

Cite this: *RSC Adv.*, 2017, 7, 31833Received 6th June 2017
Accepted 13th June 2017

DOI: 10.1039/c7ra06347d

rsc.li/rsc-advances

Simple design to achieve red-to-near-infrared emissive cationic Ir(III) emitters and their use in light emitting electrochemical cells†

 Amlan K. Pal,^a David B. Cordes,^b Alexandra M. Z. Slawin,^b Cristina Momblona,^c Antonio Pertegás,^c Enrique Ortí,^b Henk J. Bolink^{*c} and Eli Zysman-Colman^{†a}

Two cationic Ir(III) complexes bearing 2-phenylpyridinato cyclometalating ligands and bithiazole-type ancillary ligands have been synthesized and optoelectronically characterised. These emitters exhibit unusually deep red-to-near-infrared emission at room temperature, thereby rendering them as attractive emitters in solution-processed light emitting electrochemical cell (LEEC) electroluminescent devices.

Near-infrared- (NIR-) emitting compounds are an important subclass of luminescent materials due to their potential applications in sensors, night-vision displays, optical cosmetology, telecommunication and photodynamic therapies.¹ The most studied NIR-dyes in these contexts comprise semiconductor quantum dots,² lanthanoid complexes,³ organic small-molecule dyes⁴ and transition metal complexes.^{1a,b,5} With rich photo-physical behaviour, the NIR-emitting transition metal complexes are attractive candidates for solid-state lighting (SSL), as, with access to the triplet state, 100% of the electrically generated excitons can be harvested. Iridium complexes are well known phosphors for SSL due to their relatively short excited-state lifetimes, high photoluminescence quantum yields (Φ_{PL}) and remarkable colour tunability, particularly across the visible spectrum.⁶ While stable green and yellow emissive iridium complexes are relatively well documented and have been successfully incorporated into electroluminescent (EL) devices, there is a dearth of examples of red-to-NIR emissive iridium complexes.^{1a,b,7} Cationic Ir complexes of the form $[\text{Ir}(\text{C}^{\wedge}\text{N})_2(\text{N}^{\wedge}\text{N})]^+$ (where $\text{C}^{\wedge}\text{N}$ = cyclometalating ligand and $\text{N}^{\wedge}\text{N}$ = ancillary ligand) are the most widely studied class of emitters for LEECs.^{1b} In these charged complexes, the highest occupied molecular orbital (HOMO) typically resides on both the $\text{C}^{\wedge}\text{N}$ ligand and metal, while the lowest unoccupied molecular orbital (LUMO) resides on the $\text{N}^{\wedge}\text{N}$ ligand.⁸ The

common strategies of achieving red-to-NIR emission ($\lambda_{\text{em}} > 620$ nm) include: (a) destabilisation of the HOMO through introduction of electron-donating substituents^{7a,9} onto the $\text{C}^{\wedge}\text{N}$ ligands and/or stabilisation of the $\text{N}^{\wedge}\text{N}$ ligand-based LUMO by incorporation of electron-withdrawing substituents,¹⁰ and (b) introduction of π -conjugated systems either on one or both of the $\text{C}^{\wedge}\text{N}$ or $\text{N}^{\wedge}\text{N}$ ligands¹¹ thereby decreasing the HOMO–LUMO gap. Based on these principles, several charged red-to-NIR emitting Ir(III) complexes have been reported, a representative selection of which are shown in Chart 1 (sections a and b).^{7a,9}

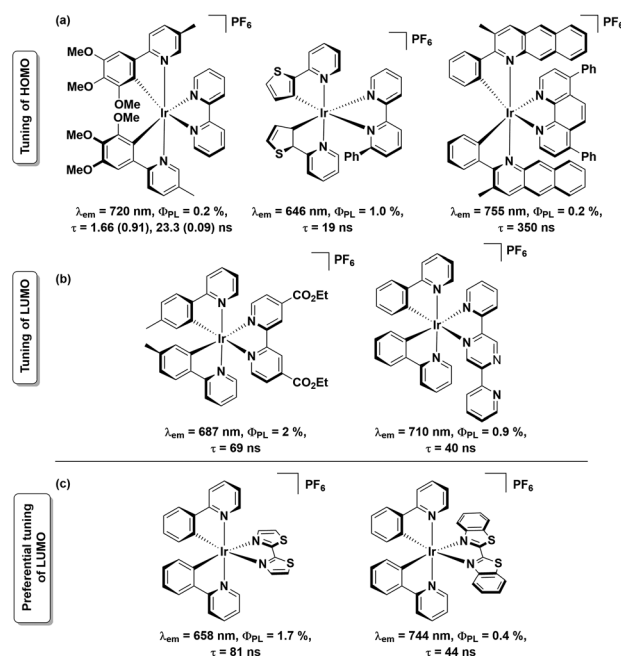


Chart 1 Strategies of achieving red-to-NIR emissive cationic Ir(III) complexes by (a) destabilisation of HOMO (top row), (b) stabilisation of LUMO (middle row). Complexes in current study are shown in (c) bottom row.

^aOrganic Semiconductor Centre, EaStCHEM School of Chemistry, University of St Andrews, St Andrews, Fife KY16 9ST, UK. E-mail: eli.zysman-colman@st-andrews.ac.uk; Web: <http://www.zysman-colman.com>

^bEaStCHEM School of Chemistry, University of St Andrews, St Andrews, Fife KY16 9ST, UK

^cInstituto de Ciencia Molecular, Universidad de Valencia, C/Catedrático José Beltrán 2, 46980 Paterna, Spain

† Electronic supplementary information (ESI) available: Synthetic, crystallographic, optoelectronic and DFT studies. CCDC 1538393 and 1538394. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c7ra06347d



Of the strategies outlined above, that involving LUMO stabilization is the most underexplored, despite its potential to produce red-to-NIR emissive cationic Ir(III) complexes.^{10,11} In this context, herein, we report the syntheses of two red-to-NIR cationic Ir(III) complexes, [Ir(ppy)₂(L1)]⁺[PF₆]⁻, **1**, and [Ir(ppy)₂(L2)]⁺[PF₆]⁻, **2**, possessing remarkably simple ligand architectures. These complexes bear 2-phenylpyridinato-C²,N' as the C²N ligand and strongly π-accepting ligands 2,2'-bithiazole (L1) and 2,2'-bibenzo[d]thiazole (L2) as the N²N ligands (Chart 1c). Coincident with the present study, complex **2** was recently reported by Ertl *et al.*,¹² though no Light Emitting Electrochemical Cell (LEEC) device data was reported. Herein, the optoelectronic properties of complexes **1** and **2** are reported along with the LEEC device performance.

Ligands L1 and L2 were synthesized by Pd-catalysed homocoupling of 2-bromothiazole or 2-bromobenzothiazole, respectively, in moderate yields (41% for L1 and 51% for L2, ESI†).¹³ Complexes **1** and **2** were obtained by cleavage of the [Ir(ppy)₂Cl]₂ dimer in the presence of 2.3 equiv. of the corresponding ancillary ligands followed by an anion metathesis with aqueous NH₄PF₆ in 91% and 95%, respectively (ESI†). Ligands L1 and L2 and complexes **1** and **2** were characterised by ¹H and ¹³C NMR spectroscopy, HRMS, melting point determination and elemental analyses. The ¹H and ¹³C NMR spectra confirmed the inherent C₂ symmetry present in both **1** and **2** (Fig. S1–S4, ESI†). Single crystal XRD studies of **1** and **2** corroborated the microanalysis and the geometry of the complexes (Fig. S5, Tables S1 and S2, ESI†).

In both crystal structures, the coordinatively saturated Ir(III) ions occupy a distorted octahedral geometry. The C²N carbon atoms are mutually *cis* disposed, similar to the solid-state structure of [Ir(ppy)₂(bpy)]⁺, **R1** (where bpy = 2,2'-bipyridine).¹⁴ The ancillary ligands coordinate in a bidentate N',N'-mode through the hard N-donors instead of the soft S-donors. The Ir–C²N bond distances [ranging from 2.003(7) to 2.007(6) Å in **1** and 2.009(4) to 2.009(5) Å in **2**] are similar to those observed in **R1** [ranging from 2.004(4) to 2.025(4) Å]. The Ir–N_{C²N} bond distances are likewise similar across the three complexes [ranging from 2.046(5) to 2.048(5) Å in **1**, 2.049(4) to 2.058(4) Å in **2** and 2.042(3) to 2.048(3) Å in **R1**]. The Ir–N_{N²N} length [2.148(6) to 2.149(6) Å in **1** and 2.180(4) to 2.195(4) Å in **2**] resemble closely those in **R1** [2.129(3) to 2.137(3) Å in **2**].

The electrochemical properties of **1** and **2** were monitored by cyclic- (CV) and differential pulse- (DPV) voltammetry and the CV and DPV traces in MeCN are shown in Fig. S6 (ESI†). The detailed CV data *vs.* SCE (Fc/Fc⁺ = 0.38 V in MeCN)¹⁵ are

summarized in Table S3 (ESI†) while only first redox potentials are listed in Table 1. In the anodic scan, complexes **1** and **2** exhibit quasi-reversible mono-electronic oxidation processes at 1.24 V and 1.37 V, respectively. These oxidation processes are assigned to the Ir(III)/IV redox couple with contributions from the C²N ligands, assignments supported by DFT calculations (Fig. S7†). While the oxidation potential of complex **1** ($E_{1/2}^{\text{Ox}} = 1.24$ V) is similar to that of **R1** ($E_{1/2}^{\text{Ox}} = 1.27$ V),¹⁶ that of complex **2** ($E_{1/2}^{\text{Ox}} = 1.37$ V) is anodically shifted by 100 mV compared to **R1**, which aligns with the stabilized nature of the calculated HOMO of **2** ($E_{\text{HOMO}} = -5.80$ eV) compared to that found in **1** ($E_{\text{HOMO}} = -5.58$ eV) and **R1** ($E_{\text{HOMO}} = -5.56$ eV). The more positive oxidation potential accounted for complex **2** is probably due to the modulation of the electron density of the Ir(III) ion by L2 as the latter is a better π-acceptor than L1 or 2,2'-bpy and thereby rendering the Ir(III) centre more electron deficient compared to that in complex **1**.

Upon scanning to negative potential, complex **1** exhibits two quasi-reversible one-electron reduction processes up to -2 V, while three other irreversible and quasi-reversible reduction waves could also be observed at higher negative potentials. Complex **2**, on the other hand, displayed four mono-electronic quasi-reversible waves up to -3 V (Table S3 and Fig. S5, ESI†). DFT calculations show that the calculated LUMOs of **1** and **2** are localized on the N²N ligand and therefore the first reduction processes are ascribed to radical anion formation on the ancillary ligand. The first reduction potentials for **1** and **2** ($E_{1/2}^{\text{Red}} = -1.15$ V for **1** and -0.82 V for **2**) are anodically shifted by 230 mV and 560 mV, respectively, compared to that for **R1**, demonstrating the significantly stronger π-accepting nature of ligands L1 and L2 compared to bpy. The CV data mirrors the DFT results, which show significant LUMO stabilization from **R1** ($E_{\text{LUMO}} = -2.31$ eV) to **1** ($E_{\text{LUMO}} = -2.76$ eV) to **2** ($E_{\text{LUMO}} = -3.07$ eV). The trend in the calculated decreasing HOMO–LUMO gap from **R1** (3.25 eV) to **1** (2.82 eV) to **2** (2.73 eV) is in good agreement to the decreasing redox gap observed for **R1** (2.65 V) to **1** (2.38 V) to **2** (2.19 V). The second quasi-reversible and third irreversible reductions (Table S3, ESI†) of complex **1** could be ascribed to the reductions of the pyridine moiety of the C²N ligand based on the coarse approximation that both the LUMO+1 and LUMO+2 reside on the C²N ligand as predicted by DFT calculations (Fig. S7, ESI†) and similar assignments are also valid for the second and third quasi-reversible reductions of complex **2** (Table S3, ESI†).

Table 1 Optoelectronic properties of **1**, **2** and **R1** in degassed MeCN^a

	$\lambda_{\text{em}}/\text{nm}$	$\Phi_{\text{PL}}/\%$	$\tau_{\text{e}}/\mu\text{s}$	$k_{\text{r}}/10^5 \text{ s}^{-1}$	$k_{\text{nr}}/10^5 \text{ s}^{-1}$	$E_{1/2}^{\text{Ox}}/\text{V} (\Delta E_{\text{p}}, \text{mV})$	$E_{1/2}^{\text{Red}}/\text{V} (\Delta E_{\text{p}}, \text{mV})$
1	658	1.72	0.081	0.21	123.2	1.24 (110)	-1.15 (74)
2	744	0.35	0.044	0.08	227.2	1.37 (169)	-0.82 (88)
R1 ^b	602	9.29	0.275	3.4	33	1.27	-1.38

^a Φ_{PL} measured at 298 K using [Ru(bpy)₃]²⁺Cl₂ ($\Phi_{\text{PL}} = 4\%$ in H₂O, aerated) as the reference.¹⁹ $k_{\text{r}} = \Phi_{\text{PL}}/\tau_{\text{e}}$ and $k_{\text{nr}} = (1 - \Phi_{\text{PL}})/\tau_{\text{e}}$. CV values are *vs.* SCE.¹⁵ $\Delta E_{\text{p}} = |E_{\text{pa}} - E_{\text{pc}}|$, where E_{pa} = anodic peak potential and E_{pc} = cathodic peak potential; $E_{1/2} = (E_{\text{pa}} + E_{\text{pc}})/2$. A silver wire was used as the pseudoreference electrode; a glassy-carbon electrode was used for the working electrode and a Pt electrode was used as the counter electrode.

^b Φ_{PL} measured at 298 K using [Ru(bpy)₃](PF₆)₂ ($\Phi_{\text{PL}} = 9.5\%$ in MeCN) as the reference.²⁰



The UV-vis absorption spectra of complexes **1** and **2** were obtained in dry MeCN and the values are tabulated in Table S4.† Overlays of experimentally observed and theoretically predicted absorption data by TD-DFT are shown in Fig. S8 (ESI†). The UV region is dominated by spin allowed $^1\pi \rightarrow \pi^*$ transitions in both ligand moieties centered around 240–280 nm (Tables S4 and S5, ESI†). For complex **2**, singlet ligand-centered $\pi \rightarrow \pi^*$ transitions (1LC) extend up to 407 nm, as predicted by TD-DFT, due to increased conjugation in **L2** (Table S5†). The set of absorption bands clustered around 315 nm in **1** are significantly red-shifted in **2** at ca. 363 nm. The character of these band in **1** is an admixture of singlet metal-to-ligand charge transfer (1MLCT) from $Ir(d\pi) \rightarrow ppy(\pi^*)$ and ppy -based 1LC transitions while in **2** the nature of these transitions is more CT in character and comprises 1MLCT from $Ir(d\pi) \rightarrow L2(\pi^*)$ and singlet ligand-to-ligand charge transfer (1LLCT) from $ppy(\pi) \rightarrow L2(\pi^*)$, as suggested by TD-DFT. The greater CT character coupled with the lower lying LUMO is responsible for this large bathochromic shift. The lowest energy absorption maxima for **1** and **2** appear at 500 nm and 599 nm, respectively. These bands are assigned as a mixture of spin-allowed and spin-forbidden $^1MLCT/{}^3MLCT$ and ligand-to-ligand charge transfer ($^1LLCT/{}^3LLCT$) transitions. These hypochromic bands at lower energy are the result of poor spatial overlap between the HOMO and LUMO and are typical for cationic Ir complexes.^{6,17} Concomitant to the larger calculated HOMO–LUMO gap for **1** ($\Delta E = 2.82$ eV) than **2** ($\Delta E = 2.73$ eV) (Fig. S7†), the absorption onset of **2** is also red-shifted by 99 nm (3.31×10^3 cm⁻¹) compared to that of **1**. The UV-vis absorption spectrum of complex **2** in CH₂Cl₂ solution exhibits intense absorption bands extending into the visible region (around 450 nm), as observed by Ertl *et al.*¹² The observed red-shift of 149 nm of the lowest energy transition of complex **2** in comparison to the study of Ertl *et al.* is presumably because of the solvent polarity where the MeCN used in this study stabilises both Frontier MOs, perhaps more pronounced for the LUMO. In comparison to **R1** (lowest-energy $\lambda_{max}^{abs} = 420$ nm),¹⁶ both complexes show significantly red-shifted absorption spectra.

The steady-state emission spectra of **1** and **2** were recorded in degassed MeCN at 298 K and are shown in Fig. S9 (ESI†). The photophysical data are summarised in Table 1. The emission profiles are broad and featureless, indicative of mixed 3CT emission, as also observed by Ertl *et al.*¹² The spin density distribution of the excited-state corroborates this assignment (Fig. S9 and S10†). The emission maximum ($\lambda_{max}^{em} = 744$ nm) of complex **2** is red-shifted by 86 nm (1.76×10^3 cm⁻¹) compared to that of complex **1** ($\lambda_{max}^{em} = 658$ nm). The DFT predicted emission maxima for **1** ($\lambda_{DFT}^{em} = 712$ nm) and **2** ($\lambda_{DFT}^{em} = 761$ nm) also follow the observed trend in emission maxima of complexes **1** and **2**, with relative errors of 7.6% and 2.2%, respectively. The emission maxima of **1** and **2** are significantly red-shifted compared to that of **R1** ($\lambda_{max}^{em} = 602$ nm).¹⁶ In comparison to the observed emission maximum at 686 nm in CH₂Cl₂ solution of complex **2** by Ertl *et al.*,¹² the same for complex **2** is red-shifted by 58 nm in our study, which is because of the effect of solvent polarity in emission.¹⁸ With the exception of the cationic Ir(III) complex bearing 3-methyl-2-phenylbenzo[*g*]

quinolinato C^N ligands reported by Tao *et al.*^{7a} (Chart 1a), complex **2** is, to the best of our knowledge, the most red-shifted emitting cationic Ir(III) complex.^{7a,9–11} The Φ_{PL} in degassed MeCN of **1** and **2** are 1.7% and 0.4%, respectively (Table 1). The corresponding excited-state lifetimes, τ_e , are very short at 81 and 44 ns, for **1** and **2**, respectively. The lower Φ_{PL} and shorter τ_e values compared to **R1** are a stark consequence of energy-gap law as evidence by the associated ~4-fold (for **1**) and 7-fold (for **2**) increase in the non-radiative decay rate constant, k_{nr} , compared to that in **R1**. The observed Φ_{PL} and τ_e values of complex **2** in CH₂Cl₂ solution, reported by Ertl *et al.*, are 3.6% and 126 ns, respectively.¹² The lower Φ_{PL} and τ_e values in our study compared to the study of Ertl *et al.* is again due to a consequence of energy-gap law for a NIR emitter like **2**.

The emission maxima of the thin films of **1** and **2** mixed with 1-butyl-3-methylimidazolium hexafluorophosphate, [Bmim][PF₆] are 658 and 707 nm, respectively ($\lambda_{exc} = 320$ nm). The thin film Φ_{PL} values of 14% and 5% for **1** and **2**, respectively, are enhanced compared to solution values. The PL spectra and photophysical data of the films are displayed in Fig. S11 and Table S9 (ESI†).

Light emitting electrochemical cells (LEECs) were prepared and the electroluminescence properties of the complexes were evaluated. The preparation of LEECs were done by spin-coating on pre-cleaned ITO substrates, which were first coated with a 80 nm thick film of PEDOT:PSS. The active layer was then deposited from dichloromethane solution. Recently, a first attempt was done for the LEEC fabrication with complex **2**.¹² At that time, the bad film morphology led to poor LEEC performance. In this work, a different solvent was used to improve the film morphology of the emissive layer. Here, the complex was dissolved in dichloromethane, which led to good films completely covering the substrates after spin-coating. After the active layer deposition, the top electrode (70 nm thick film of aluminium) was thermally evaporated. The fabricated LEECs were measured using a pulsed current driving (1 kHz, 50% duty cycle and block wave). The current densities applied to the LEECs were 800 and 1600 A m⁻², which corresponds to an average current density of 400 and 800 A m⁻², respectively, as a function of the duty cycle. For simplicity, the LEECs fabricated with **1** and **2** are called LEEC-1 and LEEC-2.

The luminance for LEEC-1 or light output for LEEC-2 and average voltage *versus* time measured under average current density of 400 A m⁻² is depicted in Fig. 1 while the measurements at 800 A m⁻² are shown in Fig. S12 (ESI†). In the inset of Fig. 1, the electroluminescence (EL) spectrum of each LEEC is shown and all the device data are summarized in Table S11 for LEEC-1 and Table S12 for LEEC-2, respectively (ESI).† LEEC-1 exhibits a red electroluminescence centred at 661 nm while LEEC-2 shows a near-infrared emission at 705 nm, in both cases, in good agreement with the PL spectra registered in thin film (658 nm for **1** and 707 nm for **2**). The applied voltage rapidly diminishes due to the ionic motion over operation, which decreases the electrical resistance. As this characteristic is directly related with the ionic mobility, generally, the time to reach the maximum light emission (t_{max}) in LEECs varies from seconds to hours. Despite the similarity structure of complexes



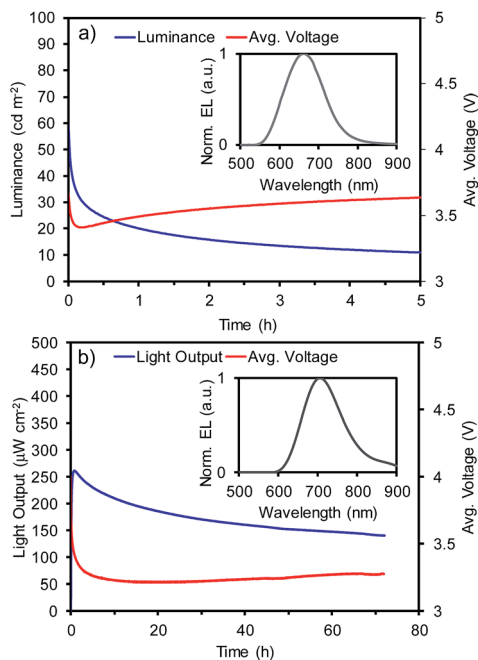


Fig. 1 Device performance versus time and electroluminescence (inset) for (a) LEEC-1 and (b) LEEC-2 under a pulsed current of 400 A m⁻² (average current density), 1 kHz and 50% of duty cycle.

1 and **2** as well as the presence of the same counter-anion (PF₆), t_{\max} is significantly shorter for LEEC-1 (<5 seconds) than LEEC-2 (0.8 and 4.3 hours for 400 and 800 A m⁻², respectively). Moreover, differences are observed between LEEC-1 and LEEC-2 when comparing the device lifetime, usually defined as the time to reach one-half of the maximum light emission ($t_{1/2}$). While LEEC-1 shows a $t_{1/2}$ of 0.3 hours, LEEC-2 was significantly more stable at above 80 hours. The differences observed seem to be related to the presence of the annulated benzenes in the ancillary ligand. A possible explanation is related to the larger cation in complex **2**, which reduces the ionic movement in the active layer during operation. LEEC-2 exhibits an improved efficiency compared to that based on **1**. While LEEC-1 achieved an external quantum efficiency (EQE) of 0.23% and 0.13%, LEEC-2 reached 0.37% and 0.30% at 400 and 800 A m⁻², respectively. Hence, the efficiency of LEEC-1 is more strongly affected by increasing the current density than LEEC-2. Thus, the slower ionic movement in LEEC-2 leads to more balanced carrier. To the best of our knowledge, only a few examples of near-infrared emitting LEECs have been reported, generally with low EQE (EQE < 0.1%)⁶ due to the energy gap law. Recently, new examples have overcome the EQE of 1%,^{5d,21} but none of these is based on a cationic iridium complex as the emitter. Considering the low emission quantum yield of complex **2** (0.4%) and the high current density applied (up to 800 A m⁻²) the achieved efficiencies for LEEC-2 are very impressive and close to the theoretical achievable one assuming a typical light outcoupling of 20%.

In summary, we demonstrate dramatic red-shifted emission for both **1** and **2** as a function of the use of strongly π -accepting bithiazole-type ancillary ligands compared to the archetype

complex **R1**. Notably, to the best of our knowledge, complex **2** displays the most red-shifted emission maximum for a cationic iridium complex, especially considering the simplicity of the ligand structure and synthesis. LEEC devices show deep red and NIR emission using **1** and **2**, respectively. LEEC-2 shows improved efficiency compared to LEEC-1 and represents the first example of a NIR LEEC employing a cationic iridium complex emitter. Current efforts are focussed on pushing the emission even further into the near-IR through addition of strongly conjugated C^N ligands.

Acknowledgements

We are grateful to the University of St Andrews and EPSRC for financial support from grant EP/M02105X/1. We thank the EPSRC UK National Mass Spectrometry Facility at Swansea University for analytical services. We thank Umicore AG for the gift of IrCl₃. We also thank the Spanish Ministry of Economy and Competitiveness (MINECO) via the Unidad de Excelencia María de Maeztu MDM-2015-0538, MAT2014-55200, CTQ2015-71154-P and PCIN-2015-255, the Generalitat Valenciana (Prometeo/2016/135) and European FEDER funds (CTQ2015-71154-P). C. M. thanks MINECO for her predoctoral contract.

References

- (a) A. Barbieri, E. Bandini, F. Monti, V. K. Praveen and N. Armaroli, *Top. Curr. Chem.*, 2016, **374**, 47; (b) A. F. Henwood and E. Zysman-Colman, *Chem. Commun.*, 2017, **53**, 807–826; (c) J. Yellol, S. A. Perez, G. Yellol, J. Zajac, A. Donaire, G. Viguera, V. Novohradsky, C. Janiak, V. Brabec and J. Ruiz, *Chem. Commun.*, 2016, **52**, 14165–14168; (d) G. Qian and Z. Y. Wang, *Chem.-Asian J.*, 2010, **5**, 1006–1029.
- C. Wang, X. Gao and X. Su, *Anal. Bioanal. Chem.*, 2010, **397**, 1397–1415.
- S. V. Eliseeva and J. C. Bunzli, *Chem. Soc. Rev.*, 2010, **39**, 189–227.
- (a) E. Zysman-Colman, S. S. Ghosh, G. Xie, S. Varghese, M. Chowdhury, N. Sharma, D. B. Cordes, A. M. Z. Slawin and I. D. W. Samuel, *ACS Appl. Mater. Interfaces*, 2016, **8**, 9247–9253; (b) J. O. Escobedo, O. Rusin, S. Lim and R. M. Strongin, *Curr. Opin. Chem. Biol.*, 2010, **14**, 64–70; (c) O. A. Melville, B. H. Lessard and T. P. Bender, *ACS Appl. Mater. Interfaces*, 2015, **7**, 13105–13118.
- (a) A. K. Pal, S. Serroni, N. Zaccheroni, S. Campagna and G. S. Hanan, *Chem. Sci.*, 2014, **5**, 4800–4811; (b) H. Xiang, J. Cheng, X. Ma, X. Zhou and J. J. Chruma, *Chem. Soc. Rev.*, 2013, **42**, 6128–6185; (c) D. A. Ross, P. A. Scattergood, A. Babaei, A. Pertegas, H. J. Bolink and P. I. Elliott, *Dalton Trans.*, 2016, **45**, 7748–7757; (d) J. H. Hsu and H. C. Su, *Phys. Chem. Chem. Phys.*, 2016, **18**, 5034–5039; (e) B. N. Bideh, C. Roldan-Carmona, H. Shahroosvand and M. K. Nazeeruddin, *J. Mater. Chem. C*, 2016, **4**, 9674–9679.
- A. F. Henwood and E. Zysman-Colman, *Top. Curr. Chem.*, 2016, **374**, 36.



- 7 (a) R. Tao, J. Qiao, G. Zhang, L. Duan, L. Wang and Y. Qiu, *J. Phys. Chem. C*, 2012, **116**, 11658–11664; (b) D. Ma, T. Tsuboi, Y. Qiu and L. Duan, *Adv. Mater.*, 2017, **29**, 1603253.
- 8 The reported complex is $[(ppy-F_2)_2Ir(dtBubpy)]^+$, where ppy- F_2 is 2,4-difluorophenylpyridinato. H. J. Bolink, E. Coronado, R. n. D. Costa, N. Lardiés and E. Orti, *Inorg. Chem.*, 2008, **47**, 9149–9151.
- 9 (a) K. Hasan, A. K. Bansal, I. D. W. Samuel, C. Roldán-Carmona, H. J. Bolink and E. Zysman-Colman, *Sci. Rep.*, 2015, **5**, 12325; (b) A. M. Bünzli, H. J. Bolink, E. C. Constable, C. E. Housecroft, J. M. Junquera-Hernandez, M. Neuburger, E. Orti, A. Pertegas, J. J. Serrano-Perez, D. Tordera and J. A. Zampese, *Dalton Trans.*, 2014, **43**, 738–750.
- 10 (a) L. Donato, C. E. McCusker, F. N. Castellano and E. Zysman-Colman, *Inorg. Chem.*, 2013, **52**, 8495–8504; (b) J. L. Rodriguez-Redondo, R. D. Costa, E. Orti, A. Sastre-Santos, H. J. Bolink and F. Fernandez-Lazaro, *Dalton Trans.*, 2009, 9787–9793; (c) D. Ma, L. Duan and Y. Qiu, *J. Mater. Chem. C*, 2016, **4**, 5051–5058.
- 11 (a) Y. Li, N. Dandu, R. Liu, Z. Li, S. Kilina and W. Sun, *J. Phys. Chem. C*, 2014, **118**, 6372–6384; (b) A. M. Bünzli, E. C. Constable, C. E. Housecroft, A. Prescimone, J. A. Zampese, G. Longo, L. Gil-Escrig, A. Pertegás, E. Orti and H. J. Bolink, *Chem. Sci.*, 2015, **6**, 2843–2852; (c) X. Zhu, L. Lystrom, S. Kilina and W. Sun, *Inorg. Chem.*, 2016, **55**, 11908–11919.
- 12 C. D. Ertl, C. Momblona, A. Pertegás, J. M. Junquera-Hernández, M.-G. La-Placa, A. Prescimone, E. Ortí, C. E. Housecroft, E. C. Constable and H. J. Bolink, *J. Am. Chem. Soc.*, 2017, **139**, 3237–3248.
- 13 B. Fu, C.-Y. Wang, B. D. Rose, Y. Jiang, M. Chang, P.-H. Chu, Z. Yuan, C. Fuentes-Hernandez, B. Kippelen, J.-L. Brédas, D. M. Collard and E. Reichmanis, *Chem. Mater.*, 2015, **27**, 2928–2937.
- 14 R. D. Costa, E. Ortí, H. J. Bolink, S. Graber, S. Schaffner, M. Neuburger, C. E. Housecroft and E. C. Constable, *Adv. Funct. Mater.*, 2009, **19**, 3456–3463.
- 15 V. V. Pavlishchuk and A. W. Addison, *Inorg. Chim. Acta*, 2000, **298**, 97–102.
- 16 For examples see: S. Ladouceur, D. Fortin and E. Zysman-Colman, *Inorg. Chem.*, 2011, **50**, 11514–11526.
- 17 Y.-C. Chiu, J.-Y. Hung, Y. Chi, C.-C. Chen, C.-H. Chang, C.-C. Wu, Y.-M. Cheng, Y.-C. Yu, G.-H. Lee and P.-T. Chou, *Adv. Mater.*, 2009, **21**, 2221–2225.
- 18 M. A. Haidekker, T. P. Brady, D. Lichlyter and E. A. Theodorakis, *Bioorg. Chem.*, 2005, **33**, 415–425.
- 19 H. Ishida, S. Tobita, Y. Hasegawa, R. Katoh and K. Nozaki, *Coord. Chem. Rev.*, 2010, **254**, 2449–2458.
- 20 S. Ladouceur and E. Zysman-Colman, *Eur. J. Inorg. Chem.*, 2013, **2013**, 2985–3007.
- 21 (a) C.-C. Ho, H.-F. Chen, Y.-C. Ho, C.-T. Liao, H.-C. Su and K.-T. Wong, *Phys. Chem. Chem. Phys.*, 2011, **13**, 17729–17736; (b) C.-L. Lee, C.-Y. Cheng and H.-C. Su, *Org. Electron.*, 2014, **15**, 711–720.

