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One-step hydrothermal synthesis and optical properties of self-quenching-resistant carbon dots towards fluorescent ink and as nanosensors for Fe³⁺ detection

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In our work, blue photoluminescent N-doped carbon dots (CDs) were developed via a green and simple hydrothermal method with citric acid and polyvinyl pyrrolidone (PVP K-30) as the carbon source and the nitrogen source, respectively. The as-prepared CDs have a high fluorescent quantum yield of 30.21% and considerable luminescence stability. The fluorescence intensity of the CDs was found to be effectively quenched when adding Fe³⁺ ions to the CDs solution. The quenching phenomenon can be used to detect Fe³⁺ ions within a linear range of 0–300 µM with a detection limit of 45.5 nmol L^{−1}, which suggested its potential application in the detection of Fe³⁺ ions. At the same time, we also noted the excellent self-quenching-resistant property of the as-prepared CDs in the solid state, and bright blue fluorescence was observed under UV excitation. What's more, the as-prepared CDs can also be used as fluorescent ink and were presented under UV excitation.

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Introduction

In the past few years, as a new kind of carbon materials, carbon dots (CDs) have great prospects in the application of luminescence, biomedical and the ion detection filed. CDs usually refer to a new class of carbon nanomaterials with a size below 10 nm, which were firstly founded by Xu through the purification of single-walled carbon nanotubes in 2004.¹ As the new member of the carbon nanomaterials family, CDs are comparable or even alternative materials to traditional semiconductor quantum dots (QDs), such as CdSe, PbS, and CdS, *etc.*, which are widely used in bioimaging and biosensing,^{2–7} patterning,^{8,9} coding,¹⁰ and optoelectronics^{11–13} for their water solubility, good stability, low toxicity, superior resistance to photobleaching, and excellent biocompatibility.^{14,15} The wide application of CDs mainly comes from their luminescent properties. Although the PL mechanism of CDs is controversial, more and more evidence shows that the (photoluminescence) PL emission of CDs is due to the synergistic effect of the non-intrinsic PL of the CDs core and surface state. In addition, a large number of studies have shown that heteroatom doping can significantly tune their electronic properties, surface and local chemistry, as well as

extend their applications. Currently, heteroatoms like nitrogen, phosphorus, sulphur, boron and fluorine have been confirmed to improve the PL quantum yield and PL intensity of the carbon dots. The doping of the heteroatom can introduce holes or electrons to the carbon dots, which distinctly improve the electronic structure and chemical activity, thus improving the optical and electrical properties for a more extensive application. Due to the N atom having a similar atomic radius to the C atom and the outer five valence electrons to bind carbon atoms, doping with nitrogen is widely used.¹⁶ For example, Wang *et al.* reported a kind of N-CDs synthesised by the hydrothermal method from streptomycin, which was used as fluorescent probes in cell imaging.¹⁷ Lu *et al.* used a mixture of oxalic acid and urea as a precursor to synthesize blue fluorescent N-doped carbon dots, which were used to detect Fe³⁺ and Ag⁺.¹⁸ Dong *et al.* used a mixture of citric and branched polyethyleneimines as a precursor for the one-step carbonization synthesis of carbon dots and used them for chemical sensing.¹⁹ However, the low-fluorescence quantum yield and self-quenching of solid-state CDs can be ascribed to excessive resonance energy transfer (RET) or the direct π–π interactions of organic molecules, which seriously limit the development of CDs.²⁰

To our knowledge, there are few reports about the one-step preparation of CDs with solid-state luminescence.^{21,22} What's more, reports that mentioned the solid-state fluorescence (SSF) of CDs usually require the recombination of CDs and a solid matrix such as a polymer matrix to obtain the SSF. However, CDs as polymer materials generally have a fixed shape, which are inconvenient to apply in many fields compared with

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phosphors. Therefore, there is an urgent need to further research CDs-based phosphors with self-quenching-resistant property using environmentally friendly and widely available raw materials through a simple, low-cost, convenient method.

As we all know, Fe^{3+} ions, as one of the most essential metal ions in biological systems, plays a crucial role in the human body and the environment.²³ However, excess amounts of Fe^{3+} ions in the human body and environment may cause some diseases and serious environmental pollution. In the past few years, many chemosensors based on organic dyes,²⁴ semiconductor quantum dots,²⁵ metal nanoparticles,²⁶ and fluorescent metal organic frameworks²⁷ have been widely used as a sensing probe of Fe^{3+} . However, most chemosensors are poisonous, biologically incompatible and water-insoluble, which greatly limited their application in detection of Fe^{3+} . So, the development of a new fluorescent Fe^{3+} sensing probe with photostability, high sensitivity and selectivity that permits naked-eye detection is truly meaningful. Carbon dots (CDs) have the advantages of water-solubility, low toxicity, good biocompatibility, and chemical stability.^{28,29} Additionally, Fe^{3+} could induce the photoluminescence quenching of carbon dots by the special coordinate interaction between Fe^{3+} and the surface groups of carbon dots. These phenomena enabled carbon dots to act as a nanosensor for Fe^{3+} detection.

In this paper, we prepared blue-emission N-CDs with a fluorescent quantum yield of 30.21% *via* a one-step hydrothermal method using citric acid and polyvinyl pyrrolidone (PVP K-30) as the carbon source and the nitrogen source. Citric acid and polyvinyl pyrrolidone (PVP K-30) are low-cost, widely available and environmentally friendly materials with excellent performance. Compared with previous reports,^{4,30} our method avoids using complicated raw materials and surface passivation, which is convenient for industrial production. During the experimental process, we found that the as-prepared CDs show great potential in the field of the detection of Fe^{3+} ions. There was a linear correlation between fluorescent intensity and the concentration of Fe^{3+} with a linear range from 0–300 μM and a detection limit of 45.5 nmol L^{-1} , which indicates that the detection sensitivity in our work is higher than the limit of detection (LOD) data reported before.^{31,32} Furthermore, the as-prepared CDs demonstrate excellent blue fluorescence in the solid state after freeze-drying to a powder, which provides more possibilities for the application of CDs in ion detection and solid-state lighting devices. Additionally, the CDs can also be used as a fluorescent ink and are presented under UV excitation.

Experimental

Chemicals and materials

Citric acid (CA), polyvinyl pyrrolidone (PVP K-30), acetone, $\text{Cu}(\text{NO}_3)_2$, $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$, CaCl_2 , ZnCl_2 , $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$, $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$, NaCl , MgCl_2 , $\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$, AgNO_3 , $\text{Pb}(\text{NO}_3)_2$, KNO_3 , $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$, $\text{Cd}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$, and $\text{Hg}(\text{NO}_3)_2 \cdot \text{H}_2\text{O}$ were all purchased from Sinopharm Chemical Reagent Co., Ltd. (Shanghai, China). Quinine sulfate was purchased from Aladdin Chemistry Co., Ltd. All the chemicals were of analytical grade and were used as received without any further purification.

High-purity deionized water with a resistivity of greater than 18 $\text{M}\Omega \text{ cm}^{-1}$ was obtained from the Millipore system and was used throughout the experiments.

Preparation of CDs

The CDs were prepared *via* a one-step hydrothermal method, 1.5 g of citric acid (CA) and 0.45 g of polyvinyl pyrrolidone (PVP K-30) were dissolved in 15 ml of deionized water to form a transparent viscous solution. Then, the mixture solution was transferred to a Teflon-lined autoclave (20 ml) and heated at 180 $^\circ\text{C}$ for 6 h. After cooling to room temperature gradually, the obtained reactor was precipitated with the addition of excess acetone. The suspension was then centrifuged twice at 8000 rpm for 20 min; a yellow solution was obtained after removing the insoluble substances. After being kept stationary for half a day to evaporate the acetone, the solution was put in a freezing chamber of the fridge and was then freeze-dried into a powder in the freeze drier. Finally, the as-prepared pale yellow CDs powder was obtained, and an aqueous solution of 3 mg ml^{-1} was also prepared for subsequent uses.

Measurements

The fluorescence spectra (PL) were obtained using an RF-5301PC fluorescence spectrophotometer (Shimadzu, Japan). Transmission electron microscopy (TEM) images were obtained on a Hitachi-7650 transmission electron microscope (Hitachi, Japan). X-ray photoelectron spectroscopic (XPS) analyses were performed on a Wscalab X-ray photoelectron spectrometer (ThermoFisher, UK). Fourier transform infrared spectra (FTIR) were recorded on a Nicolet iS10 FTIR spectrometer (Thermo Scientific, USA). Ultraviolet-visible (UV-vis) absorption spectra were performed on a UV-3600 UV-vis spectrophotometer (Shimadzu, Japan).

Fluorescent quantum yield (QY) measurement

The QY of the as-prepared CDs was measured by the slope method as done in previous research.³³ A solution of quinine sulfate in 0.1 mol L^{-1} H_2SO_4 (QY of 54% at 360 nm, $\eta = 1.33$) was used as the standard. The value of the quantum yield was calculated according to the following equation:

$$Q_x = Q_{\text{st}}(K_x/K_{\text{st}})(\eta_x/\eta_{\text{st}})^2$$

where Q is QY, K is the slope, η is the refractive index with 1.33 as the default for both quinine sulfate and CDs. The subscript 'st' refers to the quinine sulfate, while the subscript 'x' refers to the samples.

Metal ion detection

The metal ion detection procedures were performed in an aqueous solution of the as-prepared CDs at RT. 50 ml of CDs solution was added to a colorimetric tube, a blank sample was set up, and the other samples were added into salt solutions (4 mM) separately containing different metal ions of Cu^{2+} , Mn^{2+} , Ca^{2+} , Zn^{2+} , Fe^{2+} , Fe^{3+} , Na^+ , Mg^{2+} , Ni^{2+} , Ag^+ , Pb^{2+} , K^+ , Co^{2+} , Cd^{2+} , Hg^{2+} with full stirring. After stirring for 5 minutes, the



fluorescence intensity was recorded at an excitation wavelength of 360 nm. The slit for the excitation was 1.5 nm, and the emission was 1.5 nm. All the experimental procedures were carried out under the same conditions. The detection of Fe^{3+} was examined using the same procedure by adding different concentrations of Fe^{3+} into aqueous solutions of CDs.

Fluorescent ink

A filter paper upon which the CDs adhered well (the paper showed no background under UV fluorescence) was chosen as the printing paper. Colorless CDs aqueous solutions were dipped with a brush pen to write different graphic patterns on the filter paper.

Results and discussion

Composition and structure of CDs

In this work, the CDs were synthesized by a hydrothermal reaction with the raw materials of citric and polyvinyl pyrrolidone at 180 °C for 6 hours. Under the hydrothermal environment of high temperature and high pressure, the water evaporation spurs the intermolecular dehydration of citric acid

and polyvinyl pyrrolidone molecules to form “polymer-like” intermediates.³⁴ The intermediates were further carbonized and capped by polyvinyl pyrrolidone molecules simultaneously to form the scattered carbon dots.

The morphology and microstructure of the as-prepared CDs were identified by TEM. As shown in Fig. 1a and b, the obtained CDs are spherical in structure, with a uniform dispersion without apparent aggregation. In the high-resolution TEM spectrum (Fig. 1c), no obvious lattice fringes were observed, indicating that the carbon quantum dots have poor crystallinity. Fig. 1d shows the size distribution histograms of the CDs, the distribution of the diameter size of the CDs within the range of 3.1–3.9 nm, and the average particle diameter size is 3.4 nm.

The surface functional groups and chemical composition of the CDs were identified by FTIR spectra and X-ray photoelectron spectroscopy (XPS). The FTIR spectra of the as-prepared CDs and its precursors are shown in Fig. 2, the wide band within the range of 3300 cm^{-1} to 3380 cm^{-1} can be attributed to O–H and N–H stretching vibrations.³⁵ The characteristic peaks at $2970\text{--}2890\text{ cm}^{-1}$ and $1725\text{--}1680\text{ cm}^{-1}$ are the stretching vibrations of C–H and C=C, respectively. The characteristic peaks at $1450\text{--}1380\text{ cm}^{-1}$ and 1070 cm^{-1} were assigned to the bending vibrations of C–N, N–H, COO and the C–O, C–O–C groups,

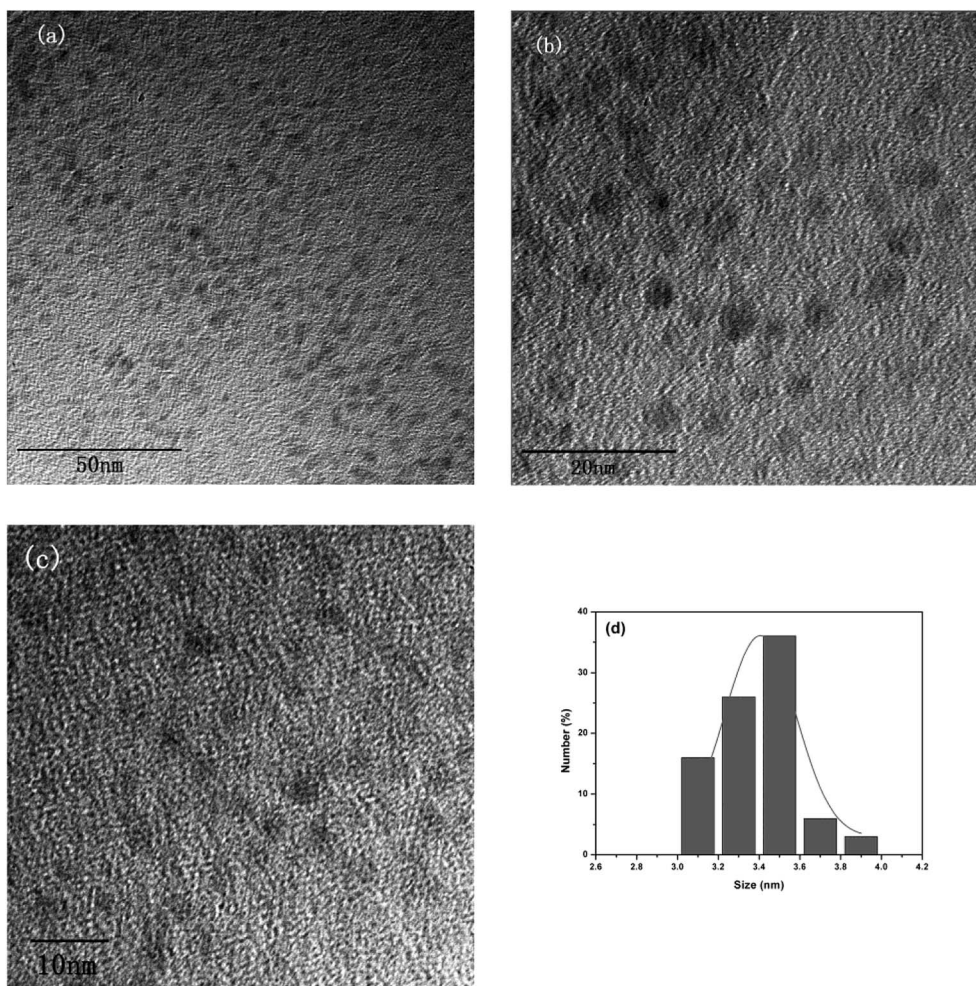


Fig. 1 (a) and (b) TEM images of the CDs; (c) high-resolution TEM images of the CDs; (d) size distribution histograms of the CDs.



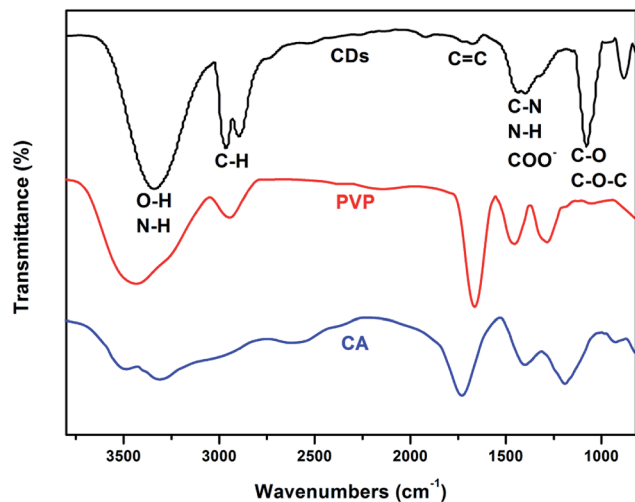


Fig. 2 FTIR spectra of the as-prepared CDs, PVP and citric acid, respectively.

respectively. What's more, compared with the FTIR spectra of the as-prepared CDs and the raw materials of PVP and citric acid (Fig. 4), we can find that the FTIR spectrum of the as-prepared CDs is almost the same as that of PVP, which can be ascribed to how the surface of CDs may be wrapped with abundant PVP long chains. However, the characteristic absorption band at 1639 and 1274 cm^{-1} corresponding to the acid amides and the

CH_2 stretching vibration of PVP are absent in the FTIR spectrum of CDs, which suggests that the PVP units in CDs were changed. Similarly, the characteristic absorption peaks of citric acid also changed in the CDs, which further confirmed that a chemical reaction took place between the reactants, not just a simple physical interaction. The synthesis process of the CDs is schematically illustrated in Scheme 1. There are abundant active groups like carboxyl and hydroxyl on the surface of the citric acid, which are beneficial for improved water-solubility and also provide more reactivity sites. The polyvinyl pyrrolidone introduces different surface functional groups that would create new surface electronic energy levels and various characteristics. Meanwhile, the polyvinyl pyrrolidone provides abundant nitrogen-containing functional groups, which increases the fluorescence quantum yield (QY) and modifies the photoelectric properties.³⁶

The surface composition of the synthesized CDs was examined with X-ray photoelectron spectroscopy (XPS) spectra, which are shown in Fig. 3. The XPS spectrum includes three peaks at 533.1, 400.1 and 285.1 eV,^{37,38} which correspond to C1s, N1s and O1s, respectively (Fig. 3a). The sample contains 69.88 at% of carbon, 7.80 at% of nitrogen and 22.32 at% of oxygen. The corresponding binding energies were given in Fig. 3a, which indicated the existence of large amounts of carbon and oxygen elements and small amounts of nitrogen elements on the surface of the CDs. The appearance of the N1s peak indicates

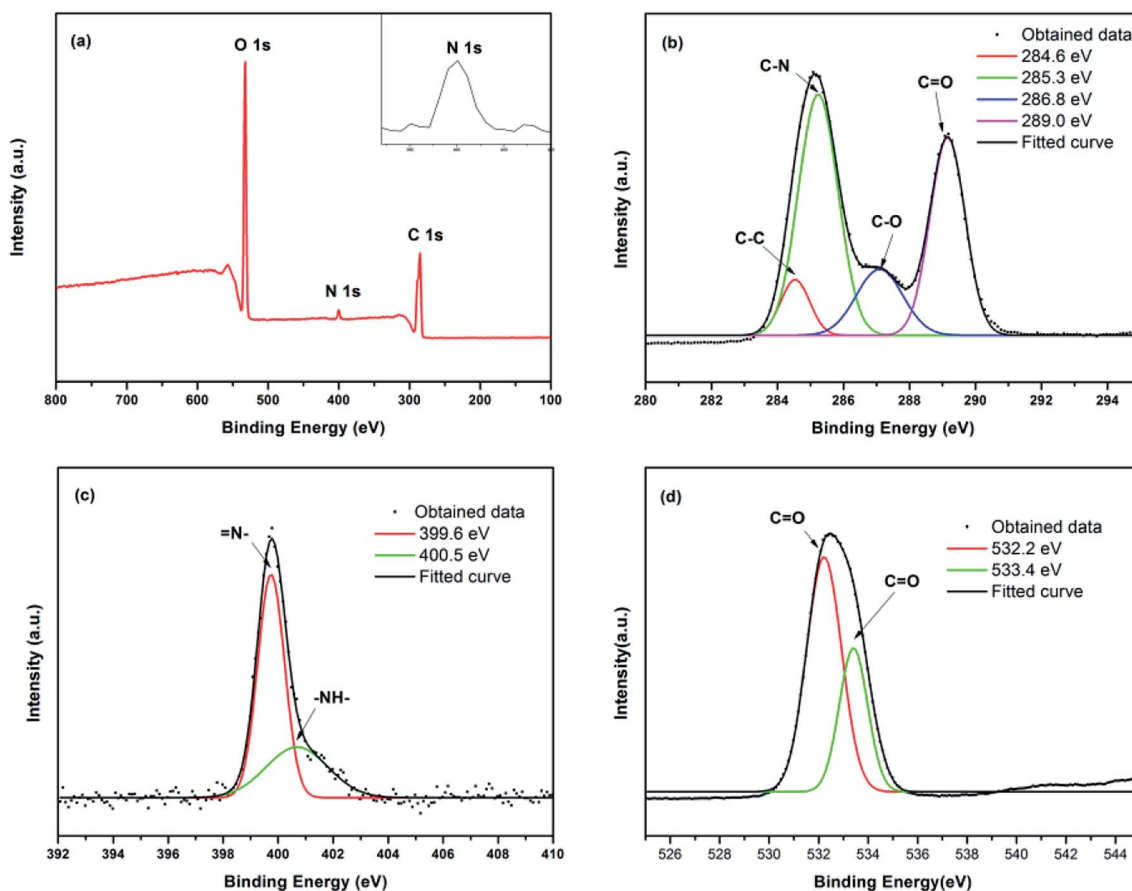


Fig. 3 (a) XPS survey spectrum, magnification and fitting curves of (b) the C1s peak, (c) the N1s peak, and (d) the O1s peak for the prepared CDs.



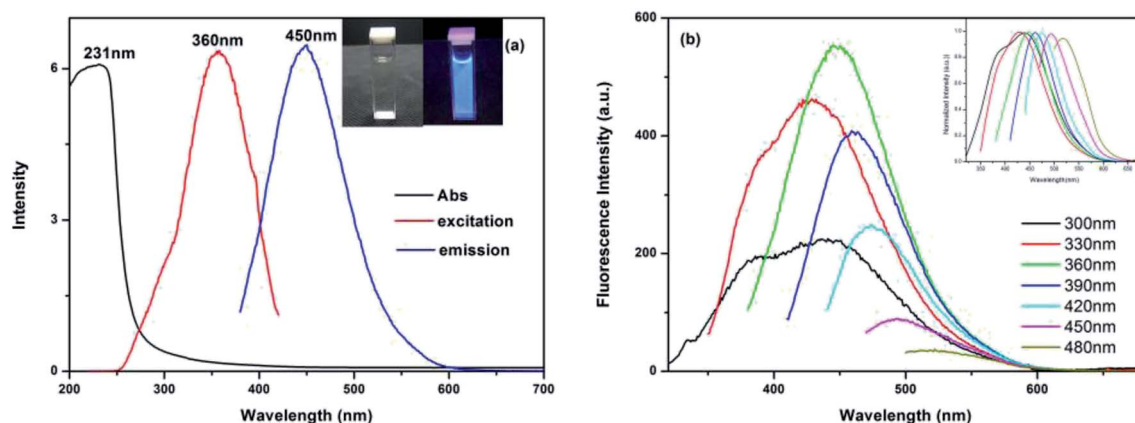


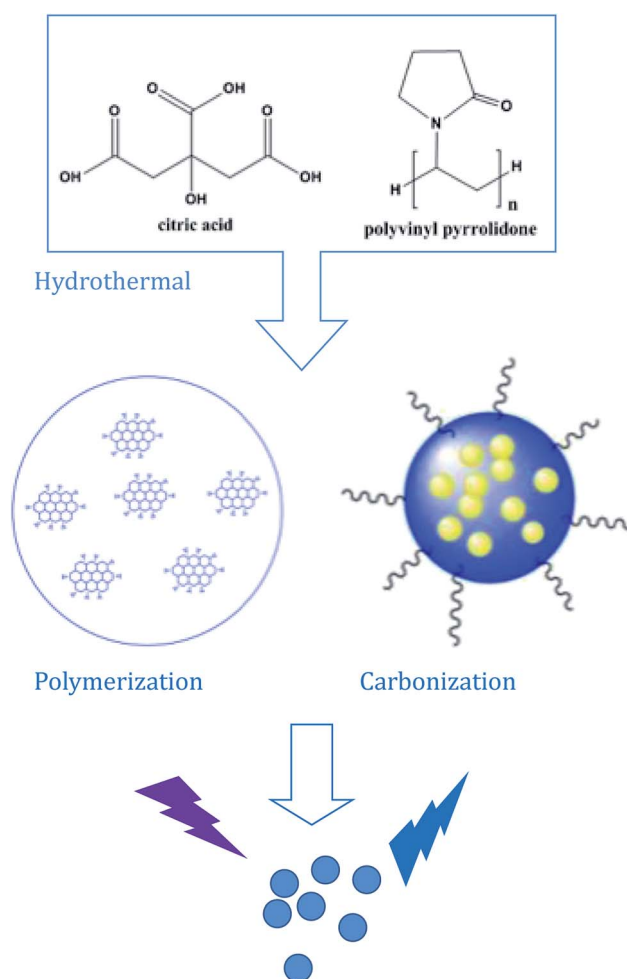
Fig. 4 (a) Absorbance, excitation and emission spectra of the CDs aqueous solution (the inset displays the fluorescent photo under a 365 nm UV lamp); (b) emission spectra of the CDs aqueous solution at different excitation wavelengths from 200 to 480 nm (the inset displays the corresponding normalized spectra).

that the N elements had successfully entered the carbon skeleton. Fig. 3b shows the C1s spectrum of CDs; the four peaks at 284.6 eV, 285.3 eV, 286.8 eV and 289.0 eV should be attributed to C–C, C–N, C–O and C=O.³⁹ The high-resolution N1s spectra (Fig. 3c) of the obtained CDs present two kinds of main peaks associated with nitrogen atoms located at 399.6 eV and 400.5 eV, which correspond to the pyridinic N of =N– and the pyrrolic N of –NH–, respectively. As previously reported, the pyrrolic N exposes the edge of the plane and the defect position, and the pyridinic N replace C atoms at the edge or plane of the defective parts.^{40,41} The defects and planes caused by N doping can significantly improve the surface chemical activity and electrical properties of CDs and then improve the QY of the CDs. The O1s spectrum of the CDs is shown in Fig. 3d. The two main peaks at 532.2 eV and 533.4 eV could be assigned to C–O and C=O. The XPS results agreed with FTIR analysis. In other words, there may be hydrophilic groups such as –OH, –NH₂ and –COOH on the surface of the obtained CDs.

Optical properties of CDs

We used the slope method to measure the QY of the as-prepared CDs. The relative fluorescence QY of the CDs was calculated to be 30.21% using quinine sulfate in 0.1 mol L^{−1} H₂SO₄ (QY of 54% at 360 nm, $\eta = 1.33$) as the reference. The UV-vis absorption spectrum is shown in Fig. 4a. The absorption band ranged from 200 to 280 nm with a peak at 231 nm, which can be assigned to the π – π^* transition of C=C sp². The obtained CDs have the strongest blue emission band at 450 nm with an excitation at 360 nm (the inset picture of Fig. 4a). The PL emission wavelength of the CDs aqueous solution has a typical excitation-dependent property upon changing the excited wavelength as shown in Fig. 4b, which is common in fluorescent carbon materials.⁴² As the excitation wavelength increases from 300 to 480 nm, the position of the emission peak shifts to long wavelengths from 425 to 520 nm. Meanwhile, the PL intensity of the CDs reaches a maximum at 450 nm when excited at 360 nm. The red shift of the emission band may be attributed to the surface state of the CDs that affect the band

gap⁴³ or the effects of different emissive sites on the different sized nanoparticles within the sample.⁴⁴ Nevertheless, the PL intensity reduces as the half-peak width increases, which may be associated with the quantum effect and surface energy trap.⁴⁵



Scheme 1 A synthetic route using citric acid and polyvinyl pyrrolidone (PVP K-30): from ionization to condensation, polymerization, and carbonization.



Furthermore, the quantum yield of the obtained CDs is as high as 32.1% (excited at 360 nm and using quinine sulfate as a reference), which is higher than in many previous reports.^{46,47}

As we all know, the fluorescence stability of CDs is a key factor in its application in various fields. Therefore, the PL stability of the CDs was also studied (Fig. 5). As shown in Fig. 5a, after irradiation under a UV lamp (365 nm) for 1800 s, the PL intensity of the CDs aqueous solution did not obviously change, which indicates that the CDs has good PL stability and have a potential application in biological detection, which may be needed for long-term radiation. What's more, in order to investigate the effect of pH on the PL intensity of CDs, we characterized the PL intensity of CDs aqueous solution with a pH value of 1, 2, 3, 4, 5, 7, 9, 11 and 12, respectively. As shown in Fig. 5b, with the increase of pH value, the PL intensity first increases and then decreases. That is, the PL intensity of the CDs dispersing in the acid medium is stronger than that in the alkaline medium, and the PL intensity reaches a maximum near the pH value of 4. By adjusting the pH value of the CDs solution, we found that the PL intensity changes with the pH value, which

demonstrates that the as-prepared CDs has pH-dependence. The result is common in the previous reports, which is relevant to the protonation/non-protonation of the surface groups ($-\text{NH}_2$, $-\text{COOH}$, $-\text{C}=\text{O}$).⁴⁸

Applications of CDs in ions detection

The influence of the quenching of different ions on the fluorescence of CDs indicates the as-prepared CDs as a promising fluorescent probe for metal ion detection. The emission spectra of the CDs under 360 nm excitation with different metal ions were recorded after incubation for 10 min, as shown in Fig. 6a. The detection solution included Cu^{2+} , Mn^{2+} , Ca^{2+} , Zn^{2+} , Fe^{2+} , Fe^{3+} , Na^+ , Mg^{2+} , Ni^{2+} , Ag^+ , Pb^{2+} , K^+ , Co^{2+} , Cd^{2+} , and Hg^{2+} ions; the emission spectra of all the metal ions did not significant changed except for the Fe^{3+} ions. Fig. 6b shows the different PL intensity ratios (F/F_0) of the CDs solutions in the presence and absence of various metal ions. From this figure, we can observe that the PL intensity ratio (F/F_0) of Fe^{3+} ions is the lowest and that other metal ions show slight changes of PL intensity, which

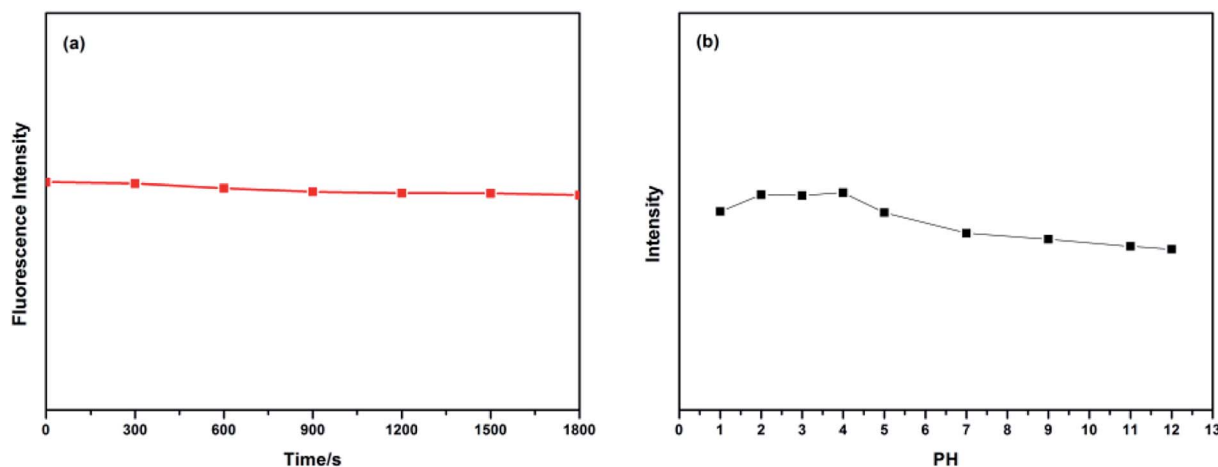


Fig. 5 (a) Irradiation time on the PL intensity of CDs and (b) the value on the PL intensity of CDs ($\lambda_{\text{ex}} = 360$ nm).

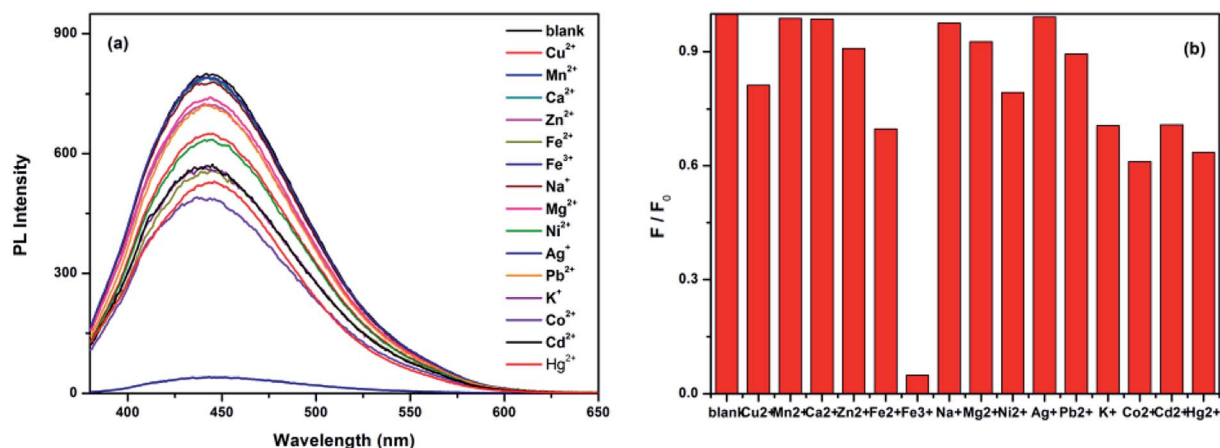


Fig. 6 (a) Emissions spectra of CDs with different ions; (b) the different PL intensity ratios (F/F_0) of the CDs solutions in the presence and absence of various metal ions (F_0 is the blank value, F is the fluorescence intensity of CDs with different ion, excitation = 360 nm, emission = 450 nm, $[\text{Mn}^{2+}] = 4 \text{ mmol L}^{-1}$).



indicated that Fe^{3+} has the largest quenching effect on the fluorescence of CDs and that the quenching effect may be due to the strong affinity of Fe^{3+} ions to CDs. The results of FTIR and XPS confirmed that part of the $-\text{COOH}$, $-\text{OH}$ and $-\text{NH}_2$ groups exists on the surface of the as-prepared CDs, which can react with Fe^{3+} ions to form a complex and effectively prevents the hydrolyzing of Fe^{3+} in solution, finally leading to the fluorescence quenching of CDs. Moreover, variations on the surface of CDs may cause the aggregation of CDs, leading to fluorescence quenching. This phenomenon is agreed with the reported work, which had revealed that after the addition of Fe^{3+} to the CDs solution, the surface of CDs might be changed, which would result in poor luminescence efficiency.^{49,50}

At room temperature, the emission spectra of the CDs in the presence of different concentrations of Fe^{3+} ions were recorded after incubation for 10 min. As shown in Fig. 7a, after excitation at 360 nm, CDs solutions with different concentrations of Fe^{3+} ions show a broad band with a peak at 450 nm, ranging from 380 to 600 nm with an increasing concentration of Fe^{3+} ions. The fluorescence intensity at 450 nm decreases gradually without any red or blue shift in position, which indicates that the quenching effect is related to the concentration of Fe^{3+} ; CDs, therefore, can be used as the probe of Fe^{3+} ions. The Fe^{3+} quenching effect can be described with the Stern–Volmer equation:

$$F_0/F - 1 = K_{\text{sv}}C$$

where K_{sv} is the Stern–Volmer quenching constant and the bimolecular quenching constant, C is the concentration of Fe^{3+} ,

and F_0 and F are the fluorescence intensity of CDs without Fe^{3+} and with different concentration of Fe^{3+} , respectively. Fig. 7b shows the standard curve of F_0/F to the concentration of Fe^{3+} ; the range of linearity is 0–300 μmol , the corresponding regression coefficient $R^2 = 0.991$, and the detection limit (LOD) for Fe^{3+} is 45.5 nmol L^{-1} based on $3\sigma/k$ (σ represents the standard deviation for the corrected blank signals with 10 parallel measurements, and k is the slope of the calibration curve). Moreover, on/off fluorescence was achieved by oxidizing Fe^{2+} to Fe^{3+} in such an imaging system (Fig. 7c and d).

Solid-state fluorescence of CDs

Furthermore, the self-quenching-resistant property of the as-prepared CDs was founded after the freeze-drying process. Excited by UV light (365 nm), the CDs powder exhibits strong blue solid-state fluorescence (SSF), as shown in Fig. 8. The SSF is attributed to the space steric hindrance caused by the PVP chains on the surface of the CDs particles, which was confirmed by FTIR spectroscopy. The long chains prevent the graphitized carbon cores from directly contacting each other, such that keeping a certain distance among the CDs particles allows for the avoidance of the self-quenching effect generated by π – π stacking (one kind of efficient energy transfer mode), which shows as the solid-state fluorescence.²¹

Furthermore, the emission spectra of the CDs powder are shown in Fig. 9. The PL emission wavelength of the CDs powder has an excitation-dependent property when changing the excitation wavelength from 330 to 600 nm, similar to the CDs aqueous solution. As the excitation wavelength increases, the

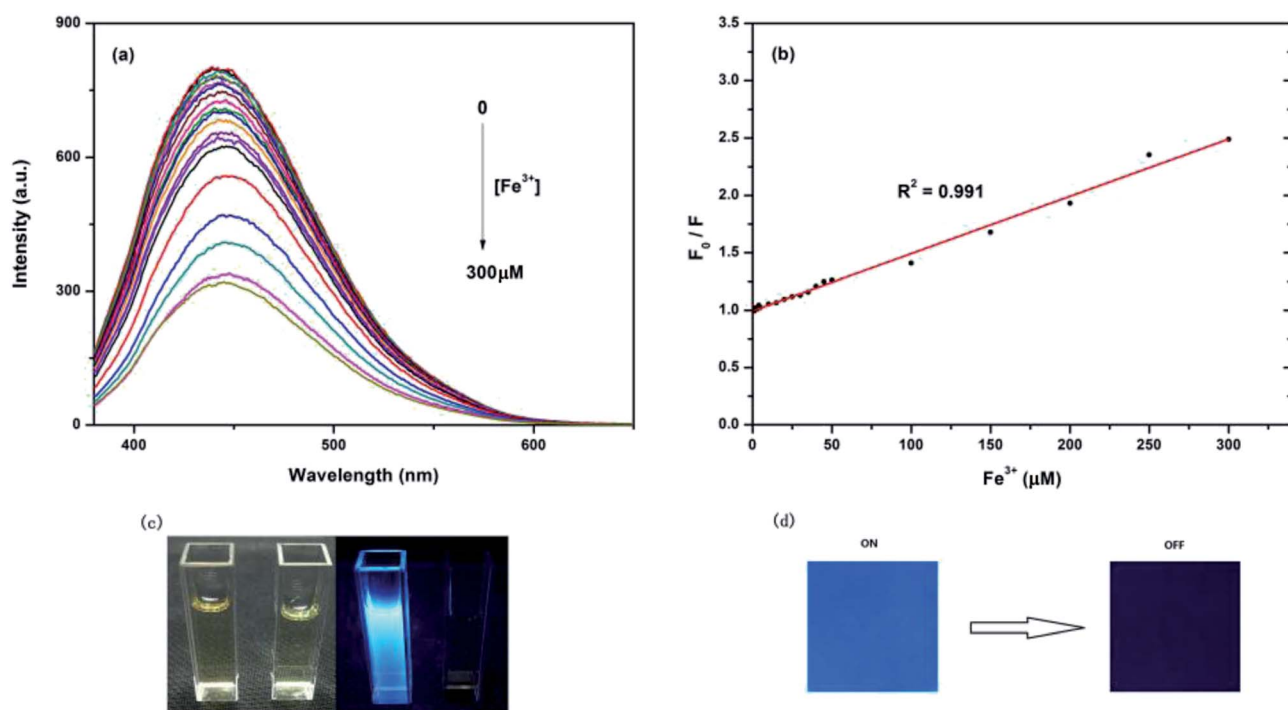


Fig. 7 (a) Quenching curves of CDs upon exposure to different concentration of Fe^{3+} ; (b) the linear region of the CDs- Fe^{3+} sensing response ($[\text{Fe}^{3+}]$: 0.0, 1.0, 2.0, 3.0, 4.0, 5.0, 10, 15, 20, 25, 30, 35, 40, 45, 50, 100, 150, 200, 250 and 300 $\mu\text{mol L}^{-1}$, respectively); (c) images of the CDs before (left) and after (right) treatment with Fe^{3+} under sunlight and UV light, respectively; (d) fluorescence images of aqueous CDs after adding Fe^{2+} and Fe^{3+} ions; the on/off fluorescence was achieved by oxidation.



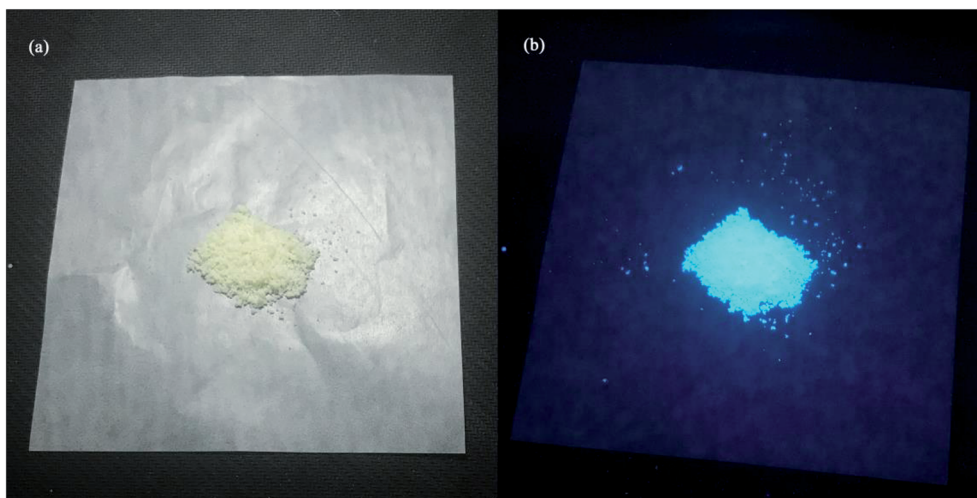


Fig. 8 Photographs of CDs were taken under (a) daylight and (b) 365 nm UV light.

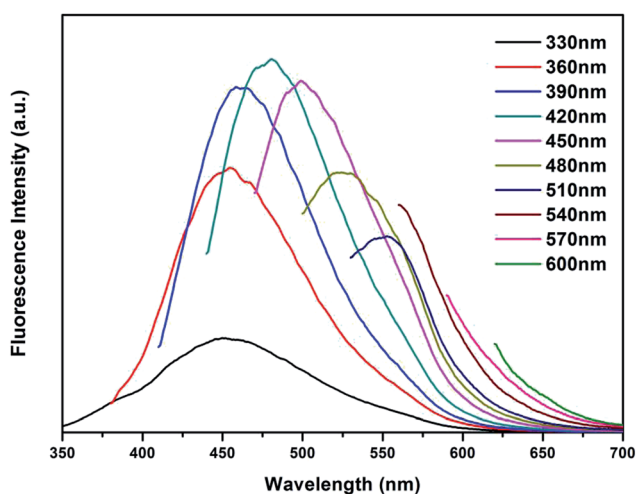


Fig. 9 SSF spectra of CDs powder at different excitation wavelengths from 330 to 600 nm.

position of the emission peak of the CDs powder shifts to longer wavelengths from 451 to 600 nm, while the PL intensity of the CDs reaches the maximum at 482 nm when excited at 420 nm. The self-quenching-resistance property of the CDs powder may provide an opportunity for the achievement of efficient SSF nanomaterials without being supported by solid matrices.

Application of CDs in fluorescent ink

For its excellent fluorescence, photobleaching resistance, good PL stability and good dispersibility, aqueous CDs can be used as fluorescent ink. The fluorescent characters written on paper by using CDs ink are not visible, except when excited under a UV lamp. As shown in Fig. 10a, “Shanghai University” CDs blue fluorescent characters can be observed under a UV light. A month later, the characters still emit a clear and intense blue light under 365 nm UV excitation. When the concentration of the CDs in aqueous solution changed, the CDs ink showed different shades of blue under the UV lamp (Fig. 10b). High-

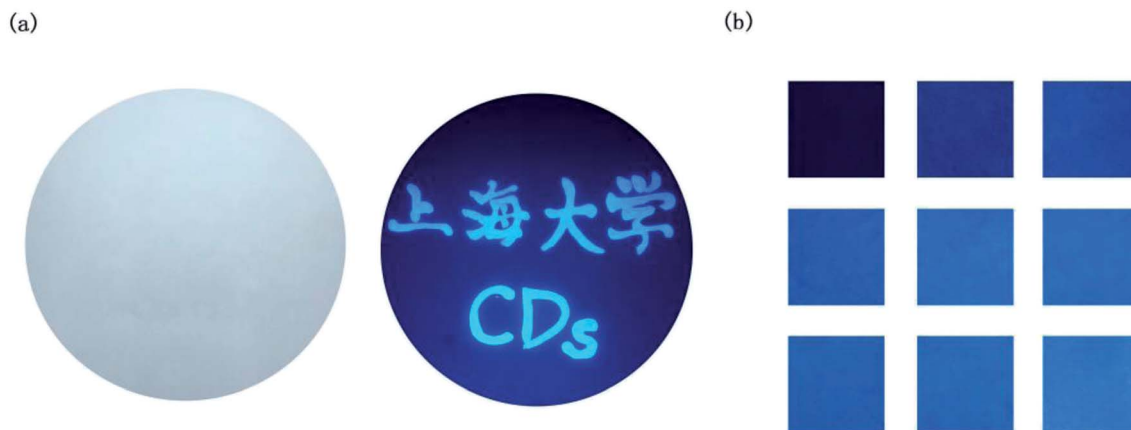


Fig. 10 (a) Different graphic patterns on paper (illuminated by the 365 nm UV lamp) and (b) ink in multiple colors, tuned by the CDs concentration in aqueous solution (illuminated by the 365 nm UV lamp).



quality fluorescent graphics and text characters can be obtained by using CDs fluorescent ink, which can replace traditional ink for anti-counterfeiting, information encryption, information storage and so on. In the near future, the CDs fluorescent ink may have great application in flexible displays.

Conclusion

In summary, the blue CDs were successfully synthesized by a simple one-pot hydrothermal treatment by using citric acid (carbon source) and polyvinyl pyrrolidone (PVP K-30) (nitrogen source) as the raw materials. The QY of the as-prepared CDs is 30.21%, and it has remarkable water-dispersibility and excellent fluorescence stability. The obtained CDs have the strongest blue emission band at 450 nm with an excitation at 360 nm. The emission spectra of the obtained CDs exhibited an excitation-dependent fluorescence behavior with a center of broad range shifts from 425 to 520 nm. The CDs could be used to detect the Fe^{3+} ions within a linear range of 0–300 μM , with a detection limit of 45.5 nmol L^{-1} , which suggested that the new CDs has great potential toward the detection of Fe^{3+} ions. What's more, the excellent self-quenching-resistant property of the as-prepared CDs was showed in the solid state, with bright blue fluorescence being observed under a 365 nm UV lamp, providing more possibilities for the solid application of CDs. Besides, the CDs could be applied as a fluorescent ink that could potentially be a substitute for commercial dyes. All of the results suggest the as-prepared CDs as a new class of luminescent carbon nanomaterials with great potential in environmental and commercial applications in the near future.

Conflicts of interest

There are no conflicts to declare.

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