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Electrochemical promotion of catalysis over Pd nanoparticles for CO_2 reduction†

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Electrochemical promotion of catalysis (EPOC) has been shown to accelerate the rate of many heterogeneous catalytic reactions; however, it has rarely been reported in low-temperature aqueous electrochemical reactions. Herein, we report a significant EPOC effect for the CO_2 reduction to generate formate over Pd nanoparticles (NPs) in a 1 M $KHCO_3$ aqueous solution. By applying a negative potential over differently-sized Pd NPs, the rate of formate production is greatly improved as compared to that at an open-circuit voltage, with a rate enhancement ratio ranging from 10 to 143. The thermocatalytic and electrocatalytic reduction of CO_2 compete with each other and are promoted by the applied negative potential and H_2 in the feeds, respectively. Inspired by the EPOC effect, a composite electrode containing Pd/C and Pt/C catalysts on different sides of a carbon paper was constructed for catalyzing the CO_2 reduction without adding H_2 to the feeds. Water electrolysis over Pt NPs generates H_2, which then effectively promotes formate production over Pd NPs.

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Introduction

Electrochemical promotion of catalysis (EPOC), discovered by M. Stoukides and C. Vayenas in the 1980s,[1,2] has been widely investigated in more than 100 heterogeneous catalytic reactions[3–6] on either metal or metal oxide surfaces, which are interfaced with a solid[7] or an aqueous electrolyte solution.[8–14] By applying an electrical current between the working electrode, coated with catalyst, and the counter electrode, the electronic properties of the supported catalyst can be tuned, accompanied by the alteration in the adsorption strength of the reactants, and in some cases, a significant enhancement in the catalytic performance can be observed. To date, only a few reports have demonstrated the EPOC effect in an aqueous electrolyte solution at ambient temperature for reactions such as the H_2 oxidation,[8,9] hydrocarbon isomerization,[10,11] CO oxidation,[12,13] and hydrazine oxidation.[14] Herein, we report that the EPOC effect can also be observed for electrochemical reduction reactions such as the reduction of CO_2 in an aqueous electrolyte solution at ambient temperature.

The reduction of carbon dioxide to produce formic acid is an attractive route to store renewable electricity and an important strategy for the utilization of carbon cycle.[7,15–19] However, CO_2 is thermodynamically stable and notoriously unreactive; therefore, high reaction temperatures[7,20–23] and high overpotentials[24–27] are usually essential to activate and transform CO_2 during thermocatalysis and electrocatalysis. Herein, we report a significant EPOC effect for the CO_2 reduction to produce formate over Pd nanoparticles (NPs) in a 1 M $KHCO_3$ aqueous solution at ambient temperature. Thermocatalytic and electrocatalytic reduction of CO_2 over Pd nanoparticles (NPs) occur simultaneously and compete with each other, which are promoted by the applied negative potential and H_2 in the feeds, respectively. The shared reaction intermediate, namely HCOO*, is formed over the Pd NPs and is proposed as the origin of the EPOC effect during the thermocatalytic and electrocatalytic reduction of CO_2. Inspired by the EPOC effect, a Pd/C–Pt/C composite electrode was constructed for the CO_2 reduction, such that the addition of H_2 to the feeds could be avoided. H_2 generated from the electrolysis of water over Pt NPs effectively promotes formate production over Pd NPs.

Results and discussion

Carbon-supported Pd NPs were prepared using sodium citrate as a stabilizing agent and $NaBH_4$ as a reducing agent.[28] Each sample was named on the basis of the average size of Pd NPs present in it, such as 3.7 nm Pd indicates that Pd NPs in the Pd/C catalyst have an average particle size of 3.7 nm (Fig. S1†). The catalyst ink, containing Pd/C and Nafion ionomer, was deposited on a piece of carbon paper (Toray TGP-H-060) with a microporous layer and dried to serve as a porous electrode for

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CO_2 reduction in a 20% H_2/CO_2-saturated 1 M $KHCO_3$ solution. To quantify the EPOC effect, the rate enhancement ratio was defined as follows:[3]

$$\rho = r/r_0 \quad (1)$$

where r and r_0 are the rates of the promoted (by applying negative potentials) and unpromoted (at open-circuit voltage, OCV) reactions, respectively. In this study, all r values measured under different atmospheres or over different electrodes are calculated by the following equation:

$$r = \frac{\text{total amount of formate generated in one hour (in mol)}}{\text{total amount of Pd in electrode (in milligram)} \times 1\ \text{h}} \quad (2)$$

CO_2 reduction experiments were conducted in an H-cell, as shown in Fig. 1a. The porous electrode coated with Pd/C catalyst was immersed in a 1 M $KHCO_3$ aqueous solution, and 20% H_2/CO_2 was fed into the cathode chamber for the CO_2 reduction reaction at OCV and different negative potentials. The maximum temperature increment of the electrolyte solution was 0.8 °C during the constant-potential electrolysis for 1 h, and the effect of temperature on the formate production rate at OCV and different negative potentials can be ignored. Fig. 1b shows the rate enhancement ratio for the formate production at different negative potentials compared to that at OCV over differently-sized Pd NPs. There are volcano-like curves for the value of ρ over differently-sized Pd NPs within the studied potential range. The value of ρ is 54 over 2.4 nm Pd at −0.1 V and reaches the maximum value of 143 at −0.2 V. Further negatively shifting the potential to −0.3 V and −0.4 V would decrease the rate enhancement ratio to 95 and 39, respectively. The ratio over 3.7 nm Pd increases from 58 to 119 when the potential is shifted from −0.1 V to −0.2 V and drops to 100 and 17 at −0.3 V and −0.4 V, respectively. The ratio over 7.8 nm Pd is obviously smaller than that over 2.4 nm Pd and 3.7 nm Pd, and the maximum ratio over 7.8 nm Pd is 23 at −0.2 V. The reaction of adsorbed hydrogen on Pd hydride surface with CO_2 to form adsorbed HCOO* is considered as the rate-determining step for CO_2 reduction.[16] Upon negatively shifting the potential from −0.1 V to −0.4 V, the hydrogen adsorption strength on the Pd hydride surface is weakened due to a favored hydrogen evolution reaction. Therefore, the optimum adsorption strength for surface-adsorbed hydrogen to react with CO_2 occurs at −0.2 V, resulting in the highest ρ value. Since small-sized Pd NPs prefer to adsorb more hydrogen over coordinatively unsaturated sites,[29] the ρ value is much higher over 2.4 nm Pd as compared to that over 3.7 nm Pd and 7.8 nm Pd at −0.4 V.

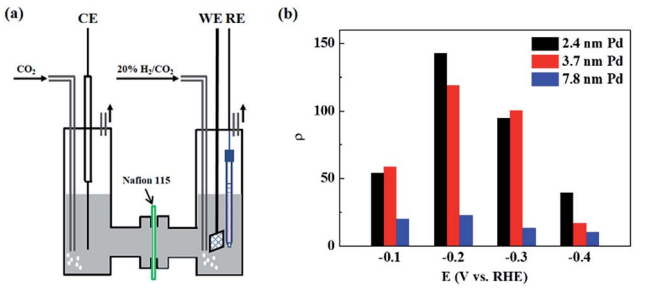

Fig. 1 (a) Schematic of the H-cell for CO_2 reduction. (b) Rate enhancement ratios for formate production at different negative potentials compared to those at OCV over differently-sized Pd NPs in 20% H_2/CO_2-saturated 1 M $KHCO_3$ solution.

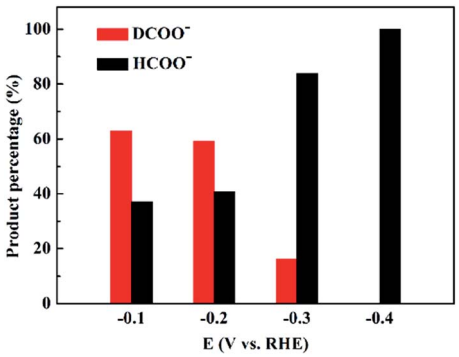

Fig. 2 $HCOO^-$ and $DCOO^-$ percentages for the CO_2 reduction over 3.7 nm Pd at different negative potentials in 20% D_2/CO_2-saturated 1 M $KHCO_3$ solution.

To investigate the origin of the EPOC effect during CO_2 reduction, an isotope-labeling experiment was conducted by replacing 20% H_2/CO_2 with 20% D_2/CO_2, and the products were analyzed by nuclear magnetic resonance (NMR) spectroscopy. PdD_x could also be generated under D_2 atmosphere, and the properties of PdD_x were close to those of PdH_x; thus, it was suggested that the electrochemical measurements under 20% D_2/CO_2 atmosphere were comparable to those under 20% H_2/CO_2 atmosphere.[30–32] $HCOO^-$ and $DCOO^-$ were quantified by ^{1}H-NMR and ^{2}H-NMR spectra, respectively. The amount of D_2O and HDO in the electrolyte solution after constant-potential electrolysis at −0.2 V for 1 h was also quantified by ^{2}H-NMR spectra, which is 0.028%, whereas the natural abundance of D is about 0.015%.[33] Therefore, non-electrochemical exchange between adsorbed D and H^+ could be ignored,[34] and $DCOO^-$ and $HCOO^-$ were considered to be produced from the $CO_2 + D_2$ thermocatalytic reaction and CO_2 electrocatalytic reaction, respectively. Fig. 2 shows the percentage of $HCOO^-$ and $DCOO^-$ formed over 3.7 nm Pd at different negative potentials. The percentage of $DCOO^-$ is higher than that of $HCOO^-$ at −0.1 V and −0.2 V, which sharply decreases at −0.3 V and reaches below the detection limit at −0.4 V. Thus, the thermocatalytic and electrocatalytic reduction of CO_2 occur simultaneously and compete with each other when applying negative potentials over Pd NPs.

We also measured the electrocatalytic reduction of CO_2 without adding H_2 to the feeds. As shown in Fig. 3, S2 and S3,† the current density becomes unstable at negatively shifted potentials, which is caused by poisoning from trace CO, a minor side product from the CO_2 electroreduction.[16,35–37] With the addition of H_2, the stability of current density increases, indicating that the electrocatalytic reduction of CO_2 is stabilized. X-ray diffraction (XRD) patterns, of 3.7 nm Pd under different atmosphere were obtained to investigate the active phase of Pd

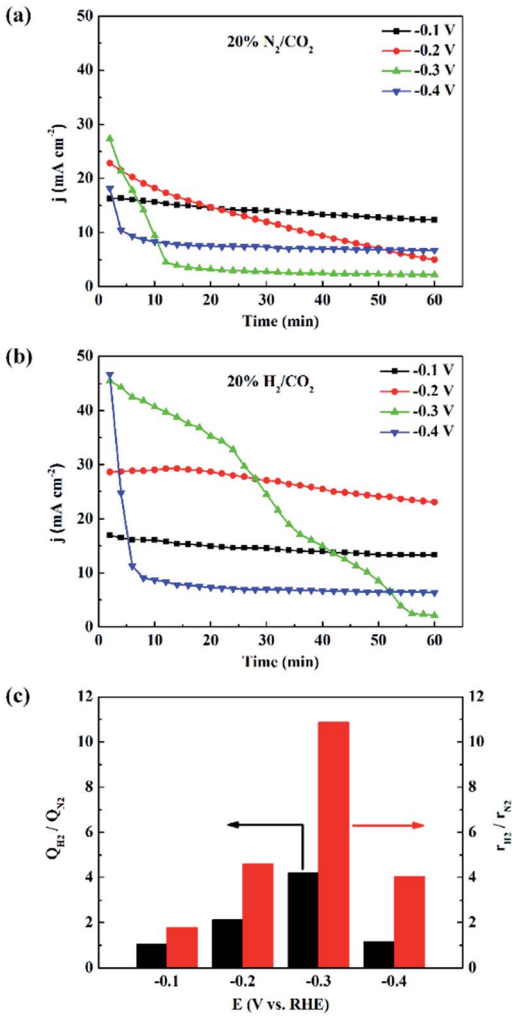

Fig. 3 Chronoamperometry curves of 3.7 nm Pd in 20% N_2/CO_2-saturated (a) and 20% H_2/CO_2-saturated (b) 1 M $KHCO_3$ solution. (c) Enhancement of electric charge (Q_{H_2}/Q_{N_2}) ratio and formate production rate ratio (r_{H_2}/r_{N_2}) over 3.7 nm Pd at different negative potentials in 20% N_2/CO_2 and 20% H_2/CO_2-saturated 1 M $KHCO_3$ solutions.

NPs, as shown in Fig. S4.† XRD pattern of 3.7 nm Pd under 20% N_2/CO_2 atmosphere matches well with that of Pd (JCPDS 46-1043), and the diffraction peak of the Pd (111) plane is located at 39.9°. When the atmosphere is switched to 20% H_2/CO_2, the diffraction peak of the Pd (111) plane quickly shifts to 38.8°, accompanied by a shift in all the other peaks. The pattern is consistent with that of $PdH_{0.706}$ (JCPDS 18-0951), which is facilely generated under H_2 atmosphere. Since the state and structure of the supported Pd NPs are not affected by water,[38] the PdH_x active phase is expected to be stable in 20% H_2/CO_2-saturated 1 M $KHCO_3$ solution.

Fig. 3c shows the electric charge ratios over 3.7 nm Pd at various potentials, calculated from the *I–t* plots, in Fig. 3a and b. At −0.1 V, the electric charge does not change over Pd NPs because surface PdH_x is stable and the rate of proton reduction can meet the requirement for the CO_2 electroreduction. At −0.2 V, CO_2 electroreduction is accelerated, and the rate of proton reduction cannot match the rate of CO_2 electroreduction, resulting in a decrease of current density. The enhancement of the electric charge ratio is the highest at −0.3 V since the addition of H_2 effectively stabilizes surface PdH_x that tends to decompose through hydrogen evolution. This enhancement of electric charge ratio indicates that the electrocatalysis is to some extent also promoted by H_2 in the feeds, which has not been reported earlier.

The rate of formate production was enhanced when the reaction atmosphere was changed from 20% N_2/CO_2 to 20% H_2/CO_2. Fig. 3c shows the enhancement ratio of r_{H_2}/r_{N_2} over 3.7 nm Pd at different negative potentials, where r_{H_2} and r_{N_2} represent the formate production rates in 20% H_2/CO_2 and 20% N_2/CO_2-saturated electrolyte solutions, respectively. The enhancement ratio of formate production rates is higher than that of electric charge accumulated in 1 h. For instance, r_{H_2} at −0.2 V reaches about 1.9 $mol_{formate}\ mg_{Pd}^{-1}\ h^{-1}$, and r_{H_2}/r_{N_2} is more than twice that of Q_{H_2}/Q_{N_2}. The additional improvement is considered as a contribution from the thermocatalytic reaction,[39] also identified in the isotope labeling experiment. The significant potential dependence of the r_{H_2}/r_{N_2} and Q_{H_2}/Q_{N_2} values over 3.7 nm Pd is attributed to the instability of Pd hydride and CO poisoning at more negative potentials during the electrocatalytic reduction of CO_2. Similar enhancement effects over 2.4 nm and 7.8 nm Pd are shown in Fig. S2 and S3.† Since the strong hydrogen adsorption on small-sized Pd NPs could stabilize the surface Pd hydride, the current densities over 2.4 nm Pd were more stable than those over 3.7 nm Pd (Fig. 3 and S2†). Therefore, the r_{H_2}/r_{N_2} and Q_{H_2}/Q_{N_2} values over 2.4 nm Pd are smaller than those over 3.7 nm Pd. The lack of potential dependence of r_{H_2}/r_{N_2} and Q_{H_2}/Q_{N_2} over 7.8 nm Pd indicates that surface Pd hydride is difficult to form over large-sized Pd NPs, as also confirmed by the unstable current densities shown in Fig. S3.†

Based on the abovementioned results, thermocatalytic and electrocatalytic reduction of CO_2 occur simultaneously as follows:[40–42]

(a) Thermocatalytic reduction reaction

$$Pd + xH_2(g) \leftrightarrow H^* \quad (3)$$

$$CO_2(g) + H^* \leftrightarrow HCOO^* \quad (4)$$

$$HCOO^* + H \leftrightarrow HCOOH^* \quad (5)$$

$$HCOOH^* \leftrightarrow HCOOH(l) + * \quad (6)$$

(b) Electrocatalytic reduction reaction

$$CO_2(g) + * + e^- + H^+(aq.) \leftrightarrow HCOO^* \quad (7)$$

$$HCOO^* + e^- + H^+(aq.) \leftrightarrow HCOOH^* + H_2O(l) \quad (8)$$

$$HCOOH^* \leftrightarrow HCOOH(l) + * \quad (9)$$

where * and H represent the palladium hydride active phase and the free H atom adsorbed on the catalyst surface, respectively. HCOO* and HCOOH* represent the corresponding

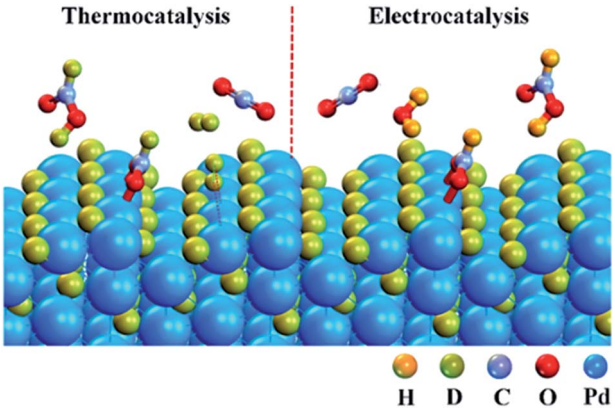

Fig. 4 Schematic of heterogeneous thermocatalytic (left) and electrocatalytic (right) reduction of CO_2 over Pd NPs conducted in a CO_2 + D_2 atmosphere.

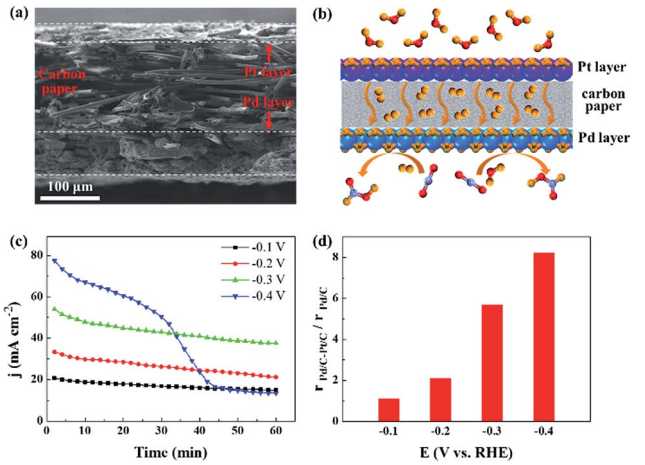

Fig. 5 CO_2 electroreduction over the Pd/C–Pt/C composite electrode in 20% N_2/CO_2-saturated 1 M $KHCO_3$ solution. (a) Cross-sectional SEM image of the Pd/C–Pt/C composite electrode. (b) Schematic for the reaction mechanism in the Pd/C–Pt/C composite electrode. (c) Chronoamperometry curves of the Pd/C–Pt/C composite electrode at different negative potentials. (d) Rate enhancement ratio for formate production over the Pd/C–Pt/C composite electrode *versus* the Pd/C electrode at different negative potentials.

intermediate species. It is clear from eqn (4) and (7) that the thermocatalytic and electrocatalytic reduction of CO_2 share the same intermediate species HCOO*. Recent research on hydrazine electrooxidation indicates that the mechanistic origin of the EPOC effect lies in structurally similar activated transition states and/or adsorbed surface intermediates arising from hydrazine oxidation and decomposition.[14] Therefore, the negative electrode potentials could affect and promote the heterogeneous catalytic reduction of CO_2.

Fig. 4 shows the schematic of CO_2 + D_2 reduction over Pd NPs. D_2 is split into D atoms on the surface of Pd NPs, which subsequently diffuse into the Pd lattice to form the PdD_x phase. The atomic D on the PdD_x surface, derived from D_2 dissociation, reacts with CO_2 to form $DCOO^-$ *via* the thermocatalytic pathway, whereas H^+ in water reacts with CO_2 to form $HCOO^-$ *via* the electrocatalytic pathway. At −0.1 V and −0.2 V, the electrocatalytic reduction rate is constrained due to the low overpotentials and the percentage of $HCOO^-$ is lower than that of $DCOO^-$. Upon increasing the overpotential, the electrocatalytic reduction rate is accelerated, exceeding the thermocatalytic rate, and thus the electrocatalytic reduction pathway dominates for CO_2 reduction at −0.4 V (Fig. 2).

H_2 generated from the electrolysis of water is promising for practical applications in CO_2 reduction. Within the potential range for the CO_2 reduction *via* CO_2 + H_2 reaction, the hydrogen evolution reaction can occur over Pt NPs, which would provide H_2 by the electrolysis of water at the same negative potentials.[43,44] We designed a composite electrode in which 3.7 nm Pd was deposited on the microporous layer, whereas commercial Pt/C catalyst was deposited on the other side of the carbon paper. The cross-sectional scanning electron microscopy (SEM) image and corresponding energy-dispersive X-ray spectroscopy mapping image of the electrode are shown in Fig. 5a and S5.† As illustrated in Fig. 5b, the Pt/C and Pd/C catalyst layers are separated by the carbon paper with a microporous layer, and H_2 generated over Pt NPs diffuses across the carbon paper or the electrolyte solution towards the Pd NPs, which promotes the thermocatalytic and electrocatalytic reduction of CO_2 over Pd NPs. The current density is clearly improved and stabilized after depositing Pt/C catalyst on the opposite side of the Pd/C catalyst layer (Fig. 3a and 5c). The activity and selectivity of the Pt/C electrode, used as a control in this study, were also measured, and only H_2 was produced (Fig. S6†). The enhancement ratio for formate production with the Pd/C–Pt/C composite electrode is shown in Fig. 5d. The maximum value reaches 8.2 at −0.4 V as a result of the combination of thermocatalytic and electrocatalytic reduction of CO_2.

Conclusions

In summary, a significant EPOC effect was observed over Pd NPs during CO_2 reduction to generate formate in 1 M $KHCO_3$ solution at ambient temperature. Both thermocatalytic and electrocatalytic reduction of CO_2 over Pd NPs were promoted by applying negative potentials and by adding H_2 to the feeds, respectively. The shared reaction intermediate HCOO* over Pd NPs was proposed as the origin of the EPOC effect during the thermocatalytic and electrocatalytic reduction of CO_2. Based on the abovementioned understanding, the Pd/C–Pt/C composite electrode was constructed for CO_2 reduction without the direct addition of H_2 to the feeds. H_2 generated through water electrolysis over Pt NPs effectively promoted the formate production over Pd NPs. The significant rate enhancement ratio for CO_2 reduction not only reveals a new example of EPOC in a low-temperature aqueous electrochemical reaction, but also provides an alternative strategy to promote the electrocatalytic reduction of CO_2.

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