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A coumarin–dihydroperimidine dye as a fluorescent chemosensor for hypochlorite in 99% water†

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The hypochlorite anion (OCl^-), a reactive oxygen species (ROS), is an important microbicidal agent in the immune system. Accurate and selective detection of OCl^- in environmental and biological samples by a fluorescent molecular sensor is an important subject. All previously reported sensors, however, have suffered from tedious multi-step synthesis for the sensors and the use of large amounts of organic solvents for the analysis. Herein, we report that a coumarin–dihydroperimidine dye prepared by facile condensation behaves as a fluorescent sensor for OCl^- in 99% water. The sensor exhibits weak fluorescence, but OCl^- -selective dehydrogenation of its dihydroperimidine unit creates a strong blue fluorescence. This turn-on fluorescence response facilitates selective and sensitive detection of OCl^- in the physiological pH range. *Ab initio* calculation revealed that the fluorescence enhancement by OCl^- is triggered by intramolecular proton transfer from the coumarin –OH to the imine nitrogen of the formed perimidine moiety.

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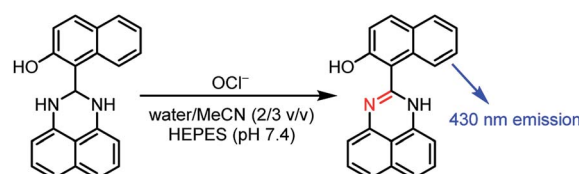
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Introduction

Reactive oxygen species (ROS) play crucial roles in several life functions.¹ Among them, hypochlorous acid (HClO) is one of the most biologically important ROS.² HClO undergoes deprotonation at physiological pH and produces the hypochlorite anion (OCl^-),³ which behaves as a microbicidal agent in the immune system.⁴ OCl^- is produced *in vivo* by the reaction of hydrogen peroxide (H_2O_2) with Cl^- via an enzymatic reaction on myeloperoxidase (MPO).⁵ Controlled generation of OCl^- is necessary to inhibit invading microbes. Uncontrolled OCl^- generation, however, causes several diseases such as neuron degeneration, arthritis, and cancer,⁶ because OCl^- reacts with several biomolecules such as amino acids, proteins, and nucleosides.⁷ In addition, HClO is widely used in daily life for sterilization and disinfection of water supplies, and high residual concentrations of OCl^- in water is hazardous to human and animal health.⁸ Analytical methods that quantitatively detect small amount of OCl^- in environmental and biological samples on inexpensive instrumentations with simple pre-treatment are necessary.

Fluorometric analysis with OCl^- -selective molecular sensors is one promising method for this purpose since this facilitates simple quantification or imaging of OCl^- with a common

fluorescence spectrometer or microscope apparatus.⁹ A number of fluorescent OCl^- sensors have been reported;^{10–17} however, many of them require tedious multi-step procedures for the synthesis of sensors or a solution containing a large amount of organic solvents for sensing due to the low solubility of the sensors in water. Among the previously reported OCl^- sensors, a “dihydroperimidine”-based sensor designed by Goswami *et al.*¹⁸ has the simplest structure, which can be prepared by a facile condensation. As shown in Scheme 1, they synthesized a naphthol–dihydroperimidine dye by the condensation of 1,8-diaminonaphthalene with 1-formyl-2-naphthol as a fluorophore. The sensor shows a sensitive turn-on fluorescence response *via* an OCl^- -selective dehydrogenation of the dihydroperimidine unit. The sensor, however, requires a solution containing 60% MeCN owing to its low solubility in water. Based on this molecular design, Fan *et al.*¹⁹ synthesized a sensor by the condensation of 1,8-diaminonaphthalene with 7-diethylamino-1,4-benzoxazin-2-one as a fluorophore. Although the sensor exhibits a selective and sensitive response towards OCl^- , it still requires a large amount of organic solvent (80%



Scheme 1 A naphthol–dihydroperimidine dye exhibiting a turn-on fluorescence response toward OCl^- .¹⁸

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† Electronic supplementary information (ESI) available: Table S1, Fig. S1–S10, and Cartesian coordinates for the molecules. See DOI: 10.1039/c9ra05533a



DMF) for the sensing. Design of a sensor that can be synthesized by a simple procedure and has a high water solubility is therefore desirable.

We used coumarin as a fluorophore due to its relatively high water solubility,²⁰ high fluorescence quantum yield,²¹ large Stokes shift,²² high stability,²³ and good cell permeability.²⁴ As shown in Scheme 2, the sensor **1**, synthesized by a simple condensation of 1,8-diaminonaphthalene with 8-formyl-7-hydroxy-4-methylcoumarin, is soluble in water containing only 1% organic solvents. The sensor shows a weak fluorescence, but OCl⁻-selective dehydrogenation of its dihydroperimidine unit creates a strong fluorescence at 462 nm. This turn-on response facilitates sensitive detection of OCl⁻. Several spectroscopic analysis and *ab initio* calculations revealed that this turn-on response by OCl⁻ is triggered by intramolecular proton transfer from the coumarin -OH to the imine nitrogen of the formed perimidine unit.

Results and discussion

Synthesis and fluorescence properties of the sensor

The sensor **1** was prepared by the reaction shown in Scheme 2. 8-Formyl-7-hydroxy-4-methylcoumarin prepared by formylation of 7-hydroxy-4-methylcoumarin (yield: 45%)²⁵ and 1,8-diaminonaphthalene were dissolved in EtOH, and the solution was stirred at 80 °C for 2.5 h in an aerated condition. The solid formed was recovered by filtration and washed thoroughly with EtOH, affording **1** as pale pink solids with 69% yield (overall yield: 31%). The purity of **1** was confirmed by ¹H NMR, ¹³C NMR and FAB-MS analysis (Fig. S1–S3, ESI†). **1** is soluble in common organic solvents such as DMSO, CHCl₃, DMF, and MeCN and in aqueous solutions with 1% organic solvents such as DMSO and MeCN. Fig. S4 (ESI†) shows the absorption spectra of 1% MeCN solutions containing different concentrations of **1**. The linear relationship between the absorbance at 325 nm and the concentration of **1** (0–20 μM) indicates that it follows the Beer's law, suggesting that **1** is fully soluble in the solutions. Note that

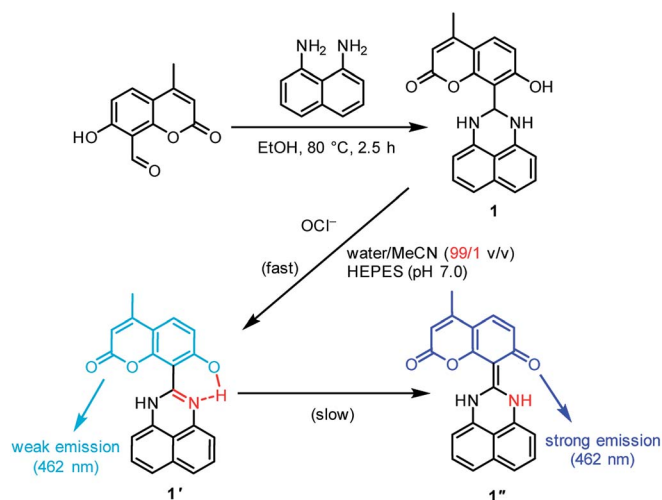
the molar extinction coefficient of **1** at 325 nm was determined to be 10 039 M⁻¹ cm⁻¹.

Fluorescence spectra of **1** (10 μM) were measured in a buffered water/MeCN mixture (99/1 v/v) with pH 7.0 (HEPES 0.1 M) at 25 °C (λ_{ex} = 344 nm). As shown in Fig. 1, **1** itself shows a very weak fluorescence (fluorescence quantum yield, Φ_F = 0.002). In contrast, addition of 50 equiv. of OCl⁻ to the solution followed by stirring for 20 min creates a strong blue fluorescence at 462 nm (Φ_F = 0.082). Other anions (F⁻, Cl⁻, AcO⁻, NO₂⁻, NO₃⁻, ClO₄⁻, and HSO₄⁻), ROS [hydroxyl radical (·OH), singlet oxygen (¹O₂), H₂O₂, superoxide radical (·O₂⁻), and *tert*-butyl hydroperoxide (*t*-BuOOH)], or RNS [NO and peroxyxynitrite (ONOO⁻)], when added to the solution containing **1**, scarcely change the fluorescence spectra, indicating that OCl⁻ selectively triggers fluorescence enhancement of **1**.

Fig. 2a shows the results of fluorescence titration of **1** with OCl⁻. Stepwise addition of OCl⁻ increases the intensity of the 462 nm fluorescence. As shown in Fig. 2b, the change in the ratio of fluorescence intensity at 462 nm (FI/FI₀) with the OCl⁻ concentrations clearly shows linear relationship, indicating that **1** facilitates accurate OCl⁻ sensing at ~100 μM. The lower detection limit was determined to be 3.3 μM based on the signal-to-noise (S/N) ratio using the equation (DL = 3 × SD/S),²⁶ where SD is the standard deviation of blank analysis (SD = 0.19, *n* = 10) and *S* is the slope of the fluorescence intensity *versus* the OCl⁻ concentrations (*S* = 0.18 μM⁻¹). This detection limit (3.3 μM) is lower than the physiological OCl⁻ concentrations (5–25 μM) in the human body,²⁷ suggesting that **1** facilitates sensitive OCl⁻ detection even in high-water-content solution.

Reaction of the sensor with OCl⁻

As shown in Scheme 2, the turn-on fluorescence response of **1** upon addition of OCl⁻ is triggered by the transformation to **1'**, *via* dehydrogenation of the dihydroperimidine moiety of **1**. This transformation is confirmed by ¹H, ¹³C NMR and FAB-MS analysis of a DMSO-d₆ solution containing **1** and OCl⁻ (Fig. S5–S7, ESI†). Partial ¹H NMR charts of **1** and **1'** measured in DMSO-d₆ are shown in Fig. 3, where the 2D COSY spectra were used for the assignment of the respective chemical shifts



Scheme 2 Synthesis of the sensor **1**, and proposed mechanism for selective turn-on fluorescence response by OCl⁻.

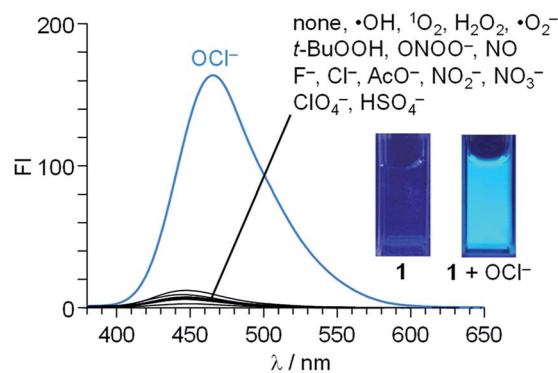


Fig. 1 Fluorescence spectra (λ_{ex} = 344 nm) of **1** (10 μM) in a buffered water/MeCN mixture (99/1 v/v; HEPES 0.1 M, pH 7.0) at 25 °C with 50 equiv. of each respective analytes. All spectra were obtained after stirring the solution for 20 min.



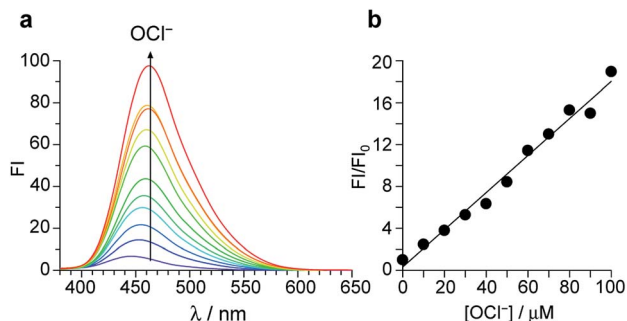


Fig. 2 (a) Change in fluorescence spectra of **1** (10 μM) upon titration with OCl^- in a buffered water/MeCN mixture (99/1 v/v; HEPES 0.1 M, pH 7.0) at 25 $^\circ\text{C}$. (b) Change in the ratio of fluorescence intensity at 462 nm (F_i/F_{i0}) versus the OCl^- concentration. The respective data were obtained after stirring the solution for 20 min.

(Fig. S8 and S9, ESI †). As shown in Fig. 3a, **1** shows an H^a proton at the 2-position of the dihydroperimidine unit at 6.0 ppm. However, as shown in Fig. 3b, addition of OCl^- to the solution leads to almost complete disappearance of the H^a proton. In addition, **1** shows two N–H protons of the dihydroperimidine moiety at 7.0 ppm. After the addition of OCl^- , its chemical shift moves to 7.1 ppm, and its integral value becomes almost 1. These data indicate that H^a and one N–H proton of **1** are removed by the reaction with OCl^- . The dehydrogenation of **1** by OCl^- is confirmed by FAB-MS analysis. As shown in Fig. S3 (ESI †), **1** shows a peak at m/z 344.1 assigned to $[\mathbf{1}]^+$. In contrast, as shown in Fig. S7 (ESI †), a solution containing **1** and OCl^-

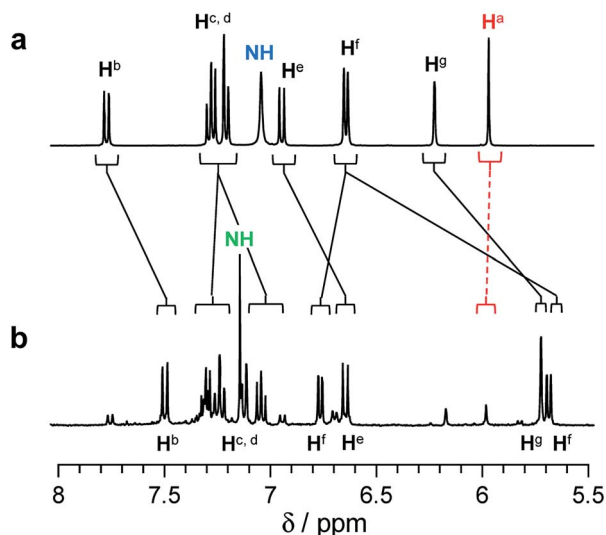
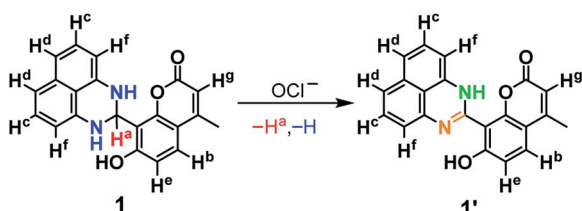


Fig. 3 ^1H NMR chart of **1** (24 mM) measured in $\text{DMSO}-d_6$ (a) without and (b) with 8 equiv of OCl^- (400 MHz, 30 $^\circ\text{C}$).

shows a peak at m/z 342.1 assigned to the dehydrogenated product $[\mathbf{1}']^+$. These NMR and FAB-MS data clearly suggest that dehydrogenation of the dihydroperimidine moiety of **1** via the oxidation by OCl^- gives **1'** containing the perimidine moiety.

As shown in Scheme 2, the enol-imine form (**1'**) is rapidly produced by dehydrogenation of **1** by OCl^- and shows a weak fluorescence ($\Phi_F = 0.009$). Then, **1'** undergoes tautomerization to the keto-amine form (**1''**) via a proton transfer of the coumarin –OH to the imine nitrogen of the perimidine unit, as often observed for similar *o*-hydroxyl Schiff bases,^{28,29} and exhibits a strong fluorescence ($\Phi_F = 0.082$). This sequence is confirmed by time-dependent changes in the absorption and fluorescence spectra of **1** monitored after addition of OCl^- . As shown in Fig. 4a, addition of OCl^- immediately increases the fluorescence intensity at 462 nm within 1 min (blue to red line), although the intensity is weak. Then, the intensity gradually increases with time and plateaus after 15 min, creating a strong fluorescence, where the emission wavelengths scarcely change during the measurements. This indicates that the reaction of **1** with OCl^- creates two different emitting species. As shown in Fig. 4b, absorption spectrum of **1** also changes immediately after the OCl^- addition within 1 min (blue to red line). Then, the spectrum changes gradually with a decrease in *ca.* 320 nm absorbance and an increase in *ca.* 375 nm absorbance. The isosbestic point at 344 nm clearly indicates that, as shown in Scheme 2, the reaction of **1** with OCl^- rapidly produces

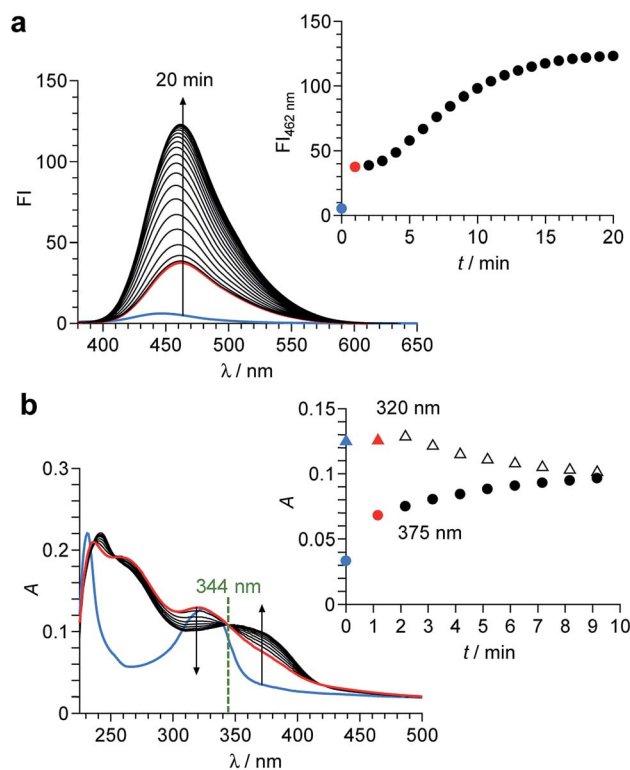


Fig. 4 (a) Time-dependent change in fluorescence spectra ($\lambda_{\text{ex}} = 344$ nm) of **1** (10 μM) in a buffered water/MeCN mixture (99/1 v/v; HEPES 0.1 M, pH 7.0) at 25 $^\circ\text{C}$ after addition of 50 equiv. of OCl^- . The inset shows change in the intensity at 462 nm. (b) Time-dependent change in absorption spectra of **1** after addition of 50 equiv. of OCl^- . The inset shows change in the absorbance at 320 nm and 375 nm.



a weakly-fluorescent enol-imine form (**1'**) and its slow tautomerization by the intramolecular proton transfer creates a strongly-fluorescent keto-amine form (**1''**).

The tautomerization of **1'** to **1''** is promoted by polar water molecules. It is well known that, for the tautomerization of *o*-hydroxy Schiff bases,^{30,31} the enol-imine form is stabilized in less polar solvents such as benzene, while the keto-amine form is stable in polar solvents such as EtOH. Fig. S10 (ESI†) shows the change in fluorescence intensity of **1** after addition of OCl⁻ in MeCN solutions with different water contents. In all solutions, the weakly-fluorescent enol-imine form (**1'**) is rapidly produced by the addition of OCl⁻. The strongly-fluorescent keto-amine form (**1''**) is not produced in low-water-content solutions (30% and 60%), whereas increasing the water content produces **1''**. This indicates that increasing water content increases the polarity of solutions and promotes **1'**-to-**1''** tautomerization. However, as shown in Fig. 3b and S5–S7 (ESI†), ¹H, ¹³C NMR and FAB-MS analysis of the product obtained by the reaction of **1** with OCl⁻ in DMSO-d₆ detected **1'**. This is because the **1'**-to-**1''** tautomerization is not promoted in less polar DMSO. These findings clearly support the **1** → **1'** → **1''** transformation by the reaction of **1** with OCl⁻ in high-water-content solutions, as shown in Scheme 2.

Ab initio calculations

The mechanism for the turn-on fluorescence response of **1** was clarified by *ab initio* calculations. The structures and optical properties of **1**, **1'**, and **1''** species were calculated by the density functional theory (DFT) and the time-dependent DFT (TD-DFT), respectively, within the Gaussian 03 program with water as a solvent. As summarized in Table S1 (ESI†), singlet electronic transition of **1** mainly consists of HOMO → LUMO+2 (S₀ → S₄) transition. Its calculated transition energy (3.76 eV, 330 nm) is

close to the absorption maximum (λ_{max}) of **1** at 323 nm (Fig. 4b, blue line). As shown in Fig. 5 (left), π-electrons of both HOMO and LUMO+2 of **1** are localized on the dihydroperimidine moiety, indicating that photoexcitation of the coumarin fluorophore is not populated. This therefore results in almost no fluorescence of **1**.

As shown in Fig. 5 (center), optimized structure of **1'** has a planar structure, where the coumarin and perimidine units lie on the same plane. It is noted that the structural optimization spontaneously creates an H-bonding interaction between the imine nitrogen and coumarin -OH units, in which the N–O distance of ~2.5 Å indicates strong electrostatic interaction between these units.³² The structural regulation by the H-bonding may create the planar structure. As shown in Table S1 (ESI†), the electronic transition of **1'** mainly consists of HOMO-1 → LUMO+1 (S₀ → S₆) transition. Its energy (3.97 eV, 312 nm) is also close to that for the absorption maximum (320 nm) of **1'** (Fig. 4b, red line). As shown in Fig. 5 (center), relatively large distribution of π-electrons are observed on both HOMO-1 and LUMO+1 for **1'**. This is because the H-bonding interaction of the coumarin -OH increases the electron density of coumarin unit.³³ The enhanced photoexcitation of the coumarin units may therefore result in weak fluorescence of **1'**.

As shown in Fig. 5 (right), optimized structure of **1''** also has a planar structure owing to the C=C bond formation between the coumarin and dihydroperimidine units. Singlet electronic transition of **1''** is mainly contributed by HOMO-1 → LUMO+1 (S₀ → S₄) transition (Table S1, ESI†). Its transition energy (3.64 eV, 340 nm) is also close to that for the absorption band (375 nm) of **1''** (Fig. 4b). As shown in Fig. 5 (right), almost all of the π-electrons of both HOMO-1 and LUMO+1 for **1''** are localized on the coumarin units because complete deprotonation of the coumarin -OH significantly increases the electron density of

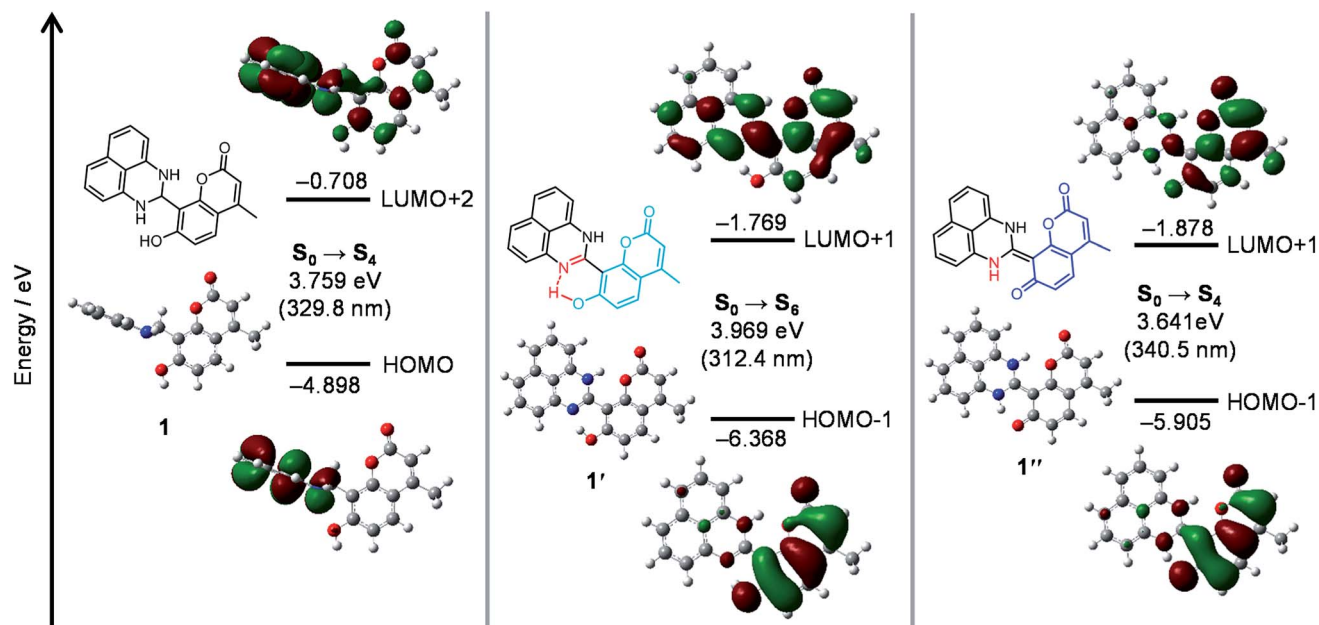


Fig. 5 Energy diagrams and interfacial plots of main molecular orbitals of (left) **1**, (center) **1'** and (right) **1''**, calculated at the DFT level (B3LYP/6-31+G*).

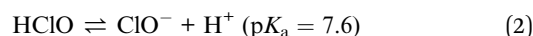
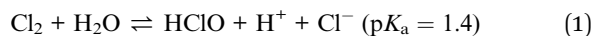


the coumarin unit.³³ This therefore results in strong coumarin fluorescence from **1''**.

The total energies of **1'** and **1''** in water were determined to be -717846.94 and -717850.01 kcal mol⁻¹, respectively. The lower energy of **1''** ($\Delta E = 3.07$ kcal mol⁻¹) indicates that the keto-amine form (**1''**) is indeed more stable in water than the enol-imine form (**1'**). This further supports the **1** \rightarrow **1'** \rightarrow **1''** transformation by the reaction of **1** with OCl⁻ in high-water-content solutions. These DFT results clearly indicate that the dihydroperimidine unit acts as a proton acceptor for the coumarin -OH (Scheme 2). The OCl⁻-triggered formation of the perimidine unit leads to H-bonding interaction between the imine nitrogen and coumarin -OH and creates weak emission (**1'**). Water-assisted tautomerization of **1'** to **1''** leads to complete proton transfer from the coumarin -OH and creates strong emission (**1''**).

Effect of pH

It is noted that pH of the solution is critical for the OCl⁻ sensing. Fig. 6 shows the fluorescence intensity of **1** at 462 nm measured at different pH with and without 50 equiv. of OCl⁻, where the mole fraction distributions of Cl₂, HClO, and OCl⁻ are also shown based on their equilibria in water,^{34,35} using the following equations:



The fluorescence enhancement of **1** by OCl⁻ occurs at neutral physiological pH (6–8), and does not occur at acidic or basic pH. In acidic media (pH < 6), protonation of OCl⁻ (HClO formation; eqn (2)) cancels the basicity of OCl⁻ and, hence, suppresses dehydrogenation of the dihydroperimidine unit of **1**. In contrast, basic media (pH > 8) stabilize OCl⁻, but the fluorescence enhancement does not occur. This is probably because, as observed for several OCl⁻ sensors,^{10,14,15} the oxidation ability of OCl⁻ decreases in basic media and inhibits

dehydrogenation of the dihydroperimidine unit. These data suggest that **1** facilitates fluorometric sensing of OCl⁻ in physiological pH media (pH 6–8).

Conclusions

We synthesized a coumarin–dihydroperimidine dye (**1**), acting as a fluorescent sensor for OCl⁻ in 99% water. **1** shows a weak fluorescence, but OCl⁻-selective dehydrogenation of its dihydroperimidine unit creates a strong blue fluorescence. **1** facilitates selective and sensitive OCl⁻ detection at physiological pH. The turn-on response of **1** occurs *via* two-step reactions. The dehydrogenation by OCl⁻ rapidly produces the enol-imine form (**1'**) involving the H-bonding interaction between the imine nitrogen and coumarin -OH. This increases the electron density of the coumarin unit, resulting in weak fluorescence. **1'** undergoes tautomerization to the keto-amine form (**1''**) due to the stabilization in polar water media. The complete proton transfer from the coumarin -OH to the imine nitrogen significantly increases the electron density of the coumarin unit, exhibiting a strong fluorescence. The molecular design based on the dihydroperimidine unit as an OCl⁻-driven proton sensor, may contribute to the design of efficient fluorescent sensors for OCl⁻ in environmental and biological samples.

Experimental

General

All chemicals were used as received. ·OH was generated by the Fenton reaction.³⁶ ¹O₂ was generated from the H₂O₂/MoO₄²⁻ system in alkaline media.³⁷ NO was generated using sodium nitroferricyanide(III) dehydrate.³⁸ ONOO⁻ was generated from the SIN-1 reagent (Dojindo Molecular Technologies, Japan). ·O₂⁻ was generated using potassium superoxide (KO₂).³⁶ Fluorescence spectra were measured on a JASCO FP-6500 fluorescence spectrophotometer with a 10 nm path length cell (both excitation and emission slit widths, 5.0 nm) at 298 ± 1 K using a temperature controller.³⁹ Absorption spectra were measured on an UV-visible photodiode-array spectrometer (Shimadzu; Multispec-1500) equipped with a temperature controller (S-1700).⁴⁰ All measurements were performed under aerated conditions. ¹H and ¹³C NMR charts were obtained using a JEOL JNM-ECS400 spectrometer. FAB-MS analysis was performed on a JEOL JMS 700 Mass Spectrometer. Fluorescence quantum yields (Φ_{F}) were determined with quinine sulfate dihydrate (in 0.1 M HClO₄ solution) as a standard.^{41,42}

Synthesis of the sensor (**1**) [8-(2,3-dihydro-1H-perimidin-2-yl)-7-hydroxy-4-methyl-2H-chromen-2-one]

8-Formyl-7-hydroxy-4-methylcoumarin (200 mg, 0.98 mmol)²⁵ and 1,8-diaminonaphthalene (188 mg, 1.20 mmol) were dissolved in EtOH (20 ml), and the solution was stirred at 80 °C for 2.5 h. The solid formed was recovered by filtration and washed thoroughly with EtOH, affording **1** as pale pink solids. Yield: 234.4 mg (69.4%). ¹H NMR (400 MHz, DMSO-d₆, TMS), δ (ppm): 10.38 (1H, s), 7.74 (1H, d, $J = 8.8$ Hz), 7.24–7.28 (2H, m), 7.18–

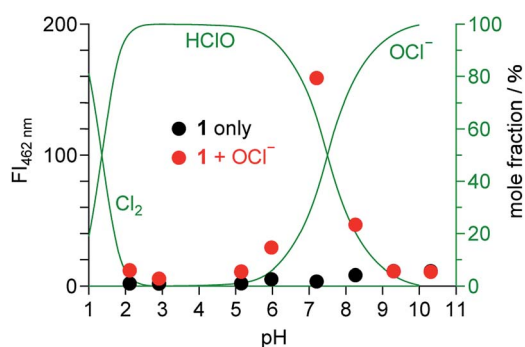


Fig. 6 Fluorescence intensity of **1** (10 μ M) monitored at 462 nm in water/MeCN mixtures (99/1 v/v) at 25 °C with different pH, (red) with and (black) without OCl⁻ (50 equiv.). The mole fraction distributions of Cl₂, HClO, and OCl⁻ calculated based on the equilibria (eqn (1) and (2)) are shown by green lines.



7.20 (2H, m), 7.03 (2H, s), 6.94 (1H, d, $J = 8.8$ Hz), 6.64 (2H, d, $J = 7.2$ Hz), 6.23 (1H, s), 5.99 (1H, s), 2.44 (3H, s). ^{13}C NMR (100 MHz, DMSO- d_6 , TMS), δ (ppm): 161.0, 159.6, 153.8, 152.4, 142.9, 134.2, 126.9, 126.7, 117.2, 113.5, 113.2, 112.0, 110.9, 110.3, 106.3, 59.9, 55.9, 18.3. FAB-MS: m/z : calcd for $\text{C}_{21}\text{H}_{16}\text{O}_3\text{N}_2^+$ (M^+) 344.1161; found (ESI $^+$): 344.1158.

Calculation details

Ab initio calculations were performed with tight convergence criteria at the DFT level within the Gaussian 03 package, using the B3LYP/6-31+G(D) basis set for all atoms. The excitation energies and oscillator strengths of the compounds were calculated by TDTFT⁴³ at the same level of optimization using the PCM with water as a solvent.⁴⁴ Cartesian coordinates are summarized at the end of ESI. †

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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