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# Determination of cesium ions in environmental water samples with a magnetic multi-walled carbon nanotube imprinted potentiometric sensor†

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A potentiometric sensor, based on the glassy carbon electrode (GCE) modified with a magnetic multi-walled carbon nanotubes/cesium ion-imprinted polymer composite (MMWCNTs@Cs-IIP), is introduced for the detection of cesium(I). The IIP was synthesized using cesium ions as the template ions, chitosan as the functional monomer and glutaraldehyde as the cross-linking agent. The membrane, which was coated on the surface of the GCE, was prepared using MMWCNTs@Cs(I)-IIP as the modifier, PVC as the neutral carrier, 2-nitrophenyloctyl ether as the plasticizer and sodium tetraphenylborate as the lipophilic salt. The proposed sensor exhibited a Nernstian slope of 0.05954 V dec<sup>-1</sup> in a working concentration range of  $1 \times 10^{-7}$  to  $1 \times 10^{-4}$  M (mol L<sup>-1</sup>) with a detection limit of  $4 \times 10^{-8}$  M. The sensor exhibited high selectivity for cesium ions and was successfully applied for the determination of Cs(I) in real samples.

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## Introduction

Recently, the demand for cesium in the international market has increased significantly because cesium plays an increasingly important role in the national defense, aerospace, energy, materials, electronics and medicine fields.<sup>1–3</sup> Salt lake brine stores a considerable amount of cesium resources, but the composition of the brine system is complex. During the utilization process, the separation of cesium from co-existing elements such as magnesium, calcium, sodium and potassium is very important.<sup>4–6</sup> On the other hand, Cs is a very dangerous radioactive source. When accidentally released into the ground and sea, it must be immediately captured for public safety.<sup>7,8</sup> However, these issues have been rather challenging because the concentration of the Cs ions is usually much lower than those of the co-existing competing cations.<sup>9</sup>

Several instrumental methods including flame atomic absorption spectrometry (FAAS), inductively coupled plasma mass spectrometry (ICP-MS) and ion chromatography (IC) have been introduced for measuring cesium.<sup>10–14</sup> These techniques require expensive instruments and time-consuming sample pretreatments. Recently, much attention has been given to electrochemical methods due to their simplicity and low cost, but some problems involving the lack of sensitivity and selectivity have not been resolved.<sup>15–17</sup>

Ion imprinted technology is a versatile approach for the preparation of synthetic materials that are capable of recognizing template ions. The high selectivity of ion imprinting

polymers (IIPs) arises from active recognition sites that are complementary to the shape and size of the imprinted ions.<sup>18–20</sup> Several IIPs-based potentiometric sensors have been reported for sensing ions.<sup>21–24</sup> However, most of the sensors are proposed for heavy metals detection, and cesium potentiometric sensors have rarely been reported.

Multiwalled carbon nanotubes (MWCNTs) and Fe<sub>3</sub>O<sub>4</sub> nanoparticles (NPs) have also been employed in sensors as modification materials due to their large surface area, good stability, conductivity and electron transfer, which can improve the analytical sensitivity of the electrochemical sensor.<sup>25–29</sup> The integration of MWCNTs and Fe<sub>3</sub>O<sub>4</sub> NPs is expected to provide a synergistic effect in electrochemical sensor construction.

In the present study, a Cs(I)-imprinted polymer-based magnetic multiwalled carbon nanotube (MMWCNTs@Cs(I)-IIP) was synthesized for the selective detection of cesium ions in aqueous solutions. The polymer could be easily separated with an external magnetic field without either additional centrifugation or filtration procedures. The MMWCNTs@Cs(I)-IIP was then used as an ionophore to prepare a PVC-based membrane covered on the surface of a glassy carbon electrode. Multi-walled carbon nanotubes are used as the substrate of cesium ion imprinted polymer, which has the advantages of high mechanical strength, stability under acidic conditions, non-swelling, large surface area, unique chemical properties and easy processing. Compared with previous work, the use of multi-walled carbon nanotubes as the base material improves the sensitivity and accuracy of ion-selective electrodes.<sup>30</sup> The performance of the prepared electrochemical sensor was investigated in detail by differential pulse voltammetry (DPV), cyclic voltammetry (CV) and chronopotentiometry. The developed sensor was effectively used for sensing of the Cs(I) ion in the collected environmental samples.

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## Experimental

### Chemicals and materials

MWCNTs (5–10 nm diameter, 5–15  $\mu\text{m}$  length), CsCl,  $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$  (AR), carboxylated chitosan (CTS, BR, water-soluble), ammonium persulfate (AR), sodium bisulfate (AR), aqueous acetic acid (AR), glutaraldehyde solution (50%), Nafion solution and tetrahydrofuran (THF, AR) were obtained from Aladdin Reagent Co., Ltd. (Shanghai, China). In addition, *o*-nitrophenyloctylether (*o*-NPOE), dibutylphthalate (DBP), sodium tetraphenylborate (NaTPB), high molecular weight polyvinylchloride (PVC) and sorbitol oleate (Span-80) were purchased from Macklin (Shanghai, China). The bare glassy carbon electrodes (GCEs) were purchased from Gaoss Union.

River water was collected from the Beichuan river in Xining. Industrial wastewater was collected from an aluminum factory in Xining. The sample of brine was collected from a salt lake in Qaidam Basin, Qinghai. All the real samples were filtered through a 0.45  $\mu\text{m}$  polycarbonate membrane, and the pH value is adjusted to 6.5 with 0.1  $\text{mol L}^{-1}$   $\text{HNO}_3$ .

### Apparatus

All the electrochemical measurements were obtained with a three-electrode system using a PGSTAT204 electrochemical workstation (Switzerland). The concentration of metal ions in the solution was determined using a flame atomic absorption spectrometer (FAAS, TAS990, Universal, China) or an inductively coupled plasma mass spectrometer (ICP-MS, Thermo, USA). Fourier transform infrared spectra in KBr were recorded in the range of 400–4000  $\text{cm}^{-1}$  using a FT-IR BXII spectrometer (Perkin-Elmer, USA). Energy dispersive spectrometry (EDS) (JSM-5610LV/INCA, United Oxford Instruments) operated at 20 kV was performed to confirm the existence of elements in the Cs(I)-IIP. Scanning Electron Microscope (SEM) images were taken using a JEOL-SEM (JSM-7001F, USA) operated at 15 kV. Transmission Electron Microscopy (TEM) images were taken using a Tecnai TF20 G2 FEG-TEM at 200 kV accelerating voltage with a standard single tilt holder. Ultra-pure water (18.25  $\text{M}\Omega\text{ cm}$ ) was used to prepare all the solutions, which was obtained from a Milli-Q Direct 16 water purification system (Millipore, Bedford, MA, USA).

**Carboxyl-functionalized MWCNTs.** The MWCNTs were heated at 300  $^\circ\text{C}$  for 75 min to remove amorphous carbon, graphite nanoparticles and catalyst impurities. Then, 0.50 g of treated MWCNTs was dispersed in 100 mL nitrifying mixture solution ( $\text{H}_2\text{SO}_4 : \text{HNO}_3 = 3 : 1$  v/v). The mixture was sonicated for 45 min and then stirred under magnetic stirring at 75  $^\circ\text{C}$  for 12 h. After the reaction was completed, the solid was filtered through a 0.45  $\mu\text{m}$  polycarbonate membrane and washed several times with ultrapure water until the pH was neutral. Finally, the filtered solid was dried at 70  $^\circ\text{C}$  overnight to obtain the carboxyl-functionalized MWCNTs (MWCNTs-COOH).

**Synthesis of the magnetic MWCNTs/ $\text{Fe}_3\text{O}_4$  nanocomposite (MMWCNTs-COOH).** In total, 0.250 g MWCNTs-COOH, 0.745 g  $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$  and 0.383 g  $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$  with 100 mL water were added into a three-neck flask in a 70  $^\circ\text{C}$  bath. After that, 10 mL 0.5 M NaOH was added dropwise to the solution under nitrogen atmosphere and stirring. The mixture was performed at 70  $^\circ\text{C}$  for 24 h. Finally, the magnetic carbon nanotubes were prepared.

### Preparation of MMWCNTs@Cs(I)-IIP and MMWCNTs@-NIP.

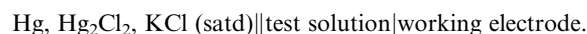
The MMWCNTs@Cs(I)-IIP were prepared by surface imprinting polymerization. In total, 0.50 g MMWCNTs-COOH, 2.0 g chitosan and 0.40 g CsCl were dissolved in 150 mL 2% acetic acid aqueous. The mixture solution was sonicated for 45 min. Then, Span-80 as a surfactant, ammonium persulfate and sodium bisulfate as initiators, and glutaraldehyde solution as the cross-linking agent were added sequentially. Upon completion of the 12 h reaction under stirring of return flow and a nitrogen environment, the mixture was cooled to room temperature, and the magnetic composite material (unleached MMWCNTs@Cs(I)-IIP) was separated with a permanent magnet. Chelated Cs(I) was removed from the MMWCNTs@Cs(I)-IIP by rinsing with HCl (1.0 M) for 24 h. The eluates were analyzed by FAAS to ensure that there was no Cs(I) left in the polymer (Table S1†). The solid was dried at 60  $^\circ\text{C}$  for 24 h to obtain the leached MMWCNTs@Cs(I)-IIP. A non-imprinted magnetic multi-walled carbon nanotube polymer (MMWCNTs@-NIP) was prepared as a blank, without the addition of CsCl.

**Preparation of the sensor MMWCNTs@Cs(I)-IIP/GCE.** Glassy Carbon Electrodes (GCEs) were polished to a bright surface with an abrasive powder, and the electrodes were rinsed repeatedly with ultrapure water and dried at room temperature. In total, 20 mg PVC, 10 mg IIP, 20 mg NaTPB, 50  $\mu\text{L}$  NPOE, 10  $\mu\text{L}$  Nafion solution and 2.5 mL THF were thoroughly mixed, and the mixture was homogenized with an ultrasonic device until viscous. The mixture was poured into a glass dish, and the polished glassy carbon electrode was immersed in the mixture for 5–10 s. It was then taken out and allowed to dry at room temperature to form a thin PVC film on the surface of the GCE. By accurately weighing the weight of the electrode before and after deposition, the amount of the Cs-IIP material deposited on the GCE is about 0.2 mg. The prepared electrodes were conditioned for 48 h in a  $1.0 \times 10^{-4}$  M Cs solution at pH 6.5.

### Analytical procedure

All the electrochemical measurements were obtained with a three-electrode system. MMWCNTs@Cs(I)-IIP/GCE were employed as the working electrode with a potassium chloride saturated calomel electrode (SCE) and a Pt sheet electrode as the reference and auxiliary electrodes, respectively.

The electrochemical cell employed for the potentiometric device add to the Fig. S1.† The response of the fabricated sensor was examined by measuring the potentials of the following electrochemical cell:



The PVC membrane electrode was scanned several times in a blank matrix electrolyte solution until a stable electrochemical signal was obtained. After each experiment, the membrane electrode was stirred in a 0.6 M HCl solution for 10 min to elute ions adsorbed on the surface of the membrane electrode. The potentials were measured by varying the concentration of Cs(I) in test solutions over a range of  $1.0 \times 10^{-8}$  to  $1.0 \times 10^{-1}$  M, and



the pH of test solutions was adjusted to 6.5. The calibration curve was determined by plotting the potential  $E$ , versus the logarithmic values of the  $\text{Cs(I)}$  concentration.

## Results and discussion

### The preparation and characterization of the MMWCNTs@Cs(I)-IIP

A feasible strategy for the preparation of the MMWCNTs@Cs(I)-IIP is shown in Scheme 1. The multi-walled carbon nanotubes were carboxylated with nitric acid and sulfuric acid system. The obtained carboxylated carbon nanotubes further interact with iron ions ( $\text{Fe}^{2+} : \text{Fe}^{3+}$  molar ratio = 2 : 1) to prepare magnetic multi-walled carbon nanotubes (MMWCNTs) in sodium hydroxide solution. Then the Cs(I) ion-imprinted polymer composite MMWCNTs@Cs(I)-IIP was synthesized by surface-imprinting technique, using MMWCNTs, Cs(I) ion, chitosan and glutaraldehyde as supporter, template ion, functional monomer and cross-linking agent respectively.<sup>31–33</sup> The  $-\text{COOH}$  groups coated on the magnetic MWCNTs' surface resulted in the imprinted layer structure being constructed. In this reaction, ammonium persulfate and sodium bisulfite together form a redox initiation system, which uses oxygen radicals generated by electron transfer between an oxidizing agent and a reducing agent to start a polymerization reaction. The advantage of this method is that the reaction rate can be increased, and the activation energy of the entire polymerization reaction can be reduced.<sup>34,35</sup>

FT-IR spectra were obtained to characterize the chemical structures of the raw-MWCNTs, MWCNTs-COOH, unleached MMWCNTs@Cs(I)-IIP and leached MMWCNTs@Cs(I)-IIP. The results are shown in Fig. 1. The spectra of the leached-IIP and unleached-IIP were basically consistent, indicating there was no significant effect on the backbone of the polymer after elution and removal of Cs(I). In the spectrum of the MWCNTs-COOH, the peak at  $1402\text{ cm}^{-1}$  was attributed to C-O stretching, while the peak at  $1639\text{ cm}^{-1}$  was attributed to the vibrations of the

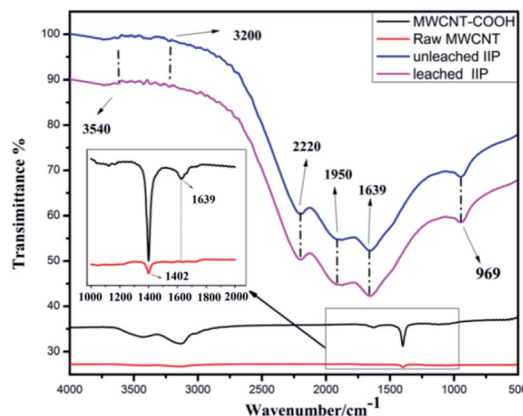
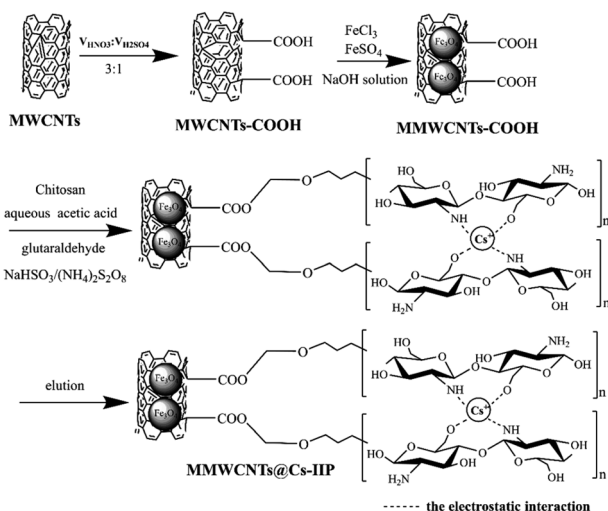


Fig. 1 The FT-IR spectra of the raw-MWCNTs, MWCNTs-COOH, unleached MMWCNTs@Cs(I)-IIP and leached MMWCNTs@Cs(I)-IIP.

COOH band.<sup>36</sup> The characteristic peaks at  $1950$  and  $969\text{ cm}^{-1}$  were attributed to the stretching vibrations of  $\text{C}=\text{O}$  and  $\text{CH}-\text{CH}_2$ . Noticeable N-H and O-H stretching vibration absorptions can be seen at  $3540\text{ cm}^{-1}$ , indicating that a polymer may be wrapped around the surface of the carbon nanotubes.

The XRD patterns of the raw MWCNTs, MMWCNTs@Cs(I)-IIP are shown in Fig. 2. The diffraction peak at  $2\theta = 26.02^\circ$  corresponds to the (002) planes of the raw MWCNTs, which was observed for both the  $\text{Fe}_3\text{O}_4$ @MWCNTs and MWCNTs@Cs(I)-IIP. The diffraction peaks at  $2\theta = 30.09^\circ, 35.46^\circ, 43.24^\circ, 53.40^\circ, 57.13^\circ$  and  $62.54^\circ$  correspond to the (220), (311), (400), (422), (511) and (440) planes.<sup>37</sup> It shows that  $\text{Fe}_3\text{O}_4$  magnetic core was successfully introduced into MWCNTs. Since the imprinted layer materials are all amorphous structures, no other diffraction peaks can be observed in the XRD spectrum.

The surface morphologies of the raw MWCNTs and MWCNTs@Cs(I)-IIP were characterized with SEM and TEM. In Fig. 3a and c, the MWCNTs exhibit a threadlike nanotube structure with a diameter ranging from 8 to 10 nm. As shown in Fig. 3b and d, after functionalization of the MWCNTs followed by the polymerization process, the  $\text{Fe}_3\text{O}_4$  nanoparticles and the imprinted polymer layer were grafted on the MWCNTs surface,



Scheme 1 The preparation process for MMWCNTs@Cs(I)-IIP.

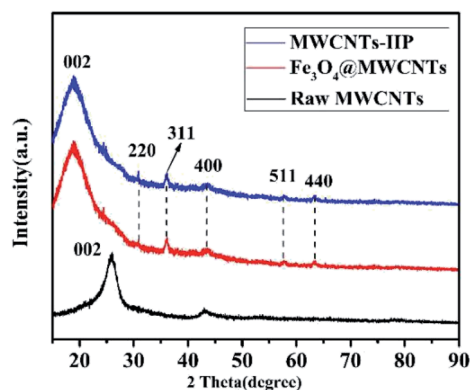


Fig. 2 The XRD patterns of the raw MWCNTs,  $\text{Fe}_3\text{O}_4$ @MWCNTs and MWCNTs@Cs(I)-IIP.





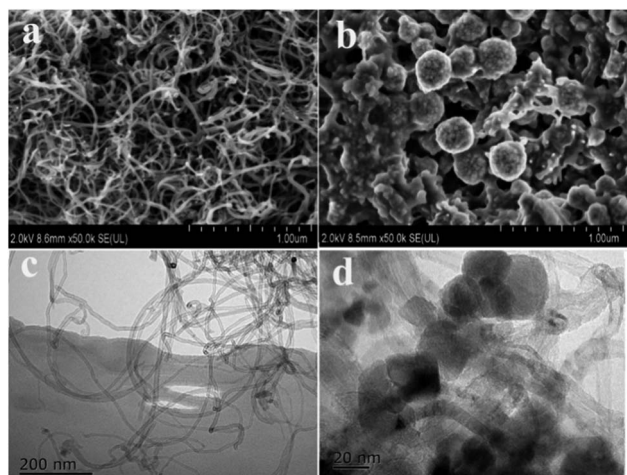


Fig. 3 SEM images of Raw MWCNTs (a), and MMWCNTs@Cs(I)-IIP (b), TEM images of raw MWCNTs (c) and MMWCNTs@Cs(I)-IIP (d).

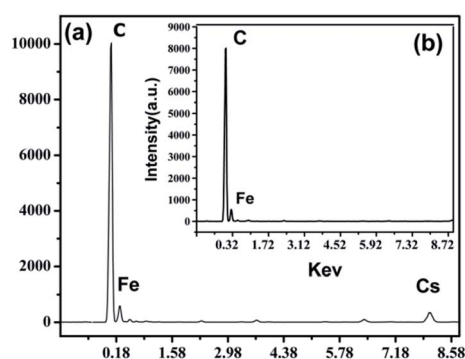


Fig. 4 EDS spectra of the unbleached MMWCNTs@Cs(I)-IIP (a) and the bleached MMWCNTs@Cs(I)-IIP (b).

and the diameter of the MWCNTs was obviously increased. The strong combination of  $\text{Fe}_3\text{O}_4$  nanoparticles and MWCNTs can be attributed to covalent interactions through functional groups on the surface of the modified MWCNTs. The MMWCNTs@Cs(I)-IIP exhibited a coagulated structure and a rough surface, which may be due to agglomeration on the surface of the polymer. As a result, the rough surface structure favors mass transfer and the formation of three-dimensional recognition sites. The results

demonstrate that the preparation process of the magnetic imprinted materials was successful.

EDS analysis is commonly performed to determine the elemental composition of materials. As shown in Fig. 4, C is the main element of the MWCNTs, and the appearance of a Fe signal indicates that the MWCNTs were successfully magnetized. The Cs signal was observed clearly in the unbleached MMWCNTs@Cs(I)-IIP sample (Fig. 4a) and disappeared after extraction (Fig. 4b). These results indicate that the Cs(I) ion was successfully imprinted into the polymer, and Cs(I) can be easily extracted.

### Electrochemical studies

The proposed ion-selective membrane consists of MMWCNTs@Cs(I)-IIP, additive (NaTPB), plasticizer (DBP or NPOE)<sup>38</sup> and PVC. In order to evaluate the effects of various components on the electrode efficiency, various kinds of GCEs coated membranes were fabricated, and their potentiometric responses in the presence of different concentrations of Cs(I) were examined. The results obtained are summarized in Table 1. The MMWCNTs@Cs(I)-IIP is necessary for a considerable sensitivity and Nernstian response of the electrode. The plasticizers strongly influenced the working concentration range of the GCEs. The membranes prepared based on the NPOE plasticizer exhibited better potentiometric characteristics in the ways of linear range and Nernstian slope. Moreover, the addition of NaTPB improved the Nernstian behavior of the electrode, which may be due to the ionic additive NaTPB reducing the anionic interference and the electrical resistance of the membrane. According to the results, the optimum potentiometric response was obtained with GCE 8 (20% PVC, 50% NPOE, 10% MMWCNTs@Cs(I)-IIP and 20% NaTPB), and this was selected as the optimum composition for the modified electrode.

### Effect of the solutions pH value

The effect of pH on the current response of the sensor was studied in the pH range of 4.0–8.5. As shown in Fig. 5, the current responses increased by increasing the pH from 4.0 to 6.5, and then decreased when the solution pH exceeded 6.5. The small current response at lower pH may be caused by the competitive adsorption of hydrogen ions, and the decrease of the current at high pH may be due to a decrease in the adsorption capacity of IIP under alkaline conditions.

Table 1 Compositions of the proposed sensor and their potentiometric response characteristics

No.	Composition (%)				Slope ( $\text{V dec}^{-1}$ )	Linear range ( $\text{mol L}^{-1}$ )	$R^2$
	PVC	Plasticizer	Cs-MIIP	NaTPB			
GCE1	20	NPOE (50)	—	20	0.02632	$1 \times 10^{-4}$ to $1 \times 10^{-1}$	0.98497
GCE 2	20	DBP (50)	—	20	0.01354	$1 \times 10^{-5}$ to $1 \times 10^{-1}$	0.94320
GCE 3	20	NPOE (50)	20	10	0.04667	$1 \times 10^{-4}$ to $1 \times 10^{-1}$	0.96916
GCE 4	20	DBP (50)	20	10	0.08469	$1 \times 10^{-6}$ to $1 \times 10^{-1}$	0.98987
GCE 5	10	DBP (50)	20	20	0.06618	$1 \times 10^{-7}$ to $1 \times 10^{-4}$	0.98772
GCE 6	10	NPOE (50)	20	20	0.04726	$1 \times 10^{-6}$ to $1 \times 10^{-1}$	0.97563
GCE 7	20	DBP (50)	10	20	0.03972	$1 \times 10^{-6}$ to $1 \times 10^{-1}$	0.99735
GCE 8	20	NPOE (50)	10	20	0.05904	$1 \times 10^{-7}$ to $1 \times 10^{-4}$	0.99725



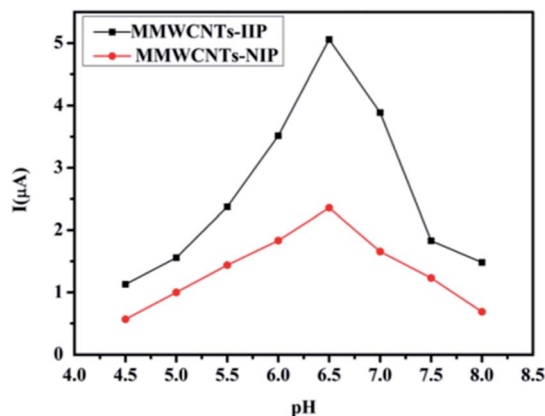


Fig. 5 Effect of pH on the current response of the optimized sensor toward  $1.0 \times 10^{-4}$  M Cs(I).

### Response time, reproducibility and lifetime of the sensor

The response time is the average time required for the electrodes to reach the potential response within  $\pm 1$  mV of the final equilibrium value. The dynamic potential difference response of the sensor is shown in Fig. 6, and the sensor had a short response time of less than 5 s. To demonstrate the reproducibility of the sensor, potential measurements were performed on Cs(I) solutions with concentration ranges from  $1 \times 10^{-7}$  to  $1 \times 10^{-4}$  M. The obtained real-time-dependent potential measurements are shown in Fig. 7. The sensor response was highly reproducible, and the proposed sensor functioned for 6 weeks without significant potential deviation.

### Analytical characteristics of the MMWCNTs@Cs(I)-IIP/GCE

A calibration curve of the MMWCNT@Cs(I)-IIP/GCE was constructed under the optimum conditions and is shown in Fig. 8. A linear relationship with a slope of  $0.05954 \text{ V dec}^{-1}$  was obtained between the response potential and the logarithm of the Cs(I) concentration over a range of  $1 \times 10^{-7}$  to  $1 \times 10^{-4}$  M. The linear regression equation is as follows:  $E(\text{V}) = 0.05954 \log[\text{Cs}^+] - 0.4804$  ( $R^2 = 0.999$ ). The limit of detection (DEC), as determined from the intersection of the two extrapolated segments of the calibration graph,<sup>39</sup> was  $4.0 \times 10^{-8}$  M. Due to the lack of

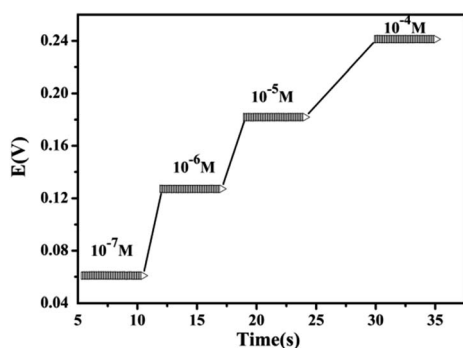


Fig. 6 Dynamic response of the MMWCNTs@Cs(I)-IIP/GCE with stepwise concentration changes of Cs(I).

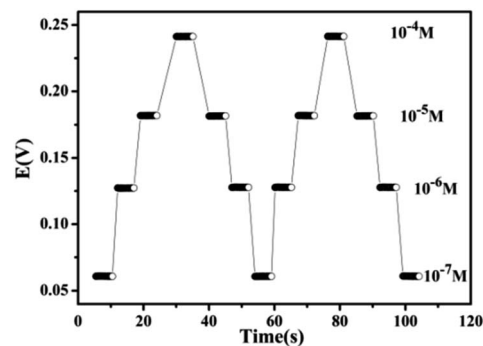


Fig. 7 Reproducibility of the MMWCNTs@Cs(I)-IIP/GCE.

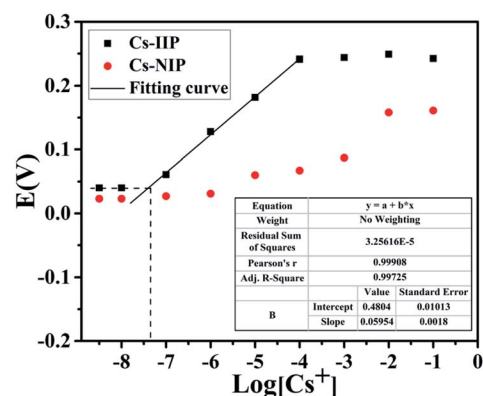


Fig. 8 Calibration curve of the developed MMWCNTs@Cs(I)-IIP/GCE and MMWCNTs@Cs(I)-NIP/GCE obtained at the optimized conditions.

specific Cs(I) recognition cavity in MNIP, MNIP/GCE does not have a good linear response and Nernst slope.

### Potentiometric selectivity

The selectivity coefficient ( $K$ ) is the ability of a given ion-selective electrode to distinguish between interfering ions and the target ion. The potentiometric selectivity coefficients of the GCE 8 for different cationic species were determined by the matched potential method (MPM), which is independent of the Nikolsky-Eisenman equation.<sup>40</sup> The selectivity coefficient values were calculated by formula (1):

$$K_{\text{CS,M}}^{\text{pot}} = \frac{a'_{\text{A}} - a_{\text{A}}}{a_{\text{B}}} \quad (1)$$

A specified activity (concentration) of the primary ion ( $a'_{\text{A}} = 1.0 \times 10^{-5}$  M for Cs(I)) was added to a reference solution ( $a_{\text{A}}$  of  $1.0 \times 10^{-6}$  M for Cs(I)), and the potential was measured.

Table 2 Potentiometric selectivity coefficients ( $K_{\text{CS,M}}^{\text{pot}}$ ) of the sensor

$M^{n+}K_{\text{CS,M}}^{\text{pot}}$	Cs-MIIP	Cs-NIP	$M^{n+}K_{\text{CS,M}}^{\text{pot}}$	Cs-IIP	Cs-NIP
$\text{Li}^+$	$1.0 \times 10^{-2}$	0.2	$\text{Mn}^{2+}$	$5.0 \times 10^{-3}$	$3.0 \times 10^{-2}$
$\text{Na}^+$	$8.0 \times 10^{-2}$	0.16	$\text{Fe}^{3+}$	$2.0 \times 10^{-3}$	$1.4 \times 10^{-2}$
$\text{K}^+$	$5.0 \times 10^{-2}$	0.13	$\text{Cu}^{2+}$	$2.0 \times 10^{-3}$	$6.3 \times 10^{-2}$
$\text{Rb}^+$	$4.0 \times 10^{-2}$	0.25	$\text{Mg}^{2+}$	$1.0 \times 10^{-3}$	$8.3 \times 10^{-2}$



Table 3 Comparison of the proposed sensor with various reported electrodes

Ionophore	Potentiometric behavior		$(-\log K_{\text{Cs,M}}^{\text{pot}})$		Application	Ref.
	Linear range	Detection limit (M)	K <sup>+</sup>	Na <sup>+</sup>		
2,3-Benzoquino-crown	$10^{-4}$ to $10^{-1}$	$2.5 \times 10^{-5}$	0.99	2.38	None	42
Calix[4]crown ether ester	$5 \times 10^{-6}$ to $10^{-1}$	$5.0 \times 10^{-6}$	2.00	1.30	None	43
5-(4-Nitrophenylazo)25,27-bis(2-propyloxy) <sup>26,28</sup>	$10^{-5}$ to $10^{-1}$	$4.6 \times 10^{-6}$	2.27	3.13	None	44
Thiacalix[4]biscrown-6,6	$10^{-6}$ to $3.6 \times 10^{-2}$	$3.8 \times 10^{-7}$	3.70	4.30	None	45
<i>P</i> -isopropylcalix[6]arene	$10^{-6}$ to $10^{-1}$	$1.0 \times 10^{-6}$	2.62	2.74	Ground waters	46
1,3-Cyclotetradecyloxy calix[4]arene crown ether	$10^{-7}$ to $10^{-2}$	$3.7 \times 10^{-8}$	2.13	4.08	Nuclear waste water streams	47
Cs-MIP (our previous publication)	$10^{-6}$ to $10^{-1}$	$3.0 \times 10^{-7}$	1.92	1.37	Tap water and brine	30
MMWCNTs@Cs(I)-IIP	$10^{-7}$ to $10^{-4}$	$4.0 \times 10^{-8}$	1.30	1.09	River; industrial waste water and brine	This work

Table 4 Determination of cesium in real samples and comparison with ICP-MS values

Sample	Spiked Cs <sup>+</sup> /(mol L <sup>-1</sup> )	Found <sup>a</sup> Cs <sup>+</sup> /(mol L <sup>-1</sup> )	RSD% ( <i>n</i> = 3)	ICP-MS/(mol L <sup>-1</sup> )	RSD% ( <i>n</i> = 3)	Recovery/(%)
Brine	—	Not detected	—	Not detected	—	—
	$3.0 \times 10^{-6}$	$3.34 \times 10^{-6}$	0.8	$3.17 \times 10^{-6}$	0.1	105.36
	$1.0 \times 10^{-5}$	$1.09 \times 10^{-5}$	0.7	$1.03 \times 10^{-5}$	0.3	105.82
Beichuan river water	—	Not detected	—	Not detected	—	—
	$3.0 \times 10^{-6}$	$2.88 \times 10^{-6}$	0.8	$2.93 \times 10^{-6}$	0.1	98.29
	$1.0 \times 10^{-4}$	$9.12 \times 10^{-5}$	0.6	$9.62 \times 10^{-5}$	0.5	94.80
Industrial waster water	—	Not detected	—	Not detected	—	—
	$3.0 \times 10^{-6}$	$3.69 \times 10^{-6}$	0.8	$3.51 \times 10^{-6}$	0.6	105.13
	$1.0 \times 10^{-5}$	$9.85 \times 10^{-6}$	0.7	$1.03 \times 10^{-5}$	0.3	95.63

<sup>a</sup> The average value of three determinations.

In a separate experiment, interfering ions ( $a_{\text{B}}$ ) were successively added to an identical reference solution, until the measured potential matched that obtained with the addition of the primary ions. The selectivity coefficient is given by the ratio of the resulting primary ion activity to the interfering ion activity (concentration).<sup>41</sup> The selectivity coefficients  $K_{\text{Cs,M}}^{\text{pot}}$  of the sensor for Cs(I) ions over other  $M^{n+}$  cations, obtained *via* the described method, are summarized in Table 2. The depicted results indicate that the sensor is noticeably selective to Cs(I) ions in the presence of most of the interfering ions tested. Therefore, the MMWCNTs@Cs(I)-IIP/GCE could be recommended for use in the determination of Cs(I) ions in samples with a complicated matrix.

### Comparison of the present sensor with previously reported Cs(I) sensors

We compared the properties of our sensor with some previously reported Cs(I) selective electrodes, as shown in Table 3. Although the linear range of our sensor was not as wide as those reported, the as-prepared sensor had a lower DEC and high selectivity. The good electrochemical performance may be attributed to the Cs(I)-IIP and the specific surface area of the MWCNTs.

### Real sample analysis

The utility of the sensor was also checked by applying the electrochemical method to measure the Cs(I) concentration in

real water samples. The contents of Cs(I) and natural mineral impurities in the samples were detected by ICP-MS, and the results are listed in Table S2.† MMWCNTs@Cs(I)-IIP/GCE was used as selective electrode to carry out standard recovery experiments on these samples. Meanwhile, ICP-MS was chosen as a comparison technique. The results obtained are shown in Table 4. The results showed that the recoveries of Cs(I) ranged from 94.80% to 105.82% with RSDs less than 1.0%, and no significant differences between the Cs(I) concentrations obtained by the developed electrode and those given by the ICP-MS method were found. This demonstrates that the MMWCNTs@Cs(I)-IIP/GCE sensor is useful for direct determination of Cs(I) in real samples.

## Conclusions

A simple, selective and sensitive ion-imprinted electrochemical sensor was developed for the rapid determination of Cs(I) in complex matrixes in the present work. The sensor was linearly dependent on cesium concentration over a range of  $1 \times 10^{-7}$  to  $1 \times 10^{-4}$  M with a detection limit of  $4.0 \times 10^{-8}$  M. The good sensitivity and lower limit of detection of the MMWCNTs@Cs(I)-IIP/GCE sensor toward cesium may be due to the synergistic effect caused by the IIP and MMWNCTs. The proposed imprinted sensor has been successfully applied for the detection of Cs(I) in complex real samples.



## Author contributions

Zhiming Wang: investigation and writing. Long Wang: investigation and writing. Cuo Zhou: investigation. Chunyan Sun: supervision.

## Conflicts of interest

There are no conflicts to declare.

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