


Cite this: *RSC Adv.*, 2021, 11, 11295

# Mechanistic insights into the C–H activation of methane mediated by the unsupported and silica-supported VO<sub>2</sub>OH and CrOOH: a DFT study†

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The direct activation and conversion of methane has been a topic of interest in both academia and industry for several decades. Deep understanding of the corresponding mechanism and reactivity mediated by diverse catalytic clusters, as well as the supporting materials, is still highly desired. In this work, the regulation mechanism of C–H bond activation of methane, mediated by the closed-shell VO<sub>2</sub>OH, the open-shell CrOOH, and their silica supported clusters, has been investigated by density functional theory (DFT) calculations. The hydrogen-atom transfer (HAT) reaction towards methane C–H bond activation is more feasible when mediated by the unsupported/silica-supported CrOOH clusters *versus* the VO<sub>2</sub>OH clusters, due to the intrinsic spin density located on the terminal O<sub>t</sub> atom. The proton-coupled electron transfer (PCET) pathways are regulated by both the nucleophilicity of the O<sub>t</sub> site and the electrophilicity of the metal center, which show no obvious difference in energy consumption among the four reactions examined. Moreover, the introduction of a silica support can lead to subtle influences on the intermolecular interaction between the CH<sub>4</sub> molecule and the catalyst cluster, as well as the thermodynamics of the methane C–H activation.

Received 23rd December 2020

Accepted 5th March 2021

DOI: 10.1039/d0ra10785a

rsc.li/rsc-advances

## 1. Introduction

Methane (CH<sub>4</sub>), as the major component of natural gas, is not only an important clean fuel, but also an abundant hydrocarbon feedstock for chemical commodities.<sup>1,2</sup> In the recent decades, with the development of modern catalytic technologies, the direct nonoxidative activation and conversion of methane to high value-added chemicals has received significant interest, from both industry and academia.<sup>3–6</sup> Although a variety of strategies have been developed, C–H bond activation, the first and also critical step for methane conversion is still challenging. This challenge arises from the fact that the extremely symmetric nonpolar methane molecule exhibits a high C–H bond dissociation energy (439 kJ mol<sup>−1</sup>).<sup>7,8</sup> To avoid the high temperature and harsh reaction conditions, efficient catalysis under mild conditions as well as the deep understanding of the

underlying mechanism for direct methane activation has always been highly desirable.<sup>9–12</sup>

Up to now, a great deal of effort has been devoted on the metal-cluster-mediated methane activation reactions, using both experimental and theoretical methods. These metal cluster species, including the metal oxide cluster,<sup>13,14</sup> metal carbides,<sup>15,16</sup> and metal cluster anions<sup>17,18</sup> *etc.*, can be regarded as simplified models of methane activation sites on heterogeneous catalysts. Meanwhile, the knowledge of the fundamental reaction mechanism towards the methane activation and conversion is continuously being improved.<sup>13–18</sup> For example, Sun *et al.* reported that the C–H bond activation may undergo a concerted pathway where the C and H atoms on the C–H bond attack two different sites, accompanied by forming two new bonds (C–Ni and O–H), or a stepwise pathway where two separated products of Ni<sub>3</sub>O<sub>3</sub>H and ethyl radical can be obtained.<sup>19</sup> Wan *et al.* found that the terminal oxygens [=O] are more active than the bridge oxygens [–O–] for C–H bond activation of methane on molybdenum oxides.<sup>20</sup> Interestingly, the adsorbed oxygen species on metal surfaces were demonstrated to exhibit promotional or inhibitory effects, which depends strongly on the nature of metals.<sup>21–26</sup> In more recent years, a series of mechanistic scenarios were explored by Schwarz *et al.*,<sup>13–16,27–29</sup> in which the classical hydrogen-atom transfer (HAT),<sup>27</sup> proton-coupled electron transfer (PCET),<sup>28</sup> and hydride transfer (HT)<sup>29</sup> pathways were explored and distinguished in various reaction systems for methane C–H activation.

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† Electronic supplementary information (ESI) available: The symmetry-adapted perturbation theory (SAPT) energy decomposition and Hirshfeld spin density and charge population analyses, and Cartesian coordinates of all optimized structures. See DOI: 10.1039/d0ra10785a

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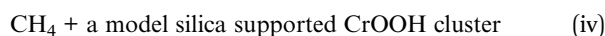


Although significant progress has been made in methane activation mediated by metal clusters, the reported gas-phase atom clusters obtained by state-of-the-art mass spectrometric are generally ionic species.<sup>13–20</sup> While, researches on the catalytic activity of the neutral metal oxyhydroxide towards C–H bond activation of methane are still sparse. Typically, the earth abundant vanadium and chromium have been demonstrated good catalytic capability for C–H activation,<sup>30–32</sup> but few studies have focused on their neutral oxyhydroxides, such as VO<sub>2</sub>OH and CrOOH. In addition, in industrial catalytic processes, catalyst supports are usually introduced to disperse and support active components, which can lead to a lower cost of the transition metal as well as a higher stability of the catalytic material compared with the unsupported ones.<sup>33</sup> However, detailed understanding of the support effect on catalyst activity, and the interaction between support and active components is yet to be more thoroughly investigated, especially at the molecular level.<sup>33,34</sup>

Herein, we report a theoretical study on C–H bond activation of methane mediated by the silica-supported/unsupported VO<sub>2</sub>OH and CrOOH clusters, through the density functional theory (DFT) calculations. The HAT and PCET reaction channels were carefully examined to obtain a better understanding towards the catalytic activity of the closed-shell VO<sub>2</sub>OH and the open-shell CrOOH, as well as the effect of silica support.

## 2. Computational methods

In the present study, DFT calculations were performed using the Gaussian 09 program package,<sup>35</sup> to investigate the reaction mechanism of methane activation catalyzed by the VO<sub>2</sub>OH and CrOOH catalysts, as well as their silica-supported complexes, as shown in the following equations:



In our calculations, a cluster containing Si<sub>4</sub>O<sub>6</sub>H<sub>3</sub> was used to model the silica support surface, denoted as “Sili” in this paper. All Si centers have a tetrahedron local structure simulating the SiO<sub>4</sub> tetrahedron moiety, and the H atoms were used to saturate the boundary dangling bonds of the Si centers. The VO<sub>3</sub>@Sili and CrO<sub>2</sub>@Sili were used to denote the complexes that the VO<sub>3</sub><sup>–</sup> and CrO<sub>2</sub><sup>–</sup> clusters supported on the silica model. In addition, a larger cluster models, Si<sub>16</sub>O<sub>30</sub>H<sub>3</sub>, denoted as “Sili2”, was used to investigate the cluster size effects (details can be seen from the Section 3.4), which showed that the smaller cluster model of Si<sub>4</sub>O<sub>6</sub>H<sub>3</sub> used in this article are as reliable as that with the larger one.

All minima of the reactants and the corresponding products for Rxn (i)–(iv) were first obtained by full system optimization, using the hybrid density functional<sup>36–38</sup> including empirical dispersion correction computed with Grimme's D3 formula (B3LYP-D3).<sup>39</sup> Note that the open-shell configurations were calculated within the

spin-unrestricted approximation, in which the broken-symmetry technique was introduced to treat the singlet (Rxn (i) and (ii)) and quartet (Rxn (iii) and (iv)) radical HAT pathway. The energy-consistent scalar-relativistic DF-adjusted 10-electron-core pseudopotential/ECP10MDF(8s7p6d1f/6s5p3d1f) basis set was used for V and Cr,<sup>40</sup> while the def-TZVP<sup>41,42</sup> basis set was applied for all other atoms. The transition states (TSs) for C–H bond activation of CH<sub>4</sub> mediated by V and Cr catalysts were obtained by employing the conventional approach of transition state optimization. Based on the optimized geometries, vibrational frequency analyses were carried out at the same level of theory to ensure that each stationary point truly represented a minimum, or a saddle point, and to obtain the thermal enthalpy corrections at 298.15 K and 1 atm. Intrinsic reaction coordinate (IRC) computations<sup>43,44</sup> were also conducted to confirm that the transition states connected to the appropriate reactants and products. Finally, the electronic energies were refined with the B3LYP-D3 functional and a larger basis set, in which the DF-adjusted 10-electron-core pseudopotential/ECP10MDF(8s7p6d2f1g)/[6s5p3d2f1g] was used for V and Cr,<sup>45</sup> while the ma-def2-TZVPP<sup>46,47</sup> basis set was used for all other atoms. To reduce the unaffordable computational cost on reactions mediated by the larger cluster model, Si<sub>16</sub>O<sub>30</sub>H<sub>3</sub>, the basis set for the reaction center was used as same as that smaller reaction models above, while a smaller def-SV basis set was applied to the remaining atoms away from the reaction center, the corresponding mixed basis set applied for the geometry optimization and frequency calculations is denoted as BSI, and that used for the refined single-point-energy calculation is denoted as BSII (see Fig. S1†).

The charge and spin density populations analyses were performed by the Hirshfeld method.<sup>48</sup> The spin-density isosurfaces (isovalue 0.02) were plotted using a multifunctional wavefunction analyzer (Multiwfn).<sup>49</sup> The symmetry-adapted perturbation theory (SAPT) energy decomposition analysis was carried out at the SAPT0/def2-TZVPP level, by using PSI4 program package.<sup>50</sup> The total SAPT0 interaction energy ( $\Delta E_{\text{int}}$ ) between the CH<sub>4</sub> and the catalyst cluster can be decomposed into contributions of the electrostatic ( $\Delta E_{\text{els}}$ ), exchange-repulsion ( $\Delta E_{\text{exc}}$ ), induction ( $\Delta E_{\text{ind}}$ ) and dispersion ( $\Delta E_{\text{disp}}$ ) terms:

$$\Delta E_{\text{tot}} = \Delta E_{\text{els}} + \Delta E_{\text{exc}} + \Delta E_{\text{ind}} + \Delta E_{\text{disp}}$$

The notations having the form of [R/TS/P]-*n* will be used in the later sections, where R/TS/P represents the reactant, transition state or product, and *n* represents one of (i)–(iv) for the Rxn numbers defined above. Meanwhile a superscript will be used before a species name to represent the spin multiplicity of this species. For example, the <sup>4</sup>TS<sub>HAT-iv</sub> denotes the quartet transition state of Rxn (iv) along the HAT pathway.

## 3. Results and discussion

### 3.1 Structures and properties of the unsupported/silica-supported VO<sub>2</sub>OH and CrOOH clusters

The stable structures, charge and spin density properties of the unsupported and silica-supported VO<sub>2</sub>OH and CrOOH clusters



were investigated at first. As shown in Fig. 1, the free  $\text{VO}_2\text{OH}$  has two terminal oxygen atoms ( $\text{O}_t$ ), which have the same V– $\text{O}_t$  bond length of 1.59 Å and equivalent charge of  $-0.31e$ . The free  $\text{CrOOH}$  have one terminal  $\text{O}_t$  atom with  $-0.36e$  charge and an O–H bond of length 1.59 Å. These calculations results show that both  $\text{VO}_2\text{OH}$  and  $\text{CrOOH}$  possess Lewis acid–base pair  $[\text{M}^{n+}(\text{V}^{5+}/\text{Cr}^{3+})-\text{O}^-]$  units,<sup>28</sup> facilitating the C–H bond activation through the PCET pathway. For the HAT mechanism for C–H activation, the  $\text{O}_t$ -center spin is necessary as the precursor state to abstract H atom.<sup>51,52</sup> Since the  $\text{VO}_2\text{OH}$  is closed-shell singlet, zero spin density were found on the  $\text{O}_t$  atom. By contrast, obvious spin density of  $-0.21$  was observed located on the  $\text{O}_t$  atom of the open-shell  $\text{CrOOH}$  cluster, which is reasonable due to the  $\text{Cr}^{3+}$  is initially a quartet, with three unpaired electrons in the 3d orbitals. In consequence, the  $\text{O}_t$  site on  $\text{CrOOH}$  is expected to be more advantageous for a radical HAT reaction; whereas, for the  $\text{O}_t$  site on the closed-shell  $\text{VO}_2\text{OH}$ , an electron reorganization is needed to accumulate O-center spin density to trigger the radical HAT process, indicating a higher energy barrier should be overcome.

In the presence of the silica support, the hydroxyl groups of both  $\text{VO}_2\text{OH}$  and  $\text{CrOOH}$  have to deprotonate to combine with the Si atom through chemisorption, forming a new Si–O bond (1.62 Å). In this condition, slight structural change and charge reorganization were observed from our computational results. As shown in Fig. 1, the  $\text{O}_t$  atom is slightly less negatively charged upon silica support for both  $\text{VO}_3@\text{Sili}$  and  $\text{CrO}_2@\text{Sili}$  clusters [ $\text{VO}_2\text{OH}$  ( $-0.31e$ )  $\rightarrow$   $\text{VO}_3@\text{Sili}$  ( $-0.29e$ );  $\text{CrOOH}$  ( $-0.36e$ )  $\rightarrow$   $\text{CrO}_2@\text{Sili}$  ( $-0.33e$ )]. Besides the charge-regulation effect, the spin density on the  $\text{O}_t$  atom of  $\text{CrOOH}$  is a little increased upon silica support [ $\text{CrOOH}$  ( $-0.21e$ )  $\rightarrow$   $\text{CrO}_2@\text{Sili}$  ( $-0.25e$ )]. Since the methane C–H activation reaction is sensitive to the charge and spin density of the  $\text{O}_t$  atom, the silica support may lead to subtle influences on the potential energy profiles (PESs) of methane activation on the  $\text{VO}_2\text{OH}$  and  $\text{CrOOH}$  clusters. Details about the mechanism of the reactions Rxn (i)–(iv) will be discussed in the Sections 3.2 and 3.3. It should be noted that, the free  $\text{CrOOH}$  in doublet state was calculated energetically  $16.4 \text{ kcal mol}^{-1}$  higher than that in quartet state. It suggests

that the electronic ground state is a quartet state with three parallel-spin electrons occupying three 3d orbitals ( $t_{2g}^3 e_g^0$ ). Thus, the doublet  $\text{CrOOH}$  were not considered in this work for the reactions Rxn (iii) and (iv).

### 3.2 Thermal activation of $\text{CH}_4$ on $\text{VO}_2\text{OH}$ clusters

#### 3.2.1 Rxn (i): reaction on an unsupported $\text{VO}_2\text{OH}$ cluster.

Fig. 2 shows the PESs of C–H activation of  $\text{CH}_4$  on a  $\text{VO}_2\text{OH}$  cluster (Rxn (i)) as well as the related critical structures. Once the separated reactants  $\text{CH}_4$  and  $\text{VO}_2\text{OH}$  encounter each other, an encounter reactant complex  $^1\text{EC-i}$  is formed with an energy decrease of  $13.1 \text{ kcal mol}^{-1}$ . The stabilization energy results mainly from the electrostatic and induction effects from the SAPT energy decomposition analysis (see Table S1 in the ESI†). The intermolecular interaction also leads to a charge transfer ( $0.21e$ ) from the  $\text{CH}_4$  moiety to the  $\text{VO}_2\text{OH}$  cluster (see Table S2 in the ESI†). Meanwhile, the  $\text{CH}_4$  molecule is slightly polarized, that two C–H bonds facing the  $\text{VO}_2\text{OH}$  cluster are stretched from 1.09 Å to 1.10–1.11 Å. In this case, the  $\text{CH}_4$  molecule is initially activated compare with its inherently nonpolar structure with four C–H bonds of equal length (1.09 Å) and zero dipole moment.

After  $^1\text{EC-i}$  is formed, one C–H bond of methane could be activated under two mechanisms in principle, that is, the HAT and the PCET mechanisms. As illustrated in Fig. 2, since the products of  $\text{CH}_4$  activation *via* a radical HAT mechanism are two separated radicals, the open-shell singlet channel and triplet channel were both examined for the HAT mechanism. Population analyses results reveal that the electron reorganization is necessary to access the transition states  $^1\text{TS}_{\text{HAT-i}}$  or  $^3\text{TS}_{\text{HAT-i}}$ , which finally generate the open-shell product  $^1\text{P}_{\text{HAT-i}}$ , consisting of a tetravalent vanadium hydroxide [ $\text{VO}(\text{OH})_2$ ] and a methyl radical  $\text{CH}_3^\cdot$ , through homolysis of C–H bond (see Fig. 2b and Table S2†). As discussed in Section 3.1, the  $\text{O}_t$  atom of  $\text{VO}_2\text{OH}$  have zero spin density, which is hard to abstract a hydrogen atom from  $\text{CH}_4$  through the radical HAT mechanism. Consistently, noticeable energy consumptions ( $38.5\text{--}42.5 \text{ kcal mol}^{-1}$ ) were found along the HAT pathways on both of the singlet and triplet PESs.

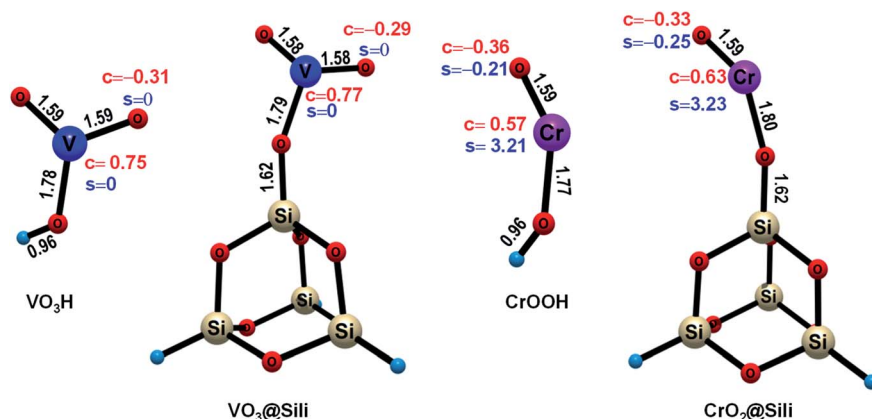


Fig. 1 The diagrams of the key bond lengths (black), the Hirshfeld charge ( $c$ , red) and spin density ( $s$ , blue) distributions for the catalyst clusters studied in this paper (the property parameters on only one side  $\text{O}_t$  of  $\text{VO}_2\text{OH}$  are labeled in this figure for clarity).

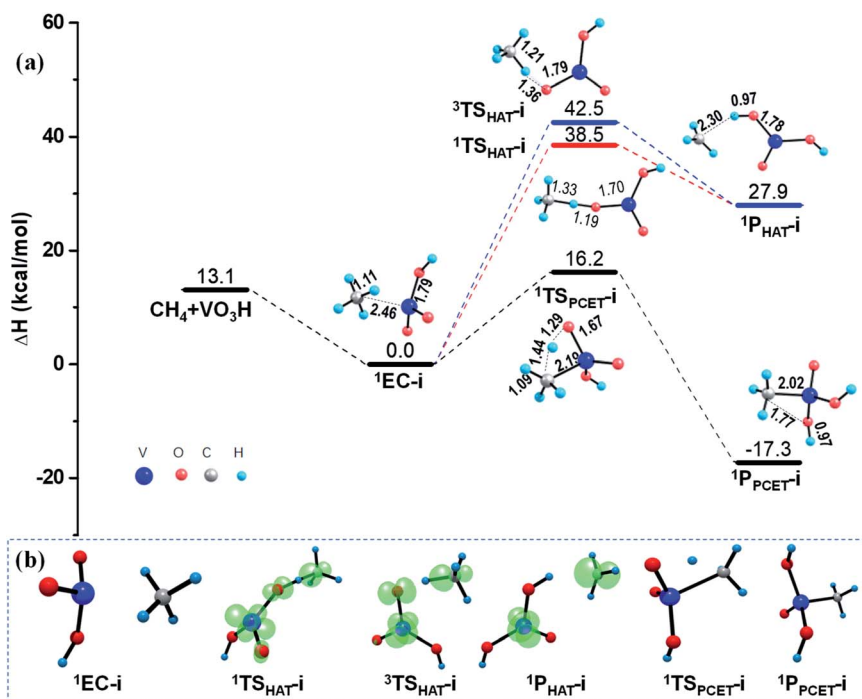


Fig. 2 (a) PESs ( $\Delta H$ , kcal mol<sup>-1</sup>) for the reaction of CH<sub>4</sub> with an unsupported VO<sub>2</sub>OH cluster, obtained at the B3LYP-D3/ma-def2-TZVPP//B3LYP-D3/TZVP level of theory. Critical structures are given with the key bond lengths in Å. (b) The evolution of the spin density distribution along the reaction pathway.

In contrast, the C–H activation of CH<sub>4</sub> is relatively easy to proceed along the PCET pathway mediated by VO<sub>2</sub>OH, by overcoming a moderate barrier of 16.4 kcal mol<sup>-1</sup>, as shown in Fig. 2. In this case, the H atom is transferred as a proton from the CH<sub>4</sub> moiety to the terminal O<sub>t</sub> atom of VO<sub>2</sub>OH, forming a new O–H bond; meanwhile the generated methyl anion CH<sub>3</sub><sup>-</sup> with an electron pair is transferred to the positive V center, forming a new V–C bond. This process eventually produces a relatively stable molecular structure (−17.3 kcal mol<sup>-1</sup>) of <sup>1</sup>P<sub>PCET-i</sub>, without change of spin density. These calculated results suggest that the PCET mechanism is predominant for the Rxn (i) mediated by a VO<sub>2</sub>OH cluster.

### 3.2.2 Rxn (ii): reaction on a silica-supported VO<sub>3</sub><sup>-</sup> cluster.

In the presence of silica support, the VO<sub>3</sub>@Sili can be formed through chemisorption of VO<sub>2</sub>OH cluster onto the silica model, and slight charge-regulation effect was observed in this case. To obtain deep understanding of the support effect on the methane C–H bond activation in Rxn (ii), we performed detailed examination of the PESs for the reaction of CH<sub>4</sub> with the silica supported VO<sub>3</sub><sup>-</sup> cluster. As shown in Fig. 3, an encounter complex <sup>1</sup>EC-ii is first formed when the CH<sub>4</sub> molecule approaches the VO<sub>3</sub>@Sili cluster, leading to an energy decrease of 20.1 kcal mol<sup>-1</sup> from the separated reactants. This means the <sup>1</sup>EC-ii is more stable than the <sup>1</sup>EC-i, due to the enhanced intermolecular interactions. More details can be seen from the results on SAPT energy decomposition analysis in Table S1.† Similar to the <sup>1</sup>EC-i, there is 0.22e charge transferred from the CH<sub>4</sub> moiety to the VO<sub>3</sub>@Sili cluster, leading to slightly polarization and initial activation of the methane molecule.

As illustrated in Fig. 3, the calculated results show negligible barrier changes of the HAT reaction paths in both singlet and triplet states, compared with the HAT reactions mediated by the free VO<sub>2</sub>OH cluster. This is reasonable since the VO<sub>2</sub>OH is intrinsically close-shell, and the spin density distribution is not affected in the presence of the silica support. While, for the PCET pathway, the silica support causes a larger energy barrier of 19.3 kcal mol<sup>-1</sup>, producing the <sup>1</sup>P<sub>PCET-ii</sub> with −15.6 kcal mol<sup>-1</sup> energy decrease. Though this cause a higher energy barrier than that in Rxn (i) from the point of view of kinetic features, both the HAT and PCET channels in Rxn (ii) are more advantageous thermodynamically due to the enhanced intermolecular interactions in the presence of silica support.

## 3.3 Thermal activation of CH<sub>4</sub> on CrOOH clusters

### 3.3.1 Rxn (iii): reaction on an unsupported CrOOH cluster.

Sections 3.2.1 and 3.2.2 present the comparisons of different reaction pathways on different PESs for the C–H activation reaction mediated by the unsupported and supported close-shell VO<sub>2</sub>OH clusters. In this section, we present further investigations on the thermal activation of C–H bond of methane mediated by the unsupported and silica-supported open-shell CrOOH clusters. As illustrated in Fig. 4, first, an encounter complex <sup>4</sup>EC-iii is formed when the separated reactants CH<sub>4</sub> and CrOOH collide with each other, leading to an energy decrease of 1.0 kcal mol<sup>-1</sup>, which is less stable compared with the <sup>1</sup>EC-i thermodynamically. It is consistent with the results of the SAPT energy decomposition analysis, that both the electrostatic and induction interactions of the <sup>4</sup>EC-iii are





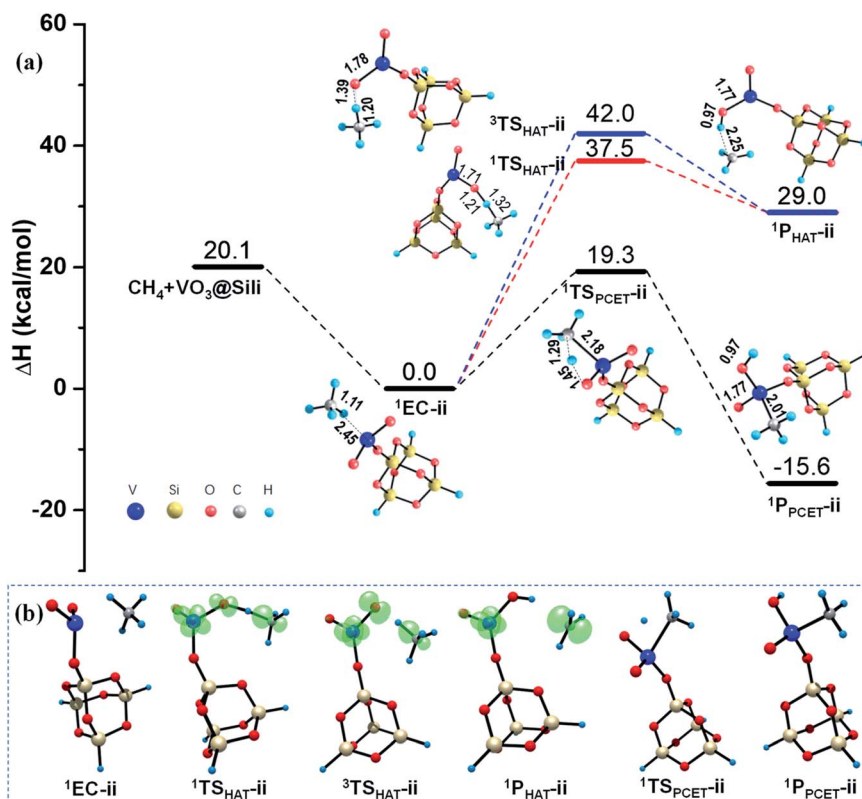


Fig. 3 (a) PESs ( $\Delta H$ , kcal mol $^{-1}$ ) for the reaction of  $\text{CH}_4$  with the silica supported  $\text{VO}_3^-$  cluster, obtained at the B3LYP-D3/ma-def2-TZVPP//B3LYP-D3/def-TZVP level of theory. Critical structures are given with the key bond lengths in Å. (b) The evolution of the spin density distribution along the reaction pathway.

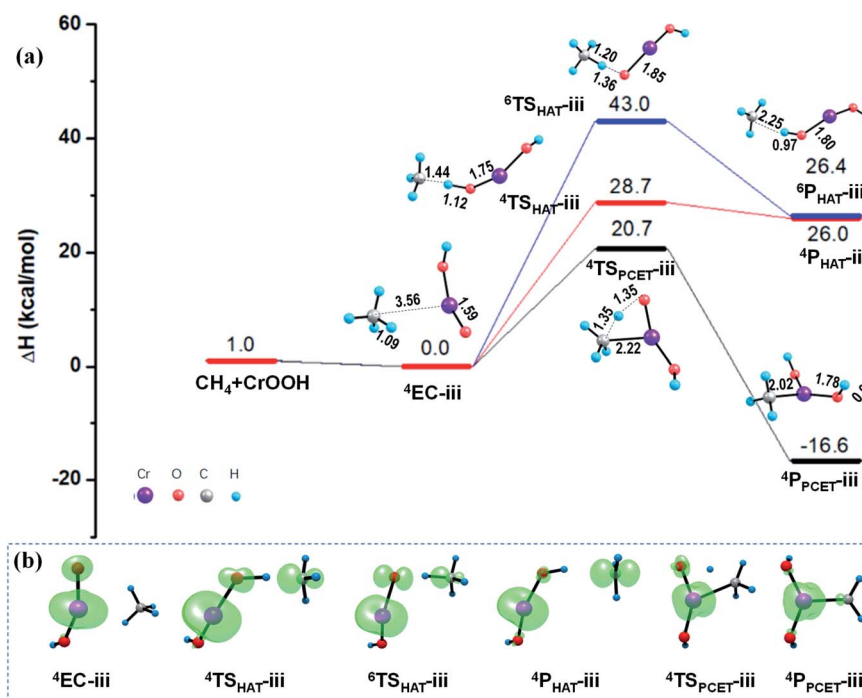


Fig. 4 (a) PESs ( $\Delta H$ , kcal mol $^{-1}$ ) for the reaction of  $\text{CH}_4$  with an unsupported  $\text{CrOOH}$  cluster, obtained at the B3LYP-D3/ma-def2-TZVPP//B3LYP-D3/def-TZVP level of theory. Critical structures are given with the key bond lengths in Å. (b) The evolution of the spin density distribution along the reaction pathway.

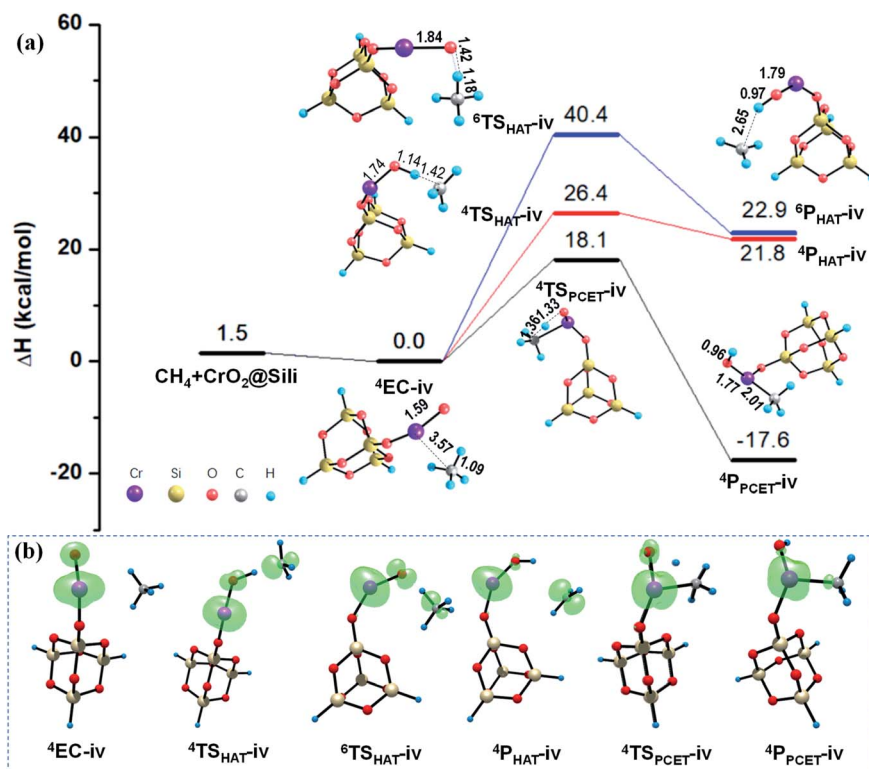


Fig. 5 (a) PESs ( $\Delta H$ , kcal mol<sup>-1</sup>) for the reaction of CH<sub>4</sub> with the silica supported CrO<sub>2</sub><sup>-</sup> cluster, obtained at the B3LYP-D3/ma-def2-TZVPP//B3LYP-D3/def-TZVP level of theory. Critical structures are given with the key bond lengths in Å. (b) The evolution of the spin density distribution along the reaction pathway.

significantly weaker than that of the <sup>1</sup>EC-i (see Table S1†). In consequence, only 0.01e is transferred from the CH<sub>4</sub> molecule to the CrOOH cluster (see Table S2†), and the CH<sub>4</sub> molecule is hardly polarized in this condition with four C-H bonds of almost equal bond length (1.09 Å). In line with expectations, a higher energy barrier of 20.7 kcal mol<sup>-1</sup> (compared with the PCET path in Rxn (i)) was found computationally along the PCET path in Rxn (iii) mediated by the free CrOOH cluster. This PCET process can be described as a heterolytic cleavage of C-H bond of the methane, in which the hydrogen atom is transferred to the O<sub>t</sub> site as a proton, and the CH<sub>3</sub> moiety is concertedly transferred to the Cr<sup>3+</sup> site as an anion (CH<sub>3</sub><sup>-</sup>). As thus, two new bonds (O<sub>t</sub>-H and Cr-C) are formed in the product P<sub>PCET</sub>, with an energy decrease of -16.6 kcal mol<sup>-1</sup>.

As mentioned in Section 3.1, unlike the PCET mechanism, the O<sub>t</sub>-center spin is necessary as a prepared site to abstract the H atom in the HAT mechanism. Different from the close-shell VO<sub>2</sub>OH cluster, a spin density of -0.21 was found located on the terminal O<sub>t</sub> atom of the quartet CrOOH cluster (see Fig. 4b and Table S2†). Correspondingly, the HAT reaction path in quartet state towards the methane C-H bond activation is more feasible with a smaller barrier of 28.7 kcal mol<sup>-1</sup>, compared with that in the Rxn (i) (38.5 kcal mol<sup>-1</sup>). The CH<sub>3</sub><sup>•</sup> is then released, and the H atom is transferred to the O<sub>t</sub> atom, forming a divalent chromium hydroxide [Cr(OH)<sub>2</sub>]. However, the HAT pathways on the sextet PESs in Rxn (iii) and (iv) are both suppressed severely by the energy consumption with more than

40.0 kcal mol<sup>-1</sup> barriers, due to significant electronic reorganizations are need to access the sextet transition states <sup>6</sup>TS<sub>HAT-iv</sub> and <sup>6</sup>TS<sub>HAT-iii</sub> (see Fig. 4b, 5b and Table S2†).

**3.3.2 Rxn (iv): reaction on a supported CrO<sub>2</sub><sup>-</sup> cluster.** To explore the silica support effect on the catalytic activity of the CrOOH cluster towards the methane C-H activation, the PESs of the Rxn (iv) were inspected as shown in Fig. 5. The intermolecular interaction was also found slightly enhanced in presence of the silica support, forming the encounter complex <sup>4</sup>EC-iv (see Table S1†). The calculated energy barrier is 18.1 kcal mol<sup>-1</sup> along the PCET pathway in Rxn (iv) (Fig. 5a), which is 2.6 kcal mol<sup>-1</sup> lower than that mediated by the free CrOOH cluster in Rxn (iii) (Fig. 4a). Though, the calculated negative charge located on the terminal O<sub>t</sub> atom of CrO<sub>2</sub>@Sili cluster is a little larger than that on the O<sub>t</sub> atom of the free CrOOH (CrOOH: -0.32e → CrO<sub>2</sub>@Sili: -0.36e), which is adverse to the proton transfer to the O<sub>t</sub> site, forming a new O<sub>t</sub>-H bond; while we found that the silica support lead to a more positive charged Cr center (CrOOH: 0.56e → CrO<sub>2</sub>@Sili: 0.62e), which is in favor of the attack of CH<sub>3</sub><sup>-</sup> to the Cr<sup>δ+</sup> site, resulting in a new Cr-C bond. Thus, it can be concluded from our computational results, that the PCET mechanism is not only related to the nucleophilicity of the O<sub>t</sub> site, but also associated with the electrophilicity of the metal center site, which is consistent with the regulation rule of Lewis acidity-basicity of the [M<sup>+</sup>-O<sup>-</sup>] unit reported by Schwarz *et al.*<sup>28</sup>



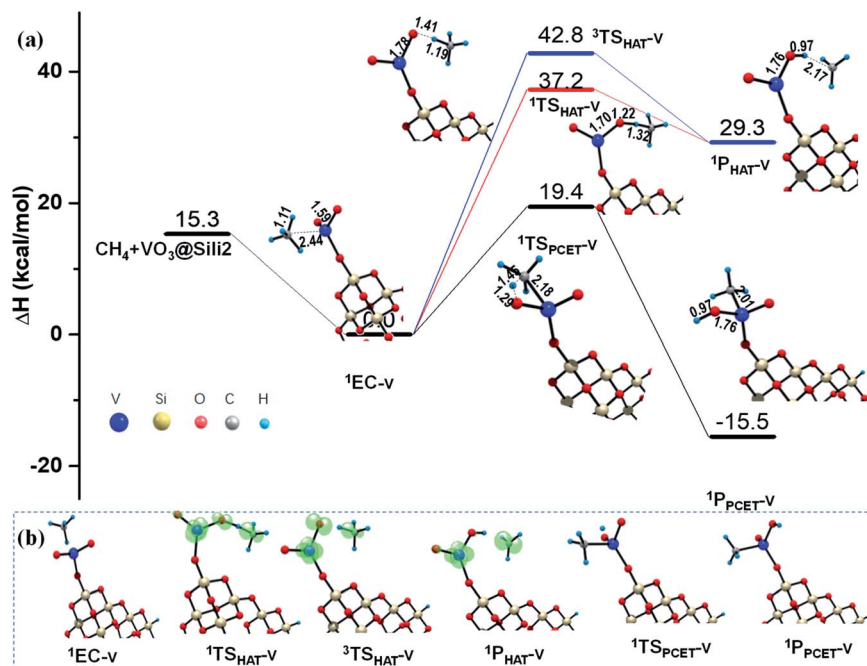


Fig. 6 (a) PESs ( $\Delta H$ , kcal mol<sup>-1</sup>) for the reaction of CH<sub>4</sub> with VO<sub>3</sub><sup>-</sup> cluster supported by the larger Si<sub>16</sub>O<sub>30</sub>H<sub>3</sub> cluster, obtained at the B3LYP-D3/BSII//B3LYP-D3/BSI level of theory. Critical structures are given with the key bond lengths in Å. (b) The evolution of the spin density distribution along the reaction pathway.

As described in Section 3.1, the spin density on the O<sub>t</sub> atom is slightly increased once the CrO<sub>2</sub>@Sili is formed through chemisorption of the CrOOH cluster onto the silica support. The larger spin density on O<sub>t</sub> atom means it is more facilitated to abstract a H atom from the CH<sub>4</sub> molecule mediated by the supported CrO<sub>2</sub><sup>-</sup> cluster *versus* that in the reaction Rxn (iii) mediated by the free CrOOH cluster. As expected, a smaller barrier (26.4 kcal mol<sup>-1</sup>) was found computationally along the singlet HAT pathway, which is 2.3 kcal mol<sup>-1</sup> lower than that under the HAT pathway mediated by the free CrOOH in Rxn (iii). Moreover, the difference of the energy barriers along HAT paths are significantly smaller mediated by open-shell CrOOH and CrO<sub>2</sub>@Sili (26.4–28.7 kcal mol<sup>-1</sup>), than that mediated by the close-shell VO<sub>2</sub>OH and VO<sub>3</sub>@Sili clusters (37.5–38.5 kcal mol<sup>-1</sup>), which is in line with the previous findings that the spin density at the hydrogen-acceptor site plays a crucial role for C–H bond activation proceeding by HAT mechanism. It is worth mentioning that, though the PCET mechanism was found more active than that of HAT mechanism in the four methane activation reactions (Rxn (i)–(iv)), it is not always true for different reaction systems.<sup>51–54</sup> For example, the HAT pathway was found more facilitated mediated by the cationic metal–carbon cluster FeC<sub>4</sub><sup>+</sup>, induced by localized spin density generated *in situ* once the reactants encounter each other.<sup>53</sup>

### 3.4 Rxn (v) vs. Rxn (ii): investigation of silica cluster size effects

The results shown in the above sections about the silica support are based a small cluster model (Sili) containing 4 Si atoms. As is known, the metal sites are always resided on a rough surface since catalytic supports are always mesoporous materials.

Furthermore, the particle size of metal or oxide clusters may possibly affect their catalytic activity, which depends on the localized or delocalized properties of the reactions toward the cluster surface.<sup>55,56</sup> In order to know whether the small model of Sili is representative for a catalytic support of silica, and also to understand as well the cluster size effect of the silica model used in this work toward the C–H bond activation of methane, a larger Si<sub>16</sub>O<sub>30</sub>H<sub>3</sub> cluster model was also examined for the reaction mediated by silica-supported VO<sub>3</sub><sup>-</sup> cluster (VO<sub>3</sub>@Sili2 in Rxn (v)). Rxn (v) differs with Rxn (ii) only in the model of the silica support.

As shown in Fig. 6, the differences of reaction barrier along the HAT reaction paths in both singlet and triplet states are within 0.6 kcal mol<sup>-1</sup>, predicted by using the two silica models (Si<sub>16</sub>O<sub>30</sub>H<sub>3</sub> in Rxn (v) vs. Si<sub>4</sub>O<sub>6</sub>H<sub>3</sub> in Rxn (ii)). The reaction barrier along the PCET pathways in Rxn (ii) and (v) show only a small difference of 0.1 kcal mol<sup>-1</sup>. The two products in Rxn (v) and (ii) also show similar thermodynamic stability with energy differences of 0.1–0.3 kcal mol<sup>-1</sup>. In addition, the changes of geometries along both the HAT and PCET pathways in Rxn (v) are very close to that in Rxn (ii), by using the larger and the smaller silica cluster models. In consequence, these calculated results show that the smaller cluster model of Si<sub>4</sub>O<sub>6</sub>H<sub>3</sub> is as reliable as the larger one of Si<sub>16</sub>O<sub>30</sub>H<sub>3</sub> at least for the current research in this work.

## 4. Conclusion

In this work, a systematic DFT computational work have been performed to provide the mechanism comparison regarding the C–H bond activation of methane through HAT and PCET



pathways mediated by the close-shell VO<sub>2</sub>OH, the open-shell CrOOH, as well as their silica supported clusters.

For the four reactions (Rxn (i)–(iv)) examined, the HAT reaction towards methane C–H bond activation is dominated by the spin density on the O<sub>t</sub> site, and the PCET reaction is code-termined by the nucleophilicity of the O<sub>t</sub> site and the electrophilicity of the metal center site. Benefiting from the obvious spin density on terminal O<sub>t</sub> atom, the calculated HAT reactions in quartet state mediated by the open-shell CrOOH and CrO<sub>2</sub>@Sili clusters are demonstrated more favorable than that in singlet state mediated by the close-shell VO<sub>2</sub>OH and VO<sub>3</sub>@Sili clusters. In contrast, there are relatively smaller differences of the energy barriers through PCET mechanism among the four reactions (Rxn (i)–(iv)). Moreover, all the HAT reactions examined are disadvantaged compared with the PCET pathways, note that the products obtained under the PCET mechanism should suffer further M(V, Cr)–C bond cleavage to generate the free CH<sub>3</sub><sup>•</sup> radical. In the presence of silica support, the intermolecular interaction between the CH<sub>4</sub> molecule and the catalyst cluster is enhanced, especially for the reactions mediated by VO<sub>2</sub>OH cluster, leading to the C–H bond activation of methane being more thermodynamically feasible.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

Support from the National Natural Science Foundation of China (21576291), and Shandong Province Natural Science Foundation (ZR2014BM002) is gratefully acknowledged.

## References

- 1 P. Tang, Q. Zhu, Z. Wu and D. Ma, *Energy Environ. Sci.*, 2014, **7**, 2580–2591.
- 2 A. Caballero and P. J. Pérez, *Chem. Soc. Rev.*, 2013, **42**, 8809–8820.
- 3 K. Otsuka and Y. Wang, *Appl. Catal., A*, 2001, **222**, 145–161.
- 4 R. Horn and R. Schloegl, *Catal. Lett.*, 2015, **145**, 23–39.
- 5 C. Karakaya and R. J. Kee, *Prog. Energy Combust. Sci.*, 2016, **55**, 60–97.
- 6 Y. Liu, D. Deng and X. Bao, *Chem*, 2020, **6**, 2497–2514.
- 7 S. J. Blanksby and G. B. Ellison, *Acc. Chem. Res.*, 2003, **36**, 255–263.
- 8 N. J. Gunsalus, A. Koppaka, S. H. Park, S. M. Bischof, B. G. Hashiguchi and R. A. Periana, *Chem. Rev.*, 2017, **117**, 8521–8573.
- 9 X. Meng, X. Cui, N. P. Rajan, L. Yu, D. Deng and X. Bao, *Chem*, 2019, **5**, 2296–2325.
- 10 U. Zavyalova, M. Holena, R. Schloegl and M. Baerns, *ChemCatChem*, 2011, **3**, 1935–1947.
- 11 H. Schwarz, *Isr. J. Chem.*, 2015, **54**, 1413–1431.
- 12 M. G. Quesne, F. Silveri, N. H. D. Leeuw and R. R. A. Catlow, *Front. Chem.*, 2019, **7**, 182.
- 13 H. Schwarz, S. Shaik and J. Li, *J. Am. Chem. Soc.*, 2017, **139**, 17201–17212.
- 14 Y. X. Zhao, X. N. Li, Z. Yuan, Q. Y. Liu, Q. Shi and S. G. He, *Chem. Sci.*, 2016, **7**, 4730–4735.
- 15 C. Geng, T. Weiske, J. Li, S. Shaik and H. Schwarz, *J. Am. Chem. Soc.*, 2018, **141**, 599–610.
- 16 H. F. Li, Z. Y. Li, Q. Y. Liu, X. N. Li, Y. X. Zhao and S. G. He, *J. Phys. Chem. Lett.*, 2015, **6**, 2287–2291.
- 17 B. A. Peter, *Chem.–Eur. J.*, 2017, **23**, 10–18.
- 18 Y. X. Zhao, Z. Y. Li, Y. Yang and S. G. He, *Acc. Chem. Res.*, 2018, **51**, 2603–2610.
- 19 X. Lin, Y. Xi and J. Sun, *J. Phys. Chem. C*, 2012, **116**, 3503–3516.
- 20 G. Fu, X. Xu, X. Lu and H. Wan, *J. Am. Chem. Soc.*, 2005, **127**, 3989.
- 21 B. Xing, X. Y. Pang and G. C. Wang, *J. Catal.*, 2011, **282**, 74–82.
- 22 B. Xing and G. C. Wang, *Phys. Chem. Chem. Phys.*, 2014, **16**, 2621–2629.
- 23 Y. Q. Wang, C. Q. Lv and G. C. Wang, *RSC Adv.*, 2015, **5**, 66221–66230.
- 24 J. J. Varghese, Q. T. Trinh and S. H. Mushrif, *Catal. Sci. Technol.*, 2016, **6**, 3984–3996.
- 25 D. Hibbitts and M. Neurock, *Surf. Sci.*, 2016, **650**, 210–220.
- 26 J. Wang and G. C. Wang, *J. Phys. Chem. C*, 2018, **122**, 17338–17346.
- 27 N. Dietl, M. Schlangen and H. Schwarz, *Angew. Chem., Int. Ed.*, 2012, **51**, 5544–5555.
- 28 J. Li, S. Zhou, J. Zhang, M. Schlangen, D. Usharani, S. Shaik and H. Schwarz, *J. Am. Chem. Soc.*, 2016, **138**, 11368–11377.
- 29 J. L. Li, S. D. Zhou, M. Schlangen, T. Weiske and H. Schwarz, *Angew. Chem., Int. Ed.*, 2016, **128**, 13266–13269.
- 30 S. Feyel, J. Dobler, D. Schroder, J. Sauer and H. Schwarz, *Angew. Chem., Int. Ed.*, 2006, **45**, 4681–4685.
- 31 R. C. Bell and A. W. Castleman Jr, *J. Phys. Chem. A*, 2002, **106**, 9893–9899.
- 32 X. Xu, F. Faglioni and W. A. Goddard III, *J. Phys. Chem. A*, 2002, **106**, 7171–7176.
- 33 P. Serna and B. C. Gates, *Acc. Chem. Res.*, 2014, **47**, 2612–2620.
- 34 Y. Xi, B. Chen, X. Lin, H. Fu and C. Wang, *Comput. Theor. Chem.*, 2016, **1076**, 65–73.
- 35 M. J. Frisch, *et al. Gaussian 09, Revision B. 01*, Gaussian, Inc., Wallingford, 2010.
- 36 A. D. Becke and D. Axel, *J. Chem. Phys.*, 1993, **98**, 1372–1377.
- 37 C. Lee, W. Yang and R. G. Parr, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1988, **37**, 785–789.
- 38 P. J. Stephens, F. J. Devlin, C. F. Chabalowski and M. J. Frisch, *J. Phys. Chem.*, 1994, **98**, 247–257.
- 39 S. Grimme, J. Antony, S. Ehrlich and H. Krieg, *J. Chem. Phys.*, 2010, **132**, 154104.
- 40 M. Dolg, U. Wedig, H. Stoll and H. Preuss, *J. Chem. Phys.*, 1987, **86**, 866–872.
- 41 A. Schäfer, C. Huber and R. Ahlrichs, *J. Chem. Phys.*, 1994, **100**, 5829–5835.
- 42 K. Eichkorn, F. Weigend, O. Treutler and R. Ahlrichs, *Theor. Chem. Acc.*, 1997, **97**, 119–124.





- 43 K. Fukui, *Acc. Chem. Res.*, 1981, **14**, 363–368.
- 44 H. P. Hratchian and H. B. Schlegel, *J. Chem. Phys.*, 2004, **120**, 9918–9924.
- 45 J. M. L. Martin and A. Sundermann, *J. Chem. Phys.*, 2001, **114**, 3408.
- 46 F. Weigend and R. Ahlrichs, *Phys. Chem. Chem. Phys.*, 2005, **7**, 3297–3305.
- 47 J. Zheng, X. Xu and D. G. Truhlar, *Theor. Chem. Acc.*, 2011, **128**, 295–305.
- 48 F. L. Hirshfeld, *Theor. Chim. Acta*, 1977, **44**, 129–138.
- 49 T. Lu and F. Chen, *J. Comput. Chem.*, 2012, **33**, 580–592.
- 50 J. M. Turney, A. C. Simmonett, R. M. Parrish, E. G. Hohenstein, F. A. Evangelista, J. T. Fermann, B. J. Mintz, L. A. Burns, J. J. Wilke and M. L. Abrams, *Wiley Interdiscip. Rev.: Comput. Mol. Sci.*, 2012, **2**, 556–565.
- 51 W. Lai, C. Li, P. D. H. Chen and P. D. S. Shaik, *Angew. Chem., Int. Ed.*, 2012, **51**, 5556–5578.
- 52 S. Ye and F. Neese, *Proc. Natl. Acad. Sci. U. S. A.*, 2011, **108**, 1228–1233.
- 53 C. Geng, J. Li, T. Weiske and H. Schwarz, *Chem.–Eur. J.*, 2019, **25**, 12940–12945.
- 54 X.-L. Ding, X.-N. Wu, Y.-X. Zhao and S.-G. He, *Acc. Chem. Res.*, 2012, **45**, 382–390.
- 55 C. S. Cooper, R. J. Oldmana and C. R. A. Catlowb, *Chem. Commun.*, 2015, **51**, 5856–5859.
- 56 Z. Hu and C. H. Turner, *J. Am. Chem. Soc.*, 2007, **129**, 3863–3878.

