



Cite this: *RSC Adv.*, 2019, 9, 30269

An insight into the intermolecular vibrational modes of dicationic ionic liquids through far-infrared spectroscopy and DFT calculations†

Luca Guglielmero,^a Lorenzo Guazzelli,^a  ^a Alessandra Toncelli,^b Cinzia Chiappe,[‡] Alessandro Tredicucci  ^{bc} and Christian Silvio Pomelli  ^{*a}

Dicationic ionic liquids (DILs) are a subclass of the ionic liquid (IL) family and are characterized by two cationic head groups linked by means of a spacer. While DILs are increasingly attracting interest due to their peculiar physico-chemical properties, there is still a lack of understanding of their intermolecular interactions. Herein, we report our investigations on the intermolecular vibrational modes of two bromide DILs and of a bistriflimide DIL. The minimal possible neutral cluster of ions was studied as a simplified model of these systems and was optimized at the DFT level. Normal modes of two sandwich-like conformers were then calculated using the harmonic approximation with analytical computation of the second derivatives of molecular energy with respect to the atomic coordinates. The calculated spectra were compared to far-infrared experimental spectra and two groups of peaks over three, for the two bromide DILs, and three over five, for the Tf₂N⁻ DIL, were described by the proposed neutral cluster model. Therefore, this model represents a reliable and computationally affordable model for the exploration of the intermolecular interactions of this kind of system.

Received 24th July 2019
 Accepted 18th September 2019

DOI: 10.1039/c9ra05735h

rsc.li/rsc-advances

Introduction

Ionic liquids (ILs) are organic salts composed of an organic cation and an organic or an inorganic anion which are liquid below 100 °C. The physico-chemical properties of these neoteric solvents (*e.g.*, negligible vapor pressure, wide electrochemical window, high thermal stability, low flammability)^{1–3} and the possibility to tweak them by changing the constituent ions⁴ allowed their use in different fields. ILs have been studied and exploited for instance in the dissolution, fractionation and functionalization of biomasses,^{5–8} in organic synthesis,^{9,10} in analytical chemistry,¹¹ in electrochemistry^{12,13} and, more recently, in biological and pharmaceutical applications.¹⁴

Dicationic ionic liquids (DILs) are a subclass of ILs which are characterized by two cationic head groups linked to one another by means of a spacer. After the pioneering studies by Ohno *et al.*^{15,16} and Armstrong *et al.*,¹⁷ DILs raised the interest of several research groups which investigated their potential field

of applications,^{18–20} also in comparison with the parental mono-cationic ILs (MILs). In fact, the most interesting aspect of DILs, which usually require a more demanding synthetic effort, is the possibility to capitalize on application-oriented advantages over the corresponding MILs. From a structural perspective, the additional variability options of DILs stem from the type and length of the linker between the two cations and/or from the nature of the cation, as asymmetric structures are also possible.

A few works indicated that DILs behave in a different way from the structurally related MILs when tested in the same application. For instance, it has been found that some DILs are characterized by lower polarity and different structural organizations,²¹ higher thermal stabilities,²² lower ecotoxicities,²³ higher activity in the cycloaddition of CO₂ to epoxides,²⁴ higher or lower cathodic reductions, depending on the type of spacer.²⁵

The physico-chemical properties of ILs, and accordingly their potential applications, are strictly governed by the type and strength of the interaction between the constituent ions. Intermolecular forces between cation and anion determine the presence of expanded network structures such as larger clusters or ion-pairs in the liquid phase. In the pursue of a deeper understanding of the interactions and structures of ILs, vibrational spectroscopy represents an invaluable option.²⁶

The frequency range between 2 and 300 cm⁻¹ involves intermolecular forces and can be studied by different experimental methods such as optical heterodyne-detected Raman-induced Kerr-effect spectroscopy (OHD-RIKES),^{27–32} dielectric relaxation spectroscopy (DRS),^{32–34} terahertz (THz)

^aDept. of Pharmacy, University of Pisa, Via Bonanno 33, 56126 Pisa, Italy. E-mail: christian.pomelli@unipi.it

^bDipartimento di Fisica “E. Fermi” and Istituto Nanoscienze CNR, Università di Pisa, Largo Bruno Pontecorvo, 56127 Pisa, Italy

[†]Laboratorio NEST, Scuola Normale Superiore, Piazza San Silvestro 12, 56127 Pisa, Italy

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c9ra05735h

‡ In beloved memory of.



spectroscopy,^{35–39} low-energy neutron scattering,⁴⁰ X-ray diffraction,⁴¹ far infrared (FIR) spectroscopy and Raman spectroscopy.^{36,38,42–47} It is worth mentioning that ILs' experimental spectra are usually difficult to interpret as a result of the complex combination of overlapping bands. Therefore, theoretical methods are often needed to assign specific vibrational motions.

For what concerns THz and FIR spectroscopy, a few studies have been performed mainly on imidazolium ILs [C_1C_2IM or C_4C_1IM and different anions]^{35–37} or ammonium protic ILs.⁴⁸ The first evidence of ${}^+C-H\cdots A^-$ hydrogen bonding was reported by Fumino *et al.*⁴² The same authors showed that the vibrational frequency shifts, observed with different ILs, cannot be explained by mass effects and are rather a result of different anion–cation force constants.^{36,42} Although the interpretation of these experimental spectra is not always univocal,²⁶ the potential of FIR-THz spectroscopy in broadening the understanding of the interactions between the ILs constituting ions has been demonstrated.

To the best of our knowledge, no data are available for the FIR-THz intermolecular vibrational modes of DILs, while some work based on OHD-RIKES has shed some light on the inter-ionic interactions of these systems.³⁷ Given the increasing interest in DILs, here we present our investigations on their intermolecular interactions by means of FIR spectroscopy, corroborated by DTF analyses. Our work focuses on one imidazolium dibistriflimide DIL ($[MIC_6IM]^{2+} + 2Tf_2N^-$) and two imidazolium dibromide DILs ($[BIC_4IB]^{2+} + 2Br^-$ and $[BIC_5IB]^{2+} + 2Br^-$), which differ for the length of the linker (Fig. 1). The former DIL has been studied before by OHD-RIKES, while the latter are model structures of dibromide DILs which are liquid at room temperature.

Results and discussion

The three dicationic ionic liquids (DILs) based on imidazolium charged centers were studied at the *ab initio* level.

The simulation of the complex phenomena that occur in such an intricate system like a dicationic ionic liquid, where charged moieties, some with several degrees of freedom, interact with each other, is a challenging task. In this paper, we

Table 1 Some geometric and energetic quantities related to the minimal neutral clusters studied here. x indicates the molar fraction. Distances are in Å, torsion angle is defined by the two bonds of N1 atoms with the first and last carbon atom of the spacer chain (like two steps of a spiral ladder). A indicates the bromine atom when the anion is Br^- and instead the nitrogen atom when the anion is Tf_2N^- . Y indicates the atom of the anion nearest to the hydrogen

	$[BIC_4IB]^{2+} + 2Br^-$		$[BIC_5IB]^{2+} + 2Br^-$		$[MIC_6IM]^{2+} + 2Tf_2N^-$	
	<i>syn</i>	<i>anti</i>	<i>syn</i>	<i>anti</i>	<i>syn</i>	<i>anti</i>
$\Delta E/(kJ\ mol^{-1})$	1.22	0.00	19.98	0.00	3.08	0.00
x (300 K)	0.38	0.62	0.0003	0.9997	0.33	0.67
μ/D	6.76	1.47	6.11	3.18	5.98	2.79
Torsion angle	−19.95	−9.82	16.94	−13.10	39.16	23.16
N1–N1	4.42	4.46	4.88	4.77	6.03	6.00
C2–C2	4.87	—	5.13	—	6.34	—
C2–C5 ^a	—	4.88	—	5.22	—	6.12
N3–N3	5.27	5.13	5.43	5.50	6.50	5.87
A–A	7.19	7.23	6.80	7.02	8.16	8.52
H2–Y	2.99	2.76	2.90	2.75	2.83	2.36
	2.69	2.59	2.59	2.62	2.23	2.29
H5–Y	2.58	2.76	2.65	2.79	2.12	2.20
	2.62	2.76	2.99	3.04	2.79	2.56

^a Average value.

will limit ourselves to a simple model system which corresponds to the minimal possible neutral cluster of ions: one dication and two single charged anions. Therefore, we will not be able to rationalize and explain all the features of the experimental spectra. However, the features predicted by this model are surely related to molecular motions of this simplified system, while the remaining ones will be likely related to motions that involve larger clusters.

When the molecular nature of these dications is taken into consideration, the most stable structure is a sandwich-like one where the two imidazolium rings assume a parallel arrangement. The distance and the relative orientation of the two parallel rings depend on the length of the linker and on the size of the non-linked sidechains. The monoatomic anions, for example bromides, can establish a larger number of

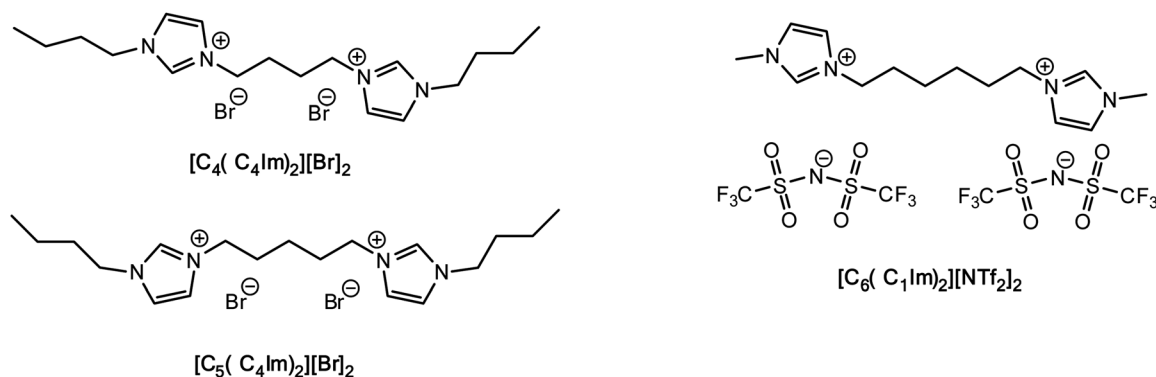


Fig. 1 Structures of the dicationic ionic liquids (DILs) studied.



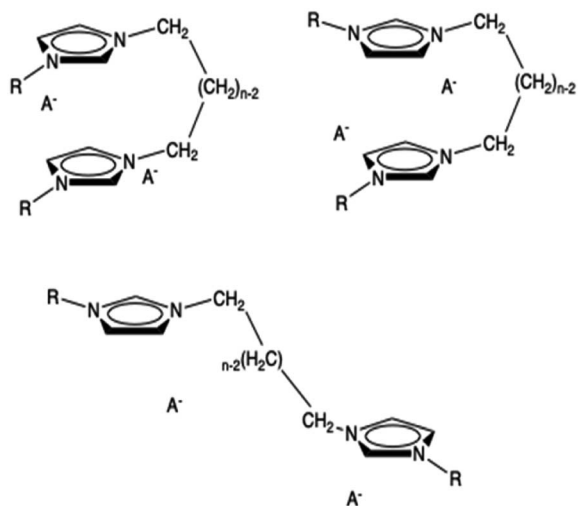


Fig. 2 Some possible conformers of a minimal neutral cluster of dicationic ionic liquids: *syn* (top left), *anti* (top right), open (bottom).

interactions with the acidic hydrogens of the dication moiety compared to the corresponding monocationic ILs (all hydrogens in α to the nitrogen atoms, including those present on the aliphatic chains). In the resulting structure, the anions are sandwiched between the imidazolium fragments and positioned in such a way to maximize their distance. Some starting geometries fulfilling these considerations with different initial positions of the anions have been prepared and optimized at the DFT level. The structures with lower energy present the two anions diametrically opposed. Thus, we focalize on them. Some higher energy structures have been discarded. The resulting local minima are, in this case, not too different, from the qualitative point of view, from the initial geometric guess.

There are two possible conformational isomers as the two imidazolium rings can be arranged in two different spatial ways: with the C2 atoms on the same side of the sandwich (*syn* conformer) or on the opposite side (*anti* conformer). The two conformations present a relevant difference in energy in the $[\text{BIC}_5\text{IB}]^{2+} + 2\text{Br}^-$ case only (See Table 1), but differ in the value of the molecular dipole. The *anti* conformer, the more energetically favored, is characterized by a smaller dipole, while the *syn* conformer is less stable with a larger dipole. The latter conformer is expected to have stronger cluster-cluster interactions. As a result of the dipole quenching in *anti* structures, this conformer presents the two rings almost parallel, while in the *syn* conformer the two aromatic moieties form a dihedral angle (C2 atoms are closer than C4 and C5 atoms).

A third possible conformation is an open one where the two rings interact with rings belonging to different dication. The anions are placed as usually near the two rings. We have tried to optimize this kind of structure but it collapses into the closed one. In the minimal cluster system, this kind of structure is not stable but it can exist in the bulk ionic liquid. Thus, this kind of extended structure, resembling a unit of a coordination polymer, cannot be studied with this approach. It is conceivable that, in organized liquids like those examined here more than

one of the structures are relevant and slowly equilibrating. All these three conformations are shown in Fig. 2.

Normal modes have been calculated using the harmonic approximation with analytical calculation of second derivatives of molecular energy with respect to the atomic coordinates.

The calculations have been performed at the B3LYP/AUG-CC-PVTZ level with empirical corrections for dispersion. The same functional and basis set has been used by Shirota *et al.*^{49–51} with good results to compute Raman spectra of similar systems. A direct comparison between our and the above cited papers is not possible. They used *ab initio* calculations on single anions and cations, not on clusters. Coherently, their dicationic structures present an open geometry due to the absence of the anions. Based on the single ion calculated Raman spectra, they parametrized a theory able to predict non-linear optical properties of the molecules. Furthermore, we do not have calculated Raman intensities and they did not report the infrared ones. Using the *ab initio* results as starting parameters, they obtained a good agreement with the experimental data, which is an implicit validation of the reliability of the calculations results.

As a final validation of the computational method, optimized geometries and normal modes have been calculated for the $[\text{BIC}_4\text{IB}]^{2+} + 2\text{Br}^-$ *anti*-conformer with two different choices of functionals and basis sets. No qualitative differences have been found. Some data about two of the normal modes discussed later in the paper is reported and discussed in ESI.†

Cluster geometries have been fully optimized. They are available as ESI.†

Even though the harmonic approximation is known to be more reliable for localized molecular motions like atom-atom stretching, and despite, the limited size of the model system used here, the absence of any representation of the solvation bulk, the calculated frequencies of the normal modes closely match the experimental results. Full detailed tables of normal modes frequencies and intensities are reported in the ESI.† In the paper body they will be commented using graphical representation and discussed with an emphasis on groups and range of frequencies.

We analyzed all the normal modes below the threshold of 700 cm^{-1} . A complete list for all the considered systems is shown in ESI.† Few of these normal modes presented a non-negligible intensity.

The match between experimental spectra and the calculated normal frequencies (we choose to not report simulated spectra based on arbitrary line shapes) is represented in Fig. 3.

The two systems with bromide anions will be analyzed first. A first noticeable feature of the experimental spectra, which is also present in the calculated frequencies, is the gap of absorption in the central part of the spectral range. The experimental spectra of these systems are composed by three large bands which are centered at about 100 , 450 and 650 cm^{-1} respectively.

The band at about 100 cm^{-1} presents a composite structure made up by at least two superimposed peaks. From the computational point of view there is a group of normal modes characterized by large intensities in this range. For both the conformers this group is divided into two subgroups, one that lies at about 80 cm^{-1} and the other one at about 130 cm^{-1} . This feature reproduces the binary nature of the experimental band.



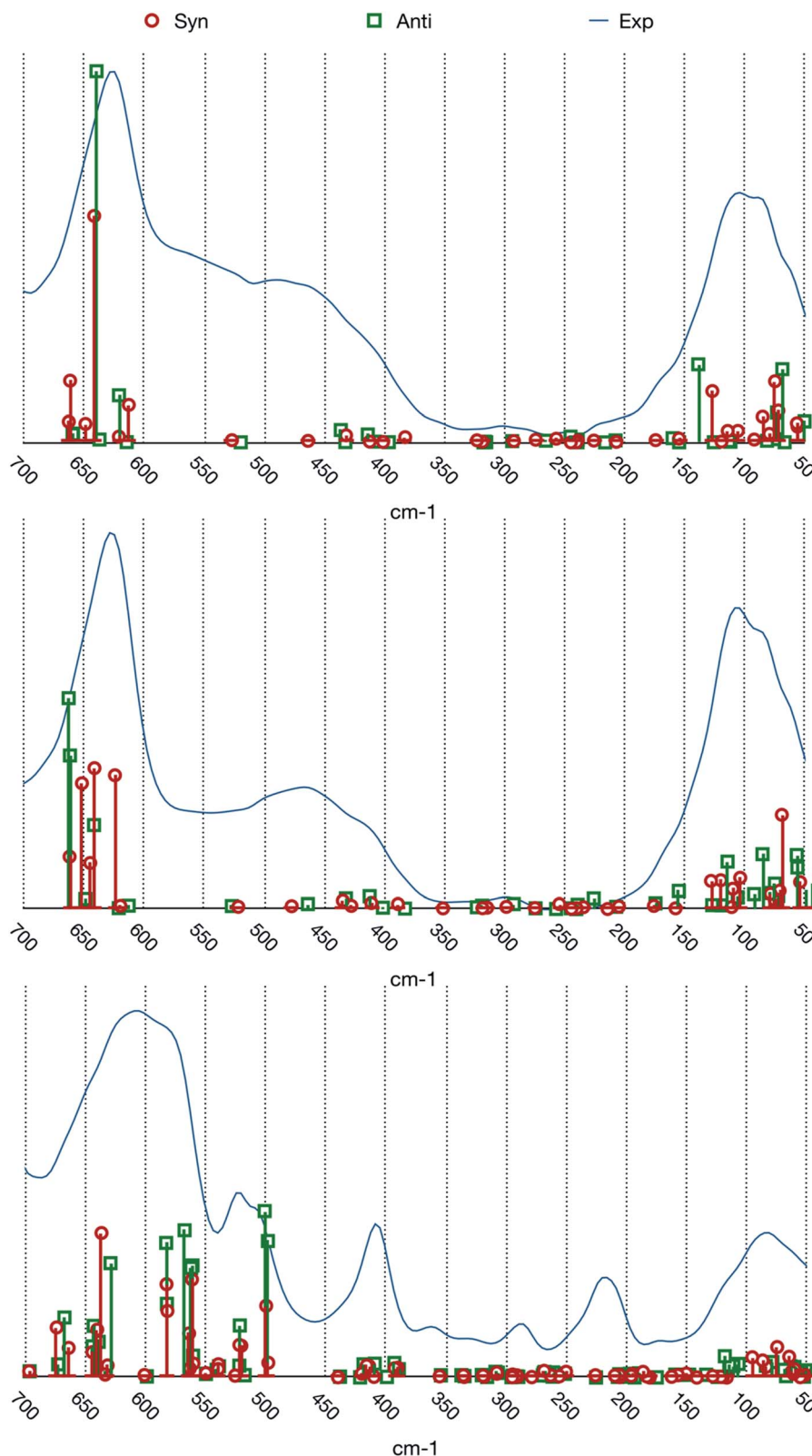


Fig. 3 Comparison between the experimental spectra and calculated normal modes frequencies and intensities. From top to bottom: $[\text{BIC}_4\text{IB}]^{2+} + 2\text{Br}^-$, $[\text{BIC}_5\text{IB}]^{2+} + 2\text{Br}^-$ and $[\text{MIC}_6\text{IM}]^{2+} + 2\text{Tf}_2\text{N}^-$.

This group is related to the relative motions between the two rings and the anions. A graphical representation of some of these normal modes is reported in Fig. 4. In all these normal

modes the two rings and the two anions display relative motions. The deformation of the rings as well as of the chain is minimal and they behave as almost rigid bodies.



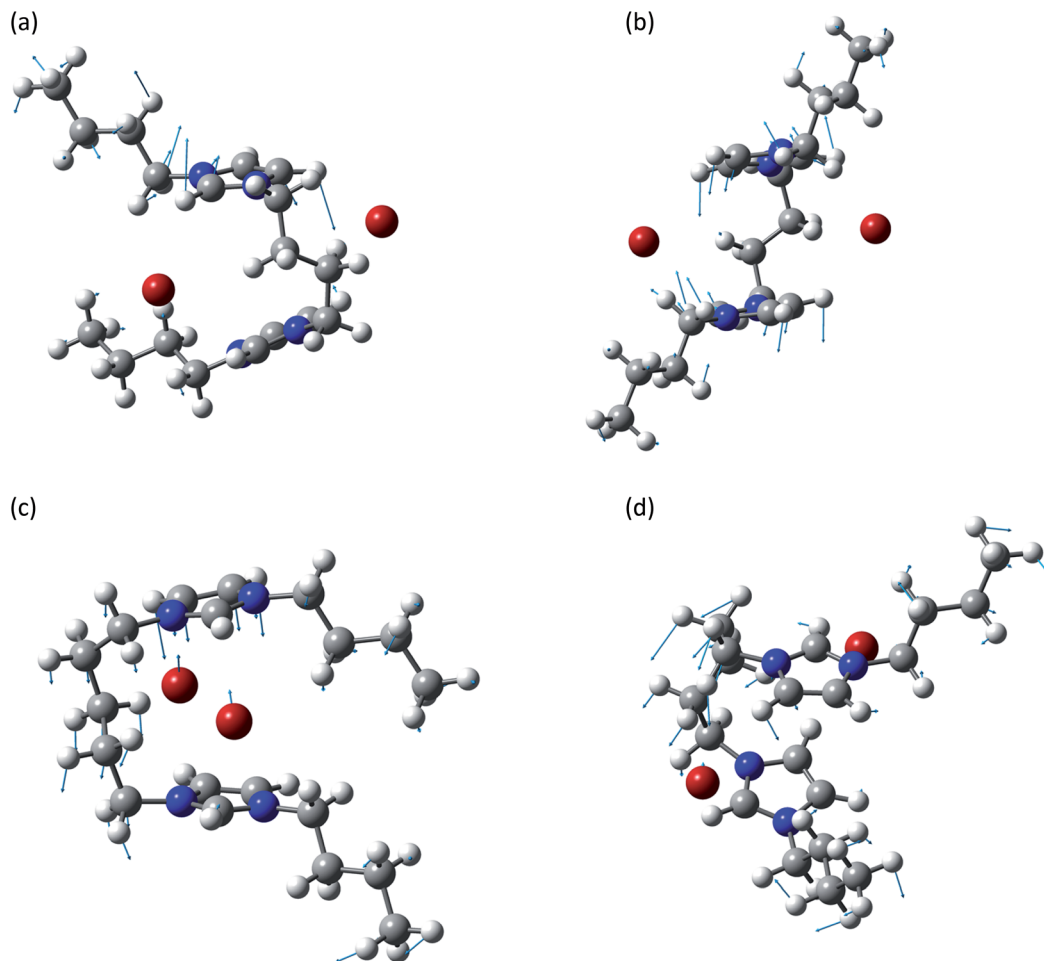


Fig. 4 Normal modes are represented by displacement vectors (light blue). Relative lengths of the vectors are in scale with respect to the displacement of the single atoms. The absolute scale is arbitrary (but is constant in all the normal modes represented in this paper). Full scale pictures are available as ESI.† (a) $[\text{BIC}_4\text{IB}]^{2+} + 2\text{Br}^-$ *syn*: normal mode at 126.9 cm^{-1} . The ring rotates about a horizontal axis. (b) $[\text{BIC}_4\text{IB}]^{2+} + 2\text{Br}^-$ *anti*: normal mode at 138.6 cm^{-1} . Synchronous rotation of one ring with respect to the other. (c) $[\text{BIC}_5\text{IB}]^{2+} + 2\text{Br}^-$ *syn*: normal mode at 69.0 cm^{-1} . Parallel displacement of the top ring with respect to the two bromine anions for $[\text{BIC}_5\text{IB}]^{2+}$. The distance between the two bromine atoms reduces while they approach the ring. (d) $[\text{BIC}_5\text{IB}]^{2+} + 2\text{Br}^-$ *anti*: normal mode at 82.8 cm^{-1} . The two rings rotated about a vertical axis in opposite directions. This corresponds to a rotation of the bromide anions.

The subgroup at lower wavenumbers corresponds to the relative motions of anions and cation. These normal modes are similar to the anion–cation stretching displayed by mono-charged ionic liquids,³⁷ but the presence of the linker introduces additional structured features like the concerted motions of the two bromide anions (Fig. 4c and d).

In the other subgroup, the displacement of the anions, already reduced due to the massive nature of bromine atoms, is basically negligible. These normal modes consist of synchronized relative motions of the two aromatic rings.

It is reasonable to assume that, if we were to search for the minimal energy structure of a model system constituted by two ionic pairs of mono-charged ionic liquids, we would find similar motions. However, the presence of the linker transforms these motions from intermolecular to intramolecular. Furthermore the linker reduces the possible relative distance and orientation of the two imidazolium groups.

The normal modes corresponding to the central band (about 450 cm^{-1}) of the experimental spectra present very small intensities. The molecular motions in this region are related to the relative torsion of the two rings about their vertical axis. This band is possibly also connected to more complex molecular motions typical of larger clusters of ions that cannot be represented by our model system.

The band at about 650 cm^{-1} corresponds to the infrared active normal modes and presents more of the usual molecular motions such as the out of plane deformation of the rings. The coordinated motions of the two rings and/or of their anions are still present. A graphical representation of some of these normal modes is shown in Fig. 5. With respect to the band at about 100 cm^{-1} , the differences between the *syn* and the *anti* structures are greater. This phenomenon arises from the primary role played in these normal modes by the hydrogen atoms bound to the C2, C4 and C5 atoms of the imidazolium



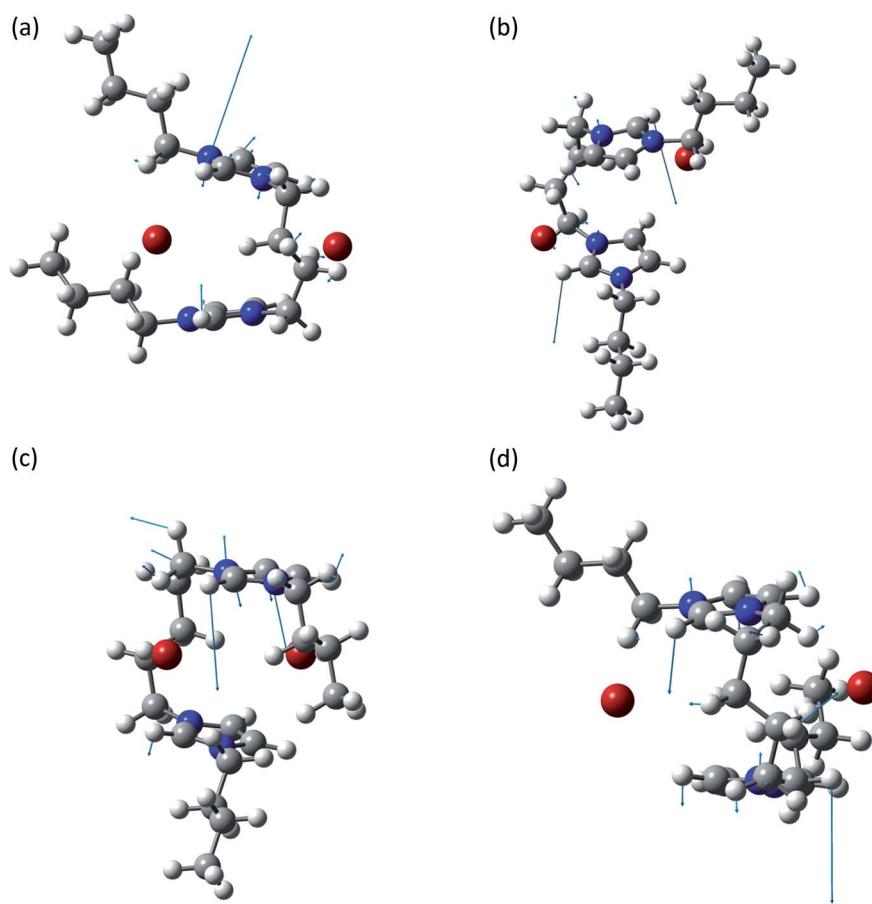


Fig. 5 Normal modes are represented by displacement vectors (light blue). Relative lengths of the vectors are in scale with respect to the displacement of the single atoms. The absolute scale is arbitrary (but is constant in all the normal modes represented in this paper). Full scale pictures are available as ESI.† (a) $[\text{BIC}_4\text{IB}]^{2+} + 2\text{Br}^-$ *syn*: normal mode at 641.8 cm^{-1} . In phase out of plane bending of the C2–H bonds. (b) $[\text{BIC}_4\text{IB}]^{2+} + 2\text{Br}^-$ *anti*: normal mode at 639.8 cm^{-1} . In phase out of plane bending of the C2–H bonds. (c) $[\text{BIC}_5\text{IB}]^{2+} + 2\text{Br}^-$ *syn*: normal mode at 624.4 cm^{-1} . Deformation of the top ring. Synchronously the bottom ring shows a similar deformation with smaller amplitude. (d) $[\text{BIC}_5\text{IB}]^{2+} + 2\text{Br}^-$ *anti*: normal mode at 651.1 cm^{-1} . Concerted deformation of both rings.

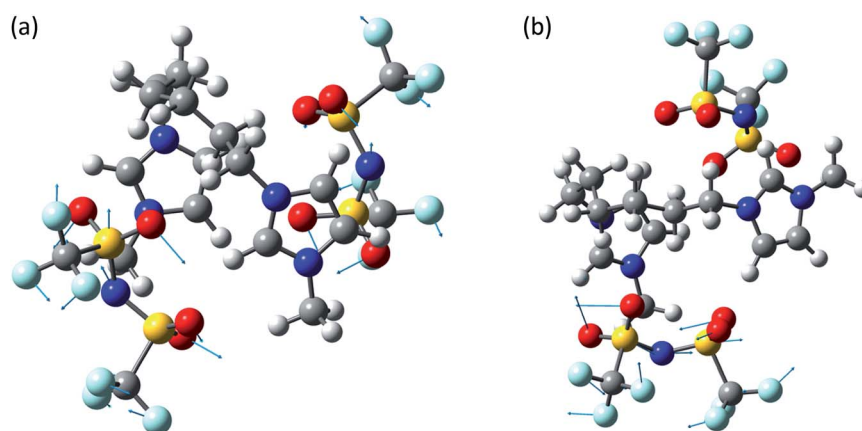


Fig. 6 Normal modes are represented by displacement vectors (light blue). Relative lengths of the vectors are in scale with respect to the displacement of the single atoms. The absolute scale is arbitrary (but is constant in all the normal modes represented in this paper). Full scale pictures are available as ESI.† (a) $[\text{MIC}_6\text{IM}]^{2+} + 2\text{Tf}_2\text{N}^-$ *syn*: normal mode at 498.2 cm^{-1} . Concerted Deformations of the coordinated anions in opposition of phase. There is also a similar, less intense, in phase normal mode at 499.7 cm^{-1} . The calculated intensities differ of about an order of magnitude. (b) $[\text{MIC}_6\text{IM}]^{2+} + 2\text{Tf}_2\text{N}^-$ *anti*: normal mode at 498.2 cm^{-1} . Deformation of a single coordinated anion. In a very similar normal mode at 500.9 cm^{-1} the other anion deforms in a nearly identical way.



rings. Almost all the motions are coordinated amongst the two rings, with variable ratios between the relative amplitudes.

The system $[\text{MIC}_6\text{IM}]^{2+} + 2\text{Tf}_2\text{N}^-$ (Fig. 6) does not present substantial differences in the above analyzed bands. The normal modes are not represented here because of their graphical complexity. However, this system presents supplementary bands at $500\text{--}550\text{ cm}^{-1}$ and $200\text{--}250\text{ cm}^{-1}$. The first band corresponds to normal modes of deformations of single anions (a motion that does not exist for monatomic anions). There is instead no computation match for the second band. The above written comments about the band at about 450 cm^{-1} are still valid.

Conclusions

With this study, we investigated the complex dynamics of dicationic ionic liquids. The minimal neutral cluster proved to be a computationally affordable and a good representation of this kind of systems.

Two groups of peaks over three, for the two systems with bromide anions, and three over five, for the Tf_2N^- anions, can be described by this model system. Therefore, the minimal cluster studied here is a realistic representation of at least a part of the dicationic ionic liquid properties.

The presence of the covalent spacer between the two imidazolium cations allowed us to focus on some motions that in simple monocationic ionic liquids are of an intermolecular nature. The cooperative molecular motions observed here can also take place in a transient mode in mono-charged ionic liquids.

Syn- and *anti*-conformers present very similar energies in two of the systems considered here. Furthermore, the less stable structure is expected to have stronger cluster–cluster interactions due to its larger dipole. Therefore, normal mode frequencies and intensities between the two conformers are not enough different to exclude the existence of one or the other conformer on the base of infrared data.

This approach can also be expanded to larger cluster of ionic liquids too. However, a different computational strategy must though be devised because of the high computational cost associated with the use of the minimal cluster model.

Experimental section

Materials

1*H*-Imidazole was purchased from Fluka, 1-methyl-1*H*-imidazole and 1-butyl-1*H*-imidazole, 1,5-dibromopentane, 1,6-dibromohexane were purchased from Alfa Aesar, whereas 1,4-dibromobutane from Sigma Aldrich. Lithium bis(trifluoromethylsulfonyl) imide was purchased from IOLITEC. All substances were used as received, without further purification, unless differently specified. 4-Methyl-2-pentanone was purchased from Alfa Aesar, all other organic solvents from Sigma Aldrich. All organic solvents were used without further purifications.

Dicationic ionic liquids preparation

3,3'-(Butane-1,4-diyl)bis(1-butyl-1*H*-imidazol-3-ium) bromide ($[\text{BIC}_4\text{IB}]^{2+} + 2\text{Br}^-$), 3,3'-(pentane-1,5-diyl)bis(1-butyl-1*H*-

imidazol-3-ium) bromide ($[\text{BIC}_4\text{IB}]^{2+} + 2\text{Br}^-$) and 3,3'-(hexane-1,6-diyl)bis(1-methyl-1*H*-imidazol-3-ium) bis(trifluoromethane) sulfonamide ($[\text{MIC}_6\text{IM}]^+ + 2\text{Tf}_2\text{N}^-$) were prepared following previously published procedures.^{52,53}

FIR measurements

FIR spectra were recorded with a FT/IR-6800 Fourier Transform Infrared Spectrometer (Jasco).

To collect the spectra an appropriate cell was developed and tested in our lab. It consists in a pair of COC polymer plates kept in position by a steel sample holder. Of the two polymer plates one is smooth and the second is hollowed in the middle, in order to guarantee a sample thickness of 100 μm . The sample holder is made by two steel plates (with a 8 mm diameter hole in the center, for transmitting the light beam) which, after inserting the COC cell between them, were tightened by screws.

DFT calculations

All calculation has been performed using the Gaussian 16 suite.⁵⁴ D3 empirical correction for dispersion has been calculated.⁵⁵ All the geometries have been fully optimized using the default setting at the B3LYP/ug-cc-pvtz level of theory. The frequencies have been calculated analytically. Two linux based workstations equipped with dual Xeon E5-2640 (10 core CPU) and 128 Gb of RAM have been used.

Conflicts of interest

The authors declare that there is no conflict of interest regarding the publication of this article.

Acknowledgements

A. T. acknowledges support from the project PRA_2018_34 (“ANISE”) from the University of Pisa.

References

- 1 Y. Cao and T. Mu, *Ind. Eng. Chem. Res.*, 2014, **53**, 8651–8664.
- 2 D. M. Fox, J. W. Gilman, A. B. Morgan, J. R. Shields, P. H. Maupin, R. E. Lyon, H. C. De Long and P. C. Trulove, *Ind. Eng. Chem. Res.*, 2008, **47**, 6327–6332.
- 3 S. Kazemiabnavi, Z. Zhang, K. Thornton and S. Banerjee, *J. Phys. Chem. B*, 2016, **120**, 5691–5702.
- 4 K. Seddon, *Chem. Eng.*, 2002, **730**, 33–35.
- 5 S. S. Silva, J. F. Mano and R. L. Reis, *Green Chem.*, 2017, **19**, 1208–1220.
- 6 I. Palazzo, A. Mezzetta, L. Guazzelli, S. Sartini, C. S. Pomelli, W. O. Parker Jr and C. Chiappe, *RSC Adv.*, 2018, **8**, 21174–21183.
- 7 C. Chiappe, M. J. Rodriguez Douton, A. Mezzetta, L. Guazzelli, C. S. Pomelli, G. Assanelli and A. R. De Angelis, *New J. Chem.*, 2018, **42**, 1845–1852.
- 8 A. Brandt-Talbot, F. J. V. Gschwend, P. S. Fennell, T. M. Lammens, B. Tan, J. Weale and J. P. Hallett, *Green Chem.*, 2017, **19**, 3078–3102.



- 9 J. P. Halett and T. Welton, *Chem. Rev.*, 2011, **111**, 3508–3576.
- 10 C. Chiappe and C. S. Pomelli, *Eur. J. Org. Chem.*, 2014, 6120–6139.
- 11 F. Ghorbanizamani and S. Timur, *Anal. Chem.*, 2018, **90**, 640–648.
- 12 M. M. Huie, R. A. Dileo, A. C. Marschilok, K. J. Takeuchi and E. S. Takeuchi, *ACS Appl. Mater. Interfaces*, 2015, **7**, 11724–11731.
- 13 M. Longhi, S. Arnaboldi, E. Husanu, S. Grecchi, I. F. Buzzi, R. Cirilli, S. Rizzo, C. Chiappe, P. R. Mussini and L. Guazzelli, *Electrochim. Acta*, 2019, **298**, 194–209.
- 14 K. S. Egorova, E. G. Gordeev and V. P. Ananikov, *Chem. Rev.*, 2017, **117**, 7132–7189.
- 15 K. Ito, N. Nishina and H. Ohno, *Electrochim. Acta*, 2000, **45**, 1295–1298.
- 16 M. Yoshizawa, K. Ito-Akita and H. Ohno, *Electrochim. Acta*, 2000, **45**, 1617–1621.
- 17 J. L. Anderson, R. Ding, A. Ellern and D. W. Armstrong, *J. Am. Chem. Soc.*, 2005, **127**, 593–604.
- 18 F. D'Anna, C. Rizzo, P. Vitale, G. Lazzara and R. Noto R, *Soft Matter*, 2014, **10**, 9281–9292.
- 19 J. Chen, L. Yang, W. Zhou, L. Zhu, Y. Zhou, Y. Xiang and D. Xia, *Energy Fuels*, 2018, **32**, 5518–5526.
- 20 Y. Ji, Y. Hou, S. Ren, C. Yao and W. Wu, *Fuel Process. Technol.*, 2018, **171**, 183–191.
- 21 D. Majhi, S. Seth and M. Sarkar, *Phys. Chem. Chem. Phys.*, 2018, **20**, 7844–7856.
- 22 C. Maton, N. De Vos and C. V. Stevens, *Chem. Soc. Rev.*, 2013, **42**, 5963–5977.
- 23 M. G. Montalbán, G. Villora and P. Licence, *Ecotoxicol. Environ. Saf.*, 2018, **150**, 129–135.
- 24 M. H. Anthofer, M. E. Wilhelm, M. Cokoja, M. Drees, W. A. Herrmann and F. E. Kuhn, *ChemCatChem*, 2015, **7**, 94–98.
- 25 I. Rocco, F. D'Anna, L. Mattiello, F. Pandolfi, C. Rizzo and M. Feroci, *ChemElectroChem*, 2019, **6**, 1–10.
- 26 V. H. Paschoal, L. F. O. Faria and M. C. C. Ribeiro, *Chem. Rev.*, 2017, **117**, 7053–7112.
- 27 H. Shirota and E. W. Castner Jr, *J. Phys. Chem. A*, 2005, **109**, 9388–9392.
- 28 D. Xiao, J. R. Rajain, A. Cady, S. Li, R. A. Bartsch and E. L. Quitevis, *J. Phys. Chem. B*, 2007, **111**, 4669–4677.
- 29 D. Xiao, L. G. Hines Jr, S. Li, R. A. Bartsch, E. L. Quitevis, O. Russina and A. Triolo, *J. Phys. Chem. B*, 2009, **113**, 6426–6433.
- 30 L. Xue, F. Bardak, G. Tamas and E. L. Quitevis, *Phys. Chem. Chem. Phys.*, 2017, **19**, 4661–4672.
- 31 S. Kakinuma, T. Ishida and H. Shirota, *J. Phys. Chem. B*, 2017, **121**, 250–264.
- 32 D. A. Turton, J. Hunger, A. Stoppy, G. Hefter, A. Thoman, M. Walther, R. Buchner and K. Wynne, *J. Am. Chem. Soc.*, 2009, **131**, 11140–11146.
- 33 C. Gaguinet, P. J. Dyson, I. Krossing, A. Oleinikowa, J. Slattery, C. Wakai and H. Weingärtner, *J. Phys. Chem. B*, 2006, **110**, 12682–12688.
- 34 K. Nakamura and T. Shikata, *ChemPhysChem*, 2010, **11**, 285–294.
- 35 K. Yamamoto, M. Tani and M. Hangyo, *J. Phys. Chem. B*, 2007, **111**, 4854–4859.
- 36 A. Wulf, K. Fumino, R. Ludwig and P. F. Taday, *ChemPhysChem*, 2010, **11**, 349–353.
- 37 T. Yamada, Y. Tominari, S. Tanaka and M. Mizuno, *J. Phys. Chem. B*, 2015, **119**, 15696–15705.
- 38 T. Yamada, Y. Tominari, S. Tanaka, M. Mizuno and K. Fukunaga, *Materials*, 2014, **7**, 7409–7422.
- 39 S. Mou, A. Rubano and D. Paparo, *J. Phys. Chem. B*, 2018, **122**, 3133–3140.
- 40 B. S. Hudson, *J. Phys. Chem. A*, 2001, **105**, 3949–3960.
- 41 C. Roth, T. Peppel, K. Fumino, M. Köckerling and R. Ludwig, *Angew. Chem., Int. Ed.*, 2010, **49**, 10221–10224.
- 42 K. Fumino, A. Wulf and R. Ludwig, *Angew. Chem.*, 2008, **120**, 3890–3894; *Angew. Chem., Int. Ed.*, 2008, **47**, 3830–3834.
- 43 K. Iwata, H. Okajima, S. Saha and H. Hamaguchi, *Acc. Chem. Res.*, 2007, **40**, 1174–1181.
- 44 A. Wulf, K. Fumino and R. Ludwig, *Angew. Chem., Int. Ed.*, 2010, **49**, 449–453.
- 45 Y. Jeon, J. Sung, C. Seo, H. Lim, H. Cheong, M. Kang, B. Moon, Y. Ouchi and D. J. Kim, *J. Phys. Chem. B*, 2008, **112**, 4735–4740.
- 46 T. Yamada, Y. Tominari, S. Tanaka and M. Mizuno, *J. Phys. Chem. B*, 2017, **121**, 3121–3129.
- 47 K. Fumino, K. Wittler and R. Ludwig, *J. Phys. Chem. B*, 2012, **116**, 9507–9511.
- 48 M. Krüger, E. Bründermann, S. Funkner, H. Weingärtner and M. Havenith, *J. Chem. Phys.*, 2010, **132**, 101101.
- 49 H. Shirota and T. Ishida, *J. Phys. Chem. B*, 2013, **117**, 1136–1150.
- 50 H. Shirota and T. Ishida, *J. Phys. Chem. B*, 2011, **115**, 10860–10870.
- 51 H. Shirota and S. Kakinuma, *J. Phys. Chem. B*, 2015, **119**, 9835–9846.
- 52 L. Guglielmero, A. Mezzetta, L. Guazzelli, C. S. Pomelli, F. D'Andrea and C. Chiappe, *Front. Chem.*, 2018, **6**, 612.
- 53 H. Shirota, T. Mandai, H. Fukazawa and T. Kato, *J. Chem. Eng. Data*, 2011, **56**, 2453–2459.
- 54 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, G. A. Petersson, H. Nakatsuji, X. Li, M. Caricato, A. V. Marenich, J. Bloino, B. G. Janesko, R. Gomperts, B. Mennucci, H. P. Hratchian, J. V. Ortiz, A. F. Izmaylov, J. L. Sonnenberg, D. Williams-Young, F. Ding, F. Lipparini, F. Egidi, J. Goings, B. Peng, A. Petrone, T. Henderson, D. Ranasinghe, V. G. Zakrzewski, J. Gao, N. Rega, G. Zheng, W. Liang, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, K. Throssell, J. A. Montgomery Jr, J. E. Peralta, F. Ogliaro, M. J. Bearpark, J. J. Heyd, E. N. Brothers, K. N. Kudin, V. N. Staroverov, T. A. Keith, R. Kobayashi, J. Normand, K. Raghavachari, A. P. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, J. M. Millam, M. Klene, C. Adamo, R. Cammi, J. W. Ochterski, R. L. Martin, K. Morokuma, O. Farkas, J. B. Foresman, and D. J. Fox, *Gaussian 16, Revision B.01*, Gaussian, Inc., Wallingford CT, 2016.
- 55 S. Grimme, J. Antony, S. Ehrlich and H. Krieg, *J. Chem. Phys.*, 2010, **132**, 15104.

