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Pyrazolo-fused 4-azafluorenones as key reagents for the synthesis of fluorescent dicyanovinylidene-substituted derivatives†

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A green process to access pyrazolo-fused 4-azafluorenones (indeno[1,2-*b*]pyrazolo[4,3-*e*]pyridines, IPP) **4a–x** via the three-component reaction between indan-1,3-dione (**1**), benzaldehydes **2** and 5-amino-1-arylpyrazoles **3** is described. These compounds were successfully used as precursors of the novel dicyanovinylidene derivatives **7a–d** containing different acceptor (A) or donor (D) aryl groups at position 4 of its fused system. The structures of products obtained (**4a–x** and **7a–d**) were determined based on NMR experiments, HRMS analysis, and X-ray diffraction studies for **7b**. The photophysical and computational studies of **7a–d** showed that these products are modulable ICT fluorophores, even some preliminary tests revealed that these compounds could be used as fluorescent chemodosimeters for cyanide detection.

1. Introduction

In recent years, the development of a highly efficient atom and step economic synthesis of fused aza-heterocycles to yield biologically active compounds has been actively pursued, and thus has become an important area of research in organic and medicinal chemistry.^{1–3} In particular, pyrazolo[3,4-*b*]pyridines (**PP**) are of biomedical importance and have been extensively studied for their broad biological activity.^{4–6} Likewise, indeno[1,2-*b*]pyridines (4-azafluorenones, IP) have shown potential as anticancer, antioxidant, antihistamine and antidepressant agents.^{7,8} Both 4-azafluorenones and pyrazolo[3,4-*b*]pyridines have found applications in materials science due to their amazing photophysical properties.^{9,10} Thus, the development of efficient methods for the synthesis and functionalization of fused systems of these two structural moieties (*i.e.*, pyrazolo-fused 4-azafluorenones) is highly desirable (Fig. 1).

Numerous approaches for the synthesis of **PP**^{11,12} and their fused derivatives have been reported,^{13,14} those reactions involved the interaction of 1,3-bis-electrophilic compounds with *N*-substituted 5-aminopyrazoles.^{11–14} Moreover, multi-component reactions (MCRs) are used in the synthesis of these compounds by the formation *in situ* of the bis-

electrophilic intermediate. This approach has been widely used in diversity-oriented synthesis (DOS) of biologically active heterocyclic compounds.^{15,16} However, there are few reports addressing the preparation of indeno-fused **PP**, and those synthetic procedures have some limitations (*e.g.*, moderate yields, the use of additives or catalysts, ionic liquids as solvent, and reduced substrate scope). Indeed, the most remarkable limitation is the substrate scope, since reactions have been restricted to the use of 5-amino-1-phenyl-3-methylpyrazole (**3a**) with only a few reported exceptions (Scheme 1a).^{17–20} One of the few examples using a different amine was done from our group, we reported an initial study of the crystal structure of indeno[1,2-*b*]pyrazolo[4,3-*e*]pyridines (IPP) **4p** and **4u** which was prepared by the three-component reaction between indan-1,3-dione (**1**), benzaldehydes **2a** or **2f**, and 5-amino-3-*tert*-butyl-1-(4-chlorophenyl)pyrazole (**3c**) in good yields. Those reactions were achieved at 80 °C under microwave irradiation (MW) using triethylamine (Et₃N) as a catalyst in water (Scheme 1b).²¹

Continuing with the development of synthetic methods to obtain pyrazole-fused aza-heterocycles^{22–25} along with our

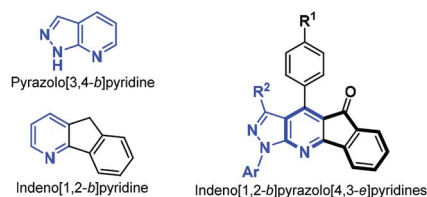
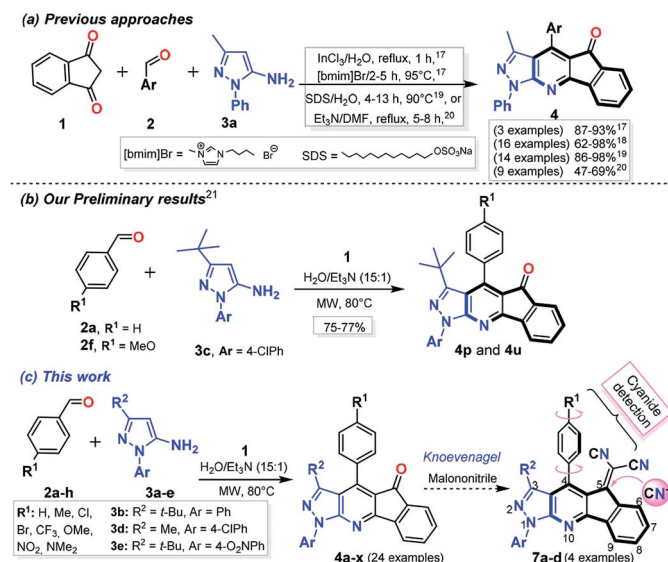


Fig. 1 Pyrazolo[3,4-*b*]pyridine, indeno[1,2-*e*]pyridine and their combined-fused derivatives. The 4-azafluorene moiety is remarked in a bold line.

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Scheme 1 Synthesis of indeno[1,2-*b*]pyrazolo[4,3-*e*]pyridine-5-ones 4 and 4a-*x*.

interest to improve the scope of our preliminary results by using different starting 5-aminopyrazoles, we report an extension of the MW-assisted synthesis of indeno[1,2-*b*]pyrazolo[4,3-*e*]pyridines 4a-*x* via the respective multicomponent reaction using the 5-amino-1-arylpyrazoles 3a-*e* of diverse reactivity. It is important to note that the electron-withdrawing aryl groups at position 1 of the starting amines (*i.e.*, 3c-*e*) decrease its reactivity towards cyclocondensation reactions,²² which would complicate the synthesis of some IPP 4a-*x* (Scheme 1c). In addition, IPP derivatives have a structural analogy with the fluorene core (Fig. 1), whose derivatives have important fluorescent properties with applications on the design of OLEDs and organic transistors.²⁶⁻²⁹ Due to the structural features of ketones 4a-*x*, the post-functionalization reactions and the physicochemical study of their respective functional products are of great interest. The synthesis of the dicyanovinylidene derivatives 7a-*d* and their photophysical properties with an initial application in fluorescent probes were also included, which has been an area that has been recently studied in our group (Scheme 1c).^{22,30-33}

2. Results and discussion

2.1. Synthesis

Starting from our preliminary and promising results of the synthesis of IPP 4p and 4u,²¹ we have studied a range of reaction using different starting 5-aminopyrazoles 3a-*e* and benzaldehydes 2a-*h* (Scheme 1c). Our study of these polycyclic compounds synthesis began by optimization of the reaction conditions (conventional or MW heating, solvent, base, temperature, and reaction time) by using as a model reaction an equimolar mixture of indan-1,3-dione (1), benzaldehyde (2a) and 5-amino-1-phenyl-3-methyl-1*H*-pyrazol (3a). As expected, the reaction gave the desired IPP 4a in good yield under the same conditions reported in our previous work (in H₂O : Et₃N at

80 °C for 10 min under MW²¹), and yield was improved by prolonging the reaction time, but not the temperature (Table 1, entries 1-3). Heating under reflux in different solvents using triethylamine as a catalyst offered the product 4a but in low to moderate yields (Table 1, entries 4-6). We found out that higher temperatures favors formations of the IPP 4a,²⁰ while lower temperatures result in low yield (Table 1, entry 5 vs. 6).³⁴

When the reaction was carried out without base or other bases (KOH or K₂CO₃) in water or ethanol under MW, the formation of 4a was diminished (Table 1, entries 7-10). Consequently, the optimal conditions were set to obtain 4a in a similar way to those reported in our preliminary study (Table 1, entry 2).²¹ In general, low yields for the formation of 4a were observed when ethanol was used as a solvent (Table 1, entries 6 and 8). These conditions are closely related to those used for pyrazolo[5,1-*b*]quinazolines synthesis reported by Chebanov *et al.* (by using dimedone, arylaldehydes and a *NH*-pyrazole).³⁵ On the other hand, when the reaction was carried out in only water (Table 1, entry 7), the dihydropyridine 5a was isolated instead of final product 4a. The structure of this intermediate was deduced by NMR spectroscopy and HPLC-HRMS spectrometry, concluding that the Et₃N not only promotes the formation of 5a, but also favors its final oxidation to afford 4a (see Fig. S1 and S2, ESI†).

Once the optimal conditions to obtain 4a were achieved, we explored a range of benzaldehydes 2a-*h* and aminopyrazoles 3a-*e* (prepared and available in our lab²³) in order to test their reactivities and produce the variously substituted IPP 4a-*x*. Thus, the reaction under MW conditions for 10-25 min of an equimolar quantity of precursors 1, 2a-*h* and 3a-*e* gave the expected products 4a-*x* in good yields. Almost all reactions showed a low electronic effect of the substituent groups on the precursor's reactivity, but longer reaction times are required when less reactive amines are used (*e.g.*, to form the products

Table 1 Reaction condition optimization for the synthesis of 4a^a

Entry	Solvent : base	<i>T</i> (°C)	Time <i>t</i>	Yield [%]
1	H ₂ O : Et ₃ N	80 ^b	10 min	68
2	H ₂ O : Et ₃ N	80 ^b	15 min	84
3	H ₂ O : Et ₃ N	100 ^b	15 min	83
4	H ₂ O : Et ₃ N	Reflux ^c	15 min	44
5	DMF : Et ₃ N	Reflux ^c	24 h	60
6	EtOH : Et ₃ N	Reflux ^c	24 h	25
7 ^d	H ₂ O	80 ^b	24 h	—
8	EtOH	80 ^b	15 min	43
9	H ₂ O : KOH	80 ^b	15 min	22
10	H ₂ O : K ₂ CO ₃	80 ^d	15 min	20

^a Reaction conditions: equimolar quantities (0.25 mmol) of 1, 2a and 3a.

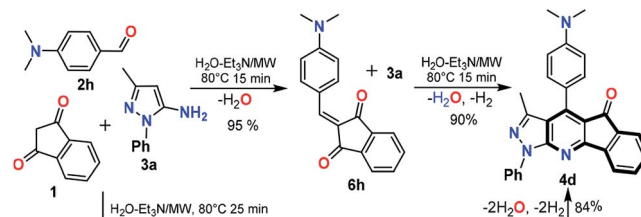
^b Run in 10 mL sealed tube under MW in 0.7 mL of solvent (0.1 equiv. of base) or a mixture H₂O : Et₃N (15 : 1 v/v). ^c Conventional heating in a solvent : Et₃N mixture (2 mL, 15 : 1 v/v). ^d A dihydropyridine intermediate was isolated, see ESI.



4p–x, where amines **3** have electron-withdrawing groups).^{22,23} Moreover, in the reactions using benzaldehydes having electron-donating groups, longer reaction times are required to give the desired products (Scheme 2). The formation of the fused compounds **4a–x** was confirmed by their complete spectral characterization (see Experimental, ESI†).

Casualy, in the reaction using the poorly electrophilic benzaldehyde **2h** under the general conditions by MW (80 °C in H₂O : Et₃N for 15 min) a different compound was formed **6h**. HRMS and NMR analysis confirmed that this product was obtained by the condensation of **1** with **2h** without the participation of the respective amine **3a** (see Experimental in ESI†). Subsequent MW reaction of the intermediate **6h** with one equiv. of **3a** (for 15 min at 80 °C) leads to the desired product **4d** in 90% yield, while a longer reaction time (25 min) directly obtain **4d** via the multicomponent reaction (Scheme 3). These results confirm that the synthesis of the IPP **4a–x** proceeds by the intermediate **6h**, which then reacts with the amine **3a** with subsequent loss of water and hydrogen molecules, in agreement with previous works.^{17–20} Broadly, this methodology was optimized and successfully tested using various substrates allowing its generality and greener approach.

With the ketones **4a–x** in hand, we carried out Knoevenagel reaction with malononitrile to produce the dicyanovinylidene derivatives **7a–d** substituted at position 4 with different donor (D) or acceptor (A) aryl groups. Products **7a–d** were prepared in good to excellent yields using an excess of malononitrile in the presence of titanium chloride (TiCl₄) and pyridine in chlorobenzene at reflux for 24 h (Scheme 4). These reaction conditions are analogous to those reported with fluorenones containing a sterically hindered carbonyl group,^{36–38} such as the structures **4a–x**. Therefore, special reaction conditions were required to

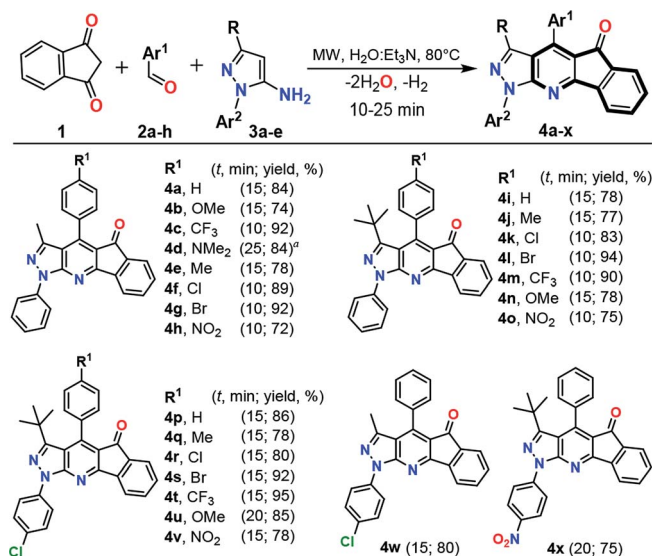


Scheme 3 Synthesis of **4d** via the intermediate **6h**; see ESI† for details.

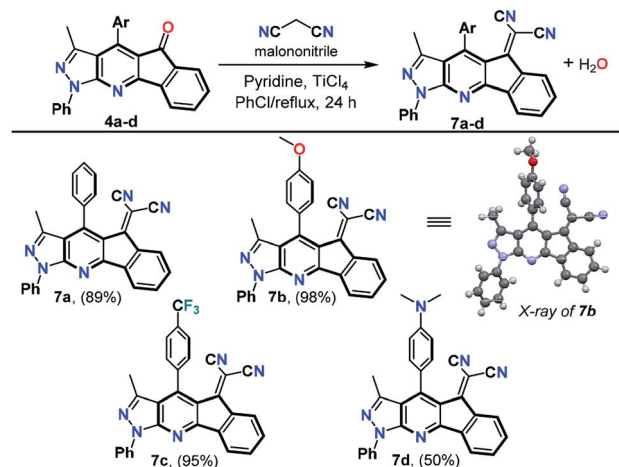
yield compounds **7a–d** since the traditional methods did not work.^{31,32,39} These compounds were synthesized considering the important photophysical properties of pyrazole derivatives^{22,30–32,40,41} and the azafluorenone moiety.⁴² In addition, the dicyanovinylidene is widely used as an acceptor moiety in the design of D–π–A dyes that exhibit intramolecular charge transfer (ICT) photophysical process.^{31,43} The structures of the products **7a–d** were determined by HRMS analysis, ¹H spectroscopy, and ¹³C NMR spectroscopy. Recrystallization of the product **7b** from *N,N*-dimethylformamide (DMF) afforded crystals of suitable size and quality for single-crystal X-ray diffraction analysis (Fig. S12, ESI†).⁴⁴

2.2. Photophysical studies

Solvatochromic studies of dicyanovinylidene derivatives **7a–d** with different electron-donor (D) and electron-acceptor (A) groups were carried out in order to establish if these products can be used as new organic fluorophores (Fig. 2). The UV-vis absorption spectra (Fig. S3†) and fluorescence emission (Fig. 3 and S4†) were taken in solvents of different polarity such as toluene (PhMe), dichloromethane (DCM), acetone, acetonitrile (ACN), and dimethylsulfoxide (DMSO) at 50 μM (Table S1†). The UV-vis spectra of **7a–d** showed two distinctive absorption bands around 300 and 430 nm, the first can be assigned to transitions $\pi \rightarrow \pi^*$ and the second (with the lower intensity) around 430 nm can be attributed to transitions $S_0 \rightarrow$ intramolecular



Scheme 2 MW-assisted synthesis of pyrazolo-fused 4-aza-fluorenones **4a–x**. Reaction conditions: equimolar quantities (0.25 mmol) of **1**, **2a–h** and **3a–e** in a mixture H₂O : Et₃N (0.7 mL, 15 : 1 v/v). Run in 10 mL sealed tube under MW. ^aIntermediate **6h** was isolated after 10 min of reaction; see ESI† for details.



Scheme 4 Synthesis of dicyanovinylidene derivatives **7a–d**. Reaction conditions: **4a–d** (1 equiv.), malononitrile (10 equiv.), pyridine (20 equiv.), and TiCl₄ (10 equiv.); see ESI† for details.



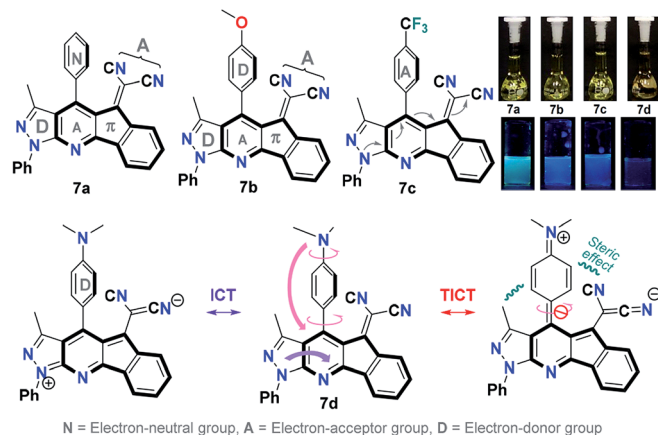


Fig. 2 Structure of 7a–d. Photographs were taken using 50 μ M solutions in DMSO. A hand-held UV lamp under long wavelength ($\lambda = 365$ nm) was used.

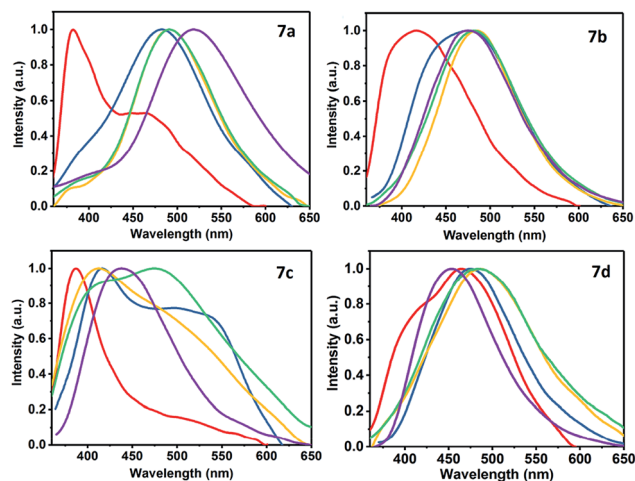


Fig. 3 Normalized fluorescence spectra of 7a–d in different solvents. Red (PhMe), blue (DCM), yellow (acetone), green (MeCN), and purple (DMSO).

charge transfer (ICT). This band is caused by an ICT from the pyrrole-type N atom of the pyrazolic ring to a C $\equiv$ N group, and its intensity reveals a weak ICT effect consistent with its poor solvatochromism in polar solvents (Fig. 2 and S3†).

Regarding the emission spectra of 7a–d, both the quantum yields and fluorescence do not have a significant dependence on the solvent polarity and the 4-aryl group. Products 7b and 7c showed a higher fluorescence emission in DMSO and a hypsochromic shift compared with 7a and 7d. The compound 7d had the higher fluorescence in acetonitrile with respect to 7a–c, showing a bathochromic shift. The compound 7b displayed the highest emission of fluorescence in acetone, whereas in dichloromethane it was the compound 7c. Finally, the use of toluene as a solvent did not show significant fluorescence emission in the heterocycles 7a–d (Fig. S4 and Table S1†). Fig. 3 shows the direct effect of the solvent in each of the products 7a–d, where 7d is the only

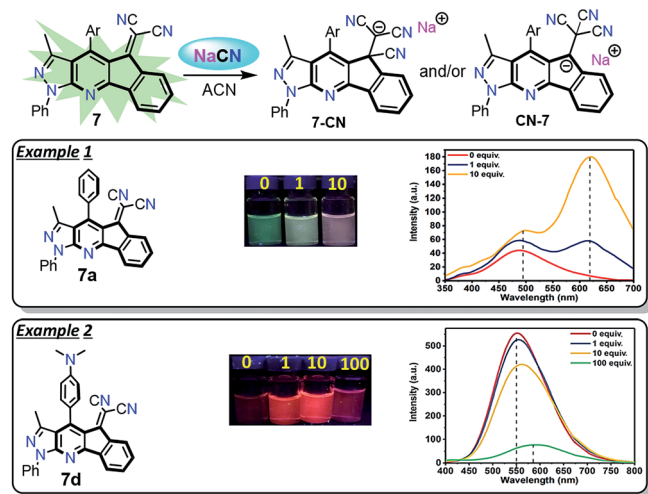
compound that shows a bathochromic shift as the polarity of the solvent increases. This result indicates that there is a more significant ICT in 7d due to the strong donor character of the dimethylamino group (NMe₂) that makes the D– π –A system more efficient.

Although 7d exhibited a greater ICT *versus* 7a–c, the fluorescence emission in DMSO ($\phi = 0.002$) is very low compared to acetonitrile ($\phi = 0.181$). This phenomenon can be the result of a positive solvatokinetic effect that consists in a reduction of the quantum yield due to a high degree of ICT that causes an increase of the speed of non-radiative relaxation of the excited state.⁴⁵ Likewise, a dual emission in the fluorescence spectra was observed mainly in toluene for 7a, 7c and 7d. The first band is assigned to the locally excited state (LE) by ICT processes, while the second to a twisted intramolecular charge transfer (TICT) processes.⁴⁶ These results agree with the structural features of 7a–d, since the TICT phenomena are very sensitive to D $\rightarrow$ A efficacy and strength, the molecular microenvironment and steric effect between groups near the D–A junction.²² The structural relaxation of excited states in weakly polar solvents (reduced interaction with 7a–d in its excited state) allows a greater freedom of rotation in the D–A junctions, thus offering a dual fluorescence (Fig. 2, 3 and S4†).²²

Preliminary UV-vis and fluorescence studies of 7a–d in the presence of different ions were carried out to identify the possible application of these compounds in chemosensors design. From qualitative test of 7a–d in acetonitrile with anions, cyanide ion (CN[−]) caused a significant change in the absorption and fluorescence emission of 7a–d. Besides, the fluorescence emission effect by adding different equivalents of CN[−] to solutions of 7a–d in acetonitrile was evaluated (Fig. S10 and S11 and Table S3, ESI†). From the UV-vis absorption spectra, a decrease of the band round 430 nm was observed after CN[−] was added, showing that a nucleophilic addition occurred on the dicyanovinylidene group (Fig. S10†).³² The emission spectra of 7a–c showed a new band round 620 nm with an increased fluorescence upon CN[−] addition (Fig. S11†). The increase in quantum yield is 6 times higher after adding 10 equiv. of CN[−] (Table S3†). An exception of this trend occurred with 7d, a decrease in the emission around 550 nm with a slight bathochromic shift was observed when CN[−] was added (Fig. S11†). After adding 100 equiv. of CN[−], it was observed that the quantum yield was 8 times lower compared with the initial fluorescence (Table S3† and Scheme 5).

The preliminary results in CN[−] sensing showed that 7d, substituted with the 4-Et₂NPh group, is the least reactive of the compounds 7a–d. Likewise, the regioisomeric addition products 7-CN or CN-7 could be obtained according to their better stabilization. Possibly, CN-7 is most favored when derivatives 7a–c were used since the emission spectra show a high bathochromic shift, which is characteristic for this type of highly-conjugated heteroaromatic anion (Scheme 5 and Fig. S11†). The study towards the design of TURN-ON and TURN-OFF sensors based on the structures of 7a–d is still ongoing.





Scheme 5 Fluorescent probes for CN^- detection. Solutions 50 μM in ACN with CN^- (0 to 100 equiv.). Photograph was taken using a hand-held UV lamp ($\lambda = 365 \text{ nm}$).

2.3. Computational calculation

In order to understand the nature of the electronic transitions present in the compounds **7a–d**, TD-DFT calculations were done in gas phase. The λ_{max} values were obtained from the theoretical UV-vis spectra with a theory level B3LYP/6-31G (d,p). The maximum calculated UV absorptions, the theoretical electronic excitation energies, the calculated oscillator forces and the electronic gas phase transitions are detailed in Table S2.† The band between 340 and 350 nm is assigned mainly to the transition HOMO–5 $\rightarrow$ LUMO in **7a**, HOMO–1 $\rightarrow$ LUMO in **7b**, HOMO–4 $\rightarrow$ LUMO in **7c**, and HOMO–5 $\rightarrow$ LUMO in **7d**, where it is predicted that these transitions are of nature $\pi \rightarrow \pi^*$. According to the diagrams of frontier orbitals of these transitions (Fig. 4 and S5–S9 of ESI†), the electronic nature of the substituent groups of the aryl ring has a significant influence on the ICT towards the dicyanovinylidene group, where **7a**

showed a charge transfer from the 4-phenyl group and the indene moiety.

Compound **7b** is carried out from the *p*-methoxyphenyl group and the phenylpyrazole fragment, while at **7c**, this phenomenon occurs from the entire molecule. Finally, for **7d** this CT is favored mainly from the *p*-dimethylaminophenyl group to dicyanovinylidene group, due to the strong donor character of this substituent. The energy levels of the frontier orbitals of compounds **7a–d** are illustrated in the Fig. 4. These results showed that an electron withdrawing group such as CF_3 could stabilize the energy levels of both the HOMO and LUMO orbitals, due to its inductive effect. On the other hand, electron-donating groups such as OMe and NMe_2 , destabilize both border orbitals.⁴⁷ Additionally, comparing the band gap energy (ΔE) of the compounds **7a–d**, it can be concluded that **7d** has a lower ΔE due to the strong electron-donor character of the dimethylamino group (NMe_2) that favors a higher CT, which makes it less reactive towards nucleophilic addition reactions.

3. Conclusions

In summary, we have developed a MW-assisted method to promote the three-component synthesis of the attractive synthetic scaffold IPP **4a–x** using diverse starting *N*-arylpyrazoles and Et_3N as catalyst. With this green approach using water as a solvent, high yields of all products with short reaction times were obtained. The IPP **4a–x** are of great interest as reagents to prepare important derivatives with biological and optical applications. Thus, novel dicyanovinylidene derivatives **7a–d** were synthesized from precursors **4a–d** and their photophysical properties were studied, which proved that the IPP core is a modular fluorophore acting *via* an ICT phenomenon. Additionally, preliminary UV-vis and fluorescence spectroscopic studies showed that the products **7a–d** could be used as fluorescent chemodosimeters for cyanide detection.

Conflicts of interest

The authors declare no competing financial interest.

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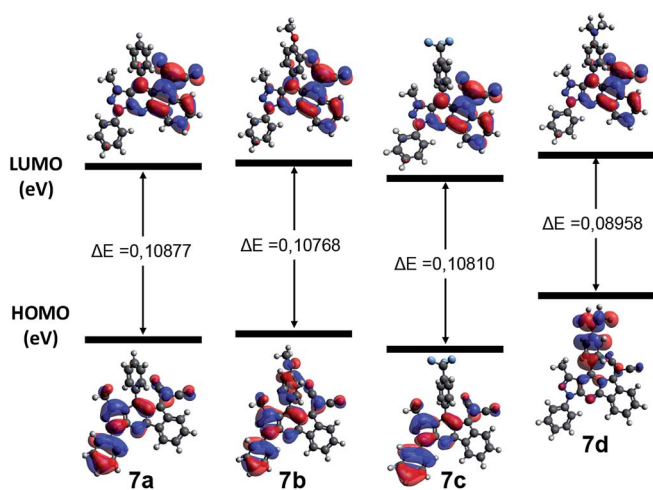


Fig. 4 Frontier orbitals of **7a–d** generated using Avogadro.⁴⁸



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