

Cite this: *RSC Adv.*, 2018, 8, 7040

Coaxial heterojunction carbon nanofibers with charge transport and electrocatalytic reduction phases for high performance dye-sensitized solar cells†

Yuan-Hua Wang,^a Hai-Qiu Fang,^a Qiang Dong,^a Duan-Hui Si,^b Xue-Dan Song,^b Chang Yu^a and Jie-Shan Qiu^{id}*^{ac}Received 7th December 2017
Accepted 31st January 2018

DOI: 10.1039/c7ra13118f

rsc.li/rsc-advances

Novel coaxial heterojunction carbon nanofibers, fabricated by electro-spinning a mixture of hydro-pitch and polyacrylonitrile, served as the counter electrode for dye-sensitized solar cells. Their high power conversion efficiency, being comparable to that of Pt CE, was achieved due to their good conductivity and high heteroatom content.

The counter electrode (CE) in dye-sensitized solar cells (DSSCs) has multiple functions, including the catalytic reduction of I_3^- to I^- and the regeneration of dye molecules, and is one of the key dominant components that governs the practical applications of DSSCs to a great degree.¹ Noble metal platinum (Pt) with its low electrical resistance and excellent electrocatalytic activity was the first to be applied as such and is now widely used as CEs in DSSCs. However, there is limited availability of Pt sources, and this has led to the high cost of CEs, which has hindered the practical applications of DSSCs.² To date, various efforts have been made to develop techniques for the production of inexpensive yet high performance CEs to replace the Pt CEs. Of the alternatives now available, carbon materials that are categorized as zero-dimensional (0D) to three-dimensional (3D) structures have attracted much attention as candidates for Pt-free electrodes in view of the advantages of low cost as well as good electrochemical stability; these include carbon black,³ carbon nanoparticles,⁴ carbon nanotubes,^{5–9} graphene,^{10–14} and graphite¹⁵ and their composites.^{16–19} It is believed that a large quantity of defects in the carbon CEs may lead to high catalytic activities, while good electric conductivity can result in fast charge transportation. Nevertheless, how to combine and integrate the high electrical conductivity and abundant defects into one carbon electrode

in a balanced way to fabricate high performance CEs with tuned structure remains a major challenge.

Recently, one dimensional (1D) core-shell nanofibers have received much attention because a good combination of electrical conductivity and catalytic activity can be realized and achieved in one electrode material.²⁰ However, the core and shell in these fibers are two separate phases, which limits the rapid charge transition and increases the electrical resistance when such nanofibers are used as CEs in DSSCs. Moreover, the tedious and complicated fabrication process greatly limits their large-scale production. As such, it is necessary to explore a new approach for this kind of material, in which electrical conductivity and catalytic activity are combined well.

Herein, we present 1D coaxial carbon nanofibers (CNFs) fabricated by the electrospinning method from two kinds of carbon precursors: hydrogenated pitch and polyacrylonitrile (PAN). The adopted hydro-pitch (HP) features planar aromatic hydrocarbon molecules and is more easily transformed into an optically anisotropic, graphitizable carbon structure,²¹ thus leading to high conductivity (Table S1, ESI†). PAN tends to form a carbonaceous structure with high nitrogen (N) content and abundant defects; it is widely acknowledged that high electrochemical activities could be achieved by the different nitrogen species and defects.^{7,16,17,22–24} Benefiting from different nucleation mechanisms, a novel coaxial core/shell structure has been produced within the matrix, in which hydro-pitch contributes to the core phase, while PAN is responsible for the shell phase. Such CNFs with two phases and heterogeneous characteristics are very unique when they are applied as CEs in DSSCs, with the core playing the role of cable tunnel for charge transporting and the shell providing active sites for catalyzing the reduction of I_3^- to I^- , as shown in Scheme 1.

The process of fabricating our CNFs (denoted as 0.15-HCNF and 1-HCNF, in which 0.15 and 1 are the weight ratios between

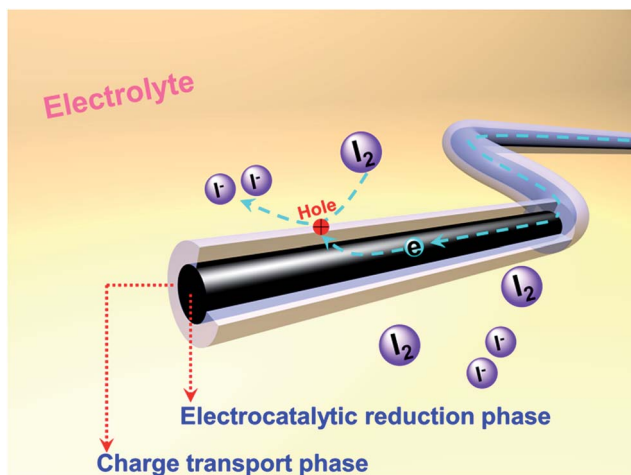
^aCarbon Research Laboratory, Liaoning Key Lab for Energy Materials and Chemical Engineering, State Key Lab of Fine Chemicals, School of Chemical Engineering, Dalian University of Technology, Dalian 116024, China. E-mail: jqiu@dut.edu.cn

^bState Key Laboratory of Fine Chemical, School of Chemistry, Dalian University of Technology, Dalian 116024, China

^cSchool of Chemical Engineering and Technology, Xi'an Jiaotong University, Xi'an, 710049, China

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c7ra13118f





Scheme 1 Illustration of the structure and functions of heterojunction carbon nanofibers as CEs in DSSCs.

hydro-pitch and PAN) is illustrated in Fig. S5,[†] and the morphologies of different CNFs and their corresponding diameter distributions are shown in Fig. 1a–c and a₁–c₁. The average diameter is 0.33 μm for PAN CNFs (PCNF), 0.36 μm for 0.15-HCNF, and 0.76 μm for 1-HCNF, indicative of a size-increased behavior with the increase in the ratio between the hydro-pitch and PAN. The obtained CNFs in sample 1-HCNF are cross-linked or joined together, as can be seen in Fig. 1c, leading to the inner parts being exposed and the observation of different phases.

The HCNFs were examined in much detail by transmission electron microscopy (TEM) to yield information about the inner structure, and the typical TEM images are shown in Fig. 2. It can

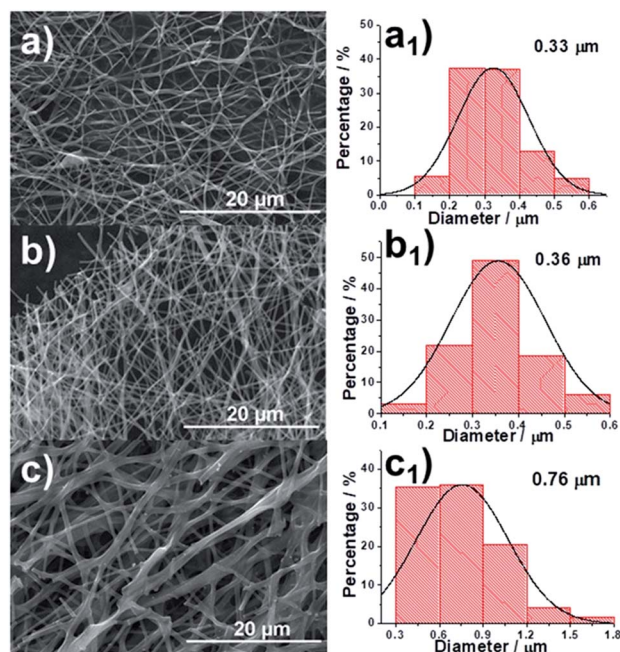


Fig. 1 Scanning electron microscopy (SEM) images of (a) PCNF, (b) 0.15-HCNF and (c) 1-HCNF.

be clearly seen that two phases in the axial direction are obviously observed in the HCNF samples (Fig. 2b and c) in comparison to PCNF (Fig. 2a). Energy-dispersive X-ray spectra (EDX) of the shell region show the high nitrogen content of 9.3 at% (Fig. 2e), which implies that the carbon shell is derived from PAN. To further gain insight into the contributions of these two precursors, sample 1-HCNF was treated with KOH at high temperature to remove the shell, and the corresponding sample was further characterized by TEM and EDX, the detailed results of which are shown in Fig. 2d and f. It can be seen that the shell layer was removed, and the content of nitrogen was only 1 at% in the core phase (Table S2, ESI[†]), in comparison to there being only 0.4 wt% (Table S3, ESI[†]) of nitrogen in the hydro-pitch.

With all of this information in mind, it was deduced that the core phase was mainly derived from the hydro-pitch, while the shell was mainly from PAN.

It is interesting that this result is different from that reported by Yang.²⁵ In general, the low molecular weight pitch tends to be pushed to the outer surface during the solvent evaporation process.^{26,27} Nevertheless, in the present system the hydro-pitch is more easily fixed in the inner part, which was evidenced by density functional theory (DFT) calculations (Fig. 3). As shown in Fig. 3a, the binding energy between the hydro-pitch and two PAN units (E_{HP}) is 5.6 kcal mol⁻¹, which is 1.3 kcal mol⁻¹ higher than that between pitch and PAN units (E_{P} , 4.3 kcal mol⁻¹). When three PAN units were applied, an increase of 25% was observed for E_{HP} , which reached 7.0 kcal mol⁻¹. The newly formed hydrogen bonding between aliphatic hydrogen in hydro-pitch and the cyano group in the PAN unit is responsible

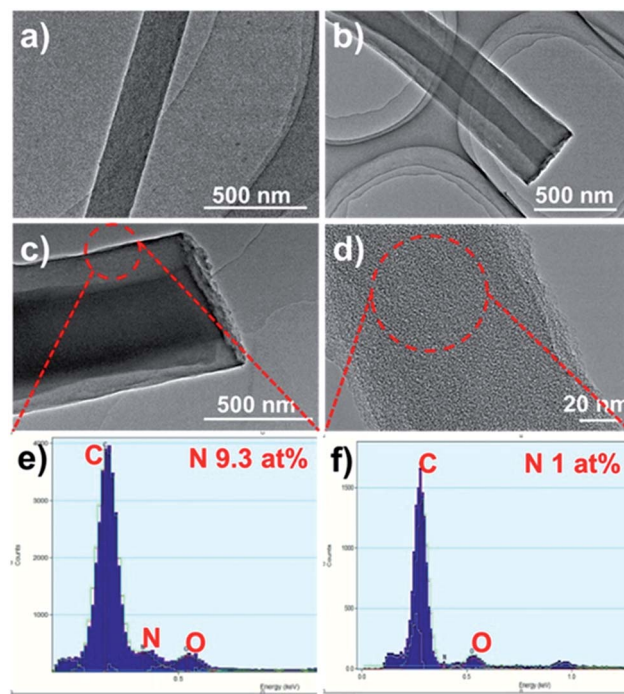


Fig. 2 TEM images of (a) PCNF, (b) 0.15-HCNF, (c) 1-HCNF, and (d) the core of the 1-HCNF fiber after KOH treatment for 1 h at 700 °C in N₂. (e and f) The EDX spectra of the fiber shell and core.



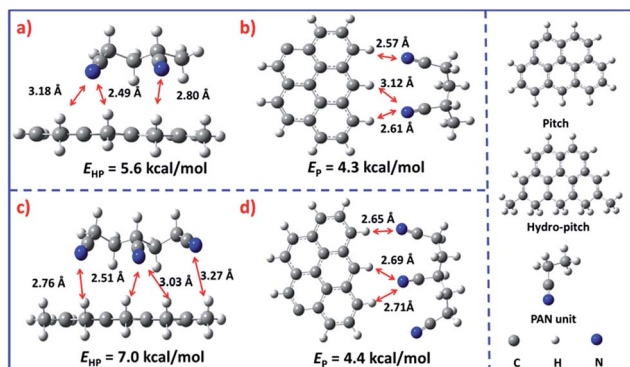


Fig. 3 The spatial configurations and binding energies between two kinds of pitch molecules and PAN units. (a) Hydro-pitch with two PAN units, (b) molecular pitch and two PAN units, (c) hydro-pitch and three PAN units, (d) molecular pitch and three PAN units.

for the remarkable increase in binding energy, while the number of E_p only slightly increased from 4.3 to 4.4 kcal mol⁻¹, and no new bonding was observed. As a result, compared to the pitch, the hydro-pitch could be more easily fixed in PAN units due to the strong binding caused by the multi-interaction between the aliphatic hydrogen and cyano groups. In addition, the hydro-pitch is always wrapped by several PAN long chain molecules, as shown in Fig. S5.† Therefore, a kind of heterojunction core-shell structure was finally produced, in which the core was attributed to the hydro-pitch while the shell was attributed to PAN.

The HCNF samples were also analyzed by X-ray photoelectron spectroscopy (XPS) and Raman spectra to reveal the changes in the surface configuration caused by the addition of hydro-pitch (Fig. S2, ESI†). Compared to PCNF (7.6 at%, Table S4, ESI†), sample 0.15-HCNF maintained 4.6 at% nitrogen content, and the same I_D/I_G ratio (1.07) was also obtained for the two samples. With an increase in hydro-pitch in the case of 1-HCNF, a slightly lower nitrogen content (4.2 at%) and I_D/I_G number (1.01) were observed, which could be attributed to the exposure of the hydro-pitch-based core, as mentioned in Fig. 1c. The XPS spectrum of N 1s is shown in Fig. S3 in the ESI† and the ratios of different nitrogen species are summarized in Table S5.† It is well known that these nitrogen species in the carbon network could produce highly electrocatalytically active sites,^{22,23} while the defects in carbon materials could also play the same role.

The HCNFs exhibited unique characteristics in which the shell retained the high nitrogen content and defects, while the core derived from hydro-pitch featured high conductivity. Such an integrated structure in one electrode is so attractive that it could be used as a high-performance CE in DSSCs, as shown in Scheme 1. In this case, it has great potential as an electrocatalyst for Pt replacement in DSSCs. Benefiting from this, the device performance for DSSCs with PCNF, Pt, and two HCNF CEs was determined and the results are shown in Table 1.

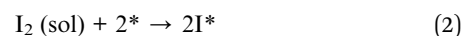
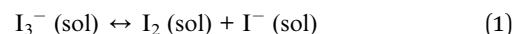
As shown in Fig. 4a, our HCNF CEs exhibit similar or even better performance compared with Pt CE, and the

Table 1 Characteristics of the $J-V$ curves for DSSCs with various working electrodes

Samples	V_{oc}/V	$J_{sc}/\text{mA cm}^{-2}$	FF	$\eta/\%$
0.15-HCNF	0.76	14.16	0.65	6.92 ± 0.15
1-HCNF	0.75	13.48	0.63	6.32 ± 0.15
PCNF	0.76	12.25	0.46	4.26 ± 0.15
Pt	0.73	13.28	0.65	6.34 ± 0.15

electrochemical performance varies with the ratio of the hydro-pitch to PAN. The sample 0.15-HCNF CE shows a higher J_{sc} of 14.16 mA cm⁻², a fill factor (FF) of 0.65 and a power conversion efficiency of 6.92%. In contrast, the J_{sc} , FF and power conversion efficiency (η) for the DSSCs with PCNF CE are only 12.25 mA cm⁻², 0.46 and 4.26%, respectively; in particular, an increase in efficiency of 62.4% was observed. Nevertheless, when the ratio of hydro-pitch and PAN increased to 1, the performance of DSSCs with 1-HCNF electrode was reduced to J_{sc} of 13.48 mA cm⁻², FF of 0.65 and η of 6.32%. The reason for this is attributed to the lower contents of nitrogen and defects, nitrogen and defects, and the loss of the Brunauer Emmet Teller (BET) surface (Table S6, ESI†).

To further demonstrate the unique effects of the coaxial CNF, to further demonstrate the unique effects of the coaxial CNF, the possible mechanism involved in the process was proposed and shown in Fig. 4b. As is known, the general consensus for the reaction mechanism can be described as follows:²⁸



After the desorption of the solvated $I^- (\text{sol})$ into the electrolyte, the activated site lost one electron and high-rate regeneration was demanded for the active site. The electrons from the external circuit are typically transferred from FTO to CE and then transmitted in CEs to the activated catalytic sites. The key in this process is the abundant supply of electrons to the activated sites within a short diffusion time. For our HCNF CE, the 1D tunnel structure minimizes the loss of electrons for its low electrical resistance, and high-rate regeneration of the

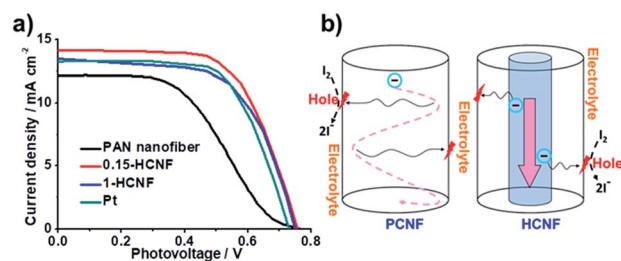


Fig. 4 (a) Current density–voltage ($J-V$) characteristics with different CEs: PCNF, Pt, 1-HCNF, and 0.15-HCNF. (b) Illustration of charge transport and diffusion in different CNFs.



active site is also realized by direct electron diffusion from the core to the activated catalytic site, due to its heterojunction structure. Benefiting from the unique structure, highly efficient utilization of the activated catalytic sites and fast charge transfer derived from short transport distance are achieved, and high power-conversion efficiency is yielded, which is comparable to or even better than other metal-free CEs reported in the literature (Table S8†).

To further investigate the electrochemical characteristics of HCNF CEs, Nyquist plots on an asymmetric dummy were obtained for Pt, PCNF and 0.15-HCNF electrodes (Fig. S4, ESI†).²⁹ In contrast to Pt, both kinds of CNF CEs have the same lower R_s value, 1.6 Ω cm² (Table S7, ESI†), indicating that good adhesion between CNFs and FTO was realized. High R_{ct} values of 6.4 Ω cm² for PCNF and 4.3 Ω cm² for 0.15-HCNF CE (Table S7, ESI†) were also observed. In the future, further surface modification needs to be carried out to reduce the R_{ct} value,^{17–19} thus optimizing and improving the electrochemical performance of CNFs.

In summary, the coaxial heterojunction CNFs with two phases were fabricated from PAN/hydro-pitch blend precursors using the electrospinning method, in which the shell and core were constructed by PAN based carbon and hydro-pitch based carbon, respectively. The as-made coaxial CNFs exhibited high conductivity derived from hydro-pitch, and high nitrogen content and defects derived from PAN. Such HCNFs were employed as CEs in DSSCs, and high power-conversion efficiencies were delivered, which are comparable to that of the Pt CE constructed under the same conditions. The tunnel along the axis direction separates the fibers into the charge transport phase and the electrocatalytic phase, and high utilization of catalytic sites is realized by the abundant charge supply and fast electron-hole recombination. In addition, this work shows that low cost and large-scale production of high performance CEs in DSSCs can be realized using the electrospinning method.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was partly supported by the NSFC of China (No. 21336001, 21522601), the Fundamental Research Funds for the Central Universities of China (DUT16ZD217), and the Education Department of the Liaoning Province of China (No. T2013001).

References

- B. O. Oregan and M. Gratzel, *Nature*, 1991, **353**, 737–781.
- S. C. Hou, X. Cai, H. W. Wu, X. Yu, M. Peng, K. Yan, *et al.*, *Energy Environ. Sci.*, 2013, **6**, 3356–3362.
- K. Imoto, K. Takahashi, T. Yamaguchi, T. Komura, J. I. Nakamura and K. Murata, *Sol. Energy Mater. Sol. Cells*, 2003, **79**, 459–469.
- R. R. Jia, J. Z. Chen, J. H. Zhao, J. F. Zheng, C. Song, L. Li, *et al.*, *J. Mater. Chem.*, 2010, **20**, 10829–10834.
- Z. B. Yang, T. Chen, R. X. He, G. Z. Guan, H. P. Li, L. B. Qiu, *et al.*, *Adv. Mater.*, 2011, **23**, 5436–5439.
- J. Han, H. Kim, D. Y. Kim, S. M. Jo and S. Y. Jang, *ACS Nano*, 2010, **4**, 3503–3509.
- W. J. Lee, E. Ramasamy, D. Y. Lee and J. S. Song, *ACS Appl. Mater. Interfaces*, 2009, **1**, 1145–1149.
- P. Dong, C. L. Pint, M. Hainey, F. Mirri, Y. J. Zhan, J. Zhang, *et al.*, *ACS Appl. Mater. Interfaces*, 2011, **3**, 3157–3161.
- G. Calogero, F. Bonaccorso, O. M. Marago, P. G. Gucciardi and G. Di Marco, *Dalton Trans.*, 2010, **39**, 2903–2909.
- H. Wang, K. Sun, F. Tao, D. J. Stacchiola and Y. H. Hu, *Angew. Chem., Int. Ed.*, 2013, **52**, 9210–9214.
- J. D. Roy-Mayhew, D. J. Bozym, C. Punckt and I. A. Aksay, *ACS Nano*, 2010, **4**, 6203–6211.
- D. W. Zhang, X. D. Li, H. B. Li, S. Chen, Z. Sun, X. J. Yin, *et al.*, *Carbon*, 2011, **49**, 5382–5388.
- L. Kavan, J. H. Yum and M. Gratzel, *Nano Lett.*, 2011, **11**, 5501–5506.
- L. Kavan, J. H. Yum and M. Gratzel, *ACS Nano*, 2011, **5**, 165–172.
- G. Veerappan, K. Bojan and S. W. Rhee, *ACS Appl. Mater. Interfaces*, 2011, **3**, 857–862.
- Y. H. Xue, J. Liu, H. Chen, R. G. Wang, D. Q. Li, J. Qu, *et al.*, *Angew. Chem., Int. Ed.*, 2012, **51**, 12124–12127.
- X. T. Meng, C. Yu, X. D. Song, Y. Liu, S. X. Liang, Z. Q. Liu, *et al.*, *Adv. Energy Mater.*, 2015, **5**, 1500180.
- C. Yu, H. Q. Fang, Z. Q. Liu, H. Hu, X. T. Meng and J. S. Qiu, *Nano Energy*, 2016, **25**, 184–192.
- H. Q. Fang, C. Yu, T. L. Ma and J. S. Qiu, *Chem. Commun.*, 2014, **50**, 3328–3330.
- X. Fang, Z. B. Yang, L. B. Qiu, H. Sun, S. W. Pan, J. Deng, *et al.*, *Adv. Mater.*, 2014, **26**, 1694–1698.
- I. Mochida, S. H. Yoon and Y. Korai, *Chem. Rec.*, 2002, **2**, 81–101.
- K. P. Gong, F. Du, Z. H. Xia, M. Durstock and L. M. Dai, *Science*, 2009, **323**, 760–764.
- R. L. Liu, D. Q. Wu, X. L. Feng and K. Müllen, *Angew. Chem., Int. Ed.*, 2010, **49**, 2565–2569.
- L. Y. Heng, A. Chou, J. Yu, Y. Chen and J. J. Gooding, *Electrochem. Commun.*, 2005, **7**, 1457–1462.
- B. H. Kim, K. S. Yang, Y. A. Kim, Y. J. Kim, B. An and K. Oshida, *J. Power Sources*, 2011, **196**, 10496–10501.
- J. S. Travis and A. R. Horst, *Biomaterials*, 2008, **29**, 1989–2006.
- S. Megelski, J. S. Stephens, D. B. Chase and J. F. Rabolt, *Macromolecules*, 2002, **35**, 8456–8466.
- Y. Hou, D. Wang, X. H. Yang, W. Q. Fang, B. Zhang, H. F. Wang, *et al.*, *Nat. Commun.*, 2013, **4**, 1583.
- J. Z. Yin, Z. N. Yu, F. Gao, J. J. Wang, H. Pang and Q. Y. Lu, *Angew. Chem., Int. Ed.*, 2010, **49**, 6328–6332.

