







Cite this: *Mater. Adv.*, 2022,
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High CO₂ separation performance on a metal–organic framework composed of nano-cages lined with an ultra-high density of dual-side open metal sites†

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Developing efficient adsorbents for CO₂ separation is the crucial step of CO₂ sequestration, which has found many applications, including in carbon capture from flue gas and natural gas purification. Although many adsorbents have been reported in the past few decades, developing efficient adsorbents with high separation selectivity and excellent capacity is technically challenging due to the trade-off effect between the adsorption capacity and selectivity. This work reports the CO₂ separation performance of a microporous MOF, which comprises nano-cages and an ultra-high density of dual-side open metal sites. The adsorption results reveal a prominent CO₂ uptake and outstanding separation selectivity for CO₂/N₂ and CO₂/CH₄. Furthermore, this MOF also exhibits the preferable adsorption of C₂–C₃ light hydrocarbons over CH₄. Simulation studies reveal that the dual-side open metal sites are the preferable adsorption sites for CO₂, C₂H₄, and C₃H₆. Therefore, the outstanding separation performance can be attributed to the combined effects of nano-cages and the high density of open metal sites.

Received 4th October 2021,
Accepted 10th November 2021

DOI: 10.1039/d1ma00919b

rsc.li/materials-advances

Introduction

With the increasing concern over global warming, the reduction of carbon emissions into the atmosphere has been a focus worldwide.¹ Generally, there are two major routes of reducing carbon dioxide emissions. First, carbon capture and sequestration (CCS) from flue gas is the primary strategy to minimize human carbon emissions in the current stage.^{2,3} Second, replacing the high-carbon energy source (such as petroleum, coal, *etc.*) with a low-carbon energy source (such as CH₄, H₂, *etc.*) is the ultimate solution to carbon reduction.^{4,5} Regarding CO₂ capture, CO₂ separation from gas mixtures

through pressure swing adsorption (PSA) has been regarded as an alternative technique of CO₂ separation to the conventional chemical absorption process, which is energy-intensive and corrosive.^{6–9} The advantages, such as high energy efficiency, low plant investments, and easy mobility, render this technique with very promising potential for CO₂ capture.^{10,11} Besides, the separation of CO₂ from natural gas is central to natural gas production due to its incredible relevance to energy value and transportation safety.^{12,13} As an alternative technique to the conventional cryogenic distillation process that is energy and cost-intensive, the physisorption-based PSA process has been the most promising strategy for CO₂/CH₄ separation due to its low energy consumption and low investments.^{14,15}

The adsorbent is the core material for the PSA process that determines the separation efficiency and energy cost. Metal–organic frameworks (MOFs) represent an emerging family of crystalline adsorbents with a uniform structure, amenable pore environments, and tuneable functionalities among various adsorbents.^{16,17} Besides their encouraging prospects in a wide range of applications, they are also considered as ideal adsorbents for CO₂ separation.^{18,19} In the past two decades, various MOFs with high CO₂ separation selectivities have been reported. However, constructing MOFs with high separation CO₂ selectivity and excellent capacity is still challenging and urgent. This work reports the separation performance of a MOF

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† Electronic supplementary information (ESI) available: Synthetic details, crystal structural parameters, PXRD pattern, calculation procedures and additional kinetic profiles. See DOI: 10.1039/d1ma00919b

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composed of nano-cages lined with a high density of dual-side open metal sites. The gas adsorption study reveals both high adsorption capacity and excellent separation selectivity for CO₂. Theoretical studies and breakthrough experiments also validate the effectiveness of this strategy in promoting the separation efficiency for CO₂ capture from flue gas and natural gas.

Results and discussion

To construct a MOF with multiple functionalities, employing a multiple-functional ligand is the most direct method. 1H-pyrazole-4-carboxylic acid (H₂PCA) is the simplest multiple-functional ligand composed of carboxylate and pyrazole groups. Through the solvothermal reaction of H₂PCA and Cu²⁺ ions, a three-dimensional porous MOF with nano-cages was obtained (see Experimental details in the ESI,† S1). Deep blue block-like single crystals with a homogeneous morphology were obtained in a high yield. Single-crystal X-ray diffraction analysis reveals a three-dimensional network with a formula of [Cu₃(μ₃-OH)(PCA)₃]. The structural retrieval shows that the crystal structure of this MOF has been reported in former literature.²⁰ However, the gas adsorption properties of this MOF have not been reported yet.

This MOF is crystallized in a high symmetrical cubic system, with a space group of *F43c*. The coordination of three pyrazole groups with three Cu²⁺ ions and one bridging oxygen atom gives rise to triangle trinuclear [Cu₃(μ₃-OH)(PCA)₃] secondary building units (SBUs) (see Fig. 1a and b). The [Cu₃(μ₃-OH)(Pz)₃] (Pz represents pyrazole) group in this SBU is a commonly encountered coordination sphere created by pyrazole and transition metal ions. Other than forming the [Cu₃(μ₃-OH)(PCA)₃] SBU, the Cu(II) atom is also coordinated to the carboxylate group on another [Cu₃(μ₃-OH)(PCA)₃] group, giving rise to a planar square coordination environment for Cu(II). Notably, a high density of open metal sites on SBUs can be formed directly without evacuating the coordinated water molecules. Structural analysis reveals that nano-cages with a diameter of 0.97 nm and windows with a diameter of 0.38 nm are involved in this MOF, with a high porosity of 47% calculated from the PLANTON²¹ program (see Fig. 1c–e).

The PXRD pattern of the sample prepared from a scale-up experiment is consistent well with the pattern simulated from the single-crystal data, suggesting that this sample possesses the same structure with [Cu₃(μ₃-OH)(PCA)₃] and has a high degree of crystallinity (as shown in Fig. 1e). Thermal gravimetric analysis (TGA) reveals that this MOF is stable up to 270 °C (see TGA curves in the ESI,† S2). After thermal activation at 130 °C, N₂ adsorption at 77 K was performed on this sample. As shown in Fig. 1f, the typical type-I isotherm indicates the microporous nature of this MOF, which is agreeable with the crystallographic data. The BET surface area derived from the N₂ isotherm is 583 m² g⁻¹, with a pore volume of 0.249 cm³ g⁻¹.

Owing to the square planar coordination geometry for Cu(II) and unique configuration, the dual sides of each Cu(II) center are exposed in the nano-cages, leading to a relatively high

density of accessible open metal sites for this MOF. The density of accessible open metal sites is calculated to be 28 nm⁻³, much higher than the famous MOFs with a high density of open metal sites, such as HKUST-1 (10.5 nm⁻³) and MOF-74 (4.6 nm⁻³). Besides, the nitrogen atoms in the ligand and oxygen atoms in the carboxylate groups could act as additional active sites for gas adsorption. The combination of nano-cages and the high density of open metal sites render this MOF a promising material for CO₂ capture.

The isotherms of CO₂, N₂, and CH₄ at ambient temperatures were measured on the activated sample of [Cu₃(μ₃-OH)(PCA)₃] to investigate its gas adsorption properties. As shown in Fig. 2a, [Cu₃(μ₃-OH)(PCA)₃] exhibits a CO₂ uptake of 2.93 mmol g⁻¹ (65.6 cm³ g⁻¹) at 1 bar and 298 K. With increasing temperature, the CO₂ uptake gradually decreases. In contrast to the high uptake of CO₂, [Cu₃(μ₃-OH)(PCA)₃] exhibits a near-linear N₂ isotherm and displays a negligible uptake of N₂ at 298 K (as shown in Fig. 2b). The significant difference between the isotherms of CO₂ and N₂ indicates the high adsorption selectivities for CO₂/N₂. Similar to N₂, [Cu₃(μ₃-OH)(PCA)₃] also displays a linear isotherm for CH₄ at 298 K, exhibiting a CH₄ uptake of 0.67 mmol g⁻¹ at 298 K and 1 bar.

Based on CO₂ isotherms at different temperatures, the CO₂ adsorption enthalpy (*Q*_{st}) was calculated by using the Clausius–Clapeyron relation (see the calculated procedure in the ESI,† S3). As shown in Fig. 2c, the *Q*_{st} of CO₂ adsorption is in the range of 31.5–36.1 kJ mol⁻¹, with minor differences with increasing adsorption amounts. These *Q*_{st} values are apparently larger than that of common MOFs, such as HKUST-1 (25.26 kJ mol⁻¹)²² and NiDABCO (20 kJ mol⁻¹).²³ The large value of *Q*_{st} suggests the high adsorption potential of [Cu₃(μ₃-OH)(PCA)₃] to CO₂ molecules. The high adsorption potentials to CO₂ molecules could be attributed to the synergistic effect of nano-cages and the high density of open metal sites.

Based on adsorption isotherms, the equilibrium adsorption selectivity for the CO₂/N₂ mixture can be evaluated by the ideal adsorbed solution theory (IAST), of which several studies have widely validated the reliability.^{10,24} As shown in Fig. 2d, the IAST selectivity for the CO₂/N₂ mixture is calculated as 70 at 1 bar and 298 K (see calculation procedures in the ESI,† S4). This value is among the high ranks of adsorbents and is significantly higher than that of the common MOFs (see comparison Table in the ESI,† S5).¹⁸ Induced by the higher adsorption potential for CO₂, [Cu₃(μ₃-OH)(PCA)₃] shows a high CO₂/CH₄ selectivity. The equilibrium CO₂/CH₄ selectivity calculated from IAST is in the range of 11.8–15.9, which is also superior to the bench-marking MOFs, such as HKUST-1 (5.5)^{18,22} and NiDABCO.²³

The outstanding CO₂ separation performance of [Cu₃(μ₃-OH)(PCA)₃] also encourages us to investigate the separation performance for light hydrocarbons, which play crucial roles in petroleum chemistry. As shown in Fig. 2e, [Cu₃(μ₃-OH)(PCA)₃] exhibits type-I isotherms for C₂H₄, C₂H₆, C₃H₆, and C₃H₈, which is in clear contrast with CH₄. At 298 K and 1 bar, adsorption uptakes of 2.2, 1.9, 2.5, and 2.1 mmol g⁻¹ are observed for C₂H₄, C₂H₆, C₃H₆, and C₃H₈, respectively. We can clearly see that the



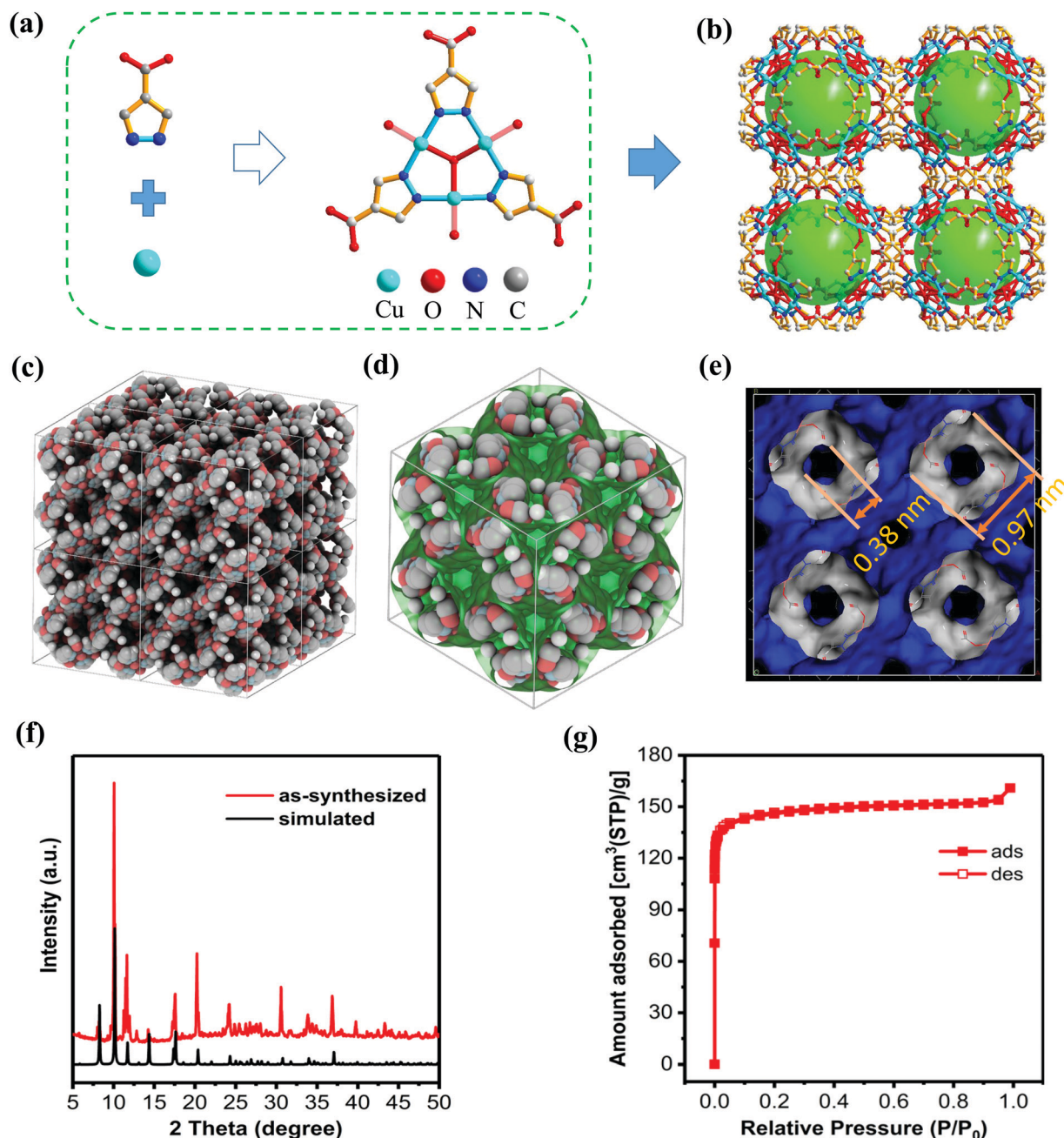


Fig. 1 (a) Illustration of the formation process, ligand and tri-nuclear secondary building units of $[\text{Cu}_3(\mu_3\text{-OH})(\text{PCA})_3]$; (b) the perspective view of the crystallographic structure of $[\text{Cu}_3(\mu_3\text{-OH})(\text{PCA})_3]$ along the c axis exhibiting the nano-cages and channels; (c) the packing mode presentation of $[\text{Cu}_3(\mu_3\text{-OH})(\text{PCA})_3]$ displaying the pore structures; (d) the adsorption surface calculated (green colour) using the He molecules as the probe viewing from the (111) direction; (e) the Connolly surface calculated using a probe radius of 1.0 Å illustrating the accessible nano-cages and windows in $[\text{Cu}_3(\mu_3\text{-OH})(\text{PCA})_3]$; (f) the comparison between the powder X-ray diffraction and simulated XRD patterns; and (g) the N_2 isotherm performed at 77 K.

adsorption uptake increases with the chain length of light hydrocarbons. Furthermore, the adsorption uptakes of C_2H_4 and C_3H_6 are higher than those of C_2H_6 and C_3H_8 , implying that alkenes are more preferentially adsorbed on $[\text{Cu}_3(\mu_3\text{-OH})(\text{PCA})_3]$ than alkanes. Based on the isotherms at different temperatures, the Q_{st} values of adsorption for C_2H_4 , C_2H_6 , C_3H_6 , and C_3H_8 are calculated as 35.4, 33.8, 54.6, and 45.8 kJ mol^{-1} , respectively. The higher Q_{st} values for C3 molecules compared to C2 molecules indicate that $[\text{Cu}_3(\mu_3\text{-OH})$

$(\text{PCA})_3]$ displays a higher interaction strength with the hydrocarbons with longer chain length. Besides, $[\text{Cu}_3(\mu_3\text{-OH})(\text{PCA})_3]$ shows a higher interaction strength with alkenes than alkanes. The preferential adsorption of alkenes to alkanes could be attributed to the π -electrons of alkenes, which could give back donations to open metal sites of MOFs. Due to the higher adsorption potentials, $[\text{Cu}_3(\mu_3\text{-OH})(\text{PCA})_3]$ exhibits higher adsorption selectivity for C2/C3 hydrocarbons than CH_4 . Based on their adsorption isotherms, the equilibrium adsorption selectivities for



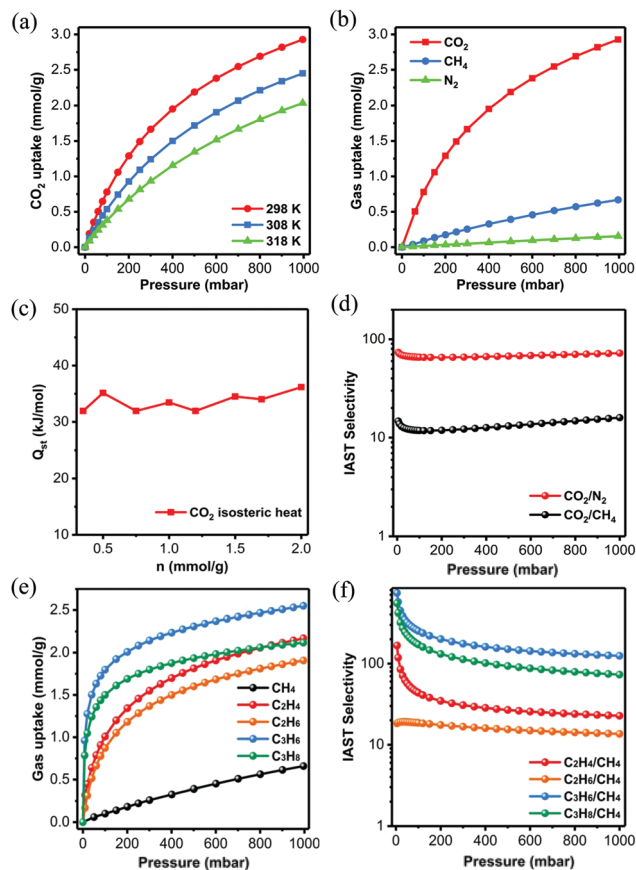


Fig. 2 (a) CO₂ adsorption isotherms on [Cu₃(μ₃-OH)(PCA)₃] at 298, 308 and 318 K and 0–1 bar; (b) the comparison of adsorption isotherms for CO₂, CH₄ and N₂ on [Cu₃(μ₃-OH)(PCA)₃] at 298 K; (c) the isosteric adsorption enthalpy for CO₂ adsorption on [Cu₃(μ₃-OH)(PCA)₃]; (d) the equilibrium separation selectivity for CO₂/N₂ and CO₂/CH₄ mixtures calculated from the IAST method at a feed gas ratio of 50/50; (e) the comparison of adsorption isotherms for CH₄, C₂H₄, C₂H₆, C₃H₆ and C₃H₈ on [Cu₃(μ₃-OH)(PCA)₃] at 298 K; (f) the equilibrium separation selectivity for equimolar C₂H₄/CH₄, C₂H₆/CH₄, C₃H₆/CH₄ and C₃H₈/CH₄ mixtures calculated from the IAST method.

C₂H₄/CH₄, C₂H₆/CH₄, C₃H₆/CH₄, and C₃H₈/CH₄ at 298 K are in the range of 22–167, 13–18, 124–737, and 73–550, respectively (as shown in Fig. 2f). Notably, at 298 K and 1 bar, high IAST selectivities of 124 and 73 have been reached for C₃H₆/CH₄ and C₃H₈/CH₄, outperforming common adsorbents and common MOFs.^{25,26} The excellent adsorption uptake and separation selectivity demonstrates the promising potential of [Cu₃(μ₃-OH)(PCA)₃] in separating light hydrocarbons.

To investigate the separation mechanism, the Grand canonical Monte Carlo (GCMC)²⁷ simulations for CO₂, CH₄, C₂H₄, and C₂H₆ adsorption on [Cu₃(μ₃-OH)(PCA)₃] were performed (as shown in Fig. 3). It is clearly shown that CO₂ molecules are concentrated near the tri-nuclear open metal sites, and the density of CO₂ molecules near open metal sites is much higher than other sites. This phenomenon demonstrates that these open metal sites are the preferred adsorption sites for CO₂. Compared to CO₂, the CH₄ molecules are more dispersed in the pore space, and the density near the open metal sites is much lower. This comparison result

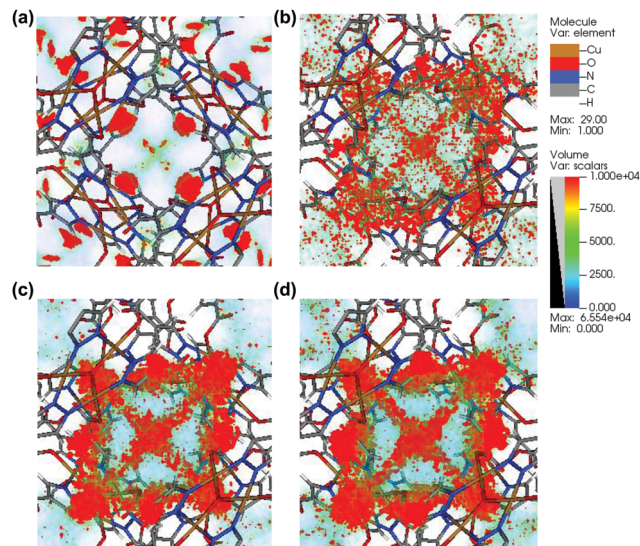


Fig. 3 (a) Density distribution of guest molecules within [Cu₃(μ₃-OH)(PCA)₃] for: (a) CO₂; (b) CH₄; (c) C₂H₄; (d) C₂H₆. (The density distribution was calculated at 298 K and 1 bar using the GCMC method.)

suggests the pronounced function of open metal sites in CO₂ adsorption. Like CO₂, the open metal sites exert high adsorption potentials to C₂H₄ and C₂H₆, which is agreeable with the experimental results. The density distributions of C₂H₄ and C₂H₆ molecules are much higher than CH₄, indicating the preferable adsorption of C₂H₄ and C₂H₆ over CH₄ molecules. The simulation results also confirm the effectiveness of incorporating a high density of open metal sites in enhancing the separation performance of MOFs.

To validate the actual separation performance of [Cu₃(μ₃-OH)(PCA)₃], the breakthrough experiments were performed on the column packed with [Cu₃(μ₃-OH)(PCA)₃] (as shown in Fig. 4). At a feed gas ratio of 50/50 and the pressure of 1 bar, N₂ quickly breaks through from the packed column, while CO₂ requires a rather long time to elute from the column (see Fig. 4a). The significant distinction in the breakthrough time of reaching the outlet of the packed column suggests that [Cu₃(μ₃-OH)(PCA)₃] is capable of separating the mixed CO₂/N₂ (50/50) gas completely. Similarly, CO₂ and CH₄ also exhibit a remarkable difference in times of breaking through from the pack column at a feed gas

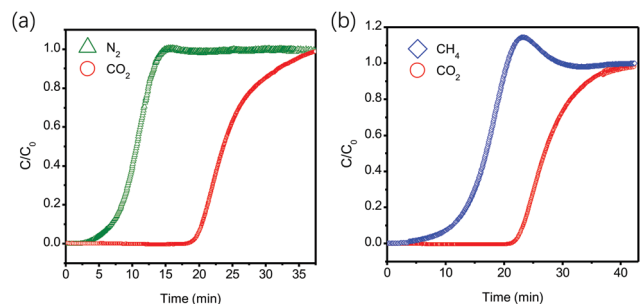


Fig. 4 (a) Experimental column breakthrough curves for (a) CO₂ and N₂ mixtures (50/50, v/v) and (b) CO₂ and CH₄ mixtures (50/50, v/v) at a feed gas rate of 2 mL min⁻¹.



ratio of 50/50, revealing a prominent separation performance for CO₂/CH₄ (see Fig. 4b). These results suggest that [Cu₃(μ₃-OH)(PCA)₃] possesses great potential in capturing CO₂ from flue gas and natural gas.

Conclusions

In summary, we herein revealed the outstanding CO₂ separation performance on a metal–organic framework ([Cu₃(μ₃-OH)(PCA)₃]) comprising nano-cages and a high density of dual-side open metal sites. Induced by the exquisite configuration of the network and tri-nuclear secondary building units, an ultra-high density of accessible open metal sites is generated in the pore channels of [Cu₃(μ₃-OH)(PCA)₃]. As a result, this MOF exhibits significant differences in adsorption uptakes for CO₂/N₂ and CO₂/CH₄, making it a promising candidate for capturing CO₂ from flue gas and natural gas. Furthermore, this MOF also shows outstanding separation performance in separating CH₄ from C₂–C₃ light hydrocarbons. Theoretical investigation reveals that the open metal sites are preferable adsorption sites for CO₂ and C₂–C₃ components. Breakthrough experiments also validate the outstanding separation performance for this MOF. The results of this work will open a new way to develop functional MOFs with high separation performance.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The authors acknowledge financial support from the Major Scientific and Technological Innovation Project of Shandong Province (2020CXGC010402), the National Natural Science Foundation of China (No. 21973254 and 21401215), and the Fundamental Research Funds for the Central Universities (No. 18CX02068A).

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