

ferroelectricity could be realized.⁶ For this purpose, liquid-crystalline (LC) assemblies are expected to have great advantage over crystalline assemblies. However, LC sumanene has not yet been developed.

Here we report the first sumanene derivatives that can form LC mesophases with a highly ordered hexagonally arranged columnar structure over a wide temperature range. The LC sumanenes carry six simple thioalkyl side chains at the peripheral aromatic positions (**1**_{C6} and **1**_{C12}, Fig. 1), which is in striking contrast to LC corannulenes.^{2f,g} The single-crystal X-ray analysis of **1**_{C1} with six thiomethyl groups (Fig. 1) showed a 1D bowl-stacking structure. Theoretical calculations suggest that hexathioalkyl sumanenes have an inversed and enhanced concave-convex polarization of the sumanene core, compared to that of non-substituted **1**_H. We also demonstrate that **1**_{C1} behaves as an electron donor due to the six thioalkyl substituents and forms a complex with C₆₀, which has never been achieved with non-substituted sumanene (**1**_H) or previously reported sumanene derivatives.

Results and discussion

Synthesis of hexathioalkyl sumanenes

Mainly due to the lack of a synthetic method for the full peripheral functionalization of sumanene,¹ a sumanene derivative that can form an LC assembly has not yet been reported. Recently, we reported the successful synthesis of 2,3,5,6,8,9-hexabromosumanene (**1**_{Br}, Fig. 1), which can readily be converted into hexa-arylated sumanene derivatives through Pd-catalyzed cross-coupling such as Suzuki–Miyaura reaction.^{1f} We found that compound **1**_{Br} also allows aromatic nucleophilic substitution with thioalkoxides. Typically, **1**_{Br} was reacted with 18 equivalents of sodium thiomethoxide (NaSCH₃) in 1,3-dimethyl-2-imidazolidinone (DMI) at 100 °C under argon to give **1**_{C1} in 35% yield (ESI†). Similarly, **1**_{C6}, **1**_{C12} and **1**_{EH} (Fig. 1) were obtained in 39–43% yields using the corresponding sodium thioalkoxide in place of NaSCH₃ (ESI†). All of these hexathioalkyl sumanenes were unambiguously characterized by ¹H and ¹³C NMR spectroscopy, IR spectroscopy, and high-resolution APCI-TOF mass spectrometry (ESI†). For instance, the ¹H NMR spectrum of **1**_{C1} in toluene-*d*₈ at 25 °C showed two doublet signals at δ = 4.50 and 3.54 ppm arising from the benzylic protons at the *exo*- and *endo*-positions, respectively (Fig. S1, ESI†). Note that these signals were not coalesced, even at elevated temperatures (*e.g.*, 100 °C, Fig. S1, ESI†). Thus, it is likely that the rate of bowl-to-bowl inversion of **1**_{C1} in solution is sufficiently slow relative to the timescale of ¹H NMR spectroscopy.

X-ray crystal structure of hexathiomethyl sumanene (**1**_{C1})

We successfully obtained needle-shaped pale-yellow single crystals of **1**_{C1}, suitable for X-ray diffraction analysis, by the slow diffusion of a hexane vapor into a dichloromethane solution of **1**_{C1} (ESI†). Single-crystal X-ray analysis revealed detailed molecular and assembled structures of **1**_{C1}. The crystal of **1**_{C1} belongs to the *P3* space group, and the

asymmetric unit in the unit cell contains four entire **1**_{C1} molecules and six fragments of one-third of **1**_{C1} (*Z* = 18). Overall, each crystallographically independent molecule is very similar in terms of bond lengths and angles. The mean bowl-depth of **1**_{C1} (1.04 Å, Fig. 2) at the peripheral aromatic carbons is slightly shallower than those observed for the crystal structures of **1**_H (1.11 Å)^{1a} and **1**_{Br} (1.08 Å).^{1f} In the crystal of **1**_{C1}, 1D columns are formed in a bowl-stack manner with a quasi-staggered stacking geometry, where the mean stacking distance (4.00 Å) is comparable to that observed for **1**_{Br} (3.94 Å),^{1f} while it is longer than that observed for **1**_H (3.86 Å).^{1a} Considering the observed stacking geometry of **1**_{C1}, along with the previously reported simulation for **1**_H,⁷ an intermolecular electrostatic force is likely responsible for the formation of the 1D column. Meanwhile, the 1D columns assemble laterally to form a 2D quasi-hexagonal structure with an inter-columnar distance of approximately 12 Å (Fig. 2). When viewed along the *c*-axis, six columns with a convex upward geometry surround a column with a convex downward geometry, so that the dipole moments generated in the column can be partially cancelled (Fig. 2a). The 2D quasi-hexagonal structure developed in the crystal of **1**_{C1} is reminiscent of the structure of a hexagonal columnar (Col_h) mesophase that is typically observed for discotic LC assemblies.⁸

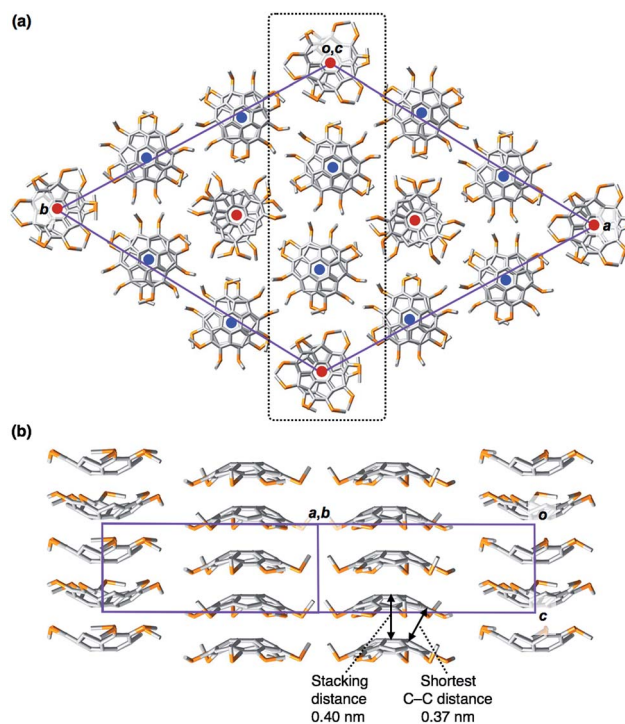


Fig. 2 (a) Crystal packing diagram of **1**_{C1} viewed along the *c* axis. In the 1D columns indicated with blue circles, **1**_{C1} molecules pile up in such a way that the convex face is oriented to the front side, while the 1D columns indicated with red circles consist of **1**_{C1} with an opposite concave orientation. (b) Packing diagram of 1D columns of **1**_{C1} viewed along the (1–10) plane, which corresponds to the part surrounded by a broken line in (a).



DFT calculations of the molecular and electronic structures of hexathiomethyl sumanene (1_{C1})

To gain insight into how the thioalkyl groups affect the electronic structure of sumanene, we performed density functional theory (DFT) calculations on the isolated molecule of 1_{C1} in vacuum at the w97BD/6-311G++(d,p) level (Fig. 3 and Table S1, ESI†). The molecular structure of 1_{C1} obtained by the single-crystal X-ray analysis was used as the initial geometry for geometry optimization. The calculated bond lengths and angles in the optimized geometries were in good agreement with those observed for the crystal structure of 1_{C1} (Fig. 2). Because of the electron-rich sulfur atoms, the calculated dipole moment of 1_{C1} along its C_3 -symmetric axis [4.7 Debye (D)] was much greater than that calculated for non-substituted sumanene 1_H (2.7 D) at the same level (Table S2, ESI†), and the directions of the dipole moments of 1_{C1} and 1_H were opposite to one another (Fig. 3). The surface electrostatic potential (ESP) diagram of 1_{C1} showed negative and positive ESP values for the concave and convex faces, respectively (Fig. 3). We suppose that intermolecular electrostatic attractive interactions between positive convex and negative concave faces of 1_{C1} could compensate for the unfavorable parallel dipole alignment, leading to the formation of the columnar structure in the crystal (Fig. 2). Note that, since the ESP diagram of 1_H illustrates negative convex and relatively positive concave faces (Fig. 3),⁴¹ the six thioalkyl groups in 1_{C1} invert and enhance the concave–convex polarization of the sumanene skeleton.

Characterization of liquid-crystalline (LC) mesophases of sumanene derivatives

Hexathioalkyl sumanenes with long alkyl side chains (1_{C6} and 1_{C12} , Fig. 1) exhibit bowl-stacking to form a highly ordered columnar LC assembly. In differential scanning calorimetry (DSC), 1_{C6} exhibited an LC mesophase over a wide temperature range (<174 °C), while 1_{C12} displayed a phase sequence involving an LC mesophase (35–117 °C) and two crystal phases (–11 to 35 °C and <–11 °C), upon cooling from the corresponding isotropic liquid phases (Fig. 4). Polarized optical

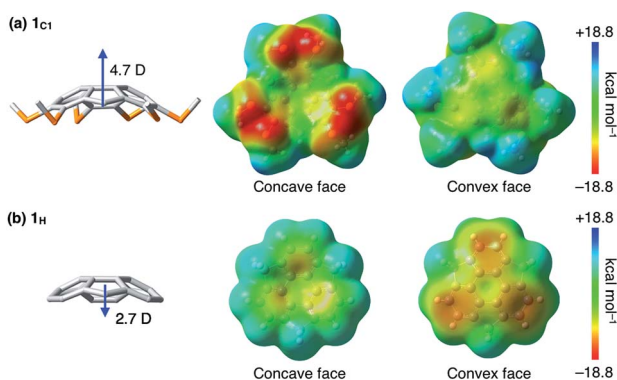


Fig. 3 Calculated electric dipole moments and electrostatic potential surfaces of (a) 1_{C1} and (b) non-substituted 1_H at the w97BD/6-311G++(d,p) level of theory. The blue arrows indicate the directions of the dipole moments.

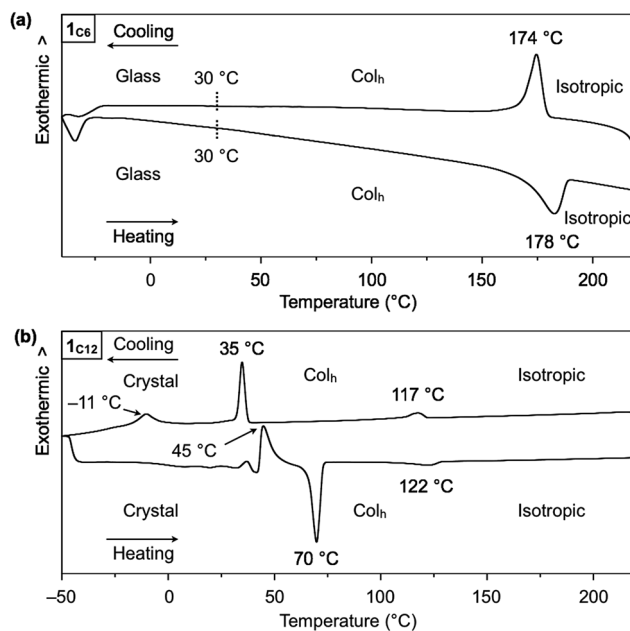


Fig. 4 DSC traces of (a) 1_{C6} and (b) 1_{C12} on second heating and cooling (scan rate = 10 °C min⁻¹).

microscopy (POM) images of the LC mesophases of 1_{C6} and 1_{C12} showed a fan-shaped texture, which is typically observed for hexagonal columnar (Col_h) LC assemblies (Fig. 5a and b).

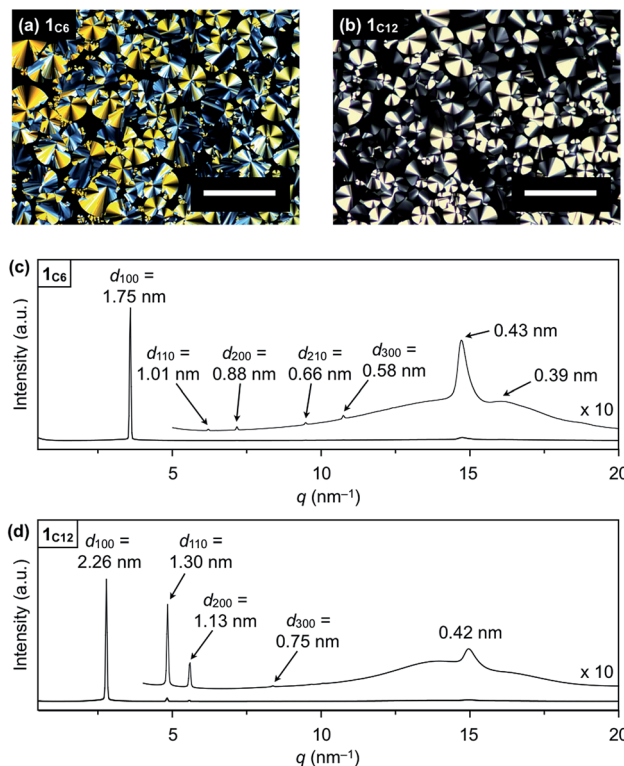


Fig. 5 POM images of (a) 1_{C6} at 140 °C on cooling and (b) 1_{C12} at 80 °C on cooling (scale bar = 200 μm). XRD patterns of (c) 1_{C6} and (d) 1_{C12} in a glass capillary at 100 °C on heating.



The powder X-ray diffraction (XRD) pattern of a bulk sample of **1**_{C6} at 100 °C upon heating displayed five diffraction peaks with *d*-spacings of 1.75, 1.01, 0.88, 0.66 and 0.58 nm. The ratio of the *d*-spacings (1 : $\sqrt{3}$: 2 : $\sqrt{7}$: 3) completely agrees with the values expected for diffractions from the (100), (110), (200), (210) and (300) planes of a 2D hexagonal lattice (Fig. 5c and S2, ESI†). The hexagonal lattice parameter (*a*), given by $2d_{100}/\sqrt{3}$, was determined to be 2.02 nm. The diffraction peak at scattering vector $q = 14.6 \text{ nm}^{-1}$ (*d*-spacing = 0.43 nm) corresponds to the core-to-core separation of the mesogen along the columnar axis, which is comparable to the bowl-stacking distance in the 1D columns of crystalline **1**_{C1} (4.0 Å, Fig. 2). A very broad diffraction peak at approximately $q = 16 \text{ nm}^{-1}$ (*d*-spacing = 0.39 nm) likely arises from the shortest C–C distance in the 1D column, judging from the crystal structure of **1**_{C1} (Fig. 2). The diffraction peak corresponding to the core-to-core separation is exceptionally strong and sharp (Fig. 5c and S2, ESI†), compared to those observed for usual discotic columnar LC assemblies.⁸ This observation suggests that the structural order of the columnar arrays of the bowl-shaped mesogen is remarkably high. Although no phase-transition feature was observed below 30 °C in the DSC analysis, temperature-dependent XRD measurements revealed that **1**_{C6} undergoes a phase transition from the LC mesophase to a glassy solid phase at approximately 30 °C. Accordingly, the XRD pattern of **1**_{C6} at *e.g.*, 20 °C upon cooling (Fig. S2, ESI†) displayed three sharp peaks at a wide-angle region ($q > 16 \text{ nm}^{-1}$). In addition to these peaks, some periodic peaks arising from the ordered Col_h structure remained, indicating that the structural feature of the LC mesophase is maintained to some extent in the glassy solid phase.

The structure of **1**_{C12} in the LC mesophase was essentially identical to that of **1**_{C6}, except for a larger hexagonal lattice parameter (*a* = 2.60 nm) due to the longer alkyl side chains. As shown in Fig. 5d and S3 (ESI†), the XRD pattern of a bulk sample of **1**_{C12} at 100 °C upon heating displayed five diffraction peaks with *d*-spacings of 2.26, 1.30, 1.13, 0.75, and 0.42 nm. The last peak, which is assignable to the core-to-core separation of the mesogen, was not affected by the difference in the side-chain lengths. We also note that when branched alkyl chains are attached to the sumanene core, formation of the LC mesophase would be suppressed, by analogy to discotic LC assemblies with a planar aromatic mesogen. For instance, compound **1**_{FH} (Fig. 1) only exhibited a phase transition from crystal to isotropic liquid, as revealed by DSC, POM and XRD analyses (Fig. S4, ESI†).

Since molecules in the LC state can behave dynamically, we supposed that the LC sumanenes such as **1**_{C12} and **1**_{C6} might exhibit a collective bowl-to-bowl inversion in response to electric fields. To explore this possibility, we measured the dielectric properties of bulk samples of **1**_{C12} and **1**_{C6} (Fig. S5, ESI†). Since dielectric properties are sensitive to the motions of polar molecular units,⁹ dielectric relaxation measurements should detect a thermally activated bowl-to-bowl inversion of **1**_{C12} and **1**_{C6}, if it occurs. However, the dielectric constants of **1**_{C12} and **1**_{C6} did not change significantly over a temperature range of crystalline and Col_h phases as well as a frequency range of 10^3 – 10^6 Hz (Fig. S5, ESI†), indicating that the LC materials have antiferroelectric properties. Based on this observation, the

bowl-stacked 1D columns of **1**_{C12} and **1**_{C6}, in the Col_h mesophases do not respond to the electric fields, but rather statically align in an antiparallel manner to cancel out the dipole generated in each column.

Complexation of hexathiomethyl sumanene (**1**_{C1}) with C₆₀

Although sumanene derivatives capable of complexation with C₆₀ have never been reported to date, we supposed that this would not be the case with hexathioalkyl sumanenes, since the electron-donating thioalkyl groups could change the inherent electronic properties of sumanene. Indeed, cyclic voltammetry of **1**_{C1} in CH₂Cl₂ in the presence of tetrabutylammonium hexafluorophosphate as a supporting electrolyte displayed a reversible oxidation wave at $E_{1/2} = 0.64 \text{ V}$ (*versus* ferrocene/ferrocenium), while non-substituted **1**_H under identical conditions showed an irreversible oxidation wave at 1.03 V (Fig. S6, ESI†). Due to its improved electron-donating properties as well as its concave structure, **1**_{C1} might show strong affinity toward electron-accepting C₆₀.¹⁰

The ¹H NMR spectrum of an equimolar mixture of **1**_{C1} and C₆₀ (**1**_{C1} = 1.0 mM, toluene-*d*₈, 25 °C) showed two doublet signals and one singlet signal at $\delta = 4.51, 3.54, \text{ and } 2.26 \text{ ppm}$ arising from the benzylic *exo*- and *endo*-protons and S–CH₃ groups of **1**_{C1}, respectively (Fig. S7, ESI†). These ¹H NMR signals were shifted slightly more downfield relative to those observed for **1**_{C1} in the absence of C₆₀ (4.50, 3.53, and 2.24 ppm, Fig. S7, ESI†). After the NMR measurement, we noticed that a considerable amount of black precipitate formed, which contained both **1**_{C1} and C₆₀ as confirmed by APCI-TOF mass spectrometric analysis. Obviously, **1**_{C1} gave a complex with C₆₀. Although the association constant between **1**_{C1} and C₆₀ could not be determined because of the low solubility of a mixture **1**_{C1} and C₆₀ in solution, Job's plot¹¹ of **1**_{C1} with C₆₀ in toluene-*d*₈ at 25 °C (total concentration $[\mathbf{1}_{\text{C1}}] + [\text{C}_{60}] = 0.1 \text{ mmol}$), based on the ¹H NMR chemical shift of the S–CH₃ signal as a reference, suggested the occurrence of an approximately 1 : 1 complexation (Fig. S8, ESI†), which is most likely due to a concave–convex interaction.⁵

The use of **1**_{C12} in place of **1**_{C1} allowed the evaluation of the association constant (K_a) with C₆₀. When C₆₀ was added to a toluene-*d*₈ solution of **1**_{C12} (1.0 mM, $[\text{C}_{60}]/[\mathbf{1}_{\text{C12}}] = 0.0\text{--}3.7$), ¹H NMR signals due to the benzylic and thioalkyl protons of **1**_{C12} gradually shifted downfield without any precipitation (Fig. S9, ESI†). From the ¹H NMR spectral change, K_a was determined to be 280 M^{-1} (Fig. S10, ESI†), which is not very high compared to those reported for curved π -systems and C₆₀.^{2,4b,f,10b} Job's plot of **1**_{C12} with C₆₀ in toluene-*d*₈ at 25 °C (total concentration $[\mathbf{1}_{\text{C12}}] + [\text{C}_{60}] = 0.1 \text{ mmol}$) confirmed the occurrence of a 1 : 1 complexation (Fig. S11, ESI†).

From a toluene solution of an equimolar mixture of **1**_{C1} and C₆₀, we successfully obtained a black-colored block single crystal suitable for single-crystal X-ray analysis (ESI†). The X-ray structure showed that **1**_{C1} and C₆₀ co-crystallized with a ratio of 2 : 3 to constitute a unit cell (space group: triclinic $P\bar{1}$), where the asymmetric unit contains one **1**_{C1} and one and a half C₆₀ molecules (Fig. 6). Thus, the complexation stoichiometry of **1**_{C1} and C₆₀ in the solid state (1 : 1.5) is different from that in



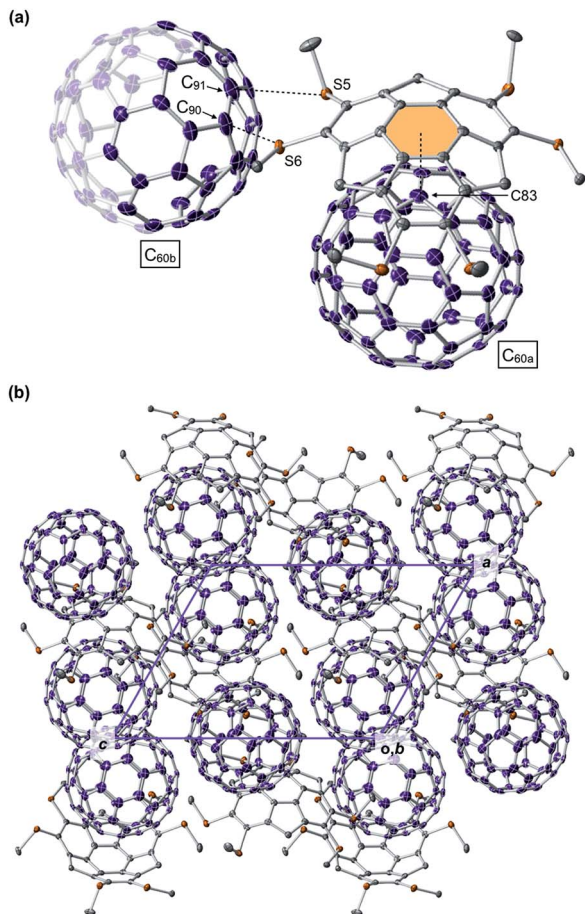


Fig. 6 (a) Asymmetric unit and (b) packing diagram in the crystal structure of $1_{C1} \cdot (C_{60})_{1.5}$. Hydrogen atoms are omitted for clarity. The asymmetric unit involves only half of a C_{60b} molecule, and the remainder is represented translucently. Colour code: carbon (1_{C1}) = grey, carbon (C_{60}) = purple and sulfur = orange.

solution ($\sim 1 : 1$). In the crystal, the two C_{60} molecules are tightly packed without disordering. One of the C_{60} molecules (C_{60a}) and 1_{C1} form a concave–convex complex, in which the shortest distance between one of the carbon atoms of C_{60a} (C_{83} , Fig. 6a) and the mean plane of the central six-membered ring of 1_{C1} is 3.370 Å. The other C_{60} molecule (C_{60b}) interacts with the thiomethyl groups of 1_{C1} (Fig. 6a), as evidenced by the intermolecular sulfur–carbon contacts, e.g., S5–C₉₀ (3.164 Å) and S6–C₉₁ (3.456 Å), which are shorter than the sum of the van der Waals radii (3.50 Å).¹² Accordingly, every C_{60b} molecule interacts with a total of 12 thiomethyl groups of six neighboring 1_{C1} molecules, leading to the formation of a two-dimensional network (Fig. S12, ESI†). This structural feature may account for the observed low solubility (i.e., facile crystallization) of the complex of 1_{C1} with C_{60} in solution, despite relatively small association constants between hexathioalkyl sumanenes and C_{60} .

Conclusions

A recently established method for full functionalization of the aromatic rings of sumanene^{1f} allowed us to investigate the

possibility that these bowl-shaped molecules could act as a mesogen for LC assemblies. As we have demonstrated in the present work, sumanene, when attached to simple thioalkyl side chains, forms a remarkably high-order columnar LC assembly, most likely due to concave–convex interactions. This result is in contrast to those with corannulene with an analogous concave–convex geometry, which requires specific side chains for the formation of a mesophase.^{2f,g} The first successful synthesis of LC sumanenes would be a first step in the development of stimuli-responsive molecular assemblies by taking advantage of the bowl-to-bowl inversion dynamics. We have also demonstrated that, unlike non-substituted sumanene, the hexathioalkyl version behaves as an electron donor, leading to the complexation with C_{60} . Since 1_{C1} shows affinity for C_{60} , which leads to facile co-crystallization from a dilute solution, hexathioalkyl sumanenes may serve as a new building block for the construction of supramolecular materials with fullerene derivatives.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by a Grant-in-Aid for Scientific Research on Innovative Areas “ π -Figuration: Control of Electron and Structural Dynamism for Innovative Functions” (JP26102008 for T. F., JP26102002 for H. S. and JP26102007 for T. A.) from the Japan Society for the Promotion of Science (JSPS). This work was also supported in part by “Dynamic Alliance for Open Innovation Bridging Human, Environment and Materials” from MEXT, Japan. The synchrotron XRD experiments were performed at the BL45XU beamline at SPring-8 with the approval of the RIKEN SPring-8 Center (proposal no. 20150068 and 20160027). Q. D. acknowledges the Program for Leading Graduate Schools “Academy for Co-Creative Education of Environment and Energy Science” from MEXT, Japan. The authors would like to thank Suzukakedai Materials Analysis Division, Technical Department, Tokyo Institute of Technology, for their support with the NMR measurements.

Notes and references

- (a) H. Sakurai, T. Daiko and T. Hirao, *Science*, 2003, **301**, 1878; (b) H. Sakurai, T. Daiko, H. Sakane, T. Amaya and T. Hirao, *J. Am. Chem. Soc.*, 2005, **127**, 11580; (c) T. Amaya, M. Hifumi, M. Okada, Y. Shimizu, T. Moriuchi, K. Segawa, Y. Ando and T. Hirao, *J. Org. Chem.*, 2011, **76**, 8049; (d) S. Higashibayashi, R. B. N. Baig, Y. Morita and H. Sakurai, *Chem. Lett.*, 2012, **41**, 84; (e) S. Higashibayashi, R. Tsuruoka, Y. Soujanya, U. Purushotham, G. N. Sastry, S. Seki, T. Ishikawa, S. Toyota and H. Sakurai, *Bull. Chem. Soc. Jpn.*, 2012, **85**, 450; (f) B. M. Schmidt, B. Topolinski, S. Higashibayashi, T. Kojima, M. Kawano, D. Lentz and H. Sakurai, *Chem.–Eur. J.*, 2013, **19**, 3282; (g) B. B. Shrestha, S. Karanjit, G. Panda, S. Higashibayashi and H. Sakurai,



- Chem. Lett.*, 2013, **42**, 386; (h) S. Higashibayashi, S. Onogi, H. K. Srivastava, G. N. Sastry, Y. Wu and H. Sakurai, *Angew. Chem., Int. Ed.*, 2013, **52**, 7314; (i) H. Toda, Y. Yakiyama, Y. Shoji, F. Ishiwari, T. Fukushima and H. Sakurai, *Chem. Lett.*, 2017, **46**, 1368; (j) B. B. Shrestha, S. Karanjit, S. Higashibayashi and H. Sakurai, *Pure Appl. Chem.*, 2014, **86**, 747; (k) N. Ngamsomprasert, G. Panda, S. Higashibayashi and H. Sakurai, *J. Org. Chem.*, 2016, **81**, 11978; (l) S. Mebs, M. Weber, P. Luger, B. M. Schmidt, H. Sakurai, S. Higashibayashi, S. Onogi and D. Lentz, *Org. Biomol. Chem.*, 2012, **10**, 2218.
- 2 (a) L. T. Scott, *Pure Appl. Chem.*, 1996, **68**, 291; (b) P. W. Rabideau and A. Sygula, *Acc. Chem. Res.*, 1996, **29**, 235; (c) Y.-T. Wu and J. S. Siegel, *Chem. Rev.*, 2006, **106**, 4843; (d) Y.-T. Wu, D. Bandera, R. Maag, A. Linden, K. K. Baldridge and J. S. Siegel, *J. Am. Chem. Soc.*, 2008, **130**, 10729; (e) A. S. Filatov, L. T. Scott and M. A. Petrukhina, *Cryst. Growth Des.*, 2010, **10**, 4607; (f) D. Miyajima, K. Tashiro, F. Araoka, H. Takezoe, J. Kim, K. Kato, M. Takata and T. Aida, *J. Am. Chem. Soc.*, 2009, **131**, 44; (g) D. Miyajima, F. Araoka, H. Takezoe, J. Kim, K. Kato, M. Takata and T. Aida, *Angew. Chem., Int. Ed.*, 2011, **50**, 7865; (h) L. N. Dawe, T. A. AlHujran, H.-A. Tran, J. I. Mercer, E. A. Jackson, L. T. Scott and P. E. Georghiou, *Chem. Commun.*, 2012, **48**, 5563; (i) B. M. Schmidt, S. Seki, B. Topolinski, K. Ohkubo, S. Fukuzumi, H. Sakurai and D. Lentz, *Angew. Chem., Int. Ed.*, 2012, **51**, 11385; (j) R. Chen, R.-Q. Lu, K. Shi, F. Wu, H.-X. Fang, Z.-X. Niu, X.-Y. Yan, M. Luo, X.-C. Wang, C.-Y. Yang, X.-Y. Wang, B. Xu, H. Xia, J. Peib and X.-Y. Cao, *Chem. Commun.*, 2015, **51**, 13768.
- 3 (a) S. Furukawa, J. Kobayashi and T. Kawashima, *J. Am. Chem. Soc.*, 2009, **131**, 14192; (b) X. Li and X. Shao, *Synlett*, 2014, **25**, 1795; (c) X. Li, Y. Zhu, J. Shao, B. Wang, S. Zhang, Y. Shao, X. Jin, X. Yao, R. Fang and X. Shao, *Angew. Chem., Int. Ed.*, 2014, **53**, 535; (d) M. Saito, S. Furukawa, J. Kobayashi and T. Kawashima, *Chem. Rec.*, 2016, **16**, 64; (e) X. Li, Y. Zhu, J. Shao, B. Wang, S. Zhang, Y. Shao, X. Jin, X. Yao, R. Fang and X. Shao, *Angew. Chem., Int. Ed.*, 2014, **53**, 535; (f) X. Hou, Y. Zhu, Y. Qin, L. Chen, X. Li, H.-L. Zhang, W. Xu, D. Zhu and X. Shao, *Chem. Commun.*, 2017, **53**, 1546; (g) S. Furukawa, Y. Suda, J. Kobayashi, T. Kawashima, T. Tada, S. Fujii, M. Kiguchi and M. Saito, *J. Am. Chem. Soc.*, 2017, **139**, 5787.
- 4 (a) Z. Wang, F. Dötz, V. Enkelmann and K. Müllen, *Angew. Chem., Int. Ed.*, 2005, **44**, 1247; (b) P. E. Georghiou, L. N. Dawe, H.-A. Tran, J. Strübe, B. Neumann, H.-G. Stammmler and D. Kuck, *J. Org. Chem.*, 2008, **73**, 9040; (c) K. Yoshida and A. Osuka, *Chem.-Asian J.*, 2015, **10**, 1526; (d) M. Yamamura, T. Saito and T. Nabeshima, *J. Am. Chem. Soc.*, 2014, **136**, 14299; (e) M. Yamamura, K. Sukegawa, D. Okada, Y. Yamamoto and T. Nabeshima, *Chem. Commun.*, 2016, **52**, 4585; (f) E. M. Pérez, M. Sierra, L. Sánchez, M. R. Torres, R. Viruela, P. M. Viruela, E. Ortí and N. Martín, *Angew. Chem., Int. Ed.*, 2007, **46**, 1847.
- 5 (a) T. Kawase and H. Kurata, *Chem. Rev.*, 2006, **106**, 5250; (b) E. M. Pérez and N. Martín, *Chem. Soc. Rev.*, 2015, **44**, 6425; (c) A. Sygula, F. R. Fronczek, R. Sygula, P. W. Rabideau and M. M. Olmstead, *J. Am. Chem. Soc.*, 2007, **129**, 3842.
- 6 (a) F. Araoka and H. Takezoe, *Jpn. J. Appl. Phys.*, 2014, **53**, 01AA01; (b) D. Miyajima, F. Araoka, H. Takezoe, J. Kim, K. Kato, M. Takata and T. Aida, *Science*, 2012, **336**, 209.
- 7 Y. Guan and S. E. Wheeler, *J. Phys. Chem. C*, 2017, **121**, 8541.
- 8 (a) T. Wöhrle, I. Wurzbach, J. Kirres, A. Kostidou, N. Kapernaum, J. Litterscheidt, J. C. Haenle, P. Staffeld, A. Baro, F. Giesselmann and S. Laschat, *Chem. Rev.*, 2016, **116**, 1139; (b) S. K. Prasad, D. S. S. Rao, S. Chandrasekhar and S. Kumar, *Mol. Cryst. Liq. Cryst.*, 2003, **396**, 121; (c) D. Adam, P. Schuhmacher, J. Simmerer, L. Häussling, K. Siemensmeyer, K. H. Etzbach, H. Ringsdorf and D. Haarer, *Nature*, 1994, **371**, 141; (d) T. Osawa, T. Kajitani, D. Hashizume, H. Ohsumi, S. Sasaki, M. Takata, Y. Koizumi, A. Saeki, S. Seki, T. Fukushima and T. Aida, *Angew. Chem., Int. Ed.*, 2012, **51**, 7990.
- 9 K. C. Cao, *Dielectric Phenomena in Solids*, Elsevier Academic Press, Boston, 2004.
- 10 (a) F. Diederich and M. Gómez-López, *Chem. Soc. Rev.*, 1999, **28**, 263; (b) S. Mizyed, P. E. Georghiou, M. Bancu, B. Cuadra, A. K. Rai, P. Cheng and L. T. Scott, *J. Am. Chem. Soc.*, 2001, **123**, 12770; K. Tashiro and T. Aida, *Chem. Soc. Rev.*, 2007, **36**, 189. (c) M.-X. Wang, *Chem. Commun.*, 2008, 4541; (d) M. Hardouin-Lerouge, P. Hudhomme and M. Sallé, *Chem. Soc. Rev.*, 2011, **40**, 30.
- 11 K. A. Connors, *Binding Constants: The Measurement of Molecular Complex Stability*, John Wiley & Sons. Inc., New York, 1987.
- 12 M. Mantina, A. C. Chamberlin, R. Valero, C. J. Cramer and D. G. Truhlar, *J. Phys. Chem. A*, 2009, **113**, 5806.

