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Cu@Co-MOFs as a novel catalyst of peroxymonosulfate for the efficient removal of methylene blue

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In this study, for the first time, we describe the single step synthesis of a Cu particle-doped cobalt-based metal-organic framework (Cu@Co-MOF) using a hydrothermal method. The as-prepared materials were characterized by powder X-ray diffraction, Fourier transform infrared spectroscopy, scanning electron microscopy-energy disperse spectroscopy, thermogravimetry, and X-ray photoelectron spectroscopy, which confirmed the incorporation of zero-valent copper on the surface of the Co-MOFs. The heterogeneous catalytic activity of Cu@Co-MOFs was tested to activate peroxymonosulfate (PMS) for the removal of methylene blue (MB). The ratio of $n(\text{Cu})/n(\text{Co})$ in the Cu@Co-MOFs showed a strong impact on the catalytic activity of the Cu@Co-MOFs, whereas a $n(\text{Cu})/n(\text{Co})$ of 1 : 1 exhibited the best catalytic performance and obtained 100% MB removal within 30 min. The effects of initial pH, reaction temperature, PMS concentration, and catalyst dosages were investigated in this study. The stability and reusability of the Cu@Co-MOFs were also investigated. The results showed a low decline in the MB removal with the increase in cycle numbers, whereas 100% MB was removed by prolonging the reaction time. Heterogeneous reactions taking place in the pores and surface of the Cu@Co-MOFs played an important role in the generation of the sulfate radicals ($\text{SO}_4^{\cdot-}$) and hydroxyl radicals ($\cdot\text{OH}$) that were the primary reactive species responsible for MB degradation.

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1. Introduction

Dye wastewater has become one of the great focus points of investigation due to the considerable usage of dyes that considerably pollute the environment and cause severe consequences. At present, more than 100 000 commercial dyes exist and over 1 million tons of dyes are manufactured each year,¹ which are toxic and harmful to the environment and humans.² Therefore, it is of vital importance to treat the water polluted with organic dyes before it is released into the environment. Many technologies such as physical adsorption,^{3,4} microbial degradation,⁵ chemical oxidation,^{6,7} ultrasonication and membrane methods have been used for the removal of such pollutants. However, the organic dyes are just transported from one media to another by activated carbon or other porous materials *via* adsorption, and they are therefore not completely eliminated from the environment. The microbial degradation process is time-consuming and inefficient because most organic dyes are non-biodegradable due to their complex structure and stability.⁸ For the ultrasonic and

membrane methods, the results are inconsistent and these methods are still at preliminary stages;⁹ in addition, these methods are expensive.

Sulfate radical ($\text{SO}_4^{\cdot-}$)-based advanced oxidation processes (SR-AOPs) are known to be the most effective method to oxidize the bio-recalcitrant organic pollutants.¹⁰⁻¹² SR-AOPs are a kind of technology for the *in situ* generation of the $\text{SO}_4^{\cdot-}$ radical. Moreover, other reactive free radicals with high oxidation potential such as hydroxyl radicals ($\cdot\text{OH}$) and superoxide radicals ($\cdot\text{O}_2^-$) could be produced by changing the reaction conditions.¹³⁻¹⁵ Peroxydisulfate (PDS, $\text{S}_2\text{O}_8^{2-}$) and peroxymonosulfate (PMS, HSO_5^-) are the common precursors of the $\text{SO}_4^{\cdot-}$ radical, but they are very stable at room temperature and unable to degrade organic pollutants. Therefore, extra energy and/or catalysts are needed to activate them for the generation of free radicals.^{16,17} Compared to the other transition metal ions, Fe^{2+} is the most efficient transition metal for the activation of PDS, while Co^{2+} is the best catalyst for the decomposition of PMS to generate $\text{SO}_4^{\cdot-}$ radicals.¹³ In addition, $\text{Ag}(i)$ also shows the best catalytic performance toward activating PDS.¹³ However, Fe^{2+} would react with $\text{SO}_4^{\cdot-}$ and result in a decline in the pollutant removal efficiency.¹⁸ The widespread use of Co^{2+} may not be conducive to the human health and the quality of the environment.¹⁹ As for $\text{Ag}(i)$, the cost is very high and it may be unrealistic to use in industry. Thus, alternative catalysts with high efficiency are urgently needed.

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Metal–organic frameworks (MOFs), a new kind of material with a periodic structure, are constructed from inorganic metal nodes and organic ligands with a topological structure.^{20–22} The promising characteristics of highly ordered structure, large specific surface areas and tunable structural features render MOFs suitable for its applications in many fields, including separation,²³ gas storage,²⁴ drug delivery²⁵ and catalysis.²⁶ As the “nodes” of the MOFs, metal ions are expected to be potential heterogeneous activators of PDS or PMS. Gao *et al.*²⁷ investigated the catalytic performance of MIL-53 as heterogeneous photocatalysts of PDS under visible LED light, their results demonstrated that the accelerated photocatalytic degradation of organic pollutants was obtained in the presence of PDS. Fe-based metal–organic frameworks including MIL-101(Fe), MIL-100(Fe), MIL-53(Fe), and MIL-88B(Fe) were examined as the activators of PDS to remove acid orange 7 (AO7) from an aqueous solution, but the degradation efficiency was low because of the weak reaction activity between Fe(III) and PDS.²⁸ Apparently, many MOFs materials can only be used in light photocatalysis because of their wide bandgap.^{7,29,30} In our previous study, the catalytic activity of the Co-based MOFs (Co₃(BTC)₂·12H₂O) was investigated as the activator of PMS for the degradation of dibutyl phthalate (DBP) in water, and it was found that the water stability of Co₃(BTC)₂·12H₂O needed to improve.³¹ Another MOFs, Cu-BTC (Cu₃(BTC)₂·12H₂O) was stable in water, but its catalytic performance for the decomposition of PDS was weak.³² Therefore, it is a great challenge to increase the water stability and catalytic performance of MOFs for further enhancement of organic pollutants degradation.

Inspired by the above results, we tried to synthesize a bimetallic-organic framework combination with Cu and Co, with the purpose of improving the water stability and catalytic performance of MOFs. In this study, we expected that the addition of Cu²⁺ and Co²⁺ into the reaction vessel would result in a bimetallic Cu–Co-MOF with better catalytic performances. However, surprisingly, we observed very few Cu²⁺ ions replace the Co²⁺ ions in the Co-MOF structure, most of them translated to zero-valent copper and grew on the surface of Co-MOF. The catalytic activity of the as-prepared materials (denoted as Cu@Co-MOFs) as the catalysts of PMS for the generation of SO₄^{•−} radicals was tested. In addition, cyclic tests for the pollutant degradation efficiency were also conducted to evaluate the materials' stability. The as-prepared Cu@Co-MOFs composites offer an excellent catalytic performance for the MB removal from water, making it a perfect candidate for heterogeneous catalysis.

2. Materials and methods

2.1 Materials

Unless otherwise specified, all the chemicals and reagents were of reagent grade and used without further purification. Terephthalic acid (BDC), cobalt nitrate hexahydrate (Co(NO₃)₂·6H₂O), cupric nitrate trihydrate (Cu(NO₃)₂·3H₂O) and methylene blue (MB, ≥90%, HPLC) were purchased from Maclean reagent Co., Ltd (Shanghai, China). Sodium persulfate (Na₂S₂O₈, 99.0%), potassium monopersulfate triple salt (KHSO₅·0.5KHSO₄·0.5K₂SO₄) (purity > 47% KHSO₅ basis), *N,N*-

dimethyl formamide (DMF) and all other chemicals such as ethanol (EtOH), *tert*-butyl alcohol (TBA), sodium hydroxide (NaOH) and sulfuric acid (H₂SO₄) were purchased from Sino-pharm Chemical Reagent Co., Ltd (Beijing, China).

2.2 Preparation of Cu@Co-MOFs

The Cu@Co-MOFs was prepared *via* a one-step hydrothermal method. Briefly, cupric nitrate trihydrate (6 mM, 1.449 g), cobalt nitrate hexahydrate (6 mM, 1.946 g), and terephthalic acid (12 mM, 1.994 g) were dissolved in 60 mL of *N,N*-dimethyl formamide (DMF). Then, the mixed solution was stirred for 30 min under vigorous magnetic stirring at room temperature, followed by transferring the solution into a Teflon-lined stainless steel autoclave with a volume capacity of 100 mL, and then 0.8 mL of hydrofluoric acid was added to the solution. The resulting mixture was heated up to 150 °C for 72 h in an oven. After the heat treatment, the autoclave was allowed to cool naturally to room temperature, and the collected products were refined through centrifugation and rinsed with DMF (3 × 20 mL) along with ethanol (3 × 20 mL) and deionized water (3 × 50 mL) alternatively, and then dried in vacuum at 100 °C for 10 h.

To assess the effect of molar ratios (including the molar ratio of (Co²⁺ + Cu²⁺): BDC and molar ratio of Co²⁺: Cu²⁺) on the structure and properties of Cu@Co-MOFs catalysts, different samples were prepared by varying the concentration of the two metal ions and BDC, the specific description is presented in Table 1.

2.3 Characterization

The morphology of the as-synthesized Cu@Co-MOFs was investigated using a thermal field emission scanning electron microscope (FESEM, Quanta 200, Holland), which was linked to the EDS/INCA 350 (energy dispersive X-ray analyser). Before testing, the sample surface was gold-coated. Powder X-ray diffraction (PXRD) patterns of the Cu@Co-MOFs were recorded on an X-ray diffractometer D8 Advance X-ray Diffraction system with a scan speed of 2° min^{−1} and a step size of 0.02° at the range of 5–80°. Fourier transform infrared (FTIR) spectra were recorded with a KBr disk containing the powder sample using an FTIR spectrometer (Nicolet Magna 550) in the range of 2000–400 cm^{−1}. Photoelectron spectroscopy (XPS) was performed using an X-ray photoelectron spectroscopy (ESCA) spectrometer with a monochromatized Al-K X-ray source. The Brunauer–Emmett–Teller specific surface area was determined through nitrogen adsorption/desorption experiments conducted using a NOVA 2000e gas adsorption analyzer (Quantachrome Corp.). Prior to the analysis, the samples were degassed at 150 °C for 1 h. The thermal stability of the Cu@Co-MOFs was determined by DSC-TGA analysis (Q1000DSC + LNCS + FACS Q600SDT) with a heating rate of 10 °C min^{−1} in flowing nitrogen gas.

2.4 Evaluation of Cu@Co-MOFs reactivity as a catalyst

Catalytic degradation experiments were performed in a 250 mL Erlenmeyer flask at 120 rpm in a rotary shaker (SHZ-C, Shanghai, China). A certain amount of PMS and MB were



Table 1 Synthetic parameters of the Cu@Co-MOFs samples

Sample	$n(\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O})/\text{mM}$	$n(\text{Cu}(\text{NO}_3)_2 \cdot 3\text{H}_2\text{O})/\text{mM}$	$n(\text{BDC})/\text{mM}$	$\text{Co}^{2+}/\text{Cu}^{2+}$	$(\text{Cu}^{2+} + \text{Co}^{2+})/\text{BDC}$
Cu@Co-MOFs-1	2.4	9.6	12	1 : 4	1 : 1
Cu@Co-MOFs-2	4.8	7.2	12	2 : 3	1 : 1
Cu@Co-MOFs-3	6	6	12	1 : 1	1 : 1
Cu@Co-MOFs-4	7.2	4.8	12	3 : 2	1 : 1
Cu@Co-MOFs-5	9.6	2.4	12	4 : 1	1 : 1
Cu@Co-MOFs-6	6	6	18	1 : 1	2 : 3
Cu@Co-MOFs-7	6	6	18	1 : 1	3 : 2

added into a final volume of 100 mL deionized water, giving initial concentrations of MB and PMS of 0.2 mM and 4.0 mM, respectively. The degradation reaction was initiated after adding the catalyst. Samples were withdrawn at the predetermined time intervals and quenched by ethanol before analysis.¹⁵ The initial pH of the reactions was adjusted by H_2SO_4 (1.0 M) or NaOH (1.0 M). For the successive recycling reaction, the used catalyst was separated by centrifugal separation after the completion of the reaction and washed with water for 4 times. Then, the catalyst was dispersed in a 100 mL solution containing the same concentration of MB and PMS. All experiments were performed in triplicates.

2.5 Analytical methods

The MB concentration during the reaction was analyzed by measuring the absorbance in the visible spectra at the characteristic wavelength of MB ($\lambda_{\text{max}} = 666$ nm) using a UV-Vis spectrophotometer (UV2301 II, Shanghai, China). The concentration of PMS was determined by spectrophotometric determination with iodometric titration.³³ The leaching metal ions (including copper ion and cobalt ion) in reaction solutions and metallic elements in Cu@Co-MOFs structure were measured using an inductively coupled plasma-atomic emission spectrometer (ICP-AES, PE OPTIMA 2100DV). The mineralization of MB was established on the basis of the total organic carbon (TOC) content, performed using a TOC analyzer (Shimadzu) after the sample was quenched by 2 M sodium thiosulfate. The pH was monitored with a pH meter (Shanghai LeiCi PHS-25) equipped with a pH electrode.

3. Results and discussion

3.1 The crystallization and morphological analyses of the Cu@Co-MOFs

X-ray diffraction (XRD) was used for the characterization of the Cu@Co-MOFs-3 structure. As shown in Fig. 1A, the clear and sharp diffraction peaks at $2\theta = 9.0^\circ, 11.0^\circ, 14.1^\circ, 17.8^\circ, 19.2^\circ, 21.0^\circ, 22.0^\circ, 23.9^\circ, 29.5^\circ, 30.8^\circ, 33.1^\circ, 34.5^\circ, 40.3^\circ$ are ascribed to the patterns of the Co-MOFs, indicating a good crystallinity of the as-prepared material. Interestingly, the diffraction peaks at $43.4^\circ, 50.4^\circ$ and 74.1° can be indexed to the indices (200), (220), and (440) of zero-valent copper.³⁴ This indicated that a part of cupric nitrate trihydrate was reduced to zero-valent copper by DMF at high temperature. In order to determine whether the

organic compounds were present, FT-IR spectroscopy was performed to identify the characteristic functional groups associated with Cu@Co-MOFs-3. As displayed in Fig. 1B, the FT-IR spectrum of Cu@Co-MOFs-3 showed characteristic peaks at $1650.1\text{ cm}^{-1}, 1576.7\text{ cm}^{-1}, 1398.1\text{ cm}^{-1}, 815.0\text{ cm}^{-1}, 757.5\text{ cm}^{-1}, 625.4\text{ cm}^{-1},$ and 521.5 cm^{-1} . In detail, the bands located at 1650 cm^{-1} was associated with the stretching vibrations of the $\nu_{\text{as}}(\text{C}=\text{O})$.³¹ The strong band of FTIR at 1398.1 cm^{-1} was ascribed to the asymmetric stretching ($\nu_{\text{as}}(\text{C}-\text{O})$), while another strong band at 1576.7 cm^{-1} was due to the presence of symmetric ($\nu_{\text{s}}(\text{C}-\text{O})$) vibrations, confirming the presence of the dicarboxylate linker within the frameworks.²⁷ The peak at 757.5 cm^{-1} reflected the presence of the C-H bending vibrations of benzene. The adsorption bands at 625.4 cm^{-1} and 521.5 cm^{-1} may be a characteristic of the Co-O stretching vibration in the Cu@Co-MOFs-3. The TGA results indicated that the Cu@Co-MOFs-3 was stable up to over 350°C , indicating that the as-prepared Cu@Co-MOFs-3 exhibited good thermal stability. Besides, the organic bridge of BDC was more stable after combination with copper and cobalt.

To validate the chemical constitution of the as-synthesized Cu@Co-MOFs-3, the Cu@Co-MOFs-3 sample was further studied *via* XPS analysis. A wide survey scan of XPS spectra was taken in the range 0–1400 eV, as shown in Fig. 2A. It can be seen that the Cu@Co-MOFs-3 mainly contained the elements C, O, Co, and Cu. In addition, the peaks of Co and Cu exhibit a similar peak area, which had confirmed that both the metal ions of Co and Cu had successfully participated in the construction of the as-prepared materials. In order to gain insights into the chemical state of Cu and Co in the composites, we had measured the Cu $2p_{3/2}$ and Co $2p_{3/2}$ spectra, which are displayed in Fig. 2B and C. In Fig. 2B, the $2p_{3/2}$ peak of Cu splits into two peaks at 932.5 eV and 935.0 eV after fitting along with the satellite peak at 942.8 eV. The peak located at 932.5 eV was attributed to the zero-valent copper Cu(0),³⁵ revealing that the surface copper was present in the Cu(0) state. The peak located at a binding energy of 935.0 eV was assigned to Cu(II). In addition, the relative area of the Cu(II) spectrum was 35.4%, while the relative area of the Cu(0) spectrum was 64.6%. This result indicated that most of the cupric nitrate trihydrate translated to zero-valent copper and a very small amount of cupric nitrate trihydrate was incorporated into Co-MOFs. As shown in Fig. 2C, after the peak fitting, the peak appeared at a binding energy of 782.0 eV along with a satellite peak (786.9 eV) that was assigned to Co(II).³⁶ In Fig. 2D, the wide and asymmetric O 1s XPS peak



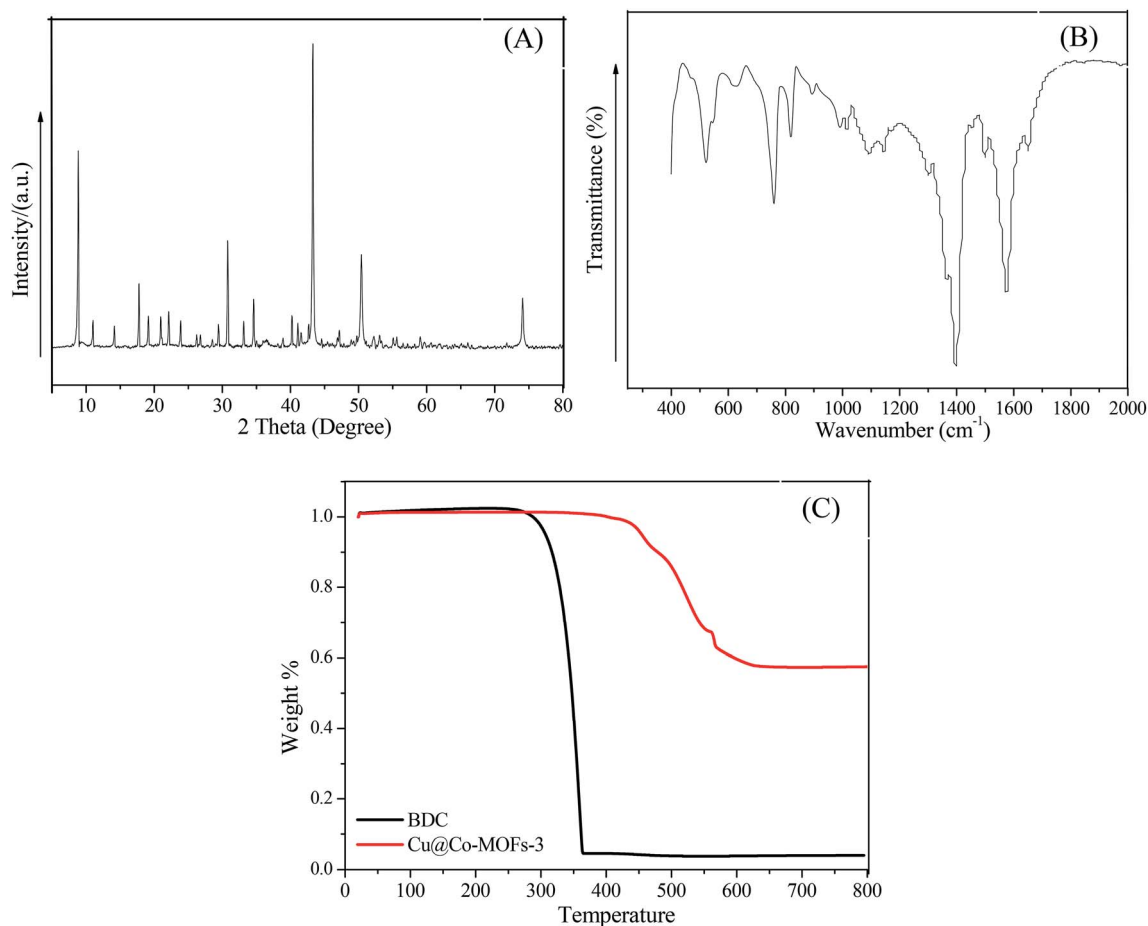


Fig. 1 PXRD pattern (A), FT-IR spectra (B) and TGA (C) of the as-synthesized Cu@Co-MOFs-3.

proved that there were more than one chemical states. The lower binding energy peak at 531.3 eV was ascribed to the Co–O bands.³⁷ The other peak located at 532.2 eV was attributed to the carboxylate groups of the BDC linkers (O–C=O). These results together with XRD and FT-IR analyses indicate the Cu@Co-MOFs-3 was synthesized successfully with a one-step hydrothermal method.

The morphology of Cu@CuCo-MOFs-3 was examined using SEM, as presented in Fig. 3. The as-prepared materials present a regular bounded polyhedron structure with an average size of *ca.* 15–20 μm . Apparently, small particles grew on the surface of the Co-MOFs. In order to confirm the composition of the particles, the as-prepared materials were investigated by EDS elemental mapping analysis. The yellow dots represented the Cu element, and the particles on the surface were confirmed to be copper. Moreover, the other elements Co, C, and O present in the sample are marked with blue, red, and green dots, respectively, and they were uniformly distributed in the as-prepared materials. Using the SEM combined with EDS mapping analysis, it was further confirmed that Cu particles were successfully grown on the surface of Co-MOFs. The permanent porosity of the Cu@Co-MOFs-3 was verified by Brunauer–Emmett–Teller (BET) analysis, which indicated that the surface area and micropore volume were 11.72 $\text{m}^2 \text{g}^{-1}$ and 0.038 $\text{cm}^3 \text{g}^{-1}$, respectively.

Elemental analysis by ICP-MS indicated that copper and cobalt loading of 0.087 mmol g^{-1} and 0.064 mmol g^{-1} , respectively, for Cu@Co-MOFs-3, indicating that both copper and cobalt coexisted in the as-prepared materials. The atomic ratios of Cu/Co in the Cu@Co-MOFs-3 were found to be 1.36, which was 1.0 in the as-synthesized reactions.

3.2 Effects of preparation condition

Fig. 4 presents the catalytic activity of Cu@Co-MOFs obtained from different preparation conditions. It was clear that all catalysts showed an excellent catalytic performance for the decomposition of PMS, resulting in more than 70% of the MB removal rate. However, for PDS, the catalytic activity of the catalysts was much lower. In Fig. 4A, it was remarkable that the catalyst (Cu@Co-MOFs-6) showed the best catalytic performance on PDS when the molar ratio of the [metal ions]/[BDC] was set to 2/3 in the preparation procedure. However, the catalyst obtained with a molar ratio of 1/1 (Cu@Co-MOFs-3) showed the highest catalytic activity on PMS and resulted in the 100% removal of MB, as shown in Fig. 4B. Apparently, the MB removal rate seemed to be well related to the $n(\text{Co})/n(\text{Cu})$ ratios of the materials in the preparation procedure. With the presence of PDS, the MB removal decreased with an increase in



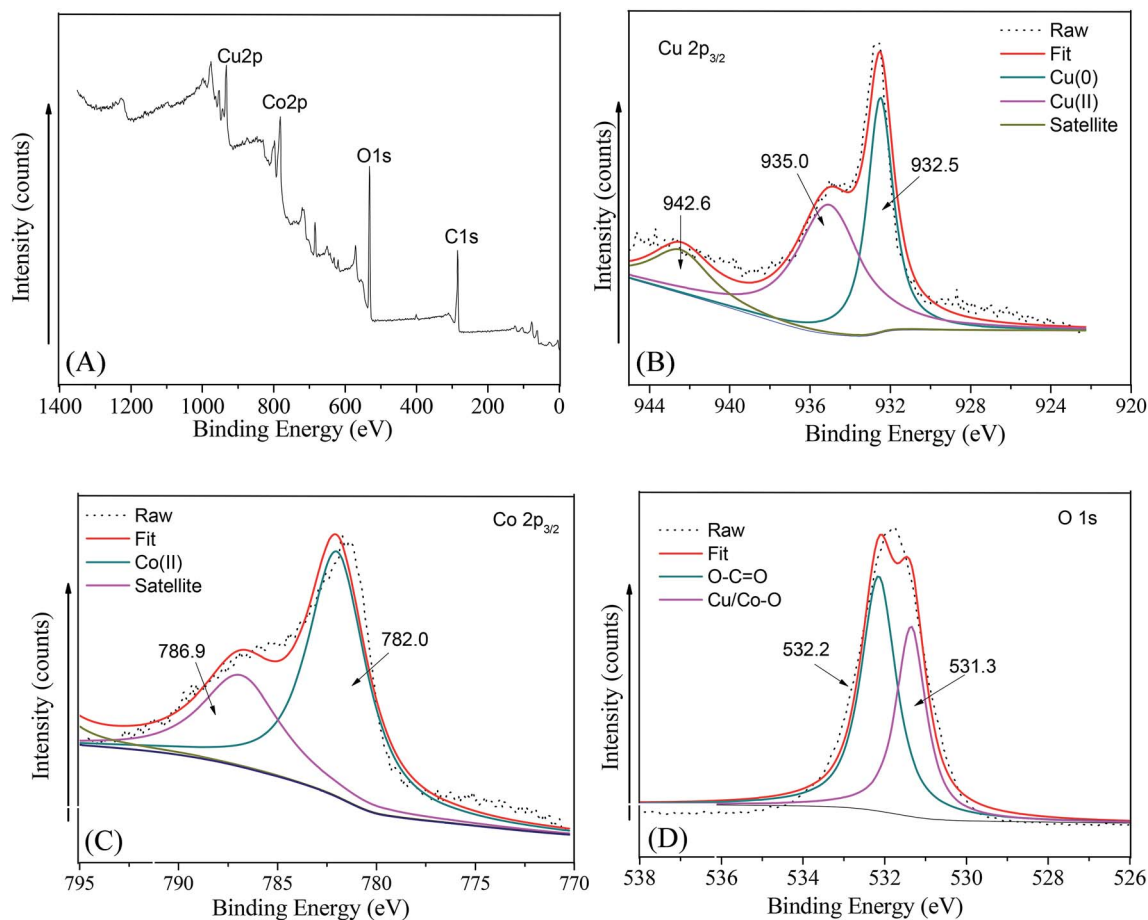
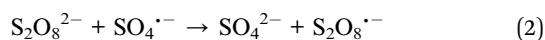
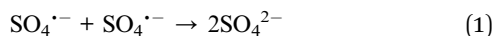


Fig. 2 The XPS spectra of the as-prepared Cu/Co-MOFs-3: (A) survey, (B) Cu 2p_{3/2}, (C) Co 2p_{3/2}, (D) O 1s.

the $n(\text{Co})/n(\text{Cu})$ ratio. Conversely, the MB removal increased when the $n(\text{Co})/n(\text{Cu})$ ratios increased from 1/4 to 3/2 in the PMS solutions. Nevertheless, a further increase in the $n(\text{Co})/n(\text{Cu})$ ratio resulted in a decrease in the MB removal rate when $n(\text{Co})/n(\text{Cu})$ exceeded 3/2. This could be elucidated by the well-known scavenging effect of $\text{SO}_4^{\cdot-}$ radical (eqn. (1) and (2)), resulting in the competitive reactions between MB and $\text{SO}_4^{\cdot-}$.



As a result, the MB removal rate in the Cu@Co-MOFs-4/PMS system reached the highest, but 100% of MB was completely removed in both the Cu@Co-MOFs-3/PMS and Cu@Co-MOFs-4/PMS systems. Hence, considering the cobalt toxicity and reducing its usage amount, the subsequent experiments were conducted using Cu@Co-MOFs-3 as catalysts.

3.3 Catalytic reactivity of Cu@Co-MOFs-3

The catalytic performance of the Cu@Co-MOFs-3 was evaluated by the MB removal efficiency and the mineralization rate of MB, the results are present in Fig. 5. As a comparison, experiments were performed in different systems (see in Fig. 4). It can be

clearly seen that less than 10% of MB was removed from the solution over 30 min in the control systems with only Cu@Co-MOFs-3, revealing that this material showed a low adsorption performance on MB. With the addition of PDS to the Cu@Co-MOFs-3/MB system, about 28% of MB was removed within 30 min, indicating that Cu@Co-MOFs-3 exhibited a considerably low catalytic activity for PDS. On the contrary, 100% of the MB removal rate was obtained under identical reaction conditions in the Cu@Co-MOFs-3/MB/PMS system. Although the TOC removal rate was only 12.4% within 30 min, it reached 62.9% after 9.0 h. This fact suggested that Cu@Co-MOFs-3 showed superior catalytic performances for the PMS oxidation of MB.

3.4 Effects of Cu/Co-MOFs-3 dosage and PMS concentration

The effect of the Cu@Co-MOFs-3 dosage on the MB removal rate was investigated with the aim to get further insight into the catalytic activity of Cu@Co-MOFs-3. As displayed in Fig. 6A, the MB removal rate increased evidently with an increase in the Cu@Co-MOFs-3 dosage from 30 mg L⁻¹ to 200 mg L⁻¹. It was noted that 0.2 mM of MB was completely removed within 30 min when the dosage was increased to 100 mg L⁻¹. Moreover, the final PMS consumption increased with an increase in the Cu@Co-MOFs-3 dosage (shown in Fig. 6B). The effect of the



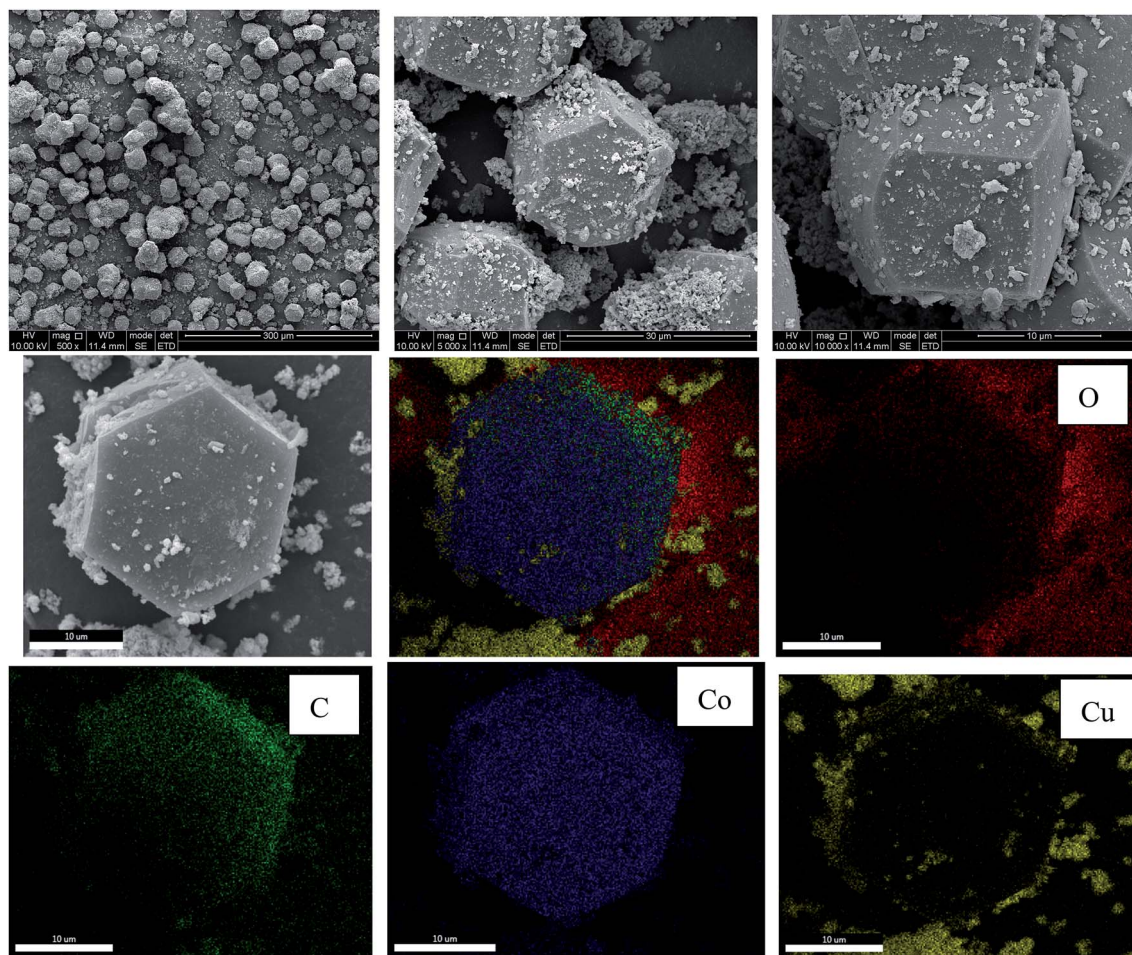


Fig. 3 SEM-EDS of as prepared Cu@Co-MOFs-3.

PMS dosage was also investigated, and the results are illustrated in Fig. 6C. It was observed that the MB removal efficiency increased evidently as the PMS concentration increased from 0.4 mM to 2.0 mM. However, a further increase in the PMS concentration resulted in a decrease in the MB removal, which could be due to the scavenging effect. Moreover, it was concluded that 2.0 mM of PMS was the optimal concentration for the MB removal in this system. These results further proved that the Cu@Co-MOFs-3 exhibited an excellent catalytic activity for the decomposition of PMS and generation of sulfate radicals.

3.5 Effect of initial pH and temperature

pH value is one of the main factors affecting the wastewater treatment using AOPs, hence, it is necessary to investigate the influence of the initial pH on the catalytic activity of Cu@Co-MOFs-3. The MB removal with an initial concentration (C_0) of 0.2 mM under different initial pH values is exhibited in Fig. 7A. It can be seen that the MB was completely removed within 30 min in the pH range of 3.0–11.0, indicating that the Cu@Co-MOFs-3 remained stable and showed an efficient catalytic activity over a wide pH range. Moreover, it was remarkable that the MB removal rate was highest at pH 11.0. This might be

attributed to the fact that the base could promote the PMS activation and the optimum initial pH is about 11 for PMS activation.³⁸ The effect of reaction temperature on the MB removal is illustrated in Fig. 7C, and it appears that the temperature has a weak influence on the MB removal rate, as MB was completely removed in the temperature of 25–60 °C. It is noteworthy that the first-order kinetic constants of MB increased evidently when the reaction temperature increased from 25 °C to 60 °C.

3.6 Stability of the Cu@Co-MOFs-3 and removal mechanism of MB

In order to evaluate the stability of Cu@Co-MOFs-3, four recycling runs using the Cu@Co-MOFs-3 were performed under the identical reaction conditions. As shown in Fig. 8A, the removal efficiency of MB showed a little decrease compared with the previous runs, where 96.4%, 94.0% and 89.6% of MB were removed within 30 min in the 2nd, 3rd, and 4th runs, respectively. However, the MB was completely removed when the reaction time was prolonged, indicating that the Cu@Co-MOFs-3 processed good stability and long lifetime. The slight decrease in the MB removal rate can be attributed to the negligible loss of Cu@Co-MOFs-3, and the decay of the active catalytic sites



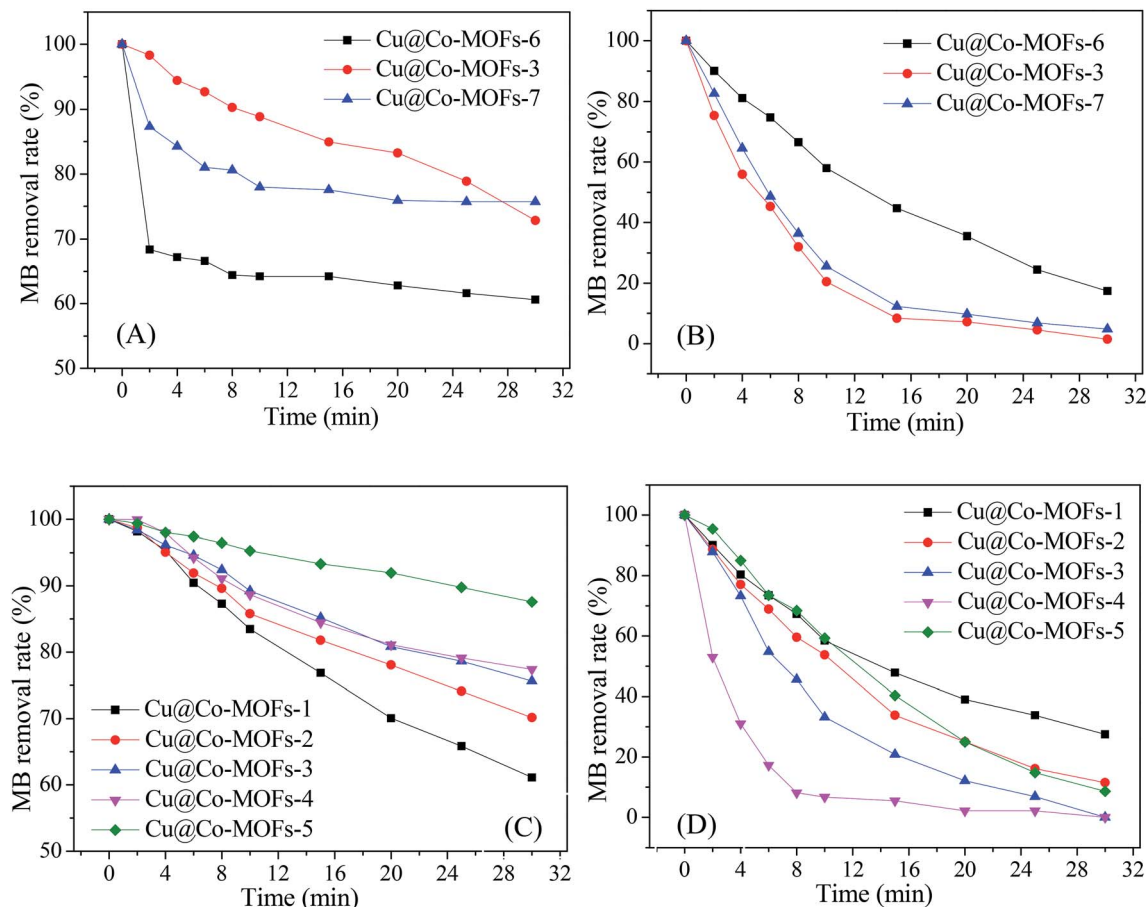


Fig. 4 (A) Effect of the molar ratio of [metal ions]/[BDC] on the MB removal rate in the presence of PDS; (B) effect of the molar ratio of [metal ions]/[BDC] on the MB removal rate in the presence of PMS; (C) effect of the molar ratio of $n(\text{Co})/n(\text{Cu})$ on the MB removal rate in the presence of PDS; (D) effect of the molar ratio of $n(\text{Co})/n(\text{Cu})$ on the MB removal rate in the presence of PMS. Experiment condition: PDS = 2.0 mM, PMS = 2.0 mM, MB = 0.2 mM, catalyst = 0.1 g L⁻¹, $T = 25\text{ }^{\circ}\text{C}$.

caused by the small amounts of leached cobalt and copper from the catalyst surface. According to the ICP-MS, the concentration of the leached cobalt was 3.89 mg L⁻¹ at pH 2.75, which was higher than the environmental quality standards for surface

water (GB 3838-2002) (less than 1.0 mg L⁻¹).³⁹ However, approximately 0.89 mg L⁻¹ cobalt was leached into the solution when the pH was set to 7.0, where 100% of the MB removal rate was observed (see in Fig. 6A).

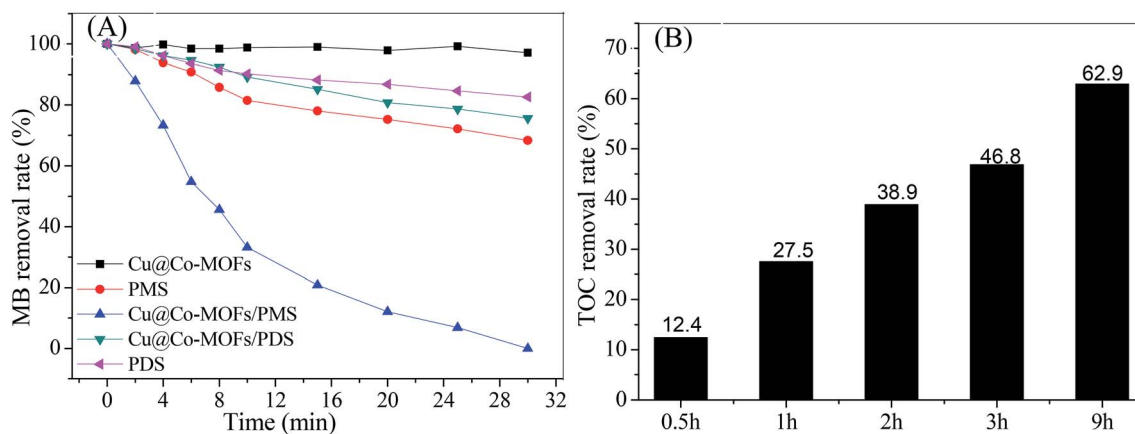


Fig. 5 (A) The degradation of MB versus time in different systems and experimental condition: PMS = 2 mM; PDS = 2 mM, MB = 0.2 mM; Cu@Co-MOFs-3 = 0.1 g L⁻¹, $T = 25\text{ }^{\circ}\text{C}$; (B) the TOC removal rate versus time, experiment condition: PMS = 4 mM, MB = 0.2 mM, Cu@Co-MOFs-3 = 0.2 g L⁻¹, $T = 25\text{ }^{\circ}\text{C}$.



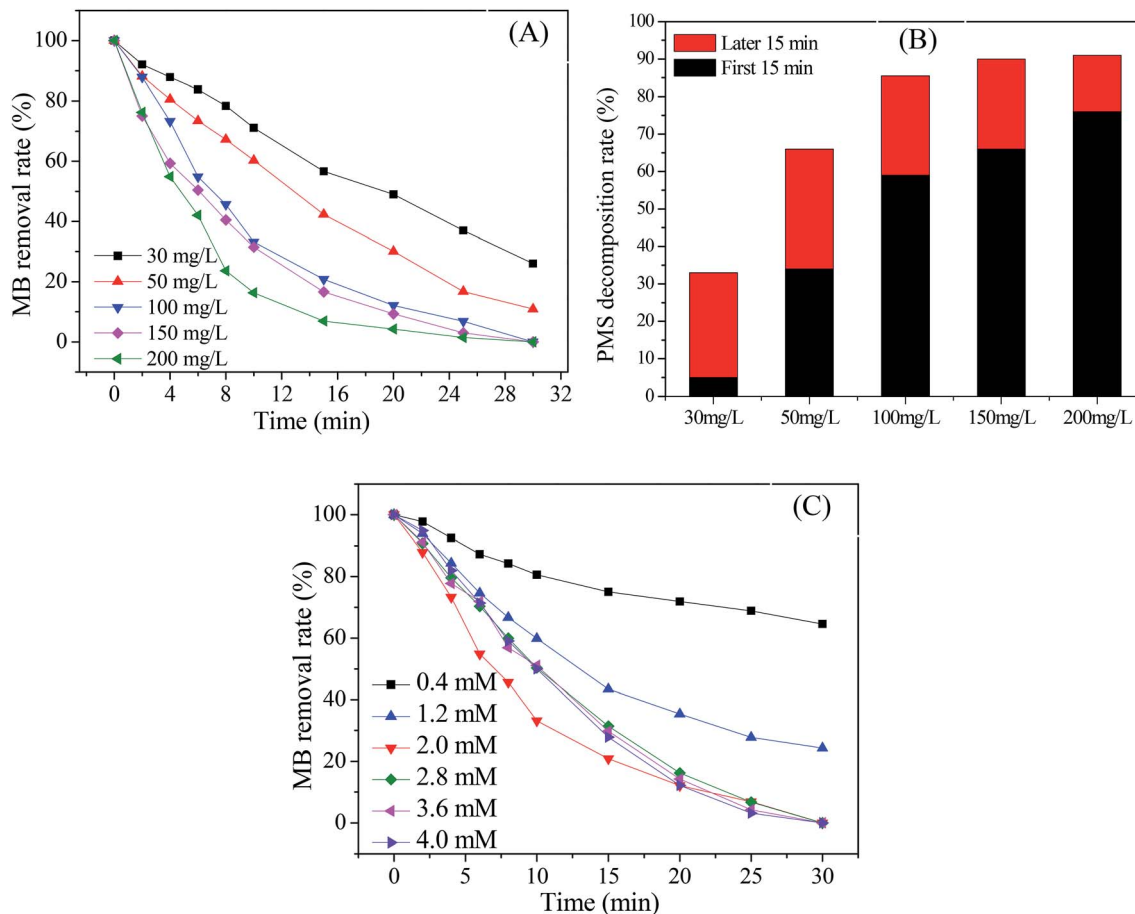


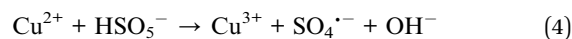
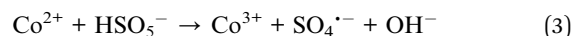
Fig. 6 (A) Effect of the catalyst dosage on the MB removal efficiency; (B) effect of the catalyst dosage on the PMS decomposition; (C) effect of the PMS concentration on the MB removal efficiency. Experiment condition: MB = 0.2 mM, PMS = 2 mM (for A and B), Cu@Co-MOFs-3 = 0.1 g L⁻¹ (only for C), T = 25 °C.

To gain insight into the pollutants removal mechanism, the scavenging experiments were investigated. Tertiary butanol (TBA) and ethanol (EtOH) are good candidates for quenching free radicals, including SO₄^{•-} and [•]OH. As expected, both TBA and EtOH exerted an inhibiting effect on the MB removal rates, as shown in Fig. 8B. The MB removal rates decreased with an increase in the concentration of TBA and EtOH. This fact indicated that SO₄^{•-} or [•]OH radical was the primary reactive species responsible for the removal of MB. Moreover, it was remarkable that the presence of EtOH led to a greater degree of inhibition than that of TBA, revealing the existence of SO₄^{•-}. According to previous studies,⁴⁰ pollutant degradation rates were inhibited by less than 10% by TBA if only SO₄^{•-} was generated for the oxidation of pollutants. The first-order kinetic constants of MB were inhibited by 24.7% and 41.3% when the molar ratio of TBA : MB were 500 : 1 and 1000 : 1, which indicated the presence of [•]OH.

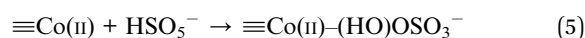
The excellent activity of Cu@Co-MOFs-3 for the activation of PMS to generate SO₄^{•-} and [•]OH radicals can be understood based on the synergistic behaviour of Co-MOFs-3 and zero-valent copper. According to the above-mentioned results, the heterogeneous catalytic reaction took place in the pores and the

surface of Cu@Co-MOFs-3 while the homogeneous reaction took place in water solutions, which resulted in the generation of SO₄^{•-} and [•]OH radicals. However, the heterogeneous reaction played an important role in the generation of SO₄^{•-} and [•]OH radicals due to the low leaching concentration of Co²⁺ and Cu²⁺. The overall heterogeneous process of the MB oxidation underwent some necessary elementary steps like PMS adsorption, free radical generation, MB adsorption and transformation into intermediates, and eventually the oxidation of the intermediates into CO₂. The overall process can be summarized by the following proposed mechanism, where the zero-valent copper, transition metal ions Co(II) and Cu(II) embedded in the framework acting as the catalytic site underwent electron translation and activated PMS to generate SO₄^{•-} (eqn. (3)–(12)).^{35,41}

In an aqueous solution:



In pores and on the surface of Cu@Co-MOFs-3:



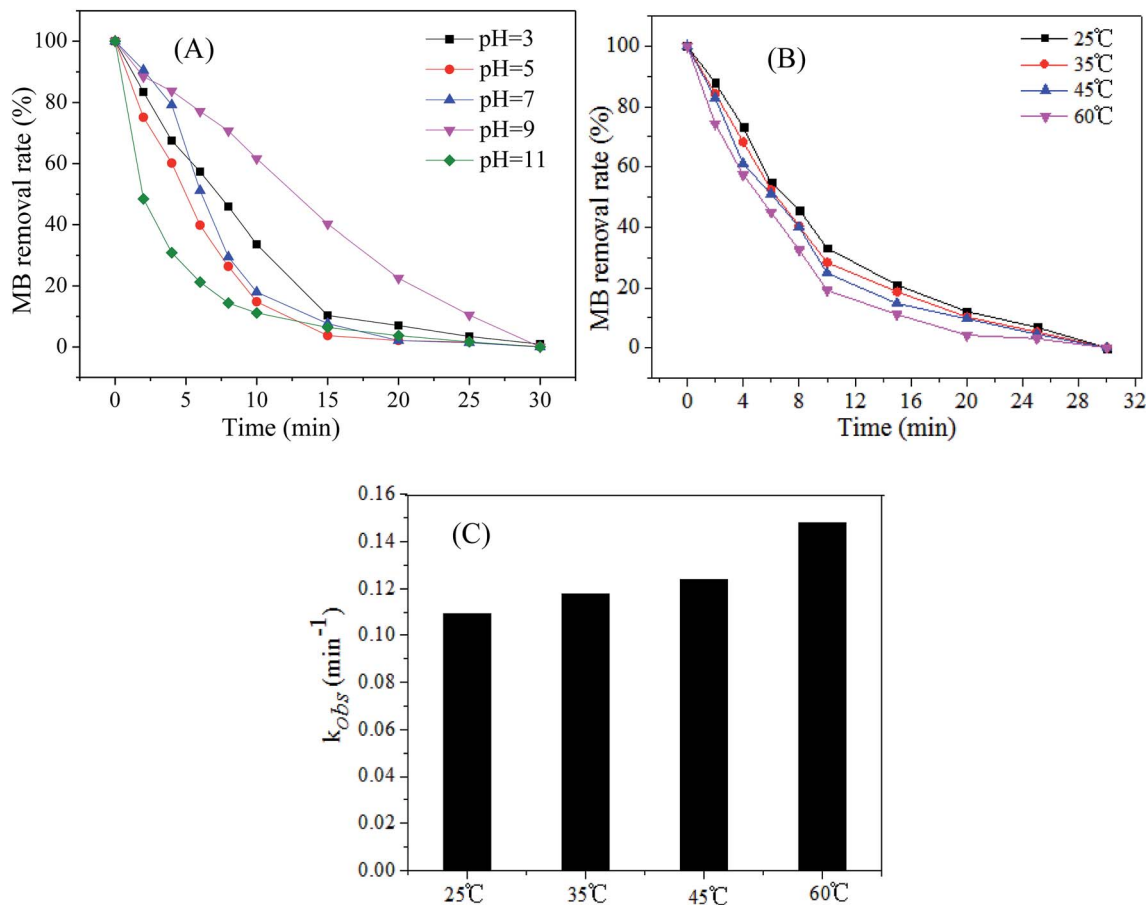


Fig. 7 (A) Effect of initial pH on the MB removal rate; (B) effect of temperature on the MB removal rate; (C) effect of temperature on the removal rate constants of MB. Experiment condition: PMS = 2 mM, MB = 0.2 mM, Cu@Co-MOFs-3 = 0.2 g L⁻¹, T = 25 °C.

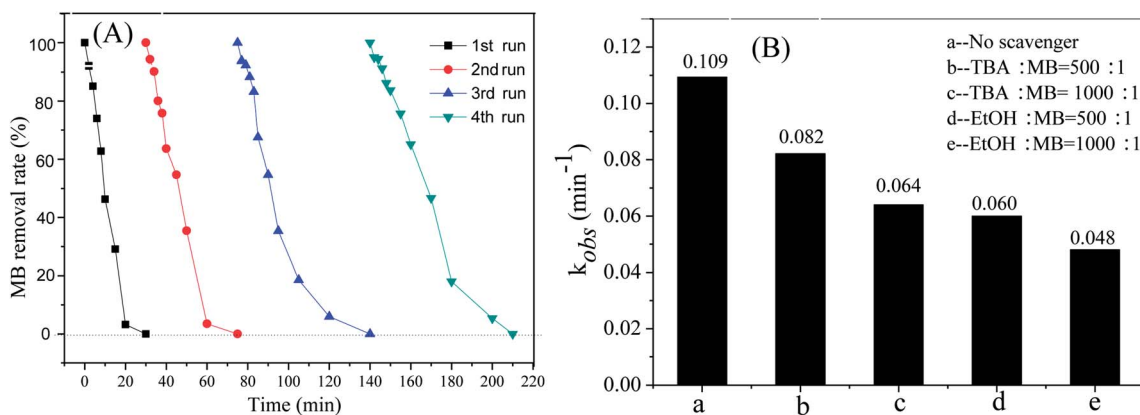
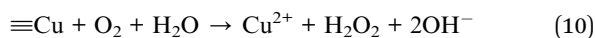
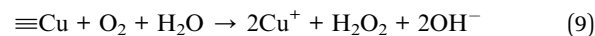
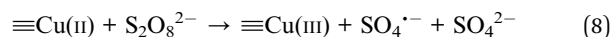
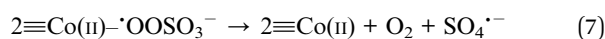
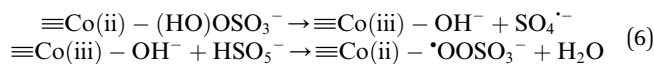
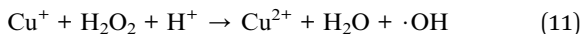


Fig. 8 (A) MB removal rate in multi-cycle batch experiments; (B) effect of scavenger on first-order kinetic constants of MB. Experiment condition: (A) PMS = 4 mM, MB = 0.2 mM, Cu@Co-MOFs-3 = 0.2 g L⁻¹, T = 25 °C; (B) PMS = 2 mM, MB = 0.2 mM, OG = 0.2 mM, Cu@Co-MOFs-3 = 0.1 g L⁻¹, T = 25 °C.





4. Conclusion

In this study, for the first time, we reported the synthesis of the Cu particle-doped Co-MOFs (Cu@Co-MOFs) by a hydrothermal method. The characterization results with XRD, FT-IR, SEM, EDS and XPS confirmed the incorporation of the zero-valent copper on the surface of the Co-MOFs structure. The catalytic performance of Cu@Co-MOFs obtained using different preparation conditions were studied for the oxidative removal of MB in the presence of PMS and PDS. It turned out that the Cu@Co-MOFs demonstrated a superior catalytic performance towards PMS for the removal of MB at a wide pH range. The best removal result was achieved with Cu@Co-MOFs-3, where 100% of MB was removed within 30 min. The MB removal rate increased with the increase in the initial pH values, reaction temperature and catalyst dosages. The percent of MB removal decreased with an increase in the PMS concentration, and the optimum PMS concentration was determined to be 2.0 mM. Moreover, the Cu@Co-MOFs-3 showed relative stability in the catalytic process with a low decline in the MB removal in four runs. The MB removal mechanism was proposed, which was ascribed to the oxidation of the primary reactive species ($\text{SO}_4^{\cdot-}$ and $\cdot\text{OH}$). The findings of this study provide important insights into the development of heterogeneous PMS catalysts for wastewater treatment.

Conflicts of interest

There are no conflicts to declare.

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