



Cite this: *RSC Adv.*, 2019, 9, 29429

Received 14th August 2019  
 Accepted 23rd August 2019

DOI: 10.1039/c9ra06323d

[rsc.li/rsc-advances](http://rsc.li/rsc-advances)

## A formation model of superoxide radicals photogenerated in nano-TiO<sub>2</sub> suspensions†

Dabin Wang,<sup>a</sup> Lixia Zhao,<sup>b</sup> Dean Song,<sup>a</sup> Jun Qiu,<sup>a</sup> Fanyu Kong<sup>\*a</sup> and Liang-Hong Guo<sup>b</sup>

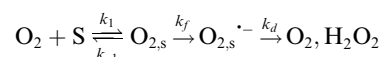
A formation model of O<sub>2</sub><sup>•−</sup> produced in TiO<sub>2</sub> photocatalysis was established, and then a custom built continuous flow chemiluminescence (CFCL) system was used to confirm the model's reliability by monitoring the O<sub>2</sub><sup>•−</sup> formation process. This model may give deeper insights into O<sub>2</sub><sup>•−</sup> formation for TiO<sub>2</sub> and other photocatalysts.

In photocatalytic reactions, such as TiO<sub>2</sub>, it is generally accepted that O<sub>2</sub><sup>•−</sup> is produced from the reduction of adsorbed oxygen by photogenerated electrons on the TiO<sub>2</sub> surface.<sup>1–5</sup> O<sub>2</sub><sup>•−</sup> formation is a rate-limiting process in TiO<sub>2</sub> photocatalytic reactions,<sup>6</sup> and thus determines the efficiency of TiO<sub>2</sub> photocatalytic reactions by promoting the separation of photogenerated electrons and holes. Moreover, O<sub>2</sub><sup>•−</sup> is an intriguing active species, attracting a great deal of attention in recent years due to its unique role. Previous studies have confirmed that O<sub>2</sub><sup>•−</sup> plays an essential role in the photodegradation of pollutants.<sup>7,8</sup> Therefore, probing the dynamic formation of O<sub>2</sub><sup>•−</sup> over the course of time under UV irradiation is conducive to better understanding TiO<sub>2</sub> photocatalytic reactions. We have successfully identified surface long-lived O<sub>2</sub><sup>•−</sup> photogenerated on TiO<sub>2</sub> surface,<sup>9</sup> however, the dynamic details of O<sub>2</sub><sup>•−</sup> formation in TiO<sub>2</sub> photocatalytic reactions still remain to be solved.

In photocatalytic reactions, O<sub>2</sub><sup>•−</sup> is continuously undergoing the processes of formation and deactivation simultaneously. Many methods have been developed to determine the O<sub>2</sub><sup>•−</sup>, including electron spin resonance (ESR),<sup>10</sup> spectrophotometric assays,<sup>11–13</sup> and fluorescence assays,<sup>14,15</sup> but the O<sub>2</sub><sup>•−</sup> could only be detected at discrete times, and thus the total quantity could only be given within the irradiation duration. Given these limitations, O<sub>2</sub><sup>•−</sup> dynamic monitoring is a great challenge. Chemiluminescence (CL) is inherently sensitive and rapid due to the relative ease with which light emission is instantly generated through a chemical reaction when two or more

reactants are mixed. These qualities of CL make it suitable for the dynamic study of O<sub>2</sub><sup>•−</sup>, despite its characteristic instability. In a previous study, we were able to successfully develop a continuous flow chemiluminescence (CFCL) method for dynamic monitoring of the formation process of O<sub>2</sub><sup>•−</sup> in TiO<sub>2</sub> photocatalytic reactions.<sup>16</sup>

For the formation mechanism of O<sub>2</sub><sup>•−</sup> in TiO<sub>2</sub> photocatalytic reactions, it is generally accepted that O<sub>2</sub><sup>•−</sup> formation occurs on TiO<sub>2</sub> surfaces by the following scheme:



where O<sub>2</sub> denotes the dissolved oxygen (DO) in solution, S denotes the oxygen adsorption site on TiO<sub>2</sub> surface, k<sub>1</sub> and k<sub>−1</sub> denote the adsorption/desorption rate constant of O<sub>2</sub> respectively, O<sub>2,s</sub> denotes the adsorbed oxygen on TiO<sub>2</sub> surface, k<sub>f</sub> denotes the formation rate constant of O<sub>2</sub><sup>•−</sup>, O<sub>2,s</sub><sup>•−</sup> denotes the O<sub>2</sub><sup>•−</sup> formed on TiO<sub>2</sub> surface, and k<sub>d</sub> denotes the rate constant of O<sub>2</sub><sup>•−</sup> decomposition. It was reported that DO was first adsorbed on the site of TiO<sub>2</sub> surface following the Langmuir isotherm,<sup>6</sup> and then the adsorbed O<sub>2</sub> was reduced to O<sub>2</sub><sup>•−</sup> by photogenerated electrons under UV irradiation. Meanwhile, the formed O<sub>2</sub><sup>•−</sup> was transformed into other species by side reactions, such as H<sub>2</sub>O<sub>2</sub> or <sup>1</sup>O<sub>2</sub> which deactivated to O<sub>2</sub> quickly. Specifically, it is suggested by some research that O<sub>2</sub><sup>•−</sup> could be produced from the oxidation of H<sub>2</sub>O<sub>2</sub> by valence-band hole (h<sup>+</sup>) or hydroxyl radical (·OH) in solution, where H<sub>2</sub>O<sub>2</sub> is produced by the two-step oxidation of water or the two-electron reduction of O<sub>2</sub>.<sup>17</sup> It is insignificant in this study because the O<sub>2</sub><sup>•−</sup> detected by CFCL method is the long-lived superoxide adsorbed on TiO<sub>2</sub> surface which is produced from the reduction of O<sub>2</sub> by photogenerated electrons, not in solution according to our previous study.<sup>9</sup> The process of O<sub>2</sub><sup>•−</sup> formation was illustrated concretely as follows: when TiO<sub>2</sub> semiconductor is irradiated, the photogenerated electrons transferring to the TiO<sub>2</sub> surface are captured by five-coordinated surface Ti<sup>4+</sup> to form the Ti<sup>3+</sup> (eqn

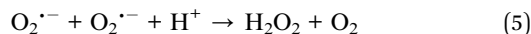
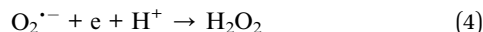
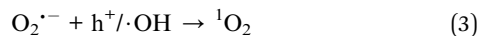
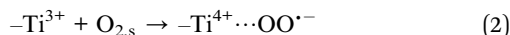
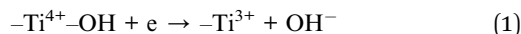
<sup>a</sup>Laboratory of Quality & Safety Risk Assessment for Tobacco, Ministry of Agriculture, Tobacco Research Institute of Chinese Academy of Agricultural Sciences, Qingdao 266101, China. E-mail: kongfanyu@caas.cn; Fax: +86-532-88703386; Tel: +86-532-88703386

<sup>b</sup>State Key Laboratory of Environmental Chemistry and Ecotoxicology, Research Center for Eco-Environmental Sciences, Chinese Academy of Sciences, P. O. Box 2871, 18 Shuangqing Road, Beijing 100085, China. E-mail: zlx@rcees.ac.cn; Fax: +86-10-62849685; Tel: +86-10-62849338

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c9ra06323d



(1)). Then the  $O_2$  adsorbed on  $TiO_2$  surface react with  $Ti^{3+}$  to form  $O_2^{\cdot-}$  (eqn (2)).<sup>1,2,18</sup>



Meanwhile,  $O_2^{\cdot-}$  formed also undergoes a decaying process possibly *via* the following three pathways (eqn (3)–(5)): (1) oxidation by  $h^+$  or  $\cdot OH$ , (2) further reduction by  $e$ , or (3) self-disproportionation in solution. Therefore, the number of  $O_2^{\cdot-}$  determined in photocatalytic reactions at any irradiation time ( $t$ ) is what the total quantity of formed  $O_2^{\cdot-}$  subtract those decomposed *via* the aforementioned side reactions from  $t_0$  s to  $t$  s upon irradiation. The formula can be expressed as follows:

$$[O_2^{\cdot-}]_t = \int_{t_0}^t d[O_2^{\cdot-}]_f/dt - \int_{t_0}^t d[O_2^{\cdot-}]_d/dt \quad (6)$$

In this equation, the  $t_0$  and  $t$  of the lower and upper limit of the definite integral represent the starting and ending time of UV irradiation respectively. Based on the formation process of  $O_2^{\cdot-}$  mentioned above, the net rate of  $O_2^{\cdot-}$  formation at any time in photocatalytic reactions can be obtained as shown in eqn (7.1). The first term of the right side of the eqn (7.1) represents the formation rate of  $O_2^{\cdot-}$ , and it is a second-order reaction with respect to the concentration of  $-Ti^{3+}$  ( $[Ti^{3+}]$ ) and adsorbed oxygen ( $O_{2,s}$ ).  $k_f$  is a second-order formation rate constant. The second term (eqn (7.1)) represents the decay rate of  $O_2^{\cdot-}$ , among which  $k_d$  is a second-order decay rate constant,  $[X]$  represents the concentration of  $e$ ,  $h^+$ ,  $\cdot OH$  or  $O_2^{\cdot-}$  in terms of eqn (3)–(5). It is generally believed to contain the aforementioned three pathways. We have previously confirmed that the  $O_2^{\cdot-}$  adsorbed on  $TiO_2$  surface is thermodynamically favored.<sup>9</sup> Therefore, the two former pathways (eqn (3) and (4)) dominated the decay process under UV irradiation. These processes were considered to be pseudo-first-order due to the constant of  $h^+$ ,  $\cdot OH$  and  $e$  under steady-state irradiation. So  $k'_d$  ( $k'_d = k_d \times [X]$ ) is the pseudo-first-order apparent rate constant (eqn (7.2)).  $[O_2]_s$  is equal to oxygen coverage on  $TiO_2$  surface ( $\theta$ ) by the number of adsorption sites ( $[S]$ ) according to Langmuir isotherm (eqn (7.2)), in which  $\theta$  is related to the DO concentration in solution ( $[O_2]$ ), adsorption constant ( $k_1$ ), desorption constant ( $k_{-1}$ ), and formation rate constant of  $O_2^{\cdot-}$  ( $k_f$ ) (eqn (7.3)).

$$\frac{d[O_2^{\cdot-}]}{dt} = k_f [Ti^{3+}] [O_2]_s - k_d [X] [O_2^{\cdot-}] \quad (7.1)$$

$$= k_f [Ti^{3+}] \theta [S] - k'_d [O_2^{\cdot-}] \quad (7.2)$$

$$= k_f [Ti^{3+}] [S] \frac{k_1 [O_2]}{k_1 [O_2] + k_{-1} + k_f} - k'_d [O_2^{\cdot-}] \quad (7.3)$$

It has been reported that the electron transfer from  $TiO_2$  to  $O_2$  is the rate-limiting step in  $TiO_2$  photocatalytic reaction.<sup>6,19–21</sup> Upon this,  $k_f$  is far less than  $k_1[O_2]$  and  $k_{-1}$ . Furthermore, the  $[O_2]$  in solution is low, and thus  $k_1[O_2]$  is far less than  $k_{-1}$ . If such speculation is true, the eqn (7.3) can be eventually transformed into the following:

$$\frac{d[O_2^{\cdot-}]}{dt} = k_f [Ti^{3+}] [S] K_1 [O_2] - k'_d [O_2^{\cdot-}] \quad (8)$$

where  $K_1$  ( $K_1 = k_1/k_{-1}$ ) is the adsorption equilibrium constant of  $[O_2]$ . Assuming that  $[O_2^{\cdot-}] = mCL + n$ , which is a linear relationship between  $O_2^{\cdot-}$  concentration and CL intensity, the eqn (8) can then be converted into the following:

$$\frac{d[mCL + n]}{dt} = k_f [Ti^{3+}] [S] K_1 [O_2] - k'_d [mCL + n] \quad (9)$$

Finally, the eqn (10) representing the formation model of  $O_2^{\cdot-}$ , with respect to time as the independent variable and CL intensity as the dependent variable, would be obtained by integrating these variables with the eqn (9) from  $t_0$  to  $t$ :

$$CL = \frac{k_f K_1 [Ti^{3+}] [S] [O_2]}{mk'_d} \left( 1 - e^{-k'_d(t-t_0)} \right) + \frac{n}{m} \quad (10)$$

In eqn (10),  $k_f$ ,  $K_1$ ,  $k'_d$ ,  $[S]$ ,  $[O_2]$ ,  $m$ , and  $n$  are constants under certain conditions. The coefficient  $((k_f K_1 [Ti^{3+}] [S] [O_2]) / (mk'_d)) + n/m$  represents the theoretical maximum CL intensity ( $CL_0$ ) corresponding to the steady-state concentration of  $O_2^{\cdot-}$  when  $t$  is infinite in irradiated  $TiO_2$  suspensions. Herein  $t_0$  is the time when irradiation starts for 50 s, due to the disturbance of background signal within the first 50 s of irradiation. According to eqn (10), given that  $k_f$ ,  $K_1$ , and  $k_d$  are determined by the intrinsic property of  $TiO_2$  photocatalyst regardless of experimental conditions,  $CL_0$  is dependent on  $[Ti^{3+}]$ ,  $[S]$ ,  $[O_2]$  and  $[X]$ , which are closely related to experimental conditions. In photocatalytic reactions,  $[Ti^{3+}]$  relies on the number of photogenerated electrons highly dependent on  $I$ ;  $[S]$  is the total surface area of  $TiO_2$  in suspension, closely related to  $[TiO_2]$ ;  $[O_2]$  is dependent on  $[DO]$  in  $TiO_2$  suspensions;  $[X]$ , the aforementioned concentration of  $h^+$ ,  $\cdot OH$  and  $e$ , is also dependent on  $I$ .

In order to verify the formation model of  $O_2^{\cdot-}$ , we fit the different CL curves with eqn (10) by changing  $[TiO_2]$ ,  $I$ , and  $[DO]$ . As shown in Fig. 1(a–i), the CL curves from different  $[TiO_2]$ ,  $I$  and  $[DO]$  were well-fit by eqn (10) with high correlation coefficients ( $R^2 > 0.99$ ), indicating that the model could simulate dynamic process of  $O_2^{\cdot-}$  formation. Furthermore, according to the  $O_2^{\cdot-}$  formation model (eqn (10)), CL intensity is linearly correlated to  $[S]$ ,  $[Ti^{3+}]$ , and  $[O_2]$ , indicating that CL intensity could increase linearly with the increase of  $[TiO_2]$ ,  $I$  and  $[DO]$  within set limits. In order to verify this assumption, CL intensity at 300 s, 600 s, 900 s, 1200 s and  $+\infty$  under different  $[TiO_2]$ ,  $I$  and  $[DO]$  conditions in  $TiO_2$  suspensions was calculated by eqn (10) (Table S1†). Then the CL intensity from different  $[TiO_2]$ ,  $I$  and  $[DO]$  at different time points were linearly fit (Fig. S1–



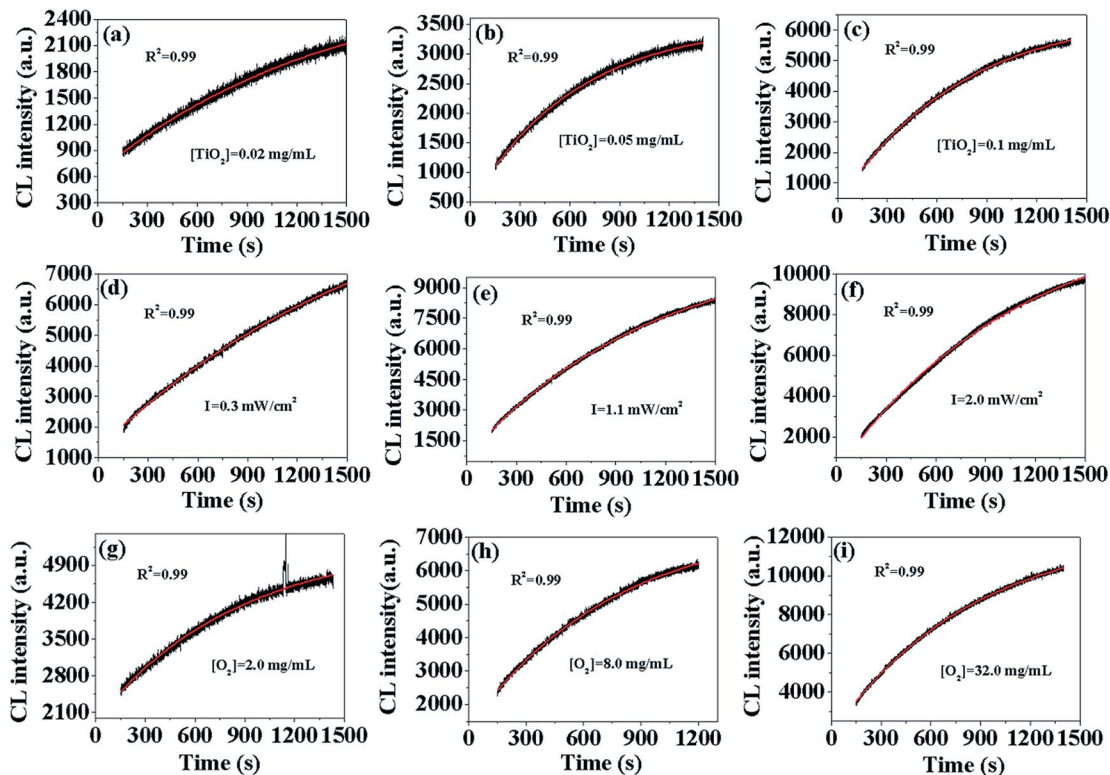


Fig. 1 CL curves of luminol (50  $\mu\text{M}$ ) with photo-irradiated  $\text{TiO}_2$  suspensions ( $\text{pH} = 11$ ) under different experimental conditions:  $\text{TiO}_2$  concentration (a–c), irradiation intensity (d–f), DO concentration (g–i). The black lines represent experimentally measured values and the red lines are the fitted values.

3 $\dagger$ ), and the  $R$  square values were calculated and summarized in Table 1. For instance, the CL intensity with  $\text{TiO}_2$  concentration of 0.02  $\text{mg mL}^{-1}$ , 0.05  $\text{mg mL}^{-1}$  and 0.1  $\text{mg mL}^{-1}$  at 300 s, 600 s, 900 s, 1200 s and  $+\infty$ , respectively, had good linear fit with high correlation coefficients ( $R^2 > 0.99$ ), except at  $+\infty$  ( $R^2 = 0.903$ ), confirming that CL intensity increased linearly with the increase of  $[\text{S}]$  as indicated by eqn (10). There was also a good linear relationship between CL intensity and  $I$  with high correlation coefficients ( $R^2 > 0.9$ ), except at  $+\infty$  ( $R^2 = 0.821$ ), indicating that CL intensity increased linearly with the increase of  $[\text{Ti}^{3+}]$ . At different  $[\text{DO}]$ , the corresponding  $R^2$  values were also high ( $R^2 > 0.95$ ), indicating that CL intensity was also linearly dependent on  $[\text{O}_2]$ . Overall, the formation model of  $\text{O}_2^{\cdot-}$  could adequately describe the dynamic process of  $\text{O}_2^{\cdot-}$  formation in photo-irradiated  $\text{TiO}_2$  suspensions.

In the present work, a formation model of  $\text{O}_2^{\cdot-}$  in  $\text{TiO}_2$  photocatalytic reactions was established. According to the model, the  $\text{O}_2^{\cdot-}$  formation was closely related with  $[\text{TiO}_2]$ ,  $[\text{DO}]$  and  $I$ , under which the dynamic process of  $\text{O}_2^{\cdot-}$  formation was successfully simulated by the model with high correlation coefficients ( $R^2 > 0.9$ ), thereby confirming the model validity. This model can explicitly provide details on  $\text{O}_2^{\cdot-}$  formation which determines the photocatalytic efficiency in  $\text{TiO}_2$  photocatalytic reactions, and give deeper insights into designing high-efficiency  $\text{TiO}_2$  photocatalysts. In accordance with this model, Feng *et al.* reported the self-doped  $\text{Ti}^{3+}$  enhanced  $\text{TiO}_2$  photocatalyst for hydrogen production through the reduction of the  $\text{TiO}_2$  surface using a one-step combustion method.<sup>22</sup> Furthermore, this model may have significant implications for other photocatalysts with respect to  $\text{O}_2^{\cdot-}$  formation.

Table 1 The calculated value of  $R$  square upon different experimental conditions at different irradiation time

Time (s)	$\text{TiO}_2$ concentration ( $\text{mg mL}^{-1}$ )	Irradiation intensity ( $\text{mW cm}^{-2}$ )	DO concentration ( $\text{mg mL}^{-1}$ )
300	0.998	0.921	0.999
600	0.999	0.946	0.982
900	0.998	0.960	0.971
1200	0.994	0.972	0.968
$+\infty$	0.903	0.821	0.977



## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

The authors gratefully acknowledge financial support from the Agricultural Science and Technology Innovation Program (ASTIP-TRIC06), National Key Research and Development Program of China (2016YFA0203102), National Natural Science Foundation of China (31601483, 31801646, 21876184, 21677152), Fundamental Research Funds for Central Non-profit Scientific Institution (1610232017011, 1610232016008, 1610232017010, 1610232018002).

## Notes and references

- 1 Y. Nakaoka and Y. Nosaka, *J. Photochem. Photobiol., A*, 1997, **110**, 299–305.
- 2 A. L. Attwood, D. M. Murphy, J. L. Edwards, T. A. Egerton and R. W. Harrion, *Res. Chem. Intermed.*, 2003, **29**, 449–465.
- 3 M. A. Henderson, W. S. Epling, C. H. F. Peden and C. L. Perkins, *J. Phys. Chem. B*, 2003, **107**, 534–545.
- 4 E. Carter, A. F. Carley and D. M. Murphy, *J. Phys. Chem. C*, 2007, **111**, 10630–10638.
- 5 J. Green, E. Carter and D. M. Murphy, *Chem. Phys. Lett.*, 2009, **477**, 340–344.
- 6 M. R. Hoffmann, S. T. Martin, W. Choi and D. W. Bahnemann, *Chem. Rev.*, 1995, **95**, 69–96.
- 7 C. Chen, P. Lei, H. Ji, W. Ma and J. Zhao, *Environ. Sci. Technol.*, 2004, **38**, 329–337.
- 8 J. Yang, C. C. Chen, H. W. Ji, W. H. Ma and J. C. Zhao, *J. Phys. Chem. B*, 2005, **109**, 21900–21907.
- 9 D. Wang, L. Zhao, D. Wang, L. Yan, C. Jing, H. Zhang, L.-H. Guo and N. Tang, *Phys. Chem. Chem. Phys.*, 2018, **20**, 18978–18985.
- 10 J. Sun, H. Zhang, L.-H. Guo and L. Zhao, *ACS Appl. Mater. Interfaces*, 2013, **5**, 13035–13041.
- 11 M. W. Sutherland and B. A. Learmonth, *Free Radical Res.*, 1997, **27**, 283–289.
- 12 B. H. J. Blelski, G. G. Shiue and S. Bajuk, *J. Phys. Chem.*, 1980, **84**, 830–833.
- 13 Y. Li, W. Zhang, J. Niu and Y. Chen, *ACS Nano*, 2012, **6**, 5164–5173.
- 14 S. T. Wang, N. G. Zhegalova, T. P. Gustafson, A. Zhou, J. Sher, S. Achilefu, O. Y. Berezin and M. Y. Berezin, *Analyst*, 2013, **138**, 4363–4369.
- 15 K.-I. Ishibashi, A. Fujishima, T. Watanabe, K. Hashimoto, K. Ishibashi, A. Fujishima and T. Watanabe, *J. Photochem. Photobiol., A*, 2000, **134**, 139–142.
- 16 D. Wang, L. Zhao, L. H. Guo and H. Zhang, *Anal. Chem.*, 2014, **86**, 10535–10539.
- 17 Y. Nosaka and A. Y. Nosaka, *Chem. Rev.*, 2017, **117**, 11302–11336.
- 18 T. Daimon, T. Hirakawa, M. Kiazawa, J. Suetake and Y. Nosaka, *Appl. Catal., A*, 2008, **340**, 169–175.
- 19 C. M. Wang, A. Heller and H. Gerischer, *J. Am. Chem. Soc.*, 1992, **114**, 5230–5234.
- 20 H. Gerischer and A. Heller, *J. Electrochem. Soc.*, 1992, **139**, 113–118.
- 21 A. M. Peiró, C. Colombo, G. Doyle, J. Nelson, A. Mills and J. R. Durrant, *J. Phys. Chem. B*, 2006, **110**, 23255–23263.
- 22 F. Zuo, L. Wang, T. Wu, Z. Zhang, D. Borchardt and P. Feng, *J. Am. Chem. Soc.*, 2010, **132**, 11856–11857.

