

Cite this: *Chem. Sci.*, 2017, 8, 3989

Adsorption and molecular siting of CO₂, water, and other gases in the superhydrophobic, flexible pores of FMOF-1 from experiment and simulation†

Peyman Z. Moghadam,^{‡a} Joshua F. Ivy,^{‡b} Ravi K. Arvapally,^{‡b} Antonio M. dos Santos,^c John C. Pearson,^{bf} Li Zhang,^{ad} Emmanouil Tylanakis,^e Pritha Ghosh,^a Iain W. H. Oswald,^b Ushasree Kaipa,^b Xiaoping Wang,^{*c} Angela K. Wilson,^{*bf} Randall Q. Snurr^{*a} and Mohammad A. Omary^{*b}

FMOF-1 is a flexible, superhydrophobic metal–organic framework with a network of channels and side pockets decorated with –CF₃ groups. CO₂ adsorption isotherms measured between 278 and 313 K and up to 55 bar reveal a maximum uptake of ca. 6.16 mol kg^{−1} (11.0 mol L^{−1}) and unusual isotherm shapes at the higher temperatures, suggesting framework expansion. We used neutron diffraction and molecular simulations to investigate the framework expansion behaviour and the accessibility of the small pockets to N₂, O₂, and CO₂. Neutron diffraction *in situ* experiments on the crystalline powder show that CO₂ molecules are favourably adsorbed at three distinct adsorption sites in the large channels of FMOF-1 and cannot access the small pockets in FMOF-1 at 290 K and oversaturated pressure at 61 bar. Stepped adsorption isotherms for N₂ and O₂ at 77 K can be explained by combining Monte Carlo simulations in several different crystal structures of FMOF-1 obtained from neutron and X-ray diffraction under different conditions. A similar analysis is successful for CO₂ adsorption at 278 and 283 K up to ca. 30 bar; however, at 298 K and pressures above 30 bar, the results suggest even more substantial expansion of the FMOF-1 framework. The measured contact angle for water on an FMOF-1 pellet is 158°, demonstrating superhydrophobicity. Simulations and adsorption measurements also show that FMOF-1 is hydrophobic and water is not adsorbed in FMOF-1 at room temperature. Simulated mixture isotherms of CO₂ in the presence of 80% relative humidity predict that water does not influence the CO₂ adsorption in FMOF-1, suggesting that hydrophobic MOFs could hold promise for CO₂ capture from humid gas streams.

Received 19th January 2017

Accepted 9th March 2017

DOI: 10.1039/c7sc00278e

rsc.li/chemical-science

Introduction

Capturing carbon from flue gas is an important challenge, as fossil fuel combustion continues to be a primary source of energy.^{1,2} The search for adsorbents capable of capturing large amounts of CO₂ has led to many studies of adsorption in

nanoporous materials known as metal–organic frameworks (MOFs).^{3–5} MOFs have been shown to be promising for a number of separation applications including CO₂ capture at the low partial pressures relevant to flue gas.^{6–11} In an early example, Yazaydin *et al.* screened 14 MOFs and concluded that the M/DOBDC series shows exceptionally high CO₂ capacity at room temperature.¹² This can be attributed to the high density of open metal sites, and other studies on CO₂ capture in MOFs have shown similar results.¹³ After the CO₂ is captured, it must be stored or used afterward. One option is permanent subterranean storage as a pressurized liquid. Alternatively, it could be converted and used as other chemical products. Darensbourg *et al.* recently reported CO₂ capture in a MOF, HKUST-1, to perform a copolymerization with propylene oxide with a 49.9% conversion rate.²

When assessing adsorbents for use with flue gas, the presence of water vapor cannot be ignored.⁹ For example, it has been shown that M/DOBDC MOFs show a significant decrease in CO₂ capture capacity under humid conditions.^{14,15} Hydrophobic MOFs could be an attractive alternative, due to their ability to

^aDepartment of Chemical & Biological Engineering, Northwestern University, 2145 Sheridan Road, Evanston, IL 60208-3120, USA. E-mail: snurr@northwestern.edu

^bDepartment of Chemistry, University of North Texas, Denton, Texas 76203, USA. E-mail: wilson@chemistry.msu.edu; omary@unt.edu

^cNeutron Sciences Directorate, Oak Ridge National Laboratory, Oak Ridge, TN 37831, USA. E-mail: wangx@ornl.gov

^dDepartment of Chemistry, Zhejiang Sci-Tech University, Hangzhou, China

^eDepartment of Materials Science & Technology, University of Crete, Voutes Campus, Heraklion, Crete GR-71003, Greece

^fDepartment of Chemistry, Michigan State University, East Lansing, MI 48824-1322, USA

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c7sc00278e

‡ P. Z. M., J. F. I., and R. K. A. contributed equally (co-1st authors).



withstand humid conditions and suppress competitive adsorption of water. A number of studies in the literature have investigated the effects of fluorination and hydrophobicity in MOFs.^{16–22} FMOF-1 is a fluorinated metal–organic framework first synthesized by Yang *et al.*²⁰ It is formed by the reaction of a perfluorinated ligand (3,5-bis(trifluoromethyl)-1,2,4-triazolate (Tz^-)) with a Ag^+ precursor, leading to $\{Ag_2[Ag_4Tz_6]\}$. FMOF-1 exhibits a perfluorinated structure, and the many CF_3 groups lining its channels and small pockets imbue it with hydrophobicity.²⁰ In principle, the CO_2 quadrupole should be able to interact with the polar C–F groups in FMOF-1, but CO_2 adsorption measurements have not been reported for FMOF-1. In addition, while FMOF-1 has been shown to experience enormous breathing behaviour as a function of temperature either under vacuum or in the presence of N_2 ,²³ the evolution of the FMOF-1 structure has not been studied previously in the presence of other guest molecules or at elevated pressures. Given that changes in framework structure can have a significant effect on the adsorption properties of a porous material, we investigate in this work whether framework flexibility is a general feature of FMOF-1 and the effect of flexibility on adsorption uptake by investigating a wide variety of guest molecules, with a particular focus on CO_2 .

In this paper, experimental CO_2 adsorption isotherms, contact angle measurement of water drops, and *in situ* neutron diffraction results during CO_2 adsorption are reported. Also reported are grand canonical Monte Carlo and quantum mechanical simulations in FMOF-1 to model the adsorption of CO_2 and other guest molecules, including adsorption of CO_2 under humid conditions. The work illustrates the power of a strong feedback loop between experiment and modeling. For example, neutron diffraction studies provided possible crystal structures for simulations of CO_2 adsorption, and measured isotherms and heats of adsorption provided validation of predictions from modeling. In turn, modeling provided insight about molecular siting in FMOF-1 and predictions about the CO_2 capture performance under humid conditions. The flexible nature of FMOF-1 was investigated *via* simulation using four different FMOF-1 structures obtained under different experimental conditions. Correlations between the framework structure and guest uptake were established for three classes of guest molecules, including diatomics at one extreme and bulky hydrocarbons at the other, with CO_2 representing an intermediate category.

Methods

FMOF-1 synthesis and adsorption measurements

FMOF-1 was prepared using previously published methods.^{20,23} Nitrogen adsorption isotherms at 77 K were measured with a Micromeritics ASAP 2020. CO_2 adsorption measurements were carried out by a VTI/TA Gravimetric High Pressure Sorption Analyzer. This VTI/TA system is equipped with ultra-high vacuum and is capable of variable temperature measurements from -196 °C to 1000 °C. It has a flow dosing manifold for high pressure studies and achieves 0.1 μg resolution with a CI Electronics microbalance. Typically, 100 mg of sample was used for

adsorption measurements. Before each measurement, the sample was purged with helium then evacuated for 60 minutes at 60 °C. Measurements were performed at 5 degree intervals from 5 ° to 40 °C and pressures up to 53 bar or the critical pressure of CO_2 at the set temperature. Isothermic heats of adsorption were derived using a set of isotherms at different temperatures and the Clausius–Clapeyron equation.²⁴ High purity CO_2 gas was used for the adsorption studies.

Contact angle measurements

Contact angle measurements were done using a Ramehart manual goniometer (Model # 50-00-1150). Static contact angle was measured. A single drop of water was added using the syringe attached to the goniometer and then the contact angle was measured on the static sessile drop with the gauge provided.

Neutron powder diffraction measurements

In situ neutron powder diffraction measurements were performed at the SNAP beamline of the Spallation Neutron Source (SNS) at Oak Ridge National Laboratory (ORNL). SNAP is a high flux and medium resolution time-of-flight diffractometer, with tunable detector placement and incident energy range. For this experiment, measurements were made with the two detector banks placed at 90 and 48 degrees, and the wavelength band used was 3.5 Å wide and centered at 6.4 Å. This configuration enabled sampling of Bragg reflections in the range 3 – 17 Å, enough to sample the longest (in d -spacing) reflections of the sample. Fully activated FMOF-1 powder sample was loaded into a gas cell fabricated with a null scattering TiZr alloy (1 : 2.08 Zr : Ti molar ratio) and warmed to 320 K in a dynamic vacuum. The gas cell was then cooled *via* 200 mbar of He exchange gas on a top loading cold cycle refrigerator cryostat to 290 K for data collection on the bare FMOF-1. CO_2 was then loaded slowly into the FMOF-1 sample at 290 K using a computer-controlled automated gas handling system and held for an hour at 61 bar to ensure the sample cell maintained over the CO_2 saturation pressure (53.2 bar) before and during data collection. High purity CO_2 gas stored at room temperature (296 K) was used directly from the cylinder with no further purification.

Neutron diffraction data were analyzed using the GSAS II package.²⁵ The locations of CO_2 molecules in FMOF-1 after CO_2 loading were obtained from difference Fourier map and refined accordingly with distance constraints.

Simulation details

Grand canonical Monte Carlo (GCMC) simulations were employed to investigate the adsorption of N_2 , O_2 , CO_2 , H_2O , n -hexane, and benzene. For each pressure point of the isotherm, 1×10^5 GCMC cycles were used for equilibration, after which another 1×10^5 cycles were used to calculate the average properties. For water simulations, we used 5×10^5 cycles each for equilibration and production runs. Each GCMC cycle is made up of N steps, where N is the number of adsorbates in the simulation box. (The number of steps per cycle is not allowed to be lower than 20; so if there are fewer than 20 adsorbates in the



simulation box, a cycle consists of 20 steps.) The TraPPE force field was used to model all adsorbates (N_2 ,²⁶ O_2 ,²⁷ CO_2 ,²⁶ *n*-hexane,²⁸ and benzene²⁹) except for water, which was described with the TIP4P model.³⁰ Lennard-Jones parameters for the framework atoms were taken from the Universal Force Field.³¹ Cross Lennard-Jones parameters were determined by Lorentz–Berthelot mixing rules. The partial charges for CF_3 groups were adopted from the work of Dalvi *et al.*,³² while partial charges for the rest of the framework were obtained from density functional theory calculations at the B3LYP level of theory using the ChelpG method.³³ Lennard-Jones parameters and partial atomic charges for the adsorbates and FMOF-1 are all listed in the ESI (Tables S3 and S4†). A cutoff distance of 12.8 Å was used for all Lennard-Jones interactions, and tail corrections were neglected. Long-range electrostatic interactions were accounted for using the Ewald summation method. The simulation box was constructed of 4 ($2 \times 2 \times 1$) unit cells with periodic boundary conditions applied in all directions.

Framework atoms were held fixed during the GCMC simulations. To explore the effect of framework flexibility on the adsorption properties, we performed simulations on three different FMOF-1 crystal structures along with the related FMOF-2 polymorph (see Fig. S12 in the ESI†) for CO_2 adsorption at room temperature. The first published crystal structure of FMOF-1 was obtained under vacuum at 100 K,²³ and this structure shall be referred to as FMOF-1a in this work. A structure obtained under a nitrogen stream at 90 K (ref. 23) shall be referred to here as FMOF-1b. The third FMOF-1 structure is a heretofore-unpublished structure obtained under a carbon dioxide stream at 61 bar and 290 K and will be referred to as FMOF-1c.

Quantum chemical methods

Density functional theory (DFT) was used to determine the binding free energy of molecules at their most likely adsorption sites. The binding sites in the cylindrical channel and small cavity were simulated individually. The cylindrical channel was truncated from the FMOF-1a structure and was composed of 11 Tz ligands and 9 Ag atoms. The small pocket was truncated from the FMOF-1b structure containing 14 Tz ligands and 10 Ag atoms. Both sites fully accounted for the correct coordination geometry of all metal centers using the neutral singlet state. Equilibrium geometries for guest molecules were found, holding the framework static. The NWChem software package was used for all quantum chemical calculations.³⁴ We chose to use the BPE0 functional with 6-311G* Pople basis sets for non-metal atoms and the Stuttgart-97 Effective Core Potential and respective basis sets for silver.^{35–38} The empirical dispersion correction DFT-D3 was added to address the long-range effects.³⁹ The inclusion of dispersion corrections with DFT has been shown to be necessary for the prediction of MOF structures, creating force field parameters, predicting changes in structure, and determining water adsorption sites.^{40–44} Orbital comparisons were made using the Mulliken population analysis. The total population density of the empty *versus* the occupied site used the same basis set and framework coordinates.

Results and discussion

CO_2 adsorption and contact angle studies

Fig. 1 shows the CO_2 adsorption isotherms for FMOF-1 at near-ambient temperatures up to 55 bar. The maximum uptake experimentally measured is *ca.* 6.16 mol kg^{-1} (11.0 mol L^{-1} ; 27.1 wt%; 483 kg m^{-3} ; 248 V STP V^{-1}) at 298 K and 55 bar. This uptake is more than 2 times higher than that predicted⁴⁵ in 2008 based on the rigid structure of FMOF-1 reported in 2007,²⁰ necessitating further modeling that takes into account the framework flexibility (*vide infra*). The isotherms do not reach a plateau and are inconsistent with a type I adsorption behaviour over this temperature and pressure range, again suggesting flexibility of the framework. This behaviour was determined not to be due to a systematic error in the experimental data, as it has been reproduced multiple times; see Fig. S2 (ESI†). Up to 30 bar, the isotherms can be fit very well with a Toth isotherm (Fig. S1†), and above this pressure the isotherms continue to rise, with a visible inflection at some temperatures.

Upon removing the sample from the instrument after the CO_2 adsorption studies, we noticed that the powder sample had transformed into a flat, yellow-coloured pellet. TGA and IR studies for pieces of such a pellet revealed essentially identical profiles to the powder sample of FMOF-1. We performed a contact angle experiment for a water droplet upon this pellet. This process proved difficult, as the water droplets tended to bounce and deflect off of the surface of the pellet rather quickly and as complete spheres, suggesting a rather extreme superhydrophobic behaviour of the material. Processing of the resulting image yielded a contact angle of ~ 158 , as shown in Fig. 2, clearly indicating a superhydrophobic behaviour and consistent with the insignificant water adsorption reported earlier for FMOF-1.^{46–48} The work herein, therefore, ascertains that FMOF-1 belongs to the “superhydrophobic” category of MOFs, similar to very few other frameworks verified as such from contact angle measurements.^{22,49,50}

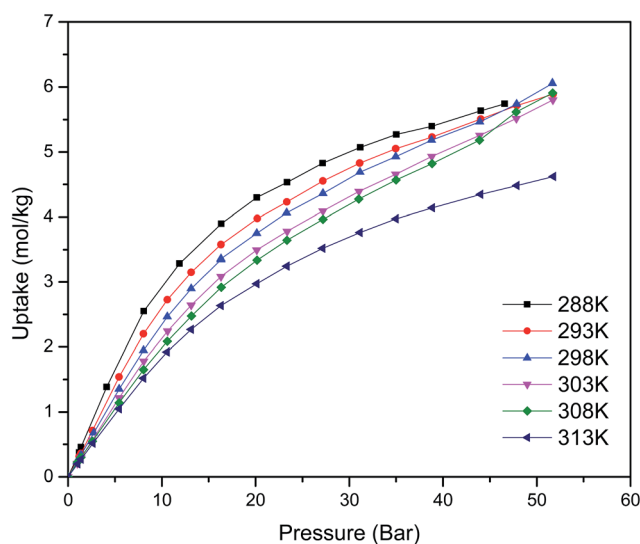


Fig. 1 CO_2 excess adsorption isotherms of FMOF-1 at various near-ambient temperatures.



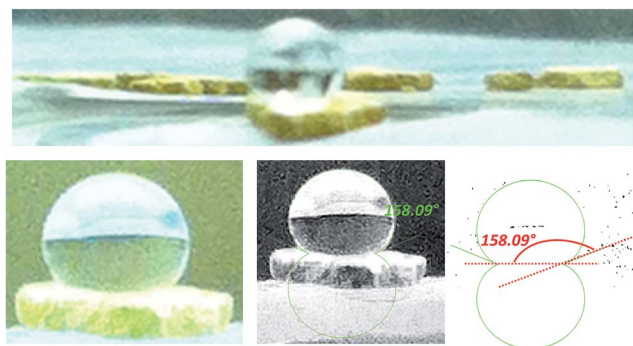


Fig. 2 Images of a water drop on a pellet of FMOF-1 that formed after the high-pressure CO₂ adsorption isotherm experiment shown in Fig. 1; most water droplets in the attempted contact angle experiments deflected off the surface of the pellet quickly as complete spheres. The bottom-right and bottom-middle images show the processing of the bottom-left raw image using the LBADSA plugin for the ImageJ software.⁵¹

Neutron diffraction studies

Neutron powder diffraction measurements were performed at the ORNL Spallation Neutrons and Pressure Diffractometer (SNAP) (Table S1[†]). Fig. 3a compares the neutron powder diffraction patterns for the evacuated FMOF-1 and the structure under 61 bar of CO₂ at 290 K. Under this oversaturated pressure condition, the free CO₂ molecules stay in the liquid state. The increase in intensity for the 011 peak at $\sim 13 \text{ \AA}$ after *in situ* CO₂ loading is clearly discernible. The FMOF-1 sample with adsorbed CO₂ was cooled to 230 K and the pressure was reduced to 4.8 bar for low-temperature measurements. The pressure in the TiZr cell was controlled *via* a computer-controlled gas handling system. Fig. S4 (ESI[†]) shows the evolution of the neutron powder diffraction patterns of the FMOF-1 sample with adsorbed CO₂. No evidence of solid CO₂ was observed in the structural refinement of the CO₂ loaded samples, indicating that all CO₂ was adsorbed into FMOF-1.

Fig. 3b shows the temperature dependence of the unit cell volume of FMOF-1 at a static pressure of 4.8 bar of CO₂ loading

and compares it with the results obtained under N₂ stream²³ and vacuum.²³ The smooth change in the unit cell volume in CO₂ (open squares) is in contrast to that of FMOF-1 with adsorbed N₂ (red circles in Fig. 3b), in which the loading of N₂ molecules into small cavities of FMOF-1 at temperatures below 119 K causes a huge negative thermal expansion in the crystal structure.²³

FMOF-1 has two types of pores. The first type are large cylindrical channels extending along both the *x*- and *y*-directions (Fig. 4a and b), with CF₃ groups protruding into the channels. The second pore type is a small cavity, and two pairs of CF₃ groups function as a gate between the small cavity and the large cylindrical channel (see the red circle in Fig. 4c). The pore size distribution and full geometric characterization of FMOF-1 can be found in Fig. S5 and Table S2.[†]

The favourable adsorption sites for CO₂ molecules in FMOF-1 were obtained from difference Fourier map and with the CO₂ molecules refined with distance constraints. Difference Fourier map from initial Rietveld refinements indicate that the CO₂ molecules are located only in the large channels of FMOF-1 at three unique sites. The FMOF-1 structure loaded with CO₂ molecules at 61 bar and 290 K is shown in Fig. 5. The oxygen atoms in the three primary CO₂ adsorption sites are shown as cyan spheres at site I near the framework –CF₃ groups at corners; as red spheres at site II near –CF₃ groups along the crystallographic *c* direction; and as orange spheres at site III in the direction of the large channels. Since there is no CO₂ in the small cavities, when all three CO₂ adsorption sites are fully occupied, there are 24 CO₂ molecules in the unit cell (3.3 mol kg^{−1}). This limits the CO₂ uptake capacity in the large cavity to 6 CO₂ molecules per {Ag₂[Ag₄Tz₆] repeat unit.

Molecular simulation studies and connection to experimental data

In order to verify our simulation parameters, we began by reproducing previously published experimental isotherms for benzene, *n*-hexane, and water in FMOF-1. Fig. 6 compares simulated results in FMOF-1a to the experimental isotherms.

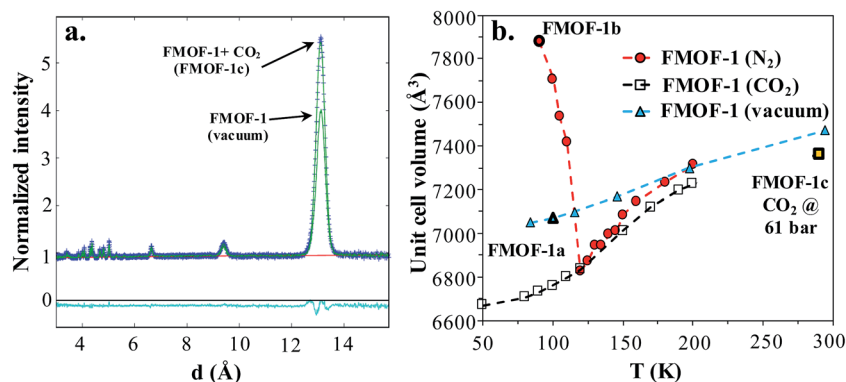


Fig. 3 (a) FMOF-1 neutron powder diffraction patterns measured at 290 K in vacuum (solid green line) and under 61 bar of CO₂ (blue crosses). The residuals (cyan) underneath the zero line are the difference of the observed and refined neutron diffraction profiles of the FMOF-1 sample with adsorbed CO₂. $wR = 1.606\%$, $GOF = 1.87$, $N_{\text{obs}} = 1658$, $N_{\text{vals}} = 14$. Space group $I42d$, $a = 13.9713(7) \text{ \AA}$, $c = 37.713(4) \text{ \AA}$, $V = 7361.4(7) \text{ \AA}^3$. (b) Temperature dependence of unit cell volume of FMOF-1 under vacuum,²³ constant stream of N₂ at atmospheric pressure²³ and CO₂ at 4.8 bar. The unit cell volume of FMOF-1 under 61 bar and 290 K (FMOF-1c) is also shown for comparison.



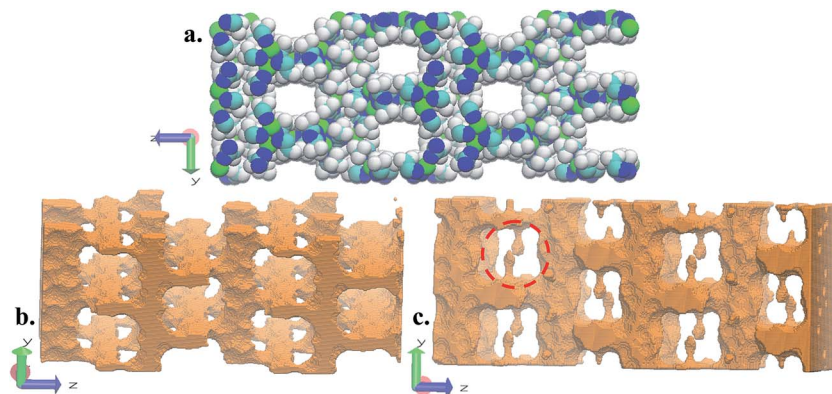


Fig. 4 Schematic views of FMOF-1 channels. (a) Front view of FMOF-1 framework. C, cyan; F, white; N, blue and Ag, green. (b) Illustration of cylindrical channel voids using a 4 Å probe. The cylindrical channels extend along the *x*- and *y*-directions. (c) Representation of channel voids and small pockets protruding alongside the channels highlighted in red circles using a 1 Å probe. Views in (b) and (c) are calculated using the method described by Sarkisov and Harrison.⁵²

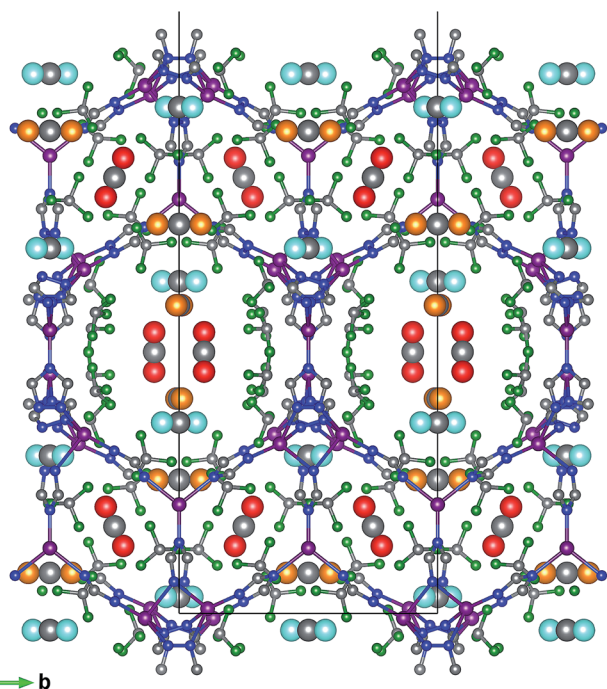


Fig. 5 Crystal structure of adsorbed CO₂ from neutron powder diffraction at 290 K and 61 bar, viewed along the crystallographic *a*-axis. The three CO₂ adsorption sites in FMOF-1 are shown as site I near CF₃ groups at corners with O atoms depicted as cyan spheres; site II near CF₃ groups along *c* direction with O atoms depicted as red spheres; and site III in the direction of the large channels with O atoms depicted as orange spheres.

We see a very good agreement between simulation and experiment for benzene and hexane through the entire pressure range, with saturation loadings around 2 mol kg⁻¹ and 1.2 mol kg⁻¹, respectively. Additionally, both experiment and simulation show no appreciable adsorption of water, thereby confirming the hydrophobic nature of FMOF-1.

However, as shown in Fig. 7a and b, simulated isotherms of N₂ and O₂ adsorption in FMOF-1a drastically underpredict the

saturation loading from experiment by almost 3 mol kg⁻¹ in both cases. The N₂ and O₂ experimental isotherms show two steps, but the FMOF-1a simulated isotherms do not exhibit this behaviour. This pronounced step can be ascribed to the flexibility of the framework upon adsorption, meaning that N₂ or O₂ adsorption leads to structural changes of the framework, allowing for more molecules to be adsorbed in the framework. Similar adsorption isotherms have been observed in some ZIFs and the MIL series of MOFs due to framework flexibility.^{53–59} Since the framework is held rigid in our simulations, we approximated the effect of framework flexibility by simulating isotherms with three different FMOF-1 structures, FMOF-1a, FMOF-1b, and FMOF-1c. FMOF-1b has a larger unit cell volume and channel size than FMOF-1a; the largest cavity diameter in FMOF-1a is 6.1 Å, while in FMOF-1b it is 6.8 Å (see Table S2†). The channel size in FMOF-1c lies between FMOF-1a and FMOF-1b, with a 6.3 Å diameter. Recall from Fig. 3 that the unit cell volume of FMOF-1 varies with temperature and CO₂ loading. Fig. 7 also shows adsorption isotherms of N₂ and O₂ in

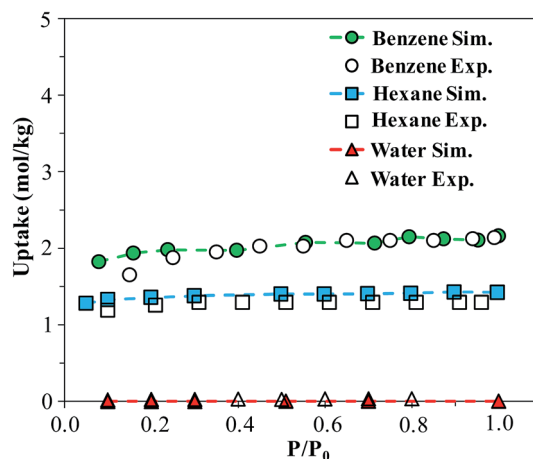


Fig. 6 Simulated and experimental adsorption isotherms for benzene, *n*-hexane, and water in FMOF-1a at 298 K. *P*₀ is the experimental saturation pressure of each adsorbate.



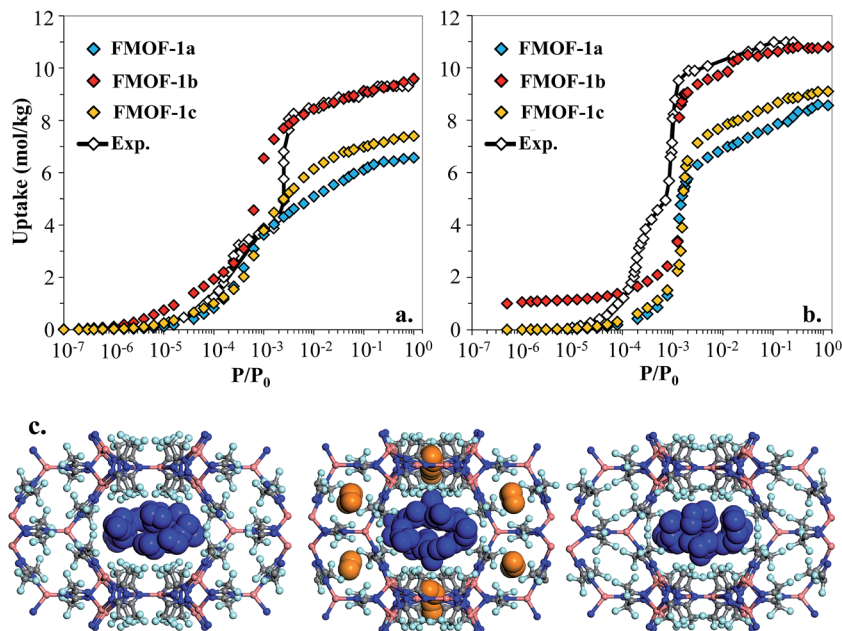


Fig. 7 Experimental adsorption isotherms for (a) N₂ and (b) O₂ at 77 K compared to simulated isotherms in three FMOF-1 structures. In (c) GCMC simulation snapshots for N₂ at saturation loading are shown for FMOF-1a (left), FMOF-1b (middle) and FMOF-1c (right) at 77 K. Nitrogen molecules adsorbed in FMOF-1 large channels and small pockets are illustrated with blue and orange vdW representation, respectively. P₀ is the experimental saturation pressure of each adsorbate.

FMOF-1b and FMOF-1c. The two-step adsorption behaviour observed experimentally can be better described when the simulated isotherms of these different structures are considered. At low loadings, the FMOF-1a and FMOF-1c results agree better with experiment, while at high loadings the saturation loading in the expanded structure (*i.e.* FMOF-1b) is closer to experiment. The maximum amounts adsorbed in FMOF-1b for N₂ and O₂ are 9.6 and 10.7 mol kg⁻¹, respectively, which are in excellent agreement with the experimental values.

The results shown in Fig. 7 corroborate that the framework of FMOF-1 goes through structural expansion upon adsorption of guest molecules. The small cavity in FMOF-1b has a diameter of *ca.* 3.4 Å, which is much larger than the small cavity in FMOF-1a or FMOF-1c (*ca.* 2.5 Å) (Table S2†). The kinetic diameters of N₂ and O₂ are 3.64 Å and 3.46 Å,³ respectively, which suggests that at most one N₂ or O₂ molecule can fit in each small pocket. Interestingly, the window sizes connecting the small pockets to the larger channels are smaller than the kinetic diameters of N₂ and O₂. However, kinetic diameters do not account for the orientation of the molecule – the cross section of a N₂ or O₂ molecule is *ca.* 2.9 Å,⁶⁰ indicating that only FMOF-1b has a window size large enough to admit an N₂ or O₂ molecule. Additionally, since the FMOF-1b structure was obtained at 90 K and the isotherms are measured at 77 K, it is possible that the window size could also be larger at the lower temperature.

To investigate the placement of N₂ within FMOF-1, we compared the position of N₂ molecules adsorbed in the FMOF-1 structures at the saturation loading in the GCMC simulations at 77 K (Fig. 7c). In FMOF-1a and FMOF-1c, N₂ molecules adsorb only in the large cylindrical channels. On the other hand, in FMOF-1b, the N₂ molecules are adsorbed first in the small

pockets at low pressures (Fig. S6†), and it is at the higher pressure that N₂ molecules adsorb in the larger cylindrical channels. In FMOF-1b, molecules prefer to adsorb in the small pores due to the strong interactions between the adsorbed molecules and the Tz ring pairs, and only as the pressure increases do the large channels fill up. Given that only one N₂ molecule can occupy each small pocket, the contribution of small pockets to the total amount adsorbed is 1.1 mol kg⁻¹ (*i.e.*, 8 molecules per unit cell of FMOF-1b). This uptake accounts for 11% of the total amount adsorbed. The remainder of the adsorption difference between FMOF-1b and FMOF-1c or FMOF-1a is due to additional adsorption in the large channels. As shown in Fig. S6,† a similar adsorption mechanism and placement was observed for O₂ at 77 K. This finding further confirms that the structural expansion of FMOF-1 occurs not only in the small pockets but also increases the capacity of the larger channels for small adsorbates. It is worth mentioning that the simulated adsorption isotherms of larger adsorbates such as *n*-hexane and benzene in FMOF-1b and FMOF-1c show no significant differences when compared to the isotherm calculated for FMOF-1a (Fig. S7†). Both fits for the *n*-hexane and benzene adsorption isotherms obtained for FMOF-1 agree very well with both the experimental data and the simulation thereof (Fig. S8a and S8b,† respectively). The kinetic diameters of *n*-hexane and benzene are 4.3 Å and 5.9 Å, respectively, both of which are much larger than the window and pore size of the small pockets. Therefore, the saturation loading of these large molecules is mainly determined by the volume of the cylindrical channels, resulting in a simple type-1 adsorption isotherm profile, with no adsorption in the smaller pores and no consequent steps in the adsorption isotherms. This finding further



demonstrates that, while the accessibility of the smaller pockets and larger main channels for different expansion levels of FMOF-1 have significant effects on the adsorption properties of small adsorbates (*e.g.*, N₂, O₂ and H₂), these structural changes have little effect on the uptake of larger adsorbates (*e.g.*, benzene, *n*-hexane, and toluene).

Fig. 8 shows the experimental CO₂ adsorption isotherms in FMOF-1 structures at 278 K and 283 K, along with simulated data for comparison. This initial comparison is restricted to 30 bar, where the experimental isotherms can be fit with a Toth isotherm and FMOF-1 shows less pronounced flexibility and expansion. The simulated isotherms in all three structures agree reasonably well with the experimental data at low pressure. The difference between experiment and simulation increases for FMOF-1b as the pressure increases. At saturation loadings, better agreement between experiment and simulation results is observed for FMOF-1c, the structure that was obtained under the stream of CO₂ (see Fig. 3). The simulation snapshots at high loading show that CO₂ molecules occupy only the large channels in FMOF-1a and FMOF-1c, since the pockets are too small for CO₂ (kinetic diameter: 3.4 Å). Interestingly, even in the expanded structure of FMOF-1b, the CO₂ molecules cannot be fully accommodated inside the small pockets and stay at the windows connecting the large channels to the small pockets as highlighted in the simulation snapshots in Fig. 8c and the CO₂ density profiles shown at high pressure in Fig. S9.†

In order to better understand the CO₂ adsorption behaviour in FMOF-1 at higher pressure (see Fig. 1), we also simulated CO₂ adsorption isotherms at 298 K up to 50 bar for different FMOF-1a–c structures and compared the results to the simulated uptakes obtained for FMOF-2 and the experimental isotherm (Fig. 9a). The FMOF-2 structure is obtained from annealing

FMOF-1 and recrystallizing from a toluene/acetonitrile solution and consists of two enlarged types of pores: hexagonal channels that are *ca.* 18 Å in width and triangular cages that are *ca.* 10 Å in diameter.⁴⁶ The maximum amount adsorbed for CO₂ at 298 K is 3.2 mol kg⁻¹ or 23.4 molecules per unit cell in FMOF-1c. This value is in excellent agreement with the result from neutron powder diffraction experiments at 290 K, where CO₂ uptake capacity of 3.3 mol kg⁻¹ (24 molecules per unit cell) can be obtained when all three CO₂ adsorption sites in the large cavity are fully occupied under 61 bar of CO₂ loading (see Fig. 5). However, with a small increase in temperature, at 298 K, the simulated CO₂ uptake is significantly lower than the experimental uptake even in the expanded FMOF-1b structure, suggesting further enlargement of the structure at this temperature. It is very clear from Fig. S2† that the CO₂ adsorption of FMOF-1 at room temperature represents a marked deviation from the typical type 1 isotherm. Generally, Toth fit modeling is carried out on type 1 isotherms. In order to obtain a good Toth fit for the CO₂ adsorption at room temperature, the pressure range is scaled down so as to compare experimental data with the simulation data. Fig. S10† shows that the Toth fit agrees well with the FMOF-1b structure. Indeed, the simulated saturation loading of CO₂ in FMOF-2 shows better agreement with the experimental data at elevated pressures around 40 bar, suggesting expansion of FMOF-1 to a more expanded polymorph at higher pressures up to 60 bar. The collective experience of the Omary group with such M-Tz(R_F)₂ compositions suggests the prevalence of polymorphism and crystallographic isomerism; this is so both with and without gas adsorption assistance. As such, the presumed significant expansions (or compressions) suggested by the closer proximity of the experimental high-pressure CO₂ uptake in FMOF-1 at

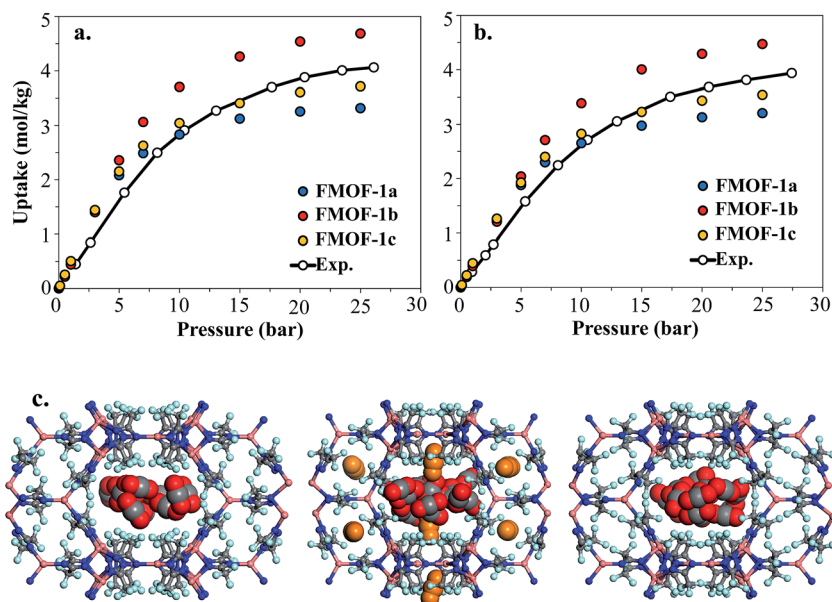


Fig. 8 Experimental adsorption isotherms for CO₂ and simulated isotherms in three FMOF-1 structures at (a) 278 K and (b) 283 K. In (c) GCMC simulation snapshots at saturation loading are shown for FMOF-1a (left), FMOF-1b (middle) and FMOF-1c (right) for CO₂ at 278 K. CO₂ molecules are illustrated in vdW representation, with those adsorbed in the large channels in red and grey and those at the entrance to the small pockets in orange. The experimental and simulated results are excess adsorption isotherms.



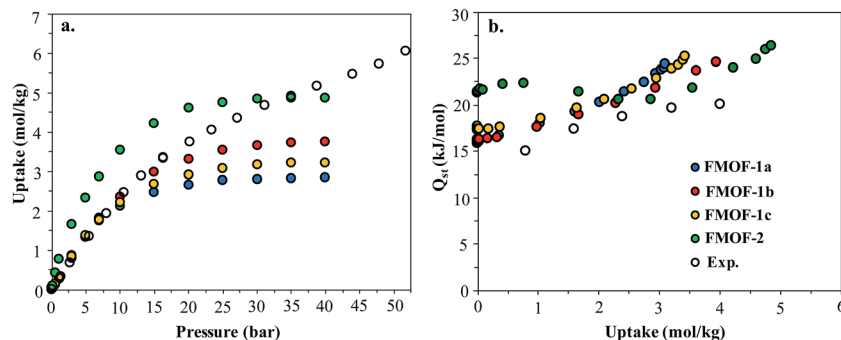


Fig. 9 (a) Experimental excess adsorption isotherms for CO_2 at 298 K in FMOF-1 and simulated excess isotherms in three FMOF-1 structures and FMOF-2. (b) Experimental and simulated heats of adsorption for CO_2 at 298 K.

ambient temperature to the simulated data for the porous-most “emptied” N_2 @FMOF-1 adsorption adduct at 90 K or – better yet – the FMOF-2 structure at 100 K should not be surprising for this highly flexible material.

Fig. 9b compares the isosteric heats of adsorption (Q_{st}) obtained from GCMC simulations to the values obtained from variable-temperature experiments. The predicted Q_{st} values were obtained from the fluctuations of the potential energy over the production cycles in the GCMC simulations for each pressure point. The experimental heats of adsorption were obtained using the Clausius–Clapeyron equation on isotherms from several temperatures (see ESI† for details). In reasonably good agreement with the experimental data, the simulated Q_{st} values for the three FMOF-1 structures rise smoothly from ~ 16 to 25 kJ mol^{-1} as the loading increases due to an increasing contribution from attractive CO_2 – CO_2 interactions (Fig. S11†). The adsorption heats for FMOF-2 are higher ($\sim 22 \text{ kJ mol}^{-1}$) at low loadings due to strong adsorption of CO_2 in the small triangular pores. In general, these moderate adsorption heats lie in the typical range for CO_2 adsorption in MOFs.^{61,62} In Fig. S11,† the average adsorbate–adsorbent and adsorbate–adsorbate energies from GCMC simulations are shown for CO_2 adsorption in FMOF-1c at 298 K as a function of pressure. The van der Waals interactions between CO_2 and FMOF-1 are clearly dominant with $\sim -12 \text{ kJ mol}^{-1}$ contribution, and the coulombic components of the total interaction energy are very small, $\sim -2 \text{ kJ mol}^{-1}$. These results suggest that the fluorine-lined cylindrical pore creates an electrostatic potential that is not very favourable for CO_2 adsorption.

To provide another perspective on guest molecule interactions with FMOF-1, QM methods (UDFT with dispersion) employing cluster models were used to determine the free energy of binding for several guest molecules. The binding sites used in these calculations are located in the cylindrical channel and the small cavity. These sites are represented in Fig. 10 in the extended MOF structure with an overlay of the guest geometry from the cluster calculations.

The 184-atom extended model for the cylindrical channel binding site is taken from the FMOF-1a structure and shown in Fig. 11a. The extended model of the cylindrical channel with coordinated Ag and Tz was truncated with complete metal coordination spheres and ligands. By preserving the metal to

ligand ratio, no atom substitutions were needed. The calculated binding free energies of CO_2 , N_2 , and H_2 in the cylindrical channel are given in Table 1. The binding free energies follow the order $\text{CO}_2 > \text{N}_2 > \text{H}_2$ in the cylindrical channel, indicating selective adsorption of CO_2 . The dispersion contribution increases the magnitude of the binding free energy by ~ 59 – 66% in these calculations. Similar results with respect to the relative contribution of dispersion were found by Neaton for carbon dioxide binding in Mg–MOF-74.⁴⁴

The binding free energies in the small cavity of FMOF-1b were also determined for CO_2 , N_2 , and H_2 using the truncated model shown in Fig. 11b. The small cavity shows stronger binding for N_2 and H_2 than in the cylindrical channel, with a similar dispersion correction making up 62–79% of the total

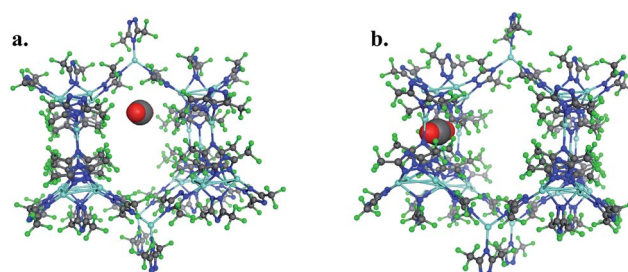


Fig. 10 (a) CO_2 in the cylindrical channel within the extended FMOF-1a structure and (b) CO_2 in the small cavity within the extended FMOF-1b structure.

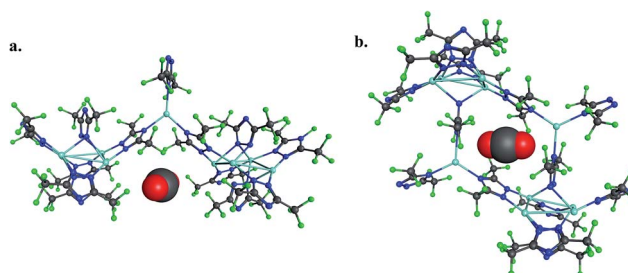


Fig. 11 (a) CO_2 interaction with the 184-atom cluster model for the cylindrical channel taken from the FMOF-1a structure. (b) CO_2 interaction with the 194-atom cluster model for the small cavity taken from the FMOF-1b structure.



Table 1 DFT binding free energies of various guest molecules with FMOF-1a cylindrical channel and FMOF-1b small cavity

| Guest | Cylindrical channel | | Small cavity | |
|-----------------|---------------------------------|------------------------------------|---------------------------------|------------------------------------|
| | PBE0 (kJ mol ⁻¹) | PBE0-D3 (kJ mol ⁻¹) | PBE0 (kJ mol ⁻¹) | PBE0-D3 (kJ mol ⁻¹) |
| CO ₂ | -9.2 | -23.2 | 138.6 | 94.5 |
| N ₂ | -7.9 | -19.3 | -9.0 | -39.6 |
| H ₂ | -2.1 | -6.6 | -3.5 | -16.8 |

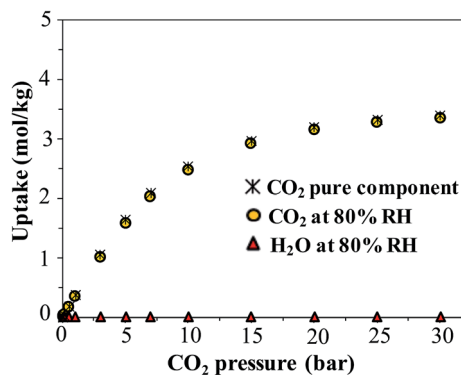


Fig. 12 Simulated excess adsorption isotherms at 298 K for pure component CO₂ and a mixture of CO₂ at different pressures with water at 80% relative humidity in FMOF-1c.

binding free energy. The binding of CO₂ in the truncated model of the small cavity is not energetically favourable, in agreement with GCMC simulations and neutron diffraction experimental results.

Effect of humidity on CO₂ uptake

Finally, we consider the effect of humidity on the CO₂ uptake of FMOF-1. Water is an ever-present component of flue gas, and we hypothesized that hydrophobic MOFs like FMOF-1 should show negligible loss of CO₂ capacity in the presence of water vapor. To test this, we carried out GCMC simulations of CO₂ adsorption in humid conditions. Fig. 12 depicts the excess amount adsorbed for a mixture of CO₂ and H₂O in FMOF-1c at 80% relative humidity and at different CO₂ pressures. Although the CO₂ uptake at 0.15 bar (the relevant condition in flue gas)⁶³ is only *ca.* 0.05 mol kg⁻¹ (0.3 molecules per unit cell), it is shown that for such a superhydrophobic MOF, the uptake of CO₂ is not influenced by the presence of water vapor at all. There is essentially no co-adsorption of water in the presence of CO₂, thereby explaining why the CO₂ adsorption amount is similar under dry or humid conditions. This result supports the hypothesis that FMOFs can be promising to diminish the effect of humidity on CO₂ adsorption performance.

Conclusions

We combined atomically detailed calculations and experiments to provide insight into the adsorption of CO₂ and other molecules

in FMOF-1. First, the force field model for FMOF-1 was verified by comparing simulated adsorption isotherms for a variety of adsorbates in FMOF-1 with existing experimental data. Adsorption isotherms of N₂ and O₂ support previous suggestions that FMOF-1 undergoes major structural changes in the presence of guest molecules. This flexibility, however, has little effect on the loading of larger adsorbates such as *n*-hexane and benzene. The two-step isotherms observed for both N₂ and O₂ can be explained by analyzing simulated adsorption isotherms in both contracted and expanded FMOF-1 structures and the accessibility of the small pockets to these smaller adsorbates. We measured CO₂ adsorption isotherms in FMOF-1 and found good agreement between experiment and simulations using the FMOF-1c structure at sub-ambient temperatures. Simulations of CO₂ adsorption in FMOF-1 along with neutron powder diffraction measurements show that CO₂ molecules cannot fit in the small pores of FMOF-1 and reside in the large channels at three distinct adsorption sites. The experimentally measured CO₂ isotherms near and above room temperature suggest significant framework expansion at high CO₂ pressures. Contact angle measurements confirm FMOF-1's previously reported hydrophobic nature, and Monte Carlo simulations predict no uptake of water even up to the vapor pressure of water. Simulations of CO₂ adsorption in the presence of 80% relative humidity show that the amount of CO₂ adsorbed is essentially the same as in the absence of humidity, validating our hypothesis that hydrophobic MOFs could hold promise for CO₂ capture from flue gas.

Conflict of interest

R. Q. S. has a financial interest in the start-up company NuMat Technologies, which is seeking to commercialize metal-organic frameworks.

Acknowledgements

M. A. O. gratefully acknowledges support to his group's contributions by the Robert A. Welch Foundation (B-1542) and the National Science Foundation (Grant CHE-1413641 and its corresponding international supplement CHE-1545934), and thanks Dr Chi Yang for his efforts to initiate the discovery of FMOFs in the Omary group. R. Q. S. acknowledges support from the National Science Foundation (CBET-1604890). Computational work was supported by Northwestern University's shared computer system, Quest (Project: P20261). Research conducted at the ORNL Spallation Neutron Source was supported by the Scientific User Facilities division, Office of Basic Energy Sciences, US Department of Energy, under Contract DE-AC05-00OR22725 with UT-Battelle, LLC. A. K. W. acknowledges the research conducted at the UNT Center for Advanced Scientific Computing and Modeling, which was supported by the U.S. Department of Energy under DE-FG02-08ER64603.

References

- 1 R. S. Haszeldine, Carbon Capture and Storage: How Green Can Black Be?, *Science*, 2009, **325**, 1647–1652.



- 2 D. J. Darensbourg, W.-C. Chung, K. Wang and H.-C. Zhou, Sequestering CO₂ for Short-Term Storage in MOFs: Copolymer Synthesis with Oxiranes, *ACS Catal.*, 2014, **4**, 1511–1515.
- 3 J.-R. Li, R. J. Kuppler and H.-C. Zhou, Selective gas adsorption and separation in metal–organic frameworks, *Chem. Soc. Rev.*, 2009, **38**, 1477–1504.
- 4 G. Férey, Hybrid porous solids: past, present, future, *Chem. Soc. Rev.*, 2008, **37**, 191–214.
- 5 S. Horike, S. Shimomura and S. Kitagawa, Soft porous crystals, *Nat. Chem.*, 2009, **1**, 695–704.
- 6 T. M. McDonald, J. A. Mason, X. Kong, E. D. Bloch, D. Gygi, A. Dani, V. Crocella, F. Giordanino, S. O. Odoh, W. S. Drisdell, B. Vlasisavljevich, A. L. Dzubak, R. Poloni, S. K. Schnell, N. Planas, K. Lee, T. Pascal, L. F. Wan, D. Prendergast, J. B. Neaton, B. Smit, J. B. Kortright, L. Gagliardi, S. Bordiga, J. A. Reimer and J. R. Long, Cooperative insertion of CO₂ in diamine-appended metal–organic frameworks, *Nature*, 2015, **519**, 303–308.
- 7 H. Furukawa, K. E. Cordova, M. O’Keeffe and O. M. Yaghi, The Chemistry and Applications of Metal–Organic Frameworks, *Science*, 2013, **341**, 1230444.
- 8 J. Liu, P. K. Thallapally, B. P. McGrail, D. R. Brown and J. Liu, Progress in adsorption-based CO₂ capture by metal–organic frameworks, *Chem. Soc. Rev.*, 2012, **41**, 2308–2322.
- 9 S. Keskin, T. M. van Heest and D. S. Sholl, Can Metal–Organic Framework Materials Play a Useful Role in Large-Scale Carbon Dioxide Separations?, *ChemSusChem*, 2010, **3**, 879–891.
- 10 T. M. McDonald, W. R. Lee, J. A. Mason, B. M. Wiers, C. S. Hong and J. R. Long, Capture of Carbon Dioxide from Air and Flue Gas in the Alkylamine-Appended Metal–Organic Framework mmen-Mg₂(dobpdc), *J. Am. Chem. Soc.*, 2012, **134**, 7056–7065.
- 11 C. E. Wilmer, O. K. Farha, Y.-S. Bae, J. T. Hupp and R. Q. Snurr, Structure–property relationships of porous materials for carbon dioxide separation and capture, *Energy Environ. Sci.*, 2012, **5**, 9849–9856.
- 12 A. Ö. Yazaydin, R. Q. Snurr, T.-H. Park, K. Koh, J. Liu, M. D. LeVan, A. I. Benin, P. Jakubczak, M. Lanuza, D. B. Galloway, J. J. Low and R. R. Willis, Screening of Metal–Organic Frameworks for Carbon Dioxide Capture from Flue Gas Using a Combined Experimental and Modeling Approach, *J. Am. Chem. Soc.*, 2009, **131**, 18198–18199.
- 13 S. R. Caskey, A. G. Wong-Foy and A. J. Matzger, Dramatic Tuning of Carbon Dioxide Uptake *via* Metal Substitution in a Coordination Polymer with Cylindrical Pores, *J. Am. Chem. Soc.*, 2008, **130**, 10870–10871.
- 14 J. Liu, A. I. Benin, A. M. B. Furtado, P. Jakubczak, R. R. Willis and M. D. LeVan, Stability Effects on CO₂ Adsorption for the DOBDC Series of Metal–Organic Frameworks, *Langmuir*, 2011, **27**, 11451–11456.
- 15 J. Canivet, A. Fateeva, Y. Guo, B. Coasne and D. Farrusseng, Water adsorption in MOFs: fundamentals and applications, *Chem. Soc. Rev.*, 2014, **43**, 5594–5617.
- 16 P. Z. Moghadam, D. Fairen-Jimenez and R. Q. Snurr, Efficient identification of hydrophobic MOFs: application in the capture of toxic industrial chemicals, *J. Mater. Chem. A*, 2016, **4**, 529–536.
- 17 P. Deria, J. E. Mondloch, E. Tylianakis, P. Ghosh, W. Bury, R. Q. Snurr, J. T. Hupp and O. K. Farha, Perfluoroalkane Functionalization of NU-1000 *via* Solvent-Assisted Ligand Incorporation: Synthesis and CO₂ Adsorption Studies, *J. Am. Chem. Soc.*, 2013, **135**, 16801–16804.
- 18 J. B. Decoste, G. W. Peterson, M. W. Smith, C. A. Stone and C. R. Willis, Enhanced Stability of Cu-BTC MOF *via* Perfluorohexane Plasma-Enhanced Chemical Vapor Deposition, *J. Am. Chem. Soc.*, 2012, **134**, 1486–1489.
- 19 P. Z. Moghadam, P. Ghosh and R. Q. Snurr, Understanding the Effects of Preadsorbed Perfluoroalkanes on the Adsorption of Water and Ammonia in MOFs, *J. Phys. Chem. C*, 2015, **119**, 3163–3170.
- 20 C. Yang, X. Wang and M. A. Omary, Fluorous Metal–Organic Frameworks for High-Density Gas Adsorption, *J. Am. Chem. Soc.*, 2007, **129**, 15454–15455.
- 21 C. Serre, Superhydrophobicity in Highly Fluorinated Porous Metal–Organic Frameworks, *Angew. Chem., Int. Ed.*, 2012, **51**, 6048–6050.
- 22 T.-H. Chen, I. Popov, O. Zenasni, O. Daugulis and O. S. Miljanic, Superhydrophobic perfluorinated metal–organic frameworks, *Chem. Commun.*, 2013, **49**, 6846–6848.
- 23 C. Yang, X. Wang and M. A. Omary, Crystallographic Observation of Dynamic Gas Adsorption Sites and Thermal Expansion in a Breathable Fluorous Metal–Organic Framework, *Angew. Chem., Int. Ed.*, 2009, **48**, 2500–2505.
- 24 B. Schmitz, U. Müller, N. Trukhan, M. Schubert, G. Férey and M. Hirscher, Heat of Adsorption for Hydrogen in Microporous High-Surface-Area Materials, *ChemPhysChem*, 2008, **9**, 2181–2184.
- 25 B. H. Toby and R. B. Von Dreele, GSAS-II: the genesis of a modern open-source all purpose crystallography software package, *J. Appl. Crystallogr.*, 2013, **46**, 544–549.
- 26 J. J. Potoff and J. I. Siepmann, Vapor–liquid equilibria of mixtures containing alkanes, carbon dioxide, and nitrogen, *AIChE J.*, 2001, **47**, 1676–1682.
- 27 L. Zhang and J. I. Siepmann, Direct calculation of Henry’s law constants from Gibbs ensemble Monte Carlo simulations: nitrogen, oxygen, carbon dioxide and methane in ethanol, *Theor. Chem. Acc.*, 2006, **115**, 391–397.
- 28 M. G. Martin and J. I. Siepmann, Transferable Potentials for Phase Equilibria. 1. United-Atom Description of *n*-Alkanes, *J. Phys. Chem. B*, 1998, **102**, 2569–2577.
- 29 N. Rai and J. I. Siepmann, Transferable Potentials for Phase Equilibria. 9. Explicit Hydrogen Description of Benzene and Five-Membered and Six-Membered Heterocyclic Aromatic Compounds, *J. Phys. Chem. B*, 2007, **111**, 10790–10799.
- 30 W. L. Jorgensen, J. Chandrasekhar, J. D. Madura, R. W. Impey and M. L. Klein, Comparison of simple potential functions for simulating liquid water, *J. Chem. Phys.*, 1983, **79**, 926–935.
- 31 A. K. Rappe, C. J. Casewit, K. S. Colwell, W. A. Goddard and W. M. Skiff, UFF, a full periodic table force field for molecular mechanics and molecular dynamics simulations, *J. Am. Chem. Soc.*, 1992, **114**, 10024–10035.



- 32 V. H. Dalvi, V. Srinivasan and P. J. Rossky, Understanding the Effectiveness of Fluorocarbon Ligands in Dispersing Nanoparticles in Supercritical Carbon Dioxide, *J. Phys. Chem. C*, 2010, **114**, 15553–15561.
- 33 C. M. Breneman and K. B. Wiberg, Determining atom-centered monopoles from molecular electrostatic potentials. The need for high sampling density in formamide conformational analysis, *J. Comput. Chem.*, 1990, **11**, 361–373.
- 34 M. Valiev, E. J. Bylaska, N. Govind, K. Kowalski, T. P. Straatsma, H. J. J. Van Dam, D. Wang, J. Nieplocha, E. Apra, T. L. Windus and W. A. de Jong, NWChem: A comprehensive and scalable open-source solution for large scale molecular simulations, *Comput. Phys. Commun.*, 2010, **181**, 1477–1489.
- 35 C. Adamo and V. Barone, Toward reliable density functional methods without adjustable parameters: The PBE0 model, *J. Chem. Phys.*, 1999, **110**, 6158–6170.
- 36 R. Krishnan, J. S. Binkley, R. Seeger and J. A. Pople, Self-consistent molecular orbital methods. XX. A basis set for correlated wave functions, *J. Chem. Phys.*, 1980, **72**, 650–654.
- 37 W. J. Hehre, R. Ditchfield and J. A. Pople, Self-Consistent Molecular Orbital Methods. XII. Further Extensions of Gaussian-Type Basis Sets for Use in Molecular Orbital Studies of Organic Molecules, *J. Chem. Phys.*, 1972, **56**, 2257–2261.
- 38 M. Kaupp, P. V. R. Schleyer, H. Stoll and H. Preuss, Pseudopotential approaches to Ca, Sr, and Ba hydrides. Why are some alkaline earth MX₂ compounds bent?, *J. Chem. Phys.*, 1991, **94**, 1360–1366.
- 39 S. Grimme, J. Antony, S. Ehrlich and H. Krieg, A consistent and accurate *ab initio* parametrization of density functional dispersion correction (DFT-D) for the 94 elements H-Pu, *J. Chem. Phys.*, 2010, **132**, 154104.
- 40 J. Joo, H. Kim and S. S. Han, Volume shrinkage of a metal-organic framework host induced by the dispersive attraction of guest gas molecules, *Phys. Chem. Chem. Phys.*, 2013, **15**, 18822–18826.
- 41 J. A. Gee and D. S. Sholl, Prediction of Adsorption Properties of Cyclic Hydrocarbons in MOFs Using DFT-Derived Force Fields, *J. Phys. Chem. C*, 2015, **119**, 16920–16926.
- 42 B. K. Chang, N. C. Bristowe, P. D. Bristowe and A. K. Cheetham, van der Waals forces in the perfluorinated metal-organic framework zinc 1,2-bis(4-pyridyl)ethane tetrafluoroterephthalate, *Phys. Chem. Chem. Phys.*, 2012, **14**, 7059–7064.
- 43 K. Lee, J. D. Howe, L.-C. Lin, B. Smit and J. B. Neaton, Small-Molecule Adsorption in Open-Site Metal-Organic Frameworks: A Systematic Density Functional Theory Study for Rational Design, *Chem. Mater.*, 2015, **27**, 668–678.
- 44 R. Poloni, B. Smit and J. B. Neaton, CO₂ Capture by Metal-Organic Frameworks with van der Waals Density Functionals, *J. Phys. Chem. A*, 2012, **116**, 4957–4964.
- 45 R. Babarao and J. Jiang, Molecular Screening of Metal-Organic Frameworks for CO₂ Storage, *Langmuir*, 2008, **24**, 6270–6278.
- 46 C. Yang, U. Kaipa, Q. Z. Mather, X. Wang, V. Nesterov, A. F. Venero and M. A. Omary, Fluorous Metal-Organic Frameworks with Superior Adsorption and Hydrophobic Properties toward Oil Spill Cleanup and Hydrocarbon Storage, *J. Am. Chem. Soc.*, 2011, **133**, 18094–18097.
- 47 N. Nijem, P. Canepa, U. Kaipa, K. Tan, K. Roodenko, S. Tekarli, J. Halbert, I. W. H. Oswald, R. K. Arvapally, C. Yang, T. Thonhauser, M. A. Omary and Y. J. Chabal, Water Cluster Confinement and Methane Adsorption in the Hydrophobic Cavities of a Fluorinated Metal-Organic Framework, *J. Am. Chem. Soc.*, 2013, **135**, 12615–12626.
- 48 N. Nijem and Y. J. Chabal, Adsorbate Interactions in Metal Organic Frameworks Studied by Vibrational Spectroscopy, *Comments Inorg. Chem.*, 2014, **34**, 78–102.
- 49 J. G. Nguyen and S. M. Cohen, Moisture-Resistant and Superhydrophobic Metal-Organic Frameworks Obtained via Postsynthetic Modification, *J. Am. Chem. Soc.*, 2010, **132**, 4560–4561.
- 50 K. P. Rao, M. Higuchi, K. Sumida, S. Furukawa, J. Duan and S. Kitagawa, Design of Superhydrophobic Porous Coordination Polymers through the Introduction of External Surface Corrugation by the Use of an Aromatic Hydrocarbon Building Unit, *Angew. Chem., Int. Ed.*, 2014, **53**, 8225–8230.
- 51 A. F. Stalder, T. Melchior, M. Müller, D. Sage, T. Blu and M. Unser, Low-bond axisymmetric drop shape analysis for surface tension and contact angle measurements of sessile drops, *Colloids Surf., A*, 2010, **364**, 72–81.
- 52 L. Sarkisov and A. Harrison, Computational structure characterisation tools in application to ordered and disordered porous materials, *Mol. Simul.*, 2011, **37**, 1248–1257.
- 53 D. Fairen-Jimenez, S. A. Moggach, M. T. Wharmby, P. A. Wright, S. Parsons and T. Düren, Opening the Gate: Framework Flexibility in ZIF-8 Explored by Experiments and Simulations, *J. Am. Chem. Soc.*, 2011, **133**, 8900–8902.
- 54 F.-X. Coudert, A. Boutin, A. H. Fuchs and A. V. Neimark, Adsorption Deformation and Structural Transitions in Metal-Organic Frameworks: From the Unit Cell to the Crystal, *J. Phys. Chem. Lett.*, 2013, **4**, 3198–3205.
- 55 W. Cai, T. Lee, M. Lee, W. Cho, D.-Y. Han, N. Choi, A. C. K. Yip and J. Choi, Thermal Structural Transitions and Carbon Dioxide Adsorption Properties of Zeolitic Imidazolate Framework-7 (ZIF-7), *J. Am. Chem. Soc.*, 2014, **136**, 7961–7971.
- 56 L. Sarkisov, R. L. Martin, M. Haranczyk and B. Smit, On the Flexibility of Metal-Organic Frameworks, *J. Am. Chem. Soc.*, 2014, **136**, 2228–2231.
- 57 K. Zhang, R. P. Lively, C. Zhang, R. R. Chance, W. J. Koros, D. S. Sholl and S. Nair, Exploring the Framework Hydrophobicity and Flexibility of ZIF-8: From Biofuel Recovery to Hydrocarbon Separations, *J. Phys. Chem. Lett.*, 2013, **4**, 3618–3622.
- 58 N. Nijem, H. Wu, P. Canepa, A. Marti, K. J. Balkus, T. Thonhauser, J. Li and Y. J. Chabal, Tuning the Gate Opening Pressure of Metal-Organic Frameworks (MOFs)



- for the Selective Separation of Hydrocarbons, *J. Am. Chem. Soc.*, 2012, **134**, 15201–15204.
- 59 A. Schneemann, V. Bon, I. Schwedler, I. Senkovska, S. Kaskel and R. A. Fischer, Flexible metal–organic frameworks, *Chem. Soc. Rev.*, 2014, **43**, 6062–6096.
- 60 C. E. Webster, R. S. Drago and M. C. Zerner, Molecular Dimensions for Adsorptives, *J. Am. Chem. Soc.*, 1998, **120**, 5509–5516.
- 61 J.-R. Li, Y. Ma, M. C. McCarthy, J. Sculley, J. Yu, H.-K. Jeong, P. B. Balbuena and H.-C. Zhou, Carbon dioxide capture-related gas adsorption and separation in metal–organic frameworks, *Coord. Chem. Rev.*, 2011, **255**, 1791–1823.
- 62 J. M. Simmons, H. Wu, W. Zhou and T. Yildirim, Carbon capture in metal–organic frameworks—a comparative study, *Energy Environ. Sci.*, 2011, **4**, 2177–2185.
- 63 Y.-S. Bae and R. Q. Snurr, Development and Evaluation of Porous Materials for Carbon Dioxide Separation and Capture, *Angew. Chem., Int. Ed.*, 2011, **50**, 11586–11596.

