



Cite this: *RSC Adv.*, 2018, 8, 26318

## An alternative method to enhance w/o emulsion stability using modified dimer acid and its application in oil based drilling fluids

Xianbin Huang, \* Jinsheng Sun, Kaihe Lv, Jingping Liu and Haokun Shen

This study presents an alternative method to enhance the emulsion stability of oil based drilling fluids (OBMs). Modified dimer acid (MDA) was synthesized with a molecular structure having two hydrophilic heads and two hydrophobic tails. Theoretically, the adsorption of MDA on an oil–water interface makes it possible to increase hydrogen bonding between water droplets and form three dimensional networks which benefit emulsion stability. The influence of MDA on the stability of base emulsions was studied by visual observation. Then the stabilization mechanism of MDA was analyzed from the micro and macro points of view by morphology study using a cryo-scanning electron microscope (cryo-SEM) and rheological measurements including viscosity vs. shear rates, zero-shear viscosity ( $\eta_0$ ), and creep and recovery tests. Experimental results showed that a substantial improvement in emulsion stability was visually observed when the MDA concentration was 2 g L<sup>-1</sup>. From cryo-SEM observation, a honeycomb structure was observed in the emulsion containing 2 g L<sup>-1</sup> MDA, which can provide a physical barrier to restrain the movement of water droplets. In comparison with the rheological behaviors of the emulsion without MDA, a remarkably larger zero-shear viscosity, a solid-like behavior and a greater elasticity were observed when 2 g L<sup>-1</sup> MDA was present. Finally, the application in OBMs shows that MDA can largely enhance electrical stability (ES) and reduce the filtration volume. The method proposed in the paper could be used to enhance the stability of w/o emulsions in a variety of fields.

Received 15th March 2018  
Accepted 13th July 2018

DOI: 10.1039/c8ra02293c

rsc.li/rsc-advances

### 1. Introduction

During drilling of subterranean oil and gas resources, drilling fluid<sup>1</sup> plays an important role in carrying drill cuttings, lubricating and cooling drill bits, balancing formation pressure, *etc.* Oil based drilling fluids (OBMs) also known as invert emulsion drilling fluids, typically consisting of an internal water phase, continuous oil phase, emulsifiers, fluid loss additives and weighting materials, are often used in the drilling of deep wells, high temperature wells, shale gas wells, and horizontal and complicated wells for their superior performances<sup>1</sup> over water based drilling fluids. Emulsion stability is the basis to obtain other functional properties of OBMs. OBMs are typically formulated using two kinds of emulsifiers to achieve a better stability: primary emulsifiers (*e.g.* calcium soaps) and secondary emulsifiers (*e.g.* polyamides). OBMs do not easily lose emulsion stability during normal drilling operations due to continuous shearing and stirring. However, the longer-time emulsion stability is highly demanded when the circulation of OBMs is ceased, such as running or pulling drill pipe into or out of a wellbore that may take days for deep wells.

Since emulsions are thermodynamically unstable, various breakdown phenomena<sup>2–4</sup> may occur such as creaming and sedimentation, flocculation, coalescence, *etc.* How to maximize emulsion stability has been intensively studied in various industries, such as food,<sup>5,6</sup> cosmetics,<sup>7</sup> petroleum development,<sup>8</sup> and pharmaceuticals.<sup>9</sup>

The common methods to enhance emulsion stability include use of fine particles<sup>8,10</sup> and high-molecular-weight polymers.<sup>11,12</sup> Fine particles can stabilize emulsions by providing steric hindrance to drop-drop coalescence and modifying the rheological properties of the interfacial region. It has been demonstrated that fine particles stabilized emulsion, often known as “Pickering emulsion<sup>13,14</sup>”, can have better stability by the use of partially hydrophobic particles<sup>15</sup> with contact angle close to 90°. High-molecular-weight polymers can act as thickeners to increase the viscosity of continuous phase to overcome gravity stress and modify the rheology of emulsions. Specifically, amphiphilic polymers<sup>16</sup> can also act as polymeric emulsifiers that can adsorb on the oil–water interface.

The structuring of internal phases using structuring agent<sup>17–19</sup> to form three dimensional networks in emulsions has received intensive attention. According to many studies, the network structures are beneficial to achieve better emulsion stability due to enhanced interactions between internal droplets. For example, Eric Dickinson and Karin Pawlowsky<sup>20</sup> used

Department of Petroleum Engineering, China University of Petroleum (East China), Qingdao, Shandong 266580, P. R. China. E-mail: 634640236@qq.com; Tel: +8615621081336



a kind of polysaccharide ( $\epsilon$ -carrageenan) to form a gel-like emulsion network in a protein-stabilized emulsion. Utai Klin-kesorn and co-workers<sup>21</sup> observed a decrease in creaming rate when studying the influence of maltodextrin on stability of corn-oil-in-water emulsions. The enhancement in emulsion stability is due to network formation of droplets according to his analysis. However, most of previous studies are concentrated on stability of oil in water emulsions. And no reference has been conducted to study the influence of structuring agents on the performance of OBMs.

In this paper, we firstly synthesized the modified dimer acid having two hydrophilic heads and two hydrophobic tails in one molecule. The performance of MDA to enhance emulsion stability was then visually observed using bottle tests. The emulsion stabilizing mechanism was studied on the micro level in terms of morphology and macro level using rheological measurements. Finally, the influence of MDA on the property of OBMs was studied.

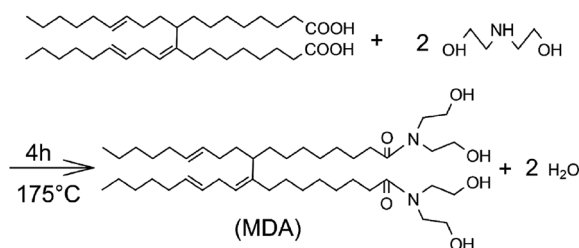
## 2. Experimental section

### 2.1 Reagents

Octadecadienoic acid dipolymer (OAD, dibasic acid > 98 wt%) was purchased from Jiangxi Aturex Co., Ltd. Diethanolamine (DEA, 99.5 wt%) and sorbitan monooleate (Span 80, 99 wt%) were obtained from Aladdin Company (Shanghai, China). Mineral oil (density = 0.88 g cm<sup>-3</sup> at 25 °C, viscosity = 244 mPa s at 20 °C) was purchased from J&K Scientific Company (Beijing, China). Drilling fluid additives include primary emulsifier (modified fatty acid), secondary emulsifier (polyaminated fatty acid), lecithin (soybean lecithin), organo-bentonite (modified natural bentonite with quaternary ammonium salt), and blown asphalt were all obtained from Halliburton Company. Barite (BaSO<sub>4</sub>,  $\rho$  = 4.3 g cm<sup>-3</sup>) was obtained from Sichuan Zhengrong Company, China.

### 2.2 MDA synthesis

MDA was synthesized as shown in Scheme 1 using OAD and DEA at a mole ratio of 1 : 2.05. The reaction was carried at 175 °C for 4 hours in an oil bath with a continuous stirring rate of 250 rpm. The water generated in the reaction was removed by a Dean–Stark trap.



Scheme 1 The synthesis process of MDA using octadecadienoic acid dipolymer and diethanolamine.

### 2.3 Preparation and characterization of base emulsions

#### (1) Preparation of base emulsions and stability observation.

3 g Span 80 and different amount of MDA (0, 0.1 g, 0.2 g, 0.3 g, 0.4 g, 0.5 g, 0.6 g, 0.8 g and 1.0 g) were initially dissolved in 80 mL mineral oil at room temperature to form homogeneous oil solutions. Then, 20 mL deionized water was added to each oil solution. Emulsions were prepared by a high-shear mixer at 5000 rpm for 5 min. A degassing process was then carried out by low-speed stirring at 200 rpm for 12 h. The 20 mL sample bottles were charged with 20 mL emulsion samples with MDA concentrations of 0 g L<sup>-1</sup>, 1 g L<sup>-1</sup> to 10 g L<sup>-1</sup>, respectively. The concentrations of MDA are based on the sum of oil volume and water volume.

(2) **Morphology.** Cryo-SEM was utilized to study the morphology of the emulsion samples with and without MDA. A small amount of an emulsion sample was placed onto a sample holder and immediately immersed into a liquid nitrogen bath (< -140 °C) for 2 min. The sample was then transferred to the high-vacuum chamber, freeze-fractured, and spray-coated with platinum. Finally, the freeze-fractured surface was observed by a Hitachi S-4300 microscope (Japan) at -135 °C.

(3) **Rheological measurements.** Rheological properties can provide information on the stability of emulsions. This part used rheology to study the stability of base emulsions. A Haake MARS III rheometer (Thermo Scientific, Karlsruhe, Germany) equipped with a cone and plate geometry (C35/1°, inclination angle 1° and gap 0.053 mm) was used to measure rheological performance of the base emulsions.

The following rheological measurements were performed: viscosity at different shear rates, zero-shear viscosity ( $\eta_0$ ) measurements, and creep and recovery tests. All the rheological measurements were carried out at 20 °C. The viscosities of the prepared base emulsion samples were measured under different shear rates from 1 s<sup>-1</sup> to 200 s<sup>-1</sup>. A constant stress of 0.1 Pa, which is inside the linear viscoelastic region for all studied emulsion samples, was applied to the following rheological measurements. (1)  $\eta_0$  was measured by oscillation method and approximated as the complex viscosity at 0.001 Hz as recommended in the literature.<sup>22</sup> (2) For creep and recovery tests, both the creep time and the recovery time were 2 min.

### 2.4 Application of MDA in OBMs

(1) **Preparation of OBMs.** Five OBMs were formulated using a FANN Hamilton Beach Mixer (Fann Instrument Company, USA) at 5000 rpm. The OBMs were composed of mineral oil, brine (25 wt% CaCl<sub>2</sub>), emulsifiers, MDA (0–5 g L<sup>-1</sup>), lecithin, organo-bentonite, blown asphalt, CaO and weighting agent barite. The adding order of the components and the amounts used in the formulations are presented in Table 1. The density of the prepared OBMs is 2.0 g cm<sup>-3</sup>.

(2) **Electricity stability (ES) test and aging treatment.** ES of the newly formulated OBMs were determined at 50 °C using a Model 23E Electrical Stability Tester (FANN) according to API RP 13B-2: Recommended Practice for Field Testing Oil-based Drilling Fluids (5th edition).<sup>23</sup>



Table 1 Formulation of oil based fluids

Adding order	Components	Amounts	Function
1	Mineral oil	320 mL	Base oil
2	Emulsifiers	14 g	Two emulsifiers
3	MDA	0 g, 0.4 g, 0.8 g, 0.12 g, 0.20 g	Structuring agent
4	Lecithin	2 g	Wetting agent
5	Brine	80 mL	Internal phase
6	Organo-bentonite	4 g	Thickener
7	Blown asphalt	8 g	Fluid loss additive
8	CaO	12 g	pH enhancer
9	Barite	856 g	Weighting agent

After that, the OBMs containing no MDA (control group) and 2 g L<sup>-1</sup> MDA were transferred into two high-temperature aging cells (175 mL) in a roller oven (Model 705ET, FANN). The temperature was set at 150 °C. Roller power was kept closed so the OBMs can be kept static in the oven to simulate the situation when the circulation stops. ES was measured every 12 hours at 50 °C. The concentrations of MDA in Sections 2.4 and 3.6 are based on the sum of oil volume and brine volume in OBMs.

**(3) High temperature high pressure (HTHP) static filtration volume.** Another five high-temperature aging cells were used. The five newly formulated OBMs were aged for 16 h at 150 °C in a roller oven where roller power was kept opened. After aging treatment, HTHP filtration volumes of OBMs were tested by a HTHP filter press (175 mL) at 150 °C under a differential pressure of 3.5 MPa according to procedures described in API RP 13B-2.<sup>23</sup>

## 3. Results and discussion

### 3.1 Three-day stability of emulsions

Fig. 1 visualizes the stability of emulsions with different MDA concentrations after the emulsions were let stand for three days. As shown in Fig. 1, the emulsion without MDA (pure emulsion) displayed a very serious sedimentation phenomenon. While the emulsions containing MDA showed a slight sedimentation phenomenon. The optimum emulsion stability was achieved when MDA concentration was 2 g L<sup>-1</sup>. But when MDA



Fig. 1 Image of emulsion samples with various MDA concentrations after standing for 3 days.

concentration exceeded 2 g L<sup>-1</sup>, the stability of the emulsions was starting to decline.

Flocculation was observed for all the emulsions containing MDA. For pure emulsion + 2 g L<sup>-1</sup> MDA, flocculation was weak. This phenomenon could be interpreted in terms of attractive interactions. When MDA concentration was low, attractive interaction was low and flocculation was weak. As MDA concentration increased, the attractive interactions increased, leading to stronger aggregation of the water droplets.

Morphology and rheological tests were conducted to explore the stabilization mechanisms at the optimal MDA concentration from micro and macro perspective in Sections 3.3 and 3.4, respectively.

### 3.2 Morphology

Cryo-SEM was utilized to visualize two emulsion samples: pure emulsion and pure emulsion + 2 g L<sup>-1</sup> MDA. Fig. 2 shows the cryo-SEM micrographs of pure emulsion (a) and pure emulsion + 2 g L<sup>-1</sup> MDA (b–d). Fig. 2a is the micrograph of pure emulsion where water droplets are embedded in continuous phase, which is similar to the structure of most ordinary emulsions described in many literatures.<sup>24,25</sup> The images of Fig. 2b–d were taken from the same freeze-fractured surface of pure emulsion + 2 g L<sup>-1</sup> MDA. Fig. 2b has the similar structure to Fig. 2a but with a larger droplet size, one reason of which is that MDA enhance attractions of water droplets and lead to weak flocculation and therefore larger particle size. Both oil and water at the surface of emulsion samples sublimated during sample preparation and SEM observation processes. Water was easier to sublimate than oil. So the sublimation of both oil and water and the heavier sublimation of water made it possible to see the honeycomb structure of continuous phase as shown in Fig. 2c and d. The holes in Fig. 2c are the positions of water droplets before sublimation. For clarity, the distribution of water droplets was depicted in Fig. 2c. However, the honeycomb structure was not observed in freeze-fractured surface of pure emulsion.

The honeycomb structure indicates the existence of “gel-like” structures of emulsions as described in many studies.<sup>6,26,27</sup> The formation of honeycomb structure in continuous phase needs further explanation because MDA did not affect the viscosity of mineral oil as shown in Section 3.4. On the other hand, as is explained in these studies,<sup>6,26,27</sup> the honeycomb structure observed in continuous phase can provide a physical



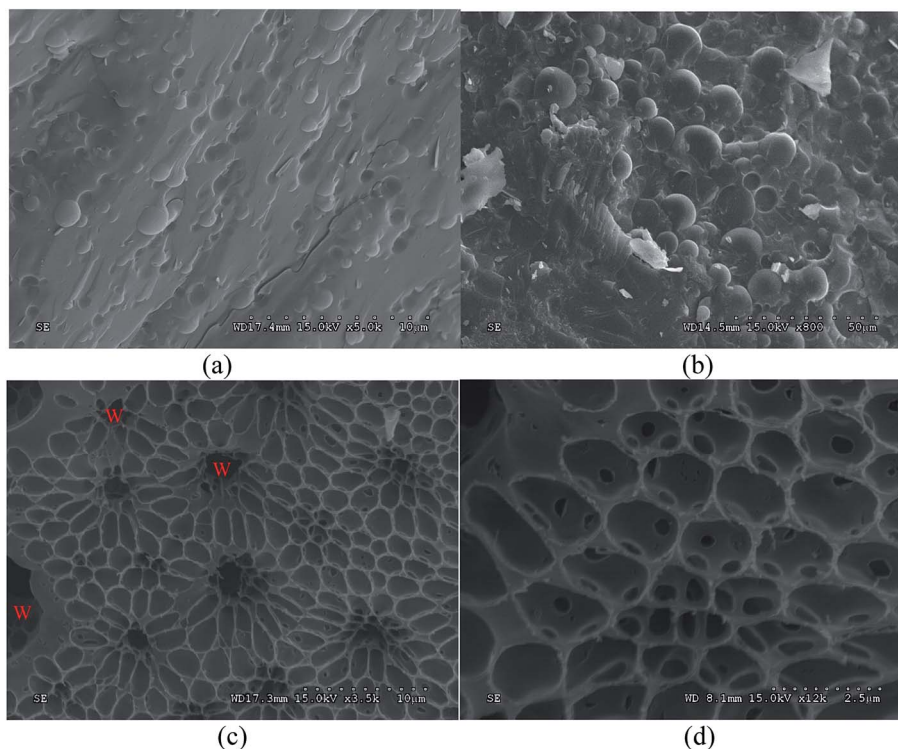


Fig. 2 Cryo-SEM micrographs of pure emulsion (a) and pure emulsion + 2 g L<sup>-1</sup> MDA (b–d).

barrier to restrain the movement of internal droplets, which contributes to emulsion stability. Thus, the cryo-SEM results confirmed MDA's contribution in structuring the studied w/o emulsions.

### 3.3 Rheological measurements

**(1) Viscosity measurement.** Aiming at evaluating the effect of MDA concentration on the viscosity of oil, the viscosities of five oil samples (80 mL oil + 3 g Span 80, 80 mL oil + 3 g Span 80 + 0.1 g MDA, 80 mL oil + 3 g Span 80 + 0.2 g MDA, 80 mL oil + 3 g Span 80 + 0.3 g MDA and 80 mL oil + 3 g Span 80 + 0.5 g MDA) were measured. Experimental results indicate that all the oil

samples showed a newtonian characteristic that viscosity didn't change under different shear rates. According to Fig. 3a, MDA had a negligibly small effect on viscosity of continuous phase within the studied concentrations. Thus, the mechanism of enhancing the emulsion stability using MDA is different from thickeners. Thickeners increase the viscosity of continuous phase to overcome gravity stress to obtain better emulsion stability.

As shown in Fig. 3b, the emulsions containing various amounts of MDA showed a non-newtonian shear-thinning behavior that the viscosity decreased with increasing shear rates. When the shear rates were relatively high (150–200 s<sup>-1</sup>),

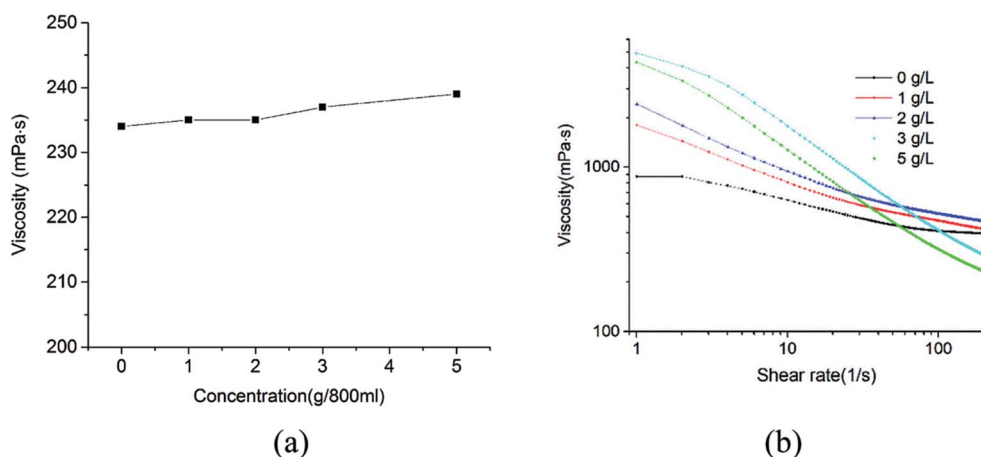


Fig. 3 The influence of MDA concentration on viscosity of oil (a) and emulsion (b).





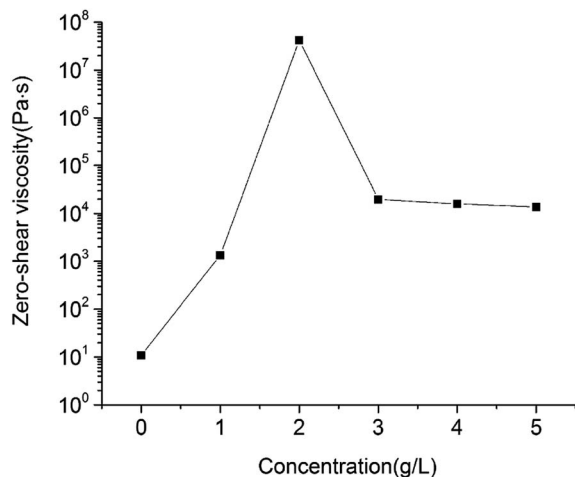


Fig. 4 Zero-shear viscosities of the emulsions with various concentrations of MDA.

the viscosities of the emulsions containing 0–2 g L<sup>-1</sup> MDA did not show remarkable changes. During normal drilling process, shear rates in a drill pipe generally encompass the range from 511 s<sup>-1</sup> to 1022 s<sup>-1</sup>, so the effect of MDA (0–2 g L<sup>-1</sup>) on the high-shear-rate viscosity of drilling fluid is slight. When circulation of drilling fluid is ceased (shear rates approach zero), the viscosities for all the samples are very large, which is beneficial to suspension of water droplets.

(2) **Approximated zero-shear viscosity.** Zero-shear viscosity was approximated as the complex viscosity at 0.001 Hz. The results are shown in Fig. 4. According to Fig. 4,  $\eta_0$  increased remarkably from 10.89 Pa s to  $4.19 \times 10^7$  Pa s when concentration of MDA increased from 0 to 2 g L<sup>-1</sup>. As MDA concentration continued to increase to 5 g L<sup>-1</sup>,  $\eta_0$  decreased to  $1.37 \times 10^4$  Pa s. The trend of  $\eta_0$  is in accordance with that of visual observation.

It is well documented that zero-shear viscosity has a good correlation with the rate of sedimentation.<sup>28,29</sup> In our study, maximum  $\eta_0$  was remarkably more than 6 orders of magnitude larger than  $\eta_0$  of pure emulsion, which is the reason why stability of pure emulsion + 2 g L<sup>-1</sup> was substantially increased compared with that of pure emulsion. The decrease of  $\eta_0$  when MDA concentration exceeded 2 g L<sup>-1</sup> could be due to the stronger flocculation and less emulsion stability.

(3) **Creep and recovery tests.** Unwanted sedimentary effects relate to processes at low shear rates which have gravity as the driving force. In order to simulate the processes, creep and recovery tests were performed at a constant stress of 0.1 Pa. The results of creep and recovery tests are shown in Fig. 5. Two parameters as shown in Table 2, peak strain (strain at the end of creep phase) and recovery ratio (recovery strain at the end of recovery phase divided by the corresponding peak strain), were used to analyze the experimental results.

During the creep phase, the strain increased over time for all the emulsions. The peak strain decreased remarkably from 0.88757 to 0.02083 as the concentration of MDA increased from 0 to 5 g L<sup>-1</sup>. In comparison with pure emulsion, other

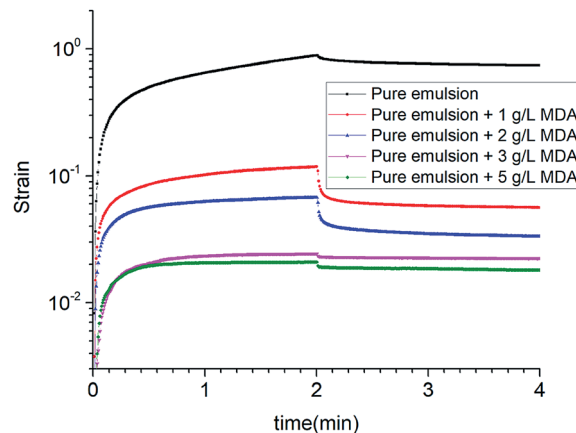


Fig. 5 Creep and recovery curves of emulsions with various concentrations of MDA. A constant shear stress of 0.1 Pa was applied for the first two minutes and then removed.

Table 2 Parameters of the creep and recovery curves in Fig. 5

Emulsion samples	Peak strain	Recovery ratio
Pure emulsion	0.88757	0.1621
Pure emulsion + 1 g L <sup>-1</sup> MDA	0.11748	0.5209
Pure emulsion + 2 g L <sup>-1</sup> MDA	0.06761	0.5082
Pure emulsion + 3 g L <sup>-1</sup> MDA	0.02431	0.0802
Pure emulsion + 5 g L <sup>-1</sup> MDA	0.02083	0.1330

emulsions showed great decreases in peak strain. A lower peak strain indicates a less flexible structure that is harder to deform under the controlled stress. Thus, as the MDA concentration increased, the emulsion samples had a more “solid like” structure.

During recovery phase, the elimination of the stress resulted in decrease in strain. Pure emulsion + 1 g L<sup>-1</sup> MDA and pure emulsion + 2 g L<sup>-1</sup> MDA had the similar recovery ratios which were much larger than those of the other emulsions. Recovery ratio is an indicator to elastic response. A larger recovery ratio indicates a greater elasticity. But compared with pure emulsion + 1 g L<sup>-1</sup> MDA, pure emulsion + 2 g L<sup>-1</sup> MDA had a much lower peak strain. So taking into account both peak strain and recovery ratio, it is possible to determine that pure emulsion + 2 g L<sup>-1</sup> MDA was best structured. The results of creep and recovery tests are in accordance with the results of zero-shear viscosity measurements.

### 3.4 Possible emulsion stabilizing theory by MDA

As can be seen in Scheme 1, MDA has two hydrophobic tails and two hydrophilic heads. The amphiphilic nature of MDA makes it to adsorb on oil–water interface, which might increase the intermolecular hydrogen bonding of water droplets. Consequently, a three dimensional networks might form because of the weak attraction force.

On one hand, MDA as a structuring agent can establish the network structures in emulsions, which can provide a resistant force to slow down the rate of sedimentation and aggregation,



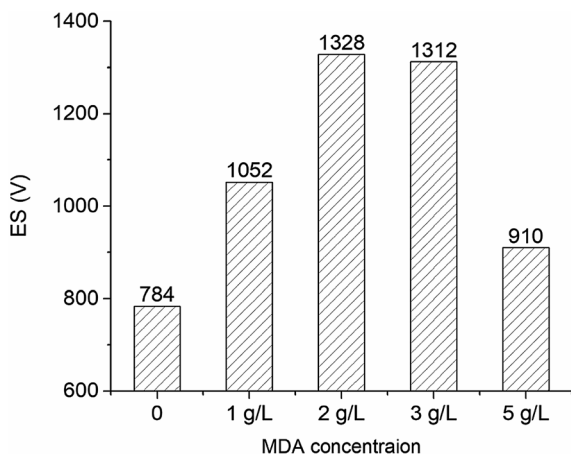


Fig. 6 ES values of OBMs containing 0 g L<sup>-1</sup>, 1 g L<sup>-1</sup>, 2 g L<sup>-1</sup>, 3 g L<sup>-1</sup>, 5 g L<sup>-1</sup> MDA.

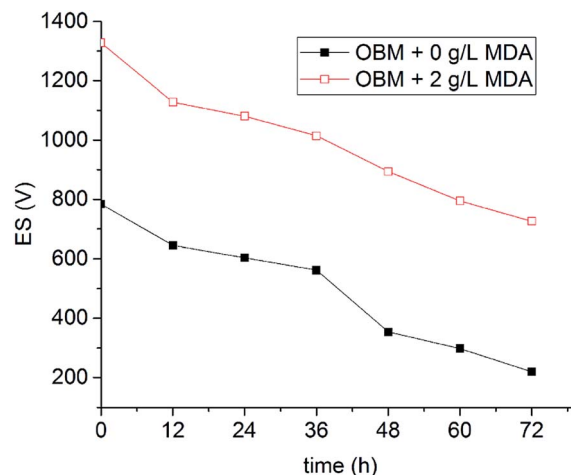


Fig. 7 ES vs. time plots of OBMs containing no MDA and 2 g L<sup>-1</sup> MDA aged at 150 °C.

benefiting the emulsion stability. But on the other hand, the negative effect is that the attraction force strengthened by MDA can cause flocculation. As MDA concentration increases in the emulsions, flocculation is becoming increasingly obvious.

At an appropriate concentration (2 g L<sup>-1</sup>), the positive effect of MDA on emulsion stability is significantly more obvious than the negative effect caused by flocculation. At this point, the presence of MDA in w/o emulsions can greatly enhance the emulsion stability. In fact, the formation of the network structures in emulsions is due to weak flocculation.<sup>20,30</sup> In other words, the mechanism of MDA to enhance emulsion stability can be concluded that the network structures in the emulsion formed by weak flocculation is beneficial to emulsion stability, which is in accordance with other work.<sup>20,30</sup>

### 3.5 Application of MDA in OBMs

(1) **Electricity stability.** ES is one of vital properties for OBMs, which shows the voltage of the current to flow in the drilling fluids. ES is related to the stability of an OBM. Generally, the higher the values, the more stable the emulsion.<sup>31</sup> Fig. 6 shows the ES values of newly formulated OBMs with different amounts of MDA. According to Fig. 6, the ES values of OBMs containing MDA were all larger in comparison with that of the control group. Especially when MDA concentrations were 2 g L<sup>-1</sup> and 3 g L<sup>-1</sup>, the ES values were both larger than 1300 V and nearly twice as large as that of the control group, indicating fairly stable emulsions.

The OBM with 2 g L<sup>-1</sup> MDA having the highest ES value and the OBM without MDA were aged statically at 150 °C for three days. The ES values of OBMs containing no MDA and 2 g L<sup>-1</sup> MDA changing with time were shown in Fig. 7. It is apparent that the ES values of the two OBMs decreased with time because the OBMs are thermodynamically unstable. After three-day aging, the ES value of the OBM containing 2 g L<sup>-1</sup> MDA was more than 700 V which was far larger than that of the control group (220 V). This made it possible to keep a longer-term emulsion stability of OBMs.

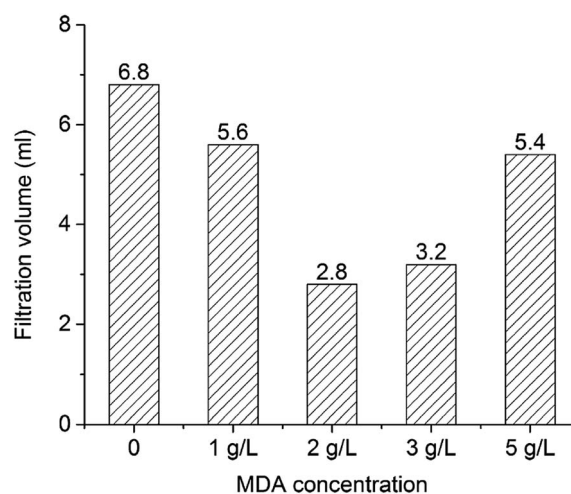


Fig. 8 Effect of MDA concentration on filtration volume.

(2) **Influence of MDA on fluid loss volume.** Filtration property is of great importance to drilling fluids since fluid loss into surrounding permeable formation should be minimized. The effect of MDA on the HTHP filtration volume was explored. As shown in Fig. 8, the OBMs with MDA at studied concentrations had lower filtration volumes compared with the control group. During drilling operations, whether drilling fluids in the well-bore are circulated or not, the shear rate of drilling fluids is minimal near the wall of a well. So near the well wall, the OBM with 2 g L<sup>-1</sup> MDA can be better structured and have highest viscosity, reducing filtration volume.

## 4. Conclusions

This paper presents an alternative method to enhance w/o emulsion stability using MDA which was synthesized by reacting OAD with DEA. Bottle tests showed that a substantial



improvement in emulsion stability was observed after addition of MDA, and the optimum concentration of MDA was  $2 \text{ g L}^{-1}$ .

A honeycomb structure was observed in continuous phase of pure emulsion +  $2 \text{ g L}^{-1}$  MDA from cryo-SEM, which indicates a “gel-like” structure. According to rheological experiments of pure emulsion +  $2 \text{ g L}^{-1}$  MDA, a remarkably larger zero-shear viscosity, a solid-like behavior and a greater elasticity were observed in comparison with pure emulsion. The “gel-like” structure and improvements in rheological properties are both beneficial to emulsion stability.

MDA can largely enhance electrical stability and reduce HTHP filtration volume of OBMs.

The possible mechanism of emulsion stabilization was proposed as follows: theoretically, the amphiphilic nature of MDA makes it possible to adsorb on oil–water interface, which might enhance the interaction force of water droplets and establish three dimensional network structures in emulsions. The network structures in the emulsion is beneficial to emulsion stability.

## Conflicts of interest

None.

## Acknowledgements

The authors are thankful to China Postdoctoral Science Foundation (2018M630812), the Fundamental Research Funds for the Central Universities (No. 18CX02171A), Scientific Research Foundation for the Introduction of Talents (YJ20170014) and Joint Funds of the National Natural Science Foundation of China (U1762212).

## References

- 1 R. Caenn, H. C. H. Darley and G. R. Gray, *Composition and properties of drilling and completion fluids*, Gulf Professional Publishing, 2011.
- 2 C. K. Reiffers-Magnani, J. L. Cuq and H. J. Watzke, *Food Hydrocolloids*, 2000, **14**, 521–530.
- 3 M. Moradi, V. Alvarado and S. Huzurbazar, *Energy Fuels*, 2010, **25**, 260–268.
- 4 L. G. Torres, R. Iturbe, M. J. Snowden, B. Z. Chowdhry and S. A. Leharne, *Colloids Surf., A*, 2007, **302**, 439–448.
- 5 E. Dickinson, *Food Hydrocolloids*, 2009, **23**, 1473–1482.
- 6 I. Heertje, *Food Struct.*, 2013, **1**, 3–23.
- 7 M. Kowalska, M. Ziomek and A. Żbikowska, *Int. J. Cosmet. Sci.*, 2015, **37**, 408–416.
- 8 A. Perino, C. Noik and C. Dalmazzone, *Energy Fuels*, 2013, **27**, 2399–2412.
- 9 N. Nihant, C. Schugens, C. Grandfils, R. Jérôme and P. Teyssié, *Pharm. Res.*, 1994, **11**, 1479–1484.
- 10 D. E. Tambe and M. M. Sharma, *Adv. Colloid Interface Sci.*, 1994, **52**, 1–63.
- 11 E. Kettler, C. B. Müller, R. Klemp, M. Hloucha, T. Döring, W. Von Rybinski and W. Richtering, in *Progress in Colloid and Polymer Science*, Springer, 2007, vol. 134, pp. 90–100.
- 12 D. Krajisnik and J. Milić, *Drug Dev. Ind. Pharm.*, 2003, **29**, 701–711.
- 13 Y. Tan, K. Xu, C. Liu, Y. Li, C. Lu and P. Wang, *Carbohydr. Polym.*, 2012, **88**, 1358–1363.
- 14 Y. Jiang, X. Liu, Y. Chen, L. Zhou, Y. He, L. Ma and J. Gao, *Bioresour. Technol.*, 2014, **153**, 278–283.
- 15 S. Lam, K. P. Velikov and O. D. Velev, *Curr. Opin. Colloid Interface Sci.*, 2014, **19**, 490–500.
- 16 F. Marchal, A. Roudot, N. Pantoustier, P. Perrin, J. Daillant and P. Guenoun, *J. Phys. Chem. B*, 2007, **111**, 13151–13155.
- 17 Unilever BCS US Inc, *US Pat.* 8940355B2, 2011.
- 18 E. Dickinson, *J. Colloid Interface Sci.*, 2015, **449**, 38–45.
- 19 A. Bot, Y. S. J. Veldhuizen, R. den Adel and E. C. Roijers, *Food Hydrocolloids*, 2009, **23**, 1184–1189.
- 20 E. Dickinson and K. Pawlowsky, *J. Agric. Food Chem.*, 1997, **45**, 3799–3806.
- 21 U. Klinkesorn, P. Sophanodora, P. Chinachoti and D. J. McClements, *Food Res. Int.*, 2004, **37**, 851–859.
- 22 J. De Visscher, H. Soenen, a. Vanelstraete and P. Redelius, in *Proceedings of the 3rd Eurasphalt and Eurobitume Congress Held Vienna*, 2004, vol. 2, pp. 1501–13.
- 23 A. P. Institute, *API RP 13B-2 Recommended Practice for Field Testing of Oil-based Drilling Fluids*, 5th edn, 2014.
- 24 J. O'Sullivan, M. Arellano, R. Pichot and I. Norton, *Food Hydrocolloids*, 2014, **42**, 386–396.
- 25 M. Reger, T. Sekine, T. Okamoto, K. Watanabe and H. Hoffmann, *Soft Matter*, 2011, **7**, 11021.
- 26 A. Beri, J. E. Norton and I. T. Norton, *Int. J. Cosmet. Sci.*, 2013, **35**, 613–621.
- 27 J. Y. Kim, J. Y. Song, E. J. Lee and S. K. Park, *Colloid Polym. Sci.*, 2003, **281**, 614–623.
- 28 M. V. Tzoumaki, T. Moschakis, V. Kiosseoglou and C. G. Biliaderis, *Food Hydrocolloids*, 2011, **25**, 1521–1529.
- 29 M. C. García, M. C. Alfaro and J. Muñoz, *Colloids Surf., B*, 2015, **135**, 465–471.
- 30 T. Tadros, *Adv. Colloid Interface Sci.*, 2004, **108**, 227–258.
- 31 F. B. Growcock, C. F. Ellis, D. D. Schmidt and J. J. Azar, *SPE Drill. Completion*, 1994, **9**, 39–46.

