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Regulation of coin metal substituents and cooperativity on the strength and nature of tetrel bonds†

Yuanxin Wei, Jianbo Cheng, Wenzuo Li and Qingzhong Li **D**

Ab initio calculations have been performed for the tetrel-bonded dyad MCN···TF₄ (M = Cu, Ag, and Au; T = C, Si, Ge, and Sn). A coin metal substituent greatly strengthens the tetrel bond, depending on the nature of the coin metal and tetrel atoms. The Ag substituent has the largest enhancing effect with the interaction energy approaching -16 kcal mol $^{-1}$. The strength of the tetrel bond in MCN···TF₄ can be further regulated by the cooperative effect with a π ····metal interaction. It is strengthened in C₂H₄····MCN···TF₄ but is weakened in C₂(CN)₄····MCN···TF₄. The interaction energy of the tetrel bond amounts to about -16 kcal mol $^{-1}$ in C₂H₄····CuCN···SiF₄, becoming a strong interaction. The nature of the tetrel bond is changed due to the metal substituents and cooperativity.

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1. Introduction

By plotting molecular electrostatic potential (MEP) maps of halogen-containing molecules, Clark *et al.*, found a positive MEP on the outermost portion of the halogen's surface, named the σ-hole.¹ This σ-hole is helpful in explaining the formation of a halogen bond and its directionality.² Such a σ-hole is also observed on many covalently-bonded atoms of Groups IV–VI.³-5 The corresponding intermolecular interactions are called tetrel,⁶ pnicogen,⁷ and chalcogen bonds.⁸ These intermolecular interactions have a uniform name: σ-hole bonding.⁹ Very recently, Grabowski and Sokalski performed a comparative study on some Lewis acid–Lewis base complexes linked by different σ-hole bonds and proposed that they have numerous correlations similar to hydrogen-bonded systems.¹⁰ More importantly, these σ-hole bonds have similar applications with hydrogen bonds in crystal materials, chemical reactions, and biological systems.^{67,11-15}

Recently, much attention has been paid to tetrel bonding ¹⁶⁻²⁶ since Bauzá *et al.*, pointed out that it might serve as a new possible molecular linker in supramolecular materials. ⁶ In addition, tetrel bonding is also confirmed to be important in chemical reactions ^{14,15} and molecular recognition. ^{27,28} Tetrel atoms are usually tetravalent in tetrel bonding, thus the formation of tetrel bonding suffers steric constraints if the tetrel atom is adjoined with steric groups. It is natural to understood that TH₄, TH₃X, and TF₄ (where T refers to a tetrel atom and X is an electron-withdrawing atom or group) are often used as the

The Laboratory of Theoretical and Computational Chemistry, School of Chemistry and Chemical Engineering, Yantai University, Yantai 264005, People's Republic of China. E-mail: liqingzhong1990@sina.com; Fax: +86 535 6902063; Tel: +86 535 6902063 † Electronic supplementary information (ESI) available. See DOI: 10.1039/c7ra09881b

tetrel donors in studying tetrel bonds. When a strong Lewis base attacks the tetravalent tetrel atom, the three covalent bonds facing the Lewis base are repelled to reduce steric constraints, and a trigonal bipyramid structure is finally favorable. This shows that deformation of a tetravalent tetrel molecule is crucial in stabilizing strong tetrel-bonded complexes.²⁶

For tetrel bonds, there is a tendency to be strengthened as the T atom moves down this column of the periodic table since the σ-hole on the T atom is enlarged.²⁹ CH₄ seldom participates in tetrel bonding due to the absence of a σ -hole on the carbon atom. However, it can do if the carbon atom binds with an electron-withdrawing atom or group.16 Scheiner performed a systematic study of factors that influence the strength of tetrel bonds.30 With NH3 as a universal Lewis base, unsubstituted TH4 molecules form the weakest tetrel bonds, and this bonding is strengthened gradually from TH₃F through THF₃ to TF₄.30 Additionally, this influence is also dependent on the nature of the T atom.30 HCN is a weaker Lewis base than NH3, while LiCN is a stronger Lewis base than NH₃ in forming a tetrel bond due to the strong electron-donating ability of Li atom.14 This indicates that metal substituents in the Lewis bases are effective in strengthening tetrel bonds like that in halogen bonds.31

In most cases, the applications of tetrel bonds in crystal materials and molecular recognition are reached by cooperative effect with itself and other interactions. The addition of a phenyl ring to a single imidazolium perturbs the binding very little with halide anions, while placement of a second imidazolium on the benzene connector group markedly enhances binding energies. Some supramolecular assemblies involving lead(π) complexes have been constructed by means of tetrel bonding and other interactions. Therefore, there are many theoretical studies reported for cooperativity involving tetrel bonds. $^{22,36-44}$ F₂CX (X = Se and Te) can simultaneously form a tetrel bond with the carbon

RSC Advances

atom and a chalcogen bond with the X atom, thus both interactions display anticooperativity.²² Both anion- π and tetrel-bond interactions are enhanced in the same multicomponent complex of X···pyrazine/1,4-dicyanobenzene···TH₃F/F₂TO.⁴⁴ In all, the strength of tetrel bonding can be tuned by cooperative effect.

Interestingly, Grabowski¹⁴ claimed that tetrel bond is not formed in $F_4Si\cdots NCH$ due to the small $Lp_N \to \sigma_{Si-F}^*$ orbital interaction and negligible charge transfer. This means that the tetrel bond in F4Si···NCH is weak with interaction energy of -2.8 kcal mol⁻¹. We plan to strengthen the tetrel bond in $F_4Si\cdots NCH$ by using MCN (M is a coin metal including Cu, Ag, and Au) and we also compare the effect of different coin metals on the strength of tetrel bonding. AuCN···TF₄ (T = C, Si, Ge, and Sn) is used to analyze the dependence of coin metal effect on the nature of the tetrel atom. To further strengthen the tetrel bond in MCN···TF₄, we introduce a third molecule C₂H₄ to bind with the coin metal in this complex. For comparison, C₂H₂···AuCN··· SiF_4 is also studied. In the π ···coin metal interaction, the π -electrons transfer from the occupied π -orbital into the empty σ -type metal orbital, while there is also a π -electron backdonation from the occupied d orbital of metal into the empty π anti-bonding orbital. To regulate the role of coin metals in the π ···coin metal interaction and its effect on the strength of tetrel bond, the four hydrogen atoms in C₂H₄ are replaced by strong electron-withdrawing group CN.

Theoretical methods

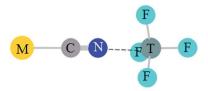
The structures of complexes and monomers were fully optimized using the Moller-Plesset perturbation (MP2) theory. The basis set adopted was aug-cc-pVTZ for the respective atoms except Sn and coin metal atoms, for which aug-cc-pVTZ-PP was used to account for relativistic effects. 46 Harmonic frequency calculations at the same level were carried out to confirm that the structures obtained corresponded to energy minima. The interaction energies using supermolecular method (difference between the energy of the complex and the energies of the monomers with their geometries within the complex) were corrected for the basis set superposition error (BSSE) by the counterpoise method of Boys and Bernardi.⁴⁷ All calculations were performed using the Gaussian 09 package. 48

The MEP analysis at the molecular surface was performed by the Wave Function Analysis-Surface Analysis Suite (WFA-SAS) program.49 The bonding characteristics were analyzed at the MP2/aug-cc-pVDZ(PP) level by means of Atoms-in-Molecules (AIM) theory⁵⁰ with the help of AIM 2000 software.⁵¹ The second-order perturbation energy and charge transfer was obtained at the HF/aug-cc-pVTZ level with natural bond orbital (NBO) method⁵² implemented in Gaussian 09. Interaction energy was decomposed using the LMOEDA method⁵³ at the same level by the GAMESS program.54

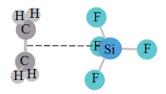
3. Results and discussion

Substituent effect of coin metal 3.1.

Fig. 1 shows the diagrams of some dyads studied, and the corresponding interaction energies are given in Table 1. The



MCN···TF₄



C2H4···SiF4

Fig. 1 Structures of MCN···TF₄ and C₂H₄···SiF₄.

Table 1 Interaction energy corrected for BSSE (ΔE , kcal mol⁻¹) and binding distance (R, Å) in the tetrel-bonded dyads^a

Dyads	ΔE	$\Delta \Delta E$	%	R
$AuCN\cdots CF_4$ $AuCN\cdots SiF_4$	-1.32(-1.09) $-4.95(-3.45)$	-0.23 -1.50	21.10 43.48	3.3268 2.8598
$AuCN\cdots GeF_4$ $AuCN\cdots SnF_4$	-19.60(-10.70) $-26.96(-12.94)$	-8.90 -14.02	83.18 108.35	2.1818 2.2343
AgCN···SiF ₄ CuCN···SiF ₄	-15.80 -6.12	-12.35 -2.67	357.97 77.39	2.2812 2.6338
$C_2H_4\cdots SiF_4$ $C_2H_4\cdots SiH_3F$	-1.93 -2.68	_	_	3.4880 3.2473

^a Note: data in parentheses are from HCN···TF₄ dyads. $\Delta\Delta E$ is the difference of ΔE in between MCN···TF₄ dyads and the corresponding HCN dyads. % is the percentage of $\Delta\Delta E$ to the ΔE of HCN···TF₄ dyads.

strength of tetrel bond has great dependence on the nature of tetrel atom since its interaction energy extensively varies from -1.32 kcal mol⁻¹ in AuCN···CF₄ to -26.96 kcal mol⁻¹ in AuCN···SnF₄. Clearly, the interaction energy of tetrel bond is more negative with the increase of the tetrel atomic mass, consistent with the magnitude of σ-hole on the tetrel atom (Fig. 2). Compared with HCN···TF₄, the tetrel bond is stronger in AuCN···TF₄ with the more negative interaction energy. The electron-donating nature of metal atom results in an increase of the negative MEP on the nitrogen atom in MCN relative to HCN (Fig. 2) and thus MCN forms a stronger tetrel bond with TF₄ than HCN. It is obvious that the enhancing effect of coin metal on the tetrel bond is related with the nature of the tetrel atom. Generally, the heavier the tetrel atom, the larger the enhancing effect of coin metal on the tetrel bond. The interaction energy of tetrel bond in AuCN···SnF4 is more than doubled, becoming a very strong tetrel bond. On the other hand, the enhancing effect of coin metal is also dependent on the nature of coin metal. The increased percentage of tetrel bonding interaction energy is larger in the order Au < Cu < Ag. The Ag substituent makes the interaction energy of tetrel bond increase from

Paper

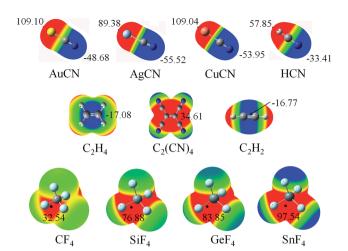


Fig. 2 MEP maps of the monomers studied, color ranges, in kcal mol^{-1} , are: red, greater than 9.5; yellow, between 9.5 and 0; green, between 0 and -9.5; blue, smaller than -9.5.

-3.45 kcal mol⁻¹ in HCN···SiF₄ to -15.80 kcal mol⁻¹ in AgCN··· SiF₄. Clearly, a strong tetrel bond is formed in AgCN···SiF₄. Moreover, the negative MEP on the nitrogen atom is largest in AgCN and smallest in AuCN (Fig. 2). Au has the greatest electronegativity among three coin metal elements, thus its electron-donating power is smallest. Additionally, Au-CN bears strong covalence interaction,55 which also reduces the negative charge density on the nitrogen atom. However, Ag-CN bears stronger ionic interaction than Au-CN,55 which is partly responsible for the larger negative MEP on the nitrogen atom of AgCN. For $HCN \cdots SiF_4$, our value $(-3.45 \text{ kcal mol}^{-1})$ of interaction energy is more negative than that in the previous value $(-2.8 \text{ kcal mol}^{-1})^{14}$ since our calculations were obtained on base of the geometries of the monomers in the complexes. This shows that deformation has some important contributions to the stability of tetrel-bonded complexes.

The N···C distance in AuCN···CF₄ is 3.3268 Å, which almost equals to the sum of the van der Waals (vdW) radii of both atoms. Although both molecules arrange in a line in AuCN··· CF₄, this complex can be taken as a van der Waals complex. The separation in other complexes is smaller than the sum of vdW Radii of the corresponding atoms. The N···Si distance in MCN···SiF₄ is smaller in the sequence AuCN > CuCN > AgCN, consistent with the change of the interaction energy.

Fig. 3 is the AIM diagrams of MCN···TF₄. In the weak complexes of AuCN···CF₄ and AuCN···SiF₄, the tetrel bond is characterized with three N···F paths. The N···F path is linear in AuCN···CF₄ but is curve in AuCN···SiF₄. The similar path is also found in CuCN···SiF₄. When the tetrel bond is strong enough, the N···F paths are changed to be a linear N···T path in AgCN···SiF₄, AuCN···GeF₄, and AuCN···SnF₄. Three fluorine atoms in the strong tetrel-bonded complexes are far away from the nitrogen atom in MCN. The topological parameters at these paths are listed in Table 2. The electron density at the N···C BCP is very small in AuCN···CF₄. Besides, the charge transfer and orbital interaction are also very small in this complex (Table 3). Thus the interaction in AuCN···CF₄ is not a tetrel bond. The

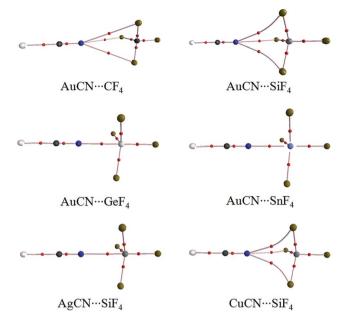


Fig. 3 AIM diagrams of MCN···TF₄.

Table 2 Electron density ($\rho_{\rm b}$, au), Laplacian ($\nabla^2 \rho_{\rm b}$, au), kinetic energy density ($G_{\rm b}$, au), potential energy density ($V_{\rm b}$, au), and total energy density ($H_{\rm b}$, au) at the intermolecular BCP in the tetrel-bonded dyads^a

^a Note: the bond path is N···F in AuCN···CF₄, AuCN···SiF₄, and CuCN··· SiF₄, N···T in AuCN···GeF₄, AuCN···SnF₄, and AgCN···SiF₄, and C···F in C_2H_4 ···SiF₄. All are average values for the N···F and C···F paths.

electron density at the N···Si BCP is increased in AuCN···SiF₄, and its Laplacian and energy density are positive. So the tetrel bond in AuCN···SiF₄ is a purely close-shell interaction according to the classification for intermolecular interactions by Arnold and Oldfield.⁵⁶ With the enhancement of tetrel bond, the electron density at the N····T BCP grows up and the similar result is found for the corresponding Laplacian. At the same time, the energy density is changed to be negative. These results show that the tetrel bond is a partially covalent interaction in CuCN····SiF₄, AgCN···SiF₄, AuCN···GeF₄, and AuCN···SnF₄.

The charge transfer in AuCN···CF₄ is not only very small but also it moves from CF₄ to AuCN. This further confirms no formation of a tetrel bond in AuCN···CF₄. In other complexes, the charge transfer moves from MCN to TF₄. The charge transfer in AuCN···SiF₄ is still small and the single Lp_N $\rightarrow \sigma_{T-F}^*$ orbital interaction is not large. Thus this charge transfer interaction has a small contribution to the enhancement of tetrel bond in AuCN···SiF₄. The charge transfer and orbital interaction have

RSC Advances

Table 3 Charge transfer (Q_{CT} , e) and second-order perturbation energy ($E^{(2)}$, kcal mol⁻¹) due to the four Lp_N $\rightarrow \sigma_{T-F}^*$ orbital interactions in the tetrel-bonded dyads a

Dyads	Q_{CT}	$E^{(2)}$
AuCN···CF ₄	0.0015	0.09
$AuCN\cdots SiF_4$	-0.0010	3.23
$AuCN\cdots GeF_4$	-0.1001	72.17
$AuCN\cdots SnF_4$	-0.1090	82.27
$AgCN \cdots SiF_4$	-0.0616	40.16
CuCN···SiF ₄	-0.0113	_
$C_2H_4\cdots SiF_4$	-0.0006	0.97

^a Note: Q_{CT} is the sum of charge on all atoms of TF₄. $E^{(2)}$ corresponds to the $\pi_{C=C} \to \sigma_{T=F}^*$ orbital interactions. $E^{(2)}$ in CuCN···SiF₄ is not shown owing to the abnormal value in the NBO analysis.

a larger increase in AgCN···SiF4 and CuCN···SiF4, indicating that charge transfer contribution cannot be ignored in strengthening the tetrel bond. The charge transfer amounts to about 0.1e in AuCN···GeF₄ and AuCN···SnF₄, and the corresponding orbital interaction is also very large.

The weak interaction in AuCN···CF₄ is still dominated by electrostatic energy and dispersion energy is more than half of electrostatic energy (Table 4). In AuCN···SiF4, electrostatic energy has a prominent increase but dispersion energy is almost not changed with respect to AuCN···CF4. The easier polarization of silicon atom leads to a relatively large increase of polarization energy in AuCN···SiF₄. For the strong tetrel bond in AuCN···GeF₄ and AuCN···SnF₄, electrostatic energy is very large $(>-50 \text{ kcal mol}^{-1})$, and polarization energy is more than half of electrostatic energy. The relatively large polarization energy is mainly responsible for the deformation of TF4 in AuCN···GeF4 and AuCN···SnF₄.54 The similar result is found in AgCN···SiF₄. The contribution of polarization energy relative to that of electrostatic energy in CuCN···SiF₄ is smaller than that in AgCN··· SiF₄. The interaction energies in Tables 1 and 4 are almost equal each other for most complexes with an exception in CuCN··· SiF_4 , where a difference of 2 kcal mol^{-1} is found. The main reason is that dispersion energy is obtained by the difference of MP2 and CCSD(T) energies.54 The strong tetrel bond in AuCN··· GeF₄ and AuCN···SnF₄ brings out a positive dispersion energy, which is primarily caused by the difference in the intra- and intermolecular correlation energy on going from noninteracting to interacting molecules.54

Table 4 Electrostatic (E^{ele}), exchange (E^{ex}), repulsion (E^{rep}), polarization (E^{pol}), dispersion (E^{disp}) energies, and total interaction energy (ΔE). All are in kcal mol⁻¹

Dyads	$E^{ m ele}$	$E^{\rm ex}$	E^{rep}	E^{pol}	E^{disp}	ΔE
$AuCN\cdots CF_4$	-1.89	-2.82	4.87	-0.40	-1.09	-1.33
AuCN···SiF ₄	-10.64	-2.62 -11.86	21.48	-2.67	-1.30	-4.99
$AuCN\cdots GeF_4$	-50.58	-56.10	113.75	-27.93	1.17	-19.69
$AuCN\cdots SnF_4$	-52.16	-53.59	107.96	-33.03	3.75	-27.08
$AgCN \cdots SiF_4$	-39.56	-46.01	89.66	-18.81	-1.41	-16.13
$CuCN\cdots SiF_4$	-18.72	-21.16	39.31	-5.90	-1.66	-8.12
$C_2H_4\cdots SiF_4$	-4.54	-7.22	12.96	-0.68	-2.44	-1.92
$C_2H_4{\cdots}SiH_3F\\$	-5.32	-12.4	20.7	-1.99	-3.71	-2.72

The interaction energy at the same level is -8.13, -8.46, and -6.77 kcal mol⁻¹ in CuCN···ClF, AgCN···ClF, and AuCN···ClF, respectively.57 The interaction energy is -6.12, -15.80, and -4.95 kcal mol⁻¹ in CuCN···SiF₄, AgCN···SiF₄, and AuCN···SiF₄, respectively. A comparative analysis indicates that the dependence of tetrel bond on the coin metal is larger than that for the halogen bond. It is mainly due to the easier polarization of silicon than chlorine caused by the smaller electronegativity and bigger atomic radius of the silicon atom. The Ag substituent causes the interaction energy of halogen bond increased by \sim 63% with respect to -5.19 kcal mol⁻¹ in HCN···ClF,⁵⁸ while this atom strengthens the interaction energy of tetrel bond with 358%. Obviously, the Ag substituent is very efficient in enhancing the strength of tetrel bond.

It was demonstrated that C_2H_4 molecule acts as the π -electron donor to form a π-tetrel bond with SiH₃F and the corresponding interaction energy is -2.63 kcal mol^{-1.59} When SiF₄ acts as the σ -hole donor in the π -tetrel bond, the interaction energy is -1.93 kcal mol⁻¹, which is smaller than that in $C_2H_4\cdots SiH_3F$. However, the σ -hole on the silicon atom of SiF_4 is larger than that of SiH₃F. This inconsistence is mainly attributed to the nonzero dipole moment of SiH3F, which can greatly polarize the C₂H₄ molecule. In addition, the repulsion force between the π electrons of C_2H_4 and three fluorine atoms of SiF₄ has some contribution to this inconsistence. For the weak π -tetrel bond in $C_2H_4\cdots SiF_4$, both charge transfer and orbital interaction are very small, but dispersion energy is relatively large enough not to be ignored.

3.2. Cooperative effects

Fig. 4 shows the structures of triads, where both a tetrel bond and a π ···metal interaction coexist. The interaction in C_2H_4 ··· AuCN is very strong (>-50 kcal mol⁻¹).⁵⁷ Here we plan to strengthen the tetrel bond by cooperativity with a strong π ... metal interaction. The weak interaction strength is almost not changed in C2H4····AuCN···CF4. This implies that a van der Waals interaction is not strengthened prominently even in the presence of a very strong interaction. In C₂H₄···MCN···TF₄ (T = Si, Ge, and Sn), the tetrel bond is further strengthened, evidenced by the shorter binding distance and the more negative interaction energy (Table 5). The increase of interaction energy is related with the nature of coin metal and tetrel atom in $C_2H_4\cdots MCN\cdots TF_4$, and its increase is larger in the order AuCN < AgCN < CuCN and $CF_4 < SnF_4 < SiF_4 < GeF_4$. The former order is like that in halogen bonds.⁵⁶ The N···Si binding distance is shortened up to about 0.4 Å in C₂H₄···CuCN···SiF₄, and the interaction energy has the largest increase $(-9.60 \text{ kcal mol}^{-1})$ in this triad. The interaction energy of tetrel bond amounts to -7.12 kcal mol⁻¹ in C₂H₄····AuCN···SiF₄, where the strength of tetrel bond is doubled under the combinative influence of metal substituent and cooperativity. The interaction energy of tetrel bond is up to -15.72 kcal mol⁻¹ in $C_2H_4\cdots CuCN\cdots SiF_4$, where the tetrel bond varies from a moderate one (CuCN···SiF₄) to a strong one. The tetrel bond in C2H2···AuCN···SiF4 has a similar variation with that in C2H4···AuCN···SiF4. This indicates that the N···Si tetrel bond is greatly enhanced by

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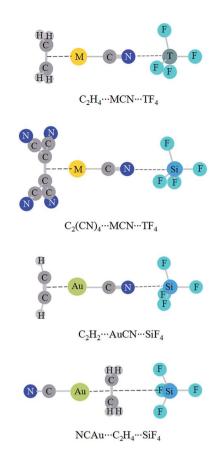


Fig. 4 Structures of triads studied.

Table 5 Binding distance (R_{TB} , Å) and interaction energy (ΔE_{TB} , kcal mol^{-1}) of tetrel bond in the triads and their difference (Δ) relative to the corresponding dyads

Triads	R_{TB}	ΔR_{TB}	ΔE_{TB}	$\Delta \Delta E_{\mathrm{TB}}$
arr a arr an	2.2440	0.0440		
$C_2H_4\cdots AuCN\cdots CF_4$	3.3119	-0.0149	-1.34	-0.02
$C_2H_4\cdots AuCN\cdots SiF_4$	2.5919	-0.2679	-7.12	-2.17
$C_2H_4\cdots AuCN\cdots GeF_4$	2.1169	-0.0649	-22.27	-2.67
$C_2H_4\cdots AuCN\cdots SnF_4$	2.2050	-0.0293	-28.34	-1.38
$C_2H_4\cdots AgCN\cdots SiF_4$	2.1415	-0.1397	-22.23	-6.43
$C_2H_4\cdots CuCN\cdots SiF_4$	2.2129	-0.4209	-15.72	-9.60
$C_2(CN)_4\cdots AuCN\cdots SiF_4$	2.9541	0.0943	-4.41	0.54
$C_2(CN)_4 \cdots AgCN \cdots SiF_4$	2.8102	0.5290	-6.37	9.43
$C_2(CN)_4\cdots CuCN\cdots SiF_4$	2.9483	0.3145	-3.31	2.81
$C_2H_2\cdots AuCN\cdots SiF_4$	2.6100	-0.2498	-6.92	-1.97
$NCAu \cdots C_2H_4 \cdots SiF_4$	3.5234	0.0354	-1.65	0.28

cooperativity. The influence of π ···metal interaction on the strength of tetrel bond is more prominent than that on halogen bond, where the increased percentage of interaction energy is less than 22%. The halogen than the four hydrogen atoms in C_2H_4 are replaced with electron-withdrawing group CN, the tetrel bond is weakened in $C_2(CN)_4$ ···MCN···SiF $_4$ and the largest weakening is found in $C_2(CN)_4$ ···AgCN···SiF $_4$. In NCAu···C $_2H_4$ ···SiF $_4$, where the middle molecule C_2H_4 plays a double Lewis base in the Au··· π interaction and π -tetrel bond, a weakening effect is also found.

The change in the strength of tetrel bond is also estimated with the electron density at the intermolecular BCP (Table 6). Against the electron density at the N···C BCP is almost not changed in C_2H_4 ···AuCN···CF₄. The electron density of tetrel bond is increased in C_2H_4 ···MCN···TF₄ (M = Cu, Ag, and Au; T = Si, Ge, and Sn) and C_2H_2 ···AuCN···SiF₄, while it is decreased in $C_2(CN)_4$ ···MCN···SiF₄ and NCAu···C₂H₄···SiF₄. With the change of tetrel bonding strength, its nature is also changed. For example, the tetrel bond has a nature of partially covalent interaction in C_2H_4 ···AuCN···SiF₄ and C_2H_2 ···AuCN···SiF₄, evidenced by the negative energy density, while it is a purely close-shell interaction in $C_2(CN)_4$ ···MCN···SiF₄ and NCAu···C₂H₄··· SiF₄, evidenced by the positive energy density.

Based on the dominant role of electrostatic interaction in the tetrel bond, it is natural to analyze the change of tetrel bonding strength in the triads with electrostatic potentials. Table 7 presents the most negative MEP on the N atom of MCN in $\pi\cdots$ MCN and on the C=C bond in NCAu···C₂H₄. This value is more negative in C₂H₄···MCN and C₂H₂···AuCN with respect to MCN but is less negative in C₂(CN)₄···MCN and NCAu···C₂H₄ with respect to MCN and C₂H₄. This shows that the former is a stronger Lewis base and the latter is a weaker Lewis base. As a result, the former forms a stronger tetrel bond and the latter forms a weaker tetrel bond.

The charge transfer of tetrel bond increases in $C_2H_4\cdots$ MCN···TF₄ (M = Cu, Ag, and Au; T = Si, Ge, and Sn) and $C_2H_2\cdots$ AuCN···SiF₄ but decreases in $C_2(CN)_4\cdots$ MCN···SiF₄ and NCAu··· $C_2H_4\cdots$ SiF₄ (Table S1†). The charge transfer has the largest increase in $C_2H_4\cdots$ CuCN···SiF₄ and the largest decrease in $C_2(CN)_4\cdots$ AgCN···SiF₄. The relationship between the change of charge transfer and the change of interaction energy is shown in Fig. 5. They display a good linear relationship with a correlation coefficient of 0.987. This indicates that charge transfer is also responsible for the change of tetrel bonding strength. Similarly, the orbital interaction has a similar change with the charge transfer.

Table 6 Electron density $(\rho_{\rm b})$ and energy density $(H_{\rm b})$ in the triads as well as the difference of electron density $(\Delta \rho_{\rm b})$ relative to the corresponding dyads, all are in au^a

Triads	$ ho_{ m b}$	$\Delta ho_{ m b}$	$H_{ m b}$
$C_2H_4\cdots AuCN\cdots CF_4$	0.0052	0.0002	0.0010
$C_2H_4\cdots AuCN\cdots SiF_4$	0.0193	0.0070	-0.0005
$C_2H_4\cdots AuCN\cdots GeF_4$	0.0619	0.0092	-0.0854
$C_2H_4\cdots AuCN\cdots SnF_4$	0.0647	0.0047	-0.0859
$C_2H_4\cdots AgCN\cdots SiF_4$	0.0416	0.0102	-0.0073
$C_2H_4\cdots CuCN\cdots SiF_4$	0.0351	0.0170	-0.0096
$C_2(CN)_4 \cdots AuCN \cdots SiF_4$	0.0105	-0.0018	0.0011
$C_2(CN)_4\cdots AgCN\cdots SiF_4$	0.0135	-0.0179	0.0008
$C_2(CN)_4\cdots CuCN\cdots SiF_4$	0.0107	-0.0074	0.0010
$C_2H_2\cdots AuCN\cdots SiF_4$	0.0187	0.0064	-0.0003
$NCAu \cdots C_2H_4 \cdots SiF_4$	0.0050	0.0200	0.0011

^a Note: the bond path is N···F in most complexes, but N···T in C_2H_4 ··· AuCN···GeF₄, C_2H_4 ···AuCN···SnF₄, and C_2H_4 ···AgCN···SiF₄, and C···F in NCAu···C₂H₄···SiF₄. All are average values for the N···F and C···F paths.

Table 7 The most negative MEP (V_{min}) on the N atom of MCN in $\pi \cdots$ MCN dyads and on the C=C bond in NCAu···C₂H₄ as well as its difference (ΔV_{min}) relative to the corresponding monomers, all are in leaf mol⁻¹

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C2(CN)4···AuCN

 $C_2H_2\cdots AuCN$

NCAu···C₂H₄

in kcal mol ⁻¹				
Dyads	$V_{ m min}$	$\Delta V_{ m min}$		
$C_2H_4\cdots CuCN$	-57.72	-3.77		
$C_2H_4\cdots AgCN$	-59.36	-3.84		
$C_2H_4\cdots AuCN$	-54.58	-5.90		
$C_2(CN)_4\cdots CuCN$	-43.90	10.05		
$C_2(CN)_4\cdots AgCN$	-47.86	7.66		

-38.88

-54.14

18.70

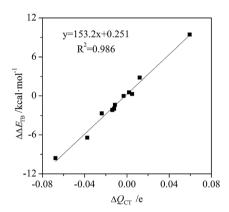


Fig. 5 Change of charge transfer ($\Delta Q_{\rm CT}$) versus change of interaction energy ($\Delta \Delta E$).

Table S2† presents the energy components of the tetrel bond in the triads to have a further insight into the origin of its interaction energy change. In most interactions, the main attractive contributions are from electrostatic and polarization, thus only their change is depicted in Fig. 6. For $C_2H_4\cdots MCN\cdots$

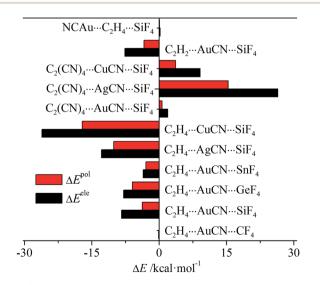


Fig. 6 Change of electrostatic and polarization energies in the triads relative to the respective dyads.

 TF_4 (M = Cu, Ag, and Au; T = Si, Ge, and Sn) and $C_2H_2\cdots AuCN\cdots SiF_4$, both terms are more negative, while they are more positive in $C_2(CN)_4\cdots MCN\cdots SiF_4$ and $NCAu\cdots C_2H_4\cdots SiF_4$. This indicates that both terms have important contributions in tuning the strength of tetrel bond through cooperativity. In all systems, electrostatic has the larger change than polarization, showing the larger contribution from electrostatic. Both terms are almost not changed in $C_2H_4\cdots AuCN\cdots CF_4$ and $NCAu\cdots C_2H_4\cdots SiF_4$, consistent with the small change of interaction energy in both triads.

4. Conclusions

9.80

-5.46

35.78

The triads of $C_2H_4\cdots AuCN\cdots TF_4$ (T=C, Si, Ge, and Sn), $C_2H_4\cdots MCN\cdots SiF_4$ (M=Cu, Ag, and Au), $C_2(CN)_4\cdots MCN\cdots SiF_4$, $C_2H_2\cdots AuCN\cdots SiF_4$, and $NCAu\cdots C_2H_4\cdots SiF_4$ as well as their dyads have been studied. The following conclusions are reached.

- (1) Coin metal substituents have a prominent enhancing effect on the strength of tetrel bond. This enhancing effect is related with the nature of coin metal and tetrel atoms, which is larger in the order C < Si < Ge < Sn and Au < Cu < Ag. The tetrel bond varies from a weak interaction in $HCN\cdots SiF_4$ (-3.45 kcal mol^{-1}) to a strong one in $AgCN\cdots SiF_4$ (-15.80 kcal mol^{-1}).
- (2) A π ···metal interaction can regulate the strength of tetrel bond. The tetrel bond is strengthened in C_2H_4 ···AuCN···TF $_4$ (T = C, Si, Ge, and Sn), C_2H_4 ···MCN···SiF $_4$ (M = Cu, Ag, and Au), and C_2H_2 ···AuCN···SiF $_4$ but is weakened in $C_2(CN)_4$ ···MCN···SiF $_4$ and NCAu···C $_2H_4$ ···SiF $_4$. Accompanied with the strengthening/weakening of tetrel bond, its nature could also be changed.

Conflicts of interest

There are no conflicts to declare.

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