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Advanced technologies for the fabrication of MOF thin films

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Metal–organic framework (MOF) thin films represent a milestone in the development of future technological breakthroughs. The processability of MOFs as films on surfaces together with their major features (i.e. tunable porosity, large internal surface area, and high crystallinity) is broadening their range of applications to areas such as gas sensing, microelectronics, photovoltaics, and membrane-based separation technologies. Despite the recent attention that MOF thin films have received, many challenges still need to be addressed for their manufacturing and integrability, especially when an industrial scale-up perspective is envisioned. In this brief review, we introduce several appealing approaches that have been developed in the last few years. First, a summary of liquid phase strategies that comprise microfluidic methods and supersaturation-driven crystallization processes is described. Then, gas phase approaches based on atomic layer deposition (ALD) are also presented.

MOFs: moving from the bulk towards thin films

Metal–organic frameworks (MOFs) are porous hybrid crystalline materials consisting of metal-based nodes connected through organic ligands.¹ The intrinsic porous architecture of MOFs yields large surface areas, from a few hundreds of $\text{m}^2 \text{g}^{-1}$ to above ten thousand $\text{m}^2 \text{g}^{-1}$,^{2–4} resulting in excellent adsorption capabilities.^{5,6} Furthermore, the plethora of MOF lattice structures available and their exquisite chemical tunability, both associated with the ample menu of metal nodes and

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organic linkers, make them multifunctional platforms with tailorable capabilities.^{7–9} MOFs are conventionally obtained as polycrystalline micro-powders with very low processability. However, their integration onto surfaces and devices is highly desired for a large variety of applications, *e.g.* volatile organic compound (VOC) sensors, membranes for water desalination, or organic field effect transistors, to mention only a few.^{10–14} In this vein, MOF thin films are attractive systems for unprecedented technological breakthroughs as they can endow surfaces with enticing properties and advanced functionalities while minimizing the amount of material used.^{15,16} However, the general lack of understanding of growth mechanisms and the poor structural control have limited their widespread application in emerging products and/or existing markets. Also, it is necessary to say that the use of MOF thin films in applications is still not a completely mature field.¹⁷

In general, MOF thin films are prepared by the following three different approaches:

(I) Layer-by-layer (LbL) methods

These techniques are based on surface chemistry, where a self-assembled monolayer (SAM) is used as a template for the growth of MOF crystals. In this case, MOFs grow as a result of sequential reactions between the organic ligand and the inorganic nodes. Frequently, the reactions are conducted in conventional glassware (non-confined conditions) where turbulent environments dominate. Additionally, a washing step is required between each reaction to eliminate surplus reagents. Consequently, while the orientation and thickness of MOF thin films can be controlled to some extent with LbL methods, these protocols are time-consuming and labour intensive.^{15,18–20}

(II) Electrically driven synthesis

This approach consists of inducing the formation of MOFs on electrodes by means of electrochemical reactions or electrophoresis.

There are basically two electrochemical strategies. One consists of using a metal anode as the source of metal nodes to form the MOF. In this approach, the metal oxidizes and the resulting ions react with the ligands present in the electrolyte, thus leading to the formation of a MOF film on the electrode. A second approach consists of inducing the cathodic electrodeposition of the MOF by placing both the metal salts and the ligands in the electrolyte. In this method, electrogenerated hydroxyl ions at the cathode (due to the hydrogen evolution reaction) deprotonate the organic ligands, which ultimately leads to the formation of the MOF film. MOF films can also be formed on substrates by means of electrophoretic deposition of colloiddally stable charged MOF nanoparticles.^{21,22} Nevertheless, the success of this methodology is strongly determined by the intrinsic properties of the building blocks (*e.g.* electron/proton conductivity), a feature that has limited its widespread application to different MOF systems. Note that only a limited number of MOFs have been processed by electrically driven deposition approaches.^{16,23,24}

(III) Casting of MOF powders

The dispersion of MOF crystals synthesized by solvothermal methods is deposited over substrates using classical casting techniques such as spray coating, spin coating, and/or drop-casting.¹⁷ Even though these techniques can be applied to cast a broad number of MOFs, the quality of the films obtained is, in general, low. Besides, oriented thin films are extremely difficult to achieve due to the poor control of the nucleation and crystal growth offered by these approaches.

As all of these methods have been extensively reviewed in the past,^{15,19,25–30} this minireview will focus exclusively on two of the most recent approaches reported to obtain MOF thin films: liquid phase approaches operating under high surface-to-volume ratio conditions (*i.e.* microfluidic technologies and blade coating) and gas phase techniques (Fig. 1). As it will be



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understand the self-assembly processes of metal–organic crystalline materials. In 2019, he was appointed an ICREA Research Professor.

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Fig. 1 Illustration showing the different approaches discussed in this work.

discussed in the following sections, both the liquid phase and gas phase methods allow for the synthesis of MOF thin films where not only the orientation of the crystal faces can be controlled but also the film thickness. Moreover, gas phase approaches do not require the use of solvents, which makes them quite enticing for the fabrication of microelectronic devices.⁵

Liquid phase approaches

In nature, single crystals with complex and hierarchical structures can be readily obtained from solution by allocating the building block units to a desired place at the right time (*i.e.* spatiotemporal assembly).³¹ However, despite the vast number of methods performed in the liquid phase to produce MOF thin films,^{15,19,25–30} control over the crystal domain size and shape of MOFs is still missing. Accordingly, new liquid phase approaches featuring spatiotemporal control of the building

blocks during MOF synthesis have been recently considered (*vide infra*). In this vein, microfluidic technologies have gained considerable attention in the scientific community as they are able to control the spatiotemporal diffusion of building blocks;³² avoid turbulent environments, which can lead to precipitation, and thus, to uncontrolled crystal growth; and, furthermore, they can mimic nature's confined environments in the laboratory with the ambition of ultimately controlling MOF crystal growth. In this section, the discussion will mainly be focussed on MOF thin film fabrication protocols where liquid phase approaches have been exploited. Two different categories are considered: continuous-flow microfluidic synthesis and supersaturation-driven crystallization processes in confined micro- and nanoscale environments.

Continuous-flow microfluidic synthesis

Continuous-flow microfluidic devices can lead to a paradigm shift in material synthesis.³³ The laminar flow regime established





Fig. 2 (a) Schematic illustration of the interfacial microfluidic synthesis of HKUST-1 membranes either at the inner or outer surface of the polymeric hollow fibres. (b) SEM images showing HKUST-1 films synthesized using the methodology described in (a). (a and b) Reproduced from ref. 38 with permission from The Royal Society of Chemistry. (c) Schematic representation of the experimental steps followed to produce the ultrathin ZIF-8 membranes inside the polymeric hollow fibres *via* microwave assisted seeding and the microfluidic-based secondary growth method. (d and e) SEM images of the ZIF-8 membrane. (c–e). Reproduced from ref. 39 with permission from Elsevier B.V., copyright 2018.

inside microfluidic devices offers advanced synthetic conditions where constant spatiotemporal mass transport of reagents can be guaranteed.^{33,34} In a microfluidic device, the mixing of reagents only occurs through diffusion (following Fick's law), a feature that has paved the way to developing approaches for MOF thin film synthesis, and, in particular, for membrane fabrication.^{18,34}

The microfluidic synthesis of MOF membranes was first reported by Nair *et al.* in 2014,³⁵ who developed a method, known as interfacial microfluidic membrane processing (IMMP), where polymeric hollow fibres were used to control the reaction–diffusion area during the synthesis of a prototypical MOF, *i.e.* zeolitic imidazolate framework-8 (ZIF-8, zinc 2-methylimidazolate). They showed that the use of hollow fibres enables spatial control during ZIF-8 growth, which leads to the controlled fabrication of ZIF-8 membranes in the inner or at the outer surface of the fibre. These membranes could be subsequently used as molecular sieving agents. Additionally, the same group has recently shown that the thickness of the

ZIF-8 membranes can be controlled employing the IMMP approach.³⁶ In this research, for example, 5 μm thick ZIF-8 membranes could be easily obtained by performing an initial heating ramp, and engineering the pore size of the membranes. Increasing both the pore size and the temperature enhanced the diffusion of reagents and their nucleation rate, respectively, features that enabled the growth of thinner ZIF-8 membranes. Furthermore, the long-term stability and improved permeation performance of these thin ZIF-8 membranes were also demonstrated.³⁷

Additionally, Kharul and coworkers have shown that the IMMP method can also be conducted when low-boiling-point solvents are considered during MOF synthesis. This aspect is key as it enables straightforward removal of solvents, and hence mild activation of the produced MOF membranes. In their study, the authors showed the formation of both ZIF-8 and Hong Kong University of Science and Technology-1 (HKUST-1, copper(II) 1,3,5-benzenetricarboxylate) membranes on either side of the hollow fibre by controlling the positioning of metal ions and ligand precursors with respect to the polymeric hollow fibre (Fig. 2a and b).³⁸

Recently, Jeong *et al.* introduced an approach employing polymeric hollow fibres where ultrathin ZIF-8 membranes were obtained by combining microwave-assisted seeding with microfluidic-based secondary growth.³⁹ As shown in Fig. 2c, after saturating the fibre pores with metal ions, the ligand precursor solution was injected inside the hollow fibre under microwave heating to generate a densely-packed seed layer. Subsequently, a growth solution composed of the two precursors was injected and flowed until ultrathin ZIF-8 films (*ca.* 800 nm) were obtained (Fig. 2d and e).

This confined synthetic strategy of MOF thin films inside capillaries (or fibres) was previously considered employing a LbL approach. In this case the authors demonstrated the formation of a homochiral MOF thin film, *i.e.* Cu₂(D-cam)₂P (D-cam = (1R,3R)-(+)-camphorate; L-cam = (1S,3R)-(–)-camphorate; P = dabco and bipy).⁴⁰

While in previous research physical barriers (*i.e.* polymeric hollow fibres) are used to control and/or define the reaction area of precursors, recently Puigmartí-Luis and coworkers exploited the controlled diffusion of precursors inside microfluidic channels to generate MOF thin films with spatiotemporal control (Fig. 3). In this investigation, Sevim *et al.* used a reactive substrate (*i.e.* Cu₂O) from which smooth and densely-packed HKUST-1 thin films could be readily generated. This method, known as in-flow MOF lithography, enabled unprecedented control during MOF thin film synthesis (Fig. 3b and c). For example, the authors demonstrated that good control over the reaction time as well as over the diffusion of precursors onto the reactive Cu₂O substrate was key to achieve compositional gradients of HKUST-1 thin films (Fig. 3d and e). In this pioneering approach, the authors clearly demonstrated that microfluidics can also play a crucial role in surface chemistry, where spatiotemporal control during the synthesis of a functional material can be easily established.⁴¹



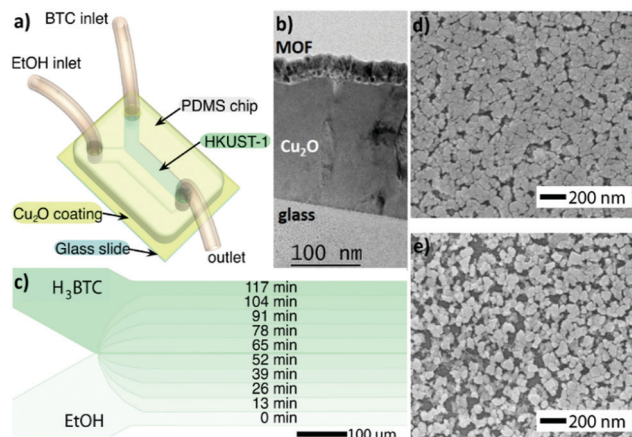


Fig. 3 (a) Schematic illustration of the in-flow MOF lithography method. (b) TEM cross-section image showing a densely-packed HKUST-1 film synthesized over a Cu₂O coated glass substrate. (c) Schematic representation of the time required to obtain compositionally graded HKUST-1 (i.e. spatiotemporal controlled synthesis of a MOF film onto a surface). (d) and (e) SEM images of a graded HKUST-1 thin film. Note that the compactness of the HKUST-1 thin film varies throughout the width of the microchannel. Reproduced from ref. 41 with permission from Wiley-VCH Verlag GmbH & Co., copyright 2019.

Supersaturation driven crystallization processes

Besides the continuous-flow microfluidic approaches, several other strategies involving confined fluidic conditions have been recently introduced to fabricate MOF thin films (*vide infra*). In this context, confined spaces featuring high surface-to-volume ratios are being used to accomplish ultimate control over MOF crystal growth. Recently, Giri and coworkers have reported a nanofluidic approach to achieve high-aspect-ratio HKUST-1 crystals.⁴² A precursor solution of HKUST-1 was confined in nanofluidic channels *via* capillary action, and then controlled crystallization was induced by heating and evaporating the solvent under a vacuum (Fig. 4a). According to their observations, the crystals initially grew perpendicularly to the surface (i.e. along the channel height), and then parallel to it until they filled the channel width, and, subsequently, along the channel length (Fig. 4b and c), hence yielding high aspect ratio HKUST-1 crystals (Fig. 4d).

Recently, confined spaces have also been used to generate conditions that mimic microgravity. These conditions have proved to be an excellent scenario to control both the growth and shape of MOF single crystals. Under these conditions, Puigmartí-Luis *et al.* have, for example, demonstrated the formation of large single crystals of a peptide-based MOF with unprecedented control over the morphology while growing non-equilibrium shapes such as thin films.⁴³

Instead of constraining MOF crystal growth into small volumes (i.e. micro/nano fluidic channels), confined spaces can also be established with common liquid phase processing techniques such as blade coating. Recently, Park *et al.* have reported the supersaturation-driven crystallization of MOF thin films in confined spaces *via* a blade coating approach.⁴⁴ Here, the supersaturation of a meniscus solution containing

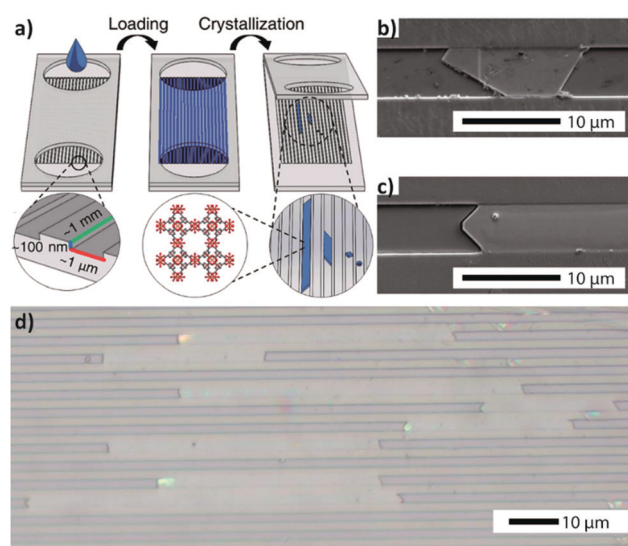


Fig. 4 (a) Schematic illustration of the nanofluidic device used for the evaporation driven crystallization approach. (b) and (c) SEM images demonstrating the confined growth of HKUST-1 crystals in confined nano-chambers. (d) An optical image of high aspect ratio HKUST-1 crystals obtained *via* the method presented in (a). Reproduced with permission from ref. 42 published by The Royal Society of Chemistry.



Fig. 5 (a) Schematic representation of the meniscus-guided crystallization method and (b) and (c) optical and corresponding SEM images of HKUST-1 films coated onto different substrates, respectively on an AAO membrane and on a flexible polyimide film. Reproduced from ref. 44 with permission from American Chemical Society, copyright 2019.

HKUST-1 precursors was finely controlled by heating the substrate and moving the blade (Fig. 5a). Changing the substrate temperature and varying the velocity of the blade enabled the fabrication of densely-packed thin films with tuneable thicknesses (down to 450 nm) on different kinds of substrates (Fig. 5b and c).

Gas phase approaches

While the liquid phase approaches discussed earlier are very attractive, the use of liquid solvents is still a drawback for some applications in digital and information technologies, where



thin films are used as dielectrics, barrier layers, and capacitors.^{45,46} MOF thin films have become an engaging alternative to the existing materials used in the microelectronics industry because in addition to displaying electric conductivity, wide dielectric constants, and high breakdown voltage, they are also recognised for their selective molecular sensing and tuneable dielectric properties.⁵

As in the case of conventional materials found in electronic devices, MOF thin films have currently been grown through a combination of well-established gas phase techniques such as physical vapour deposition (PVD, *i.e.* sputtering, evaporation, or sublimation) and chemical vapour deposition (CVD).^{5,45}

PVD techniques consist of the condensation of a vaporized form of the film on the substrate by a physical process (*e.g.* thermally or laser-induced evaporation). However, considering the high temperature usually adopted during the process (300–400 °C) to obtain the crystalline structures, it is not a technique widely used for the fabrication of MOF thin films.

In the case of CVD, a substrate is exposed to a gaseous precursor that chemically reacts with its surface. This type of deposition is appealing due to its scalability and because (in some particular cases) the reaction can take place without the use of a vacuum and at lower temperatures than those required in PVD to reach crystallinity.

An appealing variant of CVD is Atomic Layer Deposition (ALD).^{47,48} ALD is a surface-limited and self-terminating LbL deposition method, thus allowing the deposition of homogeneous thin films without the use of any solvent.^{49–54} ALD relies on alternate pulses of reactants and co-reactants in the gas phase, separated by purging steps with an inert gas (*e.g.* N₂, Ar) (Fig. 6). One of the main advantages of ALD, compared to CVD, is that the former can yield conformal films over very complex, high-aspect-ratio and even porous structures, given its surface-limited reaction nature. Also, because ALD is self-terminating, the thickness can be controlled down to the subnanometer. These two characteristics have made ALD a very appealing technique for the microelectronics industry, which has indeed

been a key driver in the development of the ALD field.^{49–54} While ALD refers to processes where only inorganic and metal-organic precursors are used, the term Molecular Layer Deposition (MLD) has been used when one or both precursors are organic, allowing the deposition of hybrid or purely organic films.^{54,55} In the case of hybrid films, such as MOFs, the organic molecules are bonded to the metal atoms and *vice versa*, forming periodic thin-film structures, where the nature of the bonds depends on the nature of the organic ligands.^{53,56,57}

The first attempt to deposit a MOF thin film by MLD was reported in 2013 by Salmi *et al.* using zinc acetate (ZnAc₂) and 1,4-benzene dicarboxylic acid (1,4-BDC) as precursors.⁵⁸ While the as-deposited films were amorphous, polycrystalline forms of MOF-5 could be obtained by placing the samples in a humidity-controlled chamber (60% humidity) for 12 h followed by annealing at 150 °C for 2 h. The conformal and uniform nature of the post-treated MOF-5 film was demonstrated by



Fig. 7 (a) Optical microscope image of a film crystallized in moist air. (b) Cross-sectional FESEM image of a MOF-5 film deposited on a silicon substrate with trenches. Reproduced from ref. 58, copyright (2013), with permission from Elsevier.

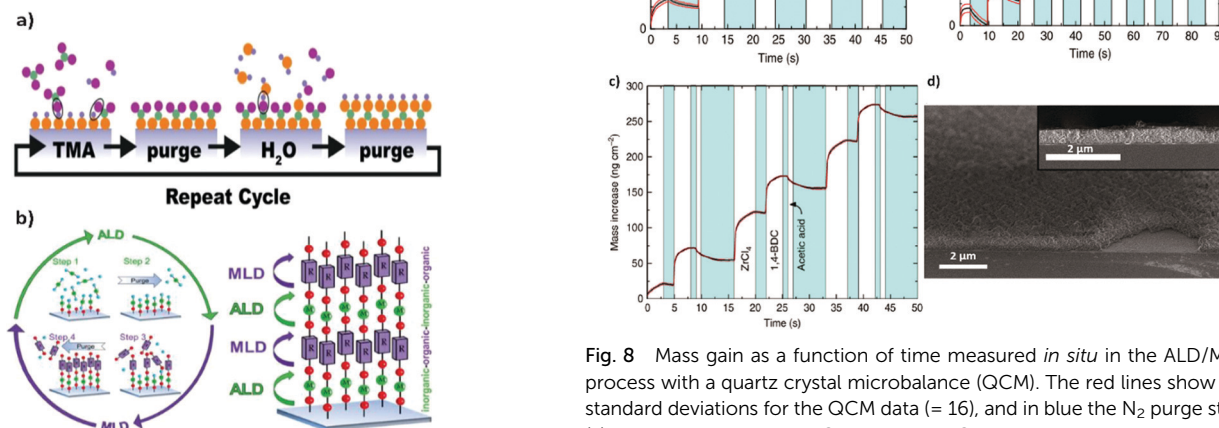


Fig. 6 (a) Schematic representation of a typical ALD cycle. Reproduced from ref. 47, copyright (2018), with permission from Elsevier. (b) Schematic representation of the ALD/MLD process where in the ALD part an inorganic precursor is used while in the MLD part an organic precursor is used. Reproduced from ref. 48, copyright (2015), with permission from Elsevier.

Fig. 8 Mass gain as a function of time measured *in situ* in the ALD/MLD process with a quartz crystal microbalance (QCM). The red lines show the standard deviations for the QCM data ($n = 16$), and in blue the N₂ purge step. (a) Alternation between ZrCl₄ and 1,4-BDC pulses separated by N₂ purges. (b) Alternation between ZrCl₄, 1,4-BDC and water. (c) ALD/MLD process in which a modulation step with acetic acid is included. (d) Cross-section SEM images of the same surface viewed at 45° and 90° angles. Reproduced from ref. 60 with permission from The Royal Society of Chemistry.



covering high aspect-ratio structures containing hard-to-access small trenches (Fig. 7).

Following the same approach, numerous groups have also reported the deposition of polycrystalline films.^{59–61} In 2016, Nilsen and his team⁶⁰ described a similar strategy that combined ALD and MLD processes to obtain University of Oslo-66 (UiO-66, zirconium terephthalate) (Fig. 8a–c). By following the reaction of zirconium tetrachloride ($ZrCl_4$) and 1,4-benzenedicarboxylic acid (1,4-BDC, commonly known as terephthalic acid) *in situ* with a quartz crystal microbalance (QCM) they demonstrated the self-limiting behaviour of the reactions and found ALD-type growth conditions. The as-obtained samples were amorphous and readily reacted with water upon exposure to the environment, causing an increase of the thickness of the film and dramatic changes in its morphology. Moreover, by comparing the X-ray patterns of the samples before and after exposure to moisture, they detected partial crystallization of unreacted 1,4-BDC. The excess linker was removed by introducing water pulses in between the $ZrCl_4$ and 1,4-BDC steps. However, the authors found that the relatively slow desorption of water from the porous film could cause undesirable side reactions with $ZrCl_4$ during subsequent pulses. To overcome this issue, the weakly-bound 1,4-BDC was removed by reacting with acetic acid, which is known to be a modulator of the crystal growth in the synthesis of UiO-66. Finally, crystalline UiO-66 films were obtained by introducing the samples in a sealed autoclave with acetic acid vapour (Fig. 8d).

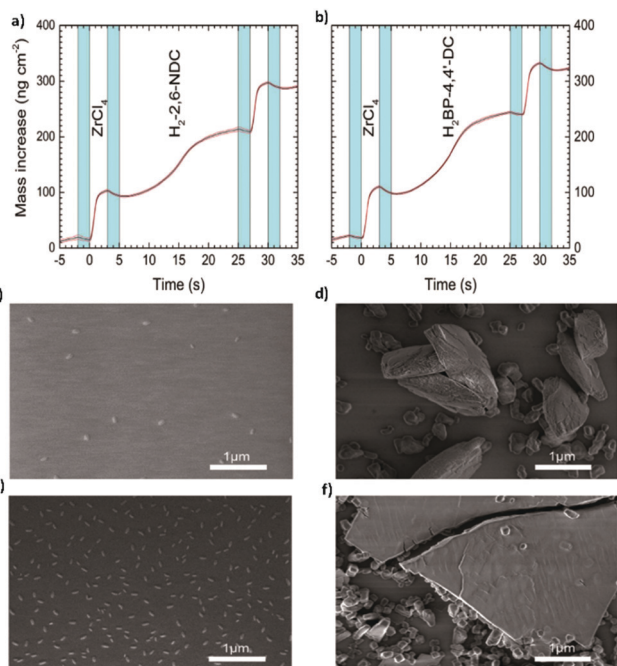


Fig. 9 Mass increase during one MLD cycle (a) of the Zr-2,6-NDC process and (b) of the Zr-BP-4,4'-DC process, measured by the *in situ* QCM technique. (c) SEM picture of the as-deposited Zr-2,6-NDC film. (d) SEM picture of the Zr-2,6-NDC film after autoclave treatment. (e) SEM picture of the as-deposited Zr-BP-4,4'-DC film. (f) SEM picture of the Zr-BP-4,4'-DC film after autoclave treatment. Reproduced from ref. 11 with permission from The Royal Society of Chemistry.

More recently, the same group¹¹ broadened their Zr-based MOF portfolio by developing a procedure involving MLD cycles of 2,6-naphthalenedicarboxylate (2,6-NDC) and biphenyl-4,4'-dicarboxylate (BP-4,4'-DC) with tetrachlorozincate ($ZnCl_4$) as a metallic precursor (Fig. 9a–c and e). The typical cycle used in this work was the following one: 3 s for the $ZrCl_4$ pulse, 2 s for purging, 20 s for the organic linkers, and 2 s of purging. The choice of using rather long pulses for the organic ligands (as compared with standard MLD processes) was due to the steric hindrance effects linked to the bulky nature of the ligands that possibly slowed down the surface kinetics.

Ahvenniemi and Karppinen also reported the deposition of Cu-TPA (copper terephthalic acid) MOF thin films using $Cu(thd)_2$ (copper 2,2,6,6-tetramethyl-3,5-heptanedione) and terephthalic acid as precursors under carefully selected conditions.^{59,62} The precursors were heated before and during the deposition to increase their vapour pressure and facilitate the reaction on top of the substrate. In this case, the possibility to directly obtain a crystalline film was demonstrated by carefully choosing the concentration of the precursor solution and the time necessary



Fig. 10 (a) Schematic representation of the CVD route in two steps. Step 1: deposition of the oxide sacrificial layer; step 2: conversion of the metal oxide into a MOF thin film. (b and c) SEM images of a ZIF-8 coating on a silicon pillar array. (d and e) High-magnification SEM pictures of the homogeneous coverage at the base of the pillar, showing that the conformality typical of the ALD process is retained after the coordination replication reaction. Reproduced from ref. 63, copyright (2016), with permission from Elsevier.



Table 1 Chronological summary of the most recent liquid phase and gas phase approaches for the deposition of MOF thin films with their possible applications

Team	Year	Method	Material	Thickness	Post-treatment	Key aspect/application	Ref.
Salmi L. D. Nair S.	2013–2016	Gas MLD <i>Liquid</i> IMMP	MOF-5 ZIF-8	— ca. 5 μm	Crystallization —	First MOF film obtained by MLD Potential control over the ZIF-8 membrane formation for gas separation	58 35–37
Kharul U. K.	2015	<i>Liquid</i> IMMP	HKUST-1, ZIF-8	10–25 μm	—	Easy removal of the solvent with a mild activation process. MOF membranes fabricated for gas separation	38
Zhang J.	2015	<i>Liquid</i> LBL epitaxial growth inside a capillary coated with 3,4-dihydroxy- <i>l</i> -phenylalanine (poly(<i>l</i> -DOPA))	Cu ₂ (<i>b</i> -cam) ₂ P (<i>b</i> -cam = (<i>1R,3</i>)(+)-camphorate; <i>l</i> -cam = (<i>1S,3R</i>)-(–)-camphorate; P = dabco and bipy)	350 nm	—	Homochiral MOF for enantiomer separation	40
Ameloot R.	2016	Gas ALD + gas phase reaction	ZIF-8	ca. 200 nm	—	First MOF-CVD reported. Well controlled thickness.	63
Karppinen M.	2016	Gas ALD/MLD	Cu-TPA	80–90 nm	—	First and only example of the direct deposition of a crystalline MOF thin film by MLD. Atomic/molecular level control of the film thickness under relatively mild conditions	62
Nilsen O.	2016	Gas ALD/MLD	UiO-66	From 130 nm as-deposited to 500 nm	Acetic acid vapour exposure	New approach to control the stoichiometry between metal cluster and organic ligands	60
Jeong H.-K.	2018	<i>Liquid</i> combination of microwave assisted seeding and microfluidic secondary growth	ZIF-8	ca. 800 nm	—	Ultrathin ZIF-8 membranes on polymeric hollow fibres for propylene/propane gas separation	39
Ameloot R.	2019	Gas ALD + gas-phase reaction	ZIF-67 and ZIF-8	ca. 30 nm	—	Gap filling of narrow spaces between the metallic trenches of a chip	46
Karppinen M.	2019	Gas ALD/MLD	Nd-TP	ca. 94 nm	Crystallization by exposure to humidity at room temperature	Atomic/molecular level control of the film thickness under relatively mild conditions	59
Puigmartí-Luis J.	2019	<i>Liquid</i> in-flow lithography	HKUST-1	ca. 15 nm	—	First controlled defect engineering of a MOF thin film <i>via</i> advanced mass transport and controlled diffusion in a confined space	41
Giri G.	2019	<i>Liquid</i> nanofluidic approach	HKUST-1	100 nm	—	High aspect ratio HKUST-1 crystals for selective small molecule uptake	42
Park S.	2019	<i>Liquid</i> supersaturation driven crystallization	HKUST-1	Down to 450 nm	—	Meniscus guided precise control of supersaturation	44
Puigmartí-Luis J.	2020	<i>Liquid</i> microfluidic approach in an advection free confined environment	Cu(II) tripeptide glycine- <i>l</i> -histidine-glycine (CuGHG)	—	—	For the first time conditions mimicking microgravity are used to grow and shape MOF single crystals	43
Nilsen O.	2020	Gas ALD/MLD	Zr MOFs with 2,6-NDC and BP-4,4'-DC	—	Acetic acid vapour exposure	Synthesis of thin films with a large range of pore size. Preliminary studies employing MOF thin films in applications such as luminescence and photocatalysis	11
Ameloot R.	2020	Gas ALD + gas phase reaction	MAF-6	ca. 90 nm	—	Synthesis of a large pore MOF material for the potential adsorption of large molecules	67
Pino V.	2020	Gas ALD + gas phase reaction + solvothermal growth step	ZIF-8	ca. 3 μm	—	Solid-phase microextraction (SPME) for analytical and pre-concentration schemes	68



for each step of the MLD cycle. To date, it remains the only example of a direct synthesis of polycrystalline MOF films without the necessity of a post-synthesis treatment. Such a difficulty could be a limitation for the application of MLD for the implementation of MOF thin films in microelectronic devices.

Similar to liquid phase approaches, the gas phase ones require spatiotemporal control of the reactants if a film with a precise degree of crystallization is desired. Although it might seem easier to control the reaction and diffusion of the precursors in liquid phase approaches, the area that can be effectively covered with a MOF film is usually very small when compared to those that can be achieved by ALD and MLD. For this reason, metal oxide films deposited by ALD have been recently used to immobilize metal ions and facilitate the MOF crystallization process. Thus, the manufacturing of well-oriented sacrificial layers that can be subsequently converted to MOF thin films through the vapor phase reaction in a closed system is possible with ALD. The first successful example was reported by Ameloot and coworkers in 2016 (Fig. 10a).^{63,64} The authors used a sacrificial ZnO thin film obtained by ALD to template the synthesis of MOFs. Thus, the ZnO film was subsequently converted to ZIF-8 through a vapor–solid reaction^{65,66} with 2-methylimidazole in a relatively short time (Fig. 10b–d). To reach the conversion of the metallic thin film into ZIF-8, a temperature higher than the 2-methylimidazole melting point was necessary to increase the amount of organic precursor close to the sacrificial metallic oxide layer. In addition, full transformation of thin (less than 15 nm) sacrificial layers into a MOF film was achieved. However, if thicker films are needed, the slow diffusion of the organic ligands through the MOF film might hamper its full conversion. Also, some problems regarding the homogeneity and continuity of the films over large areas have been noticed.⁶⁷ For instance, when very thin films (<10 nm) were left to react completely, fast crystal ripening occurred before the ZnO was completely consumed and the continuity of the films was disturbed. As pointed out by Stassin *et al.*, the need for controlling independently the temperature of the substrate and the reaction chamber walls to obtain a desired morphology remains a challenge.⁶⁷

Still, the MOF-CVD approach has recently allowed the introduction of MOF thin films as ultra-low k dielectric materials for advanced chip interconnects. As demonstrated by Krishtab *et al.*,⁴⁶ ZIF-8 and ZIF-67 (CoO_x and 2-methylimidazole) were deposited to fill the narrow gaps between the metallic trenches of a chip. The deposited MOF thin films displayed comparable performance with respect to other cutting-edge dielectrics such as porous organosilica thin films. Moreover, the accompanying volume expansion that results from the transformation of the metal oxide to a MOF enabled uniform growth over the metallic trenches and superior gap-filling. According to the existing literature, the application of MOF thin films prepared by the MOF-CVD approach is currently underway as can be seen in recent work by Pino *et al.*⁶⁸ where solid-phase microextraction (SPME) fibres were coated with ZIF-8. Nevertheless, for microelectronic applications, the MOF-CVD approach must be optimized to be compatible with automated cleanroom facilities.⁶⁴

In all the gas phase approaches reviewed above, the main characteristic has been the absence of solvents during the synthesis procedure. As previously stated, this can be advantageous because it could avoid surface-tension phenomena or eventual corrosion of the final device, and could also lead to a reduction of costs from a scale-up perspective.

Also, in the reported papers, the temperature plays a fundamental role to achieve the deposition of thin films, regardless of their crystallinity. In fact, in all the syntheses that have been reviewed here, the heating of the precursors to improve the reaction kinetics is necessary. This is because of the low volatility of the organic precursors and the competitive reactions that could take place during the process. The need for heating the precursors could be a problem since organic compounds tend to be thermally sensitive. Indeed, the range of temperatures available to tune their vapour pressure is usually quite narrow and dangerously close to the decomposition temperature of the compound itself.

Conclusions and outlook

MOF thin films have aroused growing interest in recent years due to their potential applications in microelectronics, catalysis, and gas separation. Recent developments are undeniably associated with the interesting features of MOF surfaces and their interactions, as well as the possibility of tailoring their properties while minimizing the amount of material used.

Although several solution-based techniques have been proposed for the processing of MOF thin films, the ability to carefully control their size and shape while having an efficient and environmentally friendly process remains unaddressed. The most recent liquid phase approaches reported from this perspective (*i.e.* continuous-flow fluidic approaches and supersaturation-driven crystallization) represent an important step toward the continuous production of MOF thin films with extraordinary control over different reaction parameters. At the same time, gas phase approaches (*i.e.* MOF-CVD and MLD) are arising as an alternative way, owing to their high control over the thickness and the conformality of polycrystalline films while avoiding the use of solvents. These approaches are summarized in Table 1.

Despite the advantages offered by the approaches discussed here, many of their steps demand a more fundamental understanding (*e.g.* the nucleation process, kinetics, the direction of growth), and sometimes multiple-step approaches are necessary. Thus, concerted efforts from different research and engineering fields are timely and expected. Only a multidisciplinary approach will ultimately allow for the translation of these processes to the industrial landscape. We hope that this Mini-review will engage more researchers and engineers in this technological endeavor.

Conflicts of interest

There are no conflicts to declare.



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Notes and references

- O. M. Yaghi, G. Li and H. Li, *Nature*, 1995, **378**, 703–706.
- J. Liu, F. Yang, Q. Zhang, W. Chen, Y. Gu and Q. Chen, *Inorg. Chem.*, 2019, **58**, 3564–3568.
- O. K. Farha, I. Eryazici, N. C. Jeong, B. G. Hauser, C. E. Wilmer, A. A. Sarjeant, R. Q. Snurr, S. T. Nguyen, A. Ö. Yazaydin and J. T. Hupp, *J. Am. Chem. Soc.*, 2012, **134**, 15016–15021.
- S. Joseph, D. M. Kempaiah, M. Benzigar, A. V. Baskar, S. N. Talapaneni, S. H. Jhung, D.-H. Park and A. Vinu, *J. Mater. Chem. A*, 2017, **5**, 21542–21549.
- I. Stassen, N. Burtch, A. Talin, P. Falcaro, M. Allendorf and R. Ameloot, *Chem. Soc. Rev.*, 2017, **46**, 3185–3241.
- M. A. Solomos, F. J. Claire and T. J. Kempa, *J. Mater. Chem. A*, 2019, **7**, 23537–23562.
- Y. Luo, M. Ahmad, A. Schug and M. Tsotsalas, *Adv. Mater.*, 2019, **31**, 1901744.
- L. Jiao, J. Y. R. Seow, W. S. Skinner, Z. U. Wang and H.-L. Jiang, *Mater. Today*, 2019, **27**, 43–68.
- P. Falcaro, K. Okada, T. Hara, K. Ikigaki, Y. Tokudome, A. W. Thornton, A. J. Hill, T. Williams, C. Doonan and M. Takahashi, *Nat. Mater.*, 2017, **16**, 342–348.
- Z.-G. Gu, S.-C. Chen, W.-Q. Fu, Q. Zheng and J. Zhang, *ACS Appl. Mater. Interfaces*, 2017, **9**, 7259–7264.
- K. B. Lausund, M. S. Olsen, P.-A. Hansen, H. Valen and O. Nilsen, *J. Mater. Chem. A*, 2020, **8**, 2539–2548.
- J. Liu, W. Zhou, J. Liu, Y. Fujimori, T. Higashino, H. Imahori, X. Jiang, J. Zhao, T. Sakurai, Y. Hattori, W. Matsuda, S. Seki, S. K. Garlapati, S. Dasgupta, E. Redel, L. Sun and C. Wöll, *J. Mater. Chem. A*, 2016, **4**, 12739–12747.
- A. H. Assen, O. Yassine, O. Shekhah, M. Eddaoudi and K. N. Salama, *ACS Sens.*, 2017, **2**, 1294–1301.
- S. Hurrle, S. Friebe, J. Wohlgemuth, C. Wöll, J. Caro and L. Heinke, *Chem. – Eur. J.*, 2017, **23**, 2294–2298.
- J.-L. Zhuang, A. Terfort and C. Wöll, *Coord. Chem. Rev.*, 2016, **307**, 391–424.
- W.-J. Li, M. Tu, R. Cao and R. A. Fischer, *J. Mater. Chem. A*, 2016, **4**, 12356–12369.
- J. Liu and C. Wöll, *Chem. Soc. Rev.*, 2017, **46**, 5730–5770.
- W. Li, *Prog. Mater. Sci.*, 2019, **100**, 21–63.
- J. Benito, S. Sorribas, I. Lucas, J. Coronas and I. Gascon, *ACS Appl. Mater. Interfaces*, 2016, **8**, 16486–16492.
- K. Ikigaki, K. Okada, Y. Tokudome, T. Toyao, P. Falcaro, C. J. Doonan and M. Takahashi, *Angew. Chem., Int. Ed.*, 2019, **58**, 6886–6890.
- M. J. Van Vleet, T. Weng, X. Li and J. R. Schmidt, *Chem. Rev.*, 2018, **118**, 3681–3721.
- S. D. Worrall, H. Mann, A. Rogers, M. A. Bissett, M. P. Attfield and R. A. W. Dryfe, *Electrochim. Acta*, 2016, **197**, 228–240.
- C. Warakulwit, S. Yadnum, C. Boonyuen, C. Wattanakit, A. Karajic, P. Garrigue, N. Mano, D. Bradshaw, J. Limtrakul and A. Kuhn, *CrystEngComm*, 2016, **18**, 5095–5100.
- N. Campagnol, T. R. C. Van Assche, M. Li, L. Stappers and J. Fransaer, *J. Mater. Chem. A*, 2016, **4**, 3914–3925.
- Z. G. Gu and J. Zhang, *Coord. Chem. Rev.*, 2019, **378**, 513–532.
- X. Bai, D. Chen, L. Li, L. Shao, W. He, H. Chen, Y. Li, X. Zhang, L. Zhang, T. Wang, Y. Fu and W. Qi, *ACS Appl. Mater. Interfaces*, 2018, **10**, 25960–25966.
- J. U. Balderas, D. Navarro, V. Vargas, M. M. Tellez-Cruz, S. Carmona and C. Falcony, *J. Lumin.*, 2019, **212**, 322–327.
- B. Reif, J. Somboonvong, F. Fabisch, M. Kaspereit, M. Hartmann and W. Schwieger, *Microporous Mesoporous Mater.*, 2019, **276**, 29–40.
- V. Chernikova, O. Shekhah and M. Eddaoudi, *ACS Appl. Mater. Interfaces*, 2016, **8**, 20459–20464.
- X. Liu, M. Kozłowska, T. Okkali, D. Wagner, T. Higashino, G. Brenner-Weiß, S. M. Marschner, Z. Fu, Q. Zhang, H. Imahori, S. Bräse, W. Wenzel, C. Wöll and L. Heinke, *Angew. Chem., Int. Ed.*, 2019, **58**, 9590–9595.
- G. M. Whitesides, *Science*, 2002, **295**, 2418–2421.
- J. Puigmartí-Luis, *Chem. Soc. Rev.*, 2014, **43**, 2253–2271.
- C. Echaide-Górriz, C. Clément, F. Cacho-Bailo, C. Téllez and J. Coronas, *J. Mater. Chem. A*, 2018, **6**, 5485–5506.
- S. Sevim, A. Sorrenti, C. Franco, S. Furukawa, S. Pané, A. J. Demello and J. Puigmartí-Luis, *Chem. Soc. Rev.*, 2018, **47**, 3788–3803.
- A. J. Brown, N. A. Brunelli, K. Eum, F. Rashidi, J. R. Johnson, W. J. Koros, C. W. Jones and S. Nair, *Science*, 2014, **345**, 72–75.
- K. Eum, A. Rownaghi, D. Choi, R. R. Bhave, C. W. Jones and S. Nair, *Adv. Funct. Mater.*, 2016, **26**, 5011–5018.
- K. Eum, C. Ma, A. Rownaghi, C. W. Jones and S. Nair, *ACS Appl. Mater. Interfaces*, 2016, **8**, 25337–25342.
- B. P. Biswal, A. Bhaskar, R. Banerjee and U. K. Kharul, *Nanoscale*, 2015, **7**, 7291–7298.
- M. J. Lee, M. R. Abdul Hamid, J. Lee, J. S. Kim, Y. M. Lee and H.-K. Jeong, *J. Membr. Sci.*, 2018, **559**, 28–34.
- Z.-G. Gu, W.-Q. Fu, X. Wu and J. Zhang, *Chem. Commun.*, 2016, **52**, 772–775.
- S. Sevim, C. Franco, H. Liu, H. Roussel, L. Rapenne, J. Rubio-Zuazo, X. Chen, S. Pané, D. Muñoz-Rojas, A. J. DeMello and J. Puigmartí-Luis, *Adv. Mater. Technol.*, 2019, **4**, 1800666.
- S. Guthrie, L. Huelsenbeck, A. Salahi, W. Varhue, N. Smith, X. Yu, L. U. Yoon, J. J. Choi, N. Swami and G. Giri, *Nanoscale Adv.*, 2019, **1**, 2946–2952.



- 43 A. Sorrenti, L. Jones, S. Sevim, X. Cao, A. J. DeMello, C. Martí-Gastaldo and J. Puigmartí-Luis, *J. Am. Chem. Soc.*, 2020, **142**, 9372–9381.
- 44 J.-C. Lee, J.-O. Kim, H.-J. Lee, B. Shin and S. Park, *Chem. Mater.*, 2019, **31**, 7377–7385.
- 45 I. Stassen, D. De Vos and R. Ameloot, *Chem. – Eur. J.*, 2016, **22**, 14452–14460.
- 46 M. Krishtab, I. Stassen, T. Stassin, A. J. Cruz, O. O. Okudur, S. Armini, C. Wilson, S. De Gendt and R. Ameloot, *Nat. Commun.*, 2019, **10**, 3729.
- 47 F. Grillo, J. A. Moulijn, M. T. Kreutzer and J. R. van Ommen, *Catal. Today*, 2018, **316**, 51–61.
- 48 T. Tynell and M. Karppinen, in *Oxide Thin Films, Multilayers, and Nanocomposites*, ed. P. Mele, T. Endo, S. Arisawa, C. Li and T. Tsuchiya, Springer International Publishing, Cham, 2015, pp. 159–179.
- 49 T. Muneshwar, M. Miao, E. R. Borujeny and K. Cadien, *Handbook of Thin Film Deposition*, Elsevier, 2018, pp. 359–377.
- 50 T. Tynell, T. Aizawa, I. Ohkubo, K. Nakamura and T. Mori, *J. Cryst. Growth*, 2016, **449**, 10–14.
- 51 E. Ahvenniemi, A. R. Akbashev, S. Ali, M. Bechelany, M. Berdova, S. Boyadjiev, D. C. Cameron, R. Chen, M. Chubarov, V. Cremers, A. Devi, V. Drozd, L. Elnikova, G. Gottardi, K. Grigoras, D. M. Hausmann, C. S. Hwang, S.-H. Jen, T. Kallio, J. Kanervo, I. Khmel'nitskiy, D. H. Kim, L. Klibanov, Y. Koshtyal, A. O. I. Krause, J. Kuhs, I. Kärkkäinen, M.-L. Kääriäinen, T. Kääriäinen, L. Lamagna, A. A. Łapicki, M. Leskelä, H. Lipsanen, J. Lyytinen, A. Malkov, A. Malygin, A. Mennad, C. Militzer, J. Molarius, M. Norek, Ç. Özgüt-Akgün, M. Panov, H. Pedersen, F. Pierrat, G. Popov, R. L. Puurunen, G. Rampelberg, R. H. A. Ras, E. Rauwel, F. Roozeboom, T. Sajavaara, H. Salami, H. Savin, N. Schneider, T. E. Seidel, J. Sundqvist, D. B. Suyatin, T. Törndahl, J. R. van Ommen, C. Wiemer, O. M. E. Ylivaara and O. Yurkevich, *J. Vac. Sci. Technol., A*, 2017, **35**, 010801.
- 52 D. Muñoz-Rojas and J. MacManus-Driscoll, *Mater. Horiz.*, 2014, **1**, 314–320.
- 53 J. Penttinen, M. Nisula and M. Karppinen, *Chem. – Eur. J.*, 2017, **23**, 18225–18231.
- 54 M. Nisula and M. Karppinen, *Nano Lett.*, 2016, **16**, 1276–1281.
- 55 A. Tanskanen and M. Karppinen, *Sci. Rep.*, 2018, **8**, 8976.
- 56 P. Sundberg and M. Karppinen, *Beilstein J. Nanotechnol.*, 2014, **5**, 1104–1136.
- 57 X. Meng, *J. Mater. Chem. A*, 2017, **5**, 18326–18378.
- 58 L. D. Salmi, M. J. Heikkilä, E. Puukilainen, T. Sajavaara, D. Grosso and M. Ritala, *Microporous Mesoporous Mater.*, 2013, **182**, 147–154.
- 59 E. Ahvenniemi and M. Karppinen, *Chem. Mater.*, 2016, **28**, 6260–6265.
- 60 K. B. Lausund and O. Nilsen, *Nat. Commun.*, 2016, **7**, 13578.
- 61 A. Ghazy, M. Safdar, M. Lastusaari and M. Karppinen, *Chem. Commun.*, 2020, **56**, 241–244.
- 62 E. Ahvenniemi and M. Karppinen, *Chem. Commun.*, 2016, **52**, 1139–1142.
- 63 I. Stassen, M. Styles, G. Greci, H. Van Gorp, W. Vanderlinden, S. De Feyter, P. Falcaro, D. De Vos, P. Vereecken and R. Ameloot, *Nat. Mater.*, 2016, **15**, 304–310.
- 64 A. J. Cruz, I. Stassen, M. Krishtab, K. Marcoen, T. Stassin, S. Rodríguez-Hermida, J. Teyssandier, S. Pletincx, R. Verbeke, V. Rubio-Giménez, S. Tatay, C. Martí-Gastaldo, J. Meersschaut, P. M. Vereecken, S. De Feyter, T. Hauffman and R. Ameloot, *Chem. Mater.*, 2019, **31**, 9462–9471.
- 65 Y. Zhao, N. Kornienko, Z. Liu, C. Zhu, S. Asahina, T.-R. Kuo, W. Bao, C. Xie, A. Hexemer, O. Terasaki, P. Yang and O. M. Yaghi, *J. Am. Chem. Soc.*, 2015, **137**, 2199–2202.
- 66 P. C. Lemaire, J. Zhao, P. S. Williams, H. J. Walls, S. D. Shepherd, M. D. Losego, G. W. Peterson and G. N. Parsons, *ACS Appl. Mater. Interfaces*, 2016, **8**, 9514–9522.
- 67 T. Stassin, I. Stassen, J. Marreiros, A. J. Cruz, R. Verbeke, M. Tu, H. Reinsch, M. Dickmann, W. Egger, I. F. J. Vankelecom, D. E. De Vos and R. Ameloot, *Chem. Mater.*, 2020, **32**, 1784–1793.
- 68 P. Rocío-Bautista, A. Gutiérrez-Serpa, A. J. Cruz, R. Ameloot, J. H. Ayala, A. M. Afonso, J. Pasán, S. Rodríguez-Hermida and V. Pino, *Talanta*, 2020, **215**, 120910.

