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Mathematical modeling of the reaction of metal oxides with methane†

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Methane reduction has attracted substantial interest in recent years because an abundant amount of natural gas has been found and methane possesses a strong reduction ability. However, due to the complexity of the reaction, the reductive-kinetics model has not been developed very well. This work reported a new mathematical model for methane reduction. The model is in a form of explicit functions incorporating many parameters to increase its precision. Particularly, it considers the comparison of methane cracking rate and reaction rate. Both the gas diffusion in the product layer and chemical reaction controlled-kinetics formulae were deduced by considering three kinds of shapes (spherical, cylindrical and lamellar) of particles. Also, by employing two parameters (shape coefficient S_c and equivalent diameter d_0), the formulae for the same reduction mechanism could be unified to one formula, which was easier to use. The simulation of the model also considers both isothermal and non-isothermal processes of methane reduction. Furthermore, it can describe the reduction of oxides of varied-valence metals. The kinetics of reduction of metal oxides by methane agrees with the results obtained in the practical system.

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1. Introduction

Methane (CH_4) as an alternative reducing agent to carbon is widely used in extractive metallurgy, which decreases the operating temperature, diminishes the emission of greenhouse gases and other pollutants such as heavy metals, sulfur dioxide and fly ash,^{1,2} and has received great attention in the field of metallurgy. Methane can be widely used to reduce iron oxide,^{3,4} titanium oxide⁵ and other titanium-containing materials,^{6,7} nickel oxide,¹ celestite,⁸ tungsten oxide,⁹ cobalt oxide,¹⁰ etc.

To elucidate the experimental data, we specified the technical parameters and made the reduction process more controllable in terms of design and scale-up; it is meaningful and indispensable to establish the mathematical model of the CH_4 reduction process. Although many valuable kinetic models of gas–solid reactions have been established, such as Jander's model,¹¹ Carter's model,¹² the nucleation and growth model,^{13,14} and the Ginstling–Brounshtein model,¹⁵ unfortunately, the previously mentioned treatments are greatly simplified and these models cannot meet the requirement of today's metallurgical industry development.¹⁶ Furthermore, these models cannot be employed to describe methane reduction due to its

complexity, for instance, methane cracking occurs during the reaction, especially at a higher temperature. Moreover, two or three reaction interfaces usually appear along the radial direction in the particles during the reduction of polyvalent metal oxides, such as titanium dioxide and this also increases the difficulty of the reaction simulation.

Solving differential equations is the other kind of method for analyzing gas–solid reaction kinetics. By incorporating a series of parameters in the system, the formulae become more accurate, however, the difficulty of solving them is also increased and the numerical solutions are hard to obtain.¹⁷

Therefore, a model considering the multiple reaction interfaces is proposed for simulating the reaction between metal oxides and methane in the present study. The developed formulae can be used for modeling the reaction of metal oxides with different shapes as well. In addition, for the chemical reaction controlled conditions, the comparison of the methane cracking rate and reaction rate was incorporated.

2. General mechanism for the reduction of metal oxides by methane

Some hypotheses are proposed first: raw particles possess a uniform diameter and are nonporous, and the diameter remains constant during the reduction. Furthermore, the reduction of metal oxides by methane can be divided into nine different steps

- (i) External diffusion of methane.
- (ii) Internal diffusion of methane.

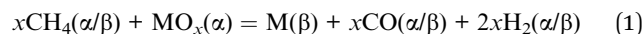
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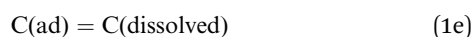
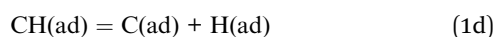
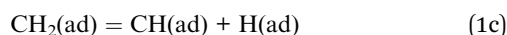
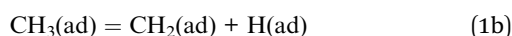
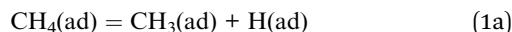


- (iii) Physical adsorption of CH₄.
- (iv) Chemical adsorption of CH₄.
- (v) Chemical reaction at the interface.



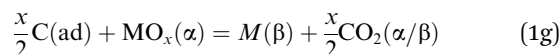
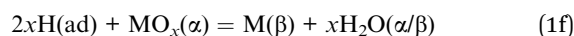
Eqn (1) shows the total reaction, in which the metal elements involved in the reaction are represented by M; the metal oxides and product phase are represented by α and β , respectively. A series of sub-steps are involved in the chemical reaction process.

- (a) Cracking of chemisorbed methane;¹⁸

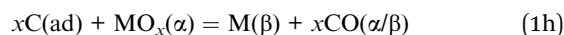


If C(ad) cannot be consumed completely, the reaction interface or blocked pores would be covered by the dissolved carbon, which is formed through eqn (1e), and retard further reduction.

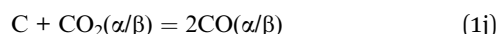
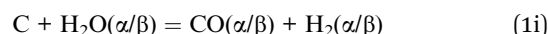
- (b) Reduction of metal oxides by H(ad) and C(ad);



or



- (c) The water-gas-shift and Boudouard reactions at the interface;



- (vi) Chemical desorption of gaseous products (CO, H₂).
- (vii) Physical desorption of gaseous products (CO, H₂).
- (viii) Internal diffusion of gaseous products (CO, H₂).
- (ix) External diffusion of gaseous products (CO, H₂).

In general, the internal diffusion of gas or interfacial chemical reaction is the restrictive step,¹⁹ which mainly depends on the properties of reaction systems. For example, the restrictive step can change if the particle size increases. Due to the similarity of formulae between the steps (ii) and (viii), this work only chooses step (ii) (internal diffusion of CH₄) as a typical example to display the derivation process and the derivation process of step (viii) is shown in the ESI.† In addition, only a brief derivation process is shown in the body of the article and the ESI† shows the detailed derivation process in the corresponding section.

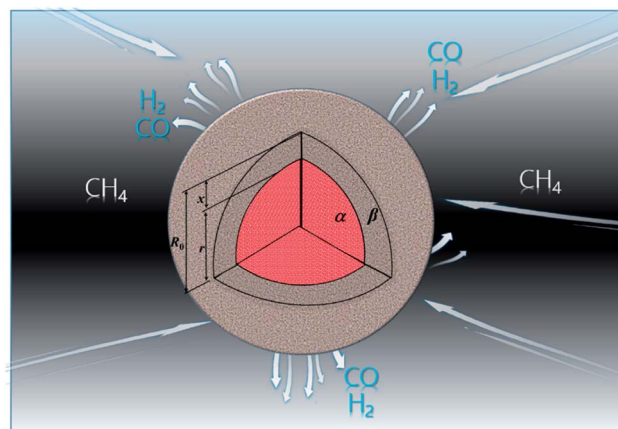


Fig. 1 Spherical particle during the reduction process with CH₄.

3. The theoretical analysis of single reaction

3.1. Interface chemical reaction

3.1.1. Spherical particles. Metal oxides with a spherical shape are shown in Fig. 1, in which x represents the thickness of the reaction product and R_0 and r represent the radii of the initial particle and unreacted metal oxides, respectively. Due to the non-porosity of the initial particle and the invariance of the overall particle shape proposed in the previous section, CH₄ reduction should obey the unreacted shrinking core mechanism,^{20,21} where the unreacted shrinking core model is commonly used in non-catalytic gas-solid reactions of metallurgy and chemical engineering.²¹ Although the unreacted shrinking core model cannot cover all the mechanisms of gas-solid reactions, it is widely accepted as the best simple model for most of the principal gas-solid reactions.^{22–24} In addition, our previous research also proved the presence of unreacted shrinking core for titania reduction.^{6,7,25,26} The ξ (reduction extent) at r can be expressed by using the following eqn (2):

$$\xi = 1 - \left(\frac{r}{R_0}\right)^3 \quad (2)$$

Eqn (3) can be obtained by the differentiation of eqn (2) with respect to time t .

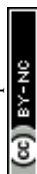
$$\frac{d\xi}{dt} = -\frac{3r^2}{R_0^3} \frac{dr}{dt} \quad (3)$$

Rates of reaction per unit area on both sides of the reversible reaction^{27–30} in eqn (1) can be calculated as follows:

$$v_r^f = K_r^f P_{\text{CH}_4}^m(\alpha/\beta) \quad (4)$$

$$v_r^b = K_r^b P_{\text{CO}}^n(\alpha/\beta) P_{\text{H}_2}^z(\alpha/\beta) \quad (5)$$

Here, reaction rate coefficients are defined by using K_r^f and K_r^b , and m , n and z are the reaction orders. For eqn (1), the overall reaction rate can be calculated per unit area by using



$$v_r = v_r^f - v_r^b = K_r^f P_{CH_4}^m(\alpha/\beta) - K_r^b P_{CO}^n(\alpha/\beta) P_{H_2}^z(\alpha/\beta) \quad (6)$$

When the interfacial chemical reaction is the control part of the reduction process, the resistance in diffusion (product layer and other steps) can be neglected, thus

$$P_{CH_4}(\alpha/\beta) = P_{CH_4} \quad (7)$$

$$P_{CO}(\alpha/\beta) = P_{CO} \quad (8)$$

$$P_{H_2}(\alpha/\beta) = P_{H_2} \quad (9)$$

The equilibrium constant of reaction (1) can be calculated as

$$K_{eq} = K_r^f / K_r^b \quad (10)$$

Substituting eqn (10) in eqn (6), eqn (11) is obtained as:

$$v_r = K_r^f \left(P_{CH_4}^m - \frac{1}{K_{eq}} P_{CO}^n P_{H_2}^z \right) \quad (11)$$

Because of the direct correlation between methane cracking and the reaction rate, two conditions are proposed: if the cracking is too fast, the deposited carbon coating on the surface of particles will retard the reaction, as discussed above; if the cracking rate is too small, the reduction will still be retarded because of the slow consumption of CO₂ and H₂O. Thus, there are two cases:

(1) When $\frac{K_C}{K_r^f} \geq 1$ (K_C is the cracking rate coefficient of methane), the growth rate of α in the radial direction, which is proportional to the overall reaction rate, can be expressed as follows:

$$\frac{dr}{dt} = -\frac{M}{\rho} \frac{K_r^f}{K_C} p_r \quad (12)$$

where M represents the molecular weight of metal oxides and ρ is the density. Thus, the reduction between the cracked products and metal oxides is the rate limiting step and the radial growth rate of α would be positively correlated with the reaction rate in this case. Due to the hindrance of excess deposited carbon, the radial growth rate of α would be negatively correlated with methane cracking. Therefore, K_C (the cracking rate coefficient of methane) and K_r^f (the reaction rate coefficient) are in the denominator and numerator, respectively. In addition, $\frac{M}{\rho}$ was obtained through the derivation of dimensions. Eqn (3) combines with eqn (11) and (12) to give the following equation.

$$\frac{d\xi}{dt} = \frac{3M(K_r^f)^2(1-\xi)^{2/3}}{R_0\rho K_C} \left(P_{CH_4}^m - \frac{1}{K_{eq}} P_{CO}^n P_{H_2}^z \right) \quad (13)$$

By integrating eqn (13) with the initial condition of $\xi = 0$ when $t = 0$, eqn (14) can be obtained.

$$\begin{aligned} \xi &= 1 - \left(1 - \frac{M(K_r^f)^2 K_r^f \left(P_{CH_4}^m - \frac{1}{K_{eq}} P_{CO}^n P_{H_2}^z \right)}{R_0\rho K_C} t \right)^3 \\ &= 1 - \left(1 - \frac{M(K_r^f)^2 \left(P_{CH_4}^m - \frac{1}{K_{eq}} P_{CO}^n P_{H_2}^z \right)}{R_0\rho K_C^0} \exp\left(-\frac{\Delta E_{app}}{RT}\right) t \right)^3 \end{aligned} \quad (14)$$

where $\Delta E_{app} = 2\Delta E_r - \Delta E_c$, K_r^0 and K_C^0 are temperature independent constants; is the apparent activation energy; ΔE_r and ΔE_c are the activation energies of the reduction reaction and methane cracking, respectively.

(2) When $\frac{K_C}{K_r^f} < 1$,

$$\frac{dr}{dt} = -\frac{M}{\rho} \frac{K_C}{K_r^f} p_r \quad (15)$$

The process of methane cracking is the rate limiting step and the radial growth rate of α would be positively correlated with the cracking rate. In this case, it makes sense to put K_C (the cracking rate coefficient of methane) in the numerator of the formula. Similarly, the relation between ξ and t can be expressed in eqn (16).

$$\xi = 1 - \left(1 - \frac{MK_C^0 \left(P_{CH_4}^m - \frac{1}{K_{eq}} P_{CO}^n P_{H_2}^z \right)}{R_0\rho} \exp\left(-\frac{\Delta E_C}{RT}\right) t \right)^3 \quad (16)$$

3.1.2. Cylindrical particles. The metal oxides with a cylindrical shape are shown in Fig. 2, in which x is the thickness of the reaction product, R_0 and r are the radii of the cylindrical particle and unreacted metal oxides, respectively. The reduction on both top and bottom surfaces is negligible when the height of the cylindrical particle is far more than its diameter. Hence, ξ can be calculated by

$$\xi = 1 - \left(\frac{r}{R_0} \right)^2 \quad (17)$$

By the differentiation of eqn (17) with respect to time t , eqn (18) is obtained as:

$$\frac{d\xi}{dt} = -\frac{2r}{R_0^2} \frac{dr}{dt} \quad (18)$$

(1) When $\frac{K_C}{K_r^f} \geq 1$,

$$\frac{dr}{dt} = -\frac{M}{\rho} \frac{K_r^f}{K_C} p_r \quad (19)$$



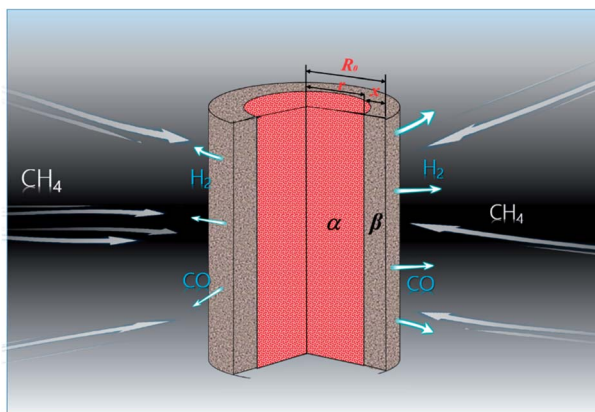


Fig. 2 Cylindrical particle during the reduction process with CH_4 .

Eqn (20) can be deduced by combining eqn (11), (17), (18) and (19) as follows:

$$\frac{d\xi}{dt} = \frac{2M(K_r^f)^2(1-\xi)^{1/2}}{R_0\rho K_C} \left(P_{\text{CH}_4}^m - \frac{1}{K_{\text{eq}}} P_{\text{CO}}^n P_{\text{H}_2}^z \right) \quad (20)$$

By substituting the initial conditions ($\xi = 0$, $t = 0$) in the equation mentioned above, eqn (21) can be derived after rearrangement

$$\xi = 1 - \left(1 - \frac{M(K_r^f)^2 \left(P_{\text{CH}_4}^m - \frac{1}{K_{\text{eq}}} P_{\text{CO}}^n P_{\text{H}_2}^z \right)}{R_0\rho K_C^0} \exp\left(-\frac{\Delta E_{\text{app}}}{RT}\right) t \right)^2 \quad (21)$$

where $\Delta E_{\text{app}} = 2\Delta E_r - \Delta E_c$.

(2) When $\frac{K_C}{K_r^f} < 1$,

$$\frac{dr}{dt} = -\frac{M}{\rho} \frac{K_C}{K_r^f} \nu_r \quad (22)$$

Similarly, the relation between ξ and t can be derived as shown in eqn (23).

$$\xi = 1 - \left(1 - \frac{MK_C^0 \left(P_{\text{CH}_4}^m - \frac{1}{K_{\text{eq}}} P_{\text{CO}}^n P_{\text{H}_2}^z \right)}{R_0\rho} \exp\left(-\frac{\Delta E_C}{RT}\right) t \right)^2 \quad (23)$$

3.1.3. Lamellar particles. The metal oxides with lamellar-shaped particles are shown in Fig. 3, where H_0 is the half-height of the lamellar particle. ξ can be calculated by:

$$\xi = \frac{x}{H_0} \quad (24)$$

It can be derived by differentiating eqn (24) with respect to time t .

$$\frac{d\xi}{dt} = \frac{1}{H_0} \frac{dx}{dt} \quad (25)$$

(1) When $\frac{K_C}{K_r^f} \geq 1$,

$$\frac{dx}{dt} = \frac{M}{\rho} \frac{K_r^f}{K_C} \nu_r \quad (26)$$

Then, combining eqn (11), (24), (25) and (26), eqn (27) can be obtained.

$$\frac{d\xi}{dt} = \frac{M(K_r^f)^2}{H_0\rho K_C} \left(P_{\text{CH}_4}^m - \frac{1}{K_{\text{eq}}} P_{\text{CO}}^n P_{\text{H}_2}^z \right) \quad (27)$$

By substituting the initial conditions ($\xi = 0$, $t = 0$) in the above equation, eqn (28) can be obtained.

$$\xi = \frac{M(K_r^f)^2 \left(P_{\text{CH}_4}^m - \frac{1}{K_{\text{eq}}} P_{\text{CO}}^n P_{\text{H}_2}^z \right)}{H_0\rho K_C^0} \exp\left(-\frac{\Delta E_{\text{app}}}{RT}\right) t \quad (28)$$

where $\Delta E_{\text{app}} = 2\Delta E_r - \Delta E_c$.

(2) When $\frac{K_C}{K_r^f} < 1$,

$$\frac{dx}{dt} = \frac{M}{\rho} \frac{K_C}{K_r^f} \nu_r \quad (29)$$

Similarly, the relation between the reaction and time can be expressed as shown in eqn (30).

$$\xi = \frac{MK_C^0 \left(P_{\text{CH}_4}^m - \frac{1}{K_{\text{eq}}} P_{\text{CO}}^n P_{\text{H}_2}^z \right)}{H_0\rho} \exp\left(-\frac{\Delta E_C}{RT}\right) t \quad (30)$$



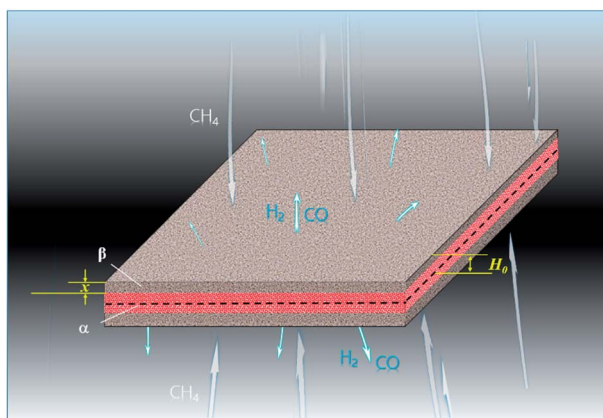


Fig. 3 Lamellar particle during the reduction process with CH₄.

Two parameters are defined to unify the form of formulae: the shape coefficient S_c , which is equal to 3, 2 and 1 when the raw materials are spherical-shaped, cylindrical-shaped and lamellar-shaped, respectively; the equivalent diameter d_0 represents the diameter of spherical particles, the diameter of cylindrical particles and the half height of lamellar particles, respectively. Under the isothermal condition, the derived formula can be rewritten as eqn (31).

When $\frac{K_C}{K_f} \geq 1$,

$$\xi = 1 - \left(1 - \frac{M(K_r^0)^2 \left(P_{CH_4}^m - \frac{1}{K_{eq}} P_{CO}^n P_{H_2}^z \right)}{d_0 \rho K_C^0} \exp\left(-\frac{\Delta E_{app}}{RT}\right) t \right)^{S_c} \quad (31)$$

When $\frac{K_C}{K_f} < 1$,

$$\xi = 1 - \left(1 - \frac{MK_C^0 \left(P_{CH_4}^m - \frac{1}{K_{eq}} P_{CO}^n P_{H_2}^z \right)}{d_0 \rho} \exp\left(-\frac{\Delta E_C}{RT}\right) t \right)^{S_c} \quad (32)$$

In industrial applications, most of the reduction processes of metal oxides are non-isothermal processes, such as the temperature of iron oxide in the shaft furnace increasing gradually due to furnace burden descent and the reverse flow of between gas and furnace burden. Furthermore, the coexistence of endothermic and exothermic reactions can also influence the temperature even in a single particle. Therefore, it is important to construct a model to predict the reduction process under non-isothermal conditions.

In this study, the rate of increase in the temperature η is assumed to be a constant value. The relationship between temperature (T) and time (t) can be calculated by,

$$T = T_0 + \eta t \quad (33)$$

in which T_0 represents the initial temperature. Here, we only discussed the simplest non-isothermal process (linear relationship). If the relationship between temperature and time is determined, any other non-isothermal formula can also be determined with the help of eqn (31) and (32). This non-isothermal model can be applied to the reactions involving powders and small pellets when they occur at moderate temperatures. Furthermore, reactions involving large pellets can also be predicted by this model when heat transfer is not the limiting factor of the reaction.

By substituting eqn (33) into eqn (31) and (32), the equations become.

When $\frac{K_C}{K_f} \geq 1$,

$$\xi = 1 - \left(1 - \frac{M(K_r^0)^2 \left(P_{CH_4}^m - \frac{1}{K_{eq}} P_{CO}^n P_{H_2}^z \right)}{d_0 \rho K_C^0} \times \exp\left(-\frac{\Delta E_{app}}{RT}\right) \frac{T - T_0}{\eta} \right)^{S_c} \quad (34)$$

When $\frac{K_C}{K_f} < 1$,

$$\xi = 1 - \left(1 - \frac{MK_C^0 \left(P_{CH_4}^m - \frac{1}{K_{eq}} P_{CO}^n P_{H_2}^z \right)}{d_0 \rho} \exp\left(-\frac{\Delta E_C}{RT}\right) \frac{T - T_0}{\eta} \right)^{S_c} \quad (35)$$

3.2. Diffusion in the product layer

3.2.1. Spherical particles. When the gas diffusion in the product layer is the control link of the whole reaction, the growth rate of α in the radial direction, which is in direct proportion to the CH₄ flux, $J_{CH_4}^\beta$, is given as:

$$\frac{dr}{dt} = -\frac{M}{\rho} J_{CH_4}^\beta \quad (36)$$



According to the Fick's first law of diffusion, CH_4 flux is expressed as

$$J_{\text{CH}_4}^{\beta} = -D_{\text{CH}_4}^{\beta} \left(\frac{P_{\text{CH}_4}(\alpha/\beta) - P_{\text{CH}_4}}{R_0 - r} \right) \quad (37)$$

where $D_{\text{CH}_4}^{\beta}$ represents the diffusion coefficient of CH_4 in the β phase;

Eqn (38) is given by integrating eqn (2), (3), (36) and (37).

$$\frac{d\xi}{dt} = -\frac{3MD_{\text{CH}_4}^{\beta}}{R_0^2\rho} (P_{\text{CH}_4}(\alpha/\beta) - P_{\text{CH}_4}) \frac{(1-\xi)^{2/3}}{1-(1-\xi)^{1/3}} \quad (38)$$

Eqn (38) is integrated and rearranged to obtain eqn (39).

$$\begin{aligned} \xi &= 1 - \left(1 - \sqrt{\frac{-2MD_{\text{CH}_4}^{\beta}}{R_0^2\rho} (P_{\text{CH}_4}(\alpha/\beta) - P_{\text{CH}_4}) t} \right)^3 \\ &= 1 - \left(1 - \sqrt{\frac{2MD_{\text{CH}_4}^{0\beta}}{R_0^2\rho} (P_{\text{CH}_4} - P_{\text{CH}_4}^{\text{eq}}) \exp\left(-\frac{\Delta E_{\text{app}}}{RT}\right) t} \right)^3 \end{aligned} \quad (39)$$

where $D_{\text{CH}_4}^{0\beta}$ represents a constant that's not affected by temperature, $P_{\text{CH}_4}^{\text{eq}}$ is the amount of CH_4 in equilibrium and ΔE_{app} is the apparent activation energy.

3.2.2. Cylindrical particles. Similarly, the growth rate of α in the radial direction should obey eqn (36). Combining eqn (17), (18), (36) and (37), eqn (40) can be obtained.

$$\frac{d\xi}{dt} = \frac{2MD_{\text{CH}_4}^{\beta}}{R_0^2\rho} (P_{\text{CH}_4}(\alpha/\beta) - P_{\text{CH}_4}) \frac{(1-\xi)^{1/2}}{1-(1-\xi)^{1/2}} \quad (40)$$

Eqn (40) can be integrated and rearranged as

$$\begin{aligned} \xi &= 1 - \left(1 - \sqrt{\frac{-2MD_{\text{CH}_4}^{\beta}}{R_0^2\rho} (P_{\text{CH}_4}(\alpha/\beta) - P_{\text{CH}_4}) t} \right)^2 \\ &= 1 - \left(1 - \sqrt{\frac{2MD_{\text{CH}_4}^{0\beta}}{R_0^2\rho} (P_{\text{CH}_4} - P_{\text{CH}_4}^{\text{eq}}) \exp\left(-\frac{\Delta E_{\text{app}}}{RT}\right) t} \right)^2 \end{aligned} \quad (41)$$

3.2.3. Lamellar particles. The growth rate of the thickness of the product layer is proportional to $J_{\text{CH}_4}^{\beta}$.

$$\frac{dx}{dt} = \frac{M}{\rho} J_{\text{CH}_4}^{\beta} \quad (42)$$

According to the Fick's first law of diffusion, eqn (43) can be given as

$$J_{\text{CH}_4}^{\beta} = -D_{\text{CH}_4}^{\beta} \left(\frac{P_{\text{CH}_4}(\alpha/\beta) - P_{\text{CH}_4}}{x} \right) \quad (43)$$

By combining eqn (24), (25), (42) and (43), eqn (44) can be obtained.

$$\frac{d\xi}{dt} = -\frac{MD_{\text{CH}_4}^{\beta}}{H_0^2\rho\xi} (P_{\text{CH}_4}(\alpha/\beta) - P_{\text{CH}_4}) \quad (44)$$

Integrating and rearranging eqn (44) yields eqn (45).

$$\begin{aligned} \xi &= \sqrt{\frac{-2MD_{\text{CH}_4}^{\beta}}{H_0^2\rho} (P_{\text{CH}_4}(\alpha/\beta) - P_{\text{CH}_4}) t} \\ &= \sqrt{\frac{2MD_{\text{CH}_4}^{0\beta}}{H_0^2\rho} (P_{\text{CH}_4} - P_{\text{CH}_4}^{\text{eq}}) \exp\left(-\frac{\Delta E_{\text{app}}}{RT}\right) t} \end{aligned} \quad (45)$$

If the shape coefficient S_c and the equivalent diameter d_0 are introduced into the model, the kinetic equation can be rewritten as

$$\xi = 1 - \left(1 - \sqrt{\frac{2MD_{\text{CH}_4}^{0\beta}}{d_0^2\rho} (P_{\text{CH}_4} - P_{\text{CH}_4}^{\text{eq}}) \exp\left(-\frac{\Delta E_{\text{app}}}{RT}\right) t} \right)^{S_c} \quad (46)$$

For spherical-shaped particles, S_c is equal to 3, 2 and 1 when the raw materials are spherical-shaped, cylindrical-shaped and lamellar-shaped, respectively; d_0 represents the diameter of spherical particles, the diameter of cylindrical particles and the half-height of lamellar particles, respectively.

For non-isothermal reduction,

$$\xi = 1 - \left(1 - \sqrt{\frac{2MD_{\text{CH}_4}^{0\beta}}{d_0^2\rho} (P_{\text{CH}_4} - P_{\text{CH}_4}^{\text{eq}}) \exp\left(-\frac{\Delta E_{\text{app}}}{RT}\right) \frac{T - T_0}{\eta}} \right)^{S_c} \quad (47)$$

4. The theoretical analysis of dual reactions

Developing a model to describe the reactions of dual interfaces or multiple interfaces is imperative and necessary. For example, when multiple valencies exist simultaneously in the reduction of metal oxides by CH_4 , the reaction process described by one interface model is inaccurate and even incorrect. When two reactions occur synchronously, the reactant in the second reaction is the product obtained in the first reaction. Two cases can be given:

(1) When the rate of product formation of the first reaction is much larger than its consumption in the second reaction, there will be two coexisting reaction interfaces in the particle. ξ can be calculated as

$$\xi = \lambda_1 \xi_1 + \lambda_2 \xi_2 \quad (48)$$

where λ_1 and λ_2 are the coefficients determined by the loss of oxygen in each single reaction. According to $\lambda = \frac{\text{oxygen loss of single reaction}}{\text{total oxygen loss}}$, the value of λ can be obtained and the order of reaction is represented by 1 or 2. The overall reduction extent is calculated by adding ξ_1 and ξ_2 with corresponding weights; ξ_1 and ξ_2 represent the reduction extent in each single reaction.

(2) There is only one interface when the rate of product formation of the first reaction is much less than its



consumption in the second reaction, and the kinetic model, in this case, should be exactly the same as that of a single reaction.

For simplicity, the condition $\frac{K_C}{K_T} \geq 1$ is taken as an example to deduce the formulae.

4.1. Interface chemical reaction

When a chemical reaction is a controlling step in the dual reactions, integrating eqn (31) and (48) yield eqn (49) and (50).

$$\xi = 1 - \lambda_1 \left(1 - \frac{M_1 (K_{r1}^0)^2 \left(P_{CH_4}^m - \frac{1}{K_{eq1}} P_{CO}^n P_{H_2}^z \right)}{d_0 \rho_1 K_{C1}^0} \exp \left(- \frac{\Delta E_{app1}}{RT} \right) t \right)^{S_c} - \lambda_2 \left(1 - \frac{M_2 (K_{r2}^0)^2 \left(P_{CH_4}^m - \frac{1}{K_{eq2}} P_{CO}^n P_{H_2}^z \right)}{d_0 \rho_2 K_{C2}^0} \exp \left(- \frac{\Delta E_{app2}}{RT} \right) t \right)^{S_c} \quad (49)$$

Under the non-isothermal condition,

$$\xi = 1 - \lambda_1 \left(1 - \frac{M_1 (K_{r1}^0)^2 \left(P_{CH_4}^m - \frac{1}{K_{eq1}} P_{CO}^n P_{H_2}^z \right)}{d_0 \rho_1 K_{C1}^0} \exp \left(- \frac{\Delta E_{app1}}{RT} \right) \frac{T - T_0}{\eta} \right)^{S_c} - \lambda_2 \left(1 - \frac{M_2 (K_{r2}^0)^2 \left(P_{CH_4}^m - \frac{1}{K_{eq2}} P_{CO}^n P_{H_2}^z \right)}{d_0 \rho_2 K_{C2}^0} \exp \left(- \frac{\Delta E_{app2}}{RT} \right) \frac{T - T_0}{\eta} \right)^{S_c} \quad (50)$$

4.2. Mixed control

When one step of chemical reaction controls a reaction and CH_4 diffusion in the product layer controls another reaction, the theoretical analysis of ξ can be done by integrating eqn (31), (46) and (48).

$$\xi = 1 - \lambda_1 \left(1 - \frac{M_1 (K_{r1}^0)^2 \left(P_{CH_4}^m - \frac{1}{K_{eq1}} P_{CO}^n P_{H_2}^z \right)}{d_0 \rho_1 K_{C1}^0} \exp \left(- \frac{\Delta E_{app1}}{RT} \right) t \right)^{S_c} - \lambda_2 \left(1 - \sqrt{\frac{2 M_2 D_{CH_4}^{0\beta}}{d_0^2 \rho_2}} \left(P_{CH_4} - P_{CH_4}^{eq} \right) \exp \left(- \frac{\Delta E_{app2}}{RT} \right) t \right)^{S_c} \quad (51)$$

For the non-isothermal reduction reaction,

$$\xi = 1 - \lambda_1 \left(1 - \frac{M_1 (K_{r1}^0)^2 \left(P_{CH_4}^m - \frac{1}{K_{eq1}} P_{CO}^n P_{H_2}^z \right)}{d_0 \rho_1 K_{C1}^0} \exp \left(- \frac{\Delta E_{app1}}{RT} \right) \frac{T - T_0}{\eta} \right)^{S_c} - \lambda_2 \left(1 - \sqrt{\frac{2 M_2 D_{CH_4}^{0\beta}}{d_0^2 \rho_2}} \left(P_{CH_4} - P_{CH_4}^{eq} \right) \exp \left(- \frac{\Delta E_{app2}}{RT} \right) \frac{T - T_0}{\eta} \right)^{S_c} \quad (52)$$



4.3. Diffusion in the product layer

If diffusion controls all reactions, it gives

$$\xi = 1 - \lambda_1 \left(1 - \sqrt{\frac{2M_1 D_{\text{CH}_4}^{0\beta}}{d_0^2 \rho_1} (P_{\text{CH}_4} - P_{\text{CH}_4}^{\text{eq1}}) \exp\left(-\frac{\Delta E_{\text{app1}}}{RT}\right) t} \right)^{S_c} - \lambda_2 \left(1 - \sqrt{\frac{2M_2 D_{\text{CH}_4}^{0\beta}}{d_0^2 \rho_2} (P_{\text{CH}_4} - P_{\text{CH}_4}^{\text{eq2}}) \exp\left(-\frac{\Delta E_{\text{app2}}}{RT}\right) t} \right)^{S_c} \quad (53)$$

Likewise, under the non-isothermal condition, the equation can be obtained as

$$\xi = 1 - \lambda_1 \left(1 - \sqrt{\frac{2M_1 D_{\text{CH}_4}^{0\beta}}{d_0^2 \rho_1} (P_{\text{CH}_4} - P_{\text{CH}_4}^{\text{eq1}}) \exp\left(-\frac{\Delta E_{\text{app1}}}{RT}\right) \frac{T - T_0}{\eta}} \right)^{S_c} - \lambda_2 \left(1 - \sqrt{\frac{2M_2 D_{\text{CH}_4}^{0\beta}}{d_0^2 \rho_2} (P_{\text{CH}_4} - P_{\text{CH}_4}^{\text{eq2}}) \exp\left(-\frac{\Delta E_{\text{app2}}}{RT}\right) \frac{T - T_0}{\eta}} \right)^{S_c} \quad (54)$$

5. Practical application of the mathematical model

To get a simpler model, heat transfer is not discussed in this work, only the processes of chemical reaction and mass transfer are considered. Thus, it can be applied to the reduction of powder and pellets with small size or in systems, such as large pellets, where heat transfer is not the controlling factor. Generally, the nucleation and growth mechanism play a dominant role when the temperature is lower and the reduction at high temperatures may obey solid-state ion diffusion because of sintering. Hence, the developed model will be appropriate when the reaction occurs at moderate temperatures.

5.1. Nickel oxide reduction predicted by the developed model

The application of the model will be very convenient because a lot of parameters are included in the derived model, such as T , R , and ρ . Rashidi *et al.*² have done experimental research on the methane reduction of nickel oxide based on thermogravimetry with a methane concentration of 10–50% (with a gas flow rate of 200 mL L⁻¹). A cylindrical sample with a diameter of 7 mm and 0.6 mm thickness was used in the experiment. Because the diameter is much larger than the thickness, the sample can be considered as lamellar-shaped. In general, the methane cracking rate can be very fast when the temperature is over 1000 °C. Therefore, eqn (30) and (45) can work in this case and a better prediction is found by using eqn (30). When methane

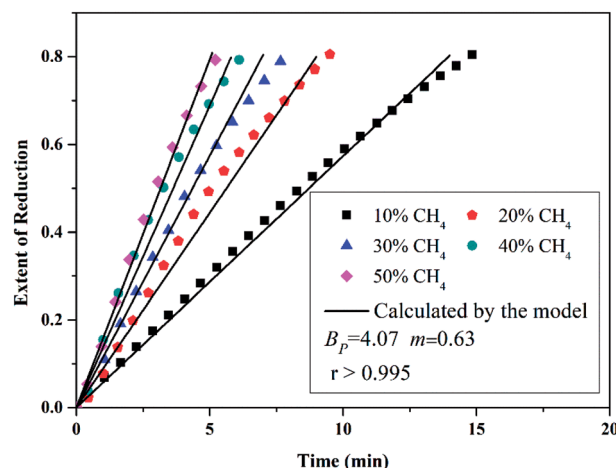


Fig. 4 Conversion extent vs. time curves for different CH₄ concentrations at 660 °C (933 K).

concentration as a single factor is studied, a parameter,

$$B_p = \frac{R_0 \rho (K_r^0)^2}{MK_C^0 \exp\left(-\frac{\Delta E_{\text{app}}}{RT}\right)},$$

is defined as convenient to describe experimental data. B_p is a function of R_0 , ρ , and M and it is affected by the substance. Once the metal oxides are fixed, B_p becomes constant. The formula can be rewritten as,

$$\xi = \frac{\left(P_{\text{CH}_4}^m - \frac{1}{K_{\text{eq}}} P_{\text{CO}}^n P_{\text{H}_2}^z\right)}{B_p} t \quad (55)$$

If the gas flow rate and K_{eq} are large enough, then the formula becomes,

$$\xi = \frac{P_{\text{CH}_4}^m}{B_p} t \quad (56)$$

Then, Fig. 4 shows the predicted curves calculated by eqn (56), in which solid lines are the values calculated by the model and the dotted lines represent experimental data. It should be noted that a thick layer of nickel will be formed around the particle (at the final stage of the reaction) and the diffusion and movement of the gases will be hindered. Therefore, the data for reduction extents less than 0.8 are considered during the fitting. By comparing the curves, it can be observed that the developed model can describe the reduction very well and the obtained kinetic equation is given as eqn (57).

$$\xi = \frac{P_{\text{CH}_4}^{0.63}}{4.07} t \quad (57)$$

5.2. Titania reduction predicted by the developed model

The reduction of titania by CH₄-containing gas (5 vol pct CH₄–75 vol pct H₂) under the isothermal condition was investigated by Zhang *et al.*⁵ In the experiment, titania powders (in the range of 212 to 355 μm or 355 to 500 μm) were used as the raw



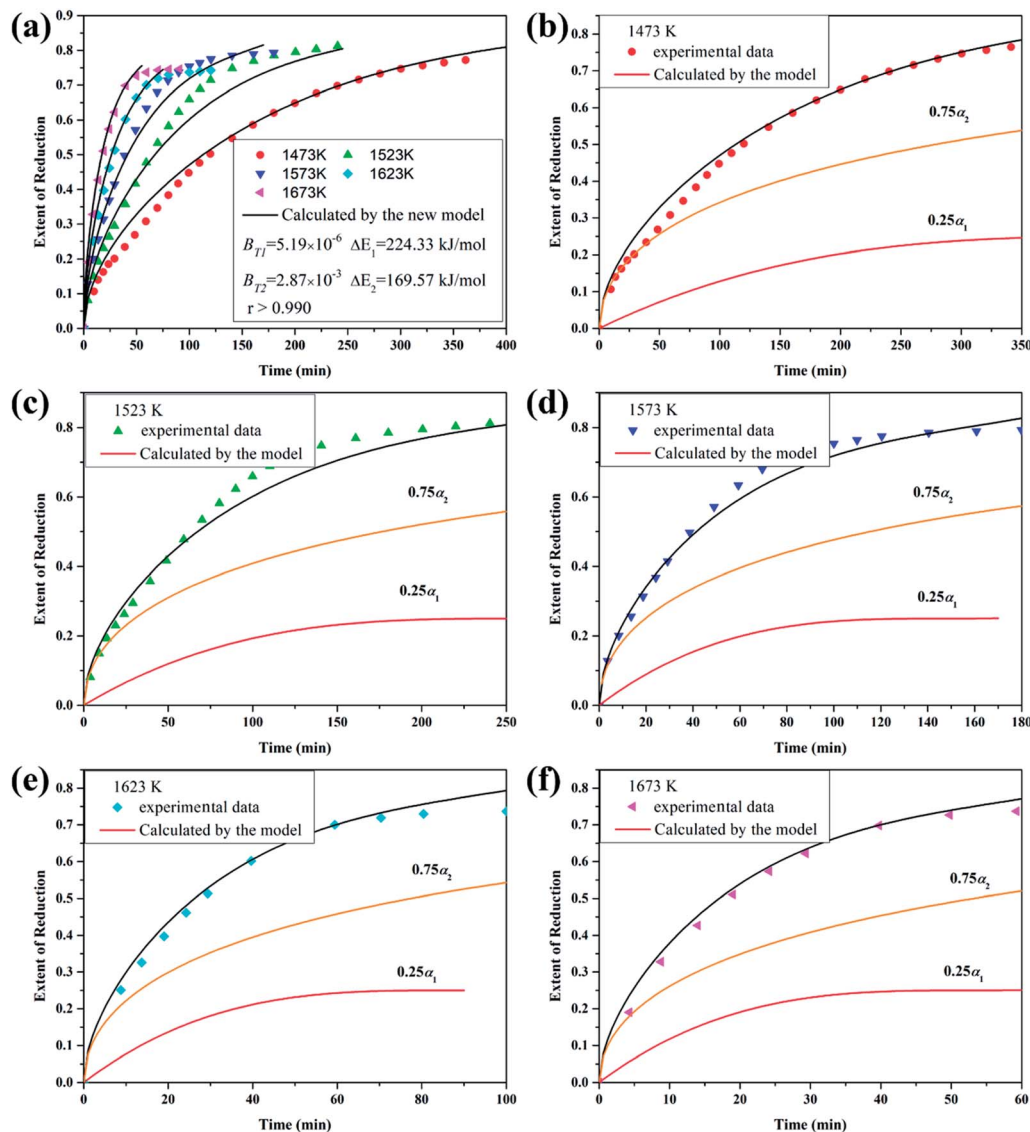
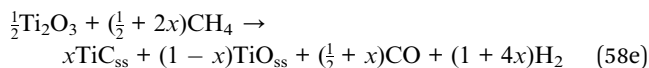
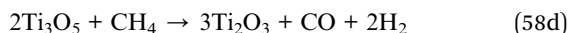
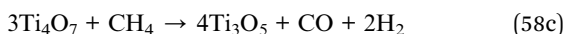
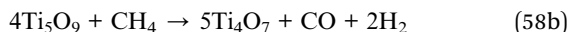


Fig. 5 (a) Theoretical and experimental curves: titania reduced by 5 vol pct CH₄-75 vol pct H₂ gas mixture at different temperatures; (b)-(f) kinetic curves (reduction extent vs. time) of the step reactions and overall reaction corresponding to (a).

materials. Fig. 5(a) shows that the dotted lines represent experimental data. It can be found that the reduction path is $\text{TiO}_2 \rightarrow \text{Ti}_5\text{O}_9 \rightarrow \text{Ti}_4\text{O}_7 \rightarrow \text{Ti}_3\text{O}_5 \rightarrow \text{Ti}_2\text{O}_3 \rightarrow (\text{TiO-TiC})_{\text{ss}}^{1,5}$ and the complex process can be described by reactions (58a-e).



Therefore, the reaction from TiO_2 to Ti_2O_3 can be considered as the first reaction and Ti_2O_3 to $(\text{TiO-TiC})_{\text{ss}}$ can be regarded as the second reaction. The dual reactions have been verified to coexist in the reduction (TiC and Ti_2O_3 were confirmed to be formed in one sample⁵). Thus, the derived dual reaction model can be applied for the prediction of experimental data. λ_1 and λ_2 have values of 0.25 and 0.75, respectively. Different dual reaction formulae are employed to model the reduction kinetics and it is found that eqn (51) gives the best fitting results. To simplify the model, two parameters are defined,

$$B_{T1} = \frac{d_0 \rho_1 K_{C1}^0}{M_1 (K_{T1}^0)^2 \left(P_{\text{CH}_4}^m - \frac{1}{K_{\text{eq1}}} P_{\text{CO}}^n P_{\text{H}_2}^z \right)} \quad \text{and} \quad B_{T2} = \frac{d_0^2 \rho_2}{2M_2 D_{\text{CH}_4}^{0\beta} (P_{\text{CH}_4} - P_{\text{CH}_4}^{\text{eq}})}.$$

Compared with the values of

P_{CH_4} , other values (P_{CO} , P_{H_2} and $P_{\text{CH}_4}^{\text{eq}}$) are very small and can be neglected, so B_{T1} and B_{T2} can be regarded as constants. Then, eqn (51) can be rewritten as

$$\xi = 1 - \lambda_1 \left(1 - \frac{1}{B_{\text{T1}}} \exp \left(- \frac{\Delta E_{\text{app1}}}{RT} \right) t \right)^3 - \lambda_2 \left(1 - \sqrt{\frac{1}{B_{\text{T2}}}} \exp \left(- \frac{\Delta E_{\text{app2}}}{RT} \right) t \right)^3 \quad (59)$$

The results predicted by the dual reaction model are shown as solid lines in Fig. 5(a) and the obtained kinetic formula is shown as eqn (60).

$$\xi = 1 - 0.25 \left(1 - \frac{1}{5.19 \times 10^{-6}} \exp \left(- \frac{224.33 \text{ kJ mol}^{-1}}{RT} \right) t \right)^3 - 0.75 \left(1 - \sqrt{\frac{1}{2.87 \times 10^{-3}}} \exp \left(- \frac{169.57 \text{ kJ mol}^{-1}}{RT} \right) t \right)^3 \quad (60)$$

Furthermore, based on the new model, the reduction kinetics for TiO_2 to Ti_2O_3 and Ti_2O_3 to $(\text{TiO}-\text{TiC})_{\text{ss}}$ are obtained successfully as shown in Fig. 5(b)–(f).

5.3. Comparison of the application of different models

Table S1† shows the earlier reported kinetic models,^{11,31,32} which are compared with the new model developed in this work with practical applications. The reduction reaction of titania by 5 vol pct CH_4 –75 vol pct H_2 gas mixture at 1473 K (investigated by Zhang *et al.*⁵) was selected to be the example calculated by different models. Obviously, the results obtained by the newly developed model are better than those obtained by the previous models, as shown in Table S1 and Fig. S1.† This is because the newly developed model considers the dual reactions, which occur simultaneously and is more close to practice.

5.4. The limitation of this model

This work assumed that all metal oxides are nonporous and these particles have a uniform diameter that remains constant during reduction. The model, developed in this work, only considers the effect of mass transfer and chemical reaction but not heat transfer. Therefore, the reactions involving powders and small pellets can be applied to this mathematical model when the reaction occurs at moderate temperatures. Furthermore, reactions involving large pellets can also be predicted by this model when the heat transfer is not the limiting factor of the reaction. However, the mixed controlling mechanism (the resistance of both chemical reaction and diffusion is large) for a single reaction is not considered in this study, which will be done in our future work.

6. Conclusions

(1) In this work, the reduction mechanism between metal oxides and methane was discussed in detail and a new mathematical model, which was in the form of explicit functions of many parameters, for methane reduction was developed. Particularly, the comparison of methane cracking rate and reaction rate was involved in the model.

(2) Both the gas diffusion in the product layer and chemical reaction controlled mechanisms were considered during the derivation. The kinetic formulae for the reduction of three shapes (spherical, cylindrical and lamellar) of particles were deduced as well and by incorporating two parameters S_c and d_o , these formulae were unified to one formula.

(3) Whether it is an isothermal or non-isothermal condition, the model can be used to simulate the reduction between methane and metal oxides. Furthermore, it can describe the reduction of varied-valence metal oxides.

(4) The experimental data obtained from literature could be feasibly predicted using the new model of methane reduction.

Conflicts of interest

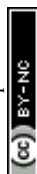
There are no conflicts to declare.

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