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Photosensitive bent-core liquid crystals based on methyl substituted 3-hydroxybenzoic acid†

Michal Kohout,^a Mohamed Alaasar,^{b,c} Anna Poryvai,^a Vladimíra Novotná,^d Silvio Poppe,^b Carsten Tschierske^b and Jiří Svoboda^a

Photosensitive liquid crystals are of contemporary interest not only from the scientific point of view but also for various applications. Herein we report the first photosensitive materials based on laterally substituted 3-hydroxybenzoic acid. The molecular self-assembly of these new materials was characterized using different physical techniques including differential scanning calorimetry (DSC), X-ray diffraction (XRD) measurements, electro-optical investigations and dielectric spectroscopy. We show that the number and position of azo units with respect to the lateral substitution on the molecular core significantly influence mesomorphic properties of the materials. Based on the position of the substituent, either non-polar or polar smectic C phases are formed. The optimum combination of both the structural elements results in an achiral material which shows a broad polymorphism and exhibits a stable dark-conglomerate crystalline phase with macroscopically chiral domains at room temperature. The structure of this phase differs from the previously described fluid sponge-like distorted smectic phases as well as from helical nano-filament phases, thus, adding new information on the possible variations in the nanostructures of this kind of phase. Moreover, the photosensitivity of these materials has been studied using ¹H NMR spectroscopy.

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Introduction

Since the discovery of polar order and macroscopic phase chirality of achiral bent molecules,^{1,2} bent-core liquid crystals (BCLCs) have become prominent materials of interest for materials science and supramolecular chemistry. Their importance lies not only in understanding the fundamentals of self-assembly in soft matter,^{3–5} but also in practical applications in non-linear optics,⁶ sensors⁷ and fast switching electro-optical devices.⁸ Nowadays, the main focus is put on advanced materials, whose properties can be tuned by different external stimuli, in particular, light. Although the first photosensitive BCLCs were prepared as early as in 1929,⁹ their application potential in areas such as optical storage,¹⁰ holographic media¹¹ and photo-alignment of LC matrixes¹² was only recognized much later. Therefore, many different azo group-containing BCLCs have recently been synthesised and their properties studied, as shown in a recent review.¹³

One of the most successful motifs in the search for light tuneable BCLCs represents resorcinol central core, which is extended with azo group-containing elongating side arms. First materials based on non-substituted resorcinol central core were either crystalline or exhibited only thermally unstable columnar mesophases.^{14–16} The real breakthrough in the area of photo-sensitive BCLCs has arrived with the introduction of a lateral substituent on the resorcinol central core. It has been shown that a cyano group in position four of the central resorcinol ring triggers the formation of SmC_SP_R (synclinal tilted smectic C phase with randomized polar order), which is capable of mirror symmetry breaking.^{17–19} Using a halogen as the substituent, a variety of dark conglomerate (DC) mesophases has been observed, in particular, for larger substituents such as bromine and iodine.^{20–22} Replacing one azo group in the molecular structure with an ester unit provided materials whose mesomorphic behaviour depended on the orientation of the ester unit as well as on the position of the halogen substituent with respect to the remaining azo group.²³ Apart from the polar electron-withdrawing substituents, also a weak electron donating methyl group has been introduced on the resorcinol central core to form materials with stable DC phases.²⁴

Unlike the resorcinol derivatives mentioned above, compounds based on 3-hydroxybenzoic acid have been studied much less. The majority of the reported materials possessed one azo group in the molecular structure and exhibited columnar B₁, lamellar B₂ or a modulated smectic B₇-type mesophases.^{14,25,26} Related materials with two azo groups have been found to be

^aDepartment of Organic Chemistry, University of Chemistry and Technology Prague, 166 28 Prague 6, Czech Republic. E-mail: michal.kohout@vscht.cz

^bInstitute of Chemistry, Martin Luther University Halle-Wittenberg, Kurt-Mothes Str. 2, Halle-Saale, D-06120, Germany

^cDepartment of Chemistry, Faculty of Science, Cairo University, Giza, Egypt

^dInstitute of Physics, Academy of Science of the Czech Republic, Na Slovance 2, 182 21 Prague 9, Czech Republic

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non-mesogenic.¹⁵ Despite the fact that BCLCs based on 3-hydroxybenzoic acid substituted in position four with all-ester side arms provided thermally more stable mesophases than the corresponding resorcinol-based analogues,²⁷ no material based on laterally substituted 3-hydroxybenzoic acid has been reported yet. Additionally, the intrinsic non-symmetry of 3-hydroxybenzoic acid central core determines different effect of the lateral substituent on the central core. Recently, it has been shown that the presence of a larger substituent in position six of the central 3-hydroxybenzoic acid unit reduces the flexibility of derived BCLCs leading to crystalline substances.²⁸ On the other hand, the presence of a larger substituent and the reduced conformational flexibility has recently been recognized as an important feature for BCLCs forming DC phases.^{13,21,23,24,29}

In this contribution, we report the effect of the lateral substituent and its location on the mesomorphic properties of the new class of photosensitive BCLCs possessing one or two azo groups in their structure. We have also varied the position of the substituent with respect to the azobenzene wing, which results in a material exhibiting rich polymorphism.

Experimental

Synthesis

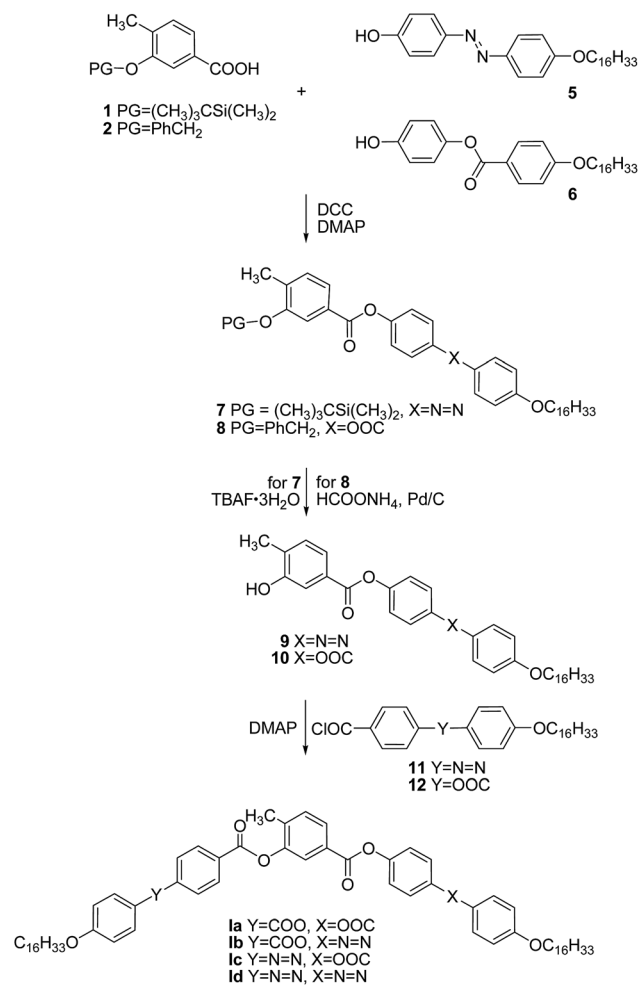
The synthesis of the new BCLCs under investigations is shown in Schemes 1 and 2. The protected central units 1–4 were prepared from commercially available starting materials in accord with previously reported methods.^{27,28} The elongating side arms were synthesised by known synthetic procedures.^{18,30,31}

A *N,N'*-dicyclohexylcarbodiimide (DCC)-mediated esterification of protected acids 1 and 2 with corresponding phenols 5 and 6 in the presence of *N,N*-dimethylaminopyridine (DMAP) as catalyst gave rise to protected derivatives 7 and 8, respectively (Scheme 1). Based on the character of the silyl (PG = (CH₃)₃CSi(CH₃)₂) and benzyl (PG = C₆H₅CH₂) protecting groups (PG), deprotection was achieved by tetrabutylammonium fluoride trihydrate (TBAF·3H₂O) and transfer-hydrogenation with ammonium formate in the presence of palladium on charcoal, respectively. In the final step, the isolated hydroxy esters 9 and 10 were acylated with acid chlorides 11 and 12 of the second lengthening arm to yield the target materials of the series I.

Similarly as above, the protected central cores 3 and 4 were esterified in a DCC-mediated reaction catalyzed with DMAP to yield protected intermediates 13 and 14. Subsequent deprotection performed with respect to the used protecting groups provided hydroxy esters 15 and 16, acylation of which afforded the target materials of the series II (Scheme 2). The analytical data and the synthetic details for the intermediates as well as for the final BCLCs are summarized in the ESI† file.

Experimental methods and set-up

The phase transition temperatures and corresponding enthalpies were determined by differential scanning calorimetry (DSC), namely Pyris Diamond Perkin-Elmer 7 calorimeter was utilized. The samples of about 2–5 mg were sealed into aluminium pans, which were put into the calorimeter chamber filled with nitrogen.



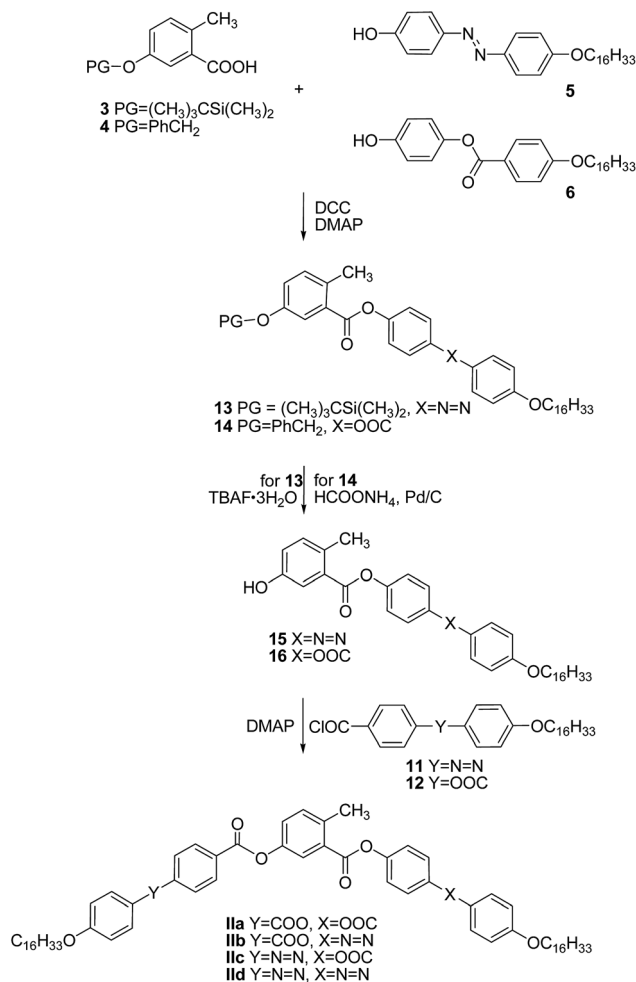
Scheme 1 Synthesis and designation of compounds of series I.

The temperature and enthalpy values were calibrated on the extrapolated enthalpy values of the melting points for water, indium and zinc. All calorimetric measurements were performed on cooling/heating runs at a rate of 5 K min⁻¹.

The textures were observed under the polarizing optical microscope Nikon Eclipse E600. The cells for texture observation and electro-optical studies were made from glasses with ITO transparent electrodes (25 mm²), separated by mylar sheets defining the cell thickness. No glass-surface treatment was utilized. The cells (usually 3 μm thick) were filled with studied compounds in the isotropic phase by the capillary action. Another sample can be prepared by spreading the droplet in the isotropic phase on the glass surface (one-free-surface sample). The Linkam LTS E350 heating/cooling stage with TMS 93 temperature programmer was used for the temperature control. The stabilization of the temperature was within ±0.1 K. The electro-optical switching characteristics were studied with the triangular-wave method using a home-made set-up in 6 μm polyimide coated ITO cells, EHC Japan. Voltage of 100 V_{pp} at a frequency 10 Hz and resistor 5 kΩ were utilized.

Dielectric properties were studied using Schlumberger 1260 impedance analyser. Measurements were performed on cooling and the temperature was stabilized during the frequency





Scheme 2 Synthesis and designation of compounds of series II.

sweeps in the range of 10 Hz to 1 MHz. The frequency dependent complex permittivity $\epsilon^* = \epsilon' - i\epsilon''$, namely the real, ϵ' , and the imaginary, ϵ'' , parts were detected in a broad temperature range on cooling from the isotropic phase.

X-ray diffraction patterns were recorded with a 2D detector (Vantec 500, Bruker). Ni filtered and pin hole collimated CuK α radiation was used. The exposure time was 30 min and the sample to detector distance was 8.95 cm and 26.7 cm respectively. Uniform orientation was achieved by alignment at the air-sample interface on top of a small droplet. The samples were held on a temperature-controlled heating stage.

Photochemical properties were studied using Agilent 400 MR DDR2 spectrometer. The selected compounds were dissolved in deuteriochloroform and irradiated at 365 nm with a common laboratory UV lamp for 8 h. The samples were placed into an NMR instrument tempered at 50 °C and the spectra were acquired with 700 s steps between measurements, duration of each measurement was 57 s.

Results and discussion

Based on the position of lateral substituent (methyl group) on the central core, the materials were divided into two series: 4-

substituted compounds are denoted as series **I** and 6-substituted compounds are denoted as series **II** (see Schemes 1 and 2). The mesomorphic behaviour of the studied materials was determined from thermal behaviour obtained by DSC, texture observation under the polarizing optical microscope and electro-optical behaviour under the applied electric field (Table 1). To support the phase identification, X-ray measurements were performed for a selected example.

Series I

Compound **Ia** with all-ester linkages exhibited a monotropic SmC_s phase. The introduction of an azo group in the elongating side arm connected to the carboxylic unit of the central core (**Ib**) led to an interesting polymorphism. On cooling, a sequence of a nematic phase, three consecutive SmC phases and a DC phase were observed. We denoted these synclincic smectic C phases as the SmC_{s1}, SmC_{s2} and SmC_{s3}, respectively, on cooling. The DC phase is stable down to room temperature and crystallises on subsequent heating. On DSC thermograph the crystallisation of the DC phase takes place at about 85 °C and is present in a form of the opposite peak on the second heating curve (Fig. 1a). On further heating, the crystal phase (Cr) transforms to the SmC at 102 °C and at 107 °C the phase transition to the isotropic phase takes place. We will describe all observed mesophases later. On changing the positions of the azo and ester linking group in compound **Ic**, the mesomorphic behaviour completely disappeared. For compound **Id** with the azo group in both the elongating side arms, a sequence of two monotropic mesophases has been observed on cooling from the isotropic liquid. Namely, a narrow temperature interval of a nematic phase is followed by a smectic C phase (Fig. 2).

Textures and their features under the applied electric field and at various temperatures were observed using a polarizing optical microscope. For compound **Id**, a very narrow nematic phase is observed above the tilted smectic phase (Fig. 2a). For all studied compounds from series **I**, the observed smectic phases reveal broken fan-shaped texture (for illustration see Fig. 2b for **Id**). The extinction position of all smectic mesophases is inclined from the layer normal by an angle of about 45 degrees, which provides evidence for the synclincic character of SmC phases. For compound **Ib**, we have observed a sequence of three SmC_s mesophases and their planar textures are differing only slightly (see Fig. S1 in ESI†). On the other hand, homeotropic textures are not so uniform and reveal specific features for various SmC_s phases of **Ib**. In Fig. 3a a phase transition from nematic to smectic phase is shown and we can observe “transition bars”, which for calamitic LC compounds can accompany N-SmC phase transition.³² Fig. 3b and c show the schlieren texture in the SmC_{s1} and the SmC_{s2} phase, respectively. The lines appearing in the SmC_{s2} phase can be connected with undulations or defects arising in smectic planes. In Fig. 3d the texture in the SmC_{s3} phase is depicted.

For all compounds of series **I**, no response to the applied electric field was obtained in the observed SmC_s mesophases, which provided further evidence on their non-polar character. Due to the narrow range of the SmC_{s3} phase exhibited by **Ib**, we



Table 1 Mesomorphic properties of materials of series I and II^a

Comp.	X	Y	Heating	Cooling
Ia	COO	OOC	Cr 102 [38.0] Iso	Iso 97 [12.4] SmC _s 95 [21.9] Cr
Ib	COO	N=N	DC 85 [26.3] Cr 102 [66.8] SmC _{s1} 107 [5.9] Iso	Iso 106 [0.92] N 105 [4.8] SmC _{s1} 92 [0.11] SmC _{s2} 82 [6.2] SmC _{s3} 80 [27.8] DC
Ic	N=N	OOC	Cr 115 [60.4] Iso	Iso 108 [59.4] Cr
Id	N=N	N=N	Cr 116 [91.7] Iso	Iso 117 [*] N 116 [6.5] SmC _s 107 [49.8] Cr
IIa	COO	OOC	Cr 107 [41.2] Iso	Iso 105 [24.4] SmC _a P _A 96 [13.5] Cr
IIb	COO	N=N	Cr 120 [38.3] Iso	Iso 116 [40.9] Cr
IIc	N=N	OOC	Cr 102 [75.6] Iso	Iso 100 [21.9] SmC _a P _A 83 [59.6] Cr
IId	N=N	N=N	Cr 116 [61.6] Iso	Iso 111 [55.7] Cr

^a The phase transition were taken from the second heating and cooling runs at a rate of 5 K min⁻¹. Phases are abbreviated: SmC_s – synclinc smectic phase; SmC_aP_A – anticlinc antiferroelectric SmC phase; Iso – isotropic liquid; Cr – crystalline phase; DC – dark conglomerate phase. * – Iso-N enthalpy could not be separated from N-SmC_s enthalpy. Temperature is given in °C and phase transition enthalpy in kJ mol⁻¹.

were not able to determine its response to the applied electric field. However, the textures of all three consecutive SmC_s phases are very similar and thus we assume that also the low temperature SmC_{s3} phase is of non-polar character. This result has further been supported by the dielectric spectroscopy measurements, see below.

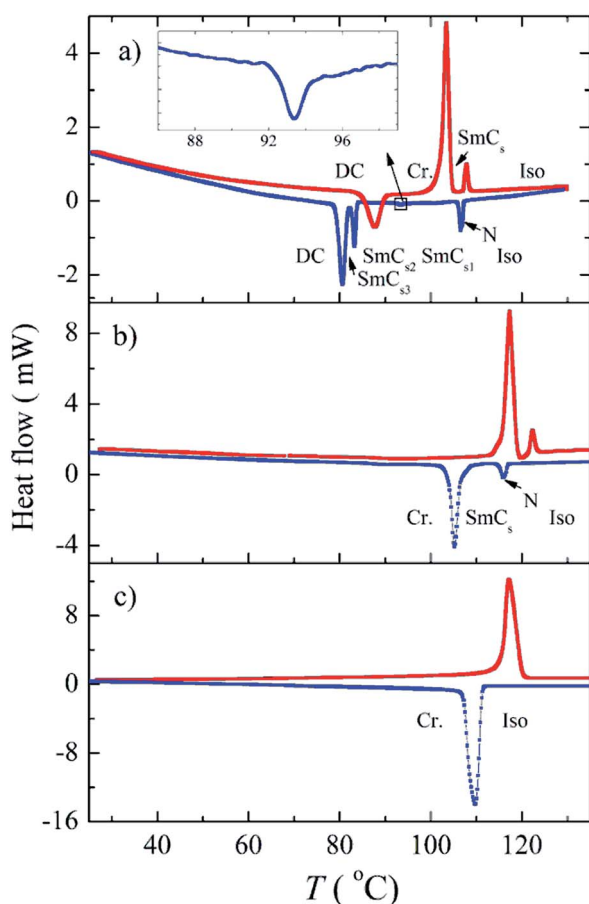


Fig. 1 DSC thermographs for compounds (a) **Ib**, (b) **Id** and (c) **IId**. The inset in the figure (a) shows the phase transition from SmC_{s1} to SmC_{s2} taken on cooling, marked with an arrow. The red upper curves show the second heating run, the lower blue lines show the second cooling, taken at a rate of 5 K min⁻¹.

In case of compound **Ib**, on further cooling from the SmC_{s3} phase, a highly viscous optically isotropic phase appears (Fig. 4b). Rotating the analyzer by a small angle leads to the appearance of dark and bright domains. Rotating the analyzer in the opposite direction reverses the dark and bright domains (Fig. 4a and c), while rotating the sample itself under crossed polarizers does not lead to any change in the dark texture. The overall distribution of the dark and bright domains is about 1 : 1 in both cases. This indicates that the domains are chiral with opposite handedness, which is a typical feature of chiral DC phases.¹³

The dielectric properties of **Ib** were investigated within the broad temperature interval of all mesophases. Fig. 5 shows the real, ϵ' , and imaginary, ϵ'' , part of permittivity with respect to the temperature, T , and frequency. There is no distinct mode, which can be attributed to the polar arrangement. The values of permittivity are rather small, which support the idea of the non-polar character of the observed SmC_s mesophases. We can identify the phase transition temperatures, which are connected with small step-like changes of the real part of permittivity, ϵ' . On the other hand, the imaginary part of permittivity reveals anomalies at the phase transition temperatures. In Fig. 5 these anomalies are marked by arrows. We can speculate that it corresponds to director fluctuations, which are more intensive in the vicinity of the phase transition. We can expect that in the

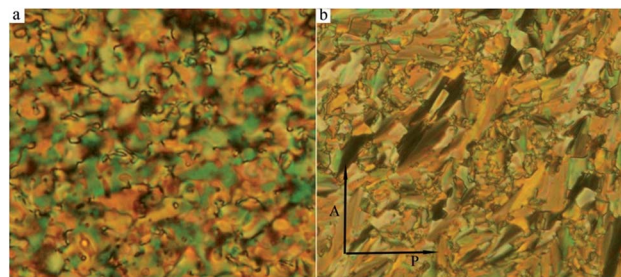


Fig. 2 Planar texture of compound **Id** at different temperature observed by polarizing microscope (a) in the N phase at $T = 116$ °C and (b) in the SmC_s phase at $T = 110$ °C. The width of the figure corresponds to 200 μ m. Orientation of the crossed polarisers (A, P) is marked.



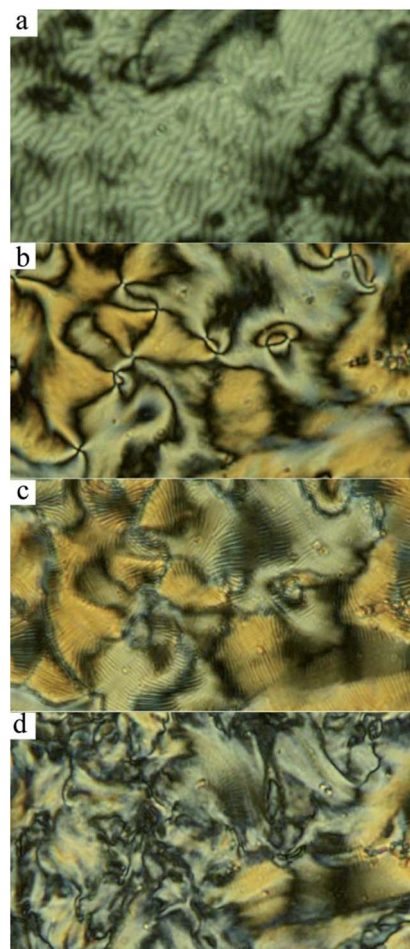


Fig. 3 Homeotropic texture for compound **1b** (a) at the N-SmC_{s1} phase transition at $T = 105$ °C, (b) in SmC_{s1} phase at $T = 100$ °C, (c) in the SmC_{s2} phase at $T = 86$ °C and (d) in the SmC_{s3} phase at $T = 81$ °C. The width of figures corresponds to 250 μm .

observed non-polar SmC_s mesophase the bent-core molecules freely rotate along their molecular axis with rather high frequency and this molecular mode contributes mostly at the phase transitions. The values of permittivity are rather small, which support the idea of the non-polar character of the observed SmC_s mesophases. On the other hand, we can identify the phase transitions because they are connected with molecular fluctuations. We can speculate that the bent-core molecules freely rotate along their molecular axis and this motion contributes mostly at the phase transitions.

X-ray scattering measurements were conducted for compound **1b** as it shows the interesting phase sequence of nematic and different SmC phases ending with the DC phase. Due to the very narrow range of its existence ~ 1 K, it was not possible to carry out XRD investigations of the nematic phase. However, within all smectic phases, the X-ray intensity exhibited a sharp peak at small scattering angles and a diffuse maximum in the wide angle region (Fig. 6a for **1b** in the SmC_{s1} phase). The layer spacing, d , can be established from the position of the intensity maximum at small scattering angles. The diffuse scattering peak at wide angles corresponds to the

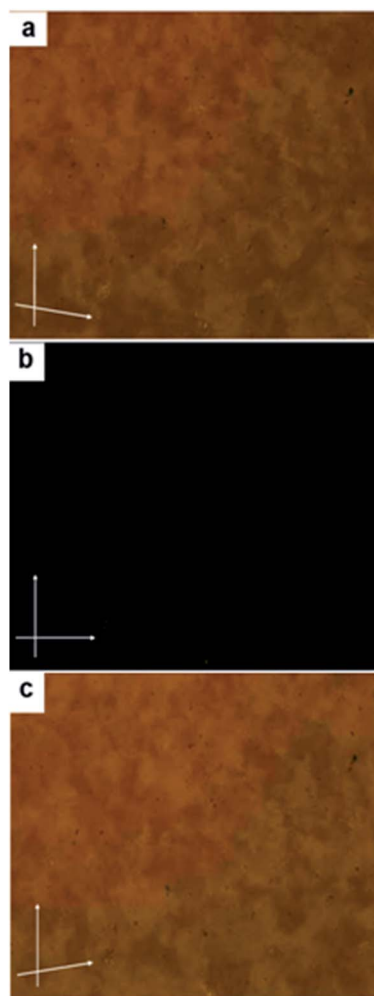


Fig. 4 Textures of the DC phase of compound **1b** at $T = 50$ °C: (a) after rotating the analyzer by 10° from the crossed position with respect to the polarizer in clock-wise direction; (b) under crossed polarizers and (c) in anticlockwise direction, showing dark and bright domains, indicating the presence of areas with opposite chirality sense.

average distance between molecules ≈ 4.5 Å. For compound **1b** the corresponding d value has been found $d = 47.8$ Å at $T = 100$ °C, which is substantially smaller than the molecular length, l , calculated in *ab initio* calculations $l = 56.3$ Å. Comparing X-ray data with calculations, the tilt angle can be estimated to be about 32 degrees. We confirmed a lamellar character of the SmC mesophases observed on cooling from the isotropic phase for other compounds. Due to the monotropic character of the mesophases and their fast crystallisation, it was not possible to finish detailed X-ray measurements.

X-ray investigations in the DC phase of **1b** (Fig. 6b) show an intense layer signal with corresponding weak harmonic reflections in small and medium angle region. All scatterings form closed rings with uniform intensity distribution, which provides evidence for the disordered meso-structure of the DC phase. The observed layer thickness points to tilted organization within layer with a molecular tilt angle of about 29° . Such a value is typical neither for DC sponge phases nor for helical-nanofilament phase (B₄ phase). Furthermore, the broad signal



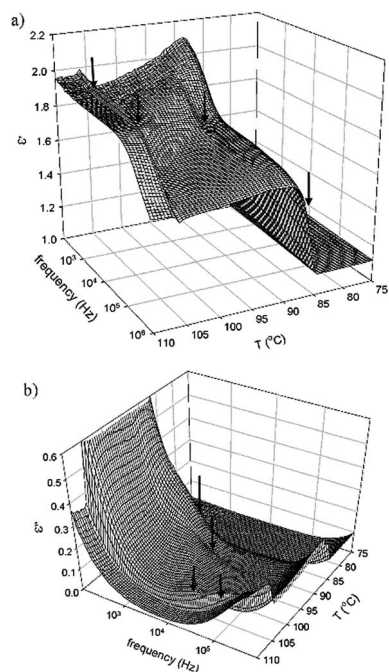


Fig. 5 3-Dimensional plots of the real, ϵ' , and imaginary, ϵ'' , part of permittivity with respect to the temperature, T , and frequency for **1b**. The arrows mark phase transition.

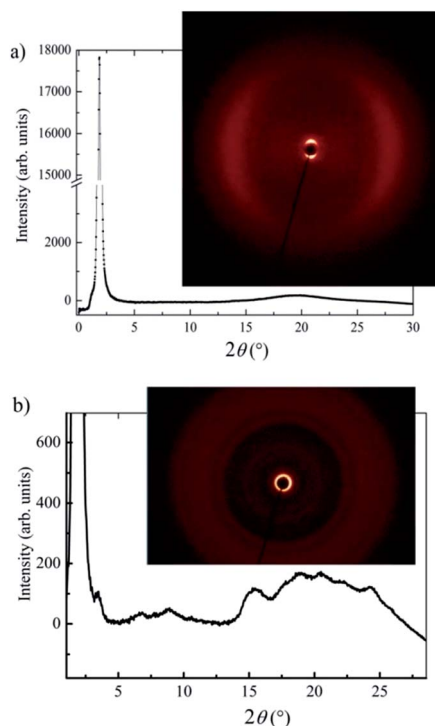


Fig. 6 X-ray intensity versus scattering angle for compound **1b**: (a) at $T = 100\text{ }^{\circ}\text{C}$ in the SmC_{s1} phase and (b) at $T = 70\text{ }^{\circ}\text{C}$ in the DC phase. The insets show the corresponding 2D X-ray patterns at the indicated temperature.

in the wide angle region could be fitted to six maxima in the 2θ range between 15 and 26° (Fig. 6b). This pattern clearly excludes fluid sponge phases that show only a very diffuse wide angle

scattering. However, such reflections can not be ascribed to a B_4 phase, which typically shows the reflection as sharp separate signals. These results indicate that the DC phase under investigation is similar to the recently described new sub-type of a DC phase exhibited by related 4-methylresorcinol-based materials with one inverted ester.²⁴ The reflections in small and medium region appear at very similar positions. The wide angle region shows more diffuse pattern with two more maxima, which could arise from slightly different molecular packing within the studied DC phase. Therefore, inverting the direction of the ester group retains the DC phase but slightly change its fine structure.

Series II

The mesomorphic properties of series **II** differ from those of series **I**. This change probably results from the different position of the lateral methyl group at the apex of the bent-core mesogen. In case of compound **IIa** with all ester linkages, a monotropic SmC_aP_A phase was identified (see Fig. 7d). Surprisingly, the introduction of the azo group in the elongating side arm connected to the carboxylic unit of the central core (**IIb**) led to the complete loss of mesomorphic behaviour. When changing the positions of the azo and ester linking groups (**IIc**), the monotropic SmC_aP_A phase was restored. Compound **IIId** possessing two azo units in the structure was crystalline only.

Compounds **IIa** and **IIc** showed two peaks at half-period under the applied triangular field (Fig. 7a and b). The switching current documents the antiferroelectric nature of the studied mesophases. Based on this observation, the mesophase for both **IIa** and **IIc** was assigned as SmC_aP_A . Due to relatively fast crystallisation of the materials under the applied electric field, we were able to record only changes of planar textures of **IIa** (Fig. 7c–e).

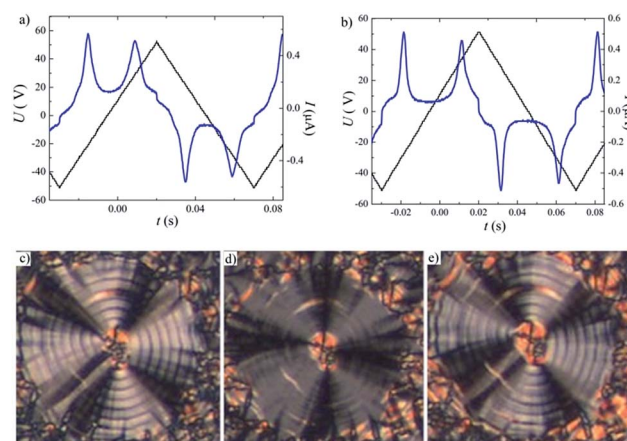


Fig. 7 Switching current response curves recorded by applying a triangular wave voltage in the SmC_aP_A phase for compounds: (a) **IIa** at $103\text{ }^{\circ}\text{C}$ and (b) **IIc** $99\text{ }^{\circ}\text{C}$; (c–e) optical textures in the SmC_aP_A phase for compound **IIa** at $102\text{ }^{\circ}\text{C}$ under an applied DC voltage between crossed polarizers in (c) field induced SmC_sP_F state at $+10\text{ V}$, (d) in the SmC_aP_A state at $V = 0$ after switching off the field and (e) the field induced SmC_sP_F state at -10 V .



The dielectric spectroscopy for compound **IIa** and **IIc** has been performed and a distinct mode observed in the SmC_aP_A phase on cooling from the isotropic phase. 3-Dimensional plot of the real and imaginary parts of the permittivity for **IIa** is illustrated in Fig. 8. The observed mode completely disappears in the isotropic as well as in the crystalline phase, so we can conclude that it is connected with collective fluctuations of molecules. Such mode is very often observed for SmCP phases and is explained by vibrations of tilted molecules in polar ordered systems with interactions between neighbouring layers.³³

Photochemical properties

The presence of the azo units in the investigated BCLCs allows the study of photoisomerisation under UV light irradiation. This phenomenon has been reported before for similar azobenzene containing BCLCs in solutions using UV-Vis spectrometry.¹³ However, herein we study this process using ^1H NMR spectroscopy. This method allows us to determine the amount of *Z*-isomer formed upon UV irradiation of the predominant *E*-isomer of an azo group containing materials. Moreover, this method is suitable for the observation of the thermal relaxation of the formed *Z*-isomer back to the more stable *E*-isomer in the dark. It is evident that with the increasing number of azo units present in a molecule, the ^1H NMR spectrum of such a compound gains in complexity. In case of two azo groups, the formation of four different chemical species can

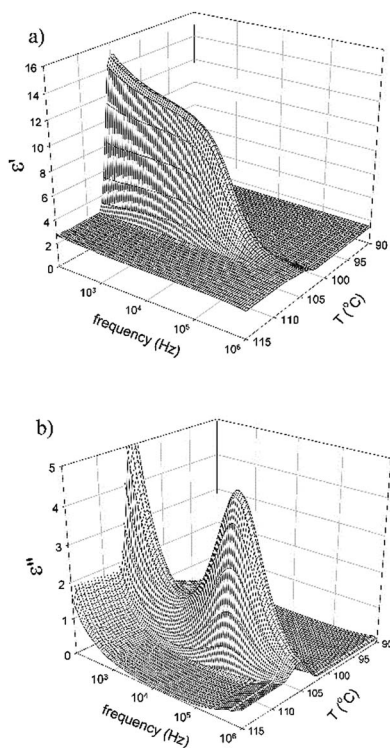


Fig. 8 3-Dimensional plot of (a) the real part of permittivity, ϵ' , and (b) the imaginary part of permittivity, ϵ'' , versus temperature, T , and frequency for **IIa**.

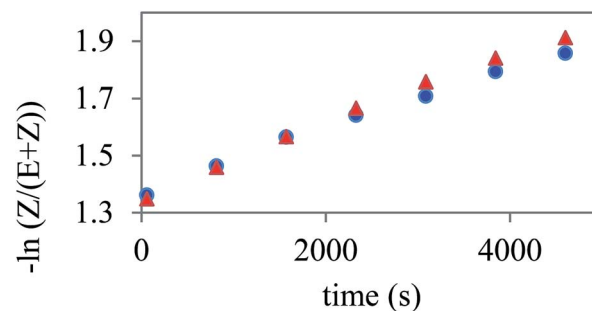


Fig. 9 The time dependent decrease of the concentration of *Z*-isomer in the solution of **Ib** (triangles) and **IIb** (dots) depicted in the logarithmic scale.

be envisioned for a non-symmetrical bent-core compound. Therefore, we decided to study substances with one azo group in the elongating side arms, namely polymorphic **Ib** and crystalline **IIb**.

From the acquired ^1H NMR spectra (see Fig. S3 and S4[†]), thermal relaxation back to the thermodynamically more stable *E*-isomer from the *Z*-isomer formed by UV-irradiation is evident. The relaxation rate constants determined from the slope of the decrease was $1.25 \cdot 10^{-4} \text{ s}^{-1}$ for **Ib** and $1.08 \cdot 10^{-4} \text{ s}^{-1}$ for **IIb** (Fig. 9). This result indicates a slightly lower flexibility of compound **IIb**, which may be caused by the steric influence of the lateral substituent in position 6 on the central core. This feature can be used for the design of the next generation of slow relaxing photosensitive bent-core compounds. Such compounds with very stable *Z*-isomer can be then used for optical data storage or photo-alignment of LC matrices.

Ab initio calculations

The differences in mesomorphic properties can also be explained on the basis of a preferred molecular conformation. The conformers with minimum energy of **Ib** showing three consecutive non-polar SmC phases and a DC phase, and **IIc** exhibiting the SmC_aP_A phase were calculated using density functional theory method at B3LYP 3-21g level in Gaussian software. Despite the fact that the calculations were performed

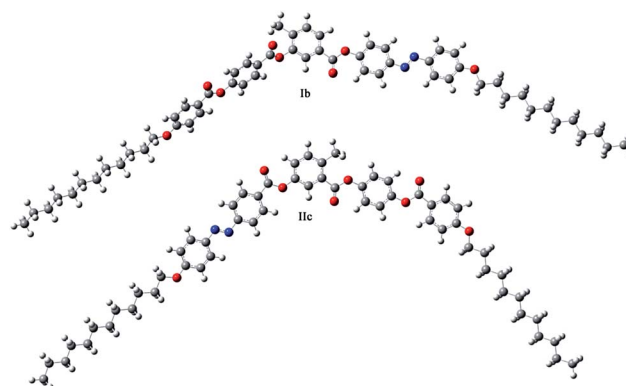


Fig. 10 Conformers with minimum energy of **Ib** and **IIc** obtained by *ab initio* calculations.



for a single molecule in vacuum, the conformers with minimum energy showed strong preference of **Ib** towards more linear arrangement, while **Ic** prefers rather a bent molecular shape (Fig. 10). The favourable molecular shapes correspond well with the formed mesophases; more linear molecules tend to organize in mesophases typical for rod-like mesogens, and bent molecules form mesophases typical of bent-core molecules.

Conclusions

New photosensitive BCLCs based on 4- and 6-methyl substituted 3-hydroxybenzoic acid have been synthesised and studied. We show that the position of the lateral substituent on the central core with respect to the position of the azo group-bearing elongating side arm strongly influences physical properties of the materials. For 4-methyl substituted materials of series **I**, generally, non-polar synclitic SmC phases have been identified. In case of **Ib**, rich polymorphism has been observed on cooling from the isotropic phase, namely, nematic phase followed by three synclitic smectic phases and ending with a dark conglomerate phase (DC phase). None of the SmC_s phase revealed switching under the applied electric field. Non-polar character is further supported by the dielectric spectroscopy data. In details, we have described the disordered meso-structure of the DC phase, which exists below the SmC_s phases at room temperature. It was confirmed that this DC phase has a different structure from those of B₄ phases and DC sponge phases, but very similar to the DC phases of the related 4-methylresorcinol derived BCLCs. This indicates that inverting the direction of one ester group in the side arm of the bent-core mesogen changes the fine structure of the formed DC phase, thus, adding a new subtype to the family of these phases.

The synclitic SmC phases observed for series **I** were replaced for series **II** derivatives either by anticlinic polar SmC phases (SmC_aP_A phases) or crystalline phases. The formation of the SmC_aP_A phase (materials **IIa** and **IIc**) is probably due to the preferred bent molecular shape of the materials. The absence of mesomorphic behaviour in the case of **IIb** and **IId** can be explained by the limited flexibility of the side chain next to the lateral substituent. The negative effect of a larger substituent in the vicinity of the carboxylic unit on the 3-hydroxybenzoic acid-based materials in combination with less flexible side chains has recently been described.²⁶ It is reasonable to assume that the introduction of the azo unit further limits the overall flexibility of the substance required for efficient molecular packing within a mesophase.

In conclusion, the mesomorphic properties of the presented materials are mainly influenced by the mutual position of the photosensitive unit and the lateral substituent. Thus, proper molecular design may provide materials exhibiting mesomorphic properties typical of calamitic materials or bent-core liquid crystals. Slower thermal relaxation of UV-induced Z-isomer of azo-containing bent-core materials may be beneficial for optical data storage and related applications.

Conflict of interest

There are no conflicts of interest to declare.

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Notes and references

- 1 D. R. Link, G. Natale, R. Shao, J. E. Maclennan, N. A. Clark, E. Körblova and D. M. Walba, *Science*, 1997, **278**, 1924.
- 2 T. Niori, T. Sekine, J. Watanabe, T. Furukawa and H. Takezoe, *J. Mater. Chem.*, 1996, **6**, 1231.
- 3 I. Dierking, *Angew. Chem., Int. Ed.*, 2010, **49**, 29.
- 4 R. A. Reddy and C. Tschierske, *J. Mater. Chem.*, 2006, **16**, 907.
- 5 C. Tschierske, *Angew. Chem., Int. Ed.*, 2013, **52**, 8828.
- 6 J. Etxebarria and M. Blanca Ros, *J. Mater. Chem.*, 2008, **18**, 2919.
- 7 W. Iglesias, N. L. Abbott, E. K. Mann and A. Jáklí, *ACS Appl. Mater. Interfaces*, 2012, **4**, 6884.
- 8 M. Nagaraj, Y. P. Panarin, J. K. Vij, C. Keith and C. Tschierske, *Appl. Phys. Lett.*, 2010, **97**, 213505.
- 9 D. Vorländer, *Ber. Dtsch. Chem. Ges. A*, 1929, **62**, 2831.
- 10 S. Kawata and Y. Kawata, *Chem. Rev.*, 2000, **100**, 1777.
- 11 P. Rochon, E. Batalla and A. Natansohn, *Appl. Phys. Lett.*, 1995, **66**, 136.
- 12 O. Yaroshchuk and Y. Reznikov, *J. Mater. Chem.*, 2012, **22**, 286.
- 13 M. Alaasar, *Liq. Cryst.*, 2016, **43**, 2208.
- 14 N. G. Nagaveni, P. Raghuvanshi, A. Roy and V. Prasad, *Liq. Cryst.*, 2013, **40**, 1238.
- 15 N. G. Nagaveni, A. Roy and V. Prasad, *J. Mater. Chem.*, 2012, **22**, 8948.
- 16 V. Prasad, *Liq. Cryst.*, 2001, **28**, 145.
- 17 M. Alaasar, M. Prehm, M. Nagaraj, J. K. Vij and C. Tschierske, *Adv. Mater.*, 2013, **25**, 2186.
- 18 M. Alaasar, M. Prehm, K. May, A. Eremin and C. Tschierske, *Adv. Funct. Mater.*, 2014, **24**, 1703.
- 19 M. Alaasar, M. Prehm, M.-G. Tamba, N. Sebastián, A. Eremin and C. Tschierske, *ChemPhysChem*, 2016, **17**, 278.
- 20 M. Alaasar, M. Prehm and C. Tschierske, *Liq. Cryst.*, 2013, **40**, 656.
- 21 M. Alaasar, M. Prehm and C. Tschierske, *Chem. Commun.*, 2013, **49**, 11062.
- 22 M. Alaasar, M. Prehm and C. Tschierske, *RSC Adv.*, 2016, **6**, 82890.
- 23 M. Alaasar, M. Prehm and C. Tschierske, *Chem.–Eur. J.*, 2016, **22**, 6583.
- 24 M. Alaasar, M. Prehm, M. Brautzsch and C. Tschierske, *J. Mater. Chem. C*, 2014, **2**, 5487.
- 25 V. Prasad, S.-W. Kang, X. Qi and S. Kumar, *J. Mater. Chem.*, 2004, **14**, 1495.
- 26 N. Trisovic, J. Antanasijevic, T. Toth-Katona, M. Kohout, M. Salamonczyk, S. Sprunt, A. Jakli and K. Fodor-Csorba, *RSC Adv.*, 2015, **5**, 64886.



- 27 M. Kohout, V. Kozmík, M. Slabochová, J. Tůma, J. Svoboda, V. Novotná and D. Pocięcha, *Liq. Cryst.*, 2015, **42**, 87.
- 28 J. Tůma, M. Kohout, J. Svoboda, V. Novotná and D. Pocięcha, *Liq. Cryst.*, 2016, **43**, 1889.
- 29 M. Alaasar, M. Prehm, M. Brautzsch and C. Tschierske, *Soft Matter*, 2014, **10**, 7285.
- 30 J. M. El Khoury, X. Zhou, L. Qu, L. Dai, A. Urbas and Q. Li, *Chem. Commun.*, 2009, 2109.
- 31 M. Kohout, J. Svoboda, V. Novotná, D. Pocięcha, M. Glogarová and E. Gorecka, *J. Mater. Chem.*, 2009, **19**, 3153.
- 32 D. Demus and L. Richter, *Textures of Liquid Crystals*, VEB Deutscher Verlag, Leipzig, 1978.
- 33 A. V. Emelyanenko and M. A. Osipov, *Phys. Rev. E: Stat., Nonlinear, Soft Matter Phys.*, 2004, **70**, 021704.

