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# Tunable preparation of highly dispersed Ni<sub>x</sub>Mn-LDO catalysts derived from Ni<sub>x</sub>Mn-LDHs precursors and application in low-temperature NH<sub>3</sub>-SCR reactions†

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A series of Ni<sub>x</sub>Mn bimixed metal oxides (Ni<sub>x</sub>Mn-LDO) were prepared *via* calcining Ni<sub>x</sub>Mn layered double hydroxides (Ni<sub>x</sub>Mn-LDHs) precursors at 400 °C and applied as catalysts in the selective catalytic reduction (SCR) of NO<sub>x</sub> with NH<sub>3</sub>. The DeNO<sub>x</sub> performance of catalysts was optimized by adjusting the Ni/Mn molar ratios of Ni<sub>x</sub>Mn-LDO precursors, in which Ni<sub>5</sub>Mn-LDO exhibited above 90% NO<sub>x</sub> conversion and N<sub>2</sub> selectivity at a temperature zone of 180–360 °C. Besides, Ni<sub>5</sub>Mn-LDO possessed considerable SO<sub>2</sub> & H<sub>2</sub>O resistance and outstanding stability. Multiple characterization techniques were used to analyze the physicochemical properties of the catalysts. The analysis results indicated that all catalysts had the same active species Ni<sub>6</sub>MnO<sub>8</sub>, while their particle sizes showed significant differences. Notably, the uniform distribution of active species particles in the Ni<sub>5</sub>Mn-LDO catalyst provided the rich surface acidity and suitable redox ability which were the primary causes for its desirable DeNO<sub>x</sub> property.

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## 1. Introduction

During the past decades, nitrogen oxide (NO<sub>x</sub>) emissions have increased dramatically, which triggered a battery of environmental problems including photochemical smog, acid rain and greenhouse effects, further endangering human health.<sup>1,2</sup> With the increasingly severe strictness of environmental regulations, it is necessary to find a valid method to eliminate NO<sub>x</sub>. Among numerous removal techniques, selective catalytic reduction of NO<sub>x</sub> with NH<sub>3</sub> (NH<sub>3</sub>-SCR) has been considered as the most effective measure for denitrification (DeNO<sub>x</sub>).<sup>3,4</sup> Currently, commercial catalysts V<sub>2</sub>O<sub>5</sub>-WO<sub>3</sub>/TiO<sub>2</sub> have been widely used in the field of industrial stationary sources, which displayed good catalytic performance at the high temperature region of 300–400 °C.<sup>5</sup> However, V-based catalysts still possess defects, such as the narrow operation temperature window, low N<sub>2</sub> selectivity at higher temperatures and inherent toxicity, which restricted its further application.<sup>6</sup> Thus, it is worthwhile to develop non-toxic DeNO<sub>x</sub> catalysts with good low temperature activity.

Based on the features of environmentally friendly and variable valence of Mn element, MnO<sub>x</sub> has been deemed as the potential catalyst for NH<sub>3</sub>-SCR reaction in the low temperature

zone.<sup>7–9</sup> Whereas there were some inevitably problems caused by its stronger oxidation capacity which easily led to the over-oxidation of NH<sub>3</sub> at higher temperature and further reduced SCR activity as well as N<sub>2</sub> selectivity.<sup>10,11</sup> Recently, it has been reported that the strong oxidation capacity of MnO<sub>x</sub> catalyst could be adapted by introducing transition metals.<sup>12–15</sup> Among all the attempts, Ni was considered to be a functional component with excellent synergistic catalysis. For instance, Wan *et al.*<sup>16</sup> and Chen *et al.*<sup>17</sup> prepared Ni–Mn bi-metal oxides and Ni–Mn spinel oxides catalysts, respectively, the catalytic performance results showed that both catalysts exhibited higher SCR activity and N<sub>2</sub> selectivity than single MnO<sub>x</sub> catalyst. Although some progress has been made, it was unavoidable that the active specie particles of Ni–Mn catalysts prepared by traditional methods were prone to agglomerate, resulting in the decline of catalytic performance.<sup>18</sup>

As investigated in the previous studies, the dispersion of active specie particles over catalyst could be effectively improved *via* regulating the morphology and structure.<sup>19–21</sup> Lately, the layered double oxides (LDO) catalysts obtained *via* calcining layered double hydroxides (LDHs) precursor attracted proverbial attention due to the excellent physical and chemical properties.<sup>22–24</sup> Moreover, our group have also found that the catalysts prepared by employing LDHs as precursor could achieve better SCR activity than other methods, it might be related with the lattice limiting effect and surface hydrazine structure existed in bulk laminate of LDHs, which were helpful to improve the dispersion and stability of the active specie particles.<sup>25</sup>

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In this work, on account of the synergistic effect of Ni and Mn as well as satisfactory function of LDHs, Ni<sub>x</sub>Mn-LDHs with different Ni/Mn molar ratios were successfully fabricated with aid of coprecipitation methods. Ni<sub>x</sub>Mn-LDO obtained by calcining Ni<sub>x</sub>Mn-LDHs at a proper temperature were hired as DeNO<sub>x</sub> catalysts in NH<sub>3</sub>-SCR reaction. In order to optimize physical and chemical properties of Ni<sub>x</sub>Mn-LDO catalysts, the relative proportions of Ni and Mn of the Ni<sub>x</sub>Mn-LDHs precursors were explored in detail. A series of characterization techniques were used to reveal the intrinsic reason for the differences in catalytic activities among catalysts. This research might expand a new horizon to develop new catalysts with outstanding efficiency.

## 2. Experimental section

### 2.1 Preparation of catalysts

The Ni<sub>x</sub>Mn-LDHs ("x" represent Ni/Mn molar ratio = 3, 4, 5, 6 and total molar amount Ni + Mn = 0.06 mol) were prepared by coprecipitation method.<sup>26</sup> To be specific, for the synthesis of Ni<sub>3</sub>Mn-LDHs, 0.045 mol (13.09 g) of Ni(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O (98.0%) and 0.015 mol (5.37 g) of Mn(NO<sub>3</sub>)<sub>2</sub> (50%) were dissolved in 100 mL deionized water to prepare metal precursor solution. The mixed alkali solution were obtained by dissolving 0.04 mol (1.6 g) of NaOH and 2.8 g of NH<sub>3</sub>·H<sub>2</sub>O (25–28%) in 100 mL deionized water. Then, the metal precursor solution and alkali solution were simultaneously added dropwise with magnetic stirring into flask at room temperature. The suspension was crystallized at 80 °C for 24 h, and then the resulting subsidence by filtration, washing and drying at 80 °C for 12 h to obtain the Ni<sub>3</sub>Mn-LDH precursor. The Ni<sub>3</sub>Mn-LDO catalyst was obtained by calcining Ni<sub>3</sub>Mn-LDHs precursor in air at 400 °C for 5 h at a heating rate of 2 °C min<sup>-1</sup>. Similarly, other samples such as Ni<sub>4</sub>Mn-LDO, Ni<sub>5</sub>Mn-LDO and Ni<sub>6</sub>Mn-LDO were obtained with the same method.

### 2.2 Characterization of catalysts

X-ray diffraction (XRD) information of catalysts was measured on a Rigaku DX-2700 instrument using Cu/Kα radiation source. The 2θ range of meterage from 5 to 85° and a scanning rate was 8° min<sup>-1</sup>. The Ni and Mn contents were determined by inductively coupled plasma-atomic (ICP) on Optima 7000 DV instrument. Fourier transform infrared spectrometer (FT-IR) experiment were tested from 400 to 3600 cm<sup>-1</sup> on a TENSOR II FT-IR (Bruker). Thermogravimetry (TG) analyzer (Setaram Evolution) equipped with quadrupole mass spectrometry (QMS) real-time monitored the thermal decomposition of LDHs sample with the temperature rising from 50 °C to 800 °C. The surface area and pore information of catalysts were obtained by using Brunauer–Emmett–Teller (BET) and Barrett–Joyner–Halenda (BJH) methods on Micromeritics ASAP-2460 apparatus. The test conditions are mainly that the catalyst samples were degassed at 250 °C for 4 h. Ammonia-temperature programmed desorption (NH<sub>3</sub>-TPD) was performed on FINESORB-3010C chemisorptions analyzer with thermal conductivity detector (TCD). About 100 mg catalysts (40–60 mesh) was pretreated at 300 °C for 1 h under 25 mL min<sup>-1</sup> He atmosphere;

subsequently, the sample adsorbed NH<sub>3</sub> to achieve saturation at 50 °C for 1 h and He was purged at the same temperature to remove weakly adsorbed NH<sub>3</sub> on the surface; finally, the desorption curve was recorded *via* increasing temperature (from 50 °C to 560 °C) with heating rate of 10 °C min<sup>-1</sup> under 25 mL min<sup>-1</sup> He atmosphere. Hydrogen-temperature programmed reduction (H<sub>2</sub>-TPR) was tested in the same way. About 80 mg catalysts was pretreated at 300 °C for 1 h under 30 mL min<sup>-1</sup> Ar atmosphere; after cooling down to 50 °C, the catalyst samples accomplished the reduction process from 80 °C to 800 °C in a flow of 30 mL min<sup>-1</sup> 10% H<sub>2</sub>/Ar. In order to explore the surface morphology and structure of the sample, scanning electron microscope (SEM) and high resolution-transmission electron microscopy (HR-TEM) were conducted on a Hitachi SU8010 apparatus and JEM-2100F (JEOL) with an accelerating voltage of 200 kV, respectively. The EDS equipped with SEM was also used to characterize the surface composition of catalyst.

### 2.3 Activity measurements of catalysts

The SCR activity measurements of catalysts were carried out on a fixed-bed quartz reactor with the internal diameter of 6 mm. About 350 mg catalyst samples (40–60 mesh) were placed in the reactor. The reaction gas composition was as follows: 600 ppm NO, 600 ppm NH<sub>3</sub>, 5 vol% O<sub>2</sub>, 10 vol% H<sub>2</sub>O (when used), 100 ppm SO<sub>2</sub> (when used) and N<sub>2</sub> as the balance gas. The total flow rate of gas was 187 mL min<sup>-1</sup> and GHSV = 45 000 h<sup>-1</sup>. The activity performance were tested at range from 90 to 390 °C and the data were collected for every 30 °C, meanwhile, the output gases were monitored by multigas 2000 FTIR. The NO<sub>x</sub> conversion and N<sub>2</sub> selectivity were calculated as follows:

$$\text{NO}_x \text{ conversion} = 100\% \times ([\text{NO}_x]_{\text{in}} - [\text{NO}_x]_{\text{out}})/[\text{NO}_x]_{\text{in}}$$

$$\text{N}_2 \text{ selectivity} = 100\% \times ([\text{NO}_x]_{\text{in}} + [\text{NH}_3]_{\text{in}} - [\text{NO}_x]_{\text{out}} - [\text{NH}_3]_{\text{out}} - 2[\text{N}_2\text{O}]_{\text{out}})/([\text{NO}_x]_{\text{in}} + [\text{NH}_3]_{\text{in}} - [\text{NO}_x]_{\text{out}} - [\text{NH}_3]_{\text{out}})$$

where "in" represents the gas concentration at the inlet, "out" represents the gas concentration at the output.

## 3. Results and discussion, experimental

### 3.1 Characterization of Ni<sub>x</sub>Mn-LDHs precursors

**3.1.1 XRD analysis.** XRD patterns of the Ni<sub>x</sub>Mn-LDHs precursors with different Ni/Mn molar ratios were displayed in Fig. 1. As the results presented, all samples showed the typical diffraction peaks at about 11.0°, 22.1°, 34.3°, 38.7°, 45.0°, 60.0° and 61.1°, which can be indexed to (003), (006), (012), (015), (018), (110) and (113) planes of Ni<sub>x</sub>Mn-LDHs,<sup>27–29</sup> and no other miscellaneous peaks were observed. That is to say, Ni<sub>x</sub>Mn-LDHs with pure phase were synthesized successfully. It can be seen that the (003) diffraction peaks of four samples appear at 11°, further the basal spacings of Ni<sub>x</sub>Mn-LDHs were calculated as 8.04 Å by Bragg's equation of  $n\lambda = 2d \sin \theta$ , which indicated that CO<sub>3</sub><sup>2-</sup> existed in the interlayers of LDHs.<sup>26,30</sup> In addition,



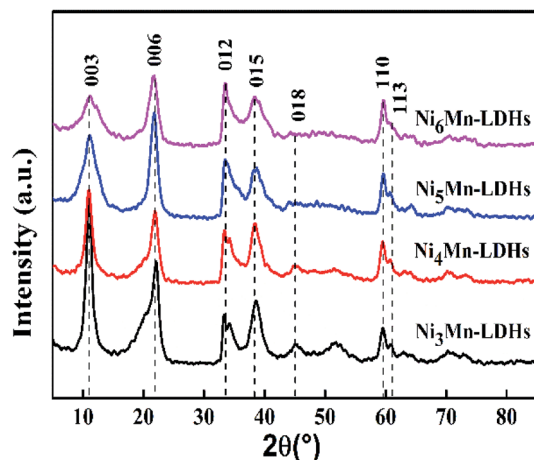


Fig. 1 XRD patterns of  $\text{Ni}_x\text{Mn-LDHs}$  precursors ( $x = 3, 4, 5, 6$ ).

with the increase of Ni/Mn molar ratios for  $\text{Ni}_x\text{Mn-LDHs}$ , it can be observed that the intensity of the (003) diffraction peaks showed a tendency to reduce, and the peaks width became wider. This implied that  $\text{Ni}_5\text{Mn-LDHs}$  and  $\text{Ni}_6\text{Mn-LDHs}$  possessed smaller LDHs crystal sizes compared with  $\text{Ni}_3\text{Mn-LDHs}$  and  $\text{Ni}_4\text{Mn-LDHs}$  (Table 1), which may change the appearance of the LDHs precursors and thus affect the dispersibility of the calcined product. Furthermore, tested Ni/Mn molar ratios were determined by ICP analysis (Table 1), and the result was consistent with feeding ratios. The composition and morphology information of the  $\text{Ni}_x\text{Mn-LDHs}$  were further investigated by FT-IR and SEM.

**3.1.2 FT-IR analysis.** Fig. 2 showed the FT-IR spectra for  $\text{Ni}_x\text{Mn-LDHs}$  ( $x = 3, 4, 5, 6$ ). It could be found that a series of infrared adsorption bands were displayed within the wave-number range of 400–3600  $\text{cm}^{-1}$ . Broad adsorption band at 3422  $\text{cm}^{-1}$  could be assigned to stretching vibration of  $-\text{OH}$  groups in the  $\text{Ni}_x\text{Mn-LDHs}$  layers and interlayer water molecules.<sup>31</sup> The adsorption band at 1641  $\text{cm}^{-1}$  might be due to the angular deformation of water.<sup>32</sup> The vibration of  $\text{CO}_3^{2-}$  and  $\text{NO}_3^-$  existed in interlayers could be observed at 1381 and 1359  $\text{cm}^{-1}$ , respectively.<sup>26</sup> Additionally, the adsorption bands around 634 and 453  $\text{cm}^{-1}$  could be attributed to the cation-oxygen vibrations, such as  $\text{O-M-O}$ ,  $\text{M-O-M}$  or  $\text{M-OH}$ .<sup>33</sup> These results further indicated that the prepared  $\text{Ni}_x\text{Mn-LDHs}$  samples were  $\text{CO}_3^{2-}$  and  $\text{NO}_3^-$  intercalated.

**Table 1** Chemical compositions and XRD parameters of  $\text{Ni}_x\text{Mn-LDHs}$  ( $x = 3, 4, 5, 6$ ) precursors

Samples	Weight content <sup>a</sup>		Ni : Mn	$\text{FWHM}_{003}$	$D_{003}^b$ (nm)
	Ni	Mn			
$\text{Ni}_3\text{Mn-LDHs}$	44.37	15.32	2.89	0.0164	8.39
$\text{Ni}_4\text{Mn-LDHs}$	47.66	12.41	3.84	0.0170	8.09
$\text{Ni}_5\text{Mn-LDHs}$	49.01	9.55	5.13	0.0216	6.40
$\text{Ni}_6\text{Mn-LDHs}$	51.22	8.71	5.88	0.0222	6.21

<sup>a</sup> Weight content of Ni and Mn was determined by ICP analysis. <sup>b</sup>  $D_{003}$  of samples was calculated by Scherrer equation.

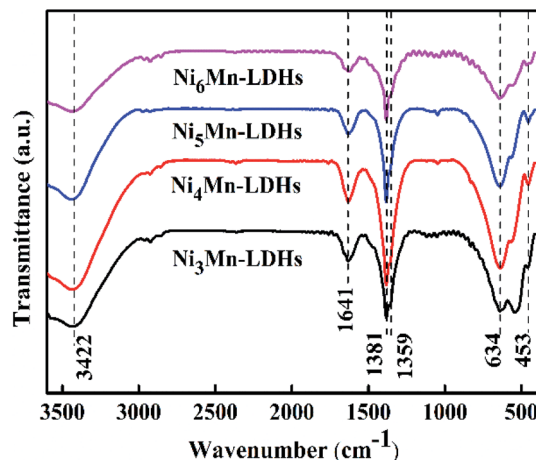


Fig. 2 FT-IR spectra of  $\text{Ni}_x\text{Mn-LDHs}$  precursors ( $x = 3, 4, 5, 6$ ).

**3.1.3 SEM analysis.** The SEM characterization was adopted to probe the microscopic morphology of the precursors, the corresponding results were shown in Fig. 3. As can be seen, the four samples exhibited typical nanosheet structure, which might further witnessed the successful synthesis of LDHs precursors.<sup>34,35</sup> Moreover, it could also be observed that surface morphology of four samples has changed obviously with the increase of Ni/Mn molar ratios. For  $\text{Ni}_3\text{Mn-LDHs}$  and  $\text{Ni}_4\text{Mn-LDHs}$ , the LDHs sheets presented a larger size and seriously stacked, which will greatly affect the exposure of the effective components on the surface of the sample. In contrast, the sheets size of  $\text{Ni}_5\text{Mn-LDHs}$  and  $\text{Ni}_6\text{Mn-LDHs}$  was significantly reduced and the degree of accumulation was greatly improved. In fact, thus morphology difference of precursor often influence the properties of the calcined product.

**3.1.4 TG-MS analysis.** The thermal decomposition of the  $\text{Ni}_x\text{Mn-LDHs}$  precursors was clearly explored by TG-MS analysis. As shown in Fig. 4, take  $\text{Ni}_5\text{Mn-LDHs}$  as an example. Three distinct weight loss peaks appeared on the curves of TG and DTG (Fig. 4(a)). The first weight loss stage around 80–130  $^{\circ}\text{C}$  was caused by the removal of physically adsorbed water, interlayer

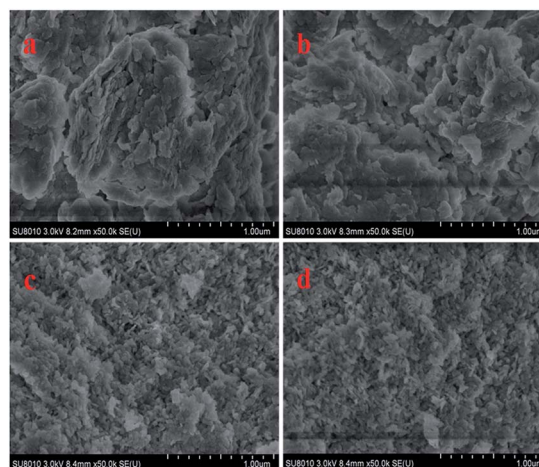


Fig. 3 SEM images of (a)  $\text{Ni}_3\text{Mn-LDHs}$ ; (b)  $\text{Ni}_4\text{Mn-LDHs}$ ; (c)  $\text{Ni}_5\text{Mn-LDHs}$  and (d)  $\text{Ni}_6\text{Mn-LDHs}$  precursors.



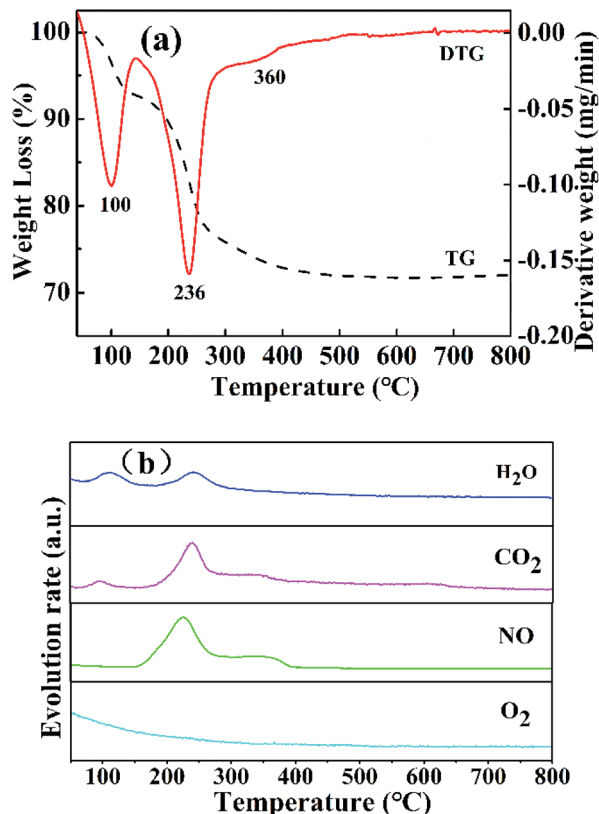


Fig. 4 Curves of TG and DTG (a) and MS (b) for  $\text{Ni}_5\text{Mn-LDHs}$  samples.

water and  $\text{CO}_2$  molecules, where the release signal of the corresponding gases ( $\text{H}_2\text{O}$  and  $\text{CO}_2$ ) was detected by MS (Fig. 4(b)).<sup>36</sup> The major weight loss process occurred at second stage, which took place at the temperature range of 150–300 °C, mainly arising from the dehydration of  $-\text{OH}$  group on the LDHs laminate as well as the loss of volatile gases  $\text{CO}_2$  and  $\text{NO}$  from the interlayer  $\text{CO}_3^{2-}$  and  $\text{NO}_3^-$  ions.<sup>37</sup> The last weight loss stage displayed a slight exothermic peak at about 360 °C, a small amount of  $\text{CO}_2$  and  $\text{NO}$  release can be still observed from MS (Fig. 4(b)), which can be assigned to decomposition of interlayer anions  $\text{CO}_3^{2-}$  and  $\text{NO}_3^-$  at higher temperature.<sup>32</sup> As the temperature further increased (>400 °C), no weight loss peak was found, indicating that the  $\text{Ni}_5\text{Mn-LDHs}$  precursor achieved complete decomposition. Based on the TG-MS analysis, it can be concluded that calcining LDHs precursors at 400 °C can form stable mixed metal oxides.

### 3.2 Characterization of $\text{Ni}_x\text{Mn-LDO}$ ( $x = 3, 4, 5, 6$ ) catalysts

**3.2.1 SCR performance of  $\text{Ni}_x\text{Mn-LDO}$  ( $x = 3, 4, 5, 6$ ) catalysts.** SCR activity and  $\text{N}_2$  selectivity of  $\text{Ni}_x\text{Mn-LDO}$  catalysts with different metal molar ratios ( $\text{Ni}/\text{Mn} = 3, 4, 5, 6$ ) were tested under a GHSV of 45 000  $\text{h}^{-1}$ . As illustrated in Fig. 5(a), the  $\text{NO}_x$  conversion of all catalysts has experienced three stages at the test temperature range of 90–390 °C. At the first stage, the  $\text{NO}_x$  conversion of  $\text{Ni}_x\text{Mn-LDO}$  catalysts increased sharply with increasing temperature and gradually reached the maximum of 98%. Followed by the second stage, the  $\text{NO}_x$  conversion

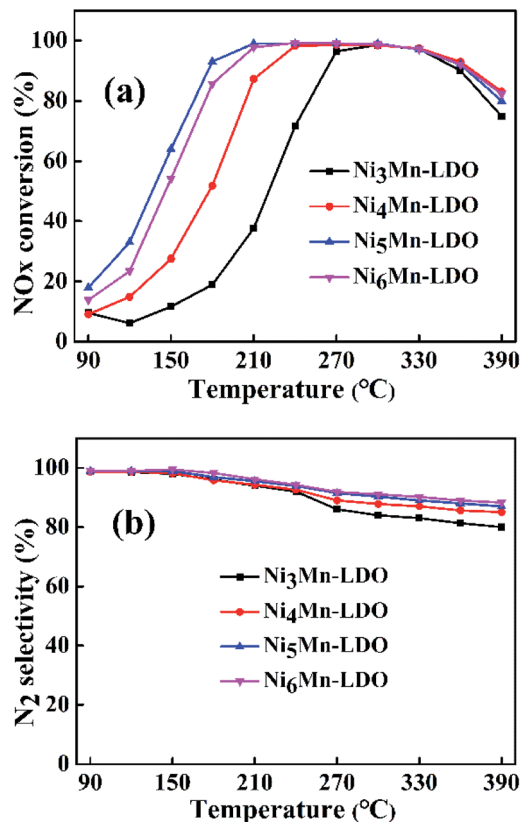


Fig. 5 (a)  $\text{NO}_x$  conversion and (b)  $\text{N}_2$  selectivity of  $\text{Ni}_x\text{Mn-LDO}$  ( $x = 3, 4, 5, 6$ ) catalysts. Reaction conditions:  $[\text{NO}] = [\text{NH}_3] = 600$  ppm,  $[\text{O}_2] = 5$  vol%,  $\text{N}_2$  as balance gas, GHSV = 45 000  $\text{h}^{-1}$ .

substantially remained at 98% without obvious changes. On the third stage, the  $\text{NO}_x$  conversion presented a declining trend. Notably, it can be observed that the  $\text{NO}_x$  conversion of  $\text{Ni}_x\text{Mn-LDO}$  catalysts were significantly influenced by  $\text{Ni}/\text{Mn}$  molar ratios on the first stage. When the  $\text{Ni}/\text{Mn}$  molar ratios rose from 3 to 5, the  $\text{NO}_x$  conversion over  $\text{Ni}_x\text{Mn-LDO}$  catalysts displayed an upward trend,  $\text{Ni}_5\text{Mn-LDO}$  catalyst afforded the best catalytic activity. However, when the  $\text{Ni}/\text{Mn}$  molar ratios was further increased,  $\text{Ni}_6\text{Mn-LDO}$  exhibited lower activity than the  $\text{Ni}_5\text{Mn-LDO}$  catalyst, which might be related to the further reduction of the Mn content. Thus, the SCR activity of different catalysts presented as the following order:  $\text{Ni}_5\text{Mn-LDO} > \text{Ni}_6\text{Mn-LDO} > \text{Ni}_4\text{Mn-LDO} > \text{Ni}_3\text{Mn-LDO}$ . In order to further understand the catalytic performance of the  $\text{Ni}_x\text{Mn-LDO}$  catalysts, the SCR activity comparison of present catalyst and the other  $\text{Ni-Mn}$  catalysts in literatures was involved in Table S1.† As the result presented that  $\text{Ni}_x\text{Mn-LDO}$  catalyst exhibited different activity temperature windows, which may be related to the structural composition of the LDHs precursors.

The  $\text{N}_2$  selectivity of  $\text{Ni}_x\text{Mn-LDO}$  catalysts was also considered. As shown in Fig. 5(b), all catalysts exhibited similar  $\text{N}_2$  selectivity about 85% or higher over the entire temperature range. Whereas, with the temperature raising and exceeding 180 °C, the  $\text{N}_2$  selectivity of each catalyst decreased to varying degrees, in which  $\text{Ni}_5\text{Mn-LDO}$  and  $\text{Ni}_6\text{Mn-LDO}$  catalysts displayed better  $\text{N}_2$  selectivity than the other two. These results





indicated that SCR performance of  $\text{Ni}_x\text{Mn-LDO}$  could be affected by the Ni/Mn molar ratios.

**3.2.2 XRD analysis.** Fig. 6 gave XRD patterns of  $\text{Ni}_x\text{Mn-LDO}$  catalysts. As can be seen from the diagram, all catalysts exhibited a similar XRD characteristic diffraction peaks. The peaks located at  $18.7^\circ$ ,  $37.3^\circ$ ,  $43.4^\circ$ ,  $63.1^\circ$  can be attributed to  $\text{Ni}_6\text{MnO}_8$  phase (JCPDS 49-1295).<sup>38</sup> The appearance of  $\text{Ni}_6\text{MnO}_8$  crystal phase can be related to the formation of Mn–Ni–O solid solution, which is mainly because that part of  $\text{Mn}^{3+}$  and  $\text{Mn}^{4+}$  can insert the lattice of NiO where the radius of  $\text{Ni}^{2+}$  (0.069 nm) is greater than  $\text{Mn}^{3+}$  (0.066 nm) and  $\text{Mn}^{4+}$  (0.056 nm).<sup>39</sup> It was reported that the phase of  $\text{Ni}_6\text{MnO}_8$  was conducive to the improvement of SCR activity.<sup>20</sup> The peaks at  $2\theta$  of  $37.3^\circ$ ,  $43.4^\circ$ ,  $63.1^\circ$ ,  $75.5^\circ$ ,  $79.7^\circ$  can be assigned to NiO phase (JCPDS 75-0197),<sup>38,39</sup> it was partially overlapped with those of  $\text{Ni}_6\text{MnO}_8$  and can be further analyzed by TEM. Meanwhile, the absence of the diffraction peaks for pure  $\text{MnO}_x$  may be due to that a part of Mn participated in the formation of  $\text{Ni}_6\text{MnO}_8$  and remained Mn existed in an amorphous state. In addition, with the increase of Ni/Mn molar ratios for  $\text{Ni}_x\text{Mn-LDO}$ , diffraction peaks of samples were broadened and its intensity descended, indicating the gradual decreasing of crystal size. Generally speaking, small crystal can expose more crystal defects, which is more beneficial to improving catalytic activity.<sup>39</sup> This well explained the difference in catalytic activity for catalysts to some extent.

**3.2.3 SEM, TEM and HRTEM analysis.** The degree of exposure of catalyst surface active species is an important factor affecting catalytic activity.<sup>40</sup> Therefore, the surface information of  $\text{Ni}_x\text{Mn-LDO}$  catalyst was investigated by SEM and TEM. Above all, the SEM images of four catalysts were shown in Fig. S1–S4.† After  $400^\circ\text{C}$  calcination, all samples exhibited a spherical morphology and the sheet structure disappeared completely, which indicated that the LDHs laminate collapsed and a stable composite metal oxide formed. To further explore the surface element composition of the catalyst, the EDS mapping of  $\text{Ni}_x\text{Mn-LDO}$  catalysts was also conducted, as shown in Fig. S1–S4.† It can be seen that the O, Ni and Mn element was well dispersed on the surface of the catalyst. In addition, based on

the relative content of Ni and Mn elements of each catalyst, surface Ni/Mn molar ratios of four samples was calculated. For  $\text{Ni}_5\text{Mn-LDO}$  catalyst, it was noticed that its Ni/Mn molar ratios (4.68) was much lower than tested value (5.13) determined by ICP analysis. This result suggested that more Mn elements can be fully exposed on the catalyst surface and further enhancing SCR activity, which may be related to the uniform distribution of metal oxide particles in the catalyst.

Then the TEM images of  $\text{Ni}_x\text{Mn-LDO}$  catalysts were shown in Fig. 7(a–d), it can be clearly observed that the active nanoparticles are anchoring on the catalyst surface, but there are some difference in dispersion with Ni/Mn molar ratios change of catalysts. As for  $\text{Ni}_3\text{Mn-LDO}$  and  $\text{Ni}_4\text{Mn-LDO}$  catalysts, a portion of active species particles have undergone significant agglomeration, with the corresponding average particle size being 17.67 nm and 15.68 nm, respectively. In contrast, the aggregation degree of active nanoparticles for  $\text{Ni}_5\text{Mn-LDO}$  and  $\text{Ni}_6\text{Mn-LDO}$  catalysts was improved significantly with the average particle size being 13.23 nm and 12.51 nm, respectively. Combined with the SEM analysis of the  $\text{Ni}_x\text{Mn-LDHs}$  samples, it can be concluded that the morphological difference of the precursors can seriously affect the exposure opportunities of catalyst surface active species. This result revealed that higher catalytic activity was achieved to  $\text{Ni}_5\text{Mn-LDO}$  and  $\text{Ni}_6\text{Mn-LDO}$  catalysts. In order to further investigate the phase composition of the catalyst, HRTEM analysis of  $\text{Ni}_x\text{Mn-LDO}$  was carried out in this part. As compiled in Fig. 7(e–h), all catalysts exhibited two different types of lattice fringes, the lattice spacings of 0.239 nm can be assigned to the (222) plane of NiO, and the lattice spacings of 0.211 nm correspond to  $\text{Ni}_6\text{MnO}_8$  (400) plane,<sup>39</sup> which is in good accordance with the XRD analysis results.

**3.2.4  $\text{N}_2$  adsorption–desorption analysis.** The  $\text{N}_2$  adsorption–desorption characterization was used to test the surface area and pore structure distribution of the  $\text{Ni}_x\text{Mn-LDO}$  catalysts. Related results were shown in Fig. 8(a and b) and Table 2. As illustrated in Fig. 8(a), four catalysts exhibit the same type of desorption curves belonging to type IV isotherms with type H2 hysteresis loops, which fully indicated the existence of mesoporous structure.<sup>34,41,42</sup> The corresponding pore diameter distribution was considered in Fig. 8(b). It can be observed that the sample of  $\text{Ni}_3\text{Mn-LDO}$  has a wide pore size distribution with 10–40 nm. Comparatively, the pore size distribution range of  $\text{Ni}_x\text{Mn-LDO}$  ( $x = 4, 5, 6$ ) was greatly reduced and more concentrated around 5–20 nm, which implies that the latter had a more uniform pore size structure.<sup>43</sup> Moreover, the surface area was obtained in Table 2 by calculating with BET method. Obviously, the surface area value of samples was sensitive to the proportion of Ni/Mn and displayed an upward trend from 41 to  $83\text{ m}^2\text{ g}^{-1}$  with enhancement of Ni/Mn molar ratios. Combined with TEM analysis, thus may be related to the uniform distribution of the active species particles. Generally, a large specific surface area can provide maximum active and adsorption sites, which might be favorable to the improvement of catalytic activity.<sup>44</sup> The  $\text{Ni}_5\text{Mn-LDO}$  catalyst presented a large specific surface area ( $70\text{ m}^2\text{ g}^{-1}$ ), suitable average pore diameter (9.35

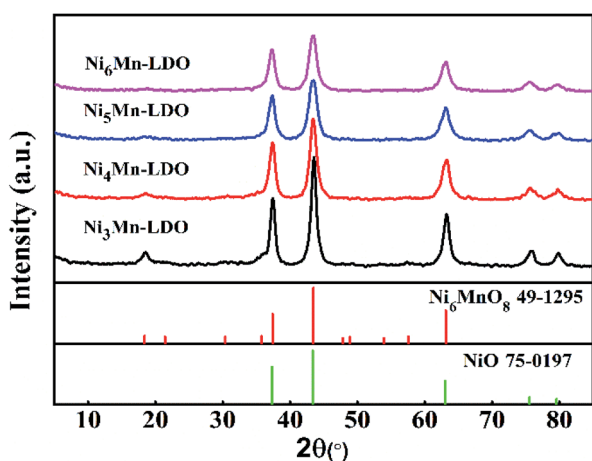


Fig. 6 XRD patterns of  $\text{Ni}_x\text{Mn-LDO}$  ( $x = 3, 4, 5, 6$ ) catalysts.



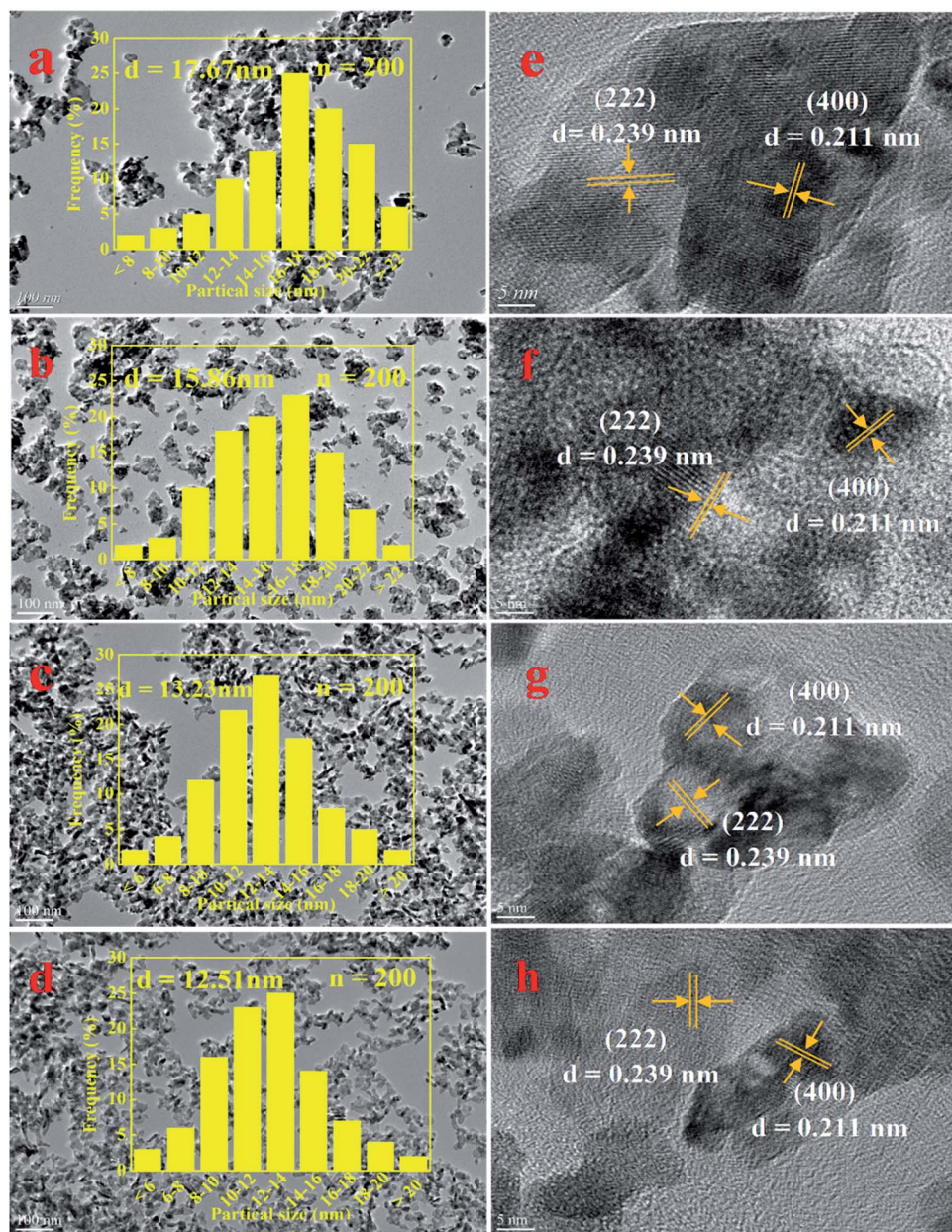


Fig. 7 (a–d) TEM and (e–h) HRTEM images of Ni<sub>x</sub>Mn-LDO catalysts: (a, e) Ni<sub>3</sub>Mn-LDO; (b, f) Ni<sub>4</sub>Mn-LDO; (c, g) Ni<sub>5</sub>Mn-LDO; (d, h) Ni<sub>6</sub>Mn-LDO.

nm) and pore volume ( $0.17 \text{ cm}^3 \text{ g}^{-1}$ ), resulting in the desirable SCR activity.<sup>43,45</sup>

**3.2.5 NH<sub>3</sub>-TPD analysis.** In NH<sub>3</sub>-SCR reaction, the adsorption and activation of NH<sub>3</sub> by catalyst is an important factor affecting the DeNO<sub>x</sub> action.<sup>41,46</sup> Therefore, surface acidity of all catalysts was measured by NH<sub>3</sub>-TPD experiment (Fig. 9). As can be seen, the four catalysts exhibited two similar NH<sub>3</sub> desorption peaks in the test temperature range with 50–560 °C. The peak at lower temperature with 50–300 °C was considered to be the desorption of weak acid sites on the catalyst surface, which can be attributed to NH<sub>4</sub><sup>+</sup> ions bound to the Brønsted acid sites.<sup>41,47,48</sup> The desorption peak at high temperature with 350–550 °C was usually caused by coordinated NH<sub>3</sub> molecular originating from the Lewis acid sites with stronger acidity, which

might be because that the complex formed by the NH<sub>3</sub> lone pair electron coordination possess stronger thermal stability than the NH<sub>4</sub><sup>+</sup> ions.<sup>49–52</sup> Since the desorption peak positions of the four catalysts were similar, so the relative acid amount of catalysts were normalized by the peak area of the desorbed peak and shown in Table 3. Noticeably, the amount of acid of Ni<sub>x</sub>Mn-LDO catalysts at the weak acid sites decreased and increased at strong acid sites gradually with the increase of Ni/Mn molar ratios, which suggest that the acid sites on the surface of catalyst are related to the relative distribution of Ni and Mn elements. In addition, Ni<sub>5</sub>Mn-LDO catalyst displayed more amount of acid than others, which implied that more NH<sub>3</sub> can be adsorbed on the catalyst surface and further promote the low-temperature NH<sub>3</sub>-SCR reaction. What's more, the acid





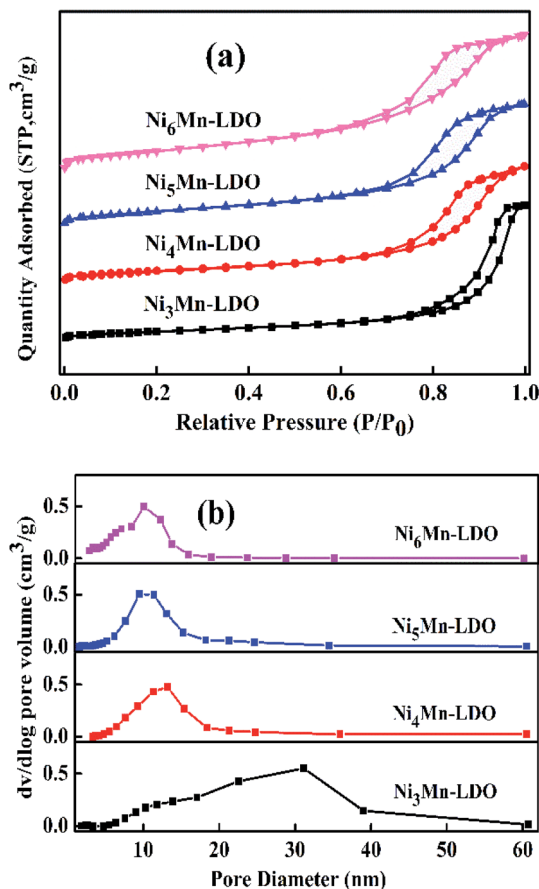


Fig. 8 (a)  $N_2$  adsorption-desorption isotherms and (b) pore diameter distribution of  $Ni_xMn-LDO$  ( $x = 3, 4, 5, 6$ ) catalysts.

Table 2 The information of surface area and pore structure distribution of  $Ni_xMn-LDO$  ( $x = 3, 4, 5, 6$ ) catalysts

Samples	$S_{BET}$ ( $m^2 g^{-1}$ )	$D_{BJH}$ (nm)	$V_{total}$ ( $cm^3 g^{-1}$ )
$Ni_3Mn-LDO$	41	17.71	0.23
$Ni_4Mn-LDO$	53	11.19	0.15
$Ni_5Mn-LDO$	70	9.35	0.17
$Ni_6Mn-LDO$	83	7.73	0.19

amount distribution of the four catalysts has a similar order to the size of the metal oxide particles. This results suggest that the uniform distribution of the active species particles is an crucial factor on the full exposure of the acid sites on the catalyst surface.

**3.2.6  $H_2$ -TPR analysis.** The redox capability of the catalysts is another indispensable indicator for evaluating catalytic activity in the  $NH_3$ -SCR reaction. The  $H_2$ -TPR results of four catalysts was demonstrated in Fig. 10. There are two reduction peaks on the curves, the first peak appears in the temperature range of 350–395 °C, corresponding to the process:  $Ni_6MnO_8 + H_2 \rightarrow MnO-NiO + H_2O$ , which represents the reduction from  $MnO_x$  species to  $MnO$ ; the second step of the reduction takes place at 405–446 °C and can be attributed to:  $MnO-NiO + H_2$

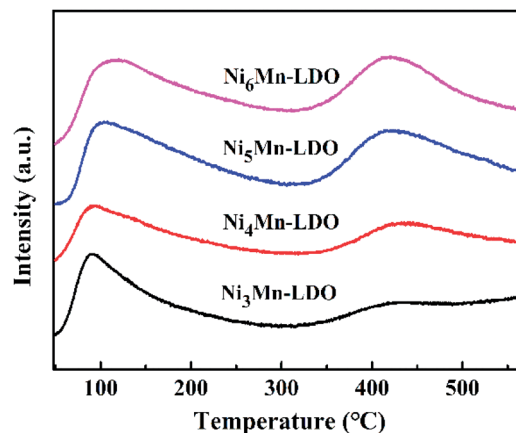


Fig. 9  $NH_3$ -TPD curves of  $Ni_xMn-LDO$  ( $x = 3, 4, 5, 6$ ) catalysts.

Table 3 Relative amount and proportion of  $NH_3$  desorption for  $Ni_xMn-LDO$  ( $x = 3, 4, 5, 6$ ) catalysts

Samples	Total amount of $NH_3$ desorption	Proportion of $NH_3$ desorption (%)	
		Weak acid site (50–300 °C)	Strong acid site (350–550 °C)
$Ni_3Mn-LDO$	1.00	88.4	11.6
$Ni_4Mn-LDO$	1.08	67.4	32.6
$Ni_5Mn-LDO$	2.10	62.2	37.8
$Ni_6Mn-LDO$	1.91	59.6	40.4

$\rightarrow Ni + MnO + H_2O$ , indicating the transformation of  $NiO$  species to  $Ni$ .<sup>20,39,53,54</sup> Likewise, it can be observed that the positions of the reduction peaks of four catalysts have changed significantly. Obviously, with the increase of  $Ni/Mn$  molar ratios, the peak position of  $Ni_xMn-LDO$  catalysts shifted to the lower temperature, which suggests that the reduction process is becoming easier. Especially,  $Ni_5Mn-LDO$  catalyst shows the lowest reduction peak temperature and correspondingly the highest catalytic activity. The reason may be related to the uniform dispersion of the reducible species on the surface of

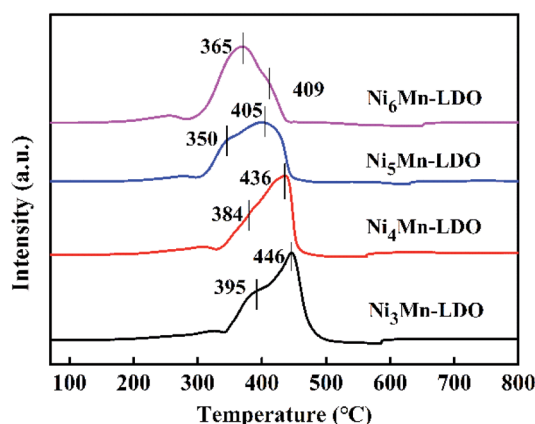


Fig. 10  $H_2$ -TPR curves of  $Ni_xMn-LDO$  ( $x = 3, 4, 5, 6$ ) catalysts.



catalyst, which can be confirmed by BET and TEM analysis results.

**3.2.7 SO<sub>2</sub> & H<sub>2</sub>O resistance and stability tests.** In actual NH<sub>3</sub>-SCR reaction conditions, the flue gas inevitably contains trace amount of SO<sub>2</sub> and H<sub>2</sub>O, which can greatly affect the DeNO<sub>x</sub> performance of catalyst at low temperature.<sup>55</sup> So, it is necessary to study the SO<sub>2</sub> & H<sub>2</sub>O resistance of catalyst. Fig. 11(a) illustrates SO<sub>2</sub> & H<sub>2</sub>O resistance tests of Ni<sub>5</sub>Mn-LDO catalyst at 240 °C under a GHSV of 45 000 h<sup>-1</sup>. When SO<sub>2</sub> was introduced into the mixture gases, NO<sub>x</sub> conversion of catalyst started to decrease from the original 98% and gradually tended to a smooth value of 88%. And then adding 10 vol% H<sub>2</sub>O into reaction system, where SO<sub>2</sub> and H<sub>2</sub>O coexisted, DeNO<sub>x</sub> performance of this catalyst was further suppressed. Upon removing SO<sub>2</sub>, NO<sub>x</sub> conversion has risen and remained at around 86%. After further closing H<sub>2</sub>O, NO<sub>x</sub> conversion can be restored to 92%, but still slightly lower than the initial value, which may be due to the deposition of ammonium sulfate on the surface of the catalyst. In general, Ni<sub>5</sub>Mn-LDO catalyst presented considerable SO<sub>2</sub> & H<sub>2</sub>O resistance.

It is also an important challenge for the catalysts to maintain high DeNO<sub>x</sub> efficiency for a long time in the NH<sub>3</sub>-SCR reaction. The stability test of Ni<sub>5</sub>Mn-LDO catalyst was depicted in

Fig. 11(b). As can be seen, NO<sub>x</sub> conversion has only a slight decrease with the extension of the test time throughout the test range, which suggest that the catalyst has good stability.

## 4. Conclusions

A series of Ni<sub>x</sub>Mn-LDO ( $x = 3, 4, 5, 6$ ) metal oxides were obtained by calcining Ni<sub>x</sub>Mn-LDHs ( $x = 3, 4, 5, 6$ ) precursors at 400 °C and were employed as NH<sub>3</sub>-SCR catalysts. The catalytic results presented that all catalysts exhibited good SCR activity and N<sub>2</sub> selectivity, which might be attributed to the formation of Ni<sub>6</sub>MnO<sub>8</sub> active species. In addition, catalytic performance of Ni<sub>x</sub>Mn-LDO catalysts were sensitive to the Ni/Mn molar ratios, which were related to the redox capacity and surface acidity caused by dispersion. In fact, it was demonstrated that the distribution of active species particles on the surface of catalyst could be improved by optimizing the morphology of the Ni<sub>x</sub>Mn-LDHs precursors. This study provides a new method to prepare catalyst with good dispersion properties by calcining LDHs precursors.

## Conflicts of interest

There are no conflicts of interest to declare.

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## Notes and references

- 1 K. Liu, H. He, Y. B. Yu, Z. D. Yan, W. W. Yang and W. P. Shan, *J. Catal.*, 2019, **369**, 372–381.
- 2 D. Damma, P. R. Ettireddy, B. M. Reddy and P. G. Smirniotis, *Catalysts*, 2019, **9**, 349.
- 3 P. Forzatti, I. Nova and E. Tronconi, *Angew. Chem., Int. Ed.*, 2009, **121**, 8516–8518.
- 4 Z. B. Zhi, X. Ning, F. Zhou, B. Yang, Y. W. Tu, J. Jin, W. Lu and Z. H. Liu, *RSC Adv.*, 2018, **8**, 21915–21925.
- 5 Z. H. Chen, Q. Yang, H. Li, X. H. Li, L. F. Wang and S. C. Tsang, *J. Catal.*, 2010, **276**, 56–65.
- 6 R. H. Gao, D. S. Zhang, X. G. Liu, L. Y. Shi, P. Maitarad, H. R. Li, J. P. Zhang and W. G. Cao, *Catal. Sci. Technol.*, 2013, **3**, 191–199.
- 7 F. Kapteijn, L. Singoredjo, A. Andreini and J. A. Moulijn, *Appl. Catal., B*, 1994, **3**, 173–189.
- 8 X. L. Tang, J. M. Hao, W. G. Xu and J. H. Li, *Catal. Commun.*, 2007, **8**, 329–334.
- 9 X. F. Tang, J. H. Li, L. Sun and J. M. Hao, *Appl. Catal., B*, 2010, **99**, 156–162.
- 10 J. H. Li, H. Z. Chang, L. Ma, J. M. Hao and R. T. Yang, *Catal. Today*, 2011, **175**, 147–156.
- 11 C. Liu, J. W. Shi, C. Gao and C. M. Niu, *Appl. Catal., A*, 2016, **522**, 54–69.

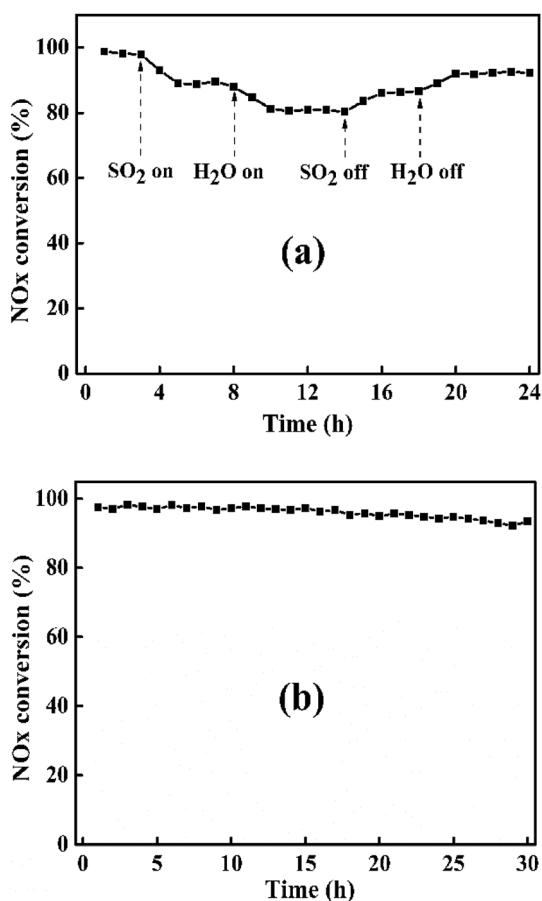


Fig. 11 (a) SO<sub>2</sub> & H<sub>2</sub>O resistance test and (b) stability test of Ni<sub>5</sub>Mn-LDO catalyst at 240 °C. Reaction conditions: [NO] = [NH<sub>3</sub>] = 600 ppm, [O<sub>2</sub>] = 5 vol%, SO<sub>2</sub> = 100 ppm (when used), 10 vol% H<sub>2</sub>O (when used), N<sub>2</sub> as balance gas, GHSV = 45 000 h<sup>-1</sup>.





- 12 M. Casapu, O. Kröcher and M. Elsener, *Appl. Catal., B*, 2009, **88**, 413–419.
- 13 M. Kang, E. D. Park, J. M. Kim and J. E. Yie, *Catal. Today*, 2006, **111**, 236–241.
- 14 Z. H. Chen, X. H. Li, X. Gao, Y. B. Jiang, Y. X. Lü, F. R. Wang and L. F. Wang, *J. Catal.*, 2009, **30**, 4–6.
- 15 Z. H. Lian, F. D. Liu, H. He, X. Y. Shi, J. S. Mo and Z. B. Wu, *Chem. Eng. J.*, 2014, **250**, 390–398.
- 16 Y. Wan, W. R. Zhao, Y. Tang, L. Li, H. J. Wang, Y. L. Cui, J. L. Gu, Y. S. Li and J. L. Shi, *Appl. Catal., B*, 2014, **148**, 114–122.
- 17 L. Q. Chen, X. Y. Niu, Z. B. Li, Y. L. Dong, Z. P. Zhang, F. L. Yuan and Y. J. Zhu, *Catal. Commun.*, 2016, **85**, 48–51.
- 18 H. R. Li, D. S. Zhang, P. P. Maitarad, L. Y. Shi, R. H. Gao, J. P. Zhang and W. G. Cao, *Chem. Commun.*, 2012, **48**, 10645–10647.
- 19 Y. L. Han, J. C. Mu, X. Y. Li, J. S. Gao, S. Y. Fan, F. Tan and Q. D. Zhao, *Chem. Commun.*, 2018, **54**, 9797–9800.
- 20 S. X. Cai, D. S. Zhang, L. Y. Shi, J. Xu, L. Zhang, L. Huang, H. R. Li and J. P. Zhang, *Nanoscale*, 2014, **6**, 7346.
- 21 J. Xiong, Q. Q. Wu, X. L. Mei, J. Liu, Y. C. Wei, Z. Zhao, D. Wu and J. M. Li, *ACS Catal.*, 2018, **8**, 7915–7930.
- 22 Y. F. Zhao, S. He, M. Wei, D. G. Evans and X. Duan, *Chem. Commun.*, 2010, **46**, 3031–3033.
- 23 H. Yan, M. Wei, J. Ma, F. Li, D. G. Evans and X. Duan, *J. Phys. Chem. A*, 2009, **113**, 6133–6141.
- 24 Q. Wang and D. O'Hare, *Chem. Rev.*, 2012, **112**, 4124–4155.
- 25 X. Wu, R. N. Wang, Y. L. Du, C. L. Zhou, H. Meng and X. M. Xie, *Mol. Catal.*, 2019, **467**, 150–160.
- 26 W. Y. Zhou, Q. Y. Tao, F. Sun, X. B. Cao, J. F. Qian, J. Xu, M. Y. He, Q. Chen and J. L. Xiao, *J. Catal.*, 2018, **361**, 1–11.
- 27 J. W. Zhao, J. L. Chen, S. M. Xu, M. F. Shao, Q. Zhang, F. Wei, J. Ma, M. Wei, D. G. Evans and X. Duan, *Adv. Funct. Mater.*, 2014, **24**, 2938–2946.
- 28 X. J. Li, M. Y. Xin, S. Guo, T. H. Cai, D. F. Du, W. Xing, L. M. Zhao, W. Y. Guo, Q. Z. Xue and Z. F. Yan, *Electrochim. Acta*, 2017, **253**, 302–310.
- 29 W. Ma, R. Z. Ma, J. H. Wu, P. Z. Sun, X. H. Liu, K. C. Zhou and T. Sasaki, *Nanoscale*, 2016, **8**, 10425–10432.
- 30 S. Miyata, *Clays Clay Miner.*, 1975, **23**, 369–375.
- 31 Q. H. Yan, S. N. Chen, L. Qiu, Y. S. Gao, D. O'Hare and Q. Wang, *Dalton Trans.*, 2017, **47**, 2992–3004.
- 32 R. N. Wang, X. Wu, C. L. Zhou, X. J. Li and Y. L. Du, *Catalysts*, 2018, **8**, 384.
- 33 Q. H. Yan, S. N. Chen, C. Zhang, D. O'Hare and Q. Wang, *J. Colloid Interface Sci.*, 2018, **526**, 63–74.
- 34 S. P. Mo, S. D. Li, W. H. Li, J. Q. Li, J. Y. Chen and Y. F. Chen, *J. Mater. Chem. A*, 2016, **4**, 8113–8122.
- 35 Z. P. Xu, L. Li, C. Y. Cheng, R. G. Ding and C. H. Zhou, *Appl. Clay Sci.*, 2013, **74**, 102–108.
- 36 X. Xiang, H. I. Hima, H. Wang and F. Li, *Chem. Mater.*, 2007, **20**, 1173–1182.
- 37 X. Wu, Y. L. Feng, Y. L. Du, X. Z. Liu, C. L. Zou and Z. Li, *Appl. Surf. Sci.*, 2019, **467**, 802–810.
- 38 J. J. Zhang, R. R. Hu, P. Dai, Z. M. Bai, X. X. Yu, M. Z. Wu and G. Li, *J. Mater. Sci.: Mater. Electron.*, 2018, **29**, 7510–7518.
- 39 Y. G. Zhang, Z. F. Qin, G. F. Wang, H. Q. Zhu, M. Dong, S. N. Li, Z. W. Wu, Z. K. Li, Z. H. Wu, J. Zhang, T. D. Hu, W. B. Fan and J. G. Wang, *Appl. Catal., B*, 2013, **129**, 172–181.
- 40 H. Meng, J. N. Liu, Y. L. Du, B. H. Hou, X. Wu and X. M. Xie, *Appl. Catal., A*, 2019, **119**, 101–105.
- 41 L. Zhang, L. Y. Shi, L. Huang, J. P. Zhang, R. H. Gao and D. S. Zhang, *ACS Catal.*, 2014, **4**, 1753–1763.
- 42 Q. Liu, J. J. Gao, F. N. Gu, X. P. Lu, Y. J. Liu, H. F. Li, Z. Y. Zhong, B. Liu, G. W. Xu and F. B. Su, *J. Catal.*, 2015, **326**, 127–138.
- 43 X. Wu, X. J. Li, Y. L. Du, R. N. Wang, X. M. Guo and B. H. Hou, *ChemCatChem*, 2019, **11**, 1993–2003.
- 44 S. Ali, L. Q. Chen, Z. B. Li, T. R. Zhang, R. Li, S. U. H. Bakhtiar, X. S. Leng, F. L. Yuan, X. Y. Niu and Y. J. Zhu, *Appl. Catal., B*, 2018, **236**, 25–35.
- 45 Y. Li, G. Q. Li, Y. Lu, W. X. Hao, Z. Wei, J. Liu and Y. F. Zhang, *Mol. Catal.*, 2018, **445**, 21–28.
- 46 L. Lietti, I. Nova, G. Ramis, L. D. Acqua, G. Busca, E. Giamello, P. Forzatti and F. Bregani, *J. Catal.*, 1999, **187**, 419–435.
- 47 Z. B. Wu, B. Q. Jiang, Y. Liu, H. Q. Wang and R. B. Jin, *Environ. Sci. Technol.*, 2007, **41**, 5812–5817.
- 48 C. Fang, D. S. Zhang, L. Y. Shi, R. H. Gao, H. R. Li, L. P. Ye and J. P. Zhang, *Catal. Sci. Technol.*, 2013, **3**, 803–811.
- 49 S. Roy, B. Viswanath, M. S. Hegde and G. Madras, *J. Phys. Chem. C*, 2008, **112**, 6002–6012.
- 50 Y. Peng, W. W. Yu, W. K. Su, X. Huang and J. H. Li, *Catal. Today*, 2015, **242**, 300–307.
- 51 M. Mhamdi, S. Khaddar-Zine and A. Ghorbel, *Appl. Catal., A*, 2009, **357**, 42–50.
- 52 L. Zhang, D. S. Zhang, J. P. Zhang, S. X. Cai, C. Fang, L. Huang, H. R. Li, R. H. Gao and L. Y. Shi, *Nanoscale*, 2013, **5**, 9821–9829.
- 53 L. Christel, A. Pierre and D. A.-M. R. Abel, *Thermochim. Acta*, 1997, **306**, 51–59.
- 54 C. Laberty, J. Pielaszek, P. Alphonse and A. Rousset, *Solid State Ionics*, 1998, **110**, 293–302.
- 55 K. Skalska, J. S. Miller and S. Ledakowicz, *Sci. Total Environ.*, 2010, **408**, 3976–3989.

