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# Effective suppression of lithium dendrite growth using fluorinated polysulfonamide-containing single-ion conducting polymer electrolytes†

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The lithium metal anode, which is considered as the “Holy Grail” anode material, possesses extra high energy density and extra low negative electrochemical potential. However, it is hardly applied in lithium batteries owing to lithium dendrite formation and interfacial instability in the electrolyte, resulting in safety problems and poor long-cycle performance. In this work, we designed and synthesized a novel artificial solid electrolyte interface (SEI) layer, which exhibits high ionic conductivity and single-ion conductive characteristics. The SEI layer named LiSFSI-PET4A-PETMP (LFPP) SEI layer was synthesized by a one-step click reaction *in situ* on the surface of lithium metals using lithium[(4-styrenesulfonyl) (fluorosulfonyl)imide] (LiSFSI), pentaerythritol tetraacrylate (PET), and pentaerythritol tetrakis (2-mercaptoacetate) (PETMP). The formed LFPP-SEI layer has a controllable 3D cross-linked network structure. It also exhibits high ionic conductivity and high lithium ion transference number near to unity, which is beneficial for preventing the nucleation of lithium dendrites, so as to suppress the formation of lithium dendrites from the source. The Li//LFP full cell using LFPP 2%-Li anode exhibits a quite high and stable capacity with a capacity retention of 75.8% and a high coulombic efficiency of 99.0% at 1C after 390 cycles. Moreover, the full battery using an LFPP 2%-Li anode and an LFPP 1%-Li anode at 0.5C for 250 cycles shows a very good coulombic efficiency of 100.0% and capacity retention values of 77.5% and 77.7%, respectively.

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## Introduction

The application of lithium metal anodes has been a research focus for its extra high energy density of 3860 mA h g<sup>-1</sup> and extra low negative electrochemical potential of -3.04 V *versus* standard hydrogen electrodes.<sup>3</sup> However, it is hardly applied in lithium batteries owing to lithium dendrite formation and interfacial instability in the electrolyte, resulting in safety problems and poor long-cycle performance.<sup>1,2</sup> In addition, it suffers from volume expansion during cycle performance, which can also weaken the cycle performance.<sup>4-6</sup>

In recent years, much effort has been devoted to suppressing the lithium dendrite formation.<sup>7-9</sup> A native solid electrolyte interface (SEI) layer between the liquid electrolyte and the

lithium metal can spontaneously shape up due to the high reactivity of the electrolyte solvent, such as cyclic carbonate (*e.g.*, ethylene carbonate, EC) and cyclic ether (*e.g.*, 1,3-dioxolane, DOL). But the native SEI layer exhibits low ionic conductivity ( $4.2 \times 10^{-8}$  S cm<sup>-1</sup>),<sup>10</sup> structural instability and chemical heterogeneity,<sup>11</sup> which induce heterogeneous electrodeposition resulting in dendrite growth. Many studies have disclosed various measures such as use of solid electrolytes or gel electrolytes,<sup>4,8,12-16</sup> construction of an artificial solid electrolyte interface (SEI) layer,<sup>17-23,54</sup> design of functionalized separators,<sup>24-30</sup> and improvement of the structure of current collectors<sup>2,31-37</sup> to suppress the formation and growth of lithium dendrites. Construction of an artificial SEI layer is one good solution to solve the challenge of suppressing the formation and growth of lithium dendrites. Many kinds of artificial SEI layers such as poly(vinylidene-co-hexa-fluoro-propylene)(PVDF-HFP)/LiF,<sup>38</sup> Al<sub>2</sub>O<sub>3</sub> sputter coating,<sup>39</sup> Li polyacrylic acid (LiPAA),<sup>40</sup> polyvinylidene difluoride (PVDF),<sup>41</sup> LLZTO/Li-Nafion dual-phase SEI,<sup>42</sup> Li polylactic acid (LiPLA),<sup>8</sup> LiAlO<sub>2</sub>,<sup>43</sup> poly(ethylene vinyl alcohol-β-acrylonitrile ether) (EVOH-β-OCH<sub>2</sub>CH<sub>2</sub>CN, abbreviated as EBC)/lithium bis-tri-fluoromethyl sulfonate (LiTFSI),<sup>44</sup> Li<sub>3</sub>PO<sub>4</sub>,<sup>45</sup> graphene oxide (GO),<sup>46</sup>

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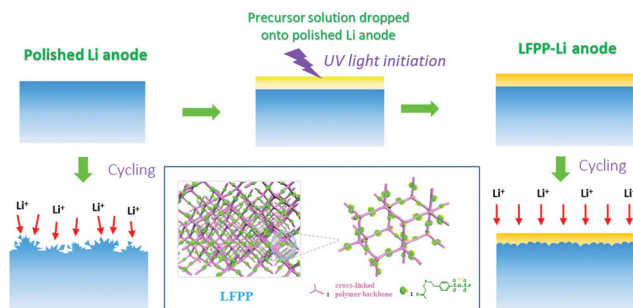


Fig. 1 Schematic of a polished Li anode and an LFPP-Li anode during cycling.

covalent organic frameworks (COFs),<sup>47</sup> and graphene/LiF<sup>48</sup> have been reported. The artificial SEI layer mentioned above can largely stabilize the active lithium metal anode; however, most of them exhibit extremely low ionic conductivity or even ionic insulation. The work by Tikekar *et al.* showed that the stable SEI layer should have the following characteristics: high ionic conductivity under strong bias, large shear modulus of the Interphase and sufficiently small size of the initial nucleates.<sup>49</sup>

Herein, we report a novel artificial SEI layer (Fig. 1) exhibiting extra high ionic conductivity ( $0.0927 \text{ mS cm}^{-1}$ ) and lithium ion transference number near to unity (0.91). The LFPP-SEI can form a uniform film on the surface of the Li anode with good flexibility and high mechanical strength. The extra high ionic conductivity and lithium ion transference number of the SEI layer can weaken or even eliminate the anion depletion-induced strong electric fields on the surface of the lithium anode to inhibit the nucleation of lithium dendrites, so as to suppress the formation of lithium dendrites from the root. The galvanostatic plating/stripping cycling performance of the symmetric cell with LFPP 1%-Li exhibits a quite stable over potential at 0.038 V for over 1000 h, indicating the superior stability of the LFPP SIPE-coated Li anode. The full LFP batteries with LFPP-Li anodes exhibit a very high coulombic efficiency of 100.0% on average and a capacity retention of 77.7% for 250 cycles.

## Results and discussion

### 1. Ionic conductivity

In our previous study, the ionic conductivity of LFPP with electrospun PVDF as a standing membrane soaked with DMC/EC (vol% = 1:1) is up to  $5.1 \times 10^{-3} \text{ S cm}^{-1}$ . To further investigate the ionic conductivity of the LFPP artificial SEI layer soaked with the solvent, the EIS analysis was carried out in symmetric coin cells assembled by two LFPP-coated stainless steel round electrodes soaked in DMC/EC/EMC (vol% = 1:1:1) using a PP membrane (Celgard 2500) as the separator.<sup>50</sup>

The LFPP SEI layer was prepared by a one-step *in situ* click reaction on the surface of the stainless steel (SS) electrode. A series of LFPP SEI layers with various thicknesses were fabricated by simply adjusting the mass concentration of precursors, consisting of LiFSI, PET4A and PETMP in DMC/EC/EMC (vol% = 1:1:1). The thickness of the LFPP SEI layers is

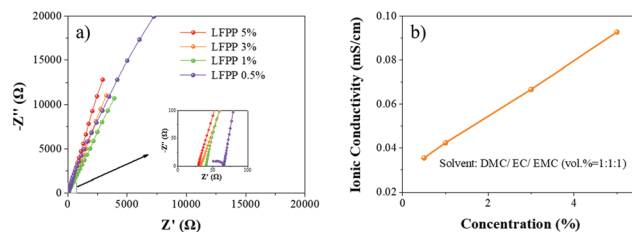


Fig. 2 EIS curves of SS//SS cells with an LFPP-coated SEI layer using DMC/EC/EMC (vol% = 1:1:1) (a) as the solvent, and the corresponding ionic conductivity versus concentration of precursors in DMC/EC/EMC (vol% = 1:1:1) (b).

not directly determined using a scanning electron microscope (SEM), because the section of SS round electrode is too thick and hard to cut off. It was calculated from the thickness of the LFPP coating on the Cu foil, which is more accessible. As the LFPP/DMC precursor solution exhibits different diffusibilities, their volumes dropped onto the Li and Cu foils are 45  $\mu\text{L}$  and 20  $\mu\text{L}$  respectively. Given that the areas of the Cu and SS foil are identical, the LFPP layer on the SS round electrode is 2.25 times as thick as that on the Cu foil. The thickness of the LFPP layer on the Cu foil can be obtained, as shown in Fig. S1 and Table S1 (ESI<sup>†</sup>).

The EIS curves and ionic conductivity calculation results are shown in Fig. 2 and Table S1 (ESI<sup>†</sup>), respectively. The ionic conductivity is very high up to  $0.0927 \text{ mS cm}^{-1}$  in DMC/EC/EMC (vol% = 1:1:1), which is much higher than that of the reported routine SEIs ( $4.2 \times 10^{-8} \text{ S cm}^{-1}$ ).<sup>10</sup> It can be seen that a higher ionic conductivity can be readily achieved by increasing the concentration of precursors, as shown in Fig. 2.

### 2. Electrochemical stability

The electrochemical stability of the electrode material determines the voltage range of lithium-ion batteries. In this work, the electrochemical stability of the LFPP SEI layer in carbonate electrolytes was measured by linear sweep voltammetry (LSV) in 2032 coin cells using stainless steel as the cathode and LFPP-Li as the anode (LFPP-Li//SS) and using a carbon cathode and LFPP-Li as the anode (Super P//LFPP-Li). The LSV measurement was carried out from 0 V to 6.0 V at a scan speed of  $1 \text{ mV s}^{-1}$ . As shown in Fig. 3a, the LFPP SEI layer is electrochemically stable up to 5.2 V. To further investigate the electrochemical stability of the LFPP SEI layer applied in carbonate electrolytes, the cyclic voltammetry was performed in a voltage range of 2.5–4.2 V at a scan speed of  $1 \text{ mV s}^{-1}$ , as shown in Fig. S4 (ESI<sup>†</sup>), indicating no obvious redox peaks. It is also confirmed that the LFPP SEI layer is electrochemically stable from 2.5 V to 4.2 V, which can then guarantee its extensive application in lithium batteries. When the electrochemical stability of the LFPP SEI layer was tested using a stainless steel foil as the working electrode, the resultant electrochemical stability window was always over-estimated. A carbon-based electrode is used in Fig. 3b. The electrochemical stability is up to 5.7 V.<sup>51</sup>

### 3. Galvanostatic Li plating/stripping cycling behaviors

To further investigate the influence of the LFPP SEI layer coated on a Li metal anode, galvanostatic cycling performance was



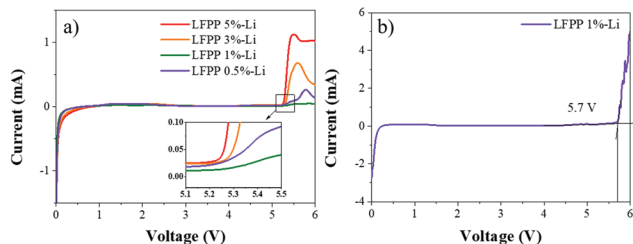


Fig. 3 Linear sweep voltammetry curves: (a) LFPP SEI layer in LFPP-Li//SS batteries and (b) LFPP SEI layer in a Super P//LFPP-Li battery.

evaluated in a symmetric LFPP-Li cell. The galvanostatic cycling of symmetric LFPP-Li||LFPP-Li cells was tested at a current density of  $0.5 \text{ mA cm}^{-2}$  with a capacity of  $1 \text{ mA h cm}^{-2}$ . As shown in Fig. 4, the over-potential of the cells with LFPP-Li increases with the increase in the concentration of LFPP precursors, owing to the increase in the thickness of the LFPP SEI layer. When the concentrations of the LFPP precursor are 5% and 2%, the galvanostatic cycling performance is stable at 0.7–1.0 V for about 950 h. After about 1100 h, the over-potential bumps up sharply, which can be attributed to the lithium dendrite formation and the increased SEI layer.<sup>40</sup> Moreover, the SFSI anion of LiSFSI salt is uniformly distributed on the surface of the lithium anode, which can guide  $\text{Li}^+$  to distribute and deposit evenly. The symmetric cell with LFPP 1%-Li exhibits the lowest and most stable over-potential at 0.038 V for over 1000 h. The outstanding cycling performance of the symmetric LFPP-Li cells with the LFPP 1%-Li SEI layer demonstrates the superior stability and lithium ion transport ability of the LFPP SEI layer.

Furthermore, XPS analysis was operated to explore the stability of the interface enhanced by LFPP SEI layers by investigating the chemical compositions on the surface of the lithium anode. Fig. 5 shows the XPS spectra for the anodes retrieved from symmetric lithium cell-based carbonate electrolytes with LFPP-Li. Comparing the anode from symmetric lithium cells with a polished Li anode after cycling for 200 h at  $0.5 \text{ mA cm}^{-2}$  with a capacity of  $1 \text{ mA h cm}^{-2}$ , there are  $\text{SO}_2\text{-F}$  (685.9 eV) and more abundant Li-F (684.4 eV) and C-O (286.3 eV) on the anode surface, which is in accordance with the chemical composition of the LFPP SEI layer.<sup>52</sup> It can be seen that the C-C, C-H peak (284.8 eV, C 1s), C-O peak (286.3 eV, C 1s),  $\text{CO}_3$  peak (288.6 eV, C 1s), poly(EC) peak (289.8 eV, C 1s),

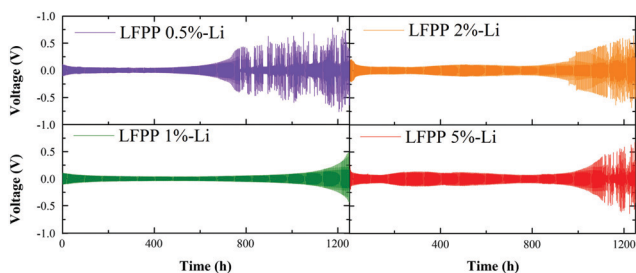


Fig. 4 Galvanostatic cycling performance of symmetric LFPP-Li||LFPP-Li cells at a current density of  $0.5 \text{ mA cm}^{-2}$  with a capacity of  $1 \text{ mA h cm}^{-2}$ .

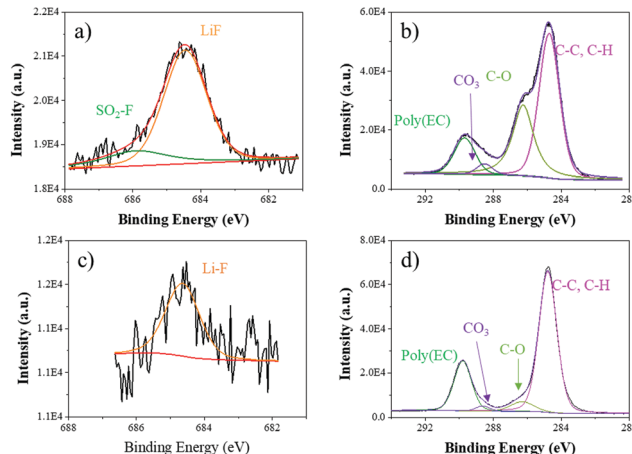


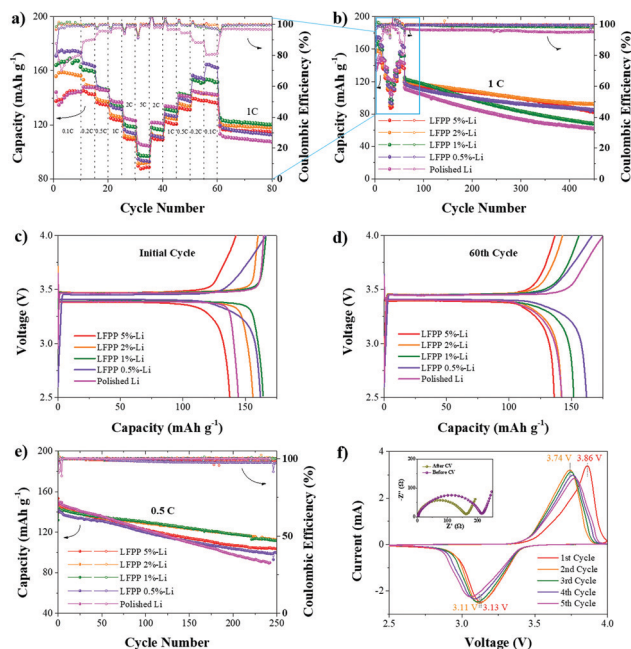
Fig. 5 XPS spectra for anodes retrieved from symmetric lithium cells based on the carbonate electrolyte with LFPP-Li after 200 h at a current density of  $0.5 \text{ mA cm}^{-2}$  with a capacity of  $1 \text{ mA h cm}^{-2}$  (a and b) and polished bare Li (c and d). (a and c) F 1s. (b and d) C 1s.

and Li-F peak (684.4 eV, F 1s) appear on the surface of the polished Li anode in a symmetric lithium cell after cycling, which are the typical composition for the native solvent-induced SEI formation. However, it has been proved that the native SEI layer induced by electrolytes is unstable, which can then tolerate chemical heterogeneity, inducing heterogeneous electrodeposition and generating dendrite growth.<sup>11</sup>

#### 4. Battery performance

The battery performance and cyclic voltammetry curves of Li//LiFePO<sub>4</sub> cells with LFPP-Li and polished Li anodes are shown in Fig. 6, respectively. Fig. 6a and b describe the rate performance of Li//LFP full batteries using the LFPP-Li anode and the polished Li anode. At a low rate, LFPP-Li//LFP batteries using LFPP-Li anodes exhibit better performance than the Li//LFP batteries using polished Li anodes. The battery using LFPP 0.5%-Li as the anode exhibits the best performance. This may be because the LFPP SEI layer exhibits good compatibility with liquid electrolytes and it has a very high lithium ion transference number near to unity, which can weaken or even eliminate the anion depletion-induced strong electric fields on the lithium anode. It is further proved that, by the charge-discharge curves of the initial cycle (Fig. 6a) and the 60th cycle of the rate performance (Fig. 6b), the full cells exhibit smaller electrode polarization ( $\Delta V$ ) and more stable voltage plateau compared with the Li//LFP battery using a polished Li anode. However, at a high rate, the battery using a polished Li anode has a higher capacity. Although the ionic conductivity of the LFPP SEI layer is much higher than the native SEI layer, it is still much lower than that of the liquid electrolyte ( $10^{-3}$ – $10^{-2} \text{ S cm}^{-1}$ ), which, in turn, restricts the battery performance at high current densities. In addition, the coulombic efficiency of the Li//LFP full cell using a polished Li anode is very low and unstable during the measurement of the rate performance at the first 60 cycles, which results from the unstable native SEI layer, lithium dendrite and dead lithium formation. This can be clearly observed from the



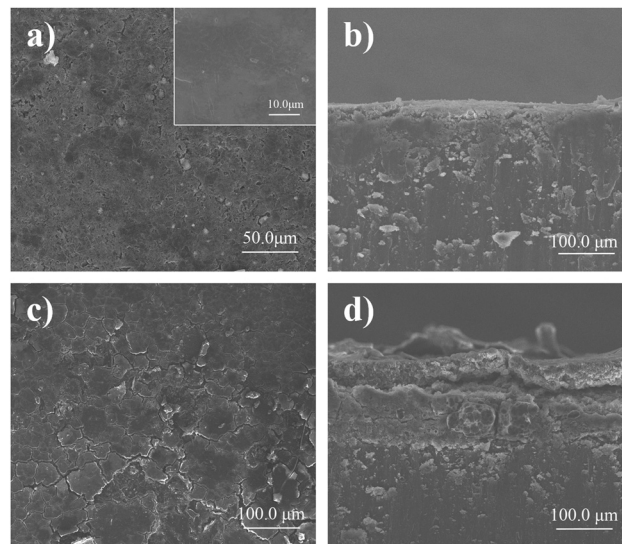


**Fig. 6** Battery performance of LFP full batteries using the LFPP-Li anode and the polished Li anode: rate performance (a and b) and the corresponding charge–discharge curves of the initial cycle (c) and the 60th cycle (d); long cycle performance at 0.5C (e); cyclic voltammograms of LFPP 1%-Li//LFP (f, inset is the EIS curve before and after CV).

SEM images shown in Fig. 7. The coulombic efficiency of LFPP-Li//LFP full cells is  $\sim 99.8\%$ , which is much higher.

The advantages of LFPP SEI layer becomes much more obvious when subjected to long cycle performance measurements at a rate of 1C, as shown in Fig. 6b, and at a rate of 0.5C in Fig. 6e. In Fig. 6b, the long cycle performance at 1C was performed from the 61st cycle to the 450th cycle. As listed in Table 1, the capacity retention and coulombic efficiency of the cell using an LFPP-Li anode is much higher than that using a polished Li anode. The cell using an LFPP 2%-Li anode exhibits the highest capacity retention of 75.8% and a high coulombic efficiency of 99.0% at 1C after 390 cycles. During the long cycle performance at 0.5C in Fig. 6e, the batteries using an LFPP 2%-Li anode and an LFPP 1%-Li anode exhibit the best performance with a very high coulombic efficiency of 100.0% and capacity retentions of 77.5% and 77.7% respectively.

The cyclic voltammogram (CV) measurement of Li LFPP 1%-Li//LFP was carried out from 2.5 V to 4.0 V at a scan rate of  $0.1 \text{ mV s}^{-1}$ . The oxidation potential and reduction potential of the first and second cycles are 3.86 V/3.13 V and 3.74 V/3.11 V, respectively, indicating the decreased polarization voltage. This is attributed to the improvement of compatibility as demonstrated further in the following EIS tests. The EIS measurements were conducted from 100 kHz to 1 Hz at a voltage of 5 mV before and after CV measurements. As shown in Fig. 6f, the interface impedance (the diameter of the semicircle) becomes smaller after 5 cycles, owing to the improvement of compatibility among the Li metal, LFPP SEI layer, and liquid electrolyte. The battery performance and CV curves of LFPP-Li



**Fig. 7** SEM images of the LFPP-Li anode (a, top view; b, side view) and the polished Li (c, top view; d, side view) anode after cycling in Li//LiFePO<sub>4</sub> cells for 500 cycles at 0.5C.

**Table 1** Coulombic efficiency and capacity retention of the long cycle performance of Li//LFP full cells using LFPP-Li anode and polished Li anode at 1C and 0.5C

	1C after rate performance		0.5C	
	CE <sup>a</sup> (%)	CR <sup>a</sup> (%)	CE (%)	CR (%)
LFPP 5%-Li	99.7	69.1	98.9	70.0
LFPP 2%-Li	99.0	75.8	100.0	77.5
LFPP 1%-Li	98.8	57.9	100.0	77.7
LFPP 0.5%-Li	99.6	71.0	98.5	70.2
Polished-Li	95.8	55.0	99.8	59.9

<sup>a</sup> CE: coulombic efficiency on average; CR: capacity retention.

applied in Li//NCM full cells are shown in Fig. S2 (ESI<sup>†</sup>). Their cycle stability is inferior to that of Li//LFP batteries.

To further evaluate the effect of the LFPP SEI layer present on the lithium metal anode surface on battery performance, the Li//LiFePO<sub>4</sub> cells with LFPP-Li and polished Li anodes were disassembled after 500 cycles, followed by the examination under a SEM. Fig. 7c and d show the top view and side view of the polished Li anode after cycling in Li//LiFePO<sub>4</sub> cells for 100 cycles at 0.5C, depicting obvious cracks from the top view and numerous lithium dendrites and dead lithium from the cross-section. Compared with the polished Li anode, the LFPP-Li anode (Fig. 7a) presents a smooth top surface with very few defects. From the cross-section in Fig. 7b, it was found that there are almost no dead lithium and very few lithium dendrites, demonstrating the effective suppression of lithium dendrite formation on the LFPP SEI layer.

## Conclusion

In summary, we report a novel artificial SEI layer, which exhibits high ionic conductivity and single-ion conductive characteristics. The LFPP SEI layer showed a very high ionic



conductivity up to 0.0927 mS cm<sup>-1</sup>. Furthermore, its lithium ion transference number is 0.91, which is quite high near to unity. These characteristics can weaken or even eliminate the anion depletion-induced strong electric fields on the lithium anode to prevent the nucleation of lithium dendrites, finally suppressing the formation of lithium dendrites from the source. The SEI layer also shows very good electrochemical stability from 0 to 5.2 V. The symmetric cell with LFPP 1%-Li exhibits a very small over-potential at 0.038 V for over 1000 h at a current density of 0.5 mA cm<sup>-2</sup> with a capacity of 1 mA h cm<sup>-2</sup>. The Li//LFP full cells using an LFPP 2%-Li anode give the most stable capacity with the highest capacity retention of 75.8% and an extra high coulombic efficiency of 99.0% at 1C after 390 cycles. During the long cycle performance at 0.5C, the batteries using LFPP 2%-Li and LFPP 1%-Li anodes exhibit the best performance with 100% coulombic efficiency on average and capacity retentions of 77.5% and 77.7%, respectively. Therefore, LFPP SEI layer is an outstanding artificial protective layer for the Li metal anode to achieve the objective of stability and dendrite growth suppression when applied in lithium metal batteries.

## Experimental section

### Chemicals and materials

Super dry solvents including acetonitrile and *N,N*-dimethylformamide were purchased from JK Chemical and used without any pretreatment. 2-Dimethoxy-2-phenylacetophenone (DMPA) (Aladdin), sodium *p*-styrenesulfonate (Aladdin), and thionyl chloride were also used as-purchased. Super P (MTI Kejing), PVDF (Arkema), and LiFePO<sub>4</sub> (MTI Kejing) were dried at 50 °C under high vacuum before use. All liquid reagents such as pentaerythritol tetraacrylate (PET4A, J&K) and pentaerythritol tetrakis (3-mercaptopropionate) (PETMP, Aldrich) were dried using 3 Å molecule sieves and stored in a glove box. Lithium[(4-styrenesulfonyl)(fluorosulfonyl)imide] (LiSFSI) was synthesized according to a procedure reported in our previous work and related studies.<sup>53</sup>

### Preparation process of LFPP-Li anodes

A precursor solution was first prepared with LiSFSI, PET4A, and PETMP using dimethyl carbonate (DMC, J&K Chemical, super dry) and DMPA as the solvent and initiator respectively. Then 70 μL precursor solution was dropped onto the top surface of a polished lithium metal foil. The polished lithium metal foil was obtained in two steps, namely, removal of the surface oxidation layer of the commercial lithium foil using a blade and rolling out using a glass rod. After thiol-ene click polymerization under 365 nm UV light for 30 min, the LFPP SEI layer took shape on the lithium foil. Although DMC is very easy to volatilize, LFPP-Li was completely dried under high vacuum at room temperature for 30 min to remove the residual solvent.

### Electrochemical measurements

An LFPP-coated stainless steel (LFPP-SS) electrode was prepared by the same method as that used for LFPP-Li preparation.

Afterwards, symmetric coin cells were assembled by two LFPP-coated stainless steel round electrodes with a diameter of 15.5 mm, using a PP (Celgard 2500) membrane as the separator and 20.0 μL DMC/EC/EMC (vol% = 1:1:1) as the solvent. Electrochemical impedance spectroscopy (EIS) was performed in the symmetric coin cells using a Solartron 1255B frequency response analyzer from 0.1 MHz to 1 Hz at an amplitude of 5 mV at 25 °C. The ionic conductivity was calculated according to the following formula:

$$\sigma = l_{\text{SEI+PP}} / (R_b \times A) \quad (1)$$

where  $l_{\text{SEI+PP}}$  is the thickness of the SEI layer and PP separator,  $R_b$  is the bulk resistance of the SEI layer and PP separator soaked with a liquid electrolyte, and  $A$  is the area of the stainless steel electrode.

The electrochemical stability of the LFPP SEI layer was tested in a standard 2032 coin cell using an LFPP-SS electrode as the cathode and LFPP-Li as the anode, and a carbon cathode and LFPP-Li as the anode (Super P//LFPP-Li) with 1 M LiPF<sub>6</sub>/DMC:EC:EMC (vol% = 1:1:1) as the liquid electrolyte. A carbon cathode was made by a casting method with a weight ratio of Super P:PVDF = 1:1. Linear scanning voltammetry (LSV) was conducted from 0 to 6.0 V (vs. Li<sup>+</sup>/Li) at a scan speed of 1 mV s<sup>-1</sup> at room temperature. Cyclic voltammogram (CV) was conducted from 2.5 V to 4.2 V at a scan speed of 1 mV s<sup>-1</sup>.

The galvanostatic cycling performance of symmetric Li//Li cells with an LFPP-Li anode or a polished Li anode was performed at a current density of 0.5 mA cm<sup>-2</sup> with a capacity of 1 mA h cm<sup>-2</sup>.

An LiFePO<sub>4</sub> (LFP) cathode was obtained by a casting method in a weight ratio of LFP:Super P:PVDF of 8:1:1. The loading density of LFP is about 1.8 mg cm<sup>-2</sup>. Cathode circles ( $\Phi = 12$  mm) were obtained using a cutting machine. The Li/LFP cells were assembled in standard CR2025 coin cells with LPFF-Li as the anode in a glove box. As control groups, batteries using a polished bare Li metal as the anode were assembled. Battery performance tests were performed in a voltage range between 2.5 V and 4.0 V at 28 °C. The Li/LFP cells using an LFPP-Li anode and a polished Li anode after 100 cycles were taken apart, and then the electrodes were washed with DMC for SEM analysis and XPS analysis.

A lithium nickel cobalt manganese oxide (NCM) cathode was fabricated by a casting method in a weight ratio of NCM 523:Super P:PVDF of 8:1:1. The loading density of NCM is 3.6 mg cm<sup>-2</sup>. Then Li/NCM cells were assembled in standard CR2025 coin cells with LFPP-Li as the anode using 1 M LiPF<sub>6</sub>/DMC:EC:EMC (vol% = 1:1:1) as the liquid electrolyte in a glove box. Battery performance tests were performed in a voltage range between 3.0 V and 4.5 V at 28 °C.

Cyclic voltammogram (CV) measurements of Li/LFP cells with an LFPP-Li anode or a bare-Li anode were operated in a voltage range between 2.5 V and 4.0 V at a scan speed of 0.1 mV s<sup>-1</sup> at room temperature. Cyclic voltammogram (CV) measurements of Li/NCM cells were conducted from 3.0 V to 4.7 V at a scan speed of 0.1 mV s<sup>-1</sup> at room temperature.



## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

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