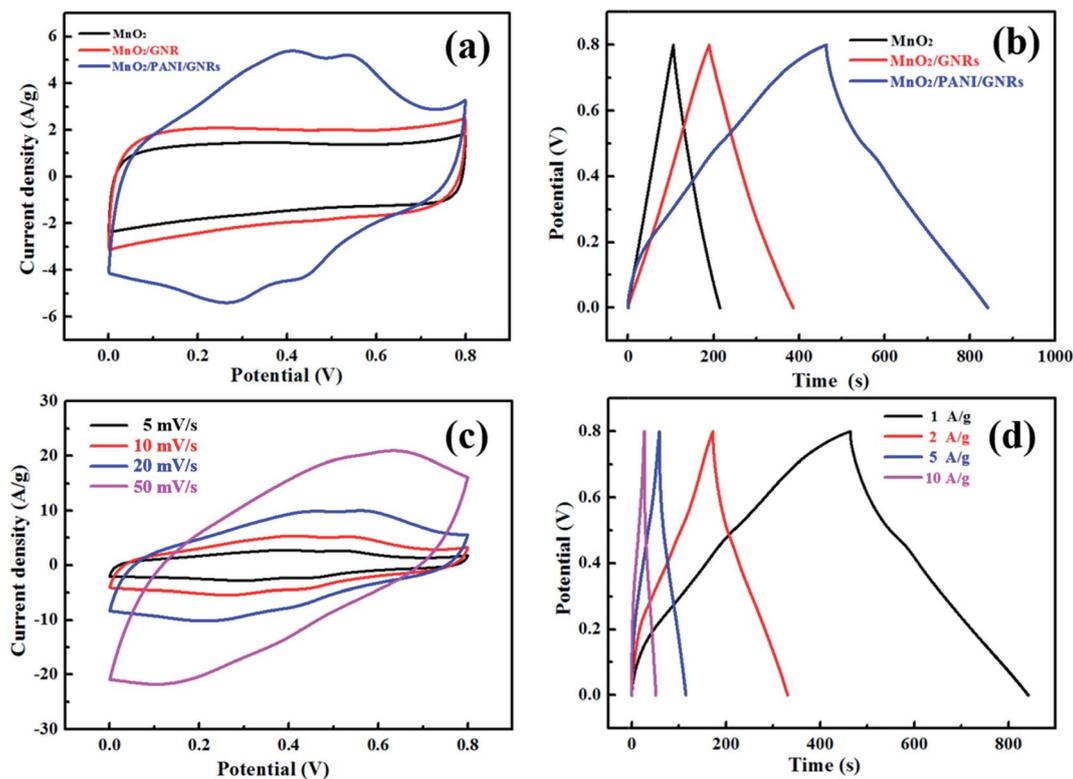


Fig. 6 (a) CV curves of MnO₂, GNRs and MnO₂/PANI/GNRs composites at the scan rate of 5 mV s⁻¹, (b) galvanostatic charge/discharge plots of MnO₂, MnO₂/GNRs and MnO₂/PANI/GNRs composites at a current density of 1 A g⁻¹, (c) CV curves of MnO₂/PANI/GNRs composites at different scan rates and (d) galvanostatic charge/discharge plots of MnO₂/PANI/GNRs composites at different current densities.

Table 1 Comparison of electrochemical performance of the other reported composites involving MnO₂, GNRs, PANI combinations (binary or ternary)^a

Materials	C (F g ⁻¹)	EI	R	T	Ref
GNR-MnO ₂ (18%)	305	PAAK/KCl	0.5 A g ⁻¹	2-electrode GCD	41
GNR/MnO ₂	266	1 M Na ₂ SO ₄	1 A g ⁻¹	3-electrode GCD	42
PANI-MnO ₂ nanowires	256	1 M KOH	1 A g ⁻¹	3-electrode GCD	43
PANI/MnO ₂	320	1 M Na ₂ SO ₄	1 A g ⁻¹	3-electrode GCD	44
PANI-GNRs-40	340	1 M H ₂ SO ₄	0.25 A g ⁻¹	3-electrode GCD	45
CGNR/PANI/MnO ₂	496	1 M H ₂ SO ₄	1 A g ⁻¹	3-electrode GCD	46
MnO ₂ /PANI/GNRs	472	1 M Na ₂ SO ₄	1 A g ⁻¹	3-electrode GCD	This work

^a EI: electrolyte type; R: current density; T: testing method.

because the diffusion of ions from the electrolyte can gain access to almost all available area of the electrode at low current density, therefore, reaching a higher capacitance. As the current density increases, the capacitance decreases because of the effective interaction between the ion and the electrode is greatly reduced and the deep inside the material does not actually contribute to the capacitance.³² The specific capacitance of MnO₂/PANI/GNRs is 472 F g⁻¹ at 1 A g⁻¹ and as high as 306 F g⁻¹ even at 10 A g⁻¹, with a capacitance retention of 65%. While the specific capacitance of MnO₂ and MnO₂/GNRs reaches 136 F g⁻¹ and 245 F g⁻¹ at a current density of 1 A g⁻¹ and remains at 75 F g⁻¹ and 163 F g⁻¹ at 10 A g⁻¹ with a capacitance retention of 55% and 63%, respectively. The MnO₂/PANI/GNRs hybrids

not only exhibit high capacitance but also high stable electrochemical performance at high current densities. This may be because the special composite structure of MnO₂/PANI/GNRs provides numerous porous which can effectively facilitate the ions access to the inside of the composites.^{14,17,35} On the other hand, the synergistic effect between GNRs, MnO₂ and PANI significantly reduce the interval conductivity of the composites. As a result, the MnO₂/PANI/GNRs composites exhibit high capacitance retention at a high current density.

Due to the synergistic effect of GNRs, MnO₂ and PANI, MnO₂/PANI/GNRs also exhibits good cyclic electrochemical performance. As shown in Fig. 7b, the galvanostatic charging-discharging cycles for MnO₂, MnO₂/GNRs and MnO₂/PANI/



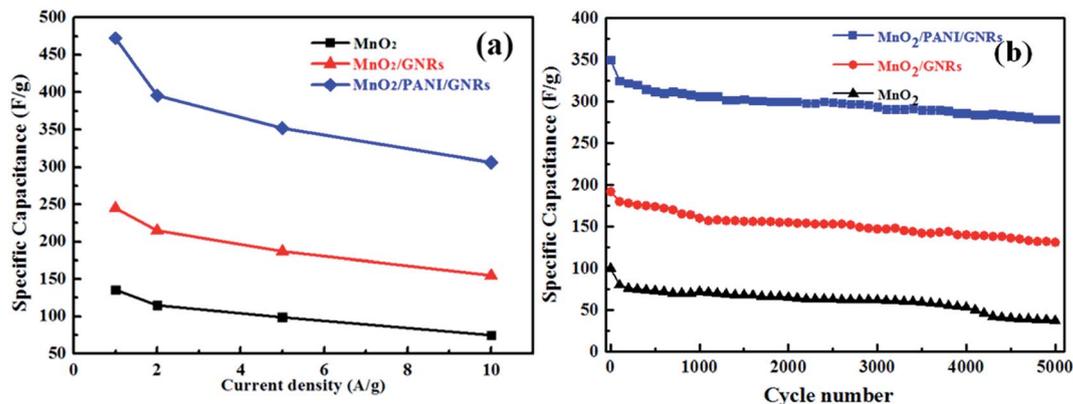


Fig. 7 (a) Specific capacitance versus current density for GNRs, MnO₂ and PANI, MnO₂/PANI/GNRs composites and (b) charge–discharge cycling performances for GNRs, MnO₂ and PANI, MnO₂/PANI/GNRs composites at a current density of 10 A g⁻¹.

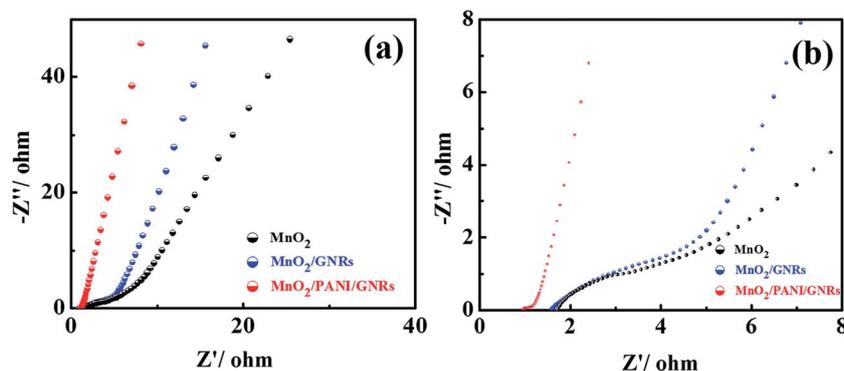


Fig. 8 (a) Nyquist plots of MnO₂, MnO₂/GNRs and MnO₂/PANI/GNRs ternary composites and (b) partial enlarged Nyquist plots of MnO₂, MnO₂/GNRs and MnO₂/PANI/GNRs.

GNRs were carried out at a current density of 10 A g⁻¹ within a voltage range from 0.0 to 0.8 V for 5000 cycles. After 5000 cycles, the MnO₂ and MnO₂/GNRs-based electrode performed only 35.5% and 68.4% capacitance retention after 1000 cycles. But for MnO₂/PANI/GNRs, there is a minor degradation of the specific capacitance with 79.7% of the initial capacitance preserved after 5000 charge–discharge cycles. The improved cycle stability may be derived from the outstanding stability of the 3D structure constituted by MnO₂ nanorods and GNRs, which exhibits excellent mechanical properties. And it effectively prevents the swelling and shrinking of MnO₂/PANI/GNRs nanostructure during the long-term cycle test. In addition, the 3D porous structure can greatly reduce the diffusion and migration length of the electrolyte ions during the fast charge/discharge process, and increases the electrochemical utilization of all the active sites.⁴⁷

To further understand the improvement of electrochemical performance of the MnO₂/PANI/GNRs ternary composite electrode, electrochemical impedance spectroscopy (EIS) was carried out within a frequency range from 100 kHz down to 0.01 Hz. The observed semicircle corresponds to a higher-frequency region, whereas the inclined straight line corresponds to the low-frequency. The depression in the semicircle

is modeled as the parallel combination of an interfacial charge transfer resistance and double-layer capacitance.⁴⁸ The inclined straight line at low-frequency represents the ion diffusion and transport due to the pseudocapacitance.⁴⁹ Fig. 8 represents the Nyquist plots for MnO₂, MnO₂/GNRs and MnO₂/PANI/GNRs electrodes. It is known that the charge transfer resistance (R_{ct}) can be directly measured by the semicircular diameter.¹¹ It can be observed that MnO₂/PANI/GNRs exhibits a smaller arc diameter than MnO₂ and MnO₂/GNRs, indicating that MnO₂/PANI/GNRs has low charge transfer resistance with faster ion diffusion rate in Na₂SO₄ solution. Compared to MnO₂/PANI/GNRs, the MnO₂/GNRs with a larger arc diameter exhibited slower diffusion rate due to the introduction of GNRs. While the MnO₂ presents a larger arc diameter, which might be attributed to the poor conductivity of MnO₂. The lowest charge transfer resistance in ternary MnO₂/PANI/GNRs implied the excellent rate capability. In addition, in the low frequency region, the slope of the curve shows the Warburg impedance, representing the diffusion and transport of ions in the electrode.⁴⁷ MnO₂/PANI/GNRs shows the most vertical line at lower frequencies, further indicating its high capacitive behavior.



4. Conclusion

In summary, unique MnO₂/PANI/GNRs composite have been successfully synthesized *via* a two-step *in situ* process. The as-prepared composites electrode exhibits an excellent specific capacitance of 472 F g⁻¹ at the current density of 1 A g⁻¹ in 1 M Na₂SO₄ aqueous solution. In addition, the MnO₂/PANI/GNRs show high electrochemical stability in cycling stability with 79.7% capacitance retention after 5000 cycles. Such intriguing electrochemical performance was because of the unique 3D porous structure constituted by MnO₂ and GNRs, leading to an excellent rate performance and cycling stability. On the other hand, the synergistic effects of GNRs, PANI and MnO₂ have overcome the disadvantages of each component to some extent, therefore, resulting in a robust capacitive performance. The facile and effective synthetic strategy, we believe, can not only be employed for the fabrication of robust high-performance energy storage material containing carbon materials, metal oxide and conducting polymers, but also can be extended to other material preparation for the development of advanced electrode materials, which may pave an avenue for the practical applications of supercapacitors.

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