


Cite this: *RSC Adv.*, 2018, 8, 37765

Facile preparation of degradable multi-arm-star-branched waterborne polyurethane with bio-based tannic acid

Shun Luo, Kai Yang, Zhen Zhong, Xiaojian Wu and Tianbin Ren *

In this research, biodegradable multi-arm-star-branched waterborne polyurethanes (MWPU) were prepared by incorporation of bio-based material (tannic acid, TA) in the structure of waterborne polyurethanes. The prepared MWPU were characterized by UV-vis spectrometry and FT-IR spectrometry, confirming the presence of TA in MWPU. The results of DSC and TGA demonstrated that the incorporation of TA remarkably enhanced the thermal stability of MWPU. The mechanical strength test indicated that the Young's modulus and tensile strength of the waterborne polyurethanes after incorporation of TA were significantly improved due to the increase of structural rigidity, hydrogen bonding and the molecular interactions of the TA-based MWPU chains. In addition, the synthesized TA-based MWPU also exhibited excellent antioxidation capacity and outstanding biodegradation property. Given these excellent properties and the sustainability of TA, the developed TA-based MWPU exhibited great potential in a wide range of practical applications.

Received 21st September 2018

Accepted 3rd November 2018

DOI: 10.1039/c8ra07875k

rsc.li/rsc-advances

Introduction

Eco-friendly waterborne polyurethanes (WPU) have gained increasing attention in a wide range of applications including coatings, adhesives, elastomers and biomedicine due to their good flexibility, non-toxicity and non-flammability as well as their reduced volatile organic compounds (VOCs) release compared with their solvent-based counterparts.^{1–3}

Recently, bio-based raw materials such as castor oil and soybean oil have been exploited as components to synthesize WPU, providing an effective and efficient approach for developing bio-based WPU.^{4–6} The biological polyurethane material is non-toxic, biodegradable and biocompatible, which is very promising in commercial applications. The degradation properties of biodegradable materials as medical materials are of interest, as a long-term material implanted in the body is generally not expected to degrade. But when used as similar tissue engineering scaffolds, medical adhesives and drug sustained-release materials, it is hoped that they can be gradually degraded in the body. Besides, biodegradable multi-arm-star-branched waterborne polyurethane can be used to make biodegradable condoms.^{7–10}

In this research, biodegradable multi-arm-star-branched waterborne polyurethanes (MWPU) were prepared by incorporation of bio-based material (tannic acid, TA) in the structure of waterborne polyurethanes. The structure of the prepared MWPU was characterized with UV-vis spectrometer and FT-IR spectrometer. The effects of incorporation of TA on the thermal stability and mechanical properties of WPU were investigated. Furthermore, the antioxidation capacity of TA-based MWPU was tested *via* DPPH and hydroxyl radical scavenging assay. Finally, the biodegradation behaviour of TA-based MWPU in simulated body fluid and natural soil was discussed.

Experimental

Materials

Bis(hydroxymethyl) propionic acid (DMPA) was purchased from Shanghai Demand Chemical Co., Ltd, China. Poly (tetramethylene ether glycol) (PTMG, Mn = 2000) and isophorone diisocyanate (IPDI) were obtained from Jining Baichuan Chemical Co., Ltd, China. L-Ascorbic acid and di(phenyl)-(2,4,6-trinitrophenyl) iminoazanium (DPPH) were obtained from Sigma Aldrich (St. Louis, MO, USA). Dibutyltin dilaurate and stannous octoate were provided by Jiangsu Yoke Technology Co., Ltd, China. Triethylamine (TEA), diethylene glycol (DEG), tannic acid (TA), tetrahydrofuran (THF), ethanol, hydrogen peroxide, 2-hydroxybenzoic acid, ferrous sulfate, sodium chloride, potassium chloride, sodium dihydrogen phosphate and potassium dihydrogen phosphate were

Institute of Nano and Biopolymeric Materials, School of Materials Science and Engineering, Key Laboratory of Advanced Civil Engineering Materials, Ministry of Education, Tongji University, 4800 Caoan Road, Shanghai 201804, China. E-mail: rentianbin008@163.com; Fax: +86-21-33515906; Tel: +86-21-33515906



obtained from Sinopharm Chemical Reagent Shanghai Co., Ltd., China and used as received.

Preparation of TA-based MWPU and films

The synthesis process of TA-based MWPU is shown in Scheme 1. A typical procedure is as follows: PTMG-2000, TA and DMPA were preprocessed to remove the moisture before using *via* drying them in a vacuum oven at 110 °C, 75 °C and 75 °C for 2 h, respectively. Then, PTMG-2000 (100 g, 0.05 mol) was placed in a 50 ml four-neck round-bottom flask under 65 °C, followed by addition of IPDI (33 g, 0.15 mol). Afterwards, the mixture was heated to 80 °C and kept for 2 h to prepare the -NCO-ended polyurethane prepolymer. DMPA (7.98 g, 0.06 mol), DEG, dibutyltin dilaurate and stannous octoate were then added for further reaction of 4 h to get the polyurethane prepolymer with -COOH groups. The mixture was cooled to 25 °C, and then TA solution dissolved in THF was added, followed by keeping the reaction at 60 °C for 5 h. After that, the mixture was placed in a water bath until the temperature dropped to 25 °C, and was then neutralized by slowly adding TEA (6.02 g, 0.06 mol) under stirring (175–200 rpm) in 45 min. Subsequently, the polymer mixture was vigorously stirred (1400–1500 rpm) for 30 min and emulsified with a certain amount of deionized water to obtain a WPU dispersion. Finally, the WPU solution with solid content of 35% was obtained by removal of THF *via* vacuum evaporation. A series of WPUs (WPU0, MWPU1, MWPU2, MWPU3) were prepared *via* adjusting the ratios of TA (Table 1).

The polyurethane films were prepared by pouring WPUs and MWPU dispersions into PTFE module, and then drying at room temperature for 24 h. Then as-prepared films with thickness of 0.5 mm was further dried in a vacuum oven to a constant weight at 60 °C and were then stored in a desiccator prior to testing and characterization.

Free radical scavenging assay

DPPH free radical scavenging assay. The mechanism of DPPH radical scavenging assay was shown in Fig. 1. Typically, DPPH[•] is an available organic nitrogen radical in ethanol solution and has strong absorption maximum at 517 nm. DPPH[•] can accept hydrogen from an antioxidant to form DPPH₂, which results in the colour of solution turning from purple to yellow and the decrease of UV absorption at 517 nm. Therefore, the antioxidant capacity can be easily evaluated by measurement of the UV absorption at 517 nm of DPPH[•] solution.

In our study, to test the radical scavenging activity, 1 ml of 0.1 M DPPH (in 50% ethanol solution) was mixed with 1 ml of tannic acid-based hyperbranched WPUs solution (0.2–2.0 mg ml⁻¹ in double distilled water). The reaction mixture was incubated at room temperature for 20 min under dark conditions. Then the absorbance was measured at 517 nm, and the scavenging activity of tannic acid-based hyperbranched WPUs was measured by using the following eqn (1),

$$\text{Scavenging efficiency (\%)} = 1 - \frac{A_1}{A_0} \times 100\% \quad (1)$$

where A_1 is the absorbance of DPPH solution with the presence of samples and A_0 is the absorbance of the control. For the control, a same volume of DI water was added into the test tube instead of polymeric solution. In addition, L-ascorbic acid was used as a positive control.^{11–13}

Hydroxyl free radical scavenging assay. To measure the hydroxyl radical scavenging, the prepared multi-arm-star-branched WPUs was dissolved in deionized water with concentration of 0.2–2.0 mg ml⁻¹. 2 ml of polymeric solution was taken out and added into the test tube. Then, ferrous sulfated aqueous solution (0.1 M) and salicylic acid solution (0.1 M) in ethanol were added into the test tube, followed by adding H₂O₂ (0.1 M, 2 ml). The mixture was shaken well and incubated in 37 °C water bath for 30 min. The absorbance was measured at 510 nm, and the scavenging ability for hydroxyl radicals was calculated by using the following eqn (2),

$$\text{Scavenging efficiency (\%)} = 1 - \frac{A_a}{A_b} \times 100\% \quad (2)$$

where A_a is the absorbance of the samples and A_b is the absorbance of the control. For the control, a same volume of DI water was added into the test tube instead of polymeric solution. In addition, L-ascorbic acid was used as a positive control.^{14–16}

Degradation behaviour of MWPU in the simulated body fluid

Before testing degradation behaviour of WPU and multi-arm-star-branched WPU films in the simulated body fluid, the film coupons with dimension of 10 mm × 10 mm × 1 mm were dried in a vacuum oven to a constant weight at 50 °C and weighed *via* digital balance (W_0). Then the coupons were immersed into a glass vial containing 10 ml of the PBS simulated body fluid (pH = 7.4) which was prepared by dissolving 0.8 g of NaCl, 0.2 g of KCl, 1.44 g of NaHPO₄, and 0.24 g of KH₂PO₄ in 1 L of DI water, followed by incubation at 37 °C. The PBS simulated body fluid was replaced once a day, and the coupons were taken out at predetermined time intervals, washed with DI water for three times and dried in the vacuum oven at 45 °C. The weight of dried coupons (W_t) were then measured. The weight loss ratio was calculated using the eqn (3),

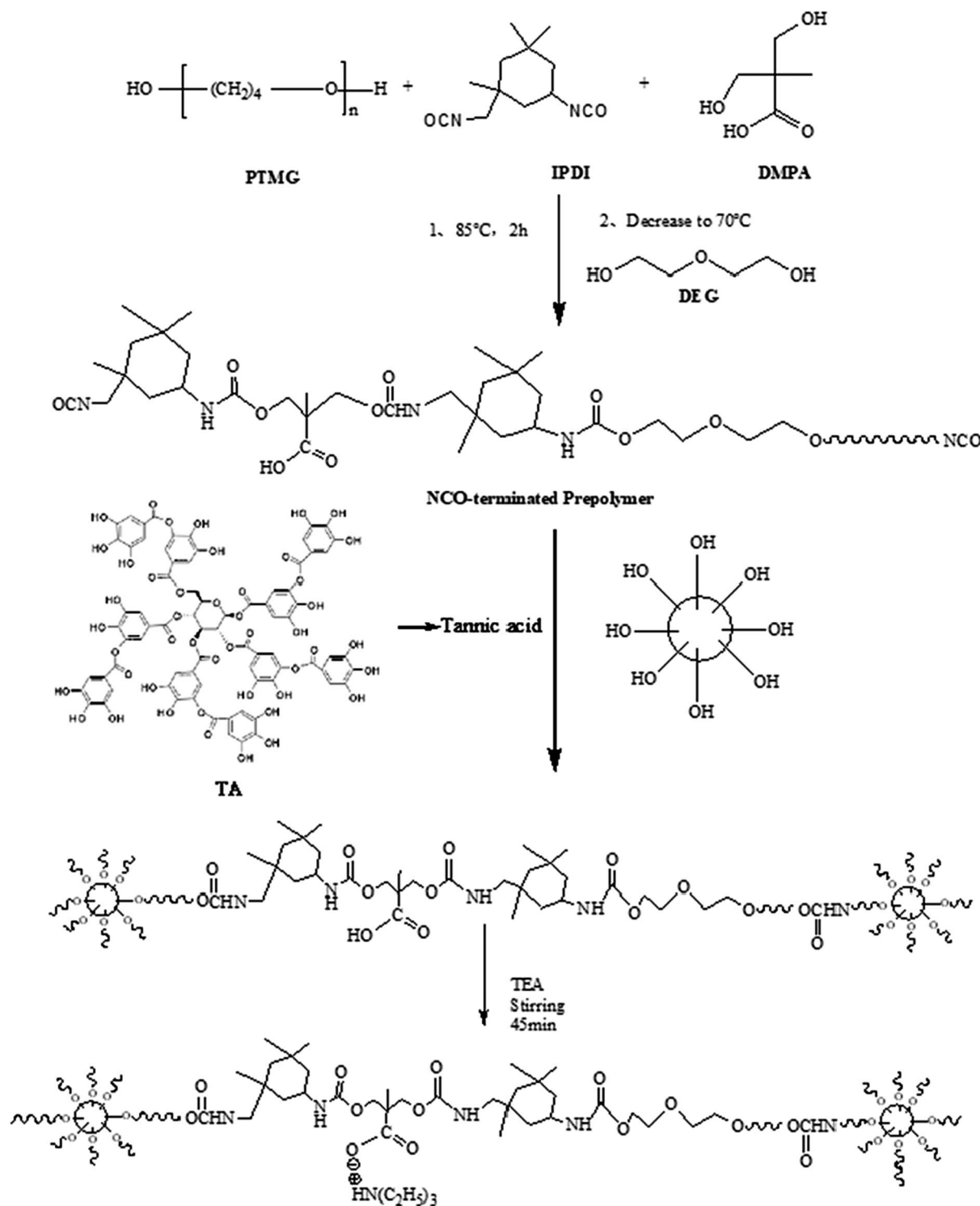
$$\text{Weight loss ratio (\%)} = \frac{W_0 - W_t}{W_0} \times 100\% \quad (3)$$

where W_0 and W_t are the weights of the samples before and after degradation.^{17,18}

Biodegradation study of multi-arm-star-branched WPU in soil

The biodegradation behaviour of multi-arm-star-branched WPU in soil was further investigated following a previously reported protocol. Typically, the soil taken from the lawn located in the Department of Material Science of Tongji University was dried under sunlight for three days. Then the soil was shattered, and the grain of sand was removed from soil. The treated soil was placed in a 250 ml beaker, and 100 ml of DI water was then





Scheme 1 Synthesis of TA-based MWPU.

added into the beaker. Before introducing the film samples into the pasty soils, each sample was cut into pieces of dimension of 10 mm × 10 mm × 1 mm and dried in a vacuum oven to a constant weight (W_i) at 50 °C. The beaker containing the film samples were buried in a soil bed at a depth of around 15–20 cm from the top of the ground surface. The biodegradation experiments were carried out under ambient conditions with regular

addition of water (except in rainy days). Samples were taken out carefully from each beaker after 30 days, 60 days, 90 days and 120 days, respectively, and washed repeatedly with water to remove the adhered soil from the surface of the films. Then the samples were dried in a vacuum oven until a constant weight was obtained. Finally, the weight of samples was measured (W_f). The weight loss ratio was calculated using the eqn (4),



Table 1 Composition of reactants for synthesis of WPU and TA-based MWPU

Materials	WPU0	MWPU1	MWPU2	MWPU3
PTMG (mol)	1.00	1.00	1.00	1.00
IPDI (mol)	3.00	3.00	3.00	3.00
DMPA (mol)	1.20	1.20	1.20	1.20
DEG (mol)	0.30	0.23	0.15	0.18
TA (mol)	0	0.05	0.10	0.15
TEA (mol)	1.20	1.20	1.20	1.20
NCO/OH (ratio)	1.20	1.20	1.20	1.20

$$\text{Weight loss ratio(\%)} = \frac{W_i - W_t}{W_i} \times 100\% \quad (4)$$

where W_i and W_t are the weights of the samples before and after degradation.^{19–22}

Characterization

The IR spectra of the prepared multi-arm-star-branched WPUs were obtained using a Nicolet 6700 FT-IR Spectrometer (Thermo Scientific), and the structure of the film samples before and after degradation was measured under the attenuated total reflection (ATR) mode. To evaluate the relative molecular weight and its distribution of the synthesized multi-arm-star-branched WPUs, gel permeation chromatography (GPC) spectrum was measured in THF with a concentration of 2 mg ml^{−1}. The flow rate and temperature were set at 1.0 ml min^{−1} and 40 °C, respectively. The intrinsic viscosity [η_r] of the synthesized multi-arm-star-branched WPUs solutions was determined at 20 °C with a Bohlin Rotational Rheometer (Malvern Instrument, UK) equipped with cone-and-plate geometry (40 mm diameter, cone angle 4°). The rotate speed was 300 radians per second. UV-vis absorption spectra of the WPUs and multi-arm-star-branched WPUs samples were recorded on a HITACHI U-1800 spectrophotometer at room temperature. The concentration of the WPUs and multi-arm-star-branched WPUs solutions was 0.1 mg ml^{−1}. The surface morphology of the film samples prepared with WPUs and multi-arm-star-branched WPUs was observed using a scanning electron microscopy (SEM, FEI, USA). All samples were coated with 10 nm thick platinum (Pt) before SEM observation. To study the thermal stabilities of the WPUs and multi-arm-star-branched WPUs, thermogravimetric analysis (TGA) was performed on TA Instruments Q500 (New Castle, DE, USA) in N₂ atmosphere with a purge rate of 50 ml min^{−1}. The data was recorded from 20 to 800 °C with a scanning rate of

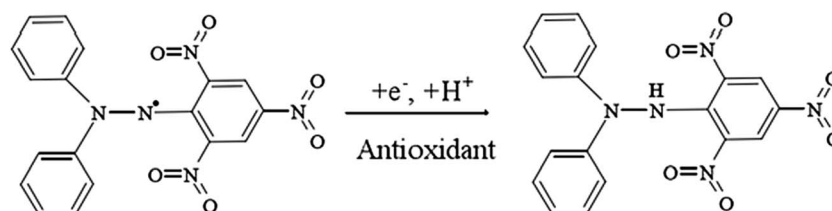
20 °C min^{−1}. Besides, differential scanning calorimetry (DSC) was also carried out on a TA Instruments DSC-Q20 from −80 to 200 °C with a heating/cooling rate of 10 °C min^{−1} in nitrogen atmosphere. To study the mechanical properties of the samples, the tensile strength, Young's modulus and elongation at break were studied by using a tensile tester (INSTRON5565, USA) at a speed of 50 mm min^{−1} at room temperature. Three specimens were tested and averaged for each sample. The specimen was molded into 45 mm long, 8 mm wide and 0.2 mm thick.

Results and discussion

Characterization of MWPU with TA

As shown in Scheme 1, there are five branches in the structure of TA, which are generating from a pyranose heterocycle, 25 hydroxyl groups, and 10 ester linkages. Therefore, in our study, bio-based tannic acid was used as an alternative to polyol to synthesize the multi-arm-star-branched WPUs. To confirm the presence of TA in the synthesized MWPU1, MWPU2 and MWPU3, the UV-vis absorption spectroscopy was used to track the characteristic absorbance peak of the TA at around 280 nm. As shown in Fig. 2a, the prominent absorption peaks were observed for MWPU1, MWPU2 and MWPU3, and the intensity of peaks of the synthesized WPUs exhibited an obvious increase ranging from MWPU1 to MWPU3, indicating the higher TA content in the structure of MWPU3 compared with MWPU1 and MWPU2.

FT-IR spectroscopy was further employed to assess the chemical structures of synthesized MWPU with different TA content. As shown in Fig. 2b, for the spectrum of MWPU1, the absorption peaks focused at 3422 cm^{−1}, 1705 cm^{−1} and 1556 cm^{−1} are mainly attributed to the symmetric stretching vibration of −N−H bond, −C=O group and the bending vibration of −N−H bond, respectively, indicating the successful formation of carbamate in MWPU1. The peaks observed at 2940 cm^{−1}, 2860 cm^{−1}, 1105 cm^{−1} and 1254 cm^{−1} are ascribed to the symmetric stretching vibration of −CH₂ group, the asymmetric stretching vibration of −CH₃ group, the symmetric bending vibration of C−O−C bond and the stretching vibration of −C−N bond. In addition, the peak positions of −C−N bonds in the spectra of MWPU2 and MWPU3 exhibited an obvious blue shift, which is due to the hydrogen bonds formed between the catechol groups of TA and −C−N bond and −C=O group of carbamates, and the number of hydrogen bond increases with the increase of TA content in MWPU2 and MWPU3. This result

**Fig. 1** Mechanism of DPPH radical scavenging assay.

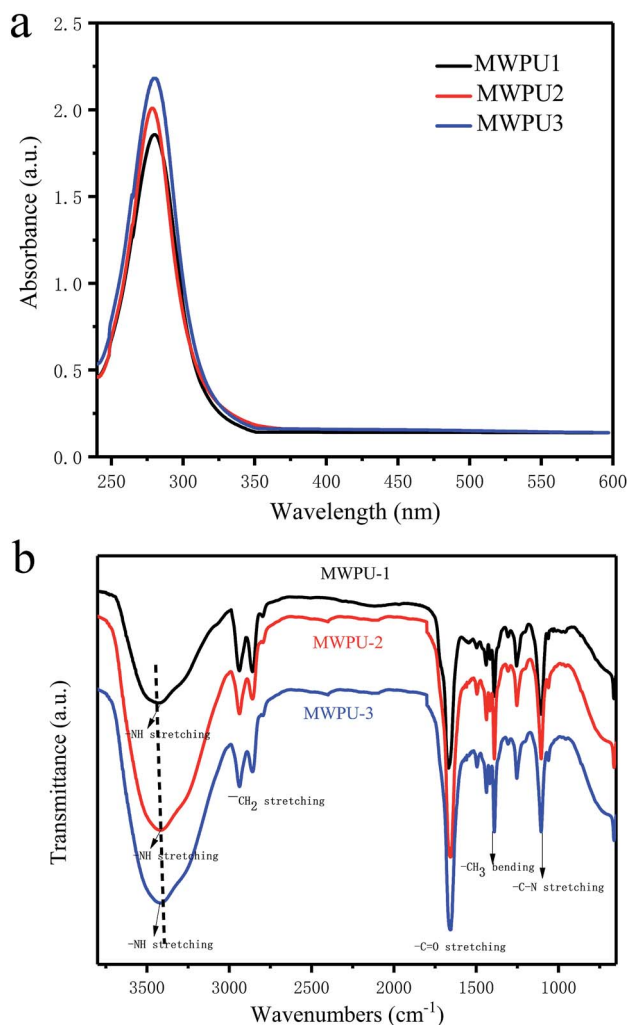


Fig. 2 UV-vis spectra (a) and FT-IR spectra (b) of synthesized MWPU1, MWPU2 and MWPU3.

further confirmed the presence of TA in the synthesized MWPU1, MWPU2 and MWPU3.

Then the molecular weight and distribution of the synthesized MWPU1, MWPU2 and MWPU3 were measured *via* GPC and the results presented in Table 2 revealed that the weight average molecular weight (M_w) of the obtained MWPU3 exhibited a significant increase to 47 200 Da from 33 100 Da of MWPU1 with the increase of the addition amount of TA in the reaction system. The polydispersity index (PDI) of the obtained MWPU1, MWPU2 and MWPU3 was measured to be 1.40, 1.34, and 1.34, respectively, indicating that all MWPU3 have narrow molecular weight distribution and the TA content in the MWPU3 did not affect the molecular weight distribution significantly. Thus, the molecular weights of the MWPU3 can be well controlled *via* adjusting the addition amount of TA in the synthesis process, without affecting the molecular weight distribution. The intrinsic viscosities of the obtained MWPU1, MWPU2 and MWPU3 solution were determined with a rotational rheometer. Compared with the reported WPU3, MWPU3 prepared with TA exhibited clearly reduced intrinsic viscosities

Table 2 Weight average (M_w) and number average (M_n) molecular weight, PDI and solution viscosity of synthesized MWPU3

Parameter	MWPU1	MWPU2	MWPU3
M_w	33 100	45 000	47 200
M_n	23 600	33 600	35 200
PDI	1.40	1.34	1.34
Viscosity (ml g ⁻¹)	0.441	0.349	0.321

(0.441, 0.349 and 0.321 mg L⁻¹ for MWPU1, MWPU2 and MWPU3, respectively), which is due to the very compact three-dimensional structures of MWPU1, MWPU2 and MWPU3 that reduces the chain interactions significantly. This result is also generally consistent with the previously reported results of other hyperbranched polymers.

The influences of TA on the thermal properties of the synthesized MWPU3

To evaluate the influences of TA on thermal properties of MWPU3, the samples of WPU0 and MWPU3 were analysed by DCS and TGA. As shown in Fig. 3, an obvious absorption peak at 15 °C appeared for the sample of WPU0 due to the melting of the soft segment, indicating a certain degree of micro-crystallization. However, the other MWPU3 samples did not have any melting absorption peaks, indicating that the introduction of rigid tannic acid disturbed the crystallization region of the pristine waterborne polyurethane.

As shown in Fig. 4, the synthesized WPU0 and MWPU3 exhibited three-stage thermal degradation process: (i) at the first stage, the decomposition of some thermolabile urethane bonds and ester bonds occurred around 250 °C; (ii) the second stage was corresponding to the decomposition of the aromatic rings from tannic acid (TA) around 300 °C; (iii) the third stage was around 420 °C, involving the decomposition of the soft segment on the polyurethane backbone. Furthermore, the temperatures at different weight loss ratios for the samples of WPU0 and MWPU3 were listed in Table 3. It could be seen that as the content of TA increases, the thermal decomposition temperature of the

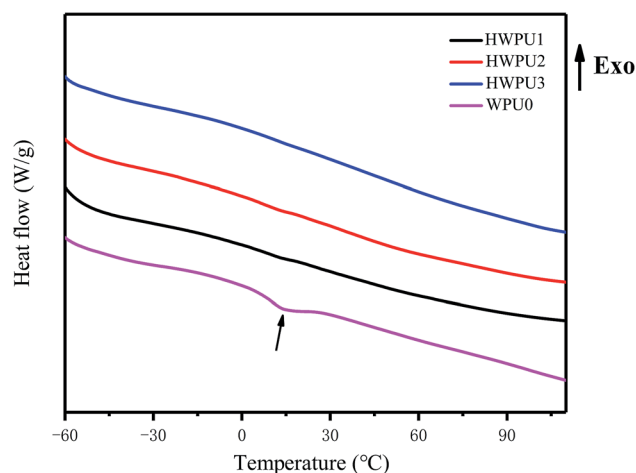


Fig. 3 DSC curves of synthesized WPU0 and MWPU3.



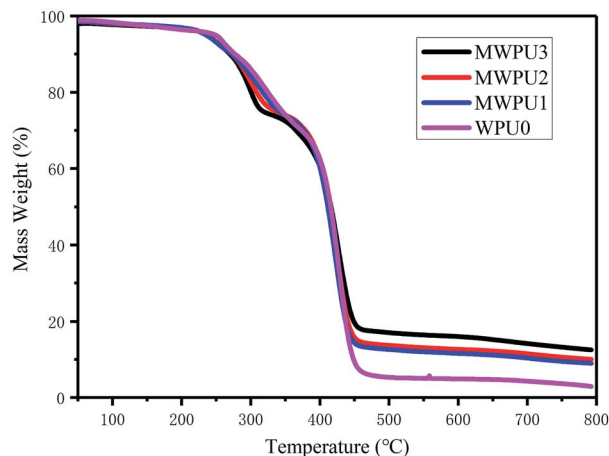


Fig. 4 TGA curves of synthesized WPU0 and MWPU1.

Table 3 Thermal decomposition temperature corresponding to different numerical Thermogravimetry

Sample	$T_{10\%}$ (°C)	$T_{20\%}$ (°C)	$T_{30\%}$ (°C)	$T_{40\%}$ (°C)	$T_{50\%}$ (°C)
WPU0	269.66	300.76	365.44	400.23	415.73
MWPU1	271.84	308.46	373.58	401.23	411.39
MWPU2	272.18	317.47	376.58	402.72	414.06
MWPU3	277.02	324.83	378.91	404.39	416.06

polyurethanes increases. For example, the $T_{10\%}$ of WPU0 and MWPU3 were 269.66 °C and 277.02 °C. The $T_{30\%}$ of WPU0 and MWPU3 were 365.44 °C and 378.91 °C. So, the introduction of TA increased the thermal decomposition temperature of the waterborne polyurethanes. This was not only because of the introductive large amount of thermally-stable benzene ring, but also due to the secondary interactions between the ester bonds and heterocycles. Besides, it should be noted that the effect of TA was mainly on the first two thermal decomposition stages. Besides, residual polyurethanes at 700 °C enhanced as the content of introductive TA increased.^{23–25}

The influences of TA on the mechanical properties of the synthesized MWPU1

The mechanical properties of the waterborne polyurethanes were analysed by the Young's modulus, tensile strength and elongation at break. The results were shown in Fig. 5 and Table 4. Young's modulus is a physical quantity that describes the deformation resistance of solid materials, which depends only on the physical properties of the material itself. The size of young's modulus indicates the rigidity of the material. The larger young's modulus is, the less likely it is to deform.

As the content of TA increases, the Young's modulus and tensile strength of the waterborne polyurethanes increased. Specifically, the Young's modulus of WPU0 and MWPU3 were 5.1 MPa and 12.4 MPa, respectively. Meanwhile, the tensile strength of WPU0 and MWPU3 were 8.5 MPa and 13.6 MPa, respectively. The increase of the Young's modulus and tensile strength of the waterborne polyurethanes after incorporation of

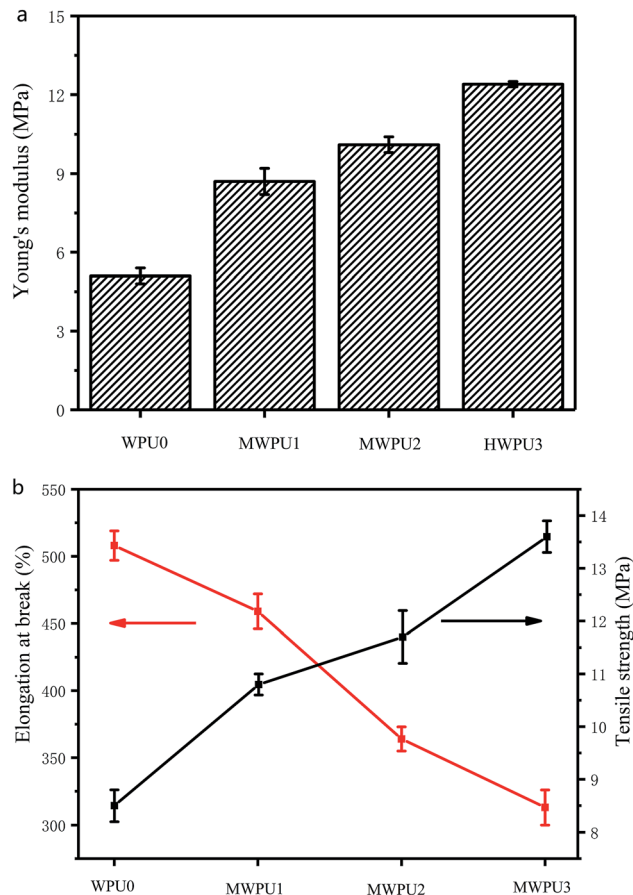


Fig. 5 Young's modulus (a), tensile strength and elongation at break (b) of synthesized WPU0 and MWPU1.

TA is attributed to the fact that the introduction of TA not only enhances the structural rigidity of waterborne polyurethanes, but also increases the hydrogen bonds and the molecular interactions of waterborne polyurethanes, thus ultimately enhancing the mechanical strength of waterborne polyurethanes. However, it is worth noting that the increase of TA content in the synthesized MWPU1 resulted in a decrease in the elongation at break of the waterborne polyurethanes (Table 4), which is due to the decrease of the flexibility of chains of MWPU1.

The antioxidation capacity of the synthesized MWPU1

DPPH radical scavenging assay. Typically, the antioxidation capacity of materials is mainly depending on their ability to donate electron to radical. To investigate the effects of

Table 4 Young's modulus, tensile strength and elongation at break of the synthesized WPU0 and MWPU1

Sample	Young's modulus (MPa)	Tensile strength (MPa)	Elongation at break (%)
WPU0	5.1 ± 0.3	8.5 ± 0.3	508 ± 11
MWPU1	8.7 ± 0.5	10.8 ± 0.2	459 ± 13
MWPU2	10.1 ± 0.3	11.7 ± 0.5	364 ± 9
MWPU3	12.4 ± 0.1	13.6 ± 0.3	313 ± 13



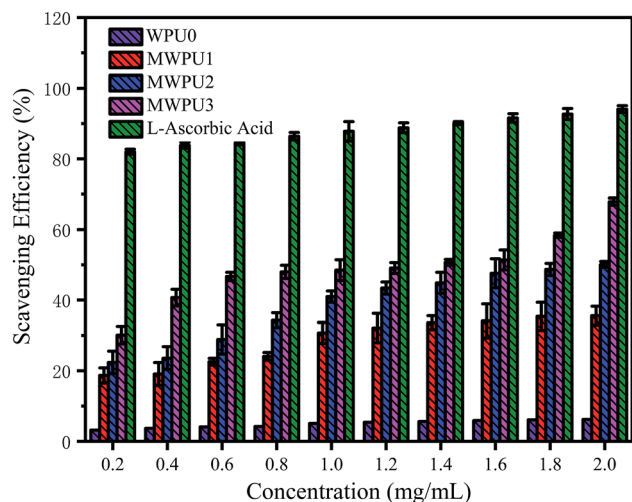


Fig. 6 DPPH radical scavenging efficiency of synthesized WPU0 and MWPU1-3.

incorporation of TA in MWPU1-3 on their antioxidation capacity, the DPPH radical scavenging activities of the synthesized WPU0 and MWPU1-3 were measured. As illustrated in Fig. 6, the WPU0 without TA did not exhibit any DPPH radical scavenging activity, whereas, the synthesized TA-based MWPU1-3 showed outstanding radical scavenging activity. Besides, the radical scavenging activity of TA-based MWPU1-3 enhanced as TA content increases. This result demonstrated that TA is the main component in MWPU1-3 contributing to the antioxidant activity. In addition, the radical scavenging activity of MWPU1-3 also gets strengthened as the concentration of MWPU1-3 increases. Particularly, when the solution concentration of MWPU1-3 is 2 mg mL^{-1} , the DPPH radical scavenging efficiency is up to 67%. This excellent DPPH radical scavenging activity is mainly attributed to the catechol groups of TA that can serve as electron donors and react with free radicals effectively to convert them into more stable products.^{26,27}

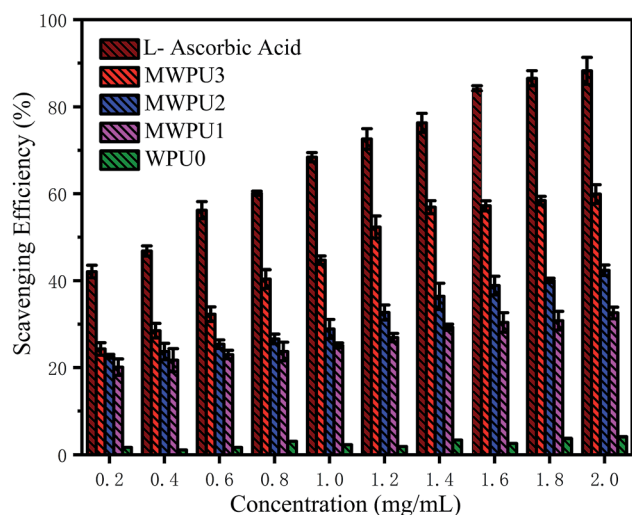


Fig. 7 Hydroxyl radical scavenging efficiency of synthesized WPU0 and MWPU1-3.

Hydroxyl radical scavenging assay. Hydroxyl radical, as the most toxic radical among the reactive oxygen species (ROS) and one of the most active radicals, can react with a range of substances in human body, thus leading to a variety of diseases. Given that, hydroxyl radical scavenging activities of the synthesized WPU0 and MWPU1-3 were further investigated. As presented in Fig. 7, WPU0 did not exhibit any scavenging activities for hydroxyl radical, and the hydroxyl radical scavenging activities of TA-based MWPU1-3 increase with the increase of TA content and the concentration of MWPU1-3 in solution. These results are generally consistent with the results of DPPH radical. Particularly, when the concentration of MWPU3 solution is 2 mg mL^{-1} , the hydroxyl radical scavenging efficiency is over 50%, further confirming the excellent antioxidation capacity of the TA-based MWPU1-3.

The degradation behaviours of WPU0 and MWPU1-3

The degradation behaviours of WPU0 and MWPU1-3 were studied by determining weight loss over time under two different conditions including in simulated body fluid and in soil. Fig. 8a represented the degradation curves of WPU0 and MWPU1-3 in

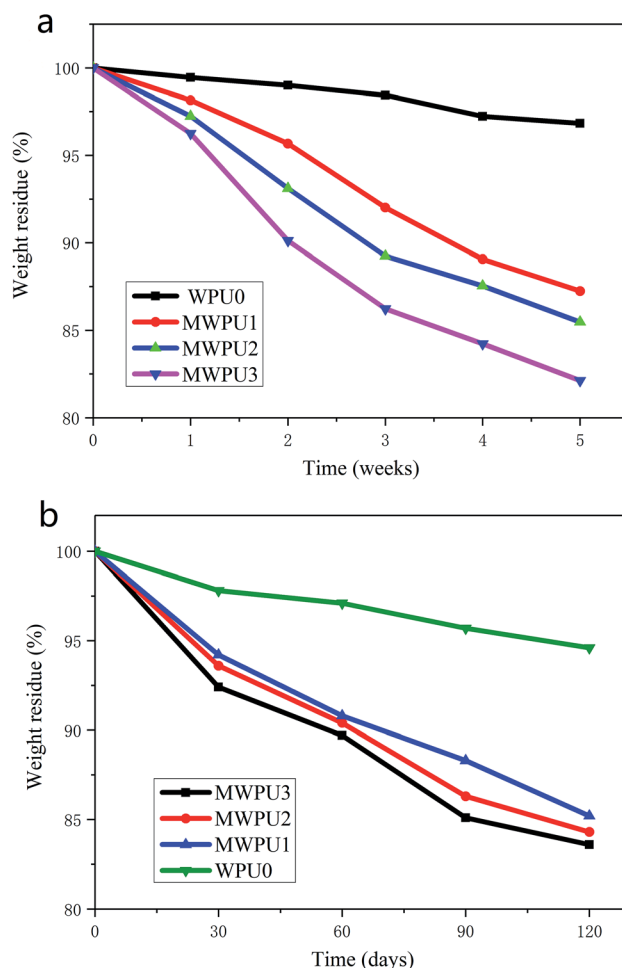


Fig. 8 Degradation curves of synthesized WPU0 and MWPU1-3 in simulated body fluid (a) and in soil (b).



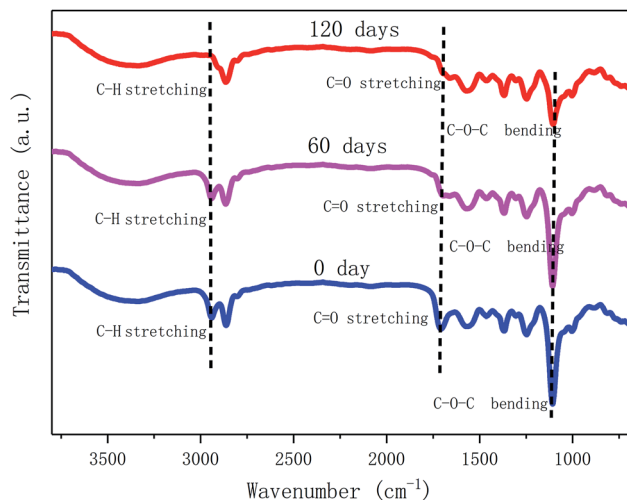


Fig. 9 FT-IR spectra of MWPU3 before and after 60 days and 120 days degradation.

simulated body fluid (PBS solution). It is found that after 5 weeks of degradation in simulated body fluid, the TA-based MWPU3 exhibited much high weight loss compared with WPU0, indicating that TA-based MWPU3 have better degradation performances. And as the increase of TA content in MWPU3, the weight loss of TA-based MWPU3 rising. For instance, the weight loss of MWPU3 after 5 weeks' degradation is up to 18%, which is higher than the performance of MWPU1 and MWPU2. The outstanding degradation performance of

MWPU3 is due to the presence of large number of ester bonds in the chains of MWPU3 that is liable to hydrolysis.

The degradation behaviours of WPU0 and MWPU3 in soil were also investigated and the results are shown in Fig. 8b. After 120 days degradation in soil, TA-based MWPU3 exhibited 15–17% weight loss, which is much higher than 5% weight loss of WPU0, further confirming the better degradation performance of TA-based MWPU3. It is worth noting that although the weight loss of TA-based MWPU3 also exhibited a same trend in soil with the increase of TA content in MWPU3, the increase of TA-based MWPU3 in soil is much lower than that in simulated body fluid.

To further investigate the degradation process of MWPU3 in soil, the structure of MWPU3 and the surface morphology of MWPU3 films were characterized with FT-IR ATR and SEM, respectively. As shown in Fig. 9, significant changes of three peaks concentrate upon $2864\text{--}2940\text{ cm}^{-1}$, 1710 cm^{-1} and 1130 cm^{-1} have been observed. Particularly, the peak at $2864\text{--}2940\text{ cm}^{-1}$ is attributed to the $-\text{CH}_2$ group and the peak intensity exhibited an obvious decrease after 120 days of degradation, which is mainly due to the degradation of C-H bond caused by the hydrocarbon-degrading bacteria. The second peak with significant change at 1710 cm^{-1} , which is ascribed to the $-\text{C}=\text{O}$ group. The peak intensity exhibited an apparent decrease after 60 days of degradation and this peak disappeared completely after 120 days. The degradation of $-\text{C}=\text{O}$ group is mainly caused by hydrolysis and biodegradation of bacteria. The peak attributed to C-O-C bond (1130 cm^{-1}) did not exhibited significant change after 60 days, whereas the peak

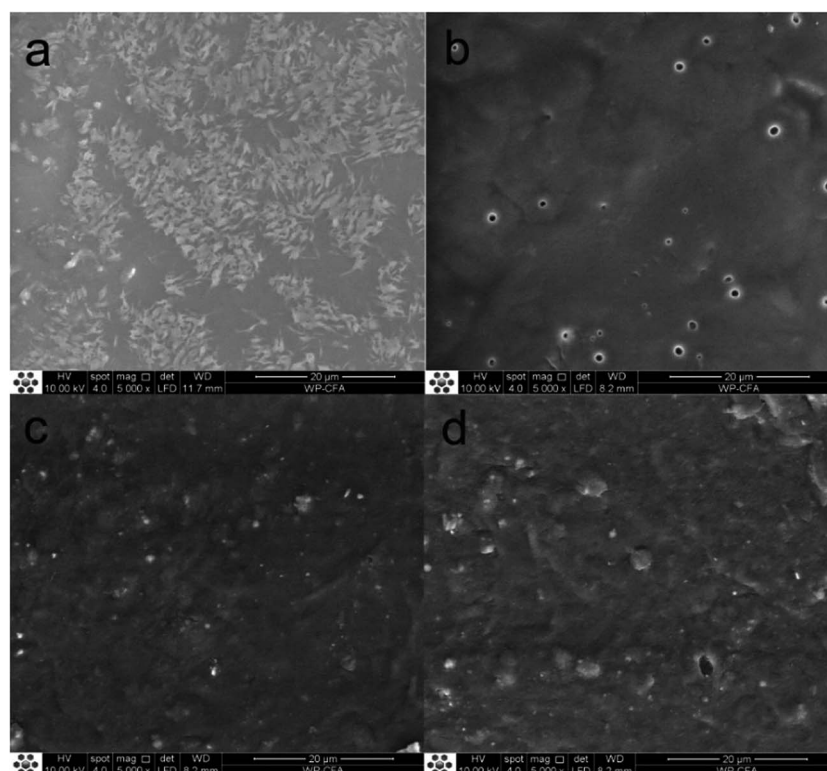


Fig. 10 Change of surface morphology of MWPU3 film before and after degradation. (a) 0 days, (b) 30 days, (c) 60 days and (d) 120 days.



intensity showed a clear decrease after 120 days of degradation, which is probably due to the degradation of enzymes or bacteria in soil. That result also indicated that the degradation of C–O–C bonds mainly occurred in last stage. According to these findings, the degradation of TA-based MWPU3 is mainly due to degradation of ester bonds, C–H bonds and C–O–C bonds caused by the bacteria, enzymes and water existed in soil.^{28,29}

The change of surface morphology of MWPU3 film during the degradation process was presented in Fig. 10. At the initial state, MWPU3 film exhibited a very smooth surface without any defect. However, as time goes, the surface of MWPU3 film become rough and some holes were observed on the film surface after 120 days of degradation, further confirmed the degradation of MWPU3.

Conclusions

In this study, biodegradable MWPU3s were prepared by incorporation of TA in the structure of waterborne polyurethane. The UV-vis spectra and FT-IR spectra of the synthesized MWPU3s confirmed the presence of TA. The result of DSC and TGA demonstrated that the incorporation of TA enhanced the thermal stability of MWPU3s, and Young's modulus and tensile strength of the waterborne polyurethanes after incorporation of TA were increased due to the increase of structural rigidity, hydrogen bonds and molecular interactions of the MWPU3 chains. Besides, the synthesized TA-based MWPU3s exhibited excellent antioxidation capacity and outstanding biodegradation property. Along with these excellent properties and sustainability of TA, the developed TA-based MWPU3s exhibited great potential in a wide range of practical applications. MWPU3 is non-toxic, biodegradable and biocompatible, which is very promising in commercial applications. The degradation property of biodegradable materials as medical materials is concerned, as a long-term material implanted in the body, it is generally not expected to degrade. But, when used as similar tissue engineering scaffolds, medical adhesives and drug sustained-release materials, it is hoped that they can be gradually degraded in the body. Besides, MWPU3 can be used to make biodegradable condoms.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was financially supported by the National Natural Science Foundation of China (81571801, 81671944), the Fundamental Research Funds for the Central Universities and the Open Funds for Characterization of Tongji University.

References

1 X. Ji, H. Wang and X. Ma, *RSC Adv.*, 2017, 7(54), 34086–34095.

- 2 A. Santamaria-Echart, L. Ugarte and C. García-Astrain, *Carbohydr. Polym.*, 2016, **151**, 1203–1209.
- 3 Z. Zhong, S. Luo and K. Yang, *RSC Adv.*, 2017, 7(67), 42296–42304.
- 4 B. B. R. Silva, R. M. C. Santana and M. M. C. Forte, *Int. J. Adhes. Adhes.*, 2010, **30**(7), 559–565.
- 5 K. Liu, Z. Su and S. Miao, *RSC Adv.*, 2016, **6**(38), 31698–31704.
- 6 Z. Luo, Y. Shi and D. Zhao, *Procedia Eng.*, 2011, **18**(18), 37–42.
- 7 I. Phiwichai, W. Yuensook and N. Sawaengsiriphon, *Eur. J. Pharm. Sci.*, 2017, **114**, 64.
- 8 B. Ghosh, S. Gogoi, S. Thakur and N. Karak, *Prog. Org. Coat.*, 2016, **90**, 324–330.
- 9 X. Lai, Z. Hu and Y. Shen, *Petrochem. Technol.*, 2017, **46**(1), 83–89.
- 10 W. Y. Li, Y. M. Cao and X. Q. Zhou, *Appl. Mech. Mater.*, 2012, **249–250**, 842–848.
- 11 D. Villano, M. S. Fernández-Pachón, M. L. Moyá, A. M. Troncoso and M. C. García-Parrilla, *Talanta*, 2007, **71**(1), 230–235.
- 12 Q. Li, L. Qiu and W. Tan, *RSC Adv.*, 2017, 7(67), 42225–42232.
- 13 X. G. Zhang, L. Tang and R. L. Shang, *Storage Process*, 2017, **17**(6), 39–46.
- 14 N. Singh and P. S. Rajini, *Food Chem.*, 2004, **85**(4), 611–616.
- 15 L. Yu, S. Haley and J. Perret, *J. Agric. Food Chem.*, 2002, **50**(6), 1619–1624.
- 16 T. Kokila, P. S. Ramesh and D. Geetha, *Appl. Nanosci.*, 2015, **5**(8), 911–920.
- 17 H. Nur Maulida, F. Qulub and A. F. Rosdiani, *J. Biomimetics, Biomater. Biomed. Eng.*, 2016, **28**, 78–84.
- 18 J. L. Ryszkowska, M. Auguścik and A. Sheikh, *Compos. Sci. Technol.*, 2010, **70**(13), 1894–1908.
- 19 J. Liu, X. Qi and D. Zhang, *Adv. Mater. Sci. Eng.*, 2017, **3**, 1–14.
- 20 B. L. Dou, A. N. Qiu-Feng and M. A. Jun-Jian, *Appl. Chem. Ind.*, 2010, **39**, 189–191.
- 21 B. Ou, M. Chen and R. Huang, *RSC Adv.*, 2016, **6**(52), 47138–47144.
- 22 J. Wang, P. Sun and Z. Zheng, *Polym. Degrad. Stab.*, 2012, **97**(11), 2294–2300.
- 23 G. Zheng, M. Lu and X. Rui, *Appl. Surf. Sci.*, 2016, **399**, 272–281.
- 24 H. Fu, Y. Wang and X. Li, *Compos. Sci. Technol.*, 2016, **126**, 86–93.
- 25 X. Li, Y. Chen and S. Wu, *Polym. Mater.: Sci. Eng.*, 2017, **33**(7), 138–143.
- 26 E. Montanari, A. Gennari and M. Pelliccia, *Macromol. Biosci.*, 2016, **16**(12), 1815.
- 27 M. Shin, K. Kim and W. Shim, *ACS Biomater. Sci. Eng.*, 2016, **2**(4), 687–696.
- 28 L. Cosgrove, P. L. Mcgeechan, G. D. Robson and P. S. Handley, *Appl. Environ. Microbiol.*, 2007, **73**(18), 5817.
- 29 J. C. Peeler, S. Schedin-Weiss and M. Soula, *J. Biol. Chem.*, 2017, **292**(52), 21623–21630.

