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Understanding the crystal structure-dependent electrochemical capacitance of spinel and rock-salt Ni–Co oxides *via* density function theory calculations†

Xuan Sun, Jinfeng Sun, Lingzhi Guo, Linrui Hou* and Changzhou Yuan*

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The spinel $NiCo_2O_4$ and rock-salt $NiCoO_2$ have been well established as attractive electrodes for supercapacitors. However, what is the intrinsic role of the congenital aspect, *i.e.*, crystal structure and the surface and/or near-surface controlled electrochemical redox behaviors, if the acquired features (*i.e.*, morphology, specific surface area, pore structure, and so on) are wholly ignored? Herein, we purposefully elucidated the underlying influences of unique crystal structures of $NiCo_2O_4$ and $NiCoO_2$ on their pseudocapacitance from mechanism analysis through the density function theory based first-principles calculations, along with the experimental validation. Systematic theoretical calculation and analysis revealed that more charge carriers near the Fermi-level, stronger affinity with OH^- in the electrolyte, easier deprotonation process, and the site-enriched characteristic for low-index surfaces of $NiCoO_2$ enable its faster redox reaction kinetics and greater charge transfer, when compared to the spinel $NiCo_2O_4$. The in-depth understanding of crystal structure–property relationship here will guide rational optimization and selection of appropriate electrodes for advanced supercapacitors.

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1 Introduction

To date, supercapacitors have shown great potential for electrochemical energy storage due to their higher power density and longer lifespan than rechargeable batteries.[1] From the intrinsic energy storage mechanism point of view, supercapacitors are typically categorized into electric double-layer capacitors (EDLCs) depending upon fast ad/desorption of ions on the electrode surface, and pseudocapacitors originating from rapid faradaic redox reactions on/near the surface of the electrodes.[2,3] However, the low specific capacitances (SCs) of EDLCs, to a great extent, limit their utilization in high energy applications. As a result, enormous amounts of research have been greatly stimulated to develop high-performance electrode materials with pseudocapacitive characteristics.[4,5]

Binary transition metal oxides (BTMOs) have been intensively investigated as active materials for supercapacitors due to their easy availability, environmental friendliness, and multiple oxidation states for efficient charge storage.[6,7] Especially, spinel $NiCo_2O_4$ and rock-salt $NiCoO_2$, stand out from others, owing to their more than two orders of magnitude higher electrical conductivity than their parent cobalt oxides.[8,9] In general, inherent electronic conductivity and ionic adsorption capability of abundant active sites are key points influencing electrochemical performance of pseudocapacitive electrodes, except for the acquired aspects including specific surface area (SSA) and pore structure.[10,11] Although extensive investigations into structural and/or compositional designs towards enhancing electrochemical properties of $NiCo_2O_4$ and/or $NiCoO_2$ electrodes have already been made,[6,12–17] no any reports can be retrieved to figure out which one possesses even better electrochemical activities. In fact, it is yet difficult to draw a definitive conclusion according to the previous contributions about the two for supercapacitors, since all the reported electrochemical properties are based on their different acquired characteristics (*i.e.*, morphology/SSA/pore structure).[6,9,12–17] In this regard, a thought-provoking question comes to the fore: what is the role of the congenital aspect, *i.e.*, crystal structures of spinel $NiCo_2O_4$ and rock-salt $NiCoO_2$, in their surface and/or near-surface controlled electrochemical redox behaviors if the aforementioned acquired features are wholly ignored? There is no doubt that the in-depth understanding of crystal structure–property relationship will render instructive insights into future optimal choice of appropriate electrodes for advanced supercapacitors.

With comprehensive considerations above in mind, herein, we elucidated the underlying influences of unique crystal structures of $NiCo_2O_4$ and $NiCoO_2$ on their pseudocapacitances from mechanism analysis through the density function theory

School of Materials Science & Engineering, University of Jinan, Jinan, 250022, P. R. China. E-mail: mse_houlr@ujn.edu.cn; mse_yuancz@ujn.edu.cn; ayuancz@163.com

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(DFT) based first-principles calculations, along with the experimental validation. Systematic theoretical calculation and analysis revealed that more charge carriers near the Fermi-level, stronger affinity with hydroxyls in the electrolyte, easier deprotonation process, and the site-enriched characteristic for low-index surfaces of $NiCoO_2$ enable its faster redox reaction kinetics and more charge transfer, when compared to the spinel $NiCo_2O_4$. As expected, the $NiCoO_2$ exhibited enhanced electrochemical behaviors especially at high rates.

2 Computational methods

Spin-polarized density functional theory calculations were carried out based on the Cambridge Sequential Total Energy Package known as CASTEP,[18] using the projector augmented wave (PAW) potentials.[19] we used the Perdew–Burke–Ernzerhof with Hubbard U (PBE + U) form within the generalized gradient approximation (GGA) for exchange–correlation energy.[19,20] Notably, different values between $U = 0$ and 8 have been used in the Ni/Co compounds in the literature. Herein, we have chosen U values of 5.5 for Ni ions, and 4.5/3.3 eV for Co ions at tetrahedral/octahedral sites, respectively, to obtain the correct electronic and magnetic state of bulk $NiCoO_2$ and $NiCo_2O_4$. The number of plane wave was collected by a 500 eV. The Brillouin zones were represented by a Monkhorst–Pack mesh of $7 \times 7 \times 7$ k-points for crystals and $3 \times 3 \times 1$ k-points for studied surface slabs.[21] The stoichiometric surfaces were modeled by a (1×1) unit cell for $NiCo_2O_4$ and (2×2) for $NiCoO_2$, which were separated in the z-direction by a 15 Å vacuum. The convergence criterion of Hellmann–Feynman forces and total energy were 0.015 eV $Å^{-1}$ and 10^{-5} eV, respectively. For accurate results, we compared band structures between $NiCo_2O_4$ and $NiCoO_2$ using HSE06 hybrid functional.[22] To model the magnetic state of $NiCo_2O_4$ and $NiCoO_2$, we used antiparallel spins for Co and Ni atoms in $NiCoO_2$ due to its high Néel temperature according to Du *et al.*[23] In addition, the Co atoms at tetrahedral sites have antiparallel state with Ni atoms for $NiCo_2O_4$, where the Co ions at octahedral sites have no spin polarization due to the fully occupied t_{2g} states.[24]

The surface energy (γ), was computed as $\gamma = (E_{slab} - nE_{bulk})/2A$, where E_{slab}, E_{bulk}, n and A were the energy of the slab supercell, the bulk energy per unit cell, the number of bulk cells contained in the surface slab, and the surface area of each side for the slab, respectively. For the surface with different hydroxyl coverage, the average adsorption energy could be defined as $E_{ads} = [E_{(n\mathrm{OH/slab})} - E_{(slab)} - nE_{(OH)}]/n$.[25,26] The deprotonation energy was calculated by $E_{de} = [E_{(O/slab)} - E_{(n\mathrm{OH/slab})} + 0.5nE_{(H_2)}]/n$, where $E_{(n\mathrm{OH/slab})}$, $E_{(O/slab)}$, $E_{(slab)}$, $E_{(OH)}/E_{(H_2)}$ and n were the energies of the adsorbed systems, the energy of the deprotonation system, the energy of the clean surface, the energies for an isolated hydroxyl/H_2 molecules, and the number of adsorbed hydroxyls or desorbed protons, respectively. To have a profound understanding about charge transfer, the Bader charge method, where an amount of each charge was considered in the region separated by the minimum of electronic charge density between each atom, was used.[27,28]

3 Results and discussion

3.1 Crystal structures

As schematically shown in Fig. 1a and b, the crystal structures of $NiCo_2O_4$ and $NiCoO_2$, typically contain 56 and 8 atoms in the unit cells, respectively. According to the site performance theory, the crystal structure of $NiCo_2O_4$ and $NiCoO_2$ can be constructed by substituting cobalt atoms located at octahedral sites in spinel Co_3O_4 and rock-salt CoO unit cells with elemental nickel, respectively.[7,24,29] Evidently, the spinel $NiCo_2O_4$ contains two kinds of Co atoms located at octahedrally coordinated sites (Co^o) and tetrahedrally coordinated sites (Co^t), respectively, coupled with alternating layers of –Co^o–O– and –Ni–O– along the z direction,[24] as described in Fig. 1a. The rock-salt $NiCoO_2$, as shown in Fig. 1b, only possesses the octahedral sites both for Co and Ni atoms. The optimized lattice constant values of $NiCo_2O_4$ and $NiCoO_2$ (Table S1, ESI†) are 8.22 and 4.26 Å, just with ~1.4% and ~0.5% error, respectively, when compared to their standard files. In addition, the computed magnetic moments of Co^t and Ni are 2.77 and −1.43 μ_B in spinel $NiCo_2O_4$, respectively, and 2.74 and −1.80 μ_B for Co and Ni in rock-salt $NiCoO_2$, which are in good agreement with the reported calculational and/or experimental values.[24,30]

3.2 Electronic properties

According to the pseudocapacitive system, the electrolyte ions diffuse from the electrolyte to the electrode/electrolyte interface under the effect of external circuit in the charging process, and then, the redox reaction occurs. Thus, the states near the Fermi-level, especially the contribution of d-orbital of metal ions, are considered to be related with the charge storage ability.[31,32] As discerned from the total band structures of $NiCo_2O_4$ (Fig. 1c) and $NiCoO_2$ (Fig. 1d), the occupied and unoccupied electronic states are visualized near the Fermi-level, favoring for their rapid electronic transport. To gain more insights into specific electronic structure around the Fermi-level, the projected band structures of d-orbital from the active Ni/Co atoms were calculated. Clearly, the bands near the Fermi-level are mainly contributed by the d-orbital of Co atoms, in which the Ni atoms contribute little. In addition, for spinel $NiCo_2O_4$, two kinds of atoms, *i.e.*, Co^o and Co^t both contribute the bands near the Fermi-level (Fig. S1, ESI†), where the contribution of Co^t can be wholly negligible. One thing worth noting is that the contributions of d-orbital near the Fermi-level are greatly conducive to the enhanced electron mobility for rapid redox process of pseudocapacitive electrodes.[33,34] On account of more contributions from d-orbital of metal atoms to the bands near the Fermi-level of rock-salt $NiCoO_2$, it can be rationally anticipated that the $NiCoO_2$ will exhibit higher electroactivities than spinel $NiCo_2O_4$.[35]

3.3 Charge-storage mechanism analysis

As for the surface modeling, we here consider three low-index, (100), (110) and (111), planes of $NiCo_2O_4$ and $NiCoO_2$ with the largest interplanar spacing, as they are predicted to be the most stable in the cubic crystals.[36,37] Among the three surfaces, the

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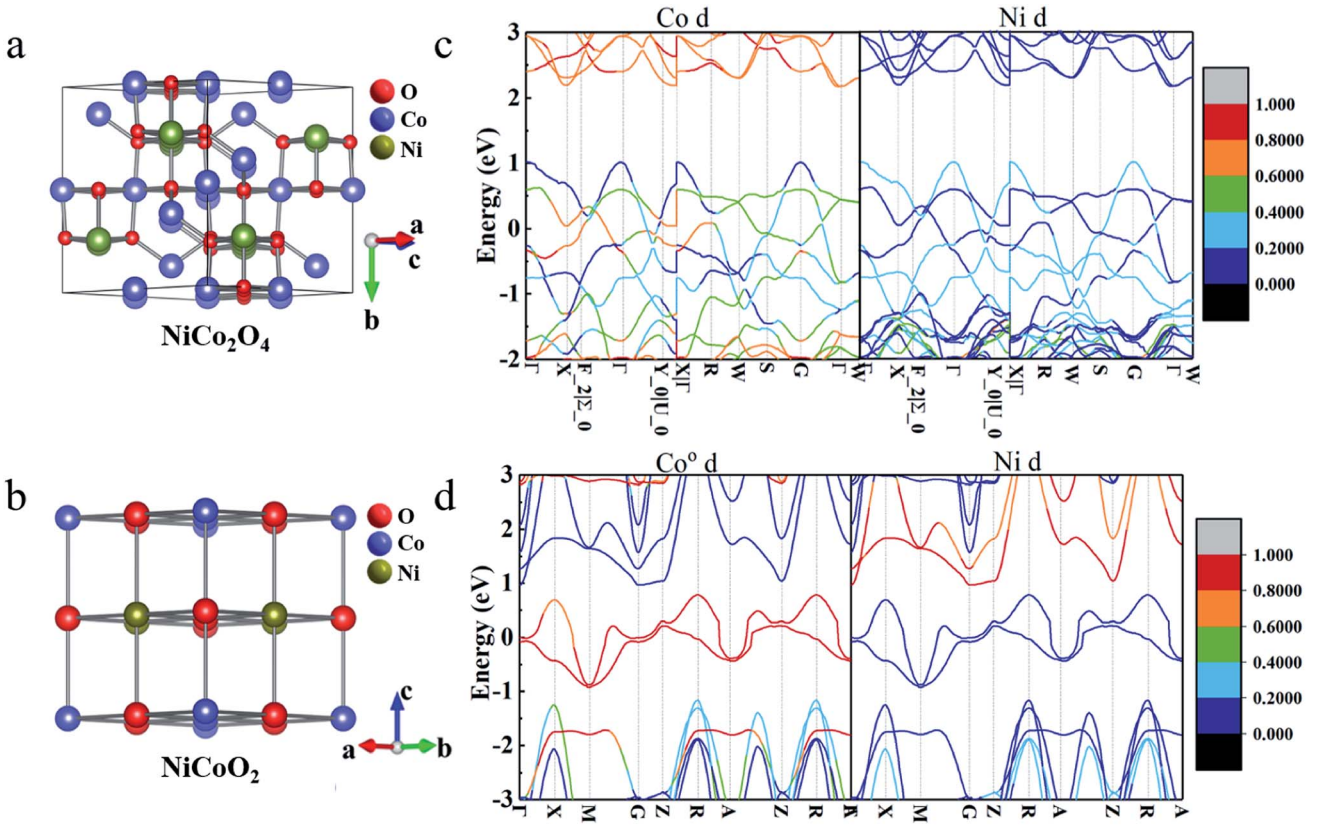

Fig. 1 Optimized crystallographic constructions of (a) $NiCo_2O_4$ and (b) $NiCoO_2$. Red, blue and green represent the oxygen, cobalt and nickel atoms. The contributions of d-orbital of all Co/Ni atoms to the band structure in (c) $NiCo_2O_4$ and (d) $NiCoO_2$. The energy zero is set at the Fermi-level.

(111) plane, both for spinel and rock salt structure, is not truly stable surfaces, which can be expected to undergo surface reconstruction.[36,38] Therefore, we finally study the (100) and (110) surfaces. The surface energy values of (100)/(110) for $NiCo_2O_4$ and $NiCoO_2$ are 1.03/1.25 and 0.95/1.99 J m^{-2}, respectively. Notably, the (110) surface, both for $NiCo_2O_4$ and $NiCoO_2$, has two different stable non-polar surface configurations (Fig. 2). Accordingly, we choose the Oh(top) (Fig. 2a) of the $NiCo_2O_4$(110) for further calculation, since its total energy is 0.23 eV, lower than those of Th/Oh(top) (Fig. 2b) according to the DFT calculation. And the Co(top) (Fig. 2c), whose total energy is 0.66 eV, lower than that of Ni(top) (Fig. 2d), is chosen for the $NiCoO_2$(110). Corresponding surface lattice constants are summarized as well (Table S1, ESI†).

According to the established charge-storage mechanism of $NiCo_2O_4$ and $NiCoO_2$ in alkaline KOH solution, the pseudocapacitance always takes place as following two steps.[10,14,39,40] First, the redox reaction occurring on the exposed surface generally starts from the OH^- adsorption, thus leading to the formation of *OH species. And then, the reaction of adsorbed *OH with the OH^- in the solution partially forms the adsorbed *O species, and releases a H_2O molecule meanwhile, which is generally ascribed as a deprotonation process. Since the pseudocapacitance originates from the fast redox reactions through the chemisorption of electrolyte ions, the binding energy can act as one of the descriptors for evaluating electrochemical performance. After simple analysis, there are two kinds of stable adsorbed sites, that is, the top sites of Co and Ni atoms both in

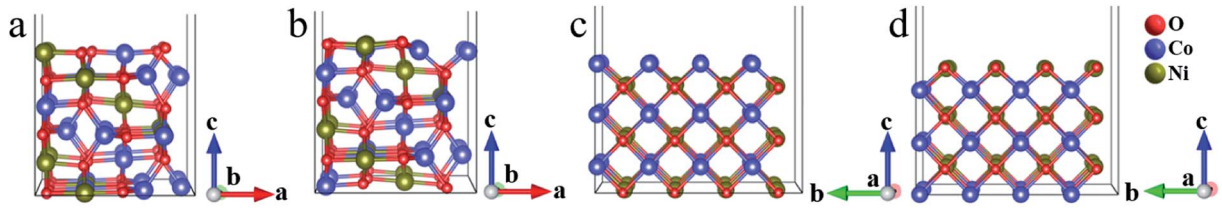

Fig. 2 Side views for the two kinds of (110) slabs of $NiCo_2O_4$: (a) Oh(top), (b) Th/Oh(top), and (110) slabs of $NiCoO_2$: (c) Co(top), (d) Ni(top). The Oh(top) contains only octahedral metal ions while Th/Oh(top) contains both tetrahedral and octahedral metal ions on the top surface. Co(top) and Ni(top) are Co atoms and Ni atoms on the top surface.

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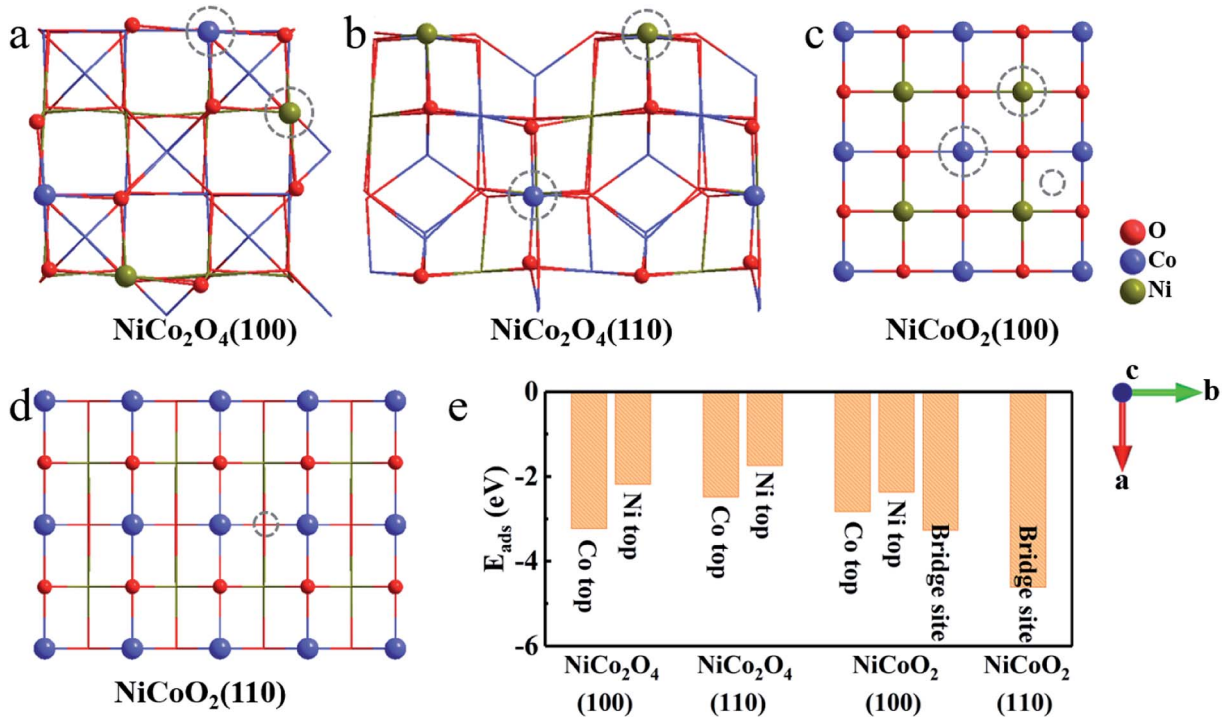

Fig. 3 Top view of the (a) $NiCo_2O_4(100)$; (b) $NiCo_2O_4(110)$; (c) $NiCoO_2(100)$ and (d) $NiCoO_2(110)$. The dotted circles represent the stable adsorbed sites for hydroxyl. (e) Corresponding energy analysis for hydroxyl adsorption at different sites.

(100) and (110) surfaces of $NiCo_2O_4$ (Fig. 3a and b). For the case of $NiCoO_2(100)$, the OH^- can be adsorbed stably both on the top sites of Ni/Co and bridge site between them (Fig. 3c). In contrast, the OH^- adsorbs only on the bridge site of two Co atoms in $NiCoO_2(110)$ (Fig. 3d). As collected in Fig. 3e, the binding energy between hydroxyls and active surfaces of $NiCo_2O_4/NiCoO_2$ is an exothermic and spontaneous process, due to their negative adsorption energy. The Co sites in $NiCo_2O_4$ display stronger absorption capacity for OH^- on both two surfaces, as indicated by even more negative E_{ads} for (100) (−3.23 eV) and (110) (−2.48 eV) than those of Ni sites (−2.18 and −1.74 eV) (Fig. S2a–d, ESI†). As regards the $NiCoO_2(100)$, the bridge site between Co and Ni (Fig. S2e, ESI†) is a more preferred site, which shows the E_{ads} of −3.26 eV for OH^-, lower than those at the top of single Co (−2.83 eV) and Ni (−2.37 eV) ions (Fig. S2f and g, ESI†). The $NiCoO_2(110)$ shows the most negative E_{ads} of −4.62 eV when adsorbed on the bridge site between two Co atoms (Fig. S2h, ESI†). Clearly, compared with those of $NiCo_2O_4$, the (100) and (110) surfaces of $NiCoO_2$ both exhibit more negative E_{ads} in the case of low coverage (Fig. 3e), which indeed favors for high faradaic reaction kinetics, thanks to the fast and stable OH^- adsorption on the two surfaces.

Then, we investigate how OH^- pattern varies when its coverage is changed. Fig. 4a presents the E_{ads} as function of hydroxyl coverage for the most stable adsorption sites of the $NiCo_2O_4$ and $NiCoO_2$. It's clear that the E_{ads} decreases as the hydroxyl coverage increases for both the (100) and (110) surfaces due to the gradual saturation of the surface. The E_{ads} values of the (100)/(110) surfaces for $NiCoO_2$ are larger than those of $NiCo_2O_4$ at all values of hydroxyls coverage. Corresponding relaxation configurations (Fig. S3a–d, ESI†) are summarized as well. According to the coordination configurations of the surface atoms, the Ni/Co atoms display octahedral coordination geometry when the OH^- coverage is full (Fig. 4b). Evidently, two hydroxyls are adsorbed on each Ni and Co sites in $NiCo_2O_4(110)$ due to two oxygen atoms loss in the coordination of octahedral Ni and Co.

Besides the binding ability between electrode surfaces and electrolyte ions, a superior electrode material fundamentally requires rich active sites and high electron-donating capability,[35] which are always related to the charge transfer originating from the hydroxyl adsorption. From the lattice constants aforementioned (Table S1, ESI†), the exposed surfaces of $NiCoO_2$ own higher concentration of active atoms (CAAs) of Ni and Co, which will bond directly to the hydroxyls in the reaction process. The CAAs are calculated as ~11.02 and ~7.57 atoms per nm for (100) and (110) surfaces of $NiCoO_2$, respectively, approximately twice of those for (100) (5.92 atoms per nm) and (110) (4.19 atoms per nm) of $NiCo_2O_4$ (Fig. S4a, ESI†). Therefore, the preponderant reaction surface of $NiCoO_2$ promotes a greater proportion of Ni and Co atoms involved in electrochemical reactions, which would provide high electron-donating capability. By performing the calculations on charge density difference (Fig. S4b–e, ESI†), we can deduce that most electrons tend to transfer from electrodes to hydroxyls when the Ni/Co TMOs are charged, indicating that the adsorption process is an oxidation process of surface metal atoms. We further explore the charge transfer (ΔQ_1) *versus* the hydroxyl coverage according

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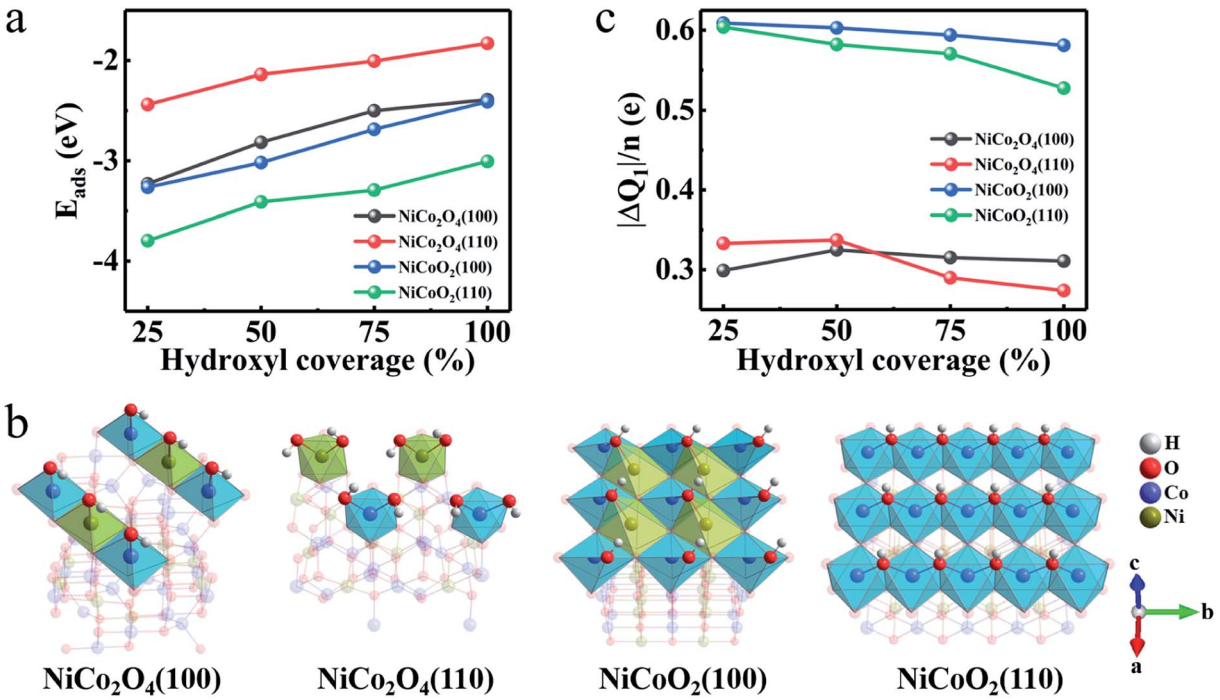

Fig. 4 (a) Hydroxyl adsorbed energy (E_{ads}) at different hydroxyl coverage. (b) Coordination configurations at full coverage of all studied surfaces. (c) Charge transfer (ΔQ_1) at different hydroxyl coverage.

to the Bader charge analysis (Fig. 4c). The average charge values ($\Delta Q_1/n$), that is the adsorbed OH^-, obtained on the $NiCoO_2$(100) and (110) surfaces are larger than those in the $NiCo_2O_4$(100) and $NiCo_2O_4$(110) at any hydroxyl coverage. It confirms a tendency that a better electron donation of surface Ni/Co atoms ensures their higher oxidation states. All analysis above therefore can deduce the exceptionally high pseudocapacitance of the rock-salt $NiCoO_2$ in alkaline KOH electrolyte.

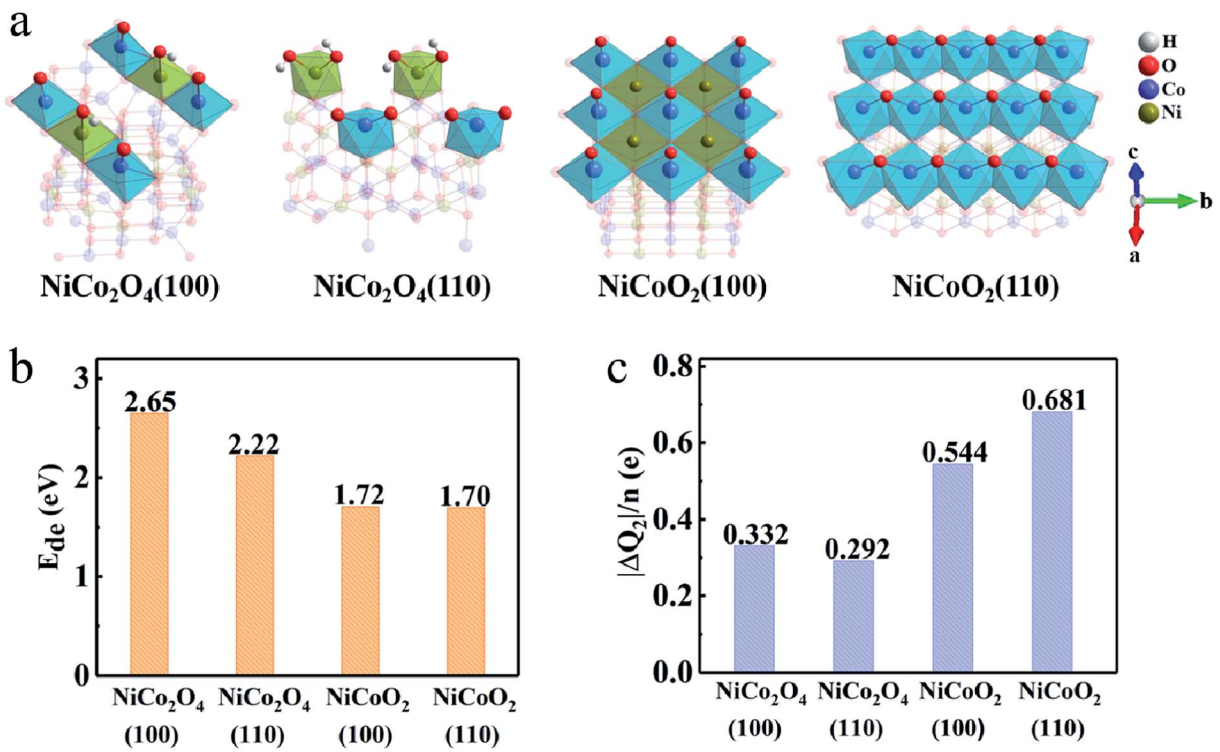

Fig. 5 (a) Surface configuration, (b) average deprotonation energy (E_{de}) and (c) charge transfer (ΔQ_2) after the complete deprotonation.

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The deprotonation process can be described as MOOH + OH^- ↔ MO_2 + H_2O (M is Co or Ni). However, the general observation is that the deprotonation process cannot occur at octahedral coordinate Ni ions, at the limited work potential window for supercapacitors, due to the higher transition potential than that of Co ions.[40,41] As shown in Fig. 5, the deprotonation, is a relatively unfavorable process compared to adsorption process due to the high deprotonation energy values (E_{de}), which means that the reaction involved in the second step needs even higher driving force. The DFT calculation also authenticates that the E_{de} values on (100) (1.72 eV) and (110) (1.70 eV) surfaces of $NiCoO_2$ are lower than those of $NiCo_2O_4$, *i.e.*, 2.65/2.22 eV for the (100)/(110) surfaces (Fig. 5b). Hence, the formed *OH has much more opportunities to desorb protons from the surfaces of $NiCoO_2$ than the $NiCo_2O_4$ to form *O. While, there are more electrons transfer from electrodes to adsorbed molecules due to the deprotonation process, which demonstrates that the surface/subsurface metal atoms are further oxidized. But for $NiCoO_2$(100), the number of charge transfer decreases after hydrogen desorption, which is related to the fact that the adsorption sites deviate from the bridge sites to the top of cobalt, decreasing the charge donation of nickel atoms (Fig. 5c).

The first-principles calculation results, as discussed above, corroborate more attractive innate merits of $NiCoO_2$ serving as the electrode material for supercapacitors as follows. First, the large electronic state contribution from Co atoms of $NiCoO_2$ near the Fermi-level provides more charge carriers for faradaic redox reactions, thus enhancing the redox activities. Second, the much easier OH^- adsorption and deprotonation processes will accelerate the surface reaction dynamics. The two aspects synergistically facilitate the rate behaviors of the $NiCoO_2$ electrode. Third, the surfaces of $NiCoO_2$ show high concentration of redox active sites and higher electron donation capability than $NiCo_2O_4$, realizing a large degree of oxidation after reaction, which is helpful to obtain large SCs. It is thanks to these appealing congenital advantages that the rock-salt $NiCoO_2$ can be reasonably anticipated with even better electrochemical performance than the spinel $NiCo_2O_4$.

3.4 Electrochemical properties

To experimentally support the above viewpoint, we purposefully chose our synthesized two Ni–Co binary oxides, *i.e.*, flower-shaped $NiCoO_2$ (ref. 14) and $NiCo_2O_4$ (Fig. S5 and S6, ESI†), and comparatively study their capacitive performance. The distinct redox peaks in cyclic voltammetry (CV) curves of both $NiCo_2O_4$ (Fig. 6a) and $NiCoO_2$ (ref. 14) verify their typical pseudocapacitance mainly, which typically results from fast and reversible redox processes of $Co^{2+}/Co^{3+}/Co^{4+}$ and Ni^{2+}/Ni^{3+} in

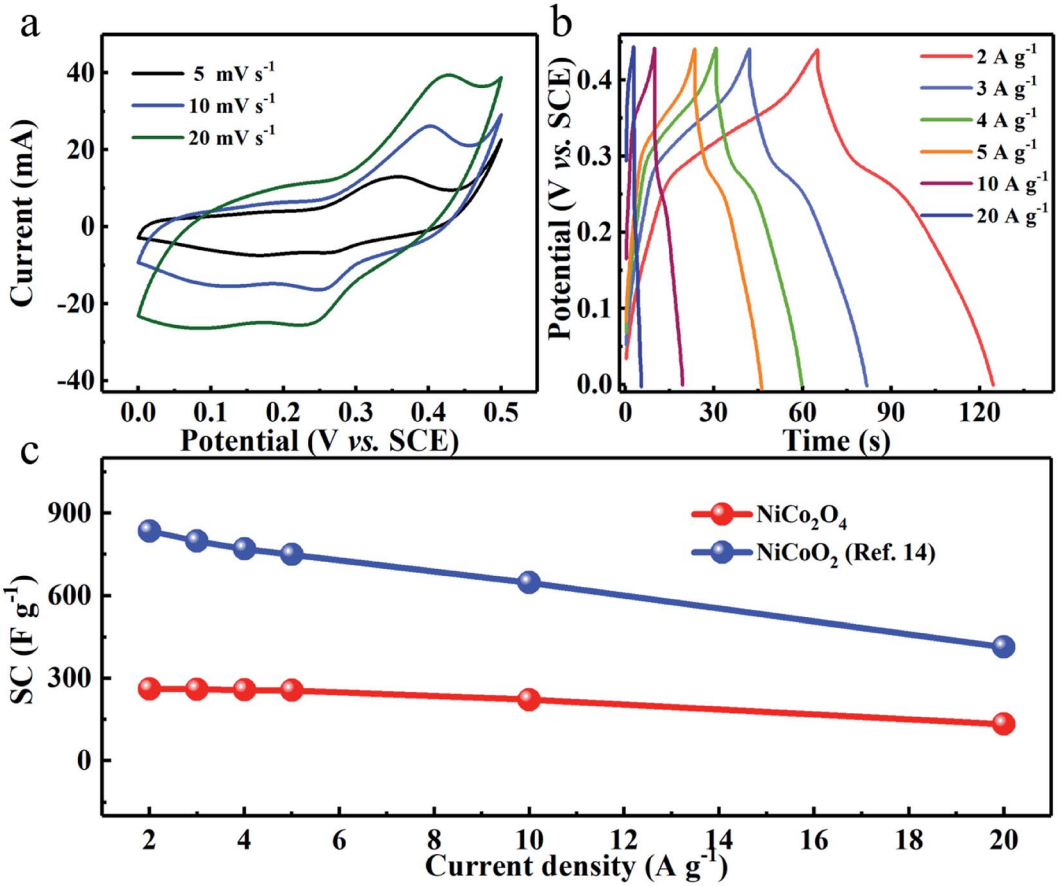

Fig. 6 (a) CV and (b) CP curves of the $NiCo_2O_4$ electrode. (c) Comparison of SCs between $NiCo_2O_4$ and $NiCoO_2$ electrodes at various current densities.

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2 M KOH aqueous solution.[9,14,15,42,43] As derived from the chronopotentiometry (CP) plots (Fig. 6b), the spinel $NiCo_2O_4$ electrode with a mass loading 5 mg cm^{-2} exhibits reversible SCs of ~262, ~260, ~258, ~256, ~222 and ~133 F g^{-1} at the current densities of 2, 3, 4, 5, 10 and 20 A g^{-1}, respectively. By sharp contrast, the rock-salt $NiCoO_2$ obtains ultrahigh SCs of ~836 and ~415 F g^{-1} at 2 and 20 A g^{-1}, respectively,[14] which are both higher than three times those of the $NiCo_2O_4$ electrode, as plotted in Fig. 6c.

4 Conclusions

In summary, we have comprehensively conducted the first-principle DFT calculations to shed light on the distinguished advantages of rock-salt $NiCoO_2$ from its crystal structure. The faster hydroxyls adsorption/deprotonation kinetics and higher electron donation capability enable the rock-salt $NiCoO_2$ as a more competitive electrode for supercapacitors when compared to its spinel counterpart of $NiCo_2O_4$. Furthermore, electrochemical evaluation experimentally authenticated the theoretical simulation and calculation. The contribution here will guide rational optimization and selection of pseudocapacitive electrodes for advanced supercapacitors.

Conflicts of interest

There are no conflicts to declare.

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References

1 J. Sun, C. Wu, X. Sun, H. Hu, C. Zhi, L. Hou and C. Yuan, *J. Mater. Chem. A*, 2017, **5**, 9443–9464.
2 J. Chen, Y. Han, X. Kong, X. Deng, H. J. Park, Y. Guo, S. Jin, Z. Qi, Z. Lee, Z. Qiao, R. S. Ruoff and H. Ji, *Angew. Chem., Int. Ed.*, 2016, **55**, 13822–13827.
3 L. Liu, H. P. Zhao and Y. Lei, *InfoMat*, 2019, **1**, 74–84.
4 H. S. Kim, J. B. Cook, H. Lin, J. S. Ko, S. H. Tolbert, V. Ozolins and B. Dunn, *Nat. Mater.*, 2017, **16**, 454–460.
5 R. Wang, M. Yao and Z. Niu, *InfoMat*, 2020, **2**, 113–125.
6 X. Li, L. Jiang, C. Zhou, J. Liu and H. Zeng, *NPG Asia Mater.*, 2015, **7**, e165.
7 S. Liu, D. Ni, H.-F. Li, K. N. Hui, C.-Y. Ouyang and S. C. Jun, *J. Mater. Chem. A*, 2018, **6**, 10674–10685.
8 C. Yuan, H. B. Wu, Y. Xie and X. Lou, *Angew. Chem., Int. Ed.*, 2014, **53**, 1488–1504.
9 B. B. Liu, J. G. Liu, T. T. Zhang, C. X. Xu and H. Liu, *J. Mater. Chem. A*, 2019, **7**, 16222–16230.
10 J. Li, Z. Liu, Q. Zhang, Y. Cheng, B. Zhao, S. Dai, H.-H. Wu, K. Zhang, D. Ding, Y. Wu, M. Liu and M.-S. Wang, *Nano Energy*, 2019, **57**, 22–33.
11 X. Li, L. Jiang, C. Zhou, J. Liu and H. Zeng, *NPG Asia Mater.*, 2015, **7**, e165.
12 D. P. Dubal, P. Gomez-Romero, B. R. Sankapal and R. Holze, *Nano Energy*, 2015, **11**, 377–399.
13 X. Xu, H. Zhou, S. Ding, J. Li, B. Li and D. Yu, *J. Power Sources*, 2014, **267**, 641–647.
14 Z. Wang, Z. Zhao, Y. Zhang, G. Pang, X. Sun, J. Zhang, L. Hou and C. Yuan, *J. Alloys Compd.*, 2019, **779**, 81–90.
15 Y. Zhu, H. Huang, G. Li, X. Liang, W. Zhou, J. Guo, W. Wei and S. Tang, *Electrochim. Acta*, 2017, **248**, 562–569.
16 J. Xiao and S. Yang, *J. Mater. Chem.*, 2012, **22**, 12253–12262.
17 H. W. Wang, Z. A. Hu, Y. Q. Chang, Y. L. Chen, H. Y. Wu, Z. Y. Zhang and Y. Y. Yang, *J. Mater. Chem.*, 2011, **21**, 10504–10511.
18 M. C. Payne, T. A. Arias and J. D. Joannopoulos, *Rev. Mod. Phys.*, 1992, **64**, 1045–1097.
19 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
20 S. L. Dudarev, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1998, **57**, 1505–1509.
21 H. J. Monkhorst, *Phys. Rev. B: Solid State*, 1977, **16**, 1748–1749.
22 J. Heyd, G. E. Scuseria and M. J. Ernzerhof, *Chem. Phys.*, 2003, **118**, 8207–8215.
23 Y.-Z. Jin, Z. Li, J.-Q. Wang, R. Li, Z.-Q. Li, H. Liu, J. Mao, C.-K. Dong, J. Yang, S.-Z. Qiao and X.-W. Du, *Adv. Energy Mater.*, 2018, **8**, 1703469.
24 X. Shi, S. L. Bernasek and A. Selloni, *J. Phys. Chem. C*, 2016, **120**, 14892–14898.
25 X. Yu, Y. Li, Y.-W. Li, J. Wang and H. Jiao, *J. Phys. Chem. C*, 2013, **117**, 7648–7655.
26 K. Shojaee, A. Montoya and B. S. Haynes, *Comput. Mater. Sci.*, 2013, **72**, 15–25.
27 E. Sanville, S. D. Kenny, R. Smith and G. Henkelman, *J. Comput. Chem.*, 2007, **28**, 899–908.
28 W. Tang, E. Sanville and G. Henkelman, *J. Phys.: Condens. Matter*, 2009, **21**, 084204.
29 J. Wu, P. Guo, R. Mi, X. Liu, H. Zhang, J. Mei, H. Liu, W. M. Lau and L.-M. Liu, *J. Mater. Chem. A*, 2015, **3**, 15331–15338.
30 K. Fominykh, G. C. Tok, P. Zeller, H. Hajiyani, T. Miller, M. Dblinger, R. Pentcheva, T. Bein and D. Fattakhova-Rohlfing, *Adv. Funct. Mater.*, 2017, **27**, 1605121.
31 Y. F. Zhang, F. Vines, Y. J. Xu, Y. Li, J. Q. Li and F. Illas, *J. Phys. Chem. B*, 2006, **110**, 15454–15458.
32 L. Zhang, J. Fischer, Y. Jia, X. Yan, W. Xu, X. Wang, J. Chen, D. Yang, H. Liu, L. Zhuang, M. Hankel, D. J. Searles, K. Huang, S. Feng, C. L. Brown and X. Yao, *J. Am. Chem. Soc.*, 2018, **140**, 10757–10763.
33 C. Xiang, Y. Liu, Y. Yin, P. Huang, Y. Zou, M. Fehse, Z. She, F. Xu, D. Banerjee, D. Hermida Merino, A. Longo, H.-B. Kraatz, D. F. Brougham, B. Wu and L. Sun, *ACS Appl. Energy Mater.*, 2019, **2**, 3389–3399.

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34 Y. F. Zhang, F. Vines, Y. J. Xu, Y. Li, J. Q. Li and F. Illas, *J. Phys. Chem. B*, 2006, **110**, 15454–15458.
35 Z. Zhang, L. Wang, W. Wang, J. Liu, Z. Hong, K. Cho and W. Wang, *J. Mater. Chem. A*, 2019, **7**, 16231–16238.
36 F. Zasada, J. Gryboś, P. Indyka, W. Piskorz, J. Kaczmarczyk and Z. Sojka, *J. Phys. Chem. C*, 2014, **118**, 19085–19097.
37 S. C. Parker, P. M. Oliver, N. H. De Leeuw, J. O. Titiloye and G. W. Watson, *Phase Transitions*, 1997, **61**, 83–107.
38 W. Zhao, M. Bajdich, S. Carey, A. Vojvodic, J. K. Nørskov and C. T. Campbell, *ACS Catal.*, 2016, **6**, 7377–7384.
39 Q. Zhang, Z. Liu, B. Zhao, Y. Cheng, L. Zhang, H.-H. Wu, M.-S. Wang, S. Dai, K. Zhang, D. Ding, Y. Wu and M. Liu, *Energy Storage Materials*, 2019, **16**, 632–645.
40 Z. Wu, Y. Zhu and X. Ji, *J. Mater. Chem. A*, 2014, **2**, 14759–14772.
41 P. Rasiyah, A. C. C. Tseung and D. B. Hibbert, *J. Electrochem. Soc.*, 1982, **129**, 1724–1727.
42 T. Y. Wei, C. H. Chen, H. C. Chien, S. Y. Lu and C. C. Hu, *Adv. Mater.*, 2010, **22**, 347–351.
43 R. Wang, M. Yao and Z. Niu, *InfoMat*, 2020, **2**, 113–125.

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