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Two-dimension on two-dimension growth: hierarchical Ni_{0.2}Mo_{0.8}N/Fe-doped Ni₃N nanosheet array for overall water splitting⁺

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Developing advanced electrocatalysts with low cost for electrocatalytic water splitting are highly desirable. Herein, we report the design of two-dimension on two-dimension growth of hierarchical Ni_{0.2}Mo_{0.8}N nanosheets on Fe-doped Ni₃N nanosheets supported on Ni foam (Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF) *via* hydrothermal reaction and nitridation treatment. In the hierarchical structures, small Ni_{0.2}Mo_{0.8}N nanosheets were uniformly anchored on Fe-Ni₃N nanosheets. Due to enhanced electron transfer between Ni_{0.2}Mo_{0.8}N and Fe-Ni₃N, Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF exhibits superior electrocatalytic activity for the oxygen evolution reaction (OER) and the hydrogen evolution reaction (HER). After stability tests for 50 h, Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF exhibits negligible degradation of the current density for the OER (91% remain) and HER (95% remain), suggesting excellent stability. Owing to the outstanding performance, Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF display a cell voltage of 1.54 V (10 mA cm⁻²) for electrocatalytic overall water splitting.

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Introduction

Hydrogen (H_2) is a promising and attractive energy source owing to its high combustion efficiency, non-polluting production and sustainability.1-3 Currently, the industrial methods to produce H₂ usually consume a lot of fossil energy, and are unfavorable for solving the problems of energy and the environment.4-6 Thus, production of H₂ through efficient and facile methods are highly desirable. Electrocatalytic water splitting is an exciting strategy to sustainably produce hydrogen energy.7-10 Electrochemical water splitting consists of two half-reaction steps, the oxygen evolution reaction (OER) and the hydrogen evolution reaction (HER), requiring higher overpotentials to obtain significant catalytic efficiency due to the existence of inherent energy barriers.11-14 Therefore, electrocatalysts are crucial to reduce the energy barrier of rate-determining steps. Pt-based noble metals are the best catalysts towards the OER and HER.15-18 However, their lower abundance and higher cost greatly suppress their wide commercial development. Therefore, exploring highly efficient and earth-abundant electrocatalysts toward water splitting are vital for large-scale practical application.

Nowadays, transition metal-based materials¹⁹⁻²¹ including the chalcogenides,²²⁻²⁴ phosphides,²⁵⁻²⁷ nitrides,²⁸⁻³⁰ carbides³¹⁻³³ and selenides³⁴⁻³⁶ have been extensively applied for electrocatalysis.

Among of these electrocatalysts, transition metal nitrides (TMNs) are the new emerging alternative electrocatalyst for water splitting due to the high electrical conductivity and corrosion-resistance.³⁷⁻³⁹ However, amounts of electrocatalysts with highly performance are only suitable for OER or HER. Therefore, developing outstanding bifunctional electrocatalysts with excellent activity for OER and HER is crucial. Integrating multiple metal into one integrated material provide a powerful way to turn the electronic structure, further improving the OER and HER activity.⁴⁰⁻⁴³

Herein, we reported the design of two-dimension on twodimension growth of hierarchical $Ni_{0.2}Mo_{0.8}N$ nanosheets on Fedoped Ni_3N nanosheets supported on Ni foam ($Ni_{0.2}Mo_{0.8}N$ /Fe- Ni_3N /NF). In the hierarchical structures, small $Ni_{0.2}Mo_{0.8}N$ nanosheets were homogenously distributed on Fe-Ni₃N nanosheets. The $Ni_{0.2}Mo_{0.8}N$ /Fe- Ni_3N /NF exhibits lowest OER overpotential of 266 mV (20 mA cm⁻²), compared with the Fe- Ni_3N /NF (292 mV), $Ni_{0.2}Mo_{0.8}N$ /Fe- Ni_3N /NF shows negligible degradation of the current density (91%) after OER stability test for 50 h. In addition, the $Ni_{0.2}Mo_{0.8}N$ /Fe- Ni_3N /NF also displays the lower overpotential (40 mV, 20 mA cm⁻²) and excellent durability (50 h, 95% current density remain). The $Ni_{0.2}Mo_{0.8}N$ /Fe- Ni_3N /NF shows a quite low cell voltage of 1.54 V (10 mA cm⁻²) for overall water splitting.

Experimental

Chemicals

Iron nitrate nonahydrate (Fe(NO₃)₃·9H₂O), sodium molybdate (Na₂(MoO₄)₂·H₂O), urea, ammonium fluoride (NH₄F), nickel

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nitrate hexahydrate (Ni(NO₃)₂· $6H_2O$), ethanol, and ruthenium dioxide (RuO₂) were obtained from Sinopharm Chemical Reagent Co. Ltd. Potassium hydroxide (KOH) and Commercial 20% Pt/C were purchased from Shanghai Macklin Biochemical Co. Ltd. The 5 wt% Nafion solution was obtained from Sigma Aldrich. Ni foam (NF) was brought from Kunshan Rongsheng company, China. Unless otherwise stated, all reagents were used without purification.

Synthesis of NiMoO₄ and Ni_{0.2}Mo_{0.8}N grown on NF (NiMoO₄/ NF and Ni_{0.2}Mo_{0.8}N/NF)

The NiMoO₄ supported on NF were synthesized by the hydrothermal reaction. 2 mmol Ni(NO₃)₂·6H₂O and 2 mmol Na₂(-MoO₄)₂·H₂O were dissolved in 17.5 mL distilled water, respectively. Then, the mixed solution were transfer to the Teflon-lined stainless steel autoclave and heated at 150 °C for 6 h. Then, after cooling down to 25 °C, the as-prepared NiMoO₄/ NF materials were cleaned by the distilled water and keep ultrasonicate for 30 min. Furthermore, the Ni_{0.2}Mo_{0.8}N/NF were prepared by the following nitridation treatment.

The dry NiMoO₄/NF was placed at furnace and treated in a flow of 20 standard cubic centimeters (sccm) NH₃ and 130 sccm Ar. The NiMoO₄/NF kept at 500 °C for 1 h (5 °C min⁻¹). Ni_{0.2}Mo_{0.8}N/NF was obtained after the treatment.

Synthesis of Fe-doped $Ni(OH)_2$ precursor grown on NF and Fe-doped Ni_3N grown on NF

Fe-doped Ni(OH)2 precursor grown on Ni foam (Fe-Ni(OH)2/NF) was obtained to prepare the Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF. Before starting the reaction, the NF $(2 \times 3 \text{ cm}^2)$ was pre-treated with HCl solution (1 M) to get rid of the oxidation layer on the NF. After 20 min, the NF was washed by distilled water. 0.2 mmol Fe(NO₃)₃·9H₂O, 3.8 mmol Ni(NO₃)₂·6H₂O, 10 mmol urea and 20 mL distilled water were mixed to form the homogeneous solution by a magnetic stirrer. The 4 mmol NH₄F and as-treated NF was immersed in the precursor solution and keep stirring for 30 min. Then the Teflon-lined stainless steel autoclave was used as the container of preparing the Fe-Ni(OH)₂/NF and heated at 120 °C for 14 h. After cooling down to 25 °C, the Fe-Ni(OH)₂/NF was cleaned by ultrasonicating with distilled water for 30 min and dried at 60 °C. Then the dry Fe-Ni(OH)₂/NF was placed at furnace and treated in a flow of 20 standard cubic centimeters (sccm) NH₃ and 130 sccm Ar. The Fe-Ni(OH)₂/NF kept at 500 °C for 1 h (5 $^{\circ}$ C min⁻¹). Fe–Ni₃N/NF was obtained after the treatment.

Fabrication of Ni_{0.2}Mo_{0.8}N/Fe-doped Ni₃N nanosheets arrays supported on NF and Ni_{0.2}Mo_{0.8}N nanosheets arrays supported on NF

The Ni_{0.2}Mo_{0.8}N/Fe-doped Ni₃N arrays supported on NF (Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF) were synthesized through the combination of hydrothermal reaction and nitridation treatment. 2 mmol Ni(NO₃)₂·6H₂O and 2 mmol Na₂(MoO₄)₂·H₂O were dissolved in 17.5 mL distilled water, respectively. The Fe-Ni(OH)₂/NF were immersed in mixed solution containing Na₂(MoO₄)₂·H₂O and Ni(NO₃)₂·6H₂O. After 30 min, the Fe-Ni(OH)₂/NF and mixed solution were transfer to the Teflon-lined stainless

steel autoclave and heated at 150 °C for 6 h. Then, after cooling down to 25 °C, the as-prepared Fe–Ni(OH)₂/NiMoO₄/NF materials were cleaned by the distilled water and keep ultrasonicate for 30 min. The Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF was obtained through the NH₃ treatment of Fe–Ni(OH)₂/NiMoO₄/NF.

Characterization methods

The morphologies of the as-synthesized materials were examined by field emission scanning electron microscope (FE-SEM, S-4800, Hitachi, Japan) and transmission electron microscope (TEM, JEM-2010plus, JEOL, Japan). The phase structures were recorded by X-ray diffractometer (D8 Discover, Bruker AXS, Germany) with a Cu K_{\alpha} radiation ($\lambda = 1.5406$ Å). Chemical states and composition of the as-synthesized materials were investigated by X-ray photoelectron spectroscopy (XPS, Axis supra, Kratos Analytical Ltd, UK). The results were obtained at 15 kV and 10 mA with monochromatic Al K_{\alpha} radiation. All the XPS spectra of materials were corrected through C 1s spectrum at 284.8 eV.

Electrochemical measurement

The electrocatalytic performances of as-synthesized materials were performed by a CHI660E electrochemical workstation (Chen Hua, Shanghai, China). All the electrochemical tests were operated in 1 M KOH solution. The three-electrode system was used to perform the OER and HER tests. The as-prepared materials (geometric area: 0.5 cm²) and a carbon rod were used as working electrode and counter electrode, respectively. Furthermore, the reference electrode was a saturated calomel electrode (SCE). The potential was corrected by the following equation: E(RHE) = E(SCE) + 0.059 pH + 0.244.

The linear sweep voltammogram (LSV) for OER was measured in a potential range of 1.0 to 1.6 V versus RHE (vs. RHE). For HER measurements, the LSV was measured in a potential range of and 0 to -0.5 V vs. RHE. The scan rate is 1 mV s^{-1} . OER and HER curves of were treated by iR correction. According to the equation: $E_{\text{corrected}} = E_{\text{uncorrected}} - iR_{\text{s}}$, the curve of *iR* correction would be finished. R_s (solution resistance) was able to observed by the electrochemical impedance spectroscopy (EIS) tests. The EIS for OER and HER was tested at 1.47 V and -0.03 V vs. RHE from 10 kHz to 0.01 Hz with a 10 mV AC amplitude. The stability test for OER was operated at 1.515 V vs. RHE for 50 h. The stability test for HER was finished at -74 mV vs. RHE for 50 h. The double-layer capacitance (C_{dl}) was obtained by the CVs cycling tests from 0.1 to 0.2 V vs. RHE with scan rates ranged from 10-60 mV s⁻¹. C_{dl} can be used to calculate the electrochemically active surface area (ECSA). Overall water splitting was measured in a two-electrode electrolyzer. The LSV was measured in a potential range of 1.0 to 1.9 V voltage. The stability test of overall water-splitting was operated at 1.59 V voltage for 20 h.

Results and discussion

As shown in Fig. 1a–c, after the hydrothermal reaction, the Fe– $Ni(OH)_2$ nanosheets were densely and vertically grown on substrate Ni foam (NF), forming a unique three-dimensional



Fig. 1 FE-SEM images of $(a-c) Fe-Ni(OH)_2/NF$, $(d-f) Fe-Ni_3N/NF$, $(g-i) Ni_{0.2}Mo_{0.8}N/NF$ and $(j-l) Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$ at different magnifications.

(3D) architectures. The thickness of the Fe–Ni(OH)₂ nanosheets ranged from 80–200 nm (Fig. S1†). After the NH₃ treatment at 500 °C, the Fe–Ni(OH)₂ were converted into Fe-doped Ni₃N (Fe– Ni₃N). The Fe–Ni(OH)₂ nanosheets were transformed into threedimensional porous network structures consist of interconnected nanospheres (Fig. 1d–f). The average diameters of Fe–Ni₃N grains ranged from 160–320 nm (Fig. S1†). As shown in Fig. S2a–c,† the NiMoO₄ precursor nanowires were synthesized on NF through the same hydrothermal reaction. After the NH₃ treatment, the as-prepared Ni_{0.2}Mo_{0.8}N/NF also indicates the nanowire morphologies and structures (Fig. 1i).

The diameters of the Ni_{0.2}Mo_{0.8}N nanowires ranged from 80 to 200 nm (Fig. S1[†]). When NiMoO₄ meets Fe-Ni(OH)₂ (Fig. S2d-f^{\dagger}), the as-synthesized NiMoO₄/Fe-Ni(OH)₂/NF exhibited a hierarchical structure consist of $Fe-Ni(OH)_2$ as substrates and NiMoO₄ vertically grown on Fe-Ni(OH)₂ nanosheets. Through the NH₃ treatment at 500 °C, the NiMoO₄/Fe-Ni(OH)₂/NF were converted into Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF (Fig. 1j-l) with two unique phases. The Fe-Ni₃N nanosheets were coated by small Ni_{0.2}Mo_{0.8}N nanosheets. It is interesting that with the $Ni_{0.2}Mo_{0.8}N$, the Fe-Ni₃N substrates still remain the morphology of NiMoO₄/Fe-Ni(OH)₂. The thickness of Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF ranged from 210-350 nm (Fig. S3a and b[†]), while the thickness range of small nanosheet is 7-15 nm (Fig. S3c and d[†]). Different from the individual Fe-Ni₃N (Fig. 1d-f), Fe-Ni₃N substrates were not also transformed into the similar structure. The small Ni_{0.2}Mo_{0.8}N nanosheets could hinder the collapse of Fe-Ni₃N from nanosheets to interconnected nanospheres. The unique structure of Ni_{0.2}Mo_{0.8}N/ Fe-Ni₃N/NF may be provide more active sites for electrocatalysis.

The crystal structures of materials were investigated by the XRD characterization. Fig. 2a and S4[†] show the sharp peaks near the 44.5°, 51.8° and 76.4°, corresponding to the metal Ni in the NF. The peaks of Fe–Ni(OH)₂/NF were observed at 33.5°, 34.4° and 38.8° (Fig. S4[†]), indicating the formation of Ni(OH)₂ phase (JCPDS 38-0715). The peaks at 38.9°, 42.1°, 58.5°, 70.6°



Fig. 2 (a) XRD patterns of $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$, $Ni_{0.2}Mo_{0.8}N/NF$ and $Fe-Ni_3N/NF$. (b) HRTEM image of $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$. (c) STEM-EDS mapping images of $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$.

and 78.4° of Fe-Ni₃N/NF (Fig. 2a) were assigned to Ni₃N phases (JCPDS 10-0280), indicating that the Ni(OH)₂/NF were transformed to Fe-Ni₃N phases after NH₃ treatment. The visible diffraction peaks of Ni_{0.2}Mo_{0.8}N/NF at 32.2°, 36.5° and 49.4° belong to the crystal planes of Ni_{0.2}Mo_{0.8}N (JCPDS 29-0931). For the Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF, the peaks for Ni₃N phases were very weak due to the low crystallinity, while the peaks for Ni_{0.2}Mo_{0.8}N phases are strong. The HRTEM image of Fe–Ni₃N/ NF reveals a lattice fringe (0.20 nm) (Fig. S5a and b⁺), corresponding to the (100) plane of Ni_3N . The TEM image of Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF (Fig. S5c⁺) indicates that small Ni_{0.2}Mo_{0.8}N nanosheets were uniformly dispersed on Fe-Ni₃N/ NF nanosheets. Fig. 2b shows the HRTEM image of Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF. The (100) and (101) planes of Ni_{0.2}Mo_{0.8}N was observed. Furthermore, the (002) plane (the interplanar distance of 0.21 nm) of Ni₃N was observed in Fig. 2b. There were no peaks of metal Fe or Fe oxides and nitrides in the measured XRD patterns and HRTEM images, indicating that the Fe elements were doped into the Ni(OH)₂ and the Ni₃N crystal matrix.44-46 The STEM-EDS mapping images of Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF were represented in Fig. 2c. As shown in new Fig. 2c, the Ni, Mo, Fe and N were uniformly distributed throughout the randomly examined Ni_{0.2}Mo_{0.8}N/ Fe-Ni₃N nanosheets. The Fe, Ni and N elements can be visible observed as substrates while the Ni, Mo and N elements also distributed as spots. The results confirmed the formation of hierarchical structures of Ni0.2Mo0.8N/Fe-Ni3N/NF using Fe-Ni₃N as support.

XPS spectra were shown in Fig. S6[†] and 3 to research the chemical compositions and surface chemical states of materials. Fig. S6[†] is the XPS survey of $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$, indicating that the Ni, Fe, Mo and N were observed. In Fig. 3a, the two peaks of the $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$ at 852.6 and 870.2 eV belong to the Ni species in Ni–N bonds.⁴⁷⁻⁴⁹ Furthermore, the peaks located at 855.6, 873.5 eV and 857.1, 876.3 eV correspond to the oxidized Ni²⁺ and Ni³⁺ species, while the remaining two peaks (880.1 and 861.8 eV) belong to satellite peaks.^{50,51} Compared with the BE of $Ni_{0.2}Mo_{0.8}N/NF$, the binding energy (BE) values of Ni–N bond of $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$ shifts negatively by about 0.2 eV (Fig. 3a). The difference in



Fig. 3 (a) Ni 2p XPS spectra of Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF, Fe-Ni₃N/NF and Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF (b) Fe 2p XPS spectra of Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF and Fe-Ni₃N/NF. (c) Mo 3d XPS spectra Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF and Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF (d-f) N 1s XPS spectra of Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF, Ni_{0.2}Mo_{0.8}N/NF and Fe-Ni₃N/NF.

BEs of the Ni–N bond between the Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF and Fe–Ni₃N/NF might be attributed to the electron interaction between the Fe–Ni₃N nanosheets and Ni_{0.2}Mo_{0.8}N nanosheets or the interaction with underlying NF substrates.

The BEs shift in Mo 3d suggests the strong electron transfer from the Ni_{0.2}Mo_{0.8}N to Fe–N₃N, leading to the strong interaction between the two phases. Fig. 3d shows the N 1s XPS spectra Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF, the peak at about 397.1 eV corresponds to the N species of transition metal–N bonds, while the peak at about 395.2 eV was attributed to Mo–N bonds.⁵³ The peak at 398.2 eV belongs to the N–H species on the surface of materials.⁵⁴ Fig. 3e and f represented the N 1s XPS spectra of Ni_{0.2}Mo_{0.8}N/NF and Fe–Ni₃N/NF. By analyzing the peak in Fig. 3d and f, the negative shift in BEs of the Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF suggests that the Fe–Ni₃N integrated with Ni_{0.2}Mo_{0.8}N can lead to the enhanced charge transfers and electron interactions.

The OER performance of samples was examined. As shown in Fig. 4a and b, the Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF displays the overpotential of 266 mV to acquire 20 mA cm⁻². The overpotentials of Ni_{0.2}Mo_{0.8}N/NF (320 mV), Fe–Ni₃N/NF (292 mV) and commercial RuO₂ (328 mV) are higher than those of the Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF. Fig. 4c shows the Tafel slopes of electrocatalysts. The Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF had the lowest Tafel slopes (54 mV dec⁻¹), indicating the best OER performance than those of the Fe–Ni₃N/NF, Ni_{0.2}Mo_{0.8}N/NF and commercial RuO₂, respectively. In addition, the performances of Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF was superior to those of recently reported OER catalysts (Table S1†).

EIS measurements were performed to analyze the OER electrode kinetics. In the Nyquist plots (Fig. 4d), the Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF reveals a small charge transfer resistance (R_{ct}) of 4.4 Ω . The R_{ct} of Fe–Ni₃N/NF and Ni_{0.2}Mo_{0.8}N/NF is 10.6 Ω and 26.9 Ω , respectively. The relatively lower Tafel slope and R_{ct} suggests the fast kinetics process of Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF. The combination of Fe–Ni₃N and Ni_{0.2}Mo_{0.8}N nanosheets lead to the superior OER activity.



Fig. 4 (a) OER polarization curves of various catalysts. (b) The overpotential of the catalysts obtained at current density of 20 mA cm⁻². (c) Tafel slopes of catalysts. (d) EIS spectra of various catalysts recorded at potential of 1.47 V vs. RHE. (e) The current-time curves of $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$ for OER. Inset is the FE-SEM image of $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$ after OER stability test. (f) The polarization curves of $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$ before and after stability test for 50 h.



Fig. 5 (a) Ni 2p, (b) Fe 2p, (c) Mo 3d and (d) N 1s XPS spectra of $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$ after OER stability test for 50 h.

The stability test was carried out by chronopotentiometry measurement to investigate the changes of the Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF catalyst after prolonged water electrolysis. Fig. 4e shows the current-time curves of OER chronopotentiometry test. The current density of Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF still retain 91% after continuous electrolysis for 50 h, and FE-SEM image of Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF demonstrates that the hierarchical and two-dimensional features of the Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF remain unchanged after the stability test. As displayed in Fig. 4f, the LSV curves before and after cycle tests also exhibit negligible decrease in current density, suggesting the excellent stability in alkaline condition.

The XRD patterns, XPS survey and XPS spectra of the $Ni_{0.2}Mo_{0.8}N$ /Fe- Ni_3N /NF after OER and HER stability test was

shown in Fig. S6-S10[†] and 5. As shown in Fig. S7 and S9,[†] the diffraction peaks for Ni_{0.2}Mo_{0.8}N (JCPDS 29-0931) both become weaker after the HER and OER stability test, respectively. The XPS survey and XPS spectra of Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF after OER and HER all indicate the existences of Ni, Mo, Fe and N elements. The chemical states display small changes, suggesting the superior stability of main structures. The N 1s XPS changes after the HER and OER suggests the surface oxidation of Ni_{0.2}Mo_{0.8}N and the formation of metal-oxyhydroxide or metal-hydroxide. In Fig. 5a, the Ni 2p XPS spectra after stability tests show the BEs of Ni^{3+} (874.5 and 856.8 eV) and Ni^{2+} (872.9 and 855.4 eV).⁵⁰ In the Fe 2p XPS spectra (Fig. 5b), the peaks of Fe³⁺ (725.5 and 713.0 eV), Fe²⁺ (722.3 and 709.7 eV) and satellite can be observed.55 Fig. 5c and d represent the Mo 3d and N 1s spectra after OER test. The peaks of Mo⁶⁺ 235.0 and Mo⁴⁺ (233.8 and 230.1 eV), Mo3+ (233.0 and 229.0 eV) and N-H bond (399.4 eV), metal-N bond (397.9 eV), Mo-N bond (394.3 eV) can still be observed.38,49 The chemical states of elements exhibit changes after the OER stability tests, suggesting that the surface of electrocatalyst might be transformed into the metaloxyhydroxide or metal-hydroxide.48

The electrochemical performance of as-synthesized materials toward HER has also been performed. As shown in Fig. 6a and b, the Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF displays the best HER activity (40 mV, 20 mA cm⁻²), which is lower than other electrocatalysts. In addition, the Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF also obtains the lowest Tafel slope of 51 mV dec⁻¹ (Fig. 6c). The Nyquist plots under HER condition were shown in Fig. 6d. The Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF displays a smaller charge-transfer resistance (R_{ct}), indicating that the Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF exhibits enhanced charge-transfer ability. The outstanding charge-transfer ability of electrocatalysts for HER lead to favorable electrocatalytic kinetics. Some recently reported HER



Fe 2

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Fig. 6 (a) HER polarization curves of various catalysts. (b) The overpotential of the catalysts obtained at current density of 20 mA cm⁻². (c) Tafel slopes of catalysts. (d) EIS spectra of various catalysts recorded at potential of -0.03 V vs. RHE. (e) The current-time curves of Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF for HER. Inset is the FE-SEM image of Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF after HER stability test. (f) The polarization curves of Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF before and after continuous stability test.



Fig. 7 (a) Ni 2p, (b) Fe 2p, (c) Mo 3d and (d) N 1s XPS spectra of $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$ after HER stability test for 50 h.

electrocatalyst was summarized in Table S2,† the overpotential of Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF can be comparable. It is noted that the outstanding HER performance of Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF could be ascribed to the strong interactions between Fe–Ni₃N and Ni_{0.2}Mo_{0.8}N/NF phases.

The durability of Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF toward HER was tested at a potential of 74 mV vs. RHE. After HER chronopotentiometry test for 50 h, FE-SEM image (Fig. 6e) shows that the hierarchical and two-dimensional features of the Ni_{0.2}Mo_{0.8}N/ Fe–Ni₃N/NF remain the same after the HER stability test. The current density of Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF retain 95% (Fig. 6e). The LSV curves (Fig. 6f) of Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF after HER stability tests both show slight difference, suggesting the excellent stability for HER in alkaline condition.

The peak for Ni_{0.2}Mo_{0.8}N phases can be observed from the Fig. S9.† In XPS survey, the signal of N was observed (Fig. S10†). By analyzing the Fig. 7a, the peaks of Ni³⁺ (875.5 and 856.4 eV), Ni²⁺ (873.1 and 855.4 eV) and Ni–N bonds (870.0 and 852.6 eV) can be observed.⁴⁵ For the Fe 2p XPS spectra (Fig. 7b), the peaks of Fe³⁺ (723.9 and 711.8 eV) and Fe⁰ (720.6 and 706.6 eV) still exist.⁵² Furthermore, Mo 3d (Fig. 7c) and N 1s (Fig. 7d) XPS spectra still show the existence of Mo⁶⁺, Mo⁴⁺, Mo³⁺ and N–H, metal–N, Mo–N.⁴⁹ XRD pattern and XPS spectra indicated that the chemical compositions of sample were almost unchanged, revealing the superior stability of Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF in HER process.

The ECSA of various electrocatalysts were roughly obtained by the C_{dl} . The C_{dl} were measured through CVs (Fig. S11[†]). In Fig. 8a, the C_{dl} of the Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF, Fe–Ni₃N/NF and Ni_{0.2}Mo_{0.8}N/ NF were 129, 11 and 70 mF cm⁻², respectively. The ECSA of the Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF, Fe–Ni₃N/NF and Ni_{0.2}Mo_{0.8}N/NF was 1612.5 cm², 137.5 cm² and 875.0 cm², respectively. The bigger C_{dl} and ECSA of Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF demonstrates that the integration of the small Ni_{0.2}Mo_{0.8}N nanosheets on Fe–Ni₃N nanosheets can expose abundant catalytically active site.

Considering the excellent performance of OER and HER and practical applications of the material, $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$ was used as the bifunctional electrocatalysts for overall waters



Fig. 8 (a) The electrochemical double-layer capacitances of $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$, Fe $-Ni_3N/NF$ and $Ni_{0.2}Mo_{0.8}N/NF$. (b) Overall water splitting polarization curves of two-electrode cell assembled by various catalysts in 1 M KOH solution. (c) The current-time curves of overall water splitting for $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$ in KOH solution. (d) Water splitting polarization curves of the initial $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$ and the $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$ after stability test.

splitting. Fig. 8b shows the overall water splitting polarization curves of electrocatalysts. The $Ni_{0.2}Mo_{0.8}N/Fe-Ni_3N/NF$ only requires a potential of 1.54 V to achieve a current density of 10 mA cm⁻², which is much smaller than that of Fe-Ni₃N/NF (1.67 V) and Ni_{0.2}Mo_{0.8}N/NF (1.59 V). Furthermore, the Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF bifunctional electrocatalyst shows strong stability, which displayed small degradation after stability test for 20 h (Fig. 8c). The Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF electrode after stability test only need 1.55 V (10 mA cm⁻²) (Fig. 8d), which further demonstrate its excellent durability.

Conclusions

In summary, we have synthesized a hierarchical, well-organized nanostructure with small Ni_{0.2}Mo_{0.8}N nanosheets grown on the Fe-Ni₃N nanosheets. Compared with the individual Fe-Ni₃N/NF and Ni_{0.2}Mo_{0.8}N/NF, the integrated Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF exhibits favorable electrical conductivity, resulting in exceptional performance of OER and HER in alkaline conditions. The experimental results indicated that Ni0.2MO0.8N/Fe-Ni3N/NF exhibit superior stability during the OER and HER process. The remarkable electrocatalytic performance and durability of the Ni_{0.2}Mo_{0.8}N/Fe-Ni₃N/NF can be attributed to the optimization of electronic properties under the interaction of Fe-Ni₃N nanosheets and small Ni_{0.2}Mo_{0.8}N nanosheets. Benefit from the excellent chemical performance, stability, and higher ECSA of Ni_{0.2}Mo_{0.8}N/Fe–Ni₃N/NF, the as-synthesis material can be used as the efficient electrocatalyst for overall water-splitting in alkaline conditions.

Conflicts of interest

The authors declare no competing financial interests.

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