




Cite this: *RSC Adv.*, 2018, 8, 28209

Three-dimensional flower-like NiCo₂O₄/CNT for efficient catalysis of the oxygen evolution reaction †

Zhaoling Ma,^a Hao Fu,^a Cibing Gu,^a Youguo Huang,^{*a} Sijiang Hu,^b Qingyu Li^a and Hongqiang Wang  ^{*ab}

The oxygen evolution reaction (OER) is an important reaction especially in water splitting and metal–air batteries. Highly efficient non-noble metal based electrocatalysts are urgently required to be developed and to replace the commercial Ru/Ir based oxide. Herein, we report the three-dimensional hierarchical NiCo₂O₄/CNT-150 composite with high activity for the OER that was synthesized *via* a hydrothermal reaction and subsequent annealing. Compared with CNTs, commercial RuO₂ catalysts, NiCo₂O₄/CNT, NiCo₂O₄/CNT-250, and NiCo₂O₄/CNT-150 exhibit enhanced electrocatalytic performance with a lower onset overpotential of 300 mV and the corresponding Tafel slope of 129 mV per decade. The flower-like NiCo₂O₄/CNT-150 shows an excellent catalysis performance with higher current density than the commercial RuO₂ catalyst. Moreover, the NiCo₂O₄/CNT-150 demonstrates the excellent long-term durability in 0.1 mol L⁻¹ KOH for the OER. The significant catalytic performances are ascribed to the excellent conductivity of CNTs and the high specific surface area of the three dimensional flower-like NiCo₂O₄.

Received 2nd July 2018
 Accepted 30th July 2018

DOI: 10.1039/c8ra05639k

rsc.li/rsc-advances

Introduction

With the increasing consumption of fossil fuels and the continuing increase in environmental pollution, the development of alternative energy conversion or storage devices, such as metal–air batteries and regenerative fuel cells *etc.*, with high power and large energy densities is of particular significance.^{1,2} The oxygen reduction reaction (ORR) and oxygen evolution reaction (OER) accompany the charge and discharge process in the cathode of metal–air batteries. In particular, production of hydrogen and oxygen from alkaline water splitting is a potential route to produce clean energy resources. However, the OERs are involved in a wide range of electrochemical processes and their kinetics are sluggish even after being facilitated by noble-metal catalysts.³

It is noted that the OER process is complex, with hydroxyl ions being consumed to produce oxygen and water molecules (4OH⁻ ↔ O₂ + 2H₂O + 4e⁻), requiring a large overpotential (η) under alkaline conditions.⁴ Up to now, the efficiency of water splitting is not satisfactory. Hence, numerous research efforts have been paid to develop more efficient and stable electrocatalysts to reduce the required overpotential for OER.

Noble-metal based catalysts such as IrO₂ and RuO₂ have emerged as the most promising catalyst candidates for the OER, but their widespread applications in real industry are limited due to the scarcity and high cost.⁵ Therefore, in order to realize overall water splitting, the electrocatalysts composing of inexpensive and more abundant materials such as manganese (Mn),⁶ and cobalt (Co)⁷ are highly desired. To date, among the various electrocatalysts for OER, cobalt-based electrocatalysts are considered to be one of the best performing materials, and the spinel oxides (Co₃O₄) have been reported to have high electrochemical performance with low overpotential and excellent chemical stability.⁸ To further improve the electrochemical activity of Co₃O₄ electrocatalysts, researchers have doped Ni atoms into the spinel structure to form NiCo₂O₄, which increased electrical conductivity of the metal oxide (at least 2 orders of magnitude higher than that of Co₃O₄) and enhanced the catalytic activity for the OER.⁹ In addition, spinel NiCo₂O₄ is easy to form a mesoporous superstructure, which has been studied extensively to investigate the morphological effect on its electrochemical activity.^{10,11} More importantly, because of the different valence states in the spinel structure where Ni occupies the octahedral sites and Co occupies both the octahedral and the tetrahedral sites, both redox couples (Co³⁺/Co²⁺ and Ni³⁺/Ni²⁺) ensure a notable electrocatalytic activity.¹² Inspired by these good features, several strategies, including sol-gel,¹³ hydrothermal,¹⁴ and potentiostatic deposition,¹⁵ have been applied to synthesize various NiCo₂O₄ nanostructures, such as nanoparticle, nanosheet, nanotube, and NiCo₂O₄-based composite

^aGuangxi Key Laboratory of Low Carbon Energy Materials, School of Chemistry and Pharmaceutical Sciences, Guangxi Normal University, Guilin, 541004, China. E-mail: whq74@126.com; whq74@mailbox.gxnu.edu.cn; Fax: +86-0773-5858562

^bHubei Key Laboratory for Processing and Application of Catalytic Materials, College of Chemical Engineering, Huanggang Normal University, Huanggang, 438000, China

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c8ra05639k



etc. However, in most cases, NiCo₂O₄ still encounters intrinsically low electrical conductivity.^{16,17} The most efficient strategies is to choose carbon material as the conductive support for NiCo₂O₄ owing to its excellent conductivity.

Carbon nanotube (CNT), as an important class of nano-carbon materials, possesses remarkable electronic conductivity,¹⁸ chemical stability and high specific surface. This enables them to be the good substrate for preparing hybrid and composite materials. Especially, many multi-walled CNT based composites have been exploited as advanced electrocatalysts.^{19,20} For OER application, various transition metal or compounds, such as the inexpensive metal compounds,^{21,22} and MnCo₂O₄,²³ Ni(OH)₂,²⁴ Co(OH)₂,²⁵ have been grown on CNTs in attempts to improve the OER performances. This has been confirmed to be a feasible and effective method. The present works on NiCo₂O₄/CNT composite with 3D hierarchical structure could arouse an attractive potential OER electrocatalyst.

In this work, we construct a three-dimensional hierarchical NiCo₂O₄/CNT network electrode for OER by a potentiostatic electrodeposition technique followed by a calcination–annealing process in the air. In this typical binder-free electrode, the CNTs vertically grown on Ti foil are used as “rod” to support the 3D flower-like NiCo₂O₄ nanosheet, which would not only remarkably improve outstanding conductivity but also notably increase the specific surface area. The excellent performance is credited to the specific micro-structure and the remarkable synergistic effects of the components. In view of the highly increased catalytic function in the OER, the NiCo₂O₄/CNT composite proves to be a very promising candidate as an electrocatalyst for water splitting.

Materials and methods

Growth of CNTs on Ti foil substrate

All reagents were of analytical grade and were directly used without further purification. All the solutions were prepared with deionized water. Ti foil (purity 99.999%, ~0.1 mm in thickness, Baoji Titanium products Co. Ltd.) were used as substrate for growth of self-propped CNTs layer. First, Ti foil was cleaned by ultrasonication respectively with absolute ethanol and deionized water. The Ti foil was further decontaminated in the solution containing 15 g L⁻¹ NaOH, 20 g L⁻¹ Na₂CO₃ and 5 mmol L⁻¹ OP emulsifier at 60 °C for 15 min. Secondly, it is subsequently immersed in a mixed acid solution containing hydrochloric acid (25% v/v) and hydrofluoric acid (5% v/v) for 2 min. Thirdly, Ni catalyst particles were deposited on Ti foil by electroless deposition in the solution containing 500 mmol L⁻¹ glycol, 20 g L⁻¹ nickel chloride and 15 g L⁻¹ ammonium fluoride at 60 °C for 5 min. CNTs were grown on Ti foil by chemical vapor deposition (CVD) process. C₂H₂ is used as carbon source under H₂ atmosphere at the pressure of 20 kPa. The growth time of CNTs was controlled for 40 min.

Preparation of NiCo₂O₄/CNT electrode

Prior to preparation, the CNT/Ti substrate (1.5 cm × 2 cm) was immersed in absolute ethanol for 10 min. The well-treated CNT/

Ti substrate was transferred into the solution containing 1 mmol Ni(NO₃)₂·6H₂O, 2 mmol Co(NO₃)₂·6H₂O, 2 mmol NH₄F, and 6 mmol CO(NH₂)₂. The solvent is comprised with 40 mL ethanol and 40 mL H₂O. The solution including CNT/Ti substrate was transferred into 100 mL Teflon-lined stainless autoclave, followed by hydrothermal synthesis at 100 °C for 10 h. The obtained precursor was washed several times by water and dried at 50 °C (labeled as NiCo₂O₄/CNT). The other samples were synthesized as similar method except for the different subsequent annealing temperature at 150 °C and 250 °C. The dried precursor was calcined for 2 h with a ramp rate of 5 °C min⁻¹. The resulted samples were denoted as NiCo₂O₄/CNT-150 and NiCo₂O₄/CNT-250. The CNTs peeled from CNT/Ti substrate was denoted as CNT. The quality of NiCo₂O₄ in NiCo₂O₄/CNT-150 and NiCo₂O₄/CNT-250 was respectively 0.55 mg (~0.19 mg cm⁻²) and 0.28 mg (~0.10 mg cm⁻²).

Materials characterization

The surface morphology and structure of the samples were analyzed by Field Emission Scanning Electron Microscopy (FE-SEM, FEI Quanta 200 FEG, Holland) and transmission electron microscope (TEM, FEI TECNAI D2 12, Holland). Electron diffraction spectroscopy (EDS) was applied to determine the elements of powders together with SEM in large field of view. X-ray photoelectron spectroscopy (XPS) was carried out on a Physical Electronics 5400 ESCA. The crystalline structure of the samples were analyzed by X-ray diffraction (XRD) on a D/Max-2500V/PC powder diffractometer equipped with Cu K α radiation ($k = 0.15406$ nm) (Rigaku, Tokyo, Japan). The Fourier transform infrared spectra were recorded using a Nicolet 6700 spectrometer (Thermo Scientific, USA). All physical characterization of CNT, NiCo₂O₄/CNT-150 and NiCo₂O₄/CNT-250 were carried out on Ti substrates.

Electrochemical measurements

All the electrochemical measurements were conducted in a three-electrode cell on a CHI 660e electrochemical workstation (Chenhua, Shanghai). All materials were peeled from Ti substrate and executed electrochemical measurements. 0.1 mol L⁻¹ KOH was used as the electrolyte. A saturated calomel electrode (SCE) and Pt sheet were respectively used as reference electrode and counter electrode. NiCo₂O₄/CNT, NiCo₂O₄/CNT-150, NiCo₂O₄/CNT-250 and NiCo₂O₄/CNT-150, NiCo₂O₄/CNT-250 and commercial RuO₂ were made into 2 mg mL⁻¹ catalyst inks, which contain 995 μ L water, 995 μ L ethanol as well as 10 μ L of 5 wt% Nafion solution. The 20 μ L catalyst inks were loaded on the glass carbon electrodes with 4 mm diameter as the working electrodes. Linear scan voltammetry (LSV) was conducted from 0 to 0.8 V with a scan rate of 5 mV s⁻¹ for the polarization curves and Tafel slopes. The stability was analyzed by chronoamperometry which was carried out at 1.63 V *versus* RHE. Electrochemical impedance spectroscopy (EIS) was performed in the frequency range of 0.1 Hz to 100 kHz at 0.63 V *versus* SCE. All electrochemical tests were collected under O₂ saturation to ensure the O₂/H₂O equilibrium at 1.23 V. The reported potentials in this work are *versus* reversible hydrogen



electrode (RHE), which is based on the Nernst equation ($E_{\text{RHE}} = E_{\text{SCE}} + 0.059 \text{ pH} + 0.2412$). The overpotential (η) was calculated according to the formula: $\eta \text{ (V)} = E_{\text{RHE}} - 1.23$. All the potential values were compensated by eliminating iR drop from ohmic resistance of the electrolytes.

Results and discussion

The synthesis process of the 3D flower-like NiCo_2O_4 on CNT is shown in Scheme 1. Ti foil was worked as the substrate. CNTs are vertically grown on the Ti foil by CVD process with the Ni-based catalysts, which could possess a high degree of graphitization. Subsequently, the 3D flower-like NiCo_2O_4 are prepared *via* hydrothermal reaction and subsequent calcination. In $\text{NiCo}_2\text{O}_4/\text{CNT}$ composite, vertically grown CNTs offer excellent electrical conductivity for rapid electron transport. Furthermore, 3D flower-like NiCo_2O_4 could show the large surface area, which could not only provide more active area for OER but also form the whole conductive networks.

The synthesized CNT, $\text{NiCo}_2\text{O}_4/\text{CNT}$, $\text{NiCo}_2\text{O}_4/\text{CNT-150}$ and $\text{NiCo}_2\text{O}_4/\text{CNT-250}$ composites were characterized by FTIR spectroscopy, and the results are shown in Fig. S1.† For CNT, the bands at $\sim 3300 \text{ cm}^{-1}$, $\sim 2955 \text{ cm}^{-1}$, $\sim 1620 \text{ cm}^{-1}$, $\sim 1260 \text{ cm}^{-1}$ and 1090 cm^{-1} represent respectively hydroxyl ($-\text{OH}$), carbon-hydrogen bonds ($-\text{CH}_3$), carbon-carbon bonds ($-\text{C}=\text{C}-$), carboxyl ($-\text{COOH}$) and carbon-oxygen bonds ($\text{C}-\text{O}$),^{26–28} whose appearance suggests the slight oxidation existed on the as-grown CNTs and it could provide the sites for the subsequent growth of NiCo_2O_4 . The bands at $\sim 810 \text{ cm}^{-1}$ and 660 cm^{-1} are attributed to the carbon-hydrogen bonds ($\text{C}-\text{H}$) of aromatic nucleus in carbon skeleton.^{28,29} In the case of the $\text{NiCo}_2\text{O}_4/\text{CNT}$, $\text{NiCo}_2\text{O}_4/\text{CNT-150}$ and $\text{NiCo}_2\text{O}_4/\text{CNT-250}$, all the characteristic peaks are same as the counterpart of CNT except of the shifts towards lower wavenumber, which suggests the successful hybrid connection of CNT and NiCo_2O_4 .

XRD measurements were used to identify the chemical composition and phase of the as-prepared electrocatalysts. Fig. 1 shows the XRD patterns obtained from CNT, $\text{NiCo}_2\text{O}_4/\text{CNT}$, $\text{NiCo}_2\text{O}_4/\text{CNT-150}$ and $\text{NiCo}_2\text{O}_4/\text{CNT-250}$. CNT shows a typical diffraction peak at $\sim 27.3^\circ$ corresponding to the (331) plane of CNT. However, five dominant diffraction peaks at $\sim 21^\circ$, $\sim 38.4^\circ$, $\sim 40^\circ$, $\sim 53^\circ$, $\sim 70.5^\circ$ and 77° corresponding to the (120), (002), (101), (102), (103) and (201) plane reflections of the Ti foil substrate. Notably, the peaks at $\sim 18.8^\circ$ and $\sim 59.1^\circ$ of $\text{NiCo}_2\text{O}_4/\text{CNT}$ and $\text{NiCo}_2\text{O}_4/\text{CNT-150}$ are respectively belonged to (111) and (511) planes of the cubic spinel NiCo_2O_4 (JCPDS no.

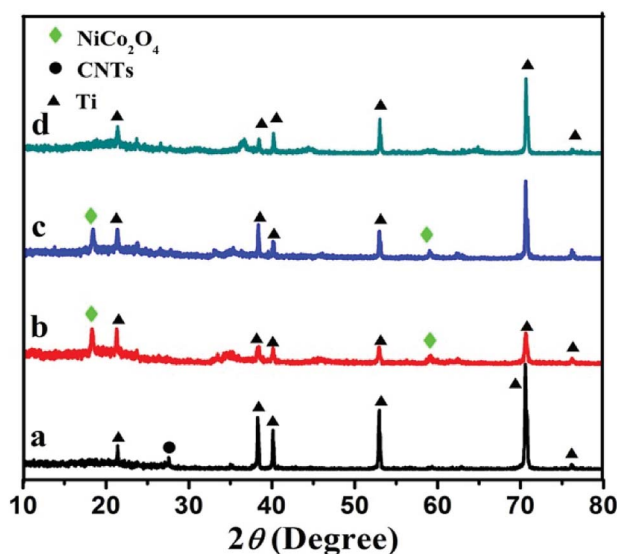
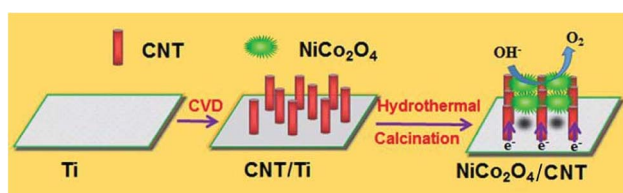


Fig. 1 X-ray powder diffraction patterns obtained for CNT (a), $\text{NiCo}_2\text{O}_4/\text{CNT}$ (b), $\text{NiCo}_2\text{O}_4/\text{CNT-150}$ (c), $\text{NiCo}_2\text{O}_4/\text{CNT-250}$ (d).

20-0781),³⁰ where nickel atoms located at octahedral sites and cobalt atoms occupied both octahedral and tetrahedral sites.³¹ Nevertheless, the two peaks are not observed in $\text{NiCo}_2\text{O}_4/\text{CNT-250}$, which means that the higher calcination temperature enabled to deteriorate the crystalline structure.

The morphology of the resulted composites was shown in Fig. 2. Fig. 2a shows the CNTs with the diameter of about 50 nm fabricated directly on Ti substrate surface by CVD. These CNTs are grown vertically and possess the diameter about 30 to 70 nm. Fig. 2b show the NiCo_2O_4 before annealing has been grown uniformly on the CNTs. The connection of nanosheets constitute into the flowerlike NiCo_2O_4 . Undergoing the annealing at 150°C , the flowerlike NiCo_2O_4 of $\text{NiCo}_2\text{O}_4/\text{CNT-150}$



Scheme 1 The synthesis processes of 3D flower-like $\text{NiCo}_2\text{O}_4/\text{CNT}$ on Ti foil.

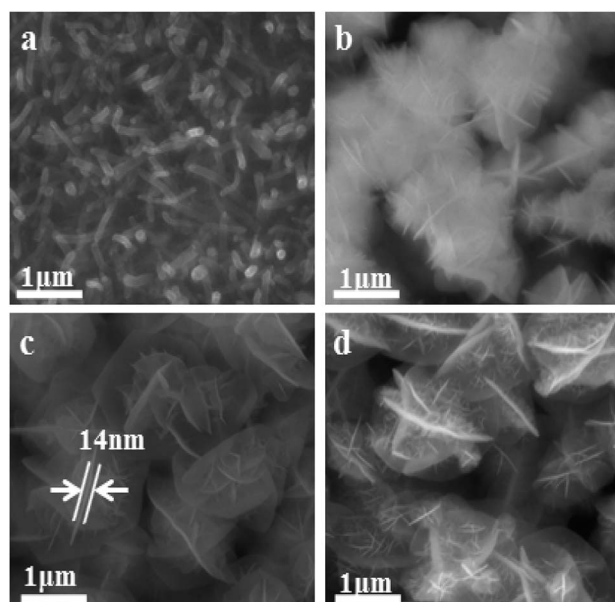


Fig. 2 SEM images of (a) CNT and (b) $\text{NiCo}_2\text{O}_4/\text{CNT}$; (c) $\text{NiCo}_2\text{O}_4/\text{CNT-150}$ and (d) $\text{NiCo}_2\text{O}_4/\text{CNT-250}$.



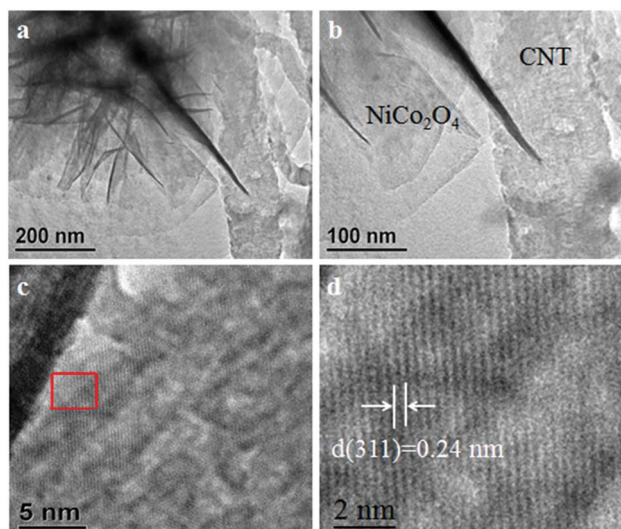


Fig. 3 (a and b) TEM images and (c and d) HRTEM images of NiCo₂O₄/CNT-150.

150 in Fig. 2c has the diameter of about 1.5 μm and exhibits clear nanosheets and surface wrinkles. Moreover, NiCo₂O₄ flower closely connected with each other, which is benefit to rapid transport electrons. The NiCo₂O₄ nanosheets show the thickness of approximately 14 nm. The surface wrinkles can afford more active sites for OER. When the annealing temperature was increased to 250 $^{\circ}\text{C}$, NiCo₂O₄ nanosheets changed into a number of nanothorns.

In order to determine the corresponding elemental mappings of NiCo₂O₄/CNT-150, elemental mapping by EDS spectrum was used to analyze elements distribution. The EDS analysis in Fig. S2† further corroborates its uniform deposition. As seen, the elements C, O, Co and Ni of NiCo₂O₄/CNT-150 are evenly distributed in the selected area, indicating the uniform NiCo₂O₄ flowers deposition on the surface of CNT.

These CNTs and 3D flower-like NiCo₂O₄ were further characterized by transmission electron microscopy (TEM) as shown in Fig. 3. It was observed that flower-like nanosheets of NiCo₂O₄/CNT-150 connect tightly to the CNTs substrate (Fig. 3a and b). This confirms the close contact between the 3D NiCo₂O₄

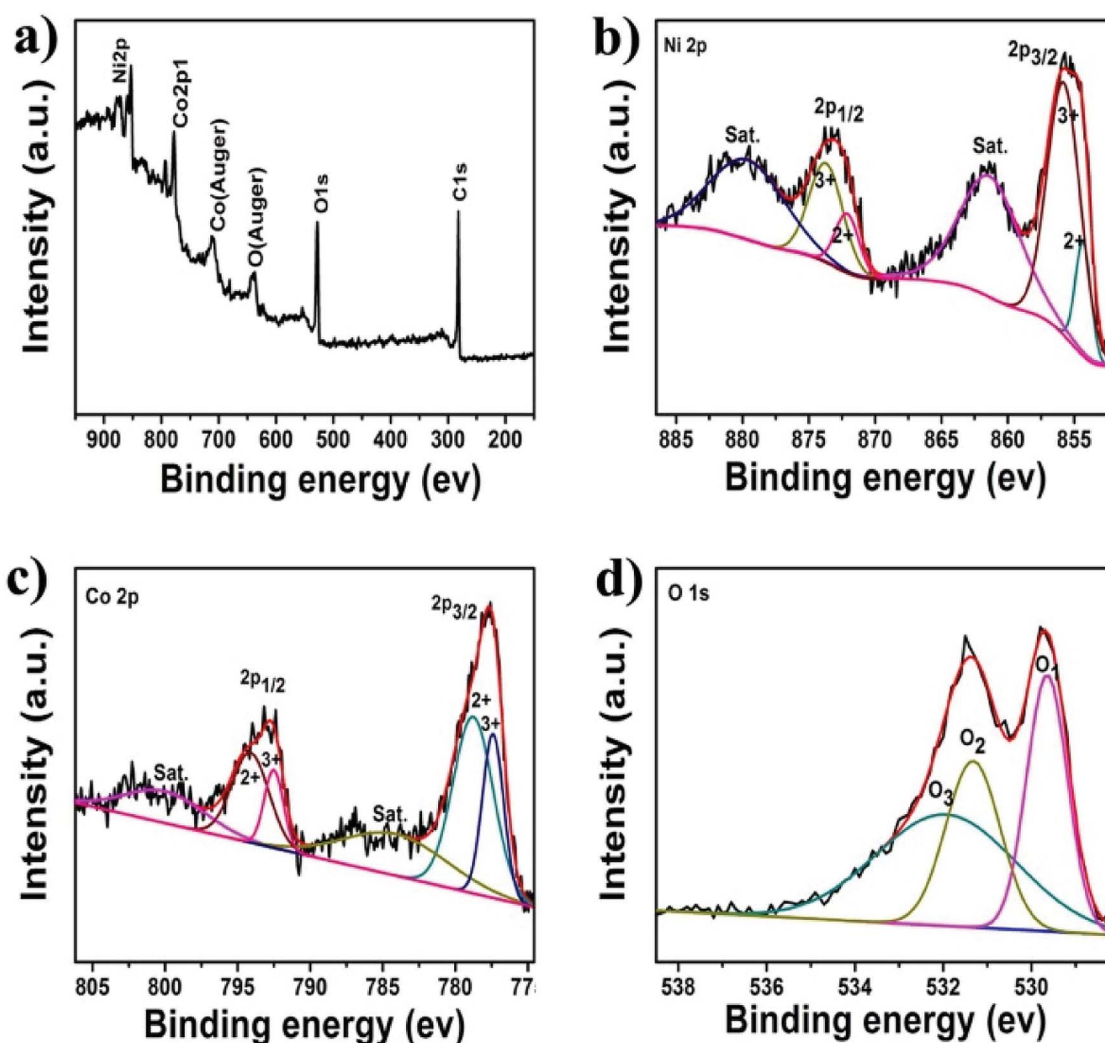


Fig. 4 XPS spectra of (a) survey spectrum, (b) Ni 2p, (c) Co 2p and (d) O 1s for NiCo₂O₄/CNT-150.



nanosheets and CNTs scaffold. The blurry lattice fringes in Fig. 3c testify the low crystallinity of NiCo₂O₄/CNT-150. The interplanar distance of 0.24 nm is corresponded to the (311) plane of NiCo₂O₄ (Fig. 3d). Clearly, the NiCo₂O₄ crystals in NiCo₂O₄/CNT-150 shows discontinuous lattice fringes. That means that abundant of crystal defects exists on the surface of NiCo₂O₄, especially for the (311) plane. It could provide a quantity of crystal defects as the new type of catalytic sites for OER.

To investigate the chemical state of the as-prepared electrocatalysts, XPS technique was used to analysis NiCo₂O₄/CNT-150. As shown in Fig. 4a, the survey spectrum indicates the presence of Ni, Co, O and C element without other impurity elements. The Ni 2p spectrum in Fig. 4b can be fitted as two shakeup satellites (indicated as "Sat.") and two spin-orbit doublets characteristic of Ni²⁺ and Ni³⁺ with Gaussian fitting.³² The peaks at 855.9 eV and 873.4 eV are assigned to Ni³⁺, while the peaks at 854.2 eV and 871.6 eV are belonged to Ni²⁺.^{33,34} It is implied that two nickel ions exist in NiCo₂O₄. The satellite peaks located at 861.5 eV and 879.3 eV are two overlapping type peaks.³⁵ Similarly, the Co 2p spectrum, as shown in Fig. 4c, also can be fitted into two satellites as well as two spin-orbit peaks, characteristics of Co²⁺ and Co³⁺.¹⁶ The O 1s spectrum in Fig. 4d, shows

three chemical states of oxygen element, respectively labeled as O₁, O₂ and O₃. The O₁ peak at 529.5 eV corresponds to metal-oxygen bonds.³⁶ The O₂ peak at 531.3 eV is assigned to OH⁻ groups, indicating that carbon nanotubes is partly hydroxylated as a result of surface hydroxides.³⁷ The O₃ peak located at 532.2 eV is identified to the physical adsorbed water.³⁸ Obviously, the XPS data illustrates that the electron couples of Ni³⁺/Ni²⁺ and Co³⁺/Co²⁺ are concomitant in the as-prepared NiCo₂O₄ sample.

The electrochemical performance of the NiCo₂O₄/CNT composite was investigated in 0.1 M KOH solution, as shown in Fig. 5. Fig. 5a shows the polarization curves of CNT, NiCo₂O₄/CNT, NiCo₂O₄/CNT-150, NiCo₂O₄/CNT-250 and commercial RuO₂ electrodes. Obviously, NiCo₂O₄/CNT-150 electrode exhibits the most superior OER performance than other electrodes. The onset overpotential of NiCo₂O₄/CNT-150 was 300 mV, which is lower than NiCo₂O₄/CNT (340 mV), NiCo₂O₄/CNT-250 (390 mV). The overpotentials for NiCo₂O₄/CNT-150 at 10 mA cm⁻² and 100 mA cm⁻² was 330 mV and 470 mV smaller than that of NiCo₂O₄/CNT, NiCo₂O₄/CNT-250 and commercial RuO₂. Comparison with other reported electrocatalyst material was summarized in Table S1.† The overpotential of NiCo₂O₄/CNT-150 is even lower than the water oxidation catalyst Au/

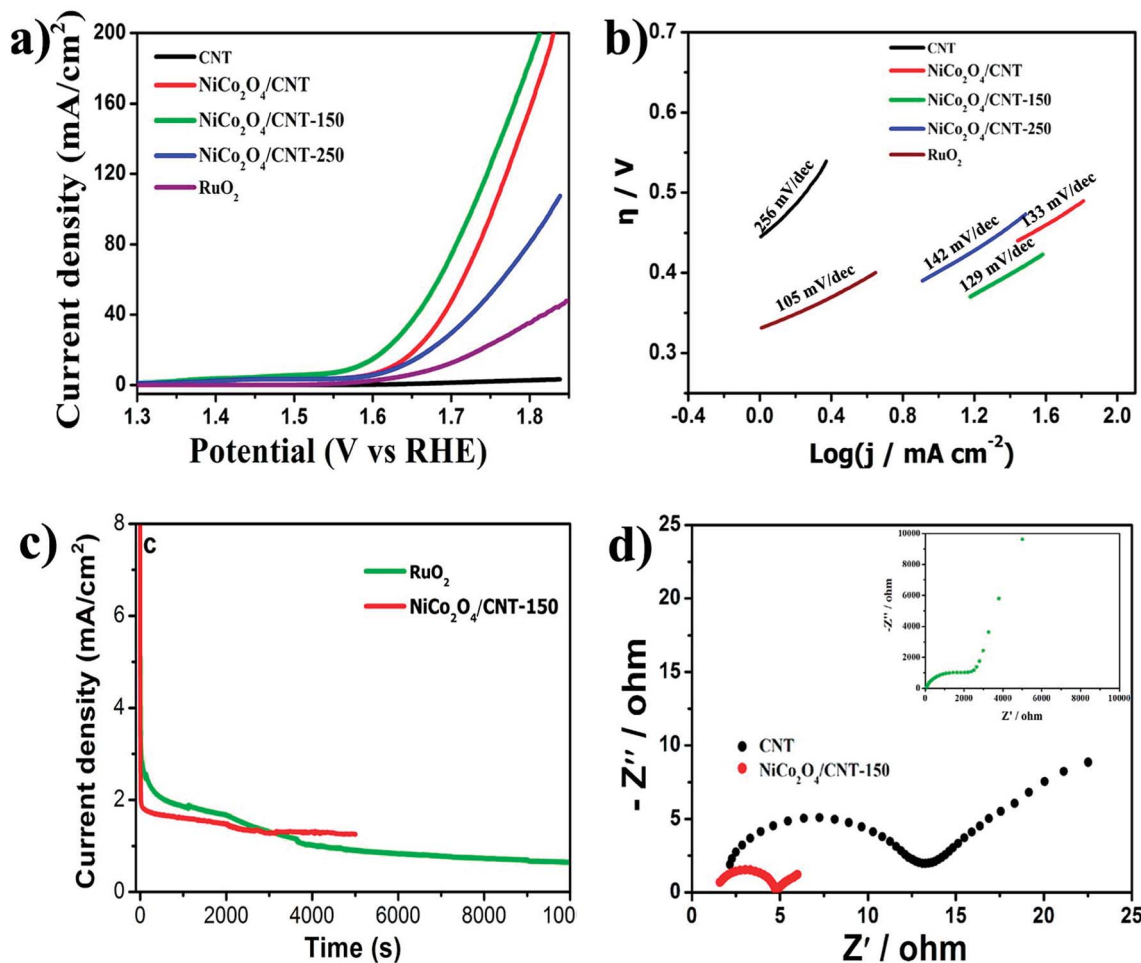


Fig. 5 (a) LSV curves, (b) Tafel plots, (c) stability measurement and (d) EIS spectra of CNT, NiCo₂O₄/CNT, NiCo₂O₄/CNT-150, NiCo₂O₄/CNT-250 and RuO₂ electrodes.



NiCo₂O₄ nanorod array with the overpotential of 360 mV at 10 mV cm⁻². These results further confirm the lattice defects on 3D flower-like NiCo₂O₄/CNT-150 may be more beneficial for affording more catalytic sites, effectively increasing catalytic activity area and further strengthening the OER performance. The Tafel behavior, especially the Tafel slope, is an important OER kinetic parameter. The superior kinetics of NiCo₂O₄/CNT-150 was confirmed by the correspondingly lower Tafel slope of 129 mV dec⁻¹, compared with that of NiCo₂O₄/CNT (133 mV dec⁻¹), NiCo₂O₄/CNT-250 (142 mV dec⁻¹), CNT (256 mV dec⁻¹) and commercial RuO₂ (105 mV dec⁻¹), as shown in Fig. 5b. The results confirm that NiCo₂O₄/CNT-150 possesses efficient kinetics for OER. The high catalytic activity is of very importance for the large scope application in the water splitting.

The stability of OER on the as-synthesized electrocatalysts was investigated at 1.63 V *versus* RHE using chronoamperometry, as presented in Fig. 5c. Due to the production of oxygen, the electrocatalysts was happened to fall off from the electrode and thus the time of the stability measurement is controlled at 5000 s for NiCo₂O₄/CNT-150 and 10000 s for the commercial RuO₂. Obviously, at initial 3000 s in Fig. 5c, the commercial RuO₂ has better stability even if the current densities of the commercial RuO₂ and NiCo₂O₄/CNT-150 both gradually decrease. The current density of NiCo₂O₄/CNT-150, however, remains hardly change after about 3000 s in contrast to the lasting decrease of the current density of the commercial RuO₂. It is demonstrated that NiCo₂O₄/CNT-150 has better electrocatalytic stability. The electrochemical impedance spectra (EIS) obtained at 1.63 V (*vs.* RHE) was shown in Fig. 5d. In the high frequency region, the solution resistance is region from the interface resistance between the electrolyte and the electrode.³⁹ NiCo₂O₄/CNT-150 shows the lowest solution resistance of only 1.3 Ω. In low frequency region, it is clear that the NiCo₂O₄/CNT-150 still shows the lowest charge transfer resistance, indicating the rapid electron transport during OER process in contrast to the CNT and the commercial RuO₂ seen in the inset of Fig. 5b. Therefore, the CVD growth enables NiCo₂O₄ and CNT to form integral electrical conductor, facilitating the accelerated electrochemical kinetics and contributing to a superior electrochemical performance for OER.

The morphology and structure changes of NiCo₂O₄/CNT-150 electrode after all the electrochemical characterization were investigated by SEM and XRD, as shown in Fig. S3 and S4.† Fig. S3† shows the unchanged morphology of the NiCo₂O₄/CNT-150. Both the size and flower shape were still remained compared with the sample before electrochemical characterization in Fig. 2c. XRD spectra in Fig. S4† further demonstrated the outstanding stability of NiCo₂O₄/CNT-150, where it held nearly the same phase composition and crystal structure. The result further confirmed an excellent stability of NiCo₂O₄/CNT-150 in electrochemical process.

Conclusions

In this work, 3D flower-like NiCo₂O₄ spheres were successfully prepared on CNTs by CVD growth combined with the

hydrothermal method. The morphology of NiCo₂O₄ on CNTs was readily controlled *via* adjusting the subsequent calcination temperature. The as-prepared NiCo₂O₄/CNT-150 composite shows lower overpotential, larger current density and higher electrochemical stability. The onset overpotential of NiCo₂O₄/CNT-150 is 300 mV, which was lower than NiCo₂O₄/CNT-250 (390 mV). Furthermore, The overpotential at 10 mA cm⁻² current density for NiCo₂O₄/CNT-150 is 330 mV smaller than that of commercial RuO₂. The superior kinetics of NiCo₂O₄/CNT-150 is confirmed by the corresponding lower Tafel slope of 129 mV dec⁻¹, compared with that of NiCo₂O₄/CNT (133 mV dec⁻¹), NiCo₂O₄/CNT-250 (142 mV dec⁻¹) as well as CNT (256 mV dec⁻¹), which are mainly ascribed to the ascendant conductivity of CNT substrate and the 3D flower shape of NiCo₂O₄. All these characteristics are greatly beneficial for oxygen evolution reaction. This work offers a strategy to enhance the OER performance of NiCo₂O₄ on CVD grown carbon by effectively increasing the electrical conductivity of NiCo₂O₄.

Conflicts of interest

This work has not been published previously. It is not under consideration for publication elsewhere, and its publication is approved by all authors and tacitly or explicitly by the responsible authorities where the work was carried out. If accepted, it will not be published elsewhere in the same form, in English or in any other language, without the written consent of the Publisher.

Acknowledgements

Financial supports from the National Natural Science Foundation of China (U1401246, 51474110, 51474077 and 21473042) are gratefully acknowledged.

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