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Chemical ionization of glyoxal and formaldehyde with H_3O^+ ions using SIFT-MS under variable system humidity

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Glyoxal ($\text{C}_2\text{H}_2\text{O}_2$) is a highly reactive molecule present at trace levels in specific gaseous environments. For analyses by chemical ionization mass spectrometry, it is important to understand the gas-phase chemistry initiated by reactions of H_3O^+ ions with $\text{C}_2\text{H}_2\text{O}_2$ molecules in the presence of water vapour. This chemistry was studied at variable humidity using a selected ion flow tube, SIFT. The initial step is a proton transfer reaction forming protonated glyoxal $\text{C}_2\text{H}_3\text{O}_2^+$. The second step, in the presence of water vapour, is the association forming $\text{C}_2\text{H}_3\text{O}_2^+(\text{H}_2\text{O})$ and interestingly also protonated formaldehyde CH_2OH^+ . Hydrated protonated formaldehyde $\text{CH}_2\text{OH}^+(\text{H}_2\text{O})$ was also observed. Relative signals of these four ionic products were studied at the end of the flow tube where the reactions took place during 0.3 ms in helium carrier gas (1.5 mbar, 300 K) as the water vapour number density varied up to 10^{14} cm^{-3} . The data were interpreted using numerical kinetics modelling of the reaction sequences and the mechanisms and kinetics of the reaction steps were characterised. The results thus facilitate SIFT-MS analyses of glyoxal in humid air whilst drawing attention to ion overlaps with formaldehyde products.

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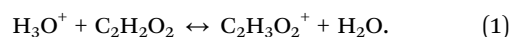
1 Introduction

Glyoxal ($\text{C}_2\text{H}_2\text{O}_2$) is the simplest dialdehyde. It can polymerize from an anhydrous monomeric form into a series of oligomers and its vapour is highly reactive, contributing to aerosol formation.^{1,2} Glyoxal plays a role in biological processes; it is known to be a cytotoxic metabolite of glucose, a product of lipid peroxidation and a marker of oxidation stress.³ In addition, recent studies have shown that glyoxal is an intermediate product in the photocatalytic reduction of carbon dioxide to methane.^{4,5} It is noteworthy that organic molecules present in the Earth's atmosphere can be transformed into glyoxal by oxidation or photo-oxidation.⁶ Such processes were observed for toluene,^{7–9} xylene,^{8–10} trimethylbenzene,^{9,10} isoprene¹¹ and α -pinene.¹² Daytime atmospheric glyoxal chemistry involves oxidation and may lead to the formation of secondary organic aerosols.¹³

Chemical ionization mass spectrometry (CIMS) is a useful tool for the detection of glyoxal since the high reactivity and

short lifetime of glyoxal impede the use of gas chromatography (GC) techniques. Nölscher *et al.* measured the total OH reactivity of tropical forest air and compared it to the reactivity expected from volatile organic compounds (VOCs) determined *via* proton-transfer-reaction mass spectrometry (PTR-MS).¹⁴ The measured OH reactivity exceeded the calculated values, indicating that either PTR-MS underestimated VOC concentrations or that not all relevant species were accounted for. This can also be caused by the different OH reactivity of isomers (which cannot be separated by PTR-MS) or the presence of molecules with high reactivity and low proton affinity (PA) compared to the PA of water (protonated water, H_3O^+ , is the most common reagent ion in PTR-MS).

The PA of glyoxal was previously estimated by HF and DFT calculations to be 675 kJ mol^{-1} and 690 kJ mol^{-1} , respectively.¹⁵ Since the PA is very similar to the PA of water (691 kJ mol^{-1}),¹⁶ proton transfer from hydronium (H_3O^+) ions may not proceed at a collisional rate.¹⁷ In addition, the reverse reaction could also take place:



CIMS analysis of atmospheric glyoxal can be further hampered by isobaric molecules, *e.g.* acetone or propanal. These are typically present at much higher concentrations, causing signal overlaps. Understanding the ion chemistry may help to distinguish these isobars.

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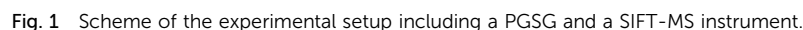
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Thus, we have carried out a detailed SIFT study with the aim to elucidate the influence of humidity on the glyoxal hydronium ion-chemistry. The experimental study was supplemented with the theoretical modelling of ion chemistry considering a sequence of individual ion-molecule reactions. We further studied the ion chemistry of formaldehyde under the same

A Profile 3 SIFT-MS instrument (Instrument Science, Crewe, UK, see Fig. 1)^{24–26} was used for the ion-chemistry study. Since SIFT-MS is described in detail in the literature, only a brief summary will be given here. A variety of reagent ions (H_3O^+ , NO^+ or O_2^+) is generated by the introduction of a water vapour/air mixture into a microwave discharge. Reagent ions are selected by a quadrupole mass filter and injected into the 5 cm long flow tube through which a He carrier gas flow was established ($p = 1.5$ mbar, $T = 300$ K). The neutral reactant gas mixture is

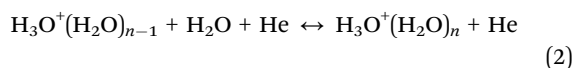


injected at a known flow rate into the flow tube, where the ion-molecule reactions take place. Ions are subsequently separated using a quadrupole mass spectrometer and detected *via* a single channel electron multiplier (SCEM).

For glyoxal experiments, 2 mg of glyoxal crystals were placed into a 2 ml glass vial closed by a septum penetrated by a 5 cm long polyether ether ketone capillary (I.D. 0.25 mm), in a diffusion tube configuration.²⁷ A precision gas standard generator (PGSG; 491M, Kin-Tek) was used at 50 °C to release glyoxal vapour and dilute it with synthetic air to reach the concentration of 15 ppmv, which was then introduced into the flow tube. The humidity of the carrier gas was varied using a reservoir containing 5 ml of water connected to the supply line, thus avoiding the interaction of water with glyoxal prior to its introduction to the SIFT. The water reservoir temperature was slowly increased from −78.5 °C to 60 °C, releasing H₂O number densities in 1.5 mbar of He between 4×10^{12} and 10^{14} cm^{−3} (corresponding to 0.3% to 7.5% sample humidity at a typical SIFT-MS sample flow). To study the influence of He carrier gas on the ion distribution, we varied the flow tube pressure from 0.5 to 2 mbar. Measurements were then carried out at two different water concentrations: $H = 0.03$ (3×10^{12} cm^{−3}) and $H = 0.08$ (1×10^{13} cm^{−3}), see (4), estimated under standard experimental conditions.

For formaldehyde experiments, a 4 l Nalophan bag containing 5 ppmv formaldehyde vapour in synthetic air was used. Humidity was varied by adding a few millilitres of water at room temperature.

The water vapour concentration in the flow tube was determined from the association reactions of (hydrated) H₃O⁺ ions with H₂O molecules,



where $n \geq 1$. The H₂O number density in the flow tube [H₂O] was calculated as

$$[\text{H}_2\text{O}] = \frac{1}{tk_{2\text{eff}}} \ln \frac{[\text{H}_3\text{O}^+] + [\text{H}_3\text{O}^+ \cdot \text{H}_2\text{O}] + [\text{H}_3\text{O}^+ \cdot (\text{H}_2\text{O})_2] + [\text{H}_3\text{O}^+ \cdot (\text{H}_2\text{O})_3]}{[\text{H}_3\text{O}^+]} \quad (3)$$

where $k_{2\text{eff}}$ is the rate coefficient for reaction (2) ($n = 1$), t is the reaction time, and the symbols in brackets correspond to respective ion count rates. Ion count rates were corrected for mass discrimination and for differential diffusion.²⁸ To compare the experimental results with the numerical model, it is useful to use the dimensionless logarithmic factor

$$H = \ln \left(\frac{[\text{H}_3\text{O}^+] + [\text{H}_3\text{O}^+ \cdot \text{H}_2\text{O}] + [\text{H}_3\text{O}^+ \cdot (\text{H}_2\text{O})_2] + \dots}{[\text{H}_3\text{O}^+]} \right) \quad (4)$$

as a relative measure of water vapour concentration.

2.3 Numerical model KIMI

To model the ion-chemistry kinetics in a flow tube, we developed the Kinetics of Ion Molecular Interaction software

(KIMI; using the Microsoft Visual Studio 2017, NET Framework 4.6.1). KIMI interactively solves sets of kinetic equations that describe the change in reagent and product ion concentrations along the axis of the flow tube and thus model ion concentrations at the end of the flow tube as a function of the reactant concentrations.

The input for the model is generated from a graphical interface that allows drawing reaction pathways and systems involving multiple ionic species. Each species has a defined initial concentration and each reaction path is described by a rate constant k . In addition, diffusion losses of ions are characterised by diffusion coefficients, D_e ,²⁹ and the characteristic radial diffusion length Λ . Concentrations of all species in the reaction system are calculated by solving a set of time dependent differential equations, exemplified by the proton transfer reaction

$$\frac{d[\text{MH}^+]}{dt} = k[\text{M}][\text{H}_3\text{O}^+] - [\text{MH}^+] \frac{D_e}{\Lambda^2} \quad (5)$$

using the Runge–Kutta method.^{30,31} In the SIFT, the time coordinate directly corresponds to the axial position due to a constant ion velocity v_i , estimated from the carrier gas velocity v_g as³²

$$v_i = 1.5 \times v_g \quad (6)$$

v_g can be calculated from the carrier gas flow rate, Φ_C , its pressure, p_g , temperature, T_g (in K) and the internal flow tube diameter d_t as

$$v_g = \frac{\Phi_C}{p_g} \frac{4}{\pi d_t^2} \frac{T_g}{273 \text{ K}} \quad (7)$$

T_g is assumed to be in equilibrium with the flow tube (usually around 300 K). The pressure is measured at the end of the flow tube. The pressure inside the flow tube is somewhat higher due to its finite conductivity according to the Hagen–Poiseuille equation

$$\Delta p = \frac{8\eta L \Phi}{\pi R^4 \bar{p}} \quad (8)$$

Under conditions of the experiments ($p = 1.5$ mbar, $T = 300$ K, $\Phi = 880$ sccm, $L = 5$ cm, flow tube radius $R = 0.5$ cm, and the dynamic viscosity of He $\eta = 2 \times 10^{-5}$ Pa s at 300 K), the pressure difference is about 0.3 mbar. To estimate the average flow tube pressure, the measured pressure was thus increased by 0.15 mbar. The total flow rate was measured for several values of pressure in the flow tube. Under typical conditions, the gas bulk velocity according to (7) is equal to 105 m s^{-1} and thus the ion velocity calculated from (6) is 157.5 m s^{-1} . The reaction time for the reaction length $L = 4$ cm corrected by a positive “end correction”³² $\varepsilon = 0.5$ cm

$$t_r = \frac{L + \varepsilon}{v_i} \quad (9)$$

is thus 286 μs .



3 Results and discussion

3.1 Experimental observations in the SIFT

The mass spectra (see Fig. 2) showed protonated glyoxal ($\text{C}_2\text{H}_3\text{O}_2^+$, m/z 59) and its water cluster ($\text{C}_2\text{H}_3\text{O}_2^+(\text{H}_2\text{O})$, m/z 77). The second water cluster ($\text{C}_2\text{H}_3\text{O}_2^+(\text{H}_2\text{O})_2$, m/z 95) appeared only at higher water vapour concentrations. The ion at m/z 31, protonated formaldehyde (CH_2OH^+),²⁰ was also observed together with its water cluster ($\text{CH}_2\text{OH}^+(\text{H}_2\text{O})$, m/z 49). A small signal at m/z 89 corresponds to the association of protonated formaldehyde and glyoxal ($\text{CH}_2\text{OH}^+\cdot\text{C}_2\text{H}_2\text{O}_2$).

The change in the relative intensities of product ions as a function of relative water vapour concentration H is shown in Fig. 3. For the lowest humidity ($H < 0.1$), protonated glyoxal was dominant ($>85\%$). With increasing humidity $H > 0.4$, protonated formaldehyde concentration rose and became dominant. It is interesting that both hydrated formaldehyde (m/z 49) and hydrated glyoxal (m/z 77) concentration remained below 20%. Water vapour thus removes the glyoxal containing ions and leads to the production of formaldehyde containing ions. This is in stark contrast to analogous acetone ion chemistry that is well understood to terminate under the same conditions in water cluster ions at m/z 77 and m/z 95.³³

Fig. 3 indicates that protonated formaldehyde is formed when H_3O^+ ions react with $\text{C}_2\text{H}_2\text{O}_2$ at 300 K in the presence of water molecules, without requiring additional collision energy. Glyoxal and formaldehyde have $\text{PA}(\text{C}_2\text{H}_2\text{O}_2) = 675$ to 690 kJ mol^{-1} ¹⁵ and $\text{PA}(\text{H}_2\text{CO}) = 713 \text{ kJ mol}^{-1}$,¹⁶ both close to the PA of water, $\text{PA}(\text{H}_2\text{O}) = 691 \text{ kJ mol}^{-1}$,¹⁶ but much lower than the PA of the water dimer, $\text{PA}(\text{H}_2\text{O})_2 = 808 \text{ kJ mol}^{-1}$.³⁴ Therefore, proton transfer is only possible from H_3O^+ (1) but not from $\text{H}_3\text{O}^+(\text{H}_2\text{O})_{1,2,3}$.

Protonated formaldehyde was not observed in the absence of H_2O and thus must be formed in a secondary reaction of protonated glyoxal with water



Alternatively, a reaction sequence can occur where an intermediate glyoxal water cluster may be formed by a ligand switching reaction

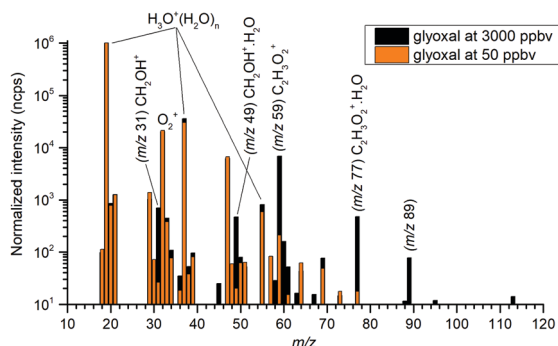
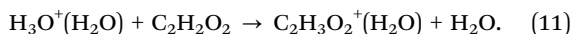


Fig. 2 Mass spectra of products of H_3O^+ reaction with glyoxal, normalized to $10^6 \text{ H}_3\text{O}^+$ reagent ions. Glyoxal concentrations are indicated by bar colour. The data were obtained for the lowest attainable water vapour concentration $H = 0.03$ ($3 \times 10^{12} \text{ cm}^{-3}$).

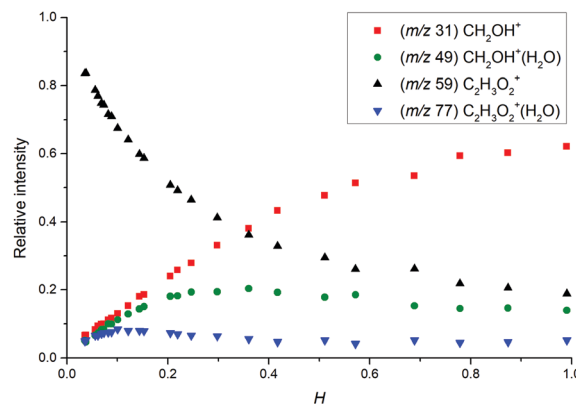
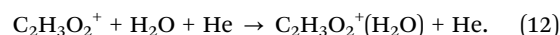
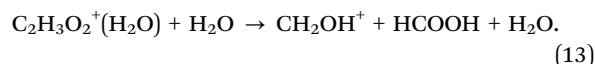


Fig. 3 Change in the relative intensities of product ions of the reaction of H_3O^+ with glyoxal for different carrier gas humidity.

It is also possible that the intermediate glyoxal water cluster is formed by a three-body association reaction



Reaction (11) and (12) may then yield protonated formaldehyde *via* dissociation in a further reaction with H_2O



Energetics of reactions (10)–(13) were calculated using the B3LYP/6-311G(p,d) level of theory in the Gaussian 09 software package as the change in the sum of the electronic and thermal free energies between optimized geometries of all reactants and products. According to this theory, reaction (10) is exergonic, $\Delta G = -64.81 \text{ kJ mol}^{-1}$, while reaction (13) is endergonic, $\Delta G = 43.85 \text{ kJ mol}^{-1}$. However, this can be overcome by internal excitation of the product $\text{C}_2\text{H}_3\text{O}_2^+(\text{H}_2\text{O})$ of the highly exergonic reaction (12), $\Delta G = -108.66 \text{ kJ mol}^{-1}$. Such excitation is unlikely for the endergonic reaction (11) ($\Delta G = 31.4 \text{ kJ mol}^{-1}$). This set of possible reactions leading to the formation of protonated formaldehyde was numerically modelled to match the experimental data (see below in Section 3.2).

Additional data were also obtained for the NO^+ and O_2^+ reactions with $\text{C}_2\text{H}_2\text{O}_2$. It was observed that the NO^+ reaction with glyoxal proceeds *via* charge transfer forming $\text{C}_2\text{H}_2\text{O}_2^+$ (m/z 58) and *via* association forming $\text{C}_2\text{H}_2\text{O}_2\text{NO}^+$ (m/z 88). For $H \sim 0.04$, the relative abundance ratio was 0.3 for $\text{C}_2\text{H}_2\text{O}_2^+$ and 0.7 for $\text{C}_2\text{H}_2\text{O}_2\text{NO}^+$. For higher humidity ($H > 0.4$), $\text{C}_2\text{H}_2\text{O}_2\text{NO}^+$ was dominant (>0.95). O_2^+ reacts by charge transfer, forming $\text{C}_2\text{H}_2\text{O}_2^+$ (m/z 58); humidity did not have a significant effect and, interestingly, the CHO^+ fragment (present in the electron ionisation mass spectra⁴⁴) was not observed.

3.2 Numerical simulation of ion chemistry

3.2.1 Distribution of reagent ions. The H_3O^+ reagent ions form $\text{H}_3\text{O}^+(\text{H}_2\text{O})_{1,2,3}$ in a sequence of three-body association reactions (2) with H_2O in helium.³⁷ The rate coefficients for these reactions obtained in previous studies are given in Table 1.



Table 1 List of the ion–molecule reactions between hydronium ions (H_3O^+) and water or formaldehyde occurring in the flow tube followed by available rate constants

Reaction		Name	k	k_{reverse}
Water				
$\text{H}_3\text{O}^+ + \text{H}_2\text{O} + \text{He} \rightarrow$	$\text{H}_3\text{O}^+(\text{H}_2\text{O}) + \text{He}$	k_{w1}	$6.65 \times 10^{-28} \text{ }^a$ $7.2 \times 10^{-28} \text{ }^b$ $5.8 \times 10^{-28} \text{ }^c$	
$\text{H}_3\text{O}^+(\text{H}_2\text{O}) + \text{H}_2\text{O} + \text{He} \rightarrow$	$\text{H}_3\text{O}^+(\text{H}_2\text{O})_2 + \text{He}$	k_{w2}	$1.51 \times 10^{-27} \text{ }^a$ $1.59 \times 10^{-27} \text{ }^c$	
$\text{H}_3\text{O}^+(\text{H}_2\text{O})_2 + \text{H}_2\text{O} + \text{He} \rightarrow$	$\text{H}_3\text{O}^+(\text{H}_2\text{O})_3 + \text{He}$	k_{w3}	$1.50 \times 10^{-27} \text{ }^a$ $1.73 \times 10^{-27} \text{ }^c$	$4 \times 10^{-15} \text{ }^a$
$\text{H}_3\text{O}^+(\text{H}_2\text{O})_3 + \text{H}_2\text{O} + \text{He} \rightarrow$	$\text{H}_3\text{O}^+(\text{H}_2\text{O})_4 + \text{He}$	k_{w4}	nd	nd
Formaldehyde				
$\text{H}_3\text{O}^+ + \text{CH}_2\text{O} \rightarrow$	$\text{CH}_2\text{OH}^+ + \text{H}_2\text{O}$	k_{f1}	$3.4 \times 10^{-9} \text{ }^{de}$ $3.3 \times 10^{-9} \text{ }^f$ $3.0 \times 10^{-9} \text{ }^{gi}$	$5.6 \times 10^{-13} \text{ }^d$ $2.3 \times 10^{-12} \text{ }^h$
$\text{H}_3\text{O}^+(\text{H}_2\text{O}) + \text{CH}_2\text{O} \rightarrow$	$\text{CH}_2\text{OH}^+(\text{H}_2\text{O}) + \text{H}_2\text{O}$	k_{f2}	$3.0 \times 10^{-9} \text{ }^d$ $2.3 \times 10^{-9} \text{ }^f$	$5.3 \times 10^{-10} \text{ }^d$ $3.5 \times 10^{-10} \text{ }^h$
$\text{CH}_2\text{OH}^+ + \text{H}_2\text{O} + \text{He} \rightarrow$	$\text{CH}_2\text{OH}^+(\text{H}_2\text{O}) + \text{He}$	k_{f3}	$2.3 \times 10^{-27} \text{ }^g$ $5.7 \times 10^{-28} \text{ }^h$	

Nd – no data. ^a From ref. 35 at 296 K. ^b From ref. 36 at 300 K. ^c From ref. 37 at 298 K. ^d From ref. 38 at 298 K. ^e From ref. 39 at 297 K in H_2 . ^f From ref. 40 at 294 K. ^g From ref. 41 at 300 K. ^h From ref. 42 at 300 K and 0.05 eV kinetic energy. ⁱ From ref. 43 at 300 K. Rate constants of bimolecular reactions are given in $\text{cm}^3 \text{s}^{-1}$ and in $\text{cm}^6 \text{s}^{-1}$ in the case of three-body reactions.

The distribution of $\text{H}_3\text{O}^+(\text{H}_2\text{O})_{1,2,3}$ in the SIFT is affected by the presence of O_2 and N_2 molecules; in the present experiment, 5% of the carrier gas. A trace amount of an organic sample (glyoxal or formaldehyde) does not affect the distribution. Effective values of $k_{w1,2,3,4}$ under experimental conditions presented herein ($p = 1.5 \text{ mbar}$, $T = 300 \text{ K}$, $[\text{H}_2\text{O}] \cong 10^{12} - 10^{14} \text{ cm}^{-3}$, $[\text{He}] = 3.62 \times 10^{16} \text{ cm}^{-3}$) were derived by fitting the data shown in Fig. 4 and are given in Table 2. These k values were then used for the modelling of formaldehyde and glyoxal ion chemistry.

3.2.2 Ion chemistry of formaldehyde. Table 1 summarizes the previous studies of the ion chemistry of formaldehyde.^{38–43} It is understood that the reverse reaction can take place due to the close PA of water and formaldehyde.⁴⁵ A protonated formaldehyde water cluster $\text{CH}_2\text{OH}^+(\text{H}_2\text{O})$ can be formed by ligand

Table 2 List of rate constants (k_n , k_{-n} for reverse reactions) relevant to the formation of hydronium clusters as given in the literature and as obtained by fitting the observed distributions in the present formaldehyde and glyoxal experiments

K	Literature	Formaldehyde	Glyoxal
k_{w1}	$5.80 \times 10^{-28} b$	6.80×10^{-28}	7.10×10^{-28}
k_{w2}	$1.59 \times 10^{-27} b$	1.59×10^{-27}	1.59×10^{-27}
k_{w3}	$1.73 \times 10^{-27} b$	1.40×10^{-27}	1.50×10^{-27}
k_{w4}	$1.00 \times 10^{-28} a$	2.50×10^{-29}	3.00×10^{-29}
k_{w-3}	$4.00 \times 10^{-15} c$	4.00×10^{-15}	4.00×10^{-15}
k_{w-4}	$1.00 \times 10^{-15} a$	1.00×10^{-15}	1.00×10^{-15}

^a Data added for the initial simulation of the $\text{H}_3\text{O}^+(\text{H}_2\text{O})_4$ cluster with m/z 91. ^b From ref. 37. ^c From ref. 35.

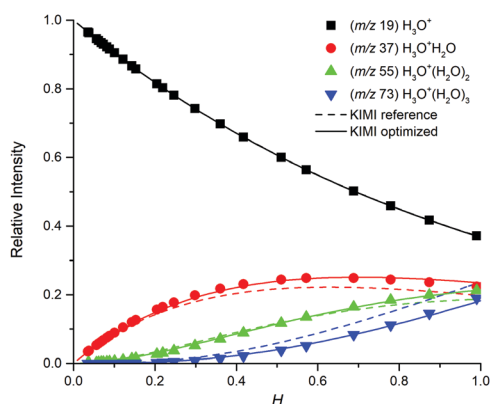


Fig. 4 Relative distribution of water clusters in the SIFT obtained by KIMI (lines) and SIFT analyses (points) of the glyoxal data set. The ion distribution calculated from literature k (dashed lines, ref. 37) does not fit the data well while the effective values given in Table 2 give good fits.

switching from the hydronium water cluster $\text{H}_3\text{O}^+(\text{H}_2\text{O})$ with reaction rate k_{f3} . As $[\text{H}_2\text{O}] \sim 10^{13} \text{ cm}^{-3}$ is much greater than $[\text{CH}_2\text{O}] \sim 10^{10} \text{ cm}^{-3}$ under the typical SIFT conditions, the reaction equilibrium favours the formation of hydronium water clusters $\text{H}_3\text{O}^+(\text{H}_2\text{O})$ even though the process is slightly endothermic. Fig. 5 illustrates experimental and model results based on the literature values and the model results fitted to the data. Note that the absolute intensities of ions decrease faster with H compared to the model prediction; this may be due to the variation of the formaldehyde concentration with time during the experiment. Thus, the fitting was done to the relative ion intensities that are not affected by this instability. The fitted values ($k_{f1} = 3.0 \times 10^{-9} \text{ cm}^3 \text{s}^{-1}$, $k_{f2} = 5.7 \times 10^{-28} \text{ cm}^6 \text{s}^{-1}$, $k_{f3} = 3.0 \times 10^{-9} \text{ cm}^3 \text{s}^{-1}$, and $k_{f-3} = 4.0 \times 10^{-10} \text{ cm}^3 \text{s}^{-1}$) were used for the modelling of formaldehyde ion chemistry observed in the glyoxal experiment.

3.2.3 Ion chemistry of glyoxal. Ion kinetics of protonated glyoxal was modelled by considering two possible mechanisms leading to the formation of protonated formaldehyde: (10), or (12) and (13).



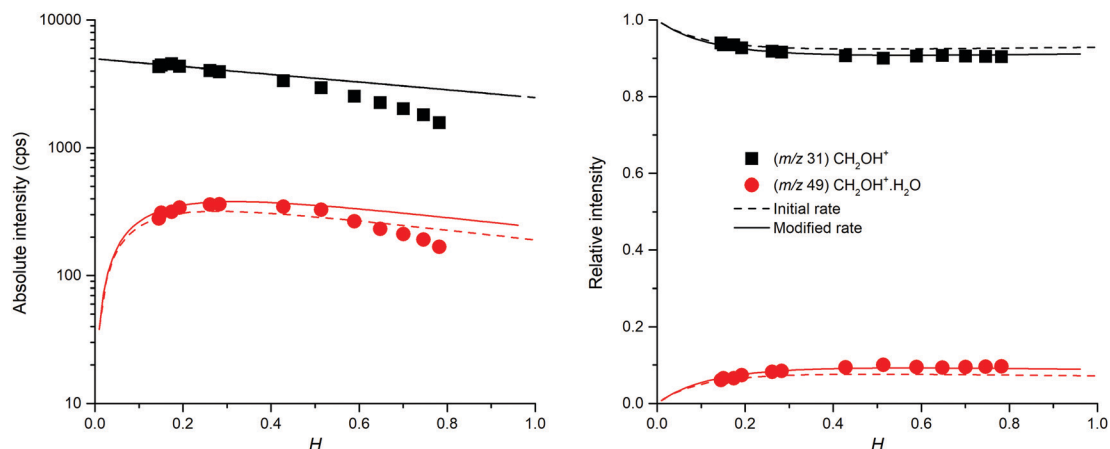


Fig. 5 Absolute (left) and relative (right) intensities of protonated formaldehyde CH_2OH^+ and its water cluster $\text{CH}_2\text{OH}^+\cdot\text{H}_2\text{O}$. The dashed line represents the reference set of selected rate constants.

We were not able to determine the distribution of vibrational energy and potential barriers of the suggested reactions. Hence,

we could not estimate whether the amount of internal energy stored within the association product of reaction (12) is

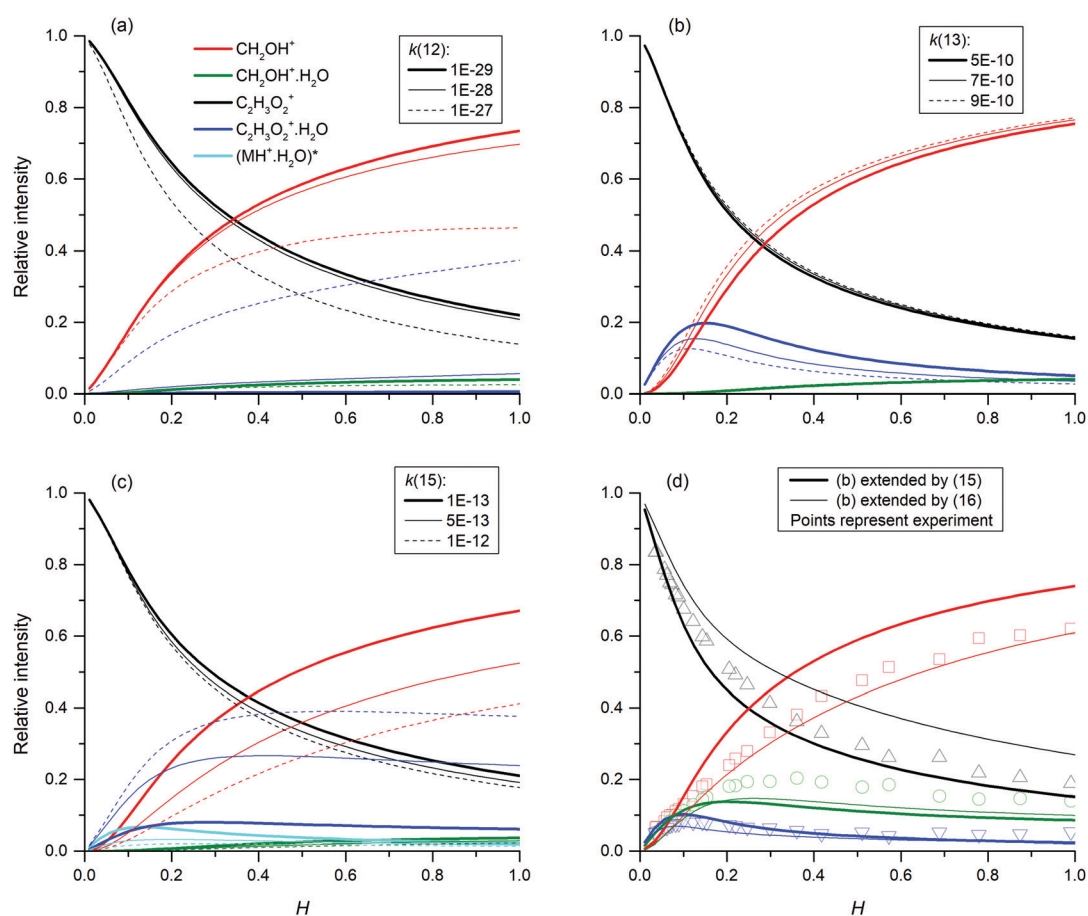
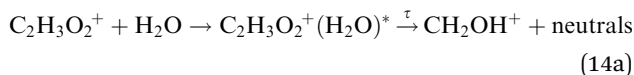


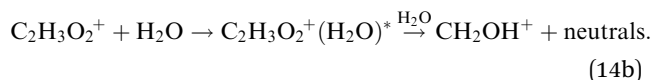
Fig. 6 Numerical modelling of possible reaction channels and their influence on relative ion distribution using different reactions and rate constants. In all presented simulations, protonated glyoxal was formed according to (1), using a reaction rate of $1.35 \times 10^{-9}\ \text{cm}^3\ \text{s}^{-1}$ in both directions. (a): reaction (10) fixed at $1 \times 10^{-10}\ \text{cm}^3\ \text{s}^{-1}$ and reaction (12) variable (units $\text{cm}^6\ \text{s}^{-1}$) (b): reaction (12) fixed at $3.7 \times 10^{-27}\ \text{cm}^6\ \text{s}^{-1}$ and reaction (13) variable (units $\text{cm}^3\ \text{s}^{-1}$) (c): reaction sequence (14b) fixed at the association rate $1 \times 10^{-10}\ \text{cm}^3\ \text{s}^{-1}$ and the dissociation rate $1 \times 10^{-9}\ \text{cm}^3\ \text{s}^{-1}$ to (15) variable. (units $\text{cm}^3\ \text{s}^{-1}$) (d): reaction (13) at $1 \times 10^{-9}\ \text{cm}^3\ \text{s}^{-1}$ followed by (16) at $k = 1 \times 10^{-10}\ \text{cm}^3\ \text{s}^{-1}$ (thick lines) and reaction (13) at $0.89 \times 10^{-9}\ \text{cm}^3\ \text{s}^{-1}$ followed by (17) at $1 \times 10^{-9}\ \text{cm}^3\ \text{s}^{-1}$.

sufficient for reaction (13) to proceed. Thus, the effect of dissociative association was explored numerically, testing whether $\text{C}_2\text{H}_3\text{O}_2^+(\text{H}_2\text{O})$ can be formed in an excited state.

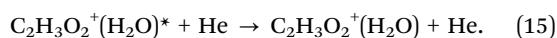
Dissociation may be unimolecular:



or induced by a collision with water molecules:



However, $\text{C}_2\text{H}_3\text{O}_2^+(\text{H}_2\text{O})^*$ can also be stabilized by collision with carrier gas or other neutral molecules and thus form a stable cluster similar to a three-body process:



The initial increase of $\text{C}_2\text{H}_3\text{O}_2^+(\text{H}_2\text{O})$ concentration with H observed in the experiments is followed by a slow decrease for $H > 0.15$ (see Fig. 3). The model explains these observations most appropriately by combining reactions (12) and (13) with rate constants of $10^{-27} \text{ cm}^6 \text{ s}^{-1}$ and $10^{-10} \text{ cm}^3 \text{ s}^{-1}$, respectively. However, the data can also be modelled by reaction (14b), forming a reaction intermediate $\text{C}_2\text{H}_3\text{O}_2^+(\text{H}_2\text{O})^*$ with a rate constant of $10^{-10} \text{ cm}^3 \text{ s}^{-1}$ followed by its reaction with water molecules at $10^{-9} \text{ cm}^3 \text{ s}^{-1}$ and (15) at $10^{-11} \text{ cm}^3 \text{ s}^{-1}$. Fig. 6 shows model results for several combinations of rate constants.

In addition, reactions (12) and (15) were studied under variable flow tube pressure (see Fig. 7) thus changing the He number density. For two different water concentrations, an increase in the carrier gas pressure leads to a decrease in the formation of protonated formaldehyde ions and a slight increase of $\text{C}_2\text{H}_3\text{O}_2^+(\text{H}_2\text{O})$, as expected from (15). However, the water concentration was influenced by He pressure due to different He/ H_2O mixing ratios and also the diffusion of ions is accelerated at lower pressures. Therefore, these experiments cannot with certainty confirm three-body reactions, however they do not exclude them, either.

The relative intensities of CH_2OH^+ and $\text{CH}_2\text{OH}^+(\text{H}_2\text{O})$ observed in the glyoxal experiment differ from those observed in the formaldehyde experiment (Section 3.2.2). The association of protonated formaldehyde and water (see k_{F3} in Table 1) alone cannot explain this observation. An additional reaction channel forming hydrated protonated formaldehyde is therefore required. The following two reaction channels describe the relative intensities observed in the glyoxal experiment:



Both reactions contribute similarly to the model curve (see Fig. 5d) using rates of $10^{-10} \text{ cm}^3 \text{ s}^{-1}$ and $10^{-9} \text{ cm}^3 \text{ s}^{-1}$ for (16) and (17), respectively. Based on numerical simulations, we cannot favour one of them and even their simultaneous contribution is possible.

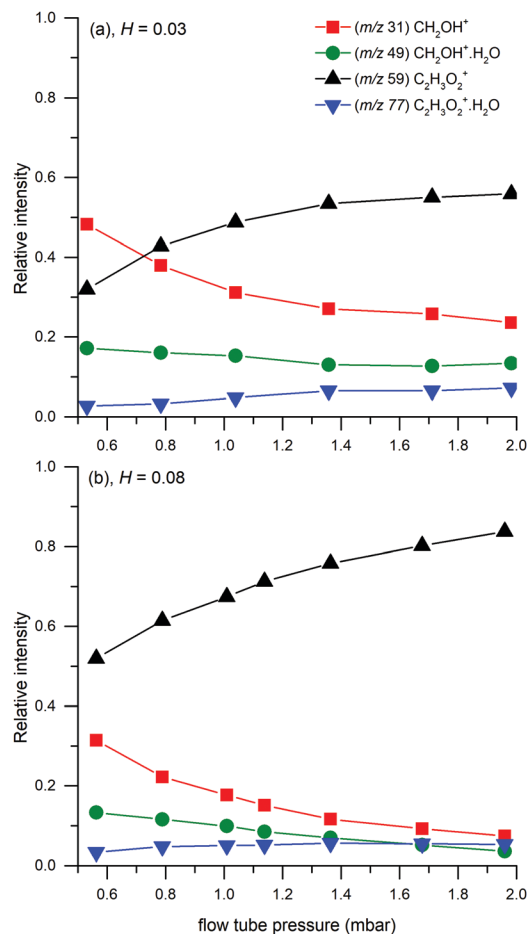


Fig. 7 Relative ion distribution as a function of the flow tube pressure. Experiments were done at $H = 0.03$ ($3 \times 10^{12} \text{ cm}^{-3}$) for (a) and $H = 0.08$ ($1 \times 10^{13} \text{ cm}^{-3}$) for (b) at 1.5 mbar.

Finally, we have modelled the endothermic ligand switching reaction (11). The high number density of neutral water molecules will shift the equilibrium in the direction of the protonated water dimer. Modelling with different reaction rate constants showed that the reaction does not affect trends in the formation of targeted species. However, it does affect the equilibrium between protonated glyoxal and protonated formaldehyde (see Fig. 8). The ligand switching reaction with water has a significant effect on the total abundance of ions, thus this effect was demonstrated together with the effect of different reaction equilibrium in (1).

The aim of the study was to describe the main reaction channels in the ion chemistry of glyoxal with hydronium in SIFT-MS. The final model of glyoxal ion chemistry consists of thirteen individual kinetic equations, combining the hydronium and formaldehyde ion chemistry given in Table 1 with eqn (1) and (11)–(13) extended by (16) and (17). Even though we carefully considered each rate constant, the final set of rate constants will not be the only solution modelling the experimental observations. It is quite possible that the resulting set can be additionally tuned to better fit the experimental data and the model does not have to be accurate on the absolute



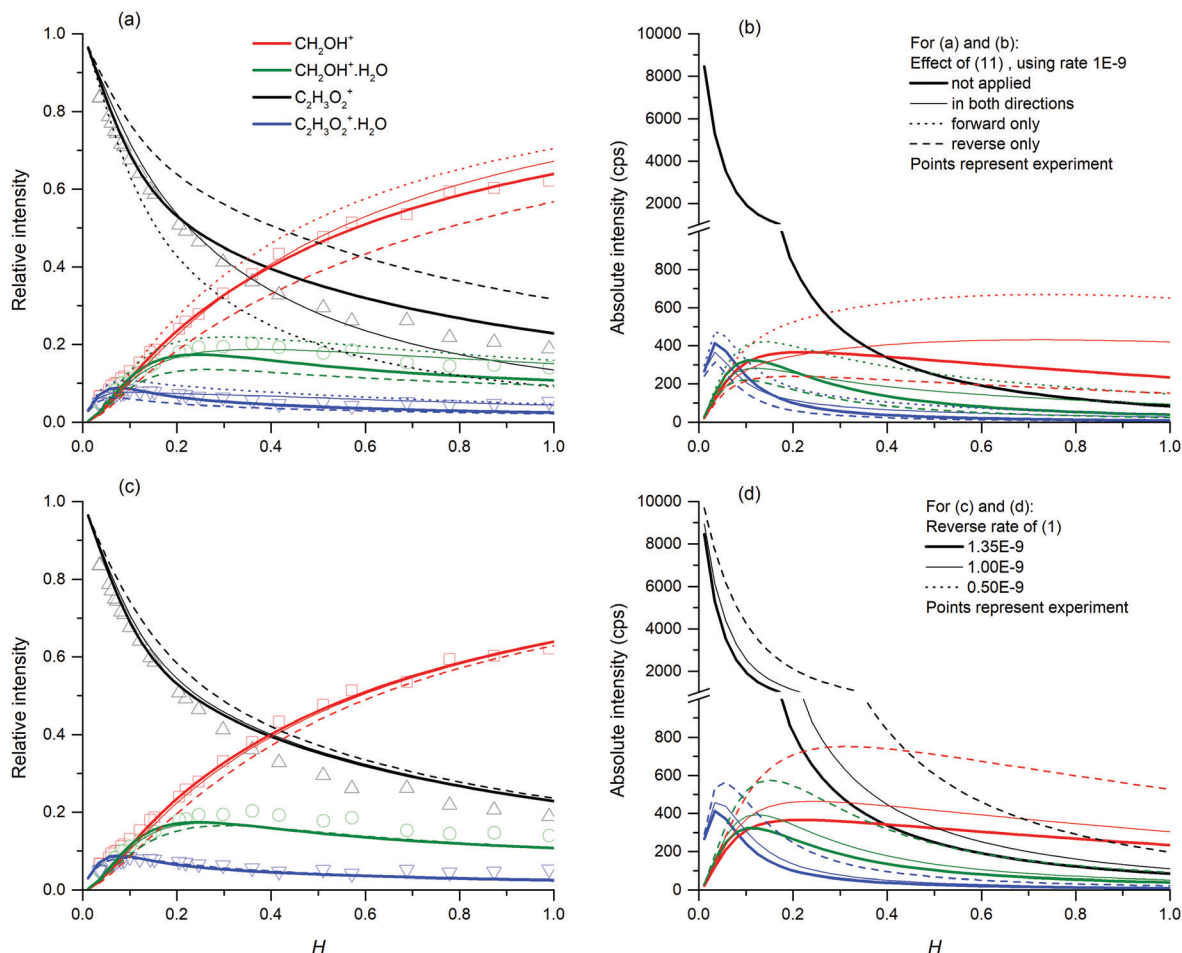


Fig. 8 Influence of (11) (top) and (1) (bottom) on the relative (left) and absolute (right) concentration of ions at the end of the flow tube. The simulation uses (13) ($k = 10^{-9} \text{ cm}^3 \text{ s}^{-1}$) and (12) ($k = 5 \times 10^{-27} \text{ cm}^6 \text{ s}^{-1}$) as the main reaction channels combined with (17) ($k = 8 \times 10^{-10} \text{ cm}^3 \text{ s}^{-1}$). The demonstration of (11) (top) uses a reverse rate of (1) with $k = 1.35 \times 10^{-9} \text{ cm}^3 \text{ s}^{-1}$. The variation of (1) (bottom) does not contain (11). The absolute scale simulation was provided using $10^{10} \text{ molecules cm}^{-3}$ glyoxal and 2×10^7 ions per cm^3 hydronium particles as an initial condition.

scale. Nevertheless, the modelling shows the dramatic effect of neutral water concentration on the ion chemistry of glyoxal.

4 Conclusions

We investigated the H_3O^+ /glyoxal ion chemistry for variable humidity using a SIFT-MS instrument. The study confirmed the secondary formation of protonated formaldehyde with increasing water concentration. For concentrations over 10^{13} water molecules per cm^3 in the flow tube, the protonated formaldehyde became the dominant product ion. The abundance of total product ions was very sensitive to water concentration, quickly decreasing for higher humidity.

The observed experimental behaviour was described using numerical simulation of ion kinetics. An additional study of hydronium water clusters and formaldehyde ion chemistry was carried out to supplement the study of glyoxal ion formation. The numerical study of glyoxal ion chemistry showed that protonated formaldehyde is formed through the reaction of $\text{C}_2\text{H}_3\text{O}_2^+(\text{H}_2\text{O})$ with water molecules. Alternative reaction channels involving water are possible, forming a protonated formaldehyde water

cluster $\text{CH}_2\text{OH}^+(\text{H}_2\text{O})$ either from protonated glyoxal (16) or from the protonated glyoxal water cluster (17).

Modifications in the glyoxal interaction with hydronium and hydronium clusters affect the total ion abundance, while changes in the reaction rate with hydronium clusters affect the reaction equilibrium as well. However, in all cases, we may observe a quick dissipation of ion signals for higher water concentration. The ion dissipation is maintained by ligand switching between protonated formaldehyde water cluster $\text{CH}_2\text{OH}^+(\text{H}_2\text{O})$ and neutral water, forming $\text{H}_3\text{O}^+\cdot\text{H}_2\text{O}$ and neutral formaldehyde. This makes the measurement of glyoxal in moist environments using CIMS particularly difficult. In addition, measurements in moist environments are much more sensitive to the influence of formaldehyde. The amount of protonated formaldehyde produced from glyoxal at H_2O concentrations greater than 10^{13} cm^{-3} corresponds to apparent 5% formaldehyde impurity in glyoxal.

Conflicts of interest

There are no conflicts to declare.



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