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# Polyaniline coated tungsten trioxide as an effective adsorbent for the removal of orange G dye from aqueous media

 Abdelghani Hsini,<sup>†\*</sup> Yassine Naciri,<sup>†</sup> Asmae Bouziani,<sup>†</sup> Nohu Aarab,<sup>a</sup> Abdelilah Essecri,<sup>a</sup> Abdelaziz Imgharn,<sup>a</sup> Mohamed Laabd,<sup>b</sup> J. A. Navío,<sup>c</sup> F. Puga,<sup>c</sup> Rajae Lakhmiri<sup>d</sup> and Abdallah Albourine<sup>\*a</sup>

In this work, the core-shell PANI@WO<sub>3</sub> composite was obtained from the reaction of aniline monomer polymerization with WO<sub>3</sub> particles; sodium persulfate was used as an oxidant. Various analytical techniques such as scanning electron microscopy (SEM-EDS), X-ray diffraction (XRD), Fourier transform infrared spectroscopy (FTIR), Brunauer–Emmett–Teller (BET), and X-ray photoelectron spectroscopy (XPS) were used to characterize the as-prepared PANI@WO<sub>3</sub> adsorbent, which well confirmed that the WO<sub>3</sub> particles were coated by polyaniline polymer. The PANI@WO<sub>3</sub> composite was tested as an adsorbent to remove reactive orange G (OG) for the first time. pH, adsorbent dose, contact time, initial dye concentration, and temperature were systematically investigated in order to study their effect on the adsorption process. The experimental findings showed that the PANI@WO<sub>3</sub> composite has considerable potential to remove an aqueous OG dye. Langmuir and Freundlich's models were used to analyze the equilibrium isotherms of OG dye adsorption on the PANI@WO<sub>3</sub> composite. As a result, the best correlation of the experimental data was provided by the Langmuir model, and the maximum capacity of adsorption was 226.50 mg g<sup>-1</sup>. From a thermodynamic point of view, the OG dye adsorption process occurred spontaneously and endothermically. Importantly, PANI@WO<sub>3</sub> still exhibited an excellent adsorption capability after four regeneration cycles, indicating the potential reusability of the PANI@WO<sub>3</sub> composite. These results indicate that the as prepared PANI@WO<sub>3</sub> composite could be employed as an efficient adsorbent and was much better than the parent material adsorption of OG dye.

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## 1. Introduction

Water is a vital component for the existence of all living organisms. Today, most water sources are at risk of being contaminated by various organic and inorganic chemical pollutants from industrial effluents.<sup>1–4</sup> The textile industry discharges effluents containing synthetic dyes and is considered to be the most polluting of all industries. Even at low concentrations, dyes are visible and deemed to be toxic.<sup>5–7</sup> Due to reduced light penetration, accompanied by a reduced amount of dissolved oxygen, dyes have distributed the photo-synthetic activity of aquatic life.<sup>8</sup> Hence, the removal of

synthetic organic dyes from waste effluents becomes a fundamental environmental challenge. For this purpose, several purification methods including biological, ozonation, membrane filtration, physiochemical, and advanced oxidation treatment processes, are commonly applied to treat dye-laden effluents.<sup>9–14</sup> However, the significant restrictions of these technologies are long operation time, low effectiveness, high sludge production, high cost, *etc.*<sup>15</sup> Adsorption is an alternative wastewater treatment process because of its high performance and ease of operation, as well as the possibility of regeneration.<sup>16–18</sup>

Several adsorbents such as zeolite, cellulose, activated carbon, clay, and so on have been applied to remove dyes from wastewater.<sup>16,18</sup> The reported adsorbents present many drawbacks such as high cost of preparation, regeneration difficulty, and a long time to achieve adsorption equilibrium, thus restricting their large-scale application to treat textile effluents.<sup>8</sup> Therefore, investigations of new recyclable adsorbents are needed.

In recent years, polyaniline (PANI) has attracted much attention. Due to the high electrochemical activity, easiness of protonation reversibility, excellent recyclability of redox,

<sup>a</sup>Laboratory of Materials and Environment, Faculty of Sciences, Ibn Zohr University, Agadir, Morocco. E-mail: [abdelghani.hsini@edu.uiz.ac.ma](mailto:abdelghani.hsini@edu.uiz.ac.ma); [albourine.abdallah@gmail.com](mailto:albourine.abdallah@gmail.com)

<sup>b</sup>Chemical Engineering Department, Middle East Technical University, Ankara, Turkey

<sup>c</sup>Instituto de Ciencia de Materiales de Sevilla, Centro Mixto Universidad de Sevilla-CSIC, Américo Vespucio 49, 41092 Sevilla, Spain

<sup>d</sup>Laboratory of Chemical Engineering and Valorization Resources, Faculty of Sciences and Techniques, Abdelmalek Essaadi University, Tangier, Morocco

<sup>†</sup> These authors contributed equally to this work.



environmental stability, easy preparation, and low cost.<sup>19,20</sup> Moreover, owing to the presence of large amounts of amine and imine functional groups in the polymer chains, PANI and its composites can be used as adsorbents to remove anionic and cationic dyes.<sup>21–23</sup> However, the major drawback of PANI usage as an adsorbent is the low mass density making it hard to settle and limiting its usefulness as an adsorbent.<sup>23,24</sup>

Recently, metal oxides have been reported to hold immense potential for adsorptive remediation of wastewater. This can be attributed to their large surface area, thermal stability, porous structures, low toxicity, easy recovery, and the presence of Lewis acid–base sites, O-vacancy sites, and defects on the surface.<sup>25,26</sup> Among these metal oxides, tungsten trioxide (WO<sub>3</sub>) nanoparticles have attracted considerable studies properties such as small energy band gap (from 2.4 to 2.8 eV), stable physico-chemical properties, non-toxicity, and resistance to photo-corrosion. However, their adsorption properties are significantly less investigated.<sup>27,28</sup>

In general, the addition of inorganic materials to conductive polymers alters the physical and chemical properties of the components, and resulting composites display synergic or complementary behaviors.<sup>29–31</sup> For example, the results of the study conducted by S. M. Ali *et al.* showed that SiO<sub>2</sub>, TiO<sub>2</sub>, Al<sub>2</sub>O<sub>3</sub>, and Fe<sub>3</sub>O<sub>4</sub> oxides nanoparticles covered by PANI are efficient in adsorbing reactive congo red dye.<sup>32</sup> Similarly, S. Agarwal *et al.* reported using polyaniline-zirconium oxide nanocomposite as an efficient adsorbent to remove MB dye from an aqueous solution.<sup>29</sup> Hence, a composite of tungsten trioxide (WO<sub>3</sub>) with PANI is expected to display tremendous results for the effective abatement of dyes from water.

No former study has been conducted to remove OG dye by PANI@WO<sub>3</sub> composite to our best knowledge. Thus, the present study would serve as an essential contribution to filling this knowledge gap. Herein, the goals of this study are to (1) synthesize PANI@WO<sub>3</sub> composite and use it as scavenger-type material for the removal of OG dye from aqueous solution; (2) investigate the effect of various operating parameters such as adsorbent dose, pH, initial dye concentration, contact time, and temperature in batch experiments on OG removal; (3) characterize the chemical, morphological and physical properties of the PANI@WO<sub>3</sub> composite before and after adsorption of OG dye *via* FTIR, XRD, XPS, SEM, EDS, TEM, and BET techniques; (4) delineate the kinetics and thermodynamics of OG dye adsorption; and (5) discuss the adsorption mechanism of chromium(vi) onto PANI@WO<sub>3</sub> composite using XPS analysis.

## 2. Experimental

### 2.1 Materials and chemicals

Sigma-Aldrich supplied all the chemical reagents used in this paper, ammonium tungstate ((NH<sub>4</sub>)<sub>6</sub>H<sub>5</sub>[H<sub>2</sub>(WO<sub>4</sub>)<sub>6</sub>]·H<sub>2</sub>O, 99%), aniline (C<sub>6</sub>H<sub>5</sub>NH<sub>2</sub>, 99%), sodium persulfate (Na<sub>2</sub>S<sub>2</sub>O<sub>8</sub>, 99%), orange G (OG), hydrochloric acid (HCl, 36%), sodium hydroxide (NaOH), acetone (C<sub>3</sub>H<sub>6</sub>O) and ethanol (CH<sub>3</sub>CH<sub>2</sub>OH) were of analytical grade, and no extra purification was needed. Distilled water was used for the preparation of stock solutions as well as for the materials synthesis and adsorption experiments.

### 2.2 Fabrication of PANI@WO<sub>3</sub> nanocomposite

The WO<sub>3</sub> powder used in this study was synthesized according to the procedure described by Li *et al.*<sup>33</sup> The *in situ* polymerization of aniline monomer was applied to synthesize PANI@WO<sub>3</sub> nanocomposite with core–shell structure in the presence of a host matrix (WO<sub>3</sub> particles). First, 1 g of WO<sub>3</sub> powder was dissolved in 100 mL of HCl (1 M) *via* sonication for 30 min. Next, a recently distilled aniline (1 mL) was dropped into the mixture and then sonicated for 1 h. Then, under stirring, 50 mL of HCl (1 M) containing the oxidant sodium persulfate (with a monomer/oxidant molar ratio of 1 : 2) was added drop by drop and kept under stirring for 12 h at ambient temperature. The resulting mixture was filtered, and the collected product was washed numerous times with water, ethanol, and acetone in order to eliminate the excess of oligoaniline and oxidant. Finally, the prepared PANI@WO<sub>3</sub> nanocomposite material was kept overnight at 80 °C.

### 2.3 Characterization of the adsorbent

The physicochemical properties of the PANI@WO<sub>3</sub> composite were measured and evaluated by numerous characterization methods. The crystal structure of WO<sub>3</sub> and PANI@WO<sub>3</sub> composite was characterized by X-ray powder diffraction (XRD) analysis *via* an EMPYREAN PANALYTICAL diffractometer equipped with a copper X-ray (wavelength of  $1.54 \times 10^{-10}$  m, voltage of 45 kV, and current intensity of 35 mA). The XRD data were recorded over a  $2\theta$  angle range of 10–80°. The microstructure of the as-prepared adsorbent samples was recorded on a transmission electron microscope (Philips CM 200 microscope). FTIR spectroscopy (ALPHA-Bruker Optics, Germany) with KBr pellets scanning in the 400 to 4000 cm<sup>-1</sup> frequency array was used to identify the characteristic chemical bond of the WO<sub>3</sub> particles prior to and after PANI coating. The analysis of the chemical composition of the PANI@WO<sub>3</sub> surface before and after OG dye adsorption was made by X-ray photoelectron spectroscopy (XPS) (SPECS spectrometer). The obtained spectra were standardized to C 1s peak at 284.6 eV. The morphology and microstructure of WO<sub>3</sub> and PANI@WO<sub>3</sub> before and after OG adsorption were documented on a scanning electron microscope (SEM, JEOL, JSMIT200) attached with a dispersive energy X-ray (EDS) analysis. A Micromeritics ASAP 2420 accelerated surface area and porosimetry system, with a liquid-nitrogen cooling bath (–195.78 °C), was used to measure the BET specific surface area.

The point of zero charges (pH<sub>PZC</sub>) corresponds to the pH value of the solution surrounding the adsorbent in which the positively charged active sites were equal to the negative ones on the adsorbent surface. The pH<sub>PZC</sub> of PANI@WO<sub>3</sub> was calculated *via* the potentiometric titration procedure described in our previous article.<sup>34</sup>

### 2.4 Batch experiments

The adsorption experiments were performed in a batch reactor. The adsorption behavior of PANI@WO<sub>3</sub> composite toward OG dye was experimentally investigated under different operational settings, including the amount of adsorbent (0.125–1 g L<sup>-1</sup>),



solution pH (2–10), adsorption time (0–240 min), initial concentration of dye (0.025–0.400 g L<sup>-1</sup>) and temperature (298–328 K). The influence of these parameters on the efficiency of adsorption was investigated, changing one parameter at a time. After each adsorption experiment, the PANI@WO<sub>3</sub> was removed from the solution *via* a Millipore filter (0.45 μm). The OG concentration prior and after adsorption was obtained *via* a UV2300 spectrophotometer and following the band of adsorption at 476 nm. The percentage absorbed *R* (%), and the capacity of adsorption *Q<sub>e</sub>* (mg g<sup>-1</sup>) was obtained from the below formulas:<sup>34</sup>

$$R (\%) = \frac{(C_0 - C_e)}{C_0} \times 100 \quad (1)$$

$$Q_e = \frac{(C_0 - C_e)}{m} \times V \quad (2)$$

*C<sub>0</sub>* (mg L<sup>-1</sup>) is the OG initial concentration; *C<sub>e</sub>* (mg L<sup>-1</sup>): OG concentration at equilibrium; *m* (g) is the quantity of the PANI@WO<sub>3</sub> composite, and *V* (L) is the volume of OG.

### 2.5 PANI@WO<sub>3</sub> regeneration

The regeneration experiment was performed for 6 h using a 0.5 M NaOH eluent solution used for the desorption of OG dye

from PANI@WO<sub>3</sub>. A successful regeneration was obtained after washing PANI@WO<sub>3</sub> composite with distilled water and redoping with 1 M HCl for an additional adsorption cycle. In this study, four adsorption/desorption cycles were conducted.

## 3. Results and discussion

### 3.1 Characterization of the adsorbent

XRD is a central technique for the determination of the structure and composition of obtained composites. Fig. 1 displays the XRD patterns of WO<sub>3</sub> and PANI@WO<sub>3</sub> composite. The characteristic diffraction peaks appeared at the peaks at  $2\theta = 23.09, 23.51, 24.30, 26.66, 28.85, 33.35,$  and  $34.09$  refer to (002), (020), (200), (120), (112), (022) and (202) diffraction planes of WO<sub>3</sub> consistent with the monoclinic structure and are in accordance with JCPDS file 01-072-0677. The lattice constants  $a = 7.3060 \text{ \AA}$ ,  $b = 7.5400 \text{ \AA}$  and  $c = 7.6920 \text{ \AA}$  for WO<sub>3</sub>. The addition of PANI did not alter the peaks' position or shapes; thus, the XRD patterns of the PANI@WO<sub>3</sub> composite are similar to WO<sub>3</sub>. Therefore, the PANI did not change the crystal structure of WO<sub>3</sub>, and no peaks belonging to PANI were observed because of its amorphous nature in the composite. Analogous results were described for other PANI-based composites.<sup>35–37</sup>

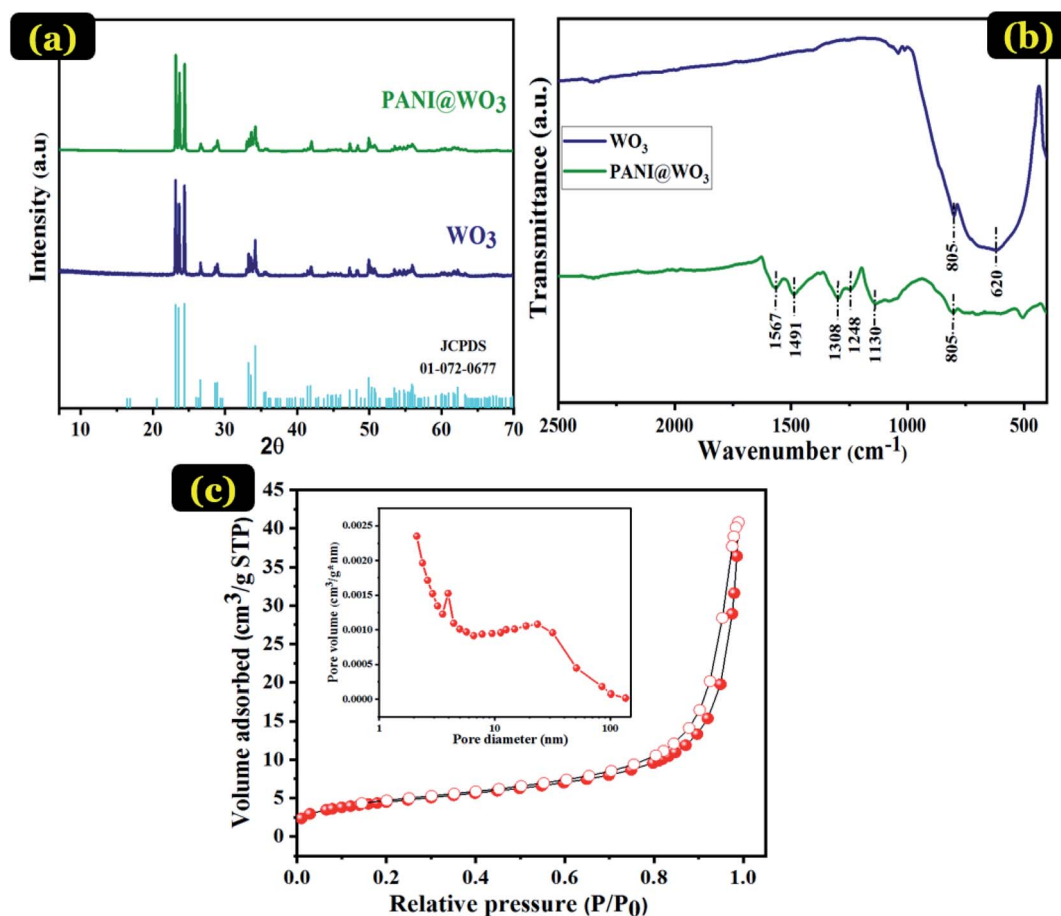


Fig. 1 XRD patterns of WO<sub>3</sub> and PANI@WO<sub>3</sub> composite (a), FTIR spectra of WO<sub>3</sub> and PANI@WO<sub>3</sub> composite (b), N<sub>2</sub> adsorption–desorption isotherms and the conforming pore size distribution plot of PANI@WO<sub>3</sub> composite (c).



The FT-IR spectroscopy further identifies the structural information and chemical component of the PANI@WO<sub>3</sub> composite. For comparison, the spectra of WO<sub>3</sub> was also recorded and shown in Fig. 1(b). The characteristic peaks of WO<sub>3</sub> situated at 620 and 805 cm<sup>-1</sup> are recognized as the stretching vibrations of O–W–O.<sup>38</sup> The adsorption peak at 805 cm<sup>-1</sup> was also observed in the spectrum of PANI@WO<sub>3</sub>, which demonstrated the presence of WO<sub>3</sub> in the PANI@WO<sub>3</sub> composite. Moreover, the FT-IR spectrum of the PANI@WO<sub>3</sub> composite comprises new bands in the range 500–1600 cm<sup>-1</sup>,

characteristic of PANI stretching. The typical peak at 1130 cm<sup>-1</sup> is related to the anion used as a dopant (Cl<sup>-</sup>PANI). The vibration peaks of C=C and C–N stretching modes in Ar–NH–Ar are situated at 1248 and 1308 cm<sup>-1</sup>, respectively. Furthermore, the bands in the frequency range of 1491–1567 cm<sup>-1</sup> are characteristic of C=C in the structure of benzenoid and quinoid stretching vibrations.<sup>39</sup> The band at approximately 805 cm<sup>-1</sup> matches the C–H bending vibration out of the plane of the *para* disubstituted benzene rings.<sup>31,40</sup> These bands confirm the structure of polyaniline in its emeraldine salt form. The FT-IR

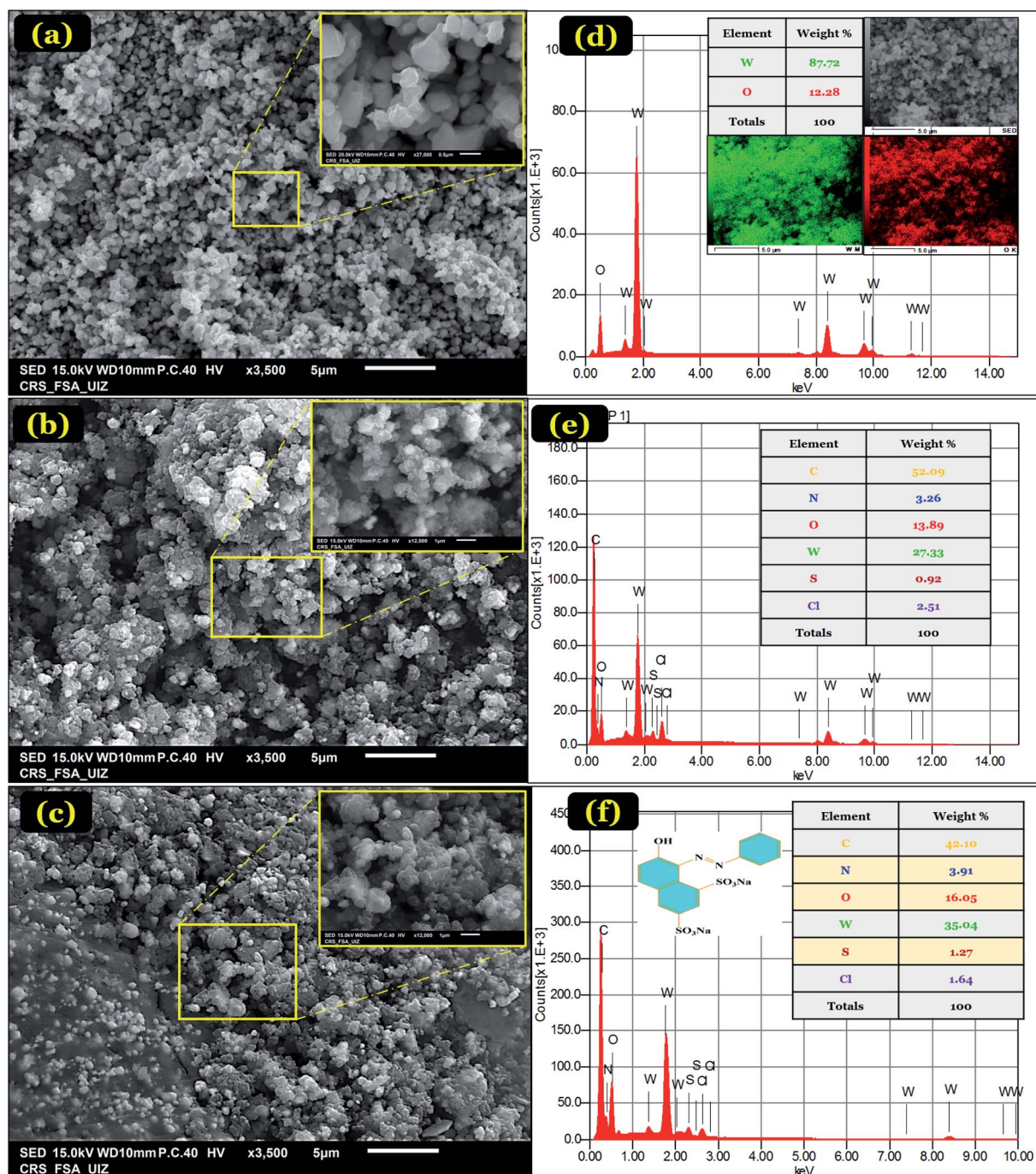


Fig. 2 SEM images of WO<sub>3</sub> (a), PANI@WO<sub>3</sub> (b), PANI@WO<sub>3</sub>-OG (c). EDS elemental analysis of WO<sub>3</sub> (d), PANI@WO<sub>3</sub> (e), and PANI@WO<sub>3</sub>-OG (f).



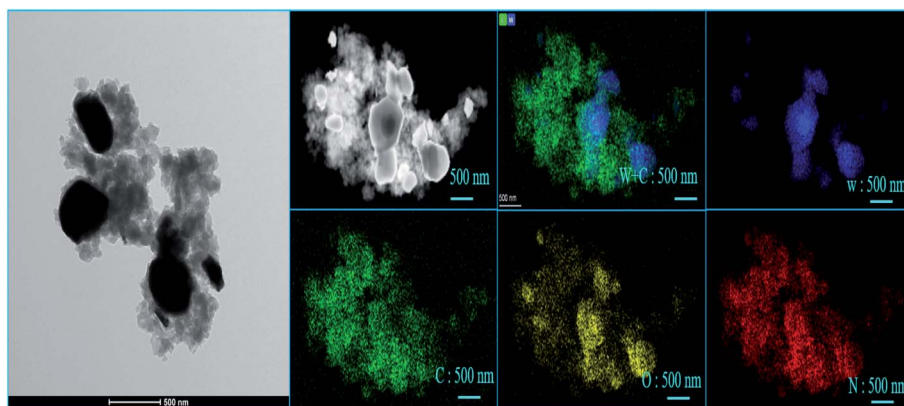


Fig. 3 TEM image, and tungsten (W), carbon (C), oxygen (O), nitrogen (N), represented by different colors, blue, green, yellow, and red, respectively, of elements mapping image of PANI@WO<sub>3</sub>.

findings provide substantial proof of the successful formation of the PANI@WO<sub>3</sub> composite.

The BET surface is an essential property of the adsorbent, which influences the adsorption performance. Pore diameter distribution also has a significant impact on the adsorption process performance. BET method can determine the specific surface area of particles, the average pore radius, and the total pore volume of the adsorbent. Fig. 1(c) exhibits the nitrogen adsorption–desorption isotherms of PANI@WO<sub>3</sub> composite with an inset showing its conforming Barrett–Joyner–Halenda (BJH) pore size distribution. From Fig. 1(c) and according to the IUPAC classification, PANI@WO<sub>3</sub> composite displays a type IV isotherm consisting of a mesoporous material.<sup>41</sup> The pore diameter is between 2 and 4 nm (Fig. 1(c)), which confirms the mesoporous structure of the PANI@WO<sub>3</sub> composite. The prepared PANI@WO<sub>3</sub> composite has shown specific surface area and total pore volume of 16.28 m<sup>2</sup> g<sup>-1</sup> and 0.04 cm<sup>3</sup> g<sup>-1</sup> using the BET method. This surface area is higher or comparable to many polymer based composite adsorbents like polyaniline/SiO<sub>2</sub> nanocomposite (13.36 m<sup>2</sup> g<sup>-1</sup>)<sup>42</sup>, [https://](https://www.sciencedirect.com/science/article/pii/S1350417718312501-b0080)

[www.sciencedirect.com/science/article/pii/S1350417718312501-b0080](https://www.sciencedirect.com/science/article/pii/S1350417718312501-b0080), polyaniline/TiO<sub>2</sub> (11.60 m<sup>2</sup> g<sup>-1</sup>).<sup>43</sup>

Fig. 2(a–c) displays the morphology of the synthesized WO<sub>3</sub>, PANI@WO<sub>3</sub>, and PANI@WO<sub>3</sub> after OG adsorption. WO<sub>3</sub> nanoparticles show a uniform structure having approximately 500 nm diameter as well as a few agglomerations. While PANI@WO<sub>3</sub> exhibits pores that resulted from PANI coating on the surface of WO<sub>3</sub> and confirms the core–shell morphology of the PANI@WO<sub>3</sub> composite. The infiltration and diffusion of adsorbate to the internal PANI@WO<sub>3</sub> surface are eased by pores formation. Fig. 2(c) shows a notable modification in the surface of the PANI@WO<sub>3</sub> composite after the adsorption of OG. The surface of the PANI@WO<sub>3</sub> composite reached the saturation point.

EDS analysis of WO<sub>3</sub>, PANI@WO<sub>3</sub>, and PANI@WO<sub>3</sub> was performed after the OG adsorption, and the outcomes are displayed in Fig. 2(d–f). The elemental compositions are specified in the corresponding tables. The WO<sub>3</sub> EDS spectrum (Fig. 2(d)) shows mostly tungsten and oxygen. Fig. 2(e) shows the presence of W and O atoms belonging to WO<sub>3</sub> and the chloride (Cl), sulfur (S), and nitrogen (N) elements corresponding to PANI.

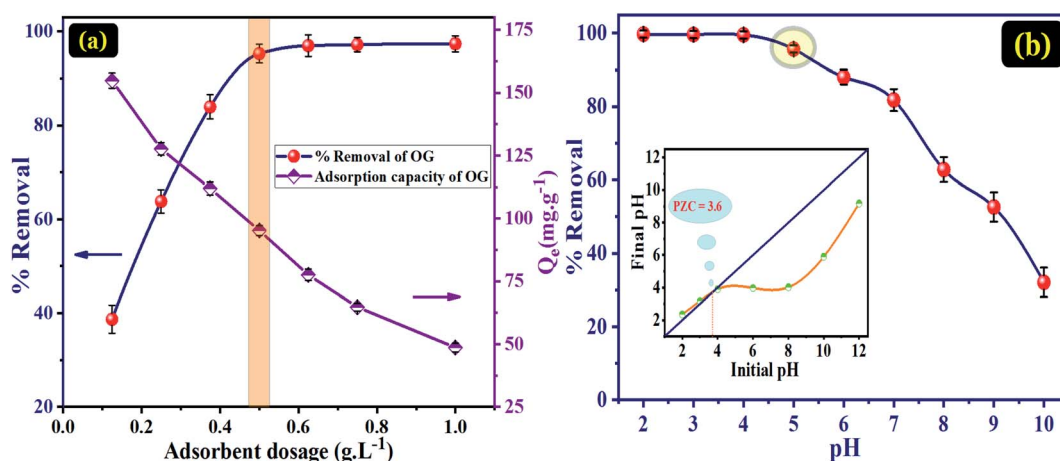


Fig. 4 (a) Effect of adsorbent dose on adsorption capacity and percentage removal of OG dye ( $C_0 = 0.050 \text{ g L}^{-1}$ ,  $\text{pH} = 5$ ,  $t = 120 \text{ min}$  and  $T = 298 \text{ K}$ ). (b) pH solution effect on the removal of OG *via* PANI@WO<sub>3</sub> ( $C_0 = 0.050 \text{ g L}^{-1}$ ,  $t = 120 \text{ min}$ , adsorbent dose =  $0.5 \text{ g L}^{-1}$  and  $T = 298 \text{ K}$ ).



Additionally, no impurity peaks were detected, confirming the purity of the synthesized PANI@WO<sub>3</sub> composite. An increase in O, S, and N was noticed upon the OG adsorption (Fig. 2(f)). In contrast, a decrease in the content of C was observed in PANI@WO<sub>3</sub>. These outcomes were anticipated as a consequence of the OG adsorption on the PANI@WO<sub>3</sub> surface.

The typical TEM images of PANI@WO<sub>3</sub> core-shell composite are given in Fig. 3. The morphological observations of the PANI@WO<sub>3</sub> composite revealed that the composite consists of well WO<sub>3</sub> particles with an average diameter of 500 nm, which are dispersed and attached to the PANI chain. It was observed that the dark core is the WO<sub>3</sub> particles, and the light colored shell is the PANI in the nanocomposite. The outcome of the TEM analysis confirms the successful synthesis of the PANI@WO<sub>3</sub> composite.

The elemental mappings for PANI@WO<sub>3</sub> composite were also performed. The elemental mapping confirmed the chemical composition of PANI@WO<sub>3</sub> composite with the presence of tungsten (W), carbon (C), oxygen (O), nitrogen (N), represented by different colors, blue, green, yellow, and red, respectively, as shown in Fig. 3. The investigation, as mentioned above, can support the successful coating of PANI on WO<sub>3</sub>.

### 3.2 Adsorption analysis

**3.2.1 Adsorbent dose and pH effects.** The influence of adsorbent dosage on adsorption capacity and percentage removal of OG dye is shown in Fig. 4(a). It has been found that an increase in the dosage of adsorbent from 0.1 to 0.5 g L<sup>-1</sup> leads to a noteworthy increase in the percentage of removal of OG dye (from 38 to 95%). This adsorption inclination may be due to the increased number of available binding sites.<sup>16</sup> Above an adsorbent of 0.5 g L<sup>-1</sup>, the removal efficiency remained constant. However, when the PANI@WO<sub>3</sub> dose increased, the capacity of adsorption decreased from 154.85 to 48.68 mg g<sup>-1</sup>, indicating that the PANI@WO<sub>3</sub> quantity had to be checked through the OG removal process.<sup>44</sup>

To some extent, the removal efficiency of OG dye could be enhanced as the number of active sites augmented, and the dose optimum of adsorbent was 0.5 g L<sup>-1</sup>. After this optimal

adsorbent dose, it was found that the efficiency of removal did not improve as more adsorbent was added. Therefore, it can be concluded that 0.5 g L<sup>-1</sup> of PANI@WO<sub>3</sub> is the optimum adsorbent dose used to perform the following batch experiments.

The pH of a solution is an essential parameter to understand the adsorption phenomenon by modifying the adsorbent surface charge and the protonation state of the adsorbate molecules. Therefore, the pH influence on the adsorption of OG and PZC of PANI@WO<sub>3</sub> was investigated in a pH array of 2–10. From Fig. 4(b), the percentage of OG dye removal remained very high (>95%) in the pH range of 2–5 and then decreased progressively as the initial solution pH increased. The pK<sub>a</sub> of OG are 1.0 and 11.5 for the deprotonation/ionization of the two sulfonic (–SO<sub>3</sub>H) groups and naphthalene hydroxy (–OH) function, respectively.<sup>34</sup> Therefore, the OG molecules are charged negatively over the entire established pH array. The pH<sub>PZC</sub> of PANI@WO<sub>3</sub> is 3.6 (Fig. 4(b)). Thus, the PANI@WO<sub>3</sub> surface charge is positive at pH < pH<sub>PZC</sub> and negative once pH > pH<sub>PZC</sub>. The OG adsorption can be due to the electrostatic interactions between OG molecules and the PANI@WO<sub>3</sub> composite surface. At pH < 4, the high adsorption performance was attributed to the electrostatic attractions between the anionic OG molecules and the positive charge of the PANI@WO<sub>3</sub> surface. Yet, at pH > 4, the adsorption decreased due to a decrease in the number of sites charged positively, inhibiting the OG adsorption on the PANI@WO<sub>3</sub> surface.<sup>45</sup> The OG adsorption efficiency observed at pH > 4 could be explained by π–π interactions between OG dye molecule and PANI. Therefore, the optimal pH value of 5 was selected and used in subsequent OG dye adsorption experiments.

**3.2.2 Kinetics study.** The influence of adsorption time on the OG removal by PANI@WO<sub>3</sub> was investigated, and the results are given in Fig. 5(a). The OG capacity of adsorption onto PANI@WO<sub>3</sub> composite increased rapidly in the first 30 min, then slowly, and the equilibrium was reached within 120 min. At the early stages of the adsorption process, the faster adsorption rate could be presumably attributed to a significant gradient of OG concentration at the solution/PANI@WO<sub>3</sub> interface. As the adsorption time extended, the active sites

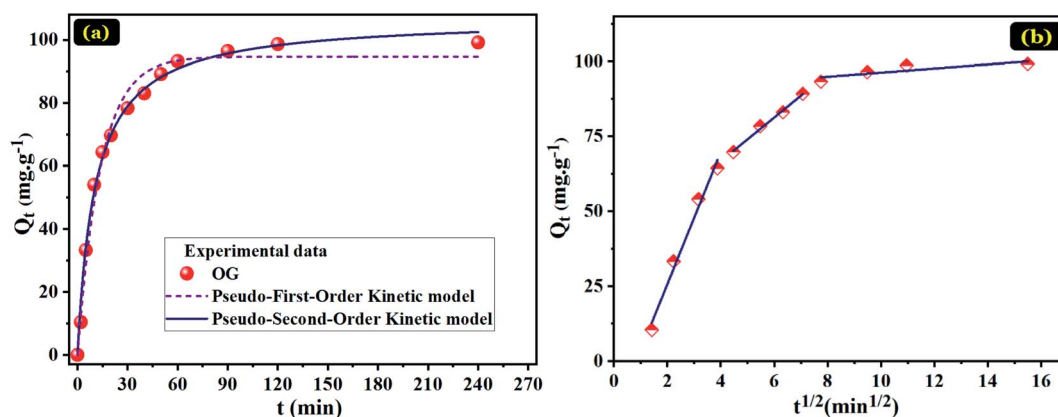


Fig. 5 Plots for OG adsorption on PANI@WO<sub>3</sub> ( $C_0 = 0.050$  g L<sup>-1</sup>, pH = 5, adsorbent dose = 0.5 g L<sup>-1</sup> and  $T = 298$  K). Pseudo-first and pseudo-second order (a) and Weber–Morris intra-particle diffusion (b).

Table 1 The adsorption of OG onto PANI@WO<sub>3</sub> kinetic parameters

	$Q_{e,exp}$ (mg g <sup>-1</sup> )	Pseudo-first order			Pseudo-second order				
		$k_1$ (min <sup>-1</sup> )	$Q_{e,1}$ (mg g <sup>-1</sup> )	$R^2$	$k_2$ (g mg <sup>-1</sup> min <sup>-1</sup> )	$Q_{e,2}$ (mg g <sup>-1</sup> )	$R^2$		
PANI@WO <sub>3</sub>	99.16	0.07203	94.64	0.9864	$8.7698 \times 10^{-4}$	106.98	0.9947		
Weber–Morris model									
	Initial linear portion			Second linear portion			Third linear portion		
	$k_{int,1}$	$C_1$	$R^2$	$k_{int,2}$	$C_2$	$R^2$	$k_{int,3}$	$C_3$	$R^2$
PANI@WO <sub>3</sub>	22.079	-18.452	0.9850	7.311	37.405	0.9935	0.686	89.385	0.7155

gradually filled, leading to decreased OG removal rate, and the adsorption efficiency plateaued (equilibrium state).

To understand the adsorption mechanisms of OG on the PANI@WO<sub>3</sub> composite, the adsorption kinetic data were fitted with pseudo-first-order and pseudo-second-order kinetic models (Fig. 5(a)). The equations of these kinetic models are given as follows:<sup>31</sup>

Pseudo-first-order model:

$$Q_t = Q_e(1 - e^{-k_1 t}) \quad (3)$$

Pseudo-second-order model:

$$Q_t = \frac{Q_e^2 k_2 t}{1 + Q_e k_2 t} \quad (4)$$

where  $k_1$  (min<sup>-1</sup>) is the rate constant of the pseudo-first-order and  $k_2$  (mg g<sup>-1</sup> min<sup>-1</sup>) the rate constant of the pseudo-second-order.  $Q_e$  (mg g<sup>-1</sup>) adsorbed amounts at equilibrium and  $Q_t$  (mg g<sup>-1</sup>) adsorbed amounts at time  $t$  (min).

The fitting parameters of the kinetic models are recorded in Table 1. On the basis of the correlation coefficient ( $R^2$ ) values, the pseudo-second-order model (0.9947) is more fitted to the

OG adsorption kinetic data than the pseudo-first-order model (0.9864). Besides, the adsorbed amount obtained from the pseudo-second-order model is nearer the experimental one. Therefore, the OG adsorption process on the PANI@WO<sub>3</sub> composite is following the pseudo-second-order kinetic rate.

The mass transfer mechanism of OG molecules between the aqueous phase and PANI@WO<sub>3</sub> surface was tested using an intra-particle diffusion model. The intraparticle diffusion model is defined as follows:<sup>31</sup>

$$Q_t = k_{int} t^{1/2} + C \quad (5)$$

where  $k_{int}$  (mg g<sup>-1</sup> min<sup>-1/2</sup>) represents the intraparticle diffusion rate constant.  $C$  (mg g<sup>-1</sup>) is a constant related to the thickness of the boundary layer.  $Q_t$  (mg g<sup>-1</sup>): the amount adsorbed at time  $t$  (min).

Fig. 5(b) shows the  $Q_t$  vs.  $t^{1/2}$  plots for OG adsorption onto PANI@WO<sub>3</sub> composite; the corresponding parameters are summarized in Table 1. The graphical representation clearly indicates that the OG dye transfer from bulk liquid to solid material surface sites is a three-stage process. The first stage corresponds to the mobility of OG from bulk liquid to the

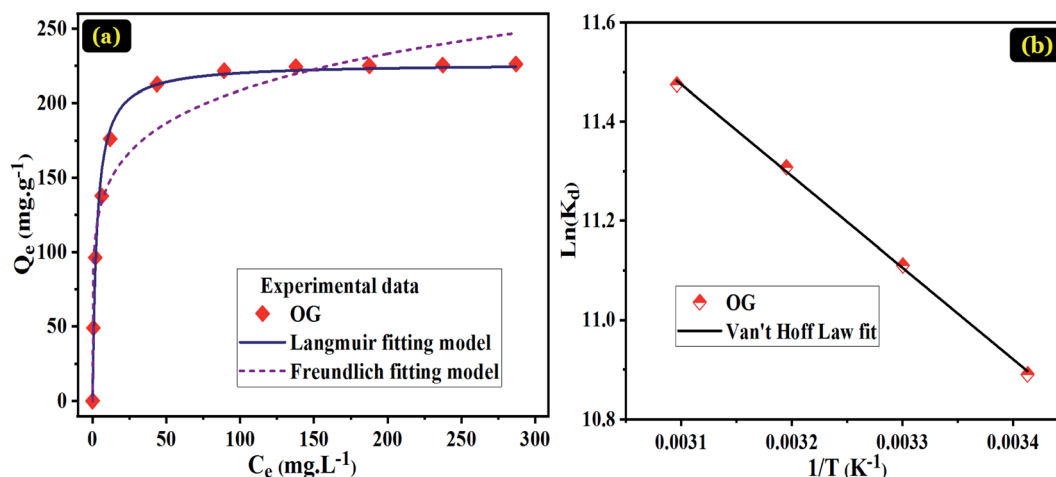


Fig. 6 (a) Non-linear Langmuir and Freundlich isotherm plots for the adsorption of OG on the PANI@WO<sub>3</sub> (pH = 5,  $t = 240$  min, adsorbent dose =  $0.5$  g L<sup>-1</sup> and  $T = 298$  K) and (b) Van't Hoff plot for thermodynamic parameters determination.



Table 2 Langmuir and Freundlich isotherm formulas and the conforming parameters for OG adsorption onto PANI@WO<sub>3</sub>

	$Q_{\text{exp}}$ (mg g <sup>-1</sup> )	Langmuir $Q_e = \frac{Q_{\text{max}}K_L C_e}{1 + K_L C_e}$			Freundlich $Q_e = K_F C_e^{1/n}$		
		$Q_{\text{max}}$ (mg g <sup>-1</sup> )	$K_L$ (L mg <sup>-1</sup> )	$R^2$	$n_F$	$K_F$ (mg g <sup>-1</sup> )	$R^2$
PANI@WO <sub>3</sub>	226.22	226.50	0.3455	0.9884	6.229	99.573	0.9345

Table 3 The uptake capacity of PANI@WO<sub>3</sub> for OG comparison with other adsorbents described in the literature

Adsorbent	Maximum adsorption capacity (mg g <sup>-1</sup> )		Reference
	OG		
ZnO-AC	153.8		47
Cts(x)-g-PNVP	63.7		48
AC from coffee grounds	100		49
PANI@AS biocomposite	190.98		34
Monoamine-modified silica	36.3		50
FSMD MNPs	109.1		51
PANI@OFI-A	303.94		52
PANI@WO <sub>3</sub> composite	226.50		Current study

external surface of PANI@WO<sub>3</sub> composite. This stage is characterized by a significant OG dye mass transfer owing to the higher mass flow rate. The second stage represents the diffusion of OG dye molecules inside the mesopores of the PANI@WO<sub>3</sub> composite. The diffusion rate of this step is inferior to that of the first one and is coordinated mainly by pore diffusion.<sup>31</sup> The third stage is credited to ultimate adsorption equilibrium due to the eventual saturation of adsorbent surface sites.<sup>46</sup>

**3.2.3 Adsorption isotherm and thermodynamic parameters.** Isotherm of adsorption study is a crucial feature to anticipate OG distribution at the interface between the liquid phase and adsorbent surface.<sup>31</sup> To understand the adsorption mechanism between OG dye and PANI@WO<sub>3</sub> surface, the experimental data of adsorption was regressively simulated with Langmuir and Freundlich isotherms models; outcomes are given in Fig. 6(a) and Table 2. On the basis of the obtained coefficient of regression ( $R^2$ ), it is clear that the Langmuir model is better suited to explain the adsorption of OG on PANI@WO<sub>3</sub> than the Freundlich model. Due to the uniform shape, the adsorption of OG is dominated by monolayer adsorption on the surface of PANI@WO<sub>3</sub>.<sup>31</sup> The adsorption capacity optimal for PANI@WO<sub>3</sub> adsorbent calculated from the Langmuir model was 226.50 mg g<sup>-1</sup>, which is near the adsorption value obtained

experimentally (226.22 mg g<sup>-1</sup>). The OG maximum capacity of adsorption on PANI@WO<sub>3</sub> was compared with already reported adsorbents, as summarized in Table 3. Obviously, the PANI@WO<sub>3</sub> displayed a remarkable OG adsorption capacity rather than earlier described adsorbents. Hence, the PANI@WO<sub>3</sub> composite is an excellent adsorbent to remove OG from an aqueous solution.

Indicative symbols:  $Q_{\text{max}}$  (mg g<sup>-1</sup>) the maximum adsorption capacity and  $C_e$  (mg L<sup>-1</sup>) concentration of OG at equilibrium;  $K_L$  (L mg<sup>-1</sup>),  $K_F$  (mg g<sup>-1</sup>), and  $n_F$  represent the constant of Langmuir isotherm associated with the adsorption heat, constant of Freundlich isotherm relating the adsorption density and the intensity of adsorption, respectively.

The thermodynamic study was achieved to investigate the nature of the OG dye adsorption process onto PANI@WO<sub>3</sub> composite (Fig. 6(b)). The thermodynamic parameters including enthalpy change ( $\Delta H^\circ$ ), Gibb's free energy change ( $\Delta G^\circ$ ), and entropy change ( $\Delta S^\circ$ ) were calculated by eqn (6) and (7):<sup>33</sup>

$$\Delta G^\circ = -RT \ln K_d \quad (6)$$

$$\ln K_d = \frac{\Delta S^\circ}{R} - \frac{\Delta H^\circ}{RT} \quad (7)$$

$T$  and  $R$  denote the absolute temperature (°K) and the constant of an ideal gas (8.314 J K<sup>-1</sup> mol<sup>-1</sup>).  $K_d$  signifies the constant of adsorption at equilibrium stated as:  $K_d = \frac{Q_e}{C_e} \times \rho$  (1000 mg L<sup>-1</sup>) denotes the density of water.

The thermodynamic parameters results are summarized in Table 4. The negative values of  $\Delta G^\circ$  suggest that the nature of the adsorption of OG over the PANI@WO<sub>3</sub> surface is a spontaneous process. Besides, the absolute values of  $\Delta G^\circ$  augmented with the increase of temperature, suggesting an enhanced spontaneity at higher temperatures. The magnitude of  $\Delta G^\circ$  (less than 40 kJ mol<sup>-1</sup>) evidenced that the OG dye adsorption is a physisorption type.<sup>34</sup> The positive  $\Delta H^\circ$  confirmed that the adsorption of OG onto PANI@WO<sub>3</sub> is an endothermic process. Moreover, the positive  $\Delta S^\circ$  value reveals a rise in the unpredictability at the solution/PANI@WO<sub>3</sub> interface through the sorption process, reflecting the excellent affinity of PANI@WO<sub>3</sub> composite towards OG dye molecules.

Table 4 The quantities of  $\Delta G^\circ$ ,  $\Delta H^\circ$ , and  $\Delta S^\circ$  for adsorption of OG on the PANI@WO<sub>3</sub>

	$\Delta H^\circ$ (kJ mol <sup>-1</sup> )	$\Delta S^\circ$ (J mol <sup>-1</sup> K <sup>-1</sup> )	$\Delta G^\circ$ (kJ mol <sup>-1</sup> )			
			298 K	308 K	318 K	328 K
PANI@WO <sub>3</sub>	15.86	143.77	-26.96	-28.43	-29.88	-31.27





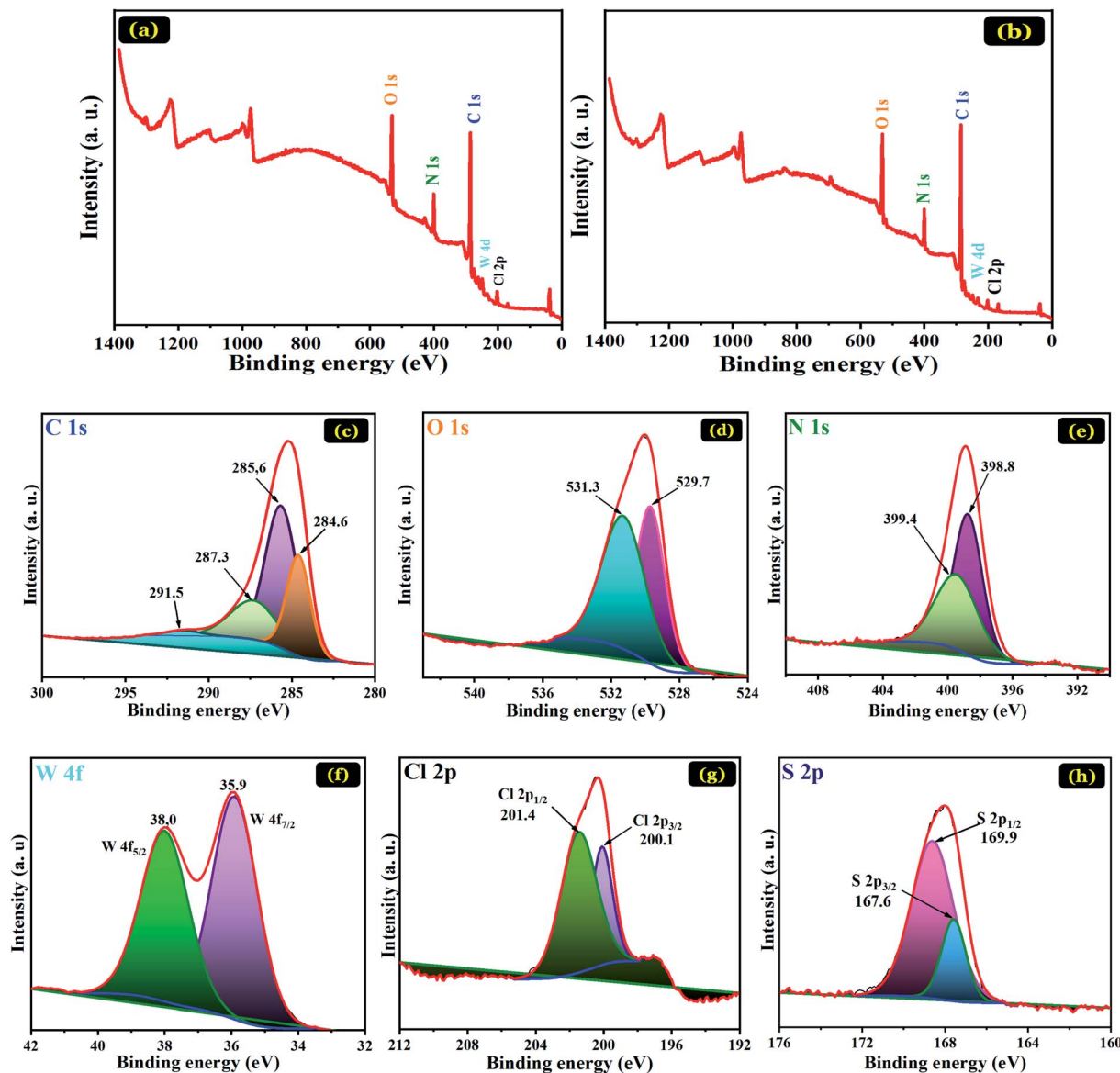


Fig. 7 XPS survey spectrum (a), high-resolution XPS spectra of C 1s (c), O 1s (d), N 1s (e), W 4f (f), Cl 2p (g) of PANI@WO<sub>3</sub> prior adsorption of OG. XPS survey spectrum (b), and high-resolution XPS spectra of S 2p (h) of PANI@WO<sub>3</sub> after OG dye adsorption.

**3.2.4 XPS analysis and mechanism of adsorption.** XPS technique was used to characterize the PANI@WO<sub>3</sub> prior and after OG adsorption. The spectral figures for C 1s, O 1s, N 1s, W

4d, Cl 2p, and S 2p of PANI@WO<sub>3</sub> are shown in Fig. 7. The conforming elemental examination is given in Table 5. The presence of PANI and WO<sub>3</sub> nanoparticles were confirmed by the presence of N 1s and W 4d peaks. Fig. 7(c) displays the C 1s spectrum where the peaks at 284.6, 285.6, 287.3, and 291.5 eV are characteristically related to the presence of C–C or C–H, C–N, C=N, and  $\pi$ - $\pi$  (satellite) links in the PANI.<sup>31</sup> The core-level spectrum of the N 1s (Fig. 7(e)) displays two peaks: 398.8 and 399.4 eV consistent respectively with the amine (–NH–) and doped amine (–NH<sup>+</sup>/–N<sup>+</sup>=) groups of the polyaniline.<sup>31</sup> Therefore, we can conclude that the doped PANI polymer (emeraldine salt form) was effectively placed on the surface of WO<sub>3</sub> particles. After OG dye adsorption, the XPS spectra signal of sulfur element was detected at a binding energy of 168 eV (Fig. 7(h)), indicating the adsorption of OG dye on the PANI@WO<sub>3</sub>

Table 5 Chemical composition of PANI@WO<sub>3</sub> prior and after adsorption of OG

Element	Atomic concentration (%)	
	PANI@WO <sub>3</sub>	PANI@WO <sub>3</sub> -OG
C	70.70	72.30
O	16.21	14.39
N	9.523	9.279
W	0.57	0.25
Cl	2.30	1.69
S	Atomic% lower than the detection limit	1.98



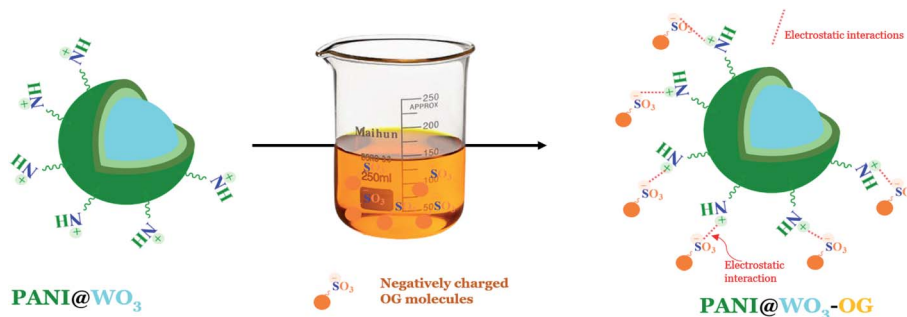


Fig. 8 Illustration of the adsorption mechanism of OG dye onto PANI@WO<sub>3</sub> composite.

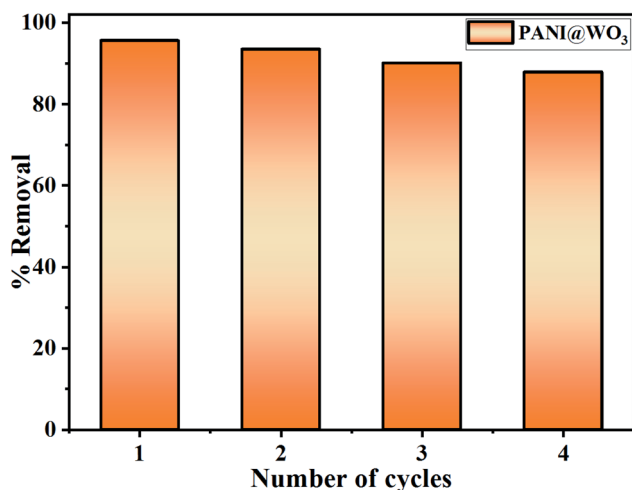


Fig. 9 Recyclability test of PANI@WO<sub>3</sub> for OG dye removal.

composite. Besides, there was no significant change in the peaks associated with the elements of PANI@WO<sub>3</sub> composite after interaction with OG dye. This means that the OG dye was adsorbed on the PANI@WO<sub>3</sub> through physical interactions, which is consistent with the thermodynamic study. Moreover, the investigation of the effect of pH on the adsorption process (Section 3.2.1. Effects of adsorbent dose and pH) further confirmed that both electrostatic interactions and  $\pi$ - $\pi$  interactions are involved in the OG dye adsorption mechanism.<sup>34,51</sup> In light of the above discussion, a proposed OG dye adsorption mechanism on the PANI@WO<sub>3</sub> composite is schematized in Fig. 8.

**3.2.5 Regeneration of PANI@WO<sub>3</sub> composite.** The regeneration of PANI@WO<sub>3</sub> using NaOH eluent for recovery of OG dye was investigated up to four times. The results of the repeated use experiments are exhibited in Fig. 9. It was noticed that the uptake ability of PANI@WO<sub>3</sub> composite to remove OG dye from aqueous solutions was still kept at about 87.90% after four cycle uses. Therefore, it was concluded that PANI@WO<sub>3</sub> composite has excellent reusability and subsequently represents an alternative adsorbent for effective clean-up of wastewater containing OG dye.

## 4. Conclusion

A new PANI@WO<sub>3</sub> composite was successfully fabricated and tested to remove OG from an aqueous solution using the adsorption properties of the composite. The PANI@WO<sub>3</sub> composite formation was confirmed by XRD, FT-IR, SEM-EDS, XPS, BET, and TEM analysis. The results confirm the quickness and efficiency of the adsorbent to adsorb OG dye (>95%). The binding mechanism of OG dye on the PANI@WO<sub>3</sub> surface was proposed to be the collaboration of electrostatic interactions and  $\pi$ - $\pi$  interactions. The maximum adsorption capacity of PANI@WO<sub>3</sub> for OG dye reached 226.5 mg g<sup>-1</sup>, which is greater than those of other adsorbents described in the literature. A table of comparative results between our results and those of other authors would be welcome. The adsorption process follows the pseudo-second-order kinetic model, and the Langmuir model is suitable to define adsorption equilibrium. The thermodynamic parameters revealed that the OG adsorption onto PANI@WO<sub>3</sub> is spontaneous and endothermic. The regeneration study demonstrated that the PANI@WO<sub>3</sub> composite has potential reusability for efficient OG removal up to four consecutive cycles. Overall, the experimental findings reported in this study demonstrated that the as-prepared PANI@WO<sub>3</sub> composite is promising adsorbent material for efficient clean-up of OG dye from textile effluents.

## Conflicts of interest

There are no conflicts to declare.

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